



Cite this: *Chem. Commun.*, 2020, 56, 12090

Metal phosphonates incorporating metalloligands: assembly, structures and properties

Song-Song Bao, Ming-Feng Qin and Li-Min Zheng *

Metal phosphonates are an important class of metal–organic hybrid materials that exhibit versatile structures, intriguing functions and high water and thermal stability. Despite a large number of metal phosphonates achieved over the past few decades, those incorporating metalloligands are rather limited. The metalloligand approach can provide a unique opportunity in constructing homo- or mixed-metal coordination polymers with rationally immobilized functional moieties for various applications in gas storage and separation, heterogeneous catalysis, sensing and multifunctional materials. In this feature article, we shall introduce the current status of a special subclass of metal phosphonates, namely, metal–metalloligand phosphonates (MMPs), including synthetic strategies, crystal structures, and properties of those based on paddlewheel diruthenium, metallo-polyazamacrocyclic, metalloporphyrin and metallo-tris-bipyridine ligands. Future challenges in this field are discussed.

Received 1st June 2020,
Accepted 27th August 2020

DOI: 10.1039/d0cc03850d

rsc.li/chemcomm

1. Introduction

Metal–organic hybrid materials, in particular metal–organic frameworks (MOFs), have been developed explosively over the past few decades.^{1–3} These materials are commonly constructed by a combination of metal-containing nodes and organic linkers, wherein a large number of organic linkers with different shapes have been designed and synthesized. Aside from enormous efforts devoted to homometallic MOFs with beautiful structures and

fascinating properties, heterometallic systems containing metalloligands have attracted more and more attention. Metalloligands are defined as metal complexes with appended functional groups that are able to coordinate with second metal ions. Within the context of the metalloligand approach, it is possible to place the auxiliary functional groups in a pre-organized conformation.⁴ Hence this approach offers opportunities to rationally immobilize functionalized or catalytically active metal sites within the frameworks. As a result, enhanced performance in gas storage and separation, heterogeneous catalysis, sensing, or multifunction can be achieved in metal–metalloligand systems.⁵

State Key Laboratory of Coordination Chemistry, School of Chemistry and Chemical Engineering, Collaborative Innovation Centre of Advanced Microstructures, Nanjing University, Nanjing 210023, China. E-mail: lmzheng@nju.edu.cn



Song-Song Bao

Song-Song Bao completed his PhD degree at Nanjing University in 2008 under the supervision of Prof. Li-Min Zheng. He then studied at Goethe University Frankfurt as a Post-Doctoral Research Associate in the research group of Prof. Matthias Wagner. He is currently working as associate professor at the School of Chemistry and Chemical Engineering, Nanjing University. His current interests involve the design and synthesis of

multifunctional metal phosphonates by using the metalloligand approach.



Ming-Feng Qin

Ming-Feng Qin received his bachelor's degree and master's degree from Jiangsu University in 2015 and 2018, respectively. He is currently working on his PhD under the supervision of Prof. Li-Min Zheng at the School of Chemistry and Chemical Engineering at Nanjing University. His current research interests focus on photocatalytic properties of metal phosphonates containing metalloligands.

Metal phosphonates are an important class of metal–organic hybrid materials, showing promising applications in ion-exchange and sorption, catalysis, and materials science.⁶ Compared with many other metal–organic systems, metal phosphonates often show high water and thermal stability,^{7,8} attributed to the presence of inorganic components in the structure such as clusters, chains or layers. The organic moieties of the phosphonate ligands RPO_3^{2-} can be modified by other functional groups resulting in a number of metal phosphonates with versatile structures and functions. For more information, the interested reader is directed to a book which covers most aspects of this field,⁹ and several review articles concerned with the synthesis and structures,¹⁰ optical,¹¹ magnetic,¹² and proton conductive properties,¹³ as well as their application in catalysis, adsorption, energy storage, drug delivery, *etc.*¹⁴ In contrast to the rapid progress of metal phosphonate chemistry, utilization of metalloligands in synthesis has not been well explored. The incorporation of metalloligands in metal phosphonates can not only confer additional function to the material but also provide possibilities to tune the properties through different metal components. However, the number of known metal–metalloligand phosphonates (MMPs) are rather limited. The obstacle could arise from the synthetic challenges in isolating phosphonate-based metalloligands, and the fast precipitation of the MMP products leading to the difficulties in analysing their structures using the conventional single crystal determination method. To solve the structures of MMPs with extremely small crystal sizes, Rietveld refinement using powder X-ray diffraction data as well as single-crystal electron diffraction method has been employed.¹⁵

In this feature article, we shall introduce the current status of metal phosphonates incorporating metalloligands including the general synthetic strategies, crystal structures and properties. It is worth mentioning that the metalloligands discussed herein are restricted to those with pendant phosphonate groups such that each phosphonate group can provide at least



Li-Min Zheng

her research group seeks to develop new crystalline and low-dimensional materials based on metal phosphonates with interesting physical and chemical properties.

Li-Min Zheng received her PhD degree from Nanjing University in 1992 under the guidance of Prof. Xinquan Xin. Then she became a lecturer (1992–1997), associate professor (1997–2002) and professor (2002–) at the same University. She spent two years with Prof. Silvio Decurtins as a postdoctoral fellow at the University of Zurich (1994–1996) and one year with Prof. Allan J. Jacobson at the University of Houston (1999–2000). Her



Scheme 1 The coordination modes of a phosphonate-based metalloligand.

one oxygen donor for further coordination (Scheme 1). As far as we are aware, there are only a few types of such metalloligands reported so far, including paddlewheel diruthenium diphosphonate, metallo-polyazamacrocyclic-phosphonate, metalloporphyrin-phosphonate and metallo-tris-bipyridine-phosphonate ligands. These metalloligands share a common feature in that they can be isolated as discrete species and used as linkers to connect second metal ions. But not all MMPs containing these metalloligands were synthesized using metalloligands as the precursor. In some cases, they were obtained through direct reaction of metal salt and phosphonic acid. In order to give readers a whole picture of these systems, all MMPs containing these metalloligands will be summarized, and their structures and properties will be described. Finally, this article will end up with a discussion of future challenges in this field.

2. Synthetic strategies

There are two general synthetic strategies to produce MMPs based on the different precursors as linkers. First, a one-step procedure using phosphonic acid as the precursor. The MMPs can be obtained by direct reaction of appropriate metal salts with phosphonic acids in solution under ambient or hydro/solvothermal conditions. This approach is simple but challenging in controlling the final products. In particular, when there are more than one kind of metal ions present, it can be difficult to obtain heterometallic MMPs in pure form. Further, the metalloligand could not be formed as anticipated under these one-pot reaction conditions. Second, a two-step procedure using metalloligands as the precursor. This approach involves the preparation of the desired metalloligands, followed by the reaction of the metalloligands with suitable metal salts. The advantage of this approach is that the pre-formed metalloligand can not only immobilize the central metal ion but also provide coordination donors with different orientations, thus allowing the formation of MMPs in a more controllable manner. However, considering the strong coordination capabilities of the phosphonate groups, it could be a challenging task to isolate these phosphonate-based metalloligands as linkers for further assembling with other metal ions. It should be emphasized that in both strategies, other reaction parameters such as the molar ratio and concentration of the metal ion and linkers, the pH of the reaction mixture, the reaction temperature and time as well as the solvents are also critical

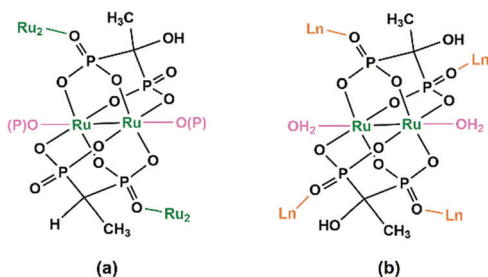
for achieving crystalline materials of MMPs suitable for single-crystal X-ray diffraction studies.

