Chemical Science



EDGE ARTICLE

View Article Online
View Journal | View Issue



Cite this: Chem. Sci., 2020, 11, 5339

dll publication charges for this article have been paid for by the Royal Society of Chemistry

Received 25th August 2019 Accepted 26th February 2020

DOI: 10.1039/c9sc04294f

rsc.li/chemical-science

Observation of binding of carbon dioxide to nitrodecorated metal-organic frameworks†‡

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Metal-organic frameworks (MOFs) functionalised with amine, amide and hydroxyl groups show great promise for CO_2 binding due to their ability to form hydrogen bonds to CO_2 . Herein we report the adsorption and selectivity of CO_2 in four iso-reticular MOFs adopting the NbO topology. Functionalisation of the parent MOF, MFM-102, with $-NO_2$, $-NH_2$ and alkyl groups leads to an enhancement of CO_2 adsorption of up to 36% for the NO_2 -decorated MOF and with raised selectivity. MFM-102- NO_2 shows the highest adsorption capacity for CO_2 (184 cm³ g⁻¹ at 273 K and 1.0 bar) within this series, comparable to the best-behaving iso-reticular MOFs. At 298 K and 1.0 bar, MFM-102- NO_2 shows a CO_2/CH_4 selectivity of 5.0. In situ inelastic neutron scattering and synchrotron FT-IR micro-spectroscopy were employed to elucidate the host-guest interaction dynamics within CO_2 -loaded MFM-102- NO_2 . Neutron powder diffraction enabled the direct observation of the preferred binding domains in MFM-102- NO_2 , and, to the best of our knowledge, we report the first example of CO_2 binding to a $-NO_2$ group in a porous MOF. Synergistic effects between the $-NO_2$ group and the open metal sites lead to optimal binding of CO_2 molecules within MFM-102- NO_2 via hydrogen bonding to C-H groups.

Introduction

Emissions of carbon dioxide (CO₂) are causing significant environmental impacts.¹ However, CO₂ is also an inexpensive, non-toxic and a potentially abundant renewable C₁ source, and, therefore, the development of cost-effective and efficient methods for carbon capture and storage (CCS) is a major target. State-of-the-art CCS is often based upon scrubbing systems that rely heavily upon the use of toxic and corrosive amines, and

isorption of CO_2 within porous materials is a promising approach with the potential of low running-cost and high efficiency. A wide spectrum of materials, such as zeolites, activated carbons, ionic liquids and silica-based materials have been investigated for CO_2 adsorption.^{2–5}

require high operational cost. In contrast, reversible phys-

Metal-organic framework (MOF) materials show potential for gas adsorption and storage owing to their high porosity and designed pore size and structure reflecting their versatility via incorporation of active binding sites by decoration of the pores with functional groups and/or by incorporation of open metal sites. 6,7 MOFs decorated with -NH2,8 -OH,9 -CONH-,10 piperazine,11 and pyridyl groups12 have been explored extensively for CO_2 adsorption. The nitro group $(-NO_2)$ is one of the most powerful electron-withdrawing groups, which can polarise -CH groups on phenyl rings. We argued that potentially this might lead to enhanced interactions between -CH groups and adsorbed CO₂ molecules, and moreover, the -NO₂ group might itself serve as a binding site for CO₂ molecules via dipole/quadrupole interactions. 13,14 However, to date, experimental evidence of CO2 binding to NO2-decorated MOFs has not been reported. Herein, we report the adsorption and selectivity of CO2 in four isoreticular NbO-type MOFs with different organic functional groups, namely -NO2, -NH2 and alkyl groups. The introduction of these functional groups inevitably reduces the pore volume of the parent MOF, but leads to notable enhancement in CO2

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[†] Electronic supplementary information (ESI) available: Synthetic procedures, characterization and additional analysis of crystal structures. Single crystal data of CO₂-loaded MFM-102-NO₂ is deposited at Cambridge Crystallographic Data Centre (CCDC) 1860478. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c9sc04294f

[‡] The crystal data for MFM-102-NO₂, MFM-102-NH₂, and bare MFM-102-NO₂, CCDC 1857304, 1857305 and 1857873, respectively, were previously reported in ref. 16.

