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Pd₃Se₁₀: a semiconducting cluster-based material†

Cullen T. Irvine,^a Chenguang Wang,^a Bowen Yu,^b Tianze Song,^a Alexander Reifsnnyder,^b Yehia Khalifa,^a Wolfgang Windl,^b David W. McComb^b and Joshua E. Goldberger^{b*}

Superatomic clusters have long attracted attention as active components in a broad range of optical and electronic technologies, making the discovery and understanding of new cluster-based materials an important frontier. Here, we establish the optimized synthesis and properties of a metastable palladium selenide Pd₃Se₁₀ phase comprised of Pd₆Se₂₀ cube-like clusters. The solution-phase electrochemically balanced reaction between Pd²⁺ and sodium polyselenide precursor mixtures, under precise reaction conditions, enables the formation of a crystalline product in high yields. The Raman fingerprint of Pd₆Se₂₀ shows that the cluster is also present in amorphous products that form at lower reaction temperatures. Pd₃Se₁₀ is stable up to 340 °C, above which it disproportionates into Se and orthorhombic PdSe₂. Finally, the optical absorption, electronic resistivity, and photoconductivity measurements indicate that crystalline Pd₃Se₁₀ is semiconducting with a 1.18 eV indirect gap. Overall, this understanding opens a pathway toward accessing a new superatomic building block which has potential in optoelectronic devices.

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Introduction

Cluster-based materials, often referred to as superatomic crystals,^{1,2} or cluster-assembled materials,^{3,4} are periodic lattices entirely or partially constructed from atomically precise ensembles of atoms in an analogous way that traditional solid-state crystals are composed of atoms or ions. Some well-established families of these cluster-based materials include Chevrel phases, which have the general formula M_xMo₆Ch₈ (M = metal, Ch = chalcogen),^{5,6} or those derived from fullerenes^{2,7,8} or Zintl ions.^{9,10} The electronic structure and properties of these materials are determined by the bonding within the cluster, between clusters, and between clusters and additional atoms or molecules that may comprise the material.^{4,11,12} With the large degree of structural tunability available by changing the type and number of atoms or ligands in a cluster, there is a huge potential to synthesize new materials with precisely tailored properties.^{2–4,12–15} Already, there exists several examples of cluster-assembled materials with intriguing electronic,^{16,17} ballistic excitonic transport,¹⁸ magnetic,^{16,19} thermal,²⁰ optical,²¹ and superconductive behavior.⁷

A small subset of these cluster-based materials are comprised of completely non-ligated clusters which are held together through only intercluster interactions, typically van der Waals forces.²² Some examples include C₆₀ and C₇₀,⁸ Pd₆Cl₁₂,²³ Sb₄O₆,²⁴ As₄O₆,²⁵ Au₆Te₁₂Se₈,²⁶ P₄Se₃,²⁷ and S₈.²⁸ This type of cluster-based material has decreased spacing between clusters due to the absence of ligands. Thus, they present an excellent opportunity to study systems in which the intercluster electronic and magnetic interactions are maximized.^{1,27} Furthermore, if the appropriate synthetic techniques can be developed, there is also a potential to explore how properties change when you create crystals composed of multiple types of non-ligated clusters or molecules, such as SbI₃·3S₈.^{22,29}

In 2023, we discovered Pd₃Se₁₀, a metastable inorganic crystal inherently composed of non-ligated Pd₆Se₂₀ clusters.³⁰ In this initial report, Pd₃Se₁₀ typically only formed transiently, within the first ten minutes before converting into more thermodynamically stable palladium selenide phases, hindering efforts to prepare large quantities and fully evaluate its properties. Here, we establish a reproducible, high-yield (71–78%) solution-phase synthesis of Pd₃Se₁₀ that prevents its conversion into more thermodynamically stable phases, and we perform a detailed experimental characterization of its physical properties. The reproducibility and yield of the synthesis was improved by developing a chemically and electrochemically balanced reaction scheme that utilizes a precise sodium polyselenide precursor mixture at an optimal temperature and

^a Department of Chemistry and Biochemistry, The Ohio State University, 151 W. Woodruff Ave, Columbus, Ohio 43210, USA. E-mail: goldberger.4@osu.edu

^b Department of Materials Science and Engineering, The Ohio State University, 140 W 19th Ave, Columbus, Ohio 43210, USA

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Materials

Sodium (Na, 99.9%), sodium borohydride (NaBH_4 , 99%), sodium tetrachloropalladate(II) (Na_2PdCl_4 , 98%), palladium (Pd, $\geq 99.9\%$), ethyl alcohol (anhydrous, $\geq 99.5\%$), and methanol ($\geq 99.8\%$) were purchased from Sigma-Aldrich. Selenium powder (200 mesh, 99.999%), platinum powder (200 mesh, 99.98%) and triethylene glycol (TREG, 99%) were purchased from Alfa Aesar. Potassium bromide (KBr) was purchased from Mallinckrodt. Gold-epoxy paste (Type GE-40) was purchased from Transene Company, Inc.

Synthesis of a 2 : 1 Na₂Se₃ and Na₂Se₄ molar mixture (Na₂Se_{3.33})

Caution! Avoid contacting sodium polyselenides with acids. $\text{Na}_2\text{Se}_{3.33}$ powder was synthesized in a similar way to our previously described method for making Na_2Se_2 powder,³⁰ with the only main difference being that the ratio of reactants was changed to form a sodium polyselenide mixture with a different Na to Se ratio. 0.356 g (0.0155 mol) of sodium, 0.084 g (0.0022 mol) of NaBH_4 , 2.094 g (0.0265 mol) of selenium, and 35 mL of anhydrous ethanol were each put into a separate flask and sealed with septa inside an argon glovebox. The flasks were removed from the glovebox and connected to a Schlenk line to perform the reaction under air-free conditions. 10 mL of ethanol was added to the three-neck flask containing Se, which was also equipped with a water-cooled condenser. 25 mL of ethanol was added to the flask with sodium to form a sodium ethoxide/ethanol solution. **Caution!** Ensure your flask can release pressure as sodium reacts with ethanol to evolve hydrogen gas. Once all the sodium had reacted, the sodium ethoxide solution was transferred to the flask with NaBH_4 and stirred until all the NaBH_4 dissolved. Then, the NaBH_4 /sodium ethoxide solution was transferred to the three-neck flask with a stirring Se and ethanol suspension to begin the reaction. The reaction solution was stirred for 30 min at room temperature and then heated to reflux. After an additional 45 min of reaction time at reflux, the solvent was removed and the flask was taken back into the glovebox. The black sodium polyselenide powder was removed from the flask and ground with a mortar and pestle. Then, the sodium polyselenide powder was

put it a new round bottom flask and sealed with a septum. The flask was removed from the glovebox and connected to the Schlenk line. The powder was annealed in a 200 °C sand bath under flowing argon for 15 hours. After the anneal, the flask was brought back into the glovebox and the powder was ground with a mortar and pestle before being stored in a vial wrapped in aluminum foil.