3. Crystal structures

3.1. Diruthenium paddlewheel units as metalloligands

Diruthenium Ru_2 (4+ to 6+) cores with paddlewheel structures are attractive metalloligands for functional MOFs owing to their variable electronic configurations and magnetic and catalytic properties.¹⁶ Successful examples include those with the paddlewheel dimers constructed from chelating and bridging O,O'-donor ligands, such as carboxylate,¹⁷ carbonate¹⁸ and phosphate.¹⁹ The Ru_2 -diphosphonate system has superiority over carbonate because the organic groups of the phosphonate ligand can be modified with additional functions. In 2003, we reported the first example of metal phosphonates with extended structures which contain a diruthenium diphosphonate paddlewheel core, $(\text{NH}_4)_3\text{Ru}_2(\text{hedp})_2 \cdot 2\text{H}_2\text{O}$ (**1**),²⁰ where hedpH_4 is 1-hydroxyethylidenediphosphonic acid $[\text{CH}_3\text{C}(\text{OH})(\text{PO}_3\text{H}_2)_2]$. This compound was synthesized by a one-pot reaction of RuCl_3 and hedpH_4 under hydrothermal conditions. It has a 2D layer structure in which the negatively charged $\text{Ru}_2^{\text{II,III}}(\text{hedp})_2^{3-}$ paddlewheels are inter-connected by filling the axial sites of ruthenium with the phosphonate oxygen atoms from the equivalent $\text{Ru}_2(\text{hedp})_2^{3-}$ units. The interlayer space is filled with NH_4^+ cations and lattice water molecules. Here, the diruthenium motif is not only an acceptor but also a donor (Scheme 2a). A similar layer topology is also found in $\text{Na}_3\text{Ru}_2(\text{hedp})_2 \cdot 4\text{H}_2\text{O}$ (**2**).²¹ In this case, the layers are further connected by $\{\text{NaO}_6\}$ octahedra, forming a 3D framework structure.

Interestingly, compound **1** is soluble in an aqueous solution of salt with the dissociation of the framework structure and the formation of stable dinuclear species of $\text{Ru}_2(\text{hedp})_2^{3-}$, confirmed by the isolation of $\text{K}_3[\text{Ru}_2(\text{hedp})_2(\text{H}_2\text{O})_2] \cdot 6\text{H}_2\text{O}$.²² The pre-formed $\text{Ru}_2(\text{hedp})_2^{3-}$ paddlewheel is stable in aqueous solution and can provide four pendant phosphonate oxygen atoms and two hydroxyl groups as potential donors. Thus it can serve as a metalloligand to assemble with second metal ions to construct MMPs. Indeed, the direct reaction of $\text{Ru}_2(\text{hedp})_2^{3-}$ solution with $\text{Ln}(\text{NO}_3)_3$ ($\text{Ln} = \text{La}, \text{Ce}, \text{Pr}, \text{Nd}, \text{Sm}, \text{Eu}, \text{Gd}, \text{Tb}, \text{Dy}, \text{Ho}, \text{Er}$) at room temperature led to two types of Ln-Ru₂ compounds with the formula $\text{Ln}(\text{H}_2\text{O})_4[\text{Ru}_2(\text{hedp})_2(\text{H}_2\text{O})_2] \cdot 5.5\text{H}_2\text{O}$ [$\text{Ln} = \text{La}$ (**3**), Ce (**4**)] and $\text{Ln}(\text{H}_2\text{O})_4[\text{Ru}_2(\text{hedp})_2(\text{H}_2\text{O})_2] \cdot 8\text{H}_2\text{O}$



Scheme 2 Coordination modes of $\text{Ru}_2(\text{hedp})_2^{n-}$ in **1** (a) and **3–15** (b).

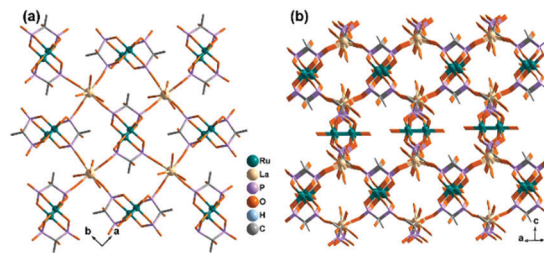


Fig. 1 (a) One square-grid layer of structure **3**. (b) 3D PtS-type open-framework structure of **5**. All H atoms are omitted for clarity. Adapted from ref. 23.

[$\text{Ln} = \text{La}$ (**5**), Ce (**6**), Pr (**7**), Nd (**8**), Sm (**9**), Eu (**10**), Gd (**11**), Tb (**12**), Dy (**13**), Ho (**14**), Er (**15**)].²³ In all cases, $\text{Ru}_2(\text{hedp})_2(\text{H}_2\text{O})_2^{3-}$ acts as a tetra-dentate metalloligand, coordinating to four Ln^{III} ions using its four pendant phosphonate oxygen atoms (Scheme 2b). Compounds **3** and **4** are isostructural. They have a square-grid layer structure in which each $\text{Ru}_2(\text{hedp})_2(\text{H}_2\text{O})_2^{3-}$ binds to four Ln^{III} ions and *vice versa* (Fig. 1a). Compounds **5–15** are also isostructural, but possess a 3D PtS-type open-framework structure (Fig. 1b). The lanthanide ions are eight-coordinated by four phosphonate oxygen (O_p) atoms and four water molecules (O_w) and the hydroxyl groups remain pendant. Despite similar connectivity of $\text{Ru}_2(\text{hedp})_2(\text{H}_2\text{O})_2^{3-}$ and Ln^{III} ions, the significant structural difference between **3–4** and **5–15** originates from the slight difference in the geometry of $\{\text{LnO}_4\text{Ow}_4\}$ polyhedra. In particular, the geometry of $\{\text{LnO}_4\}$ is essential in determining the topology of the final products, which is a distorted square plane in **3–4**, but closer to a distorted tetrahedron in **5–15**. Fig. 1 shows the single layer and the framework structures of **3** and **5**, respectively. **5** contains three kinds of channels, A ($4.3 \times 3.5 \text{ \AA}$ along the $[330]$), B ($3.6 \times 4.9 \text{ \AA}$ along the $[330]$) and C ($3.5 \times 3.5 \text{ \AA}$ along the $[001]$) (van der Waals radii accounted), which host the lattice water molecules. As far as we are aware, these are the first examples of Ln-Ru₂ heterometallic compounds containing paddlewheel Ru_2^{n+} cores.

By functionalizing the methylenediphosphonate ligand with an amino group, we obtained a 1D $\text{Ru}_2(\text{II,II})$ compound $(\text{H}_3\text{O})_2[\text{Ru}_2(\text{Hamdp})_2]$ (**16**), where amdp^{4-} represents 1-ammonium-methylene-diphosphonate $[\text{NH}_2\text{CH}(\text{PO}_3)_2^{4-}]$. **16** is soluble in aqueous solution with the formation of discrete species of $\text{Ru}_2(\text{Hamdp})_2$. The assembly of $\text{Ru}_2(\text{Hamdp})_2$ with $\text{Ln}(\text{NO}_3)_3$ ($\text{Ln} = \text{Dy}, \text{Yb}$) in aqueous solution at room temperature resulted in mixed-valent compound $\text{Dy}(\text{H}_2\text{O})_3[\text{Ru}_2(\text{Hamdp})_2][\text{Ru}_2(\text{Hamdp})_2(\text{H}_2\text{O})_2]_{0.5} \cdot 12\text{H}_2\text{O}$ (**17**) with a 2D double layer structure and $\text{Yb}[\text{Ru}_2(\text{Hamdp})_2]_2[\text{Ru}_2(\text{Hamdp})_2(\text{H}_2\text{O})_2] \cdot 15\text{H}_2\text{O}$ (**18**) with a 3D porous framework structure (Fig. 2).²⁴ The mixed-valent nature of the Ru_2 units is suggested by their Ru-Ru distances and confirmed by magnetic studies.