adsorption. In particular, the NO₂-functionalised MOF MFM-102-NO₂ shows a 36% increase in CO₂ uptake compared to the parent MFM-102, confirming a positive effect on CO₂ binding. A combination of static and dynamic experiments using *in situ* neutron powder diffraction (NPD), inelastic neutron scattering (INS) and synchrotron-based IR micro-spectroscopy have been applied to study the binding domains of adsorbed CO₂ and the host–guest binding dynamics in MFM-102-NO₂ as a function of CO₂ loading to reveal fundamental insights into the mechanisms for enhanced CO₂ adsorption.¹⁵

Experimental

Material and instrumentation

Chemicals and reagents were purchased from Fisher Scientific, Sigma Aldrich or Fluorochem and used as received without further purification. The detailed synthesis of MFM-102 has been reported previously.16 1H NMR and 13C NMR spectra were measured on a Bruker AV400 or AV500 spectrometer. Highresolution electrospray mass spectra were measured on a Bruker MicroTOF spectrometer with samples dissolved in MeOH, and scanning was conducted in both positive and negative modes. Infrared (IR) spectra were recorded in the 400-4000 cm⁻¹ range in ATR sampling mode with a Thermo Scientific iD5 diamond ATR on a Nicolet iS5 FT-IR spectrometer, and elemental analysis was carried out on a CE-440 elemental analyser (EAI Company). TGA measurements were performed using a PerkinElmer TGA 7 Gravimetric Analyser under a flow of N2 (20 ml min⁻¹) at a heating rate of 5 °C min⁻¹. Powder X-ray diffraction (PXRD) patterns were obtained on a PANalytical X'Pert Pro MPD diffractometer in Bragg-Brentano geometry with Cu-K_{α 1} radiation ($\lambda = 1.5406 \text{ Å}$) at 40 kV and 30 mA over 2θ range 3-40°. Samples were evenly dispersed on zerobackground silicon plates with a cavity depth of 0.3 mm.

Neutron powder diffraction (NPD) of gas-loaded MFM-102-NO $_2$

NPD data were collected at WISH beamline at ISIS Muon and Neutron facility. Acetone exchanged MFM-102-NO $_2$ was loaded into vanadium cans of 11 mm diameter, and the sample was activated by heating at 393 K and at 1×10^{-7} mbar for 2 days. Data for the bare framework were collected after placing the sample can into a liquid He cryostat and cooling to 7 K. CO $_2$ gas was dosed into the system after warming to 290 K. The gases were dosed volumetrically from a calibrated volume, and to ensure gases were fully adsorbed into the sample without condensation elsewhere the system was cooled to 7 K slowly over 3 h. Structure solutions of gas-loaded MFM-102-NO $_2$ were solved by sequential difference Fourier analyses followed by Rietveld refinements (see ESI† Section 4).

Inelastic neutron scattering (INS) of CO_2 -loaded MFM-102-NO₂ and MFM-102-NH₂

INS data were collected on the TOSCA beamline at ISIS Muon and Neutron facility and on the VISION spectrometer at Spallation Neutron Source, Oak Ridge National Laboratory (USA).

MFM-102-NO $_2$ and MFM-102-NH $_2$ were loaded into a vanadium can of 11 mm diameter and outgassed at 10^{-7} mbar at 393 K for 2 days. After placing the sample into a He cooled cryostat, INS data of the bare framework were collected at 7 K. CO $_2$ was dosed volumetrically from a calibrated volume at room temperature and gradually cooled to 7 K to allow the gas to fully adsorb into MFM-102-NO $_2$ and MFM-102-NH $_2$. INS data of CO $_2$ -loaded MFM-102-NO $_2$ and CO $_2$ -loaded MFM-102-NH $_2$ were collected at 7 K. Experimentally obtained INS data were compared with modelled data obtained via density functional theory (DFT) calculations (see ESI † for further details).

Synchrotron FTIR micro-spectroscopy of CO₂-loaded MFM-102-NO₂

Single crystals of MFM-102-NO $_2$ were loaded onto a ZnSe slide and placed into a Linkam FTIR600 variable temperature gastight cell fitted with ZnSe windows. The MOF sample was activated *in situ* under a flow of N $_2$ whilst heating the Linkam stage to 413 K for 6 h. Partial pressures of zeolite dried gases N $_2$ and CO $_2$ were controlled by varying the volumetric flow of the two gases via separate mass flow controllers. FTIR spectra were collected at B22 MIRIAM beamline at Diamond Light Source using a highly bright synchrotron IR source connected to a Bruker Hyperion 3000 IR microscope with a 15× objective and MCT detector (liq. N $_2$ cooled). Spectra (256 scans) were measured at room temperature with a 20 × 20 μ m beam, in the spectral range of 4000–650 cm $^{-1}$ (4 cm $^{-1}$ resolution).