Synthesis of Pd₃Se₁₀

A general procedure for the synthesis of $\text{Pd}_3\text{Se}_{10}$ is described below. However, the reaction temperature, reaction time, and Na_2PdCl_4 average injection rate were varied between reactions with the details included in Table S1 (ESI[†]).

Na₂Se_{3.33} (0.209 g) and Na₂PdCl₄ (0.199 g) were weighed out in an argon glovebox. The Na₂Se_{3.33} powder and 45 mL of TREG were put into a three-neck flask equipped with a condenser and attached to a Schlenk line so that the reaction could be conducted in an argon atmosphere. The Na₂PdCl₄ powder and 15 mL of TREG were mixed in a glass vial. Na₂PdCl₄ was dissolved by shaking manually and heating the vial to 60 °C on a hot plate. The Na₂Se_{3.33}/TREG solution was stirred and heated to the desired reaction temperature and then the Na₂PdCl₄ solution was added to the three-neck flask either all at once or at the desired average injection rate. For the slow injection rate experiments, the Na₂PdCl₄ solution was not added continuously, but rather manually in 0.25 mL or 0.5 mL portions with a set amount of time in between injections to achieve an average rate of precursor addition. **Caution!** The production of gas is sometimes observed during this reaction. Ensure your setup allows for pressure to be released. After an initial temperature drop, the reaction solution was heated back to the desired temperature and stirred for the rest of the desired reaction time. At the end of the reaction time, the flask was removed from the heating source and allowed to cool to room temperature. For the reactions that were less than 10 minutes, the flask was submerged in an ice bath to end the reaction. When cool, the reaction solution was poured into centrifuge tubes. DI water was also used to rinse the remaining solids from the flask into the centrifuge tubes. The tubes were centrifuged for 10 min at 1560g and then the TREG supernatant was discarded, leaving black solids in the tubes. For the purification process, the solids in each tube were mixed with 10 mL of a solvent (methanol or DI water). Then, the tube was centrifuged for 5 min at 1560g. The supernatant was discarded after centrifugation and the washing process was repeated multiple times with both DI water and methanol. The solids were air-dried overnight.

Characterization

Powder X-ray diffraction (PXRD) data was obtained using a Bruker D8 Advance Diffractometer with Cu K α_1 radiation. The Na₂Se_{3.33} sample was prepared in an argon glovebox and covered with Kapton tape to allow for air-free data collection. Rietveld analysis was performed using GSAS-II.³¹

For transmission electron microscopy (TEM), dilute suspensions of Pd₃Se₁₀ in methanol were prepared by bath sonication

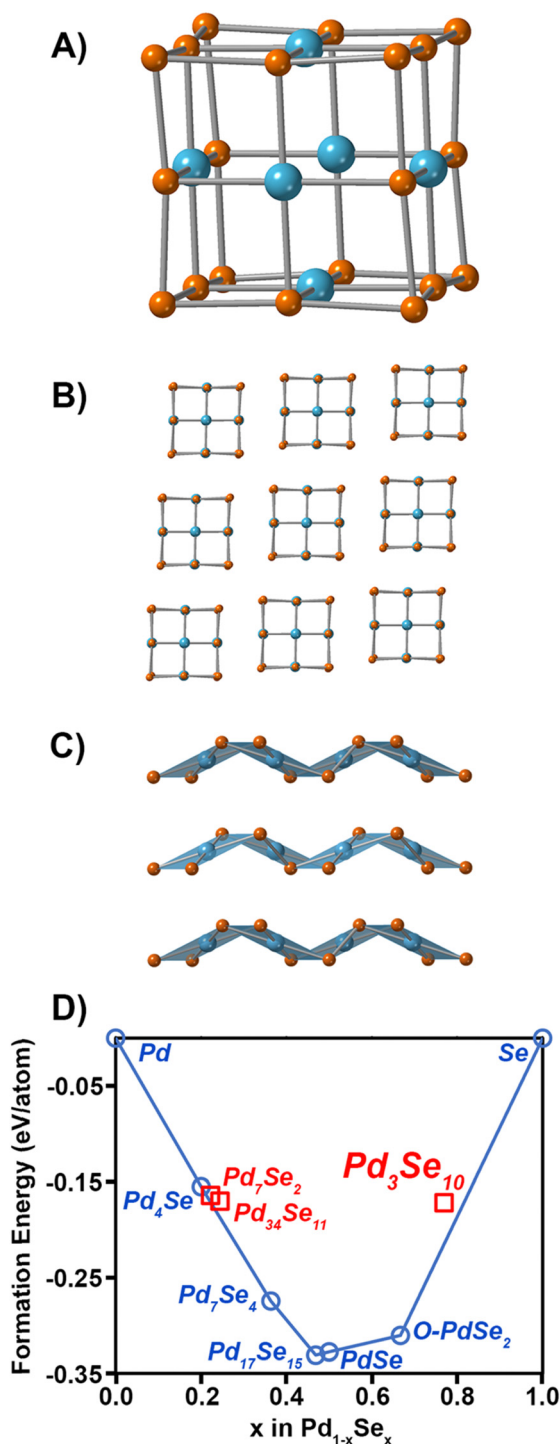


Fig. 1 (A) Structure of a Pd₆Se₂₀ cluster. Pd atoms are in blue and Se atoms are in orange. (B) Crystal structure of Pd₃Se₁₀ looking down the [111] zone axis. (C) Crystal structure of O-PdSe₂ looking down the *b* axis. (D) Pd–Se convex hull phase diagram. Red squares correspond to metastable phases, blue circles correspond to thermodynamically stable phases, and the blue line is the convex hull.