The structures of **17** and **18** are unique compared to other known diruthenium coordination polymers, and hence deserve further description. In **17**, three types of $\text{Ru}_2(\text{H}_x\text{amdp})_2$ units are present serving as bi-, tri- and tetra-dentate metalloligands, respectively (Scheme 3a). The coordination of $\text{Ru}_2(\text{Hamdp})_2^{2-}$

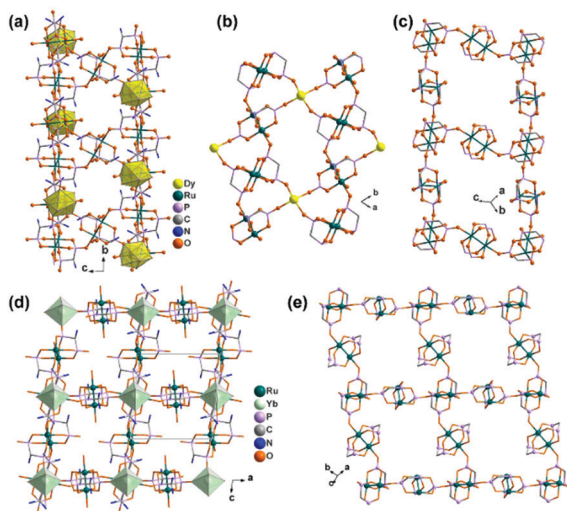
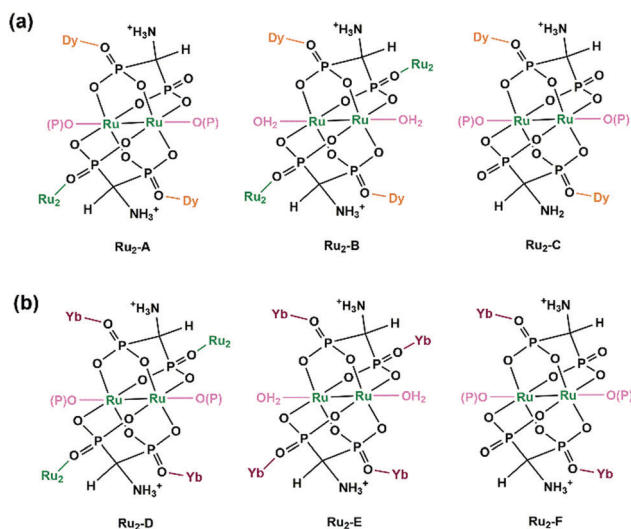


Fig. 2 (a) A double layer structure of **17**. (b) The kagomé layer made up of $\text{Ru}_2(\text{Hamdp})_2^-$ and Dy^{III} ions in structure **17**. (c) The ladder-like chain of pure Ru_2 in structure **17**. (d) 3D open-framework structure of **18**. (e) The Ru_2 layer in structure **18**. All H atoms, $-\text{NH}_3^+$ groups, uncoordinated phosphonate O atoms and coordination water molecules are omitted in (b, c and e) for clarity. Adapted from ref. 24.

(tri- and tetra-dentate) with Dy^{III} ions leads to a kagomé layer (Fig. 2b), and two such layers are further pillared by the third type of $\text{Ru}_2(\text{Hamdp})(\text{amdp})^{2-}$ (bi-dentate) forming a double layer structure (Fig. 2a and c). In **18**, there exist three types of $\text{Ru}_2(\text{H}_x\text{amdp})_2$ units acting as bi- and tetra-dentate metalloligands (Scheme 3b). Although the coordination mode of the bi-dentate Ru_2 ($\text{Ru}_2\text{-F}$, Scheme 3b) is similar to that in **17** ($\text{Ru}_2\text{-C}$, Scheme 3a), that of the tetra-dentate ones is remarkably different. There are two kinds of tetra-dentate $\text{Ru}_2(\text{Hamdp})_2^-$ metalloligands in **18**. One of them binds to two other Ru_2 units and two Yb^{III} ions and accepts two phosphonate oxygen atoms in the



Scheme 3 Coordination modes of $\text{Ru}_2(\text{H}_x\text{amdp})_2$ dimers in **17** (a) and **18** (b). Reprinted with permission from ref. 24. Copyright 2019, American Chemical Society.

axial positions ($\text{Ru}_2\text{-D}$), while the other binds to four Yb^{III} ions with the axial positions filled with water molecules ($\text{Ru}_2\text{-E}$). These three types of Ru_2 units are cross-linked to form a layer containing rhombic eight-member rings of Ru_2 (Fig. 2e), unlike four-member rings in **1**. The Ru_2 layers are stacked by six-coordinated Yb into a 3D framework (Fig. 2d). The framework contains two kinds of narrow channels, A (approximately $4.1 \times 6.7 \text{ \AA}$ along the c -axis) and B (approximately $6.6 \times 10.6 \text{ \AA}$ along the a -axis) (van der Waals radii not accounted), which host the lattice water molecules.

Apparently, paddlewheel diruthenium diphosphonates can be used as metalloligands to successfully assemble Ln-Ru_2 MMPs. Compared with the $\text{Ru}_2(\text{hedp})_2/\text{Ln}$ complexes,²³ however, compounds **17** and **18** exhibit entirely new structural topologies, attributed to the different charges of the $\text{Ru}_2(\text{H}_x\text{amdp})_2$ unit due to the presence of protonated amino group in the diphosphonate ligand. This opens a new route to design and synthesize Ru_2 -based materials with new architectures by modifying the organic groups of the diphosphonate ligand. Table 1 summarizes MMPs containing $\text{Ru}_2(\text{hedp})_2$ and $\text{Ru}_2(\text{H}_x\text{amdp})_2$ metalloligands.

3.2. Metallo-polyazamacrocycle units as metalloligands

Polyazamacrocycles functionalized by phosphonate groups offer multiple metal binding sites of two distinct types of chemical nature, one at the phosphonate pendant arms and the other within the polyazamacrocycle ring. Table 2 gives the MMPs containing metallo-polyazamacrocycle ligands. 1,4,7-Triazacyclononane-1,4,7-triyl-tris(methylenephosphonic acid) (notpH_6) contains the smallest polyazamacrocycle with three nitrogen and nine oxygen atoms as potential coordination donors. It is ideally suited to form six-coordinate complexes with two groups of facial donors: the macrocyclic nitrogen atoms on one side and the pendant phosphonate oxygen atoms on the other side, thus leading to mononuclear species such as $\text{M}^{\text{III}}(\text{notpH}_3)$ ($\text{M} = \text{Fe}, \text{Co}$)²⁵ and $\text{Cu}^{\text{II}}(\text{notpH}_4)$ ²⁶ (Scheme 4). These mononuclear species are polar and can serve as metallo-ligands, by using the remaining six uncoordinated phosphonate oxygen atoms, to connect other metal ions into extended networks. Moreover, the octahedrally coordinated Fe^{III} or Co^{III} centers are chiral showing Δ or Λ configurations. Unfortunately, enantioenriched MMPs containing $\text{M}(\text{notpH}_x)$ metallo-ligands have not been reported so far. Only in rare cases, racemic conglomerates were obtained which contain an equal amount of opposite enantiomers.