Results and discussion

Synthesis, structure and porosity

MFM-102, MFM-102-NH₂, MFM-102-NO₂ and MFM-111 were synthesised according to our previously reported methods,16-18 and the phase purity of all bulk materials was confirmed by powder X-ray diffraction (see ESI† Section 1). All four MOFs are iso-structural and crystallise in the trigonal space group $R\bar{3}m$. The framework is built from Cu₂(O₂CR)₄ paddlewheel units in which two CuII cations are bridged by four carboxylate groups surrounded the [Cu···Cu] axis and two terminal H2O molecules at axial positions to form a $[Cu_2(O_2CR)_4(OH_2)_2]$ node. These Cu₂(O₂CR)₄ units are bridged by the diisophthalate ligands to afford a 3D extended NbO-type network. Within each framework, there are two different types of metal-ligand cages (denoted as A and B) in a 1: 1 ratio (Fig. 1). Cage A consists of six Cu₂(O₂CR)₄ paddlewheels and six linkers to give a spherical shape of 15 Å diameter. Cage B is constructed of twelve Cu₂(O₂CR)₄ paddlewheels residing at the vertices and six ligands on the faces to form an ellipsoid shaped cavity of 32 Å along the c axis and 18 Å across ab plane. Due to the alternative packing of cages A and B, the alkyl, nitro, and amino groups are directly grafted into the pore walls, providing an excellent platform to study their role in guest binding. The Brunauer-Emmett-Teller (BET) surface area and pore volume of MFM-102, MFM-111, MFM-102-NH₂, MFM-102-NO₂ were determined to be 3412, 2930, 2928, 2893 m² g⁻¹ and 1.29, 1.19, 1.12, 1.14 cm³ g⁻¹, respectively, from N₂ isotherms at 77 K. As

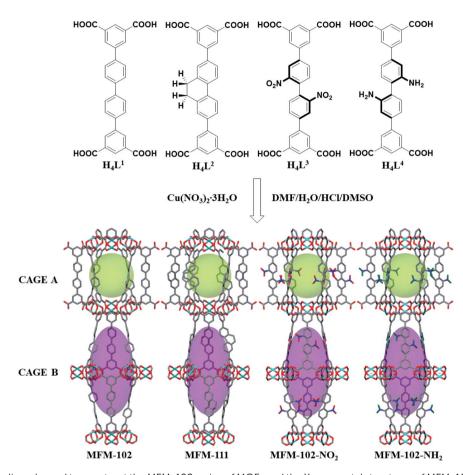


Fig. 1 Views of organic ligands used to construct the MFM-102 series of MOFs and the X-ray crystal structures of MFM-102, MFM-111, MFM-102-NO₂ and MFM-102-NH₂. The functional groups are highlighted in ball-stick mode. (C, dark grey; O, red; N, blue; Cu, turquoise; all coordinated H₂O and hydrogen atoms are omitted for clarity, except for the hydrogen atoms on the -NH₂ group, green).

expected, the porosity of decorated complexes reduces compared to the parent MOF, and so introduction of alkyl, -NH₂ and -NO₂ groups into MFM-102 leads to reduction in BET surface areas of 14%, 14% and 15%, respectively.

Gas adsorption properties

Adsorption isotherms of CO₂ and CH₄ were recorded at 273 and 298 K up to 1 bar. All isotherms are completely reversible, and show fast adsorption-desorption kinetics with equilibrium achieved typically within 3-5 min. Uptake of CO2 at 273 K and 1.0 bar were recorded as 184, 178, 165, 135 cm³ g⁻¹ for MFM-102-NO2, MFM-102-NH2, MFM-111 and MFM-102, respectively (Fig. 2a). Interestingly, despite the reduction in porosity, the introduction of alkyl, $-NH_2$ and $-NO_2$ groups into MFM-102 leads to enhancement on the CO₂ uptake of resultant MOFs by 22%, 32% and 36%, respectively. Notably, the CO₂ uptake of MFM-102-NO2 is comparable to the best-performing isoreticular MOF (NJU-Bai-14, 187 cm³ g⁻¹) under the same conditions.19 Similar enhancement of CO2 adsorption was also observed at 298 K for the decorated MOFs (Fig. 2b), and these uptakes are comparable to other leading NbO-type MOFs (Table 1). In sharp contrast, these MOFs show similarly low uptakes of CH_4 at 273 and 298 K and 1.0 bar (ranging from 30 cm³ g⁻¹ to 20 cm³ g⁻¹). This result indicates the potential of MFM-102-NO₂