DFT calculations of the convex hull predict that Pd₃Se₁₀ is a metastable phase with an energy above the hull of 0.043 eV atom^{−1} (Fig. 1D). The thermodynamically stable

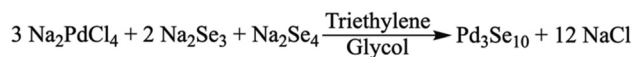
phases that sit on the hull are O-PdSe₂ and Se. A fundamentally important question is why this metastable phase forms in solution instead of O-PdSe₂. It is well established that nanoparticles of metastable phases will nucleate preferentially if their surface energies are lower than their stable counterparts.^{41–44} Thus, we calculated and compared the surface energies of O-PdSe₂ and Pd₃Se₁₀. In O-PdSe₂, the (001) facet will correspond to the lower energy basal plane of this layered material, whereas the (100) and (010) edges will be higher energy facets with dangling bonds. In O-PdSe₂, the energy of the (100) and (010) edge facets are 45.4 and 45.3 meV Å^{−2}, respectively, while the energy of the (001) facet is 27.8 meV Å^{−2}. In contrast, Pd₃Se₁₀ has six equivalent low energy surface facets: ($\bar{1}11$), ($1\bar{1}\bar{1}$), (101), ($\bar{1}0\bar{1}$), (0 $\bar{1}1$) and (01 $\bar{1}$) using hexagonal axes. The surface energies of these facets were calculated to be 15.4 meV Å^{−2}. The fact that the surface energies of Pd₃Se₁₀ are lower than those of O-PdSe₂ explains why Pd₃Se₁₀ particles can form at nanoscale dimensions.

Improving the reliability of the Pd₃Se₁₀ synthesis

Pd₃Se₁₀ was originally synthesized by reacting Na₂Se₂ and Na₂PdCl₄ in triethylene glycol.³⁰ Although this method was successful, the synthesis was highly irreproducible, with variable yields ranging from 1% to 65%. We hypothesized that the root of the issue was that the originally proposed reaction is not electrochemically balanced. If we assume Pd₃Se₁₀ is comprised of Pd²⁺ cations, each Se atom should have an overall oxidation state of −0.6 to form a charge balanced phase. Thus, the formation of Pd₃Se₁₀ from the reaction between Na₂Se₂ (which has a Se^{−1} oxidation state) and Na₂PdCl₄ would require an additional sacrificial oxidant. Instead, a new chemically and electrochemically balanced equation for the synthesis of Pd₃Se₁₀ was proposed (Scheme 1). In this reaction, Na₂Se₂ was replaced with a mixture of Na₂Se₃ and Na₂Se₄ at a 2 : 1 molar ratio, in which Se would have an average oxidation state of −0.6. For simplicity, this combination of Na₂Se₃ and Na₂Se₄ will be referred to as “Na₂Se_{3.33}”, which reflects the atomic ratio of Na to Se in the new precursor, as well as the fact that it is composed of polyselenide dianions.

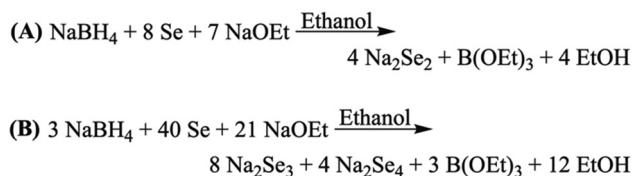
Na₂Se₂ can be made using a modified version of Klayman's method, in which Se is reduced by NaBH₄ in an ethanol solution containing sodium ethoxide (Scheme 2A).⁴⁵ To synthesize Na₂Se_{3.33}, the ratio of NaBH₄ to Se was altered so that less Se atoms would be reduced, making the average oxidation state of Se in the polyselenide dianion product less negative (Scheme 2B). The actual product was 63.9(7) mol% Na₂Se₃ and 36.1(5) mol% Na₂Se₄, as determined by Rietveld analysis (Fig. S1 and Table S2, ESI†).

In the process of synthesizing this new sodium polyselenide precursor, we also solved the crystal structure of Na₂Se₃ from



Scheme 1 Solution-phase method for synthesizing Pd₃Se₁₀ utilizing Na₂Se₃ and Na₂Se₄.





Scheme 2 Synthetic scheme for (A) sodium diselenide⁴⁵ and a (B) 2 : 1 molar mixture of sodium triselenide and sodium tetraselenide. Et = Ethyl.

the powder X-ray diffraction (PXRD) data of the product containing Na_2Se_3 and Na_2Se_4 . Although two separate papers have reported synthesizing Na_2Se_3 in liquid ammonia and also provided supporting UV-visible spectroscopy and ^{77}Se NMR evidence that what they synthesized contained triselenide dianions,^{46,47} the crystal structure of Na_2Se_3 has not been reported until now. To solve the crystal structure of Na_2Se_3 , its lattice parameters and possible space groups were first determined using the auto-indexing function in TOPAS for the peaks in the pattern that could not be attributed to Na_2Se_4 . Next, the likely number of formula units per unit cell (Z) for Na_2Se_3 was inferred from its cell volume ($502 \text{ \AA}^3$) and the assumption that its density would be similar to the two reported sodium polyselenides: Na_2Se_2 (3.38 g cm^{-3})⁴⁸ and Na_2Se_4 (3.79 g cm^{-3}).⁴⁹ Assuming a possible density range of $3\text{--}4 \text{ g cm}^{-3}$, the only possible integer Z value for Na_2Se_3 would be 4. With all of this information, a likely crystal structure was found using the Monte Carlo parallel tempering algorithm in FOX.⁵⁰ Lastly, a Rietveld refinement was performed using GSAS-II to determine the crystal structure of Na_2Se_3 as well as the weight and mole fractions of Na_2Se_3 and Na_2Se_4 in the sodium polyselenide precursor (Fig. S1 and Table S2, ESI†).³¹

Na_2Se_3 crystallizes in the monoclinic $C2/c$ space group and, as expected, it consists of sodium cations and triselenide dianions (Fig. S2 and Tables S3, S4, ESI†). It has a reduced crystal symmetry in comparison to the other alkali metal triselenides K_2Se_3 , Rb_2Se_3 , and Cs_2Se_3 , which all share the same orthorhombic K_2S_3 structure type.^{51,52} Despite its differences in crystal symmetry, the Se–Se bond length ($2.423 \text{ \AA}$) and the Se–Se–Se bond angle (102.8°) of the triselenide dianion in Na_2Se_3 are in close agreement with what is reported for the other three alkali metal triselenides (Table S5, ESI†).

Optimizing the $\text{Pd}_3\text{Se}_{10}$ synthesis

With the $\text{Na}_2\text{Se}_{3.33}$ precursor ready for use, we then sought to optimize the solution phase synthesis of $\text{Pd}_3\text{Se}_{10}$ by tuning three parameters: reaction time, temperature, and Na_2PdCl_4 injection rate. The PXRD data for the results of the $\text{Pd}_3\text{Se}_{10}$ syntheses are shown in Fig. 2 and a summary of the different reaction conditions and results is provided in Table S1 (ESI†). For all reactions, a 60°C solution of Na_2PdCl_4 was added to a hot solution of $\text{Na}_2\text{Se}_{3.33}$, which was connected to a Schlenk line to maintain air-free conditions. The Na_2PdCl_4 precursor was added either all at once (fast addition) or at a target average injection rate.

