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## Coupling of manganese oxidative crystallization and thallium immobilization under light irradiation

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**This study investigates the mutual influence of Mn(II) oxidative crystallization and thallium (Tl) immobilization in nitrate solution under light irradiation. MnO<sub>x</sub> formed by the photochemical oxidation of Mn(II) effectively removed Tl(I) through adsorption and coprecipitation, with Mn(II) inhibiting the oxidation of Tl(I) to more toxic Tl(III). In turn, Tl(I) suppressed the reductive transformation of δ-MnO<sub>2</sub>, which favored its own immobilization.**

Manganese (oxyhydr)oxides (MnO<sub>x</sub>) are ubiquitously distributed across various geological settings, including marine, soil, and sediment.<sup>1</sup> Due to their relatively high specific surface area and high redox and adsorption reactivity, MnO<sub>x</sub> exert strong control over the geochemical processes on the Earth's surface, such as the redox cycling, as well as the migration and transformation of environmental substances.<sup>2</sup> The reactivity of MnO<sub>x</sub> is intrinsically governed by their crystalline structure and valence state composition, making the understanding of their formation and transformation processes fundamentally essential.

Soluble Mn(II) ions are readily oxidized to MnO<sub>x</sub> through biotic and abiotic processes. Microbially mediated Mn(II) oxidation is recognized as the predominant pathway for natural MnO<sub>x</sub> formation, owing to its rapid oxidation kinetics (several orders of magnitude faster than abiotic processes).<sup>3,4</sup> However, emerging research has revealed that photochemically produced reactive intermediates generated from various environmental components (e.g., nitrate,<sup>5</sup> natural semiconductor minerals,<sup>6</sup> nanoplastics,<sup>7</sup> and dissolved organic matter<sup>8</sup>) can oxidize Mn(II) at rates comparable to those of currently known biotic and abiotic

oxidation processes. This finding indicates that the photochemical oxidation of Mn(II) plays a vital role in the formation of MnO<sub>x</sub> in natural systems.

As important natural metal scavengers, MnO<sub>x</sub> can effectively immobilize various metal ions through multiple mechanisms, including adsorption, oxidation, ion exchange, and surface-induced precipitation.<sup>9,10</sup> One example is thallium (Tl), a highly toxic but underestimated heavy metal.<sup>11,12</sup> Due to abundant natural (e.g., weathering of Tl-bearing rocks/minerals) and anthropogenic (e.g., mining, smelting, and coal combustion) sources, Tl is widely distributed in water, soil, sediment, and the atmosphere.<sup>13,14</sup> Monovalent Tl(I) primarily exists as free cations, while trivalent Tl(III) is readily hydrolyzed to form avicennite.<sup>15</sup> The formation and transformation of MnO<sub>x</sub> play crucial roles in the biogeochemical cycling of Tl due to the high geochemical affinity and strong correlation between MnO<sub>x</sub> and Tl.<sup>16,17</sup> For example, Tl(I) is effectively removed during biogenic Mn(II) oxidation, with 64% of Tl(I) being oxidized to Tl(III).<sup>18</sup> Furthermore, anoxic MnO<sub>2</sub> transformation driven by Mn(II)/Fe(II) can enhance Tl mobilization.<sup>19</sup> However, the coupled cycling of Mn and Tl under light irradiation in the Earth's surface environments remains poorly elucidated.

Herein, we focus on an important natural abiotic Mn–Tl process for the first time and reveal the details about the coupling between the Mn redox cycle and Tl immobilization under irradiation. The proposed light-driven Mn redox cycle is composed of three key processes: (i) the oxidation of Mn(II) to δ-MnO<sub>2</sub> by superoxide radicals (O<sub>2</sub><sup>•−</sup>) generated from the photolysis of nitrate; (ii) the reductive transformation of δ-MnO<sub>2</sub> via Mn(II) + Mn(IV) ⇌ 2Mn(III); and (iii) the photoreduction of the formed MnO<sub>x</sub> and reoxidation of Mn(II). The formed MnO<sub>x</sub> effectively removed Tl(I) through adsorption and coprecipitation, with Mn(II) inhibiting Tl(I) oxidation. In turn, Tl(I) suppressed the reductive transformation of δ-MnO<sub>2</sub> due to the competition between Mn(II) and Tl(I). Our results enhance the comprehension of the coupled environmental geochemical processes involving Mn and Tl in natural systems and offer new insights for environmental remediation.