as a promising material for selective adsorption of CO₂. The adsorption selectivities for equimolar mixtures of CO2/CH4 within these MOFs have been calculated using ideal adsorbed solution theory (IAST) using the experimental single-component isotherm data (Fig. S4 and S5†). At 298 K and 1 bar, the CO_2/CH_4 (50 : 50) selectivity is 5.0, which is comparable to those reported for other Cu-based MOFs with similar porosity, such as NJU-Bai12 (5.0),²⁰ ZJNU-81(5.5),²¹ ZJNU-82 (5.4),²¹ ZJNU-83 (4.9),21 ZJU-15 (4.4) (ref. 22) and ZJNU-57 (5.5),23 but which do not show NbO topologies.

The isosteric heats of adsorption (Q_{st}) of CO_2 were calculated using virial method based on isotherm data at 273 and 298 K (Fig. 2c). The values for $Q_{\rm st}$ at low surface coverage range from 24.0 to 24.6 kJ mol $^{-1}$, comparable to other MOFs of NbO topology with a high density of open Cu(II) sites (Table 1). With increasing CO_2 loading, the values of Q_{st} steadily reduce in all four cases. Interestingly, the Q_{st} values of MFM-102 drop more rapidly than functionalised materials with increasing loading, indicating the presence of additional binding sites in the decorated MOFs.

In situ neutron powder diffraction (NPD) of CO₂-loaded MFM-102-NO₂

In situ NPD was employed to determine the binding domains for adsorbed CO2 molecules within desolvated MFM-102-NO2 at

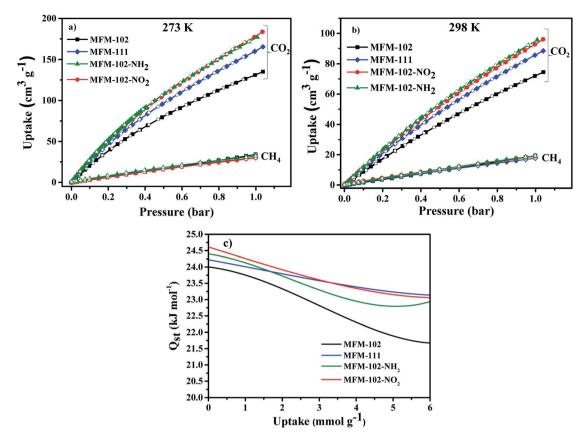


Fig. 2 Adsorption isotherms for desolvated MFM-102, MFM-111, MFM-102-NO₂ and MFM-102-NH₂ (a) 273 K, (b) 298 K. Solid and open symbols represent adsorption and desorption, respectively. (c) Heat of adsorption as a function of CO₂ uptake derived by the virial method.

a loading of 2.0 CO₂/Cu. NPD data for the desolvated complex confirm the complete removal of guest solvents, including coordinated $\rm H_2O$ molecules on the Cu sites, and no significant structural distortion is observed compared to the parent solvated material. A loading of 2.0 CO₂/Cu was used to assess the strongest binding sites within the material without involving notable adsorbate–adsorbate interactions. Fourier difference map analysis of the NPD patterns afforded the location of guest $\rm CO_2$ molecules, which after further development by Rietveld refinement, allowed

unambiguous determination of gas positions, orientations and crystallographic occupancies within each sample.