First, the effect of reaction time was investigated (Fig. 2A) for reactions conducted at the same reaction temperature (245°C)

and injection rate (fast addition). $\text{Pd}_3\text{Se}_{10}$ crystallizes within 30 s, but O- PdSe_2 and Se start to form soon after, with minor peaks corresponding to O- PdSe_2 first appearing in the pattern for the 3 min reaction (Fig. S3, ESI†). After 1 h, O- PdSe_2 had become a major component in the product, in addition to $\text{Pd}_3\text{Se}_{10}$. This trend is similar to what was observed for our original reaction using the Na_2Se_2 precursor, in which longer reactions at 245°C led to the formation of the thermodynamically stable O- PdSe_2 phase and the disappearance of the metastable $\text{Pd}_3\text{Se}_{10}$ phase.³⁰

Second, the effect of reaction temperature was investigated (Fig. 2B) for 1 h reactions with a fast addition of the Na_2PdCl_4 precursor. Since lower activation energies are needed to synthesize metastable phases,⁵⁵ our goal was to find a temperature high enough to form crystalline $\text{Pd}_3\text{Se}_{10}$, but low enough to avoid the formation of O- PdSe_2 . Only amorphous products formed for the 100°C and 150°C reactions, indicating that these temperatures were too low to provide the necessary energy to form any crystalline palladium selenide phase. However, $\text{Pd}_3\text{Se}_{10}$ did form for the 200°C reaction, in addition to selenium and an unknown minor impurity. The impurity peaks are highlighted in Fig. S4 (ESI†). Although the structure of this impurity couldn't be identified, it is likely a metastable product since it was able to form at 200°C , but not at 245°C . Additionally, the lack of peaks in the PXRD pattern corresponding to O- PdSe_2 indicate that 200°C is a low enough temperature to avoid the formation of thermodynamic phases.

For a 10 min reaction at 200°C , with a fast addition of the Na_2PdCl_4 solution, $\text{Pd}_3\text{Se}_{10}$ formed without the presence of any impurities (Fig. 2C). Although this result validated our hypothesis that using a lower temperature would allow for the synthesis of phase pure $\text{Pd}_3\text{Se}_{10}$, the product formed had broad XRD peaks indicative of small crystalline domain sizes. Since using a slower injection rate typically leads to the synthesis of larger nanoparticles,^{56–60} we hypothesized that altering this parameter could also be used to create larger crystalline domains of $\text{Pd}_3\text{Se}_{10}$.

The effect of the Na_2PdCl_4 injection rate was investigated for syntheses performed at 200°C , in an effort to improve the crystallinity of the $\text{Pd}_3\text{Se}_{10}$ product (Fig. 2C). Qualitatively, a slower injection rate produced larger crystalline domain sizes, as evidenced by the sharper peaks present in the PXRD pattern for the 10 min reaction with an average injection rate of 1.5 mL min^{-1} . Slowing the injection rate even further to 0.25 mL min^{-1} for a 1 h reaction produced slightly sharper PXRD peaks and also prevented the formation of any of the aforementioned impurities that were observed with a fast precursor addition (Fig. 2B and C). To quantify how much the crystallinity of the $\text{Pd}_3\text{Se}_{10}$ product improved by using slower injection rates, the full width at half maximum (FWHM) values of the (303) peaks were compared (Fig. 2D). Using these values, the crystalline domain sizes of the samples were calculated using the Scherrer equation (Fig. 2D inset).⁵⁴ The FWHM values for both samples made using the slower injection rates was about twice as small as the value for the sample made with a fast precursor addition. This corresponds to average crystalline domain sizes