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Fig. 3 (a) The effect of SOD on the removal kinetics of Mn. (b) The EPR spectrum of O<sub>2</sub><sup>•-</sup> trapped by DMPO. (c) Ti L<sub>III</sub>-edge XANES and (d) k<sup>3</sup>-weighted EXAFS spectra of the solid samples collected at 8 h from Ti/Mn-0.012 and Ti/Mn-0.12.

spectroscopy using DMPO as a spin-trapping agent (Fig. 3b). A distinctive absorption peak at 258 nm (Fig. S9, SI), originating from the Mn(III)-pyrophosphate complex, indicated the formation of Mn(III) intermediates.<sup>24</sup> Mn(III) is thermodynamically unstable and can be further converted into Mn(IV) through either disproportionation or oxidation. Recent research has demonstrated that the disproportionation of Mn(III) intermediates is the main pathway for the photochemical generation of Mn(IV) oxide.<sup>25</sup>

$\delta$ -MnO<sub>2</sub> is easily transformed by changes in external redox conditions. The concentration of Mn(II) serves as a critical factor governing the structural evolution of  $\delta$ -MnO<sub>2</sub>. At pH 8.0, no change in the phase of  $\delta$ -MnO<sub>2</sub> was detected when [Mn(II)]/[Mn] ≤ 2%, whereas the transformation to feitknechtite occurred when the ratio reached 10%.<sup>26</sup> In this study, the rapid decline of Mn(II) concentration during the reaction (Fig. 1a) restricted the comproportionation reactions to the initial stage, resulting in the decrease of Mn AOS (Fig. 1c). Moreover, the interaction between  $\delta$ -MnO<sub>2</sub> and Mn(II) is dependent on pH. Mn(II) only interfered with the stacking of  $\delta$ -MnO<sub>2</sub> sheets at pH < 7.0, while  $\delta$ -MnO<sub>2</sub> underwent bulk structural conversion to manganite and hausmannite through a feitknechtite reaction intermediate within the pH range of 7.0–8.5.<sup>20</sup> This transformation process elucidates the coexistence of feitknechtite and hausmannite with  $\delta$ -MnO<sub>2</sub> in the present study, in contrast to the sole presence of  $\delta$ -MnO<sub>2</sub> observed in an earlier study where the pH decreased from 8.9 to around 7.0 after 1 h.<sup>5</sup> In addition to Mn(II) concentration and pH, coexisting ions control the transformation of  $\delta$ -MnO<sub>2</sub>.<sup>27</sup> It is well established that metal cations can kinetically hinder the transformation of  $\delta$ -MnO<sub>2</sub> by competing with Mn(II) for active sites.<sup>28,29</sup> Similarly, the competitive adsorption between Mn(II) and Tl on  $\delta$ -MnO<sub>2</sub> in

this study inhibited its reductive transformation, thereby leading to an increase in the apparent oxidation rate of Mn(II) (Table S1, SI).