MFM-102-NO₂ displays six binding sites for CO₂ (denoted as I to VI with decreasing occupancies, Fig. 3 and 4). Significantly, site I (occupancy = 0.50) was found not to be at the open Cu(II) sites but is confined to the window between Cages A and B. At site I, CO_2^{I} forms supramolecular interactions to aromatic –CH group from two adjacent NO₂-substituted phenyl rings $[O_{CO_2}\cdots \underline{H}C^-=2.13(5), 2.42(8) \,\mathring{A}]$. It is noteworthy that –NO₂ is a strong electron-withdrawing groups which leads to

Table 1 Comparison of CO₂ uptakes in selected Cu-MOFs of NbO topology

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MOFs	CO ₂ uptake, 2	73 K (cm 3 g $^{-1}$) CO $_{2}$ uptake, 298	$K \text{ (cm}^3 \text{ g}^{-1}) V_p$, pore volume	(cm ³ g ⁻¹) BET area (m	$Q_{\rm st}$ (kJ mo	ol ⁻¹) Ref.
ZJNU-44	175	116	0.894	2314	_	24
ZJNU-43	180	103	0.894	2243	_	24
ZJNU-45	163	107	0.877	2232	_	24
ZJNU-40	_	86	0.997	2072	24.0	25
ZJNU-41	_	97	0.837	2530	23.5	26
ZJU-8	175	98	1.022	2501	21.9	27
NJU-Bai-14	187	100	0.942	2384	24.5	19
MFM-101	164	86	1.080	2805	_	25
MFM-102	135	74	1.297	3412	24.0	This work
MFM-102-NO ₂	184	96	1.118	2893	24.6	This work
MFM-102-NH ₂	180	97	1.140	2928	24.4	This work
MFM-111	165	88	1.190	2930	24.2	This work
MFM-130	109	59	1.000	2173	26.0	28

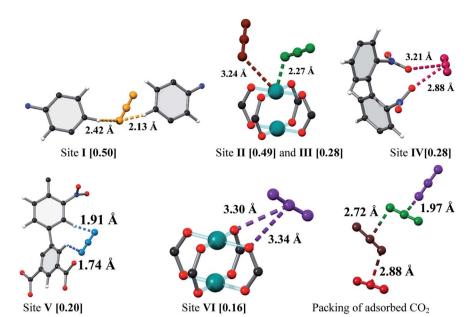


Fig. 3 View of the binding sites for adsorbed CO₂ molecules in MFM-102-NO₂ and the relative positions of these positions in the framework cages [C, dark grey; O, red; N, blue; Cu, turquoise and CO₂ molecules at site I (yellow), site II (green), site III (brown), site IV (ruby), site V (blue) and site VI (purple)].

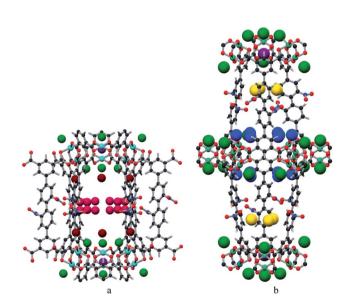


Fig. 4 The positions of the CO_2 in the framework cage A (a) and B (b) of MFM-102-NO₂. CO_2 molecules are at site I (yellow), site II (green), site III (brown), site IV (ruby), site V (blue) and site VI (purple).

polarisation of adjacent –CH moieties. These –CH groups therefore show enhanced acidity and thus act as effective hydrogen bond donors to CO_2 molecules. Site II and III (occupancy = 0.49 and 0.28, respectively) reside at the open $Cu(\pi)$ centre $[O_{CO_2}\cdots Cu=2.27(7)$ Å and 3.24(7) Å for site II and III, respectively]. Site IV (occupancy = 0.28) is stabilised by the dipole/quadrupole interaction between $O(\delta^-)$ of the –NO₂ group from 2 neighbouring phenyl rings and the $C(\delta^+)$ centre of CO_2 with distances of 3.20(6) and 2.87(9) Å. Site V (occupancy = 0.20) was found close to the isophthalate phenyl ring

 $[O_{CO_2}\cdots HC^-=1.72(9) \text{ Å}]$ and the NO_2 -decorated phenyl ring $[O_{CO_2}\cdots HC^-=1.91(4) \text{ Å}]$ stabilised by interactions with aromatic hydrogen atoms. At site VI (occupancy = 0.167), the C atom of CO_2 is close to the two oxygen atoms from carboxylate groups of the $[Cu_2(O_2CR)_4]$ paddlewheel with distances of 3.30(1)–3.34(5) Å. Overall, the effects of $-NO_2$ groups on CO_2 binding are two-fold: (i) polarising neighbouring phenyl C-H bonds via the conjugate effect of the aromatic ring bound to electron-withdrawing $-NO_2$ groups (site I), and (ii) creating additional adsorption sites via direct dipole/quadrupole interaction of $-NO_2$ with CO_2 (site IV).