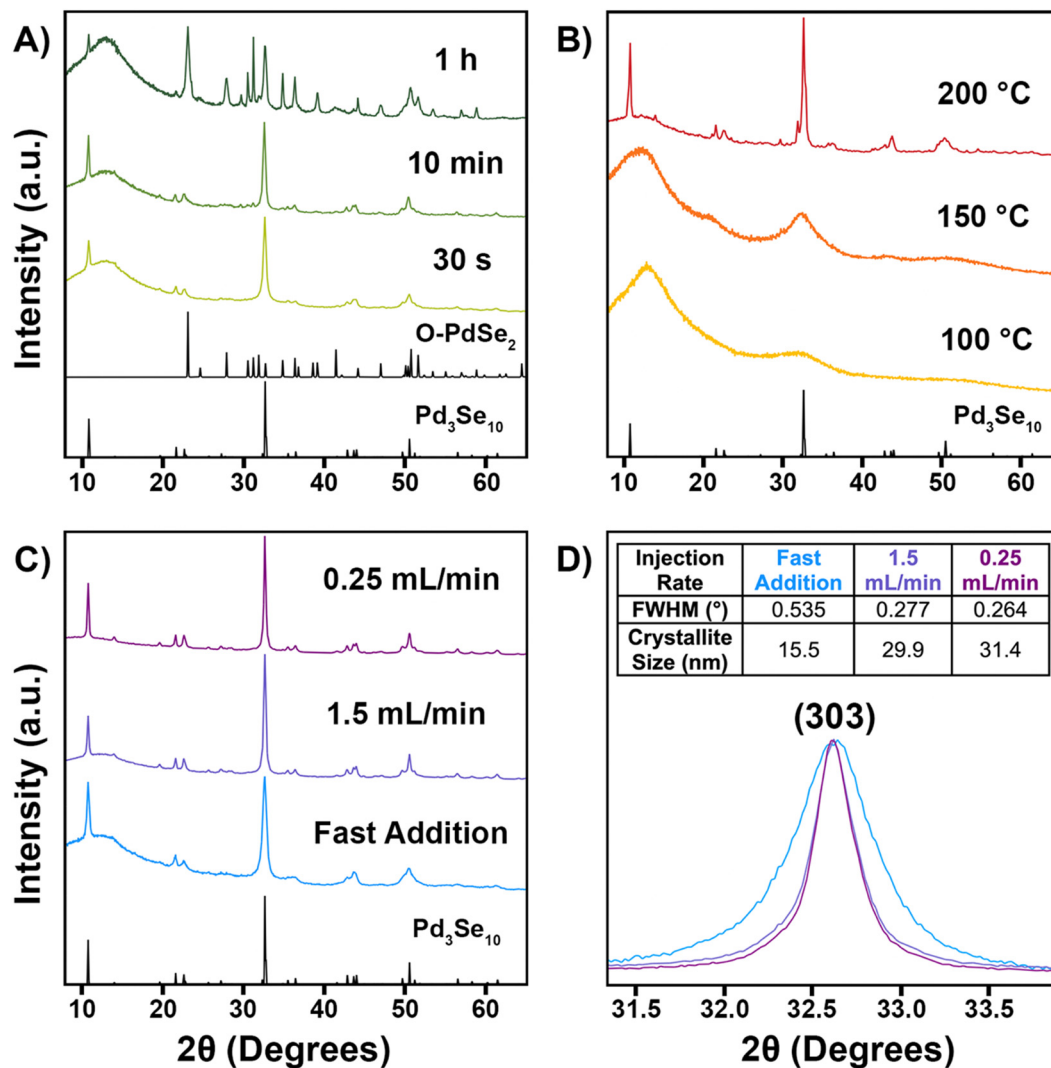


Fig. 2 PXRD patterns for various $\text{Pd}_3\text{Se}_{10}$ synthesis attempts. $\text{Pd}_3\text{Se}_{10}$ and O-PdSe_2 reference patterns are shown in black.^{30,53} (A) Results for syntheses of different reaction times, but the same temperature (245 °C) and Na_2PdCl_4 injection rate (fast addition). (B) Results for syntheses at different temperatures, but the same reaction time (1 h) and Na_2PdCl_4 injection rate (fast addition). (C) Results for syntheses at different Na_2PdCl_4 injection rates, but the same temperature (200 °C). The fast addition and 1.5 mL min^{-1} experiments were 10 min long and the 0.25 mL min^{-1} experiment was 1 h long. (D) Zoomed-in view of the (303) peak for the patterns in (C). The table inset lists the full width at half maximum (FWHM) values for the (303) peaks and the corresponding crystalline domain sizes that were calculated using the Scherrer equation.⁵⁴ The broad peak around 15° in the PXRD patterns is due to the amorphous plastic sample holder used during data collection.

of approximately 30 nm for both slow injection samples and 15 nm for the fast addition sample.

Electron microscopy was used to further characterize the structure of the $\text{Pd}_3\text{Se}_{10}$ product, as well as its particle size and morphology (Fig. 3). The atomic resolution HAADF-STEM image of a $\text{Pd}_3\text{Se}_{10}$ particle was taken to show how the $\text{Pd}_6\text{Se}_{20}$ clusters pack together and the overlay highlights the match between the atomic number (Z) contrast observed in the image and the expected crystal structure (Fig. 3A). Fig. 3B shows a typical TEM image of an aggregate of $\text{Pd}_3\text{Se}_{10}$ particles. Pure $\text{Pd}_3\text{Se}_{10}$ samples had similar particle sizes regardless of their crystalline domain sizes that were calculated using the Scherrer equation. The particles had edge lengths ranging from 10–70 nm, but over 60% of the particles for both samples that were

measured had edge lengths in the 20–40 nm range (Fig. S5, ESI†). Fig. 3C shows the corresponding selected area electron diffraction (SAED) ring pattern for the aggregate of $\text{Pd}_3\text{Se}_{10}$ particles and the indexing is shown in Fig. S6 and Table S6 (ESI†). The measured d -spacings match well with the actual $\text{Pd}_3\text{Se}_{10}$ d -spacings and the most intense ring in the SAED pattern and the most intense peak in the PXRD data correspond to the same (303) reflection. Lastly, energy-dispersive X-ray spectroscopy (EDX) for $\text{Pd}_3\text{Se}_{10}$ confirmed the expected atomic ratio of Pd to Se (Fig. S7 and Table S7, ESI†).

X-ray photoelectron spectroscopy (XPS) data of $\text{Pd}_3\text{Se}_{10}$ compared to Pd, Se, and O-PdSe_2 , supports our expectations of the Pd and Se oxidation states in $\text{Pd}_3\text{Se}_{10}$ (Fig. 4). The Pd $3d_{5/2}$ binding energy of $\text{Pd}_3\text{Se}_{10}$ (336.5 eV) is similar to that



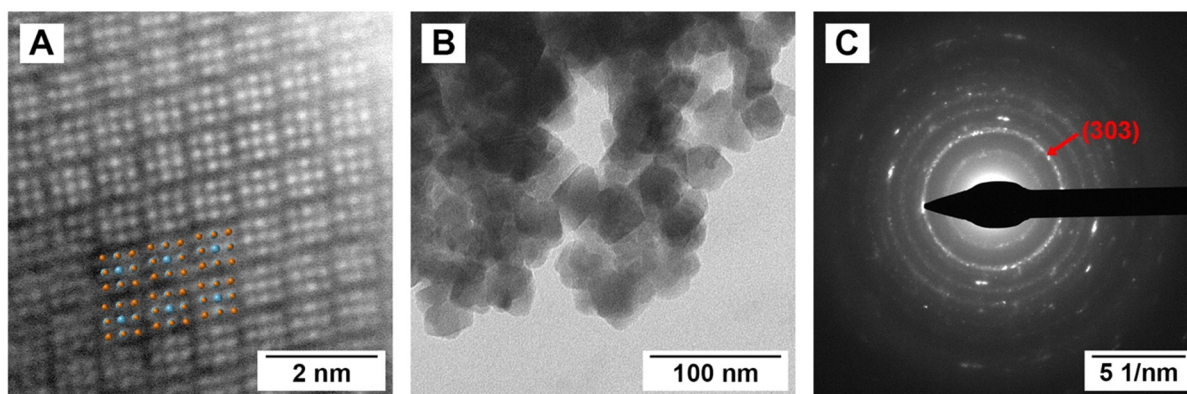


Fig. 3 (A) HAADF-STEM image of a $\text{Pd}_3\text{Se}_{10}$ particle looking down the $[\bar{1}11]$ zone axis, with the crystal structure overlayed on a section of the image. The orange spheres correspond to Se atoms, the blue spheres correspond to Pd atoms. (B) TEM image of an aggregate of $\text{Pd}_3\text{Se}_{10}$ particles and (C) the corresponding SAED pattern. The most intense ring, corresponding to the (303) reflection, is indicated with the red arrow.

of O-PdSe₂ (336.7 eV) and 1.2 eV greater than Pd (335.3 eV) (Fig. 4A). The fact that these binding energies are reasonably close in $\text{Pd}_3\text{Se}_{10}$ and O-PdSe₂ supports the assignment of a Pd^{2+} oxidation state in $\text{Pd}_3\text{Se}_{10}$. By consequence, the average oxidation state of a Se atom in the material should be -0.6 , to maintain charge neutrality. Depending on the nature of charge distribution throughout the $\text{Pd}_6\text{Se}_{20}$ clusters, this oxidation

state could occur with the negative charge delocalized across every Se, or alternatively, with the 12 edge Se atoms having a formal Se^{-1} charge and the 8 corner atoms being formally Se^0 . The Se $3d_{5/2}$ binding energy of $\text{Pd}_3\text{Se}_{10}$ (54.6 eV) is closer to that of O-PdSe₂ (54.7 eV) than Se (55.2 eV) (Fig. 4B). This indicates that the average Se atom in $\text{Pd}_3\text{Se}_{10}$ has a negative oxidation state, but it does not allow for a precise oxidation state assignment or provide information on the nature of the delocalization, especially considering the small differences in binding energies between neutral and reduced forms of Se.^{61–63} Regardless, the combination of the Pd and Se 3d XPS spectra and the square planar geometry of the Pd atoms in the material support this description of the Se oxidation state in $\text{Pd}_3\text{Se}_{10}$.