The formation and transformation of MnO<sub>x</sub> significantly influence the fate and transport of Tl. Tl and Mn showed similar distribution patterns, and the atomic ratios of Tl reached 0.20% and 1.17% in the solid samples collected from Ti/Mn-0.012 and Ti/Mn-0.12 at 8 h, respectively, confirming the effective Tl immobilization by MnO<sub>x</sub> (Fig. S10, SI). During the formation of  $\delta$ -MnO<sub>2</sub>, cations (*e.g.*, Mn, Na, Tl) are adsorbed to neutralize the negative charges generated by Mn(IV) vacancies.<sup>30</sup> The fraction of NH<sub>2</sub>OH·HCl-extractable Tl (*i.e.*, more stable Tl species) in *in situ* formed MnO<sub>x</sub> was higher than that in *ex situ* MnO<sub>x</sub> (Fig. S11, SI), indicating that partial Tl was removed through the coprecipitation process during the nucleation and growth of MnO<sub>x</sub>. Notably, unlike transition metals such as Ni(II) and Zn(II), Tl(I) has a significantly different ionic radius from Mn(III) and Mn(IV), making its incorporation into the MnO<sub>x</sub> lattice unlikely.<sup>31</sup> The Tl L<sub>III</sub>-edge X-ray absorption near-edge structure (XANES) spectra indicated that the sorbed Tl was primarily Tl(I) (Fig. 3c), consistent with the XPS results (Fig. S12, SI). Moreover, the corresponding extended X-ray absorption fine structure (EXAFS) spectra showed very low first-shell amplitudes similar to TiNO<sub>3</sub> (Fig. 3d). Tl(I) can be oxidized to Tl(III) by the reactive species generated from nitrate photolysis or the vacant octahedral sites of  $\delta$ -MnO<sub>2</sub>.<sup>32</sup> However, the oxidation of Tl(I) was significantly suppressed in the presence of Mn(II), as evidenced by the negligible detection of Tl(III) in both solution (Fig. S13, SI) and MnO<sub>x</sub> (Fig. 3c). The standard redox potential of Tl(III)/Tl(I) ( $E_0 = +1.252$  V) is slightly higher than that of Mn(IV)/Mn(II) ( $E_0 = +1.224$  V), indicating a thermodynamic preference for the oxidation of Mn(II) over Tl(I) by reactive oxygen species when they coexist. Furthermore, Mn(II) + Mn(IV) ⇌ 2Mn(III) leads to a reduction in the oxidative capacity of  $\delta$ -MnO<sub>2</sub> toward Tl(I). Similarly, partial reduction of  $\delta$ -MnO<sub>2</sub> induced by Fe(II) or HEPES effectively inhibited the oxidative uptake of Tl.<sup>33</sup> In addition, the transformation process of  $\delta$ -MnO<sub>2</sub> will lead to the repartitioning of Tl due to the distinct affinities of MnO<sub>x</sub> with different properties and structures toward Tl. Recent studies have shown that the conversion of  $\delta$ -MnO<sub>2</sub> to manganite leads to the release of Tl in highly reducing environments (10 mM Mn(II)).<sup>19</sup> Although the reductive transformation of  $\delta$ -MnO<sub>2</sub> was unfavorable for Tl immobilization (Fig. S14, SI), the concentration of Tl continued to decrease overall (Fig. 1d), which was attributed to the continuous generation of  $\delta$ -MnO<sub>2</sub> with high sorption capacity.

In conclusion, we revealed the details about the coupling of the light-driven Mn redox cycle with Tl immobilization (Fig. 4). Mn(II) was oxidized to  $\delta$ -MnO<sub>2</sub> by O<sub>2</sub><sup>•-</sup> generated during the photolysis of nitrate, while surface-adsorbed Mn(II) facilitated the reductive transformation of  $\delta$ -MnO<sub>2</sub>. Furthermore, photo-reduction of the formed MnO<sub>x</sub> produced Mn(II) for the next round of oxidation, which facilitated the continuous generation of  $\delta$ -MnO<sub>2</sub>. In the presence of Tl(I), the phase transformation process was suppressed due to the competition between Mn(II) and Tl(I). In addition, the oxidation of Tl(I) was inhibited by



Fig. 4 Schematic for the coupling of the Mn redox cycle with Tl immobilization.

Mn(II), and Tl(I) was effectively removed through adsorption and coprecipitation during the formation of MnO<sub>x</sub>. The concentrations of Mn and nitrate, as well as the Tl/Mn molar ratio employed in this study, are environmentally relevant.<sup>33–35</sup> Notably, the coupled effect of Mn(II) oxidative crystallization and Tl immobilization was also observed under natural solar irradiation, albeit proceeding at a slower rate compared to laboratory experiments (Fig. S15, SI). This study validates the crucial role of light in driving Mn redox cycling, particularly in environments unfavorable for microorganisms (e.g., extreme temperatures, strong acidity/alkalinity). This enhances our understanding of MnO<sub>x</sub> phase evolution in the environmental Mn cycle, explaining their diversity in nature. Furthermore, Mn redox cycling is often coupled with elemental cycles and organic matter degradation. To the best of our knowledge, this study is the first to uncover the complex photochemical interactions between Mn and Tl. The UV/NO<sub>3</sub><sup>-</sup> technology demonstrated here presents a practical pathway for treating industrial wastewater containing Mn and other metal contaminants, enabling their simultaneous removal and facilitating efficient metal recovery and utilization.

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## Conflicts of interest

There are no conflicts to declare.

## Data availability

The data supporting this article have been included as part of the supplementary information (SI). Supplementary information includes experimental details, additional figures, and supporting tables. See DOI: <https://doi.org/10.1039/d5cc05797c>.

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