In 2006 we reported the first example of coordination polymers containing the $\text{M}(\text{notp})$ unit with the formula $\text{Mn}_3(\text{notp})(\text{H}_2\text{O})_6 \cdot 1.5\text{H}_2\text{O}$ (**19**),²⁷ which was obtained by direct reaction of manganese salt and notpH_6 under hydrothermal conditions. This compound crystallizes in a highly symmetrical trigonal system with space group $R\bar{3}$. Each $\text{Mn}^{\text{II}}(\text{notp})^{4-}$ binds to six Mn^{II} ions *via* corner-sharing of $\{\text{PO}_3\text{C}\}$ tetrahedra and $\{\text{MnO}_6\}$ octahedra (Scheme 5a), while each $\{\text{MnO}_6\}$ octahedron is corner-shared with three $\{\text{PO}_3\text{C}\}$ tetrahedra from three equivalent $\text{Mn}^{\text{II}}(\text{notp})^{4-}$ units, thus leading to an inorganic layer containing six-member rings of Mn_3P_3 (Fig. 3a). A unique

Table 1 Metal–metalloligand phosphonates incorporating the Ru₂ paddlewheel

Compound	Ru ₂ core	Space group	Ru–Ru (Å)	N ^a	Structural feature	Magnetic property <i>D</i> (cm ⁻¹)/ <i>θ</i> (K)/ <i>z</i> <i>f</i> (cm ⁻¹)	Ref.
Na ₃ Ru ₂ (hedp) ₂ ·4H ₂ O (2)	Ru ₂ ⁵⁺	<i>C2/c</i>	2.354	3D	Open framework	96.0/0.64/n.a.	21
{La(H ₂ O) ₄ [Ru ₂ (hedp) ₂ (H ₂ O) ₂]}·5.5H ₂ O (3)	Ru ₂ ⁵⁺	<i>P2₁/n</i>	2.357	2D	Square-grid layer	101.6/n.a./–0.044	23
{Ce(H ₂ O) ₄ [Ru ₂ (hedp) ₂ (H ₂ O) ₂]}·5.5H ₂ O (4)	Ru ₂ ⁵⁺	<i>P2₁/n</i>	2.357	2D	Square-grid layer	101.6/n.a./–0.044	23
{La(H ₂ O) ₄ [Ru ₂ (hedp) ₂ (H ₂ O) ₂]}·8H ₂ O (5)	Ru ₂ ⁵⁺	<i>C2/c</i>	2.356	3D	PtS type open-framework	103.2/n.a./–0.023	23
{Ce(H ₂ O) ₄ [Ru ₂ (hedp) ₂ (H ₂ O) ₂]}·8H ₂ O (6)	Ru ₂ ⁵⁺	<i>C2/c</i>	2.356	3D	PtS type open-framework	n.a.	23
{Pr(H ₂ O) ₄ [Ru ₂ (hedp) ₂ (H ₂ O) ₂]}·8H ₂ O (7)	Ru ₂ ⁵⁺	<i>C2/c</i>	2.418	3D	PtS type open-framework	n.a.	23
{Nd(H ₂ O) ₄ [Ru ₂ (hedp) ₂ (H ₂ O) ₂]}·8H ₂ O (8)	Ru ₂ ⁵⁺	<i>C2/c</i>	2.351	3D	PtS type open-framework	n.a.	23
{Gd(H ₂ O) ₄ [Ru ₂ (hedp) ₂ (H ₂ O) ₂]}·8H ₂ O (11)	Ru ₂ ⁵⁺	<i>C2/c</i>	2.354	3D	PtS type open-framework	n.a.	23
{Tb(H ₂ O) ₄ [Ru ₂ (hedp) ₂ (H ₂ O) ₂]}·8H ₂ O (12)	Ru ₂ ⁵⁺	<i>C2/c</i>	2.346	3D	PtS type open-framework	n.a.	23
{Dy(H ₂ O) ₄ [Ru ₂ (hedp) ₂ (H ₂ O) ₂]}·8H ₂ O (13)	Ru ₂ ⁵⁺	<i>C2/c</i>	2.360	3D	PtS type open-framework	n.a.	23
{Ho(H ₂ O) ₄ [Ru ₂ (hedp) ₂ (H ₂ O) ₂]}·8H ₂ O (14)	Ru ₂ ⁵⁺	<i>C2/c</i>	2.353	3D	PtS type open-framework	n.a.	23
{Er(H ₂ O) ₄ [Ru ₂ (hedp) ₂ (H ₂ O) ₂]}·8H ₂ O (15)	Ru ₂ ⁵⁺	<i>C2/c</i>	2.342	3D	PtS type open-framework	n.a.	23
{Dy(H ₂ O) ₃ [Ru ₂ (Hamdp) ₂][Ru ₂ (Hamdp) ₂ (H ₂ O) ₂]}·12H ₂ O (17)	Ru ₂ ⁵⁺	<i>P1</i>	2.346– 2.358	2D	Pillared-bilayer	n.a./–16.2/n.a.	24
{Yb[Ru ₂ (Hamdp) ₂] ₂ [Ru ₂ (Hamdp) ₂ (H ₂ O) ₂]}·15H ₂ O (18)	Ru ₂ ⁵⁺	<i>P1</i>	2.354– 2.362	3D	Open framework	n.a./–30.0/n.a.	24

^a N: dimension of the structure, n.a.: not available.

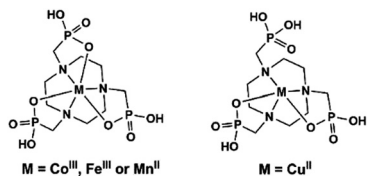
feature of the layer structure lies in the alignment of the hydrophobic triazacyclononane moieties on one side and the hydrophilic coordinated water molecules on the other side, which results in a polar layer. Two such polar layers are connected by hydrogen bond interactions between the coordination water molecules, leading to a supramolecular double layer structure which is centrosymmetric (Fig. 3b). Notably, direct reactions of uranyl or lanthanide ions with notpH₆ at low pH (*ca.* 1) resulted in UO₂(notpH₃)·H₂O²⁷ and a series of Ln-notp compounds²⁸ in which the macrocycle rings are partially protonated without the presence of M(notp) unit.

Although the overall structure of compound **19** is non-polar, the single layer polarity is retained and extended to the whole structure in compounds Zn₂Fe(notp)Cl(H₂O) (**20**) and [ZnCo(notpH)(H₂O)]·2H₂O (**21**),²⁹ where Fe^{III}(notp)^{3–} and Co^{III}(notpH)^{2–} are metalloligands. The two compounds were synthesized by a bottom-up approach using pre-formed Fe^{III}(notpH₃) or Co^{III}(notpH₃) monomers as precursors. **20** crystallizes in the hexagonal *P6₃* chiral space group which is also polar. Although the single crystal of **21** is chiral, the bulk sample is a conglomerate and is circular dichroism silent. Within the structure, the six phosphonate oxygen atoms in

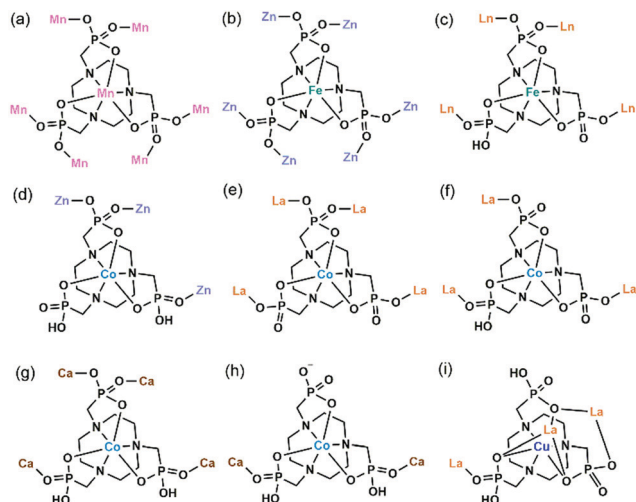
Table 2 Metal–metalloligand phosphonates incorporating polyazamacrocycles