Raman spectroscopy was also used to characterize the $\text{Pd}_3\text{Se}_{10}$ products to better understand the chemical nature of the amorphous samples grown at 100 °C and 150 °C (Fig. 5). An isolated cluster of $\text{Pd}_6\text{Se}_{20}$, assuming it has the same molecular symmetry as it has in the crystal structure, would have a D_{3d} point group and 48 vibrational modes (8 A_{1g} , 4 A_{2g} , 12 E_g , 4 A_{1u} , 8 A_{2u} and 12 E_u), of which only the 8 A_{1g} and 12 E_g modes are Raman active. $\text{Pd}_3\text{Se}_{10}$ crystallizes into the $R\bar{3}m$ space group with 3 atoms occupying different 18 h Wyckoff sites ($x, -x, z$), 1 atom occupying the 18g Wyckoff site ($x, 0, \frac{1}{2}$) and 1 atom occupying the 6c site ($0, 0, z$). Considering only the Raman active modes, each atom on the 18h site would have 2 A_{1g} and 3 E_g modes, each atom on the 18g site would have 1 A_{1g} and 3 E_g modes, and each atom on the 6c site would have 1 A_{1g} and 1 E_g mode. In total, crystalline $\text{Pd}_3\text{Se}_{10}$ should have 21 Raman active modes (8 A_{1g} and 13 E_g). 20 of these modes correspond to the internal modes of the molecular cluster, whereas the additional E_g mode would correspond to an external mode, or a collective oscillation that involves multiple clusters throughout the crystal. In the crystalline $\text{Pd}_3\text{Se}_{10}$ spectrum, at least 8 intense Raman modes are observed, which could be either individual modes or combinations of closely spaced Raman modes. Because of the weak van der Waals bonding between $\text{Pd}_6\text{Se}_{20}$ clusters in the crystal structure, one would expect the frequencies of the internal modes to be only slightly different than an

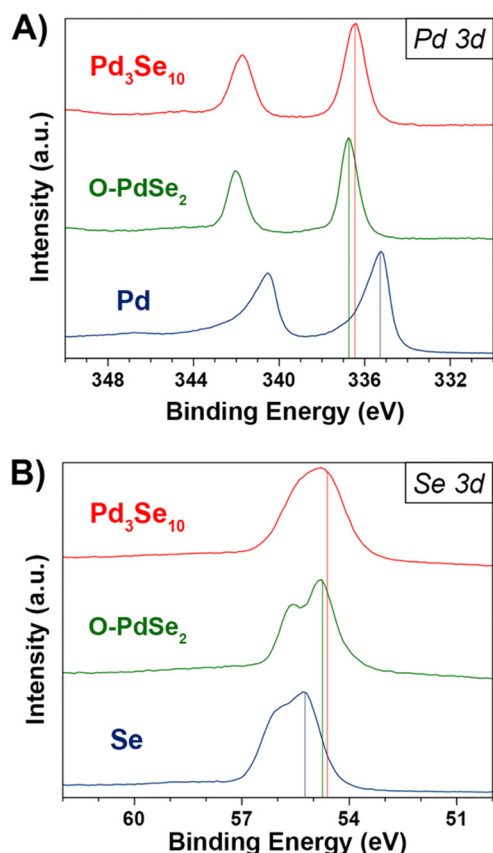


Fig. 4 (A) Pd 3d XPS spectra for $\text{Pd}_3\text{Se}_{10}$, O-PdSe₂, and palladium powders. (B) Se 3d XPS spectra for $\text{Pd}_3\text{Se}_{10}$, O-PdSe₂, and selenium powders.

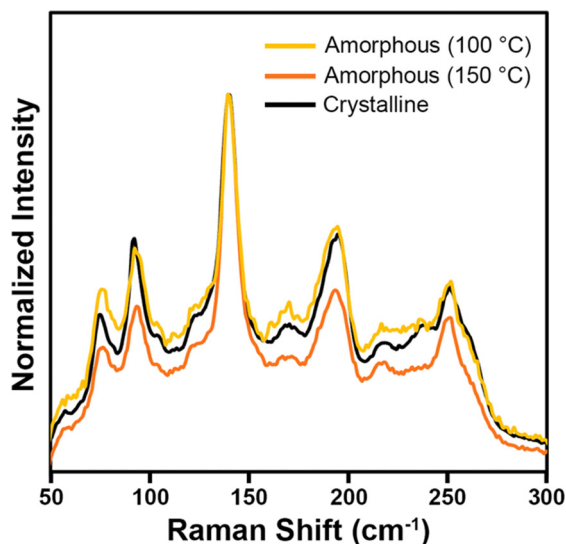


Fig. 5 Raman spectra of a crystalline $\text{Pd}_3\text{Se}_{10}$ sample and the two amorphous $\text{Pd}_3\text{Se}_{10}$ samples synthesized at 100 °C and 150 °C.

isolated cluster. The Raman spectra of the two amorphous samples (Fig. 5) both had very similar peak positions and relative intensities to the crystalline $\text{Pd}_3\text{Se}_{10}$ sample. These Raman spectra are clearly distinct from the spectra of selenium, O-PdSe_2 , and the monoclinic phase of PdSe_2 ⁶⁴ (Fig. S8, ESI†). The fact that the amorphous and crystalline spectra are nearly identical indicate that the amorphous samples are indeed composed of the same $\text{Pd}_6\text{Se}_{20}$ clusters as crystalline $\text{Pd}_3\text{Se}_{10}$ but arranged in a disordered manner.

This deeper understanding of the nature of the amorphous $\text{Pd}_3\text{Se}_{10}$ samples allows us to infer mechanistic information on the formation of crystalline $\text{Pd}_3\text{Se}_{10}$ in solution. Since an amorphous material composed of $\text{Pd}_6\text{Se}_{20}$ clusters can form at lower temperatures, this means that the activation energy of

forming $\text{Pd}_6\text{Se}_{20}$ clusters from the Pd and Se precursors is less than the activation energy needed for the clusters to assemble into an ordered extended solid. Consequently, the general mechanism of forming crystalline $\text{Pd}_3\text{Se}_{10}$ can be thought of as occurring in two steps: an initial faster step to create $\text{Pd}_6\text{Se}_{20}$ clusters, followed by a slower second step involving the clusters coming together and organizing themselves into a crystalline solid (Fig. S9, ESI†). In actuality, the first step of forming $\text{Pd}_6\text{Se}_{20}$ clusters must consist of many other elementary steps, such as substitution reactions to replace the chloride ligands in PdCl_4^{2-} with polyselenide ligands and reactions to allow for Pd atoms to be connected by bridging polyselenide ligands. Obtaining experimental evidence of these elementary steps, however, would require a much more detailed investigation that considers the air-free and high temperature reaction conditions required, as well as the short timescale of cluster formation, and is beyond the scope of this work.