In situ synchrotron FT-IR micro-spectroscopy of MFM-102-NO $_2$

To investigate further the nature of host-guest interactions, in situ synchrotron FT-IR micro-spectroscopy was conducted on a single crystal of activated MFM-102-NO2. FTIR spectra were collected as a function of CO2-loading by increasing the partial pressure of CO₂ in N₂ from 0 to 1.0 bar (ppCO₂) (Fig. 5a). Gaseous CO2 has three fundamental vibrational modes: the symmetric C=O stretching mode v_1 (1388 cm⁻¹), a doublygenerate bending mode v_2 (667 cm⁻¹) and an asymmetric stretching vibration ν_3 (2349 cm⁻¹) (Table S1†).²⁹ The $\nu_1 + \nu_3$ (3714 cm^{-1}) and $2\nu_2 + \nu_3 (3612 \text{ cm}^{-1})$ combination bands of CO₂ (Fig. 5a) were used to monitor CO₂ sorption as the fundamental anti-symmetric stretch at \sim 2348 cm⁻¹ saturates at ppCO₂ above 0.2 bar. On adsorption, a combination of electrostatic and dispersion forces between CO₂ and the pore surface weakens the C=O bonds in CO₂, leading to the observed redshift of IR bands of adsorbed CO2 to lower frequencies (3700 and 3600 cm⁻¹, respectively) compared to those observed in gas phase (Table S1†). If CO2 strongly coordinates to an open metal

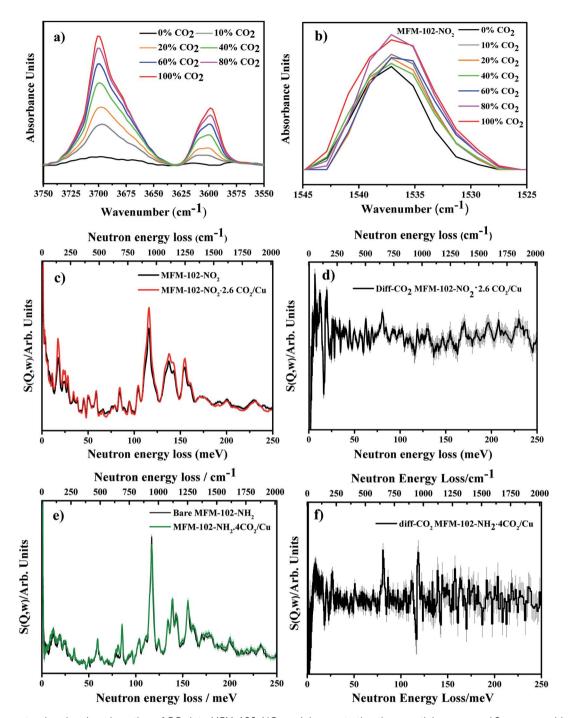


Fig. 5 (a) IR spectra showing the adsorption of CO_2 into MFM-102-NO₂ and demonstrating the growth in $\nu_1 + \nu_3$ and $2\nu_2 + \nu_3$ combination bands for CO_2 near 3700 cm⁻¹ and 3600 cm⁻¹. (b) Changes in intensity and broadening of the N=O asymmetric stretching vibration with CO_2 loading. (c) INS spectra for activated MFM-102-NO2 and MFM-102-NO₂ with CO_2 loading. (d) Difference INS plot, derived by subtracting INS spectra for CO_2 -loaded MFM-102-NO₂ and bare MFM-102-NO₂ spectra. (e) and (f) are the same as (c) and (d), respectively, but replacing MFM-102-NO₂ with MFM-102-NH₂.

site, the molecular symmetry is lowered and the degeneracy of the ν_2 bending mode is removed, splitting the band into two peaks.³⁰ Three IR spectra of the activated, 10% CO₂- and 100% CO₂-loaded MFM-102-NO₂ reveal the changes in the peak at 660 cm⁻¹. This peak shows a significant increase in intensity at 10% CO₂ loading, and broadens and splits into two peaks at