Compound	Space group	N ^a	Structure	Proton conductivity (S cm ⁻¹) and other properties ^a	Ref.
[Mn ₃ (notp)(H ₂ O) ₆]·1.5H ₂ O (19)	<i>R3̄</i>	2D	Layer containing 6-rings	Weak AF coupling	27
[Zn ₂ Fe(notp)Cl(H ₂ O)] (20)	<i>P6₃</i>	2D	Layer containing 6-rings	4.3 × 10 ⁻⁷ , 95% RH, 25 °C	29
[ZnCo(notpH)(H ₂ O)]·2H ₂ O (21)	<i>Pca2₁</i>	2D	Layer containing 10-rings	1.0 × 10 ⁻⁶ , 95% RH, 25 °C, dielectric	29
[LaCo(notp)(H ₂ O) ₄]· <i>n</i> H ₂ O (22)	<i>P2₁/n</i>	2D	Layer containing 6-, 8-rings	3.0 × 10 ⁻⁶ , 95% RH, 25 °C, magnetic	30
[LaCo(notpH)(H ₂ O) ₆]ClO ₄ ·5H ₂ O (23)	<i>P2₁/n</i>	2D	Layer containing 12-rings	3.5 × 10 ⁻⁶ , 95% RH, 25 °C	31
[NdFe(notpH)(H ₂ O) ₄]ClO ₄ ·5H ₂ O (24)	<i>P2₁/n</i>	2D	Layer containing 4-, 8-rings	Paramagnetic	32
[SmFe(notpH)(H ₂ O) ₄]ClO ₄ ·5H ₂ O (25)	<i>P2₁/n</i>	2D	Layer containing 4-, 8-rings	Paramagnetic	32
[EuFe(notpH)(H ₂ O) ₄]ClO ₄ ·5H ₂ O (26)	<i>P2₁/n</i>	2D	Layer containing 4-, 8-rings	Paramagnetic	32
[GdFe(notpH)(H ₂ O) ₄]ClO ₄ ·5H ₂ O (27)	<i>P2₁/n</i>	2D	Layer containing 4-, 8-rings	Weak AF coupling	32
[CaCo(notpH ₂)(H ₂ O) ₂]ClO ₄ ·4H ₂ O (28)	<i>P2₁/n</i>	2D	Layer containing 4-, 8-rings	1.55 × 10 ⁻⁵ , 95% RH, 25 °C	33
[CaCo(notpH ₂)(H ₂ O) ₂]Cl·5H ₂ O (30)	<i>P2₁/c</i>	2D	Layer containing 4-, 8-rings	1.40 × 10 ⁻⁴ , 95% RH, 25 °C	34
[CaCo(notpH ₂)(H ₂ O) ₂]NO ₃ ·4H ₂ O (31)	<i>P2₁/n</i>	2D	Layer containing 4-, 8-rings	4.37 × 10 ⁻⁵ , 95% RH, 25 °C	34
[CaCo(notpH ₂)(H ₂ O) ₂]PF ₆ ·2H ₂ O (32)	<i>P2₁/n</i>	2D	Layer containing 4-, 8-rings	1.39 × 10 ⁻⁵ , 95% RH, 25 °C	34
[Ca ₂ Co ₂ (notpH ₂) ₂ (H ₂ O) ₈](IrCl ₆) ₂ ·2H ₂ O (33)	<i>P1</i>	1D	{CaO ₇ } dimer linked by {PO ₃ C}	1.27 × 10 ⁻⁶ , 95% RH, 25 °C	35
[CaCo ₂ (notpH ₂) ₂ (H ₂ O) ₂](IrCl ₆) ₂ ·4H ₂ O (34)	<i>C2/c</i>	1D	{CaO ₇ } linked by {PO ₃ C}	6.1 × 10 ⁻⁶ , 95% RH, 25 °C; magnetic	35
[Co ₂ (notpH ₄) ₂](IrCl ₆) ₂ ·6H ₂ O (35)	<i>P1</i>	0D	Discrete monomer	1.43 × 10 ⁻⁵ , 95% RH, 25 °C	35
[LaCu(notpH ₂)(H ₂ O) ₂]ClO ₄ ·3H ₂ O (36)	<i>P1</i>	1D	{LaO ₉ } linked by μ-O and {PO ₃ C}	Weak AF coupling	32
[GdCu(dotpH ₄)Cl]·4.5H ₂ O (37)	<i>P1</i>	3D	Open-framework with channels	Relaxometry	37
{[Na ₁₃ (OH) ₃ (H ₂ O) ₂₉]{Gd(dotp) ₂ }}·7H ₂ O·2CH ₃ CH ₂ OH (38)	<i>P4/n</i>	2D	Double layer structure	n.a.	38
[Mn ₂ Gd(dota-4ampH)(H ₂ O) ₇]·21H ₂ O (39)	<i>P2₁/c</i>	2D	Square-grid layer	Breathing effect	40
[Cu ₂ Cl ₂ (Me ₂ do2pH ₂)(H ₂ O)]·0.5C ₃ H ₆ O·2.5H ₂ O (41)	<i>P2₁/n</i>	2D	Rhombic layer	AF coupling	41
Li[Nd(do3apH)(H ₂ O)]·11.5H ₂ O (42)	<i>P1</i>	1D	Polar linear chain	n.a.	42
Li[Tb(do3apH)(H ₂ O)]·0.5HCl·5H ₂ O (43)	<i>Iba2</i>	1D	Zigzag chain	n.a.	43
[Cu(H ₂ O) ₂ {Cu(L ¹)}](ClO ₄)·2H ₂ O·0.5Me ₂ O (44)	<i>P2₁/c</i>	1D	Zigzag chain	n.a.	45
[Cd _{2.75} (L ²)(H ₂ O) ₇]·1.5NO ₃ ·7H ₂ O·MeOH (45)	<i>P2₁/c</i>	3D	Open-framework with channels	n.a.	46

^a N: dimension of the structure, n.a.: not available, AF: antiferromagnetic.



Scheme 4 Molecular structures of metalloligands $M^{\text{III}}(\text{notpH}_3)$ and $M^{\text{II}}(\text{notpH}_4)$.



Scheme 5 Coordination modes of $M(\text{notpH}_x)$ metalloligands.

$\text{Fe}(\text{notp})_3^-$ are each coordinated to a Zn atom (Scheme 5b), whereas each Zn atom is tetrahedrally coordinated by three phosphonate oxygen atoms from three equivalent $\text{Fe}(\text{notp})_3^-$ ligands and one Cl^- anion. Such arrangement results in a polar layer containing six-member rings of Zn_2FeP_3 with the organic components all residing on the same side of the layer. These polar layers are stacked unidirectionally in an ABAB manner, leading to a polar 3D network structure (Fig. 3c).

The polar structure is also observed in **21**. This compound crystallizes in an orthorhombic system, polar space group $Pca2_1$. The $\text{Co}(\text{notpH})^{2-}$ metalloligand is tri-dentate (Scheme 5d). Only three of its six phosphonate oxygen atoms are involved in coordination with three equivalent Zn atoms, while each Zn is tetrahedrally coordinated by three phosphonate oxygen atoms from three $\text{Co}(\text{notpH})^{2-}$ ligands and one water molecule. The resulted layer is undulated along the a -axis containing ten-member rings of $\text{Zn}_3\text{Co}_2\text{P}_5$. The layer is racemic due to the presence of both Δ and Λ configurations of the cobalt centre. It is polar, however, with the organic components aligning in the same direction along the c -axis. The layers are packed unidirectionally to form a polar 3D network (Fig. 3d). The polar nature is confirmed by the dielectric measurements on single crystals of **21**.