Thermal stability

Since $\text{Pd}_3\text{Se}_{10}$ is a metastable material, it is critical to know its decomposition temperature, as this sets an upper limit on the possible temperatures it can be synthesized at. To determine the decomposition temperature of $\text{Pd}_3\text{Se}_{10}$ and gain an understanding of how it decomposes, a simultaneous TGA/DSC experiment was performed for a crystalline $\text{Pd}_3\text{Se}_{10}$ sample (Fig. 6A). The first significant feature in the DSC curve is an exothermic peak with an onset at 340 °C and a maximum at 350 °C. Although some sample mass loss occurs before 350 °C, the majority happens after the exotherm, as indicated by the large decrease in weight percent in the TGA data. The second feature in the DSC curve is an endotherm that immediately follows the exothermic process. This peak ends at 470 °C, which corresponds to the temperature that the TGA curve begins to flatten out. After the $\text{Pd}_3\text{Se}_{10}$ sample was heated to 500 °C, a PXRD pattern of the decomposition product was obtained

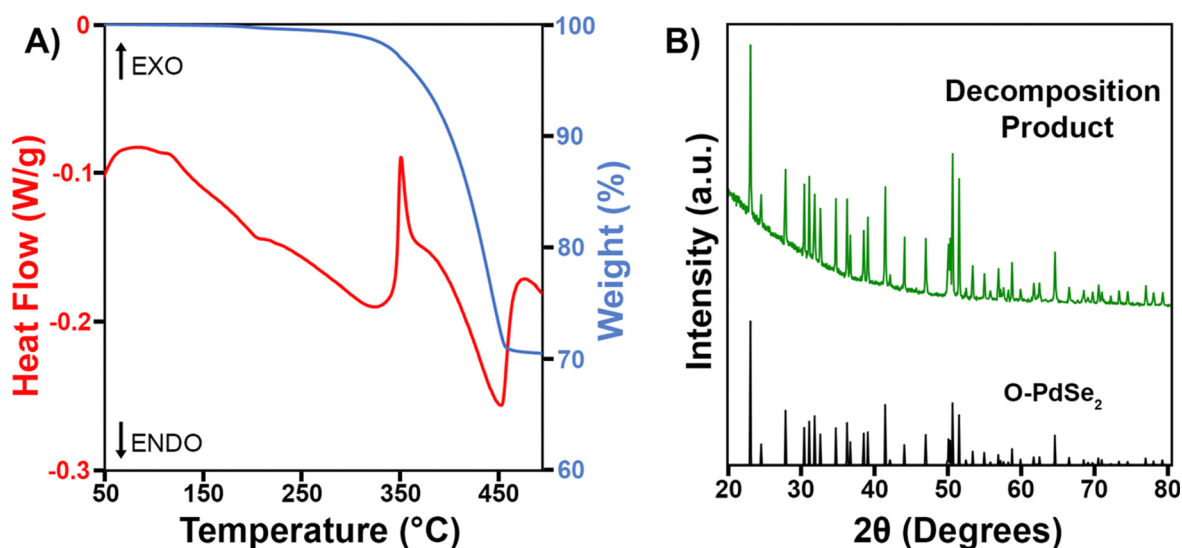


Fig. 6 (A) TGA/DSC plot for a crystalline $\text{Pd}_3\text{Se}_{10}$ sample. (B) PXRD pattern of the decomposition product obtained after heating $\text{Pd}_3\text{Se}_{10}$ to 500 °C in the TGA instrument.



(Fig. 6B). The pattern reveals that O-PdSe₂ formed due to the heat treatment and no Pd₃Se₁₀ remains.

Taking both the TGA/DSC and PXRD data in Fig. 6 into account, the Pd₃Se₁₀ decomposition temperature can be determined and a balanced chemical reaction for the thermal decomposition can be written (Scheme 3). First, we confirmed that the gradual downward slope of the DSC curve before the exotherm does not correspond to a phase transition by acquiring a PXRD pattern of the Pd₃Se₁₀ sample after heating it to 300 °C in the TGA instrument (Fig. S10, ESI†). This PXRD pattern shows that heating to 300 °C will not decompose Pd₃Se₁₀. However, it does cause a slight decrease in crystalline domain size as indicated by the broadening of the diffraction peaks after heating. Knowing this, the exothermic process at 350 °C can confidently be assigned to the transformation of Pd₃Se₁₀ into two more thermodynamically stable phases: O-PdSe₂ and selenium. The subsequent endothermic process and sample weight loss can be attributed to the sublimation of selenium, which leaves O-PdSe₂ as the only phase remaining after the 500 °C heat treatment. The total weight loss for the TGA experiment is 29.5%, which is very close to the 28.5% weight loss expected for the complete sublimation of the selenium produced by the equation in Scheme 3. The 1% discrepancy in weight percent could be due to the loss of residual solvent from the Pd₃Se₁₀ sample.

To compare, the thermal stability of an amorphous Pd₃Se₁₀ sample was also investigated using TGA/DSC and PXRD (Fig. S11 and S12, ESI†). Amorphous Pd₃Se₁₀ undergoes a nearly identical decomposition process as crystalline Pd₃Se₁₀. The DSC curve features an exothermic peak with a maximum at 330 °C, immediately followed by an endotherm and an increased rate of weight loss in the TGA plot. The total weight loss for the amorphous sample was 31.0% and the PXRD revealed that the decomposition product was O-PdSe₂. The similarity in this decomposition process provides supporting evidence that the amorphous sample is composed of Pd₆Se₂₀ clusters.

Optical and electronic properties

Diffuse reflectance spectroscopy (DRS) data of different Pd₃Se₁₀ samples was acquired to better understand the material's optical properties (Fig. 7). Fig. 7A shows the Tauc plot of Pd₃Se₁₀, assuming an indirect transition, for the sample with a crystalline domain size of 30 nm. This data indicates that Pd₃Se₁₀ is an indirect semiconductor with a 1.18 eV band gap. The direct gap occurs at 1.40 eV (Fig. S13, ESI†).