657 cm⁻¹ and 663 cm⁻¹ at 100% CO₂ loading (Fig. S7a†). The IR band of the host framework reveals some interesting features. There is a continuous red shift of the band at 1670 cm⁻¹ (the C=O stretching mode of -COO groups³¹) with increasing CO₂ loading, consistent with the interactions between CO₂ molecules and $[Cu_2(O_2CR)_4]$ paddlewheel units (site VI, Fig. S7b†). An

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apparent change in the peak shape and intensity at 1537 cm⁻¹ (N-O asymmetric stretching mode in the -NO₂ group³²) suggests dipole/quadrupole interaction between the -NO2 groups and adsorbed CO₂ molecules (Fig. 5b). In the low energy regions, the peak at 1025 cm⁻¹ (corresponding to C-H out-of-plane wagging³³) gradually decreases in intensity with increasing CO₂ loading, while the peak at 1005 cm⁻¹ grows and shifts to lower energy (Fig. S7c†). These changes in FTIR spectra are in excellent agreement with the presence of multiple interactions between the framework and CO2 molecules as observed by the NPD experiment.

Inelastic neutron scattering (INS) studies of CO₂-loaded MFM-102-NO2 and MFM-102-NH2

To gain further insight into the host-guest dynamics of CO₂loaded MFM-102-NO₂, in situ INS has been investigated. The calculated INS spectrum of bare MOF was found to be in good agreement with the experimental INS data (Fig. S8†) leading to assignment of all the vibrational modes. In the INS spectrum of CO2-loaded MFM-102-NO2, all the INS peaks shift slightly to higher energy, attributed to a stiffening effect of the host-guest lattice on CO₂ adsorption (Fig. 4c). There are changes in the peak intensity at 164 and 193 meV, corresponding to the symmetrical and asymmetrical stretching of -NO₂ groups. The various distortion motions of the phenyl rings (83 meV), out-ofplane C-H bending on the isophthalate rings (116 meV) and inplane C-H bending on phenyl rings (136 meV) increase in intensity. These results confirm the important role of -NO2 and -CH groups on the adsorption of CO₂. Similarly, the difference INS spectra of CO₂-adsorbed MFM-102-NH₂ also show changes in intensity of the peaks at 80, 117, 119 meV, originating from various in-plane deformational vibrations of the phenyl rings and the wagging motion of aromatic C-H groups. These changes are attributed to the reduction of the motion of aromatic C-H groups with the introduction of CO2 into the framework.

Conclusions

We report a comprehensive investigation of the adsorption and binding of CO₂ in four iso-reticular MOFs with varying functional groups. Introduction of -NO₂ group to the parent MOF, MFM-102, leads a 15% reduction on the BET surface area but a 36% enhancement on the CO2 adsorption capacity at 273 K and 1.0 bar. NPD analysis of CO2-loaded MFM-102-NO2 reveals cooperative binding of CO2 to -CH hydrogen bond donors that is enhanced by the electron withdrawing effects of the -NO₂ groups that decorate pore walls. In addition, direct interaction of CO2 with the open metal sites and the -NO2 groups is observed, thus enhancing further interaction of the MOF with CO2. The in situ spectroscopic studies using INS and FTIR spectroscopy establish that adsorbed CO₂ molecules interact strongly with the -NO2 and -CH groups of MFM-102-NO2, with apparent shifts and changes of intensity of N=O and C-H vibrational bands on CO2 loading. This work provides a fundamental understanding of the effect of the -NO2 group on CO2

adsorption, and the information gained here provides further insights into the development of materials showing improved gas binding via specific interaction with ligand sites within the MOF structure.

Conflicts of Interest

The authors declare no competing interests.

Acknowledgements

We thank EPSRC (EP/I011870), the Royal Society and University of Manchester for funding. This project has received funding from the European Research Council (ERC) under the European Union's Horizon 2020 research and innovation programme (grant agreement no 742401, NANOCHEM). We thank ISIS/ STFC, Diamond Light Source and Oak Ridge National Laboratory for access to Beamlines WISH, B22 and VISION, respectively. Neutron scattering experiments were performed at ORNL's Spallation Neutron Source, supported by the Scientific User Facilities Division, Office of Basic Energy Sciences, US DOE, under Contract No. DE-AC0500OR22725 with UT Battelle, LLC. The computing resources were made available through the VirtuES and the ICE-MAN projects, funded by Laboratory Directed Research and Development program and Compute and Data Environment for Science (CADES) at ORNL.

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