In contrast, the bottom-up assembly of $M^{\text{III}}(\text{notpH}_3)$ ($M = \text{Co}, \text{Fe}$) metalloligands with lanthanide or calcium ions resulted in products crystallizing in centrosymmetric space groups, possibly due to the fact that both Ln^{III} and Ca^{II} ions have no significant stereochemical influence on coordination. $[\text{LaCo}(\text{notp})(\text{H}_2\text{O})_4] \cdot n\text{H}_2\text{O}$ (**22**)³⁰ and $[\text{LaCo}(\text{notpH})(\text{H}_2\text{O})_6]\text{ClO}_4 \cdot$

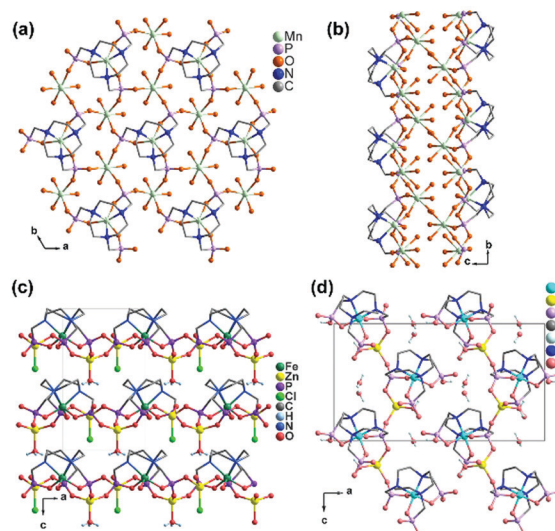


Fig. 3 (a) Single layer and (b) supramolecular double layer of structure **19**. (c) Packing diagram of structure **20**. (d) Packing diagram of structure **21**. All H atoms except those attached to oxygen in (d) are omitted for clarity. Adapted from ref. 27 and reproduced from ref. 29 with permission from the Royal Society of Chemistry.

$5\text{H}_2\text{O}$ (**23**)³¹ can be isolated by reacting $\text{Co}^{\text{III}}(\text{notpH}_3)$ and $\text{La}(\text{ClO}_4)_3$ at different pH, more acidic in the latter case. They crystallize in monoclinic space groups $P2_1/c$ and $P2_1/n$, respectively. In **22**, La^{III} is eight-coordinated by four phosphonate oxygen atoms and four water molecules. Each $\text{Co}(\text{notp})_3^{3-}$ acts as a tetradentate metalloligand and links to four La^{III} ions through four phosphonate oxygen atoms (Scheme 5e), while each La is connected to four $\text{Co}(\text{notp})_3^{3-}$ and an equivalent La atom through phosphonate linkages. This leads to a 2D waved layer containing 6-member rings of CoLa_2P_3 and 8-member rings of $\text{Co}_2\text{La}_2\text{P}_4$ (Fig. 4a). The interlayer space is filled with lattice water molecules.

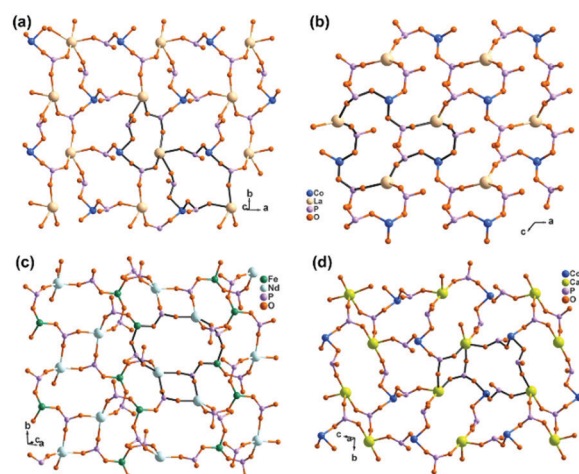


Fig. 4 Layer structures of compounds (a) $[\text{LaCo}(\text{notp})(\text{H}_2\text{O})_4] \cdot n\text{H}_2\text{O}$ (**22**), (b) $[\text{LaCo}(\text{notpH})(\text{H}_2\text{O})_6]\text{ClO}_4 \cdot 5\text{H}_2\text{O}$ (**23**), (c) $[\text{NdFe}(\text{notpH})(\text{H}_2\text{O})_4]\text{ClO}_4 \cdot 5\text{H}_2\text{O}$ (**24**), and (d) $[\text{CaCo}(\text{notpH}_2)(\text{H}_2\text{O})_2](\text{ClO}_4) \cdot 4\text{H}_2\text{O}$ (**28**). All C, N and H atoms as well as coordination water molecules are omitted for clarity. Adapted from ref. 30–33.

In **23**, La^{III} is nine-coordinated by three phosphonate oxygen atoms from three Co(notpH)²⁻, unlike that in **22**. Further, one phosphonate group in Co(notpH)²⁻ is protonated. Each Co(notpH)²⁻ serves as a tridentate metalloligand connecting to three La^{III} ions (Scheme 5f). As a result, a layer structure containing 12-member rings of Co₃La₃P₆ is constructed, which is completely different from that of **22** (Fig. 4b). The layer is positively charged with the ClO₄⁻ counter-anions and lattice water molecules residing in the interlayer space.

Interestingly, similar reactions using Fe^{III}(notpH₃) instead of Co^{III}(notpH₃) with lanthanide perchlorate resulted in layered compounds [LnFe(notpH)(H₂O)₄](ClO₄)·5H₂O [Ln = Nd (**24**), Sm (**25**), Eu (**26**), Gd (**27**)] with different structures.³² Compounds **24**–**27** are isostructural, crystallizing in the monoclinic *P*2₁/*n* space group. The coordination environment of the Ln^{III} ion is similar to that in **22**, which is eight-coordinated by four phosphonate oxygen atoms and four water molecules. While the Fe(notpH)²⁻ metalloligand is analogous to Co(notpH)²⁻ in **23** with one phosphonate group singly protonated, their coordination modes are different. Fe(notpH)²⁻ is tetra-dentate (Scheme 5c), whereas Co(notpH)²⁻ is tri-dentate. Consequently, a different type of layer structure containing 4- and 8-member rings is built up, as seen in Fig. 4c.

The replacement of La^{III} by Ca^{II} leads to a new layered compound [CaCo(notpH₂)(H₂O)₂](ClO₄)·4H₂O (**28**),³³ where the two phosphonate groups are each singly protonated in order to maintain the overall charge balance. The Ca^{II} ion prefers an octahedral coordination geometry with the six sites provided by four phosphonate oxygen atoms and two water molecules. Each Ca is connected to four [Co(notpH₂)]⁻ and *vice versa*, leading to a 2D layer with similar topology to that in **24**–**27** (Fig. 4d and Scheme 5g). Humidity dependent single crystal to single crystal structural transformation is observed for **28**, forming a new phase [CaCo(notpH₂)(H₂O)₂](ClO₄)·2H₂O (**29**) at low atmosphere humidity, accompanied by a glide of the layer relative to the adjacent layers and a re-arrangement of the hydrogen bond network. The ClO₄⁻ in **28** can be replaced by other counter-anions forming [CaCo(notpH₂)(H₂O)₂](X)_n·*n*H₂O [X = Cl⁻ (**30**, *n* = 5), NO₃⁻ (**31**, *n* = 4), PF₆⁻ (**32**, *n* = 2)] with the same layer topology but different number of lattice water molecules.³⁴

The counter-anion can be paramagnetic. By incorporating IrCl₆²⁻, trimetallic Co^{III}–Ca^{II}–Ir^{IV} compounds [Ca₂Co₂(notpH₂)₂(H₂O)₈](IrCl₆)·2H₂O (**33**) and [CaCo₂(notpH₃)₂(H₂O)₂](IrCl₆)·4H₂O (**34**) as well as [Co₂(notpH₄)₂](IrCl₆)·6H₂O (**35**) were isolated at pH 3.5, 2.5 and 1.7, respectively.³⁵ In **33**, edge-sharing dimers of {CaO₇} are connected by the bidentate [Co(notpH₂)]⁻ metalloligands (Scheme 5h) into infinite chains. **34** also shows a chain structure, made up of {CaO₆} octahedra and Co(notpH₃) linkages. In the case of **35**, discrete Co(notpH₄)⁺ species is found together with the IrCl₆²⁻ anion.

Despite the dominant octahedral geometries of the metal centres in M(notpH_{*x*}) (M = Co^{III}, Fe^{III}, Mn^{II}), the Cu^{II} ion favours a square or square pyramidal environment due to the Jahn–Teller distortion. In [LaCu(notpH₂)(H₂O)₂](ClO₄)·3H₂O (**36**),³² the Cu^{II} ion in Cu(notpH₂)²⁻ has a distorted square pyramidal geometry

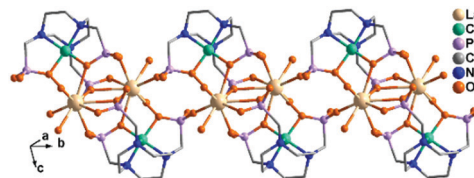


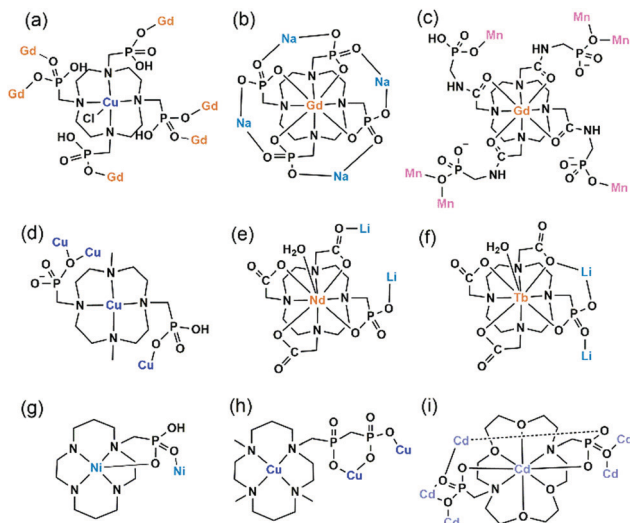
Fig. 5 Chain structure of **36**. All H atoms are omitted for clarity. Adapted from ref. 32.

with the basal sites occupied by two O and two N donors from the same notpH₂⁴⁻, and axial site by the third nitrogen atom. Cu(notpH₂)²⁻ serves as a hepta-dentate metalloligand binding to three equivalent La atoms (Scheme 5i). Each La atom is nine-coordinated by seven phosphonate oxygen and two water molecules. Neighbouring La atoms are bridged by Cu(notpH₂)²⁻ metalloligands through μ₃-O and O–P–O units alternatively, forming an infinite chain (Fig. 5).

Other polyazamacrocyclic phosphonate ligands that have been commonly studied are 1,4,7,10-tetraazacyclododecane (cyclen) and 1,4,8,11-tetraazacyclotetradecane (cyclam) with one-, two-, three- or four-pendant phosphonate arms. However, most studies were limited to their coordination chemistry in solution due to their utilization in medicine and molecular biology as contrast agents in magnetic resonance imaging (MRI) or metal radioisotope carriers for diagnostic and/or therapeutic purposes.³⁶ Coordination polymers incorporating cyclen- or cyclam-phosphonate metalloligands are rare.

An interesting example is [GdCu(dotpH₄)Cl]·4.5H₂O (**37**)³⁷ reported by Maspoch and co-workers, where dotpH₈ represents 1,4,7,10-tetraazacyclododecane-1,4,7,10-tetramethylenephosphonic acid. The compound was synthesized by direct reaction of CuCl₂, Gd(NO₃)₃ and dotpH₈ in aqueous solution at 85 °C. Within the structure, there are two crystallographically distinguished Cu atoms (Cu1, Cu2), each having a square pyramidal geometry. The basal sites are provided by four N atoms from the dotpH₄⁴⁻ ligand and the axial site by the Cl⁻ anion. Each Cu(dotpH₄)²⁻ serves as a hexa-dentate metalloligand binding to six Gd atoms (Scheme 6a). There are also two kinds of Gd atoms, Gd1 and Gd2, each having a distorted octahedral coordination environment. The equivalents of each Gd are doubly bridged by O–P–O units forming Gd₁₂ and Gd₂₂ dimers. These dimers are further connected by O–P–O units alternatively forming infinite chains. The Gd chains are cross-linked by Cu(dotpH₄)²⁻ to construct a 3D open framework structure with one dimensional channels (*ca.* 5 × 5 Å, van der Waals radii not accounted) (Fig. 6a and b). Notably, the same structure cannot be obtained by replacing Cu(II) by Zn(II), Ni(II), Co(II), Mn(II), Fe(III) and Mg(II), likely because they have less preference for nitrogen donors compared to Cu(II).

Based on the same dotp ligand, Vecilla and co-workers reported a Na–Gd compound, [{Na₁₃(OH)₃(H₂O)₂₉}{Gd(dotp)}₂]·7H₂O·2CH₃CH₂OH (**38**).³⁸ As shown in Fig. 6c, the Gd(dotp) metalloligands in **38** are dispersed in the {NaO_{*x*}} polyhedral network by coordinating to four Na atoms through their eight pendant phosphonate oxygen atoms forming a layer (Scheme 6b). Two such layers are connected by {NaO_{*x*}} polyhedra to form a



Scheme 6 Coordination modes of metalloligands based on cyclen-, cyclam- and 1,10-diaza-18-crown-6-phosphonate.

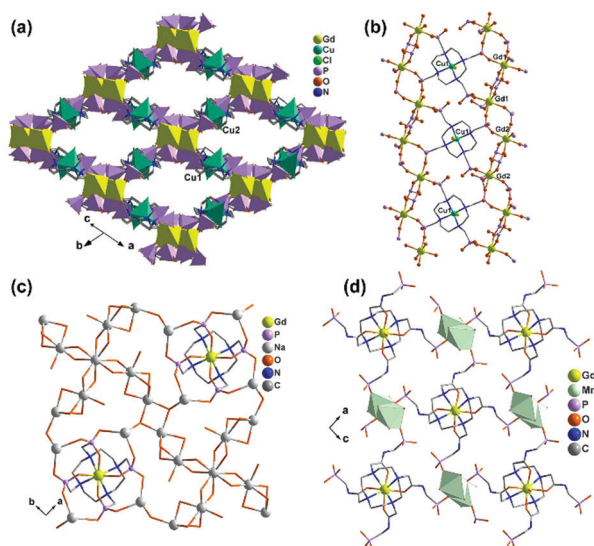


Fig. 6 (a) The 3D open framework structure and (b) the Gd chain connection of **37**. (c) Single layer in structure **38**. (d) Layer structure made up of $\text{Mn}_2(\text{PO}_3)_4(\text{H}_2\text{O})_6$ dimers and $\text{Gd}(\text{DOTA-4ampH})(\text{H}_2\text{O})_7 \cdot 2.1\text{H}_2\text{O}$ (**39**). Adapted from ref. 37, 38 and 40.

double layer structure. If notpH_8 reacted with lanthanide or transition metal salts at low pH, discrete or polymeric compounds can be isolated in which the macrocycle nitrogen atoms are partially protonated and not involved in coordination with metal ions.³⁹

1,4,7,10-Tetraazacyclododecane-1,4,7,10-tetraacetamido methylenephosphonic acid (DOTA-4ampH_8) has four phosphonic acid arms with an additional acetamido moiety. It is ideally suited to wrap up the lanthanide ion using four N and four acetoamido O atoms to form a metalloligand, leaving the four phosphonate groups for further coordination. Maspoeh *et al.* successfully isolated a two-dimensional Mn–Gd compound $[\text{Mn}_2\text{Gd}(\text{DOTA-4ampH})(\text{H}_2\text{O})_7] \cdot 2.1\text{H}_2\text{O}$ (**39**),⁴⁰ using the pre-formed $\text{Gd}(\text{DOTA-4ampH}_x)$ metalloligand as the precursor.

It crystallizes in monoclinic space group $P2_1/c$. The Gd(III) ion is nine-coordinated by four N, four acetoamido O atoms and one water molecule. The metalloligand $\text{Gd}(\text{DOTA-4ampH})$ is hexa-dentate, using four phosphonate oxygen atoms, each from one phosphonate group (Scheme 6c). Two of the four phosphonate oxygen atoms act as $\mu_3\text{-O}$ bridging two Mn atoms, thus forming $\text{Mn}_2(\text{PO}_3)_4(\text{H}_2\text{O})_6$ dimers. The dimers are connected by the metalloligands in four directions, leading to a layer structure in the *ac* plane (Fig. 6d). The layers are stacked along the *c*-axis generating 1D channels with dimensions of approximately $14.8 \times 9.5 \text{ \AA}$ (van der Waals radii not accounted) where the lattice water molecules reside. Compound **39** shows interesting humidity dependent single-crystal to single-crystal structural transformation, forming $[\text{Mn}_2\text{Gd}(\text{DOTA-4ampH})(\text{H}_2\text{O})_7] \cdot 5\text{H}_2\text{O}$ (**40**) at RH < 40%. The transformation causes the layers changing from corrugated to flatter, but the coordination network remains almost the same. Water sorption isotherms show a gate-opening effect at a RH of *ca.* 85% upon water adsorption, and a gate-closure effect at RH 55–77% upon water desorption.

Vizza and co-workers synthesized a cyclen derivative with two phosphonic acid arms, *e.g.* 4,10-dimethyl-1,4,7,10-tetraazacyclododecane-1,7-bis(methanephosphonic acid) ($\text{Me}_2\text{do2pH}_4$). The reaction of $\text{Me}_2\text{do2pH}_4$ with $\text{Cu}(\text{OAc})_2$ at room temperature results in the formation of compound $\{[\text{Cu}_2\text{Cl}_2(\text{Me}_2\text{do2pH}_2)(\text{H}_2\text{O})] \cdot 0.5\text{C}_3\text{H}_6\text{O} \cdot 2.5\text{H}_2\text{O}\}$ (**41**).⁴¹ Analogous to that in **37**, the Cu(II) ion also has a distorted square-pyramidal coordination geometry with the four basal positions occupied by macrocycle N atoms and the axial position by Cl^- anion. However, the $\text{Cu}(\text{Me}_2\text{do2pH}_2)$ metalloligand is tri-dentate (Scheme 6d). Its two phosphonate arms each provide one phosphonate oxygen to bind three Cu(II) ions. One phosphonate oxygen serves as $\mu_3\text{-O}$ bridging two equivalent Cu atoms to form the $\{\text{Cu}_2\text{O}_2\}$ dimer. The dimers are connected by $\text{Cu}(\text{Me}_2\text{do2pH}_2)$ in four directions forming a 2D layer containing rhombic rings made up of four metalloligands, four Cu atoms from $\{\text{Cu}_2\text{O}_2\}$ dimers, and eight $\{\text{PO}_3\text{C}\}$ groups (Fig. 7a). The layers are stacked forming a 3D supramolecular network, generating 1D channels along the *a*-axis. The channel size is *ca.* $5.2 \times 5.8 \text{ \AA}$ (van der Waals radii not accounted).

By employing a mono(methylphosphonate)-cyclen derivative, 1,4,7,10-tetraazacyclododecane-4,7,10-tris(carboxymethyl)-1-methylphosphonic acid ($\text{H}_5\text{do3aP}$), Lukeš and co-workers isolated two Li–Ln compounds with chain structures, *e.g.* $\text{Li}[\text{Nd}(\text{Hdo3aP})(\text{H}_2\text{O})] \cdot 11.5\text{H}_2\text{O}$ (**42**)⁴² and $\text{Li}[\text{Tb}(\text{Hdo3aP})(\text{H}_2\text{O})] \cdot 0.5\text{HCl} \cdot 5\text{H}_2\text{O}$ (**43**).⁴³ In both cases, the Ln(III) ions are nine-coordinated by four macrocycle N atoms, three carboxylate O atoms and one phosphonate O atom as well as one coordination water molecule. However, the metalloligand $\text{Nd}(\text{Hdo3aP})^-$ in **42** is bi-dentate linking to two Li atoms *via* one carboxylate O atom and one phosphonate O atom (Scheme 6e). But the $\text{Tb}(\text{Hdo3aP})^-$ in **43** is tri-dentate, chelating and bridging to two equivalent Li atoms (Scheme 6f). As a consequence, the tetrahedrally coordinated Li(I) ions in **42** are connected by $\text{Nd}(\text{Hdo3aP})^-$ into a linear chain which is polar (Fig. 7b). On the other hand, the Li(I) ions in **43** are connected by $\text{Tb}(\text{Hdo3aP})^-$ into a zigzag chain *via* two phosphonate oxygen

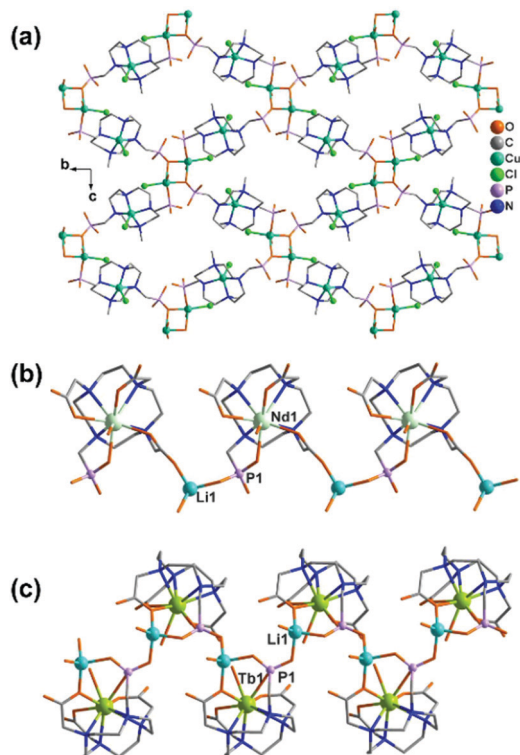


Fig. 7 (a) Single layer of structure **41**. (b) Chain structure of **42**. (c) Chain structure of **55**.

atoms which is non-polar (Fig. 7c). It should be mentioned that the overall structure of **42** is non-polar due to the centrosymmetric packing of the polar chains.

As far as we are aware, there are only two coordination polymers containing cyclam-phosphonate metalloligands. One is Ni(cyclam1p),⁴⁴ where cyclam-p represents [(1,4,8,11-tetraazacyclotetradecan-1-yl)methylene]phosphonic acid. Within the structure, Ni(cyclam1p) serves as both donor and acceptor using one phosphonate oxygen to bind the Ni atom from the neighbouring unit (Scheme 6g), thus forming an infinite chain. The other is $\{Cu(H_2O)_2[Cu(L^1)]\}(ClO_4) \cdot 2H_2O \cdot 0.5acetone$ (**44**),⁴⁵ isolated by Kotek and co-workers, where the cyclam derivative H_3L^1 contains a single diphosphonate arm. In this case, $Cu(L^1)$ serves as a tri-dentate metalloligand (Scheme 6h), chelating and bridging the Cu(II) ions into infinite zigzag chains. The chain consists of corner sharing $\{CuO_5\}$ square pyramids and $\{PO_3C\}$ tetrahedra. The cyclam moieties are pendant on the two sides of the inorganic chain (Fig. 8a).

Clearfield and co-workers employed N,N' -bis(phosphonomethyl)-1,10-diaza-18-crown-6 (H_4L^2), and isolated compound $Cd_{2.75}(L^2)(H_2O)_7 \cdot 1.5NO_3 \cdot 7H_2O \cdot MeOH$ (**45**) with an open framework structure.⁴⁶ Within the metalloligand $Cd(L^2)^{2-}$, the Cd atom is eight-coordinated by four O and two N atoms from the crown ring and two O atoms from the phosphonate arms. The metalloligand is hepta-dentate, chelating and bridging five Cd atoms (Scheme 6i). A pentanuclear Cd_5 cluster is found, where the central $\{CdO_8\}$ is edge-shared with four $\{CdO_6\}$ octahedra. Each Cd_5 cluster is connected to four $Cd(L^2)^{2-}$, forming a $Cd_5\{Cd_4(L^2)_4\}$ cluster