Probing deeper, the DRS data for the Pd₃Se₁₀ sample with a 30 nm domain size was compared to the sample with a 15 nm domain size and an amorphous sample (Fig. 7B). While the indirect optical gap for the 15 nm sample (1.14 eV) (Fig. S14, ESI†) was very similar to the 30 nm sample's gap, the 15 nm

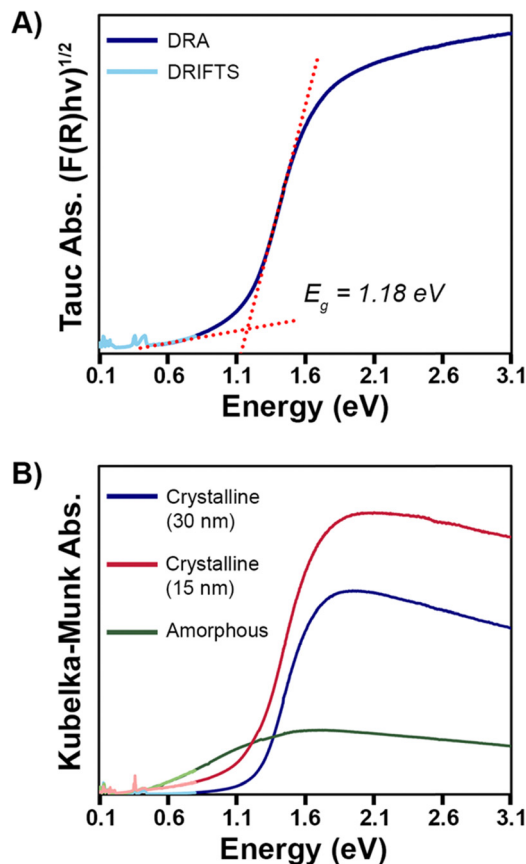
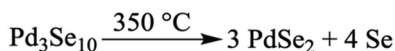


Fig. 7 (A) Tauc plot, assuming an indirect transition, for the Pd₃Se₁₀ sample with a crystalline domain size of 30 nm, as calculated from the Scherrer equation. (B) Diffuse reflectance spectra, transformed by the Kubelka–Munk function, for amorphous Pd₃Se₁₀ and crystalline Pd₃Se₁₀ samples with 15 nm and 30 nm domain sizes. The darker-colored (higher energy) part of each spectrum is data collected using diffuse reflectance absorbance (DRA) and the lighter-colored (lower energy) part of each spectrum is data collected using diffuse reflectance infrared Fourier transform spectroscopy (DRIFTS).

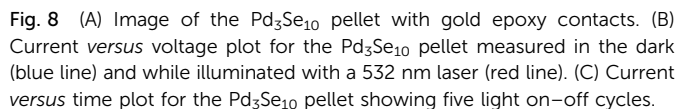
sample's DRS plot featured a broader absorption edge in the Urbach tail region. The Urbach tail is the exponential region of the band edge in between the baseline and linear region in a semiconductor's absorption plot. It is attributed to sub-bandgap absorption caused by a lack of crystallinity, defects, or impurities. The steepness of the Urbach tail can be quantified by the Urbach energy (E_U), which helps to assess the relative degree of structural disorder between different samples.^{35,65–69}

The Urbach energies for the 15 and 30 nm Pd₃Se₁₀ samples were extracted from the $\ln(F(R))$ vs. energy plots (Fig. S15 and Table S8, ESI†). The E_U for the 15 nm sample (0.346 eV) was 1.6 times greater than the E_U for the 30 nm sample (0.217 eV). This analysis helps quantitatively show that the 15 nm sample has a larger degree of structural disorder and consequently more sub-bandgap absorption than the 30 nm sample, as would be expected for a material with a smaller crystalline grain size. For the amorphous Pd₃Se₁₀ sample, an optical gap was observed at 0.25 eV (Fig. S16, ESI†) and there wasn't an



Scheme 3 Thermal decomposition reaction of Pd₃Se₁₀.





The electronic resistivity and photoconductivity measurements on a polycrystalline $\text{Pd}_3\text{Se}_{10}$ pellet provided further evidence of the material's semiconducting behaviour. $\text{Pd}_3\text{Se}_{10}$ was pressed into a pellet and sintered at 120 °C to produce a sample with an estimated density of 75%. Electrical contacts on the pellet were made using gold epoxy (Fig. 8a). Two-probe current-voltage measurements at room temperature and in the dark showed ohmic behaviour (Fig. 8b), and based on the pellet dimensions, the material's very high resistivity was estimated to be $3.4 \times 10^7 \Omega \text{ m}$. Additionally, when the pellet was illuminated with 532 nm above band gap laser light, the slope of the current-voltage plot increased, indicative of an increase

Conclusions

Pd₃Se₁₀ is a new Pd₆Se₂₀ cluster-based semiconducting material that shows intriguing optoelectronic properties. To overcome reproducibility issues that plagued its initial synthetic method, we developed a solution-phase electrochemically balanced reaction utilizing Na₂PdCl₄ and a 2 : 1 molar mixture of Na₂Se₃ and Na₂Se₄. Under optimized conditions, this reaction allowed for the reliable and high-yield production of crystalline Pd₃Se₁₀. The reason that Pd₃Se₁₀ can form in solution was also a point of interest, considering that convex hull calculations predict it to be a metastable phase. This prediction was supported experimentally by TGA/DSC and PXRD data which show that it decomposes at 350 °C into the more thermodynamically stable phases O-PdSe₂ and Se. Further calculations suggest that the ability of Pd₃Se₁₀ to form in solution is likely due to its lower surface energies relative to O-PdSe₂. Raman spectroscopy showed that Pd₆Se₂₀ clusters were present in both the crystalline sample and the amorphous samples which formed below the optimal synthesis temperature. From this, it was inferred that Pd₆Se₂₀ clusters form in solution first before assembling into a crystalline lattice. Lastly, DRS, electrical resistivity, and photoconductivity measurements show that Pd₃Se₁₀ is semiconducting, with an indirect band gap of 1.18 eV. Overall, Pd₆Se₂₀ is an exciting new entry into the library of superatomic clusters and materials.

Author contributions

C. T. I. and J. E. G. designed and organized the project. C. T. I. and C. W. conducted the synthesis and most characterization of the materials. A. R. and D. W. M. conducted and analysed the HAADF-STEM. Y. K. assisted with the collection and analysis of XPS data. B. Y. and W. W. performed and analysed theoretical computations. T. S. conducted and helped analyse the resistivity and photoconductivity measurements. C. T. I. and J. E. G. wrote the manuscript with contributions from all authors.

Data availability

The data supporting this article have been included as part of the ESI.†

Conflicts of interest

There are no conflicts to declare.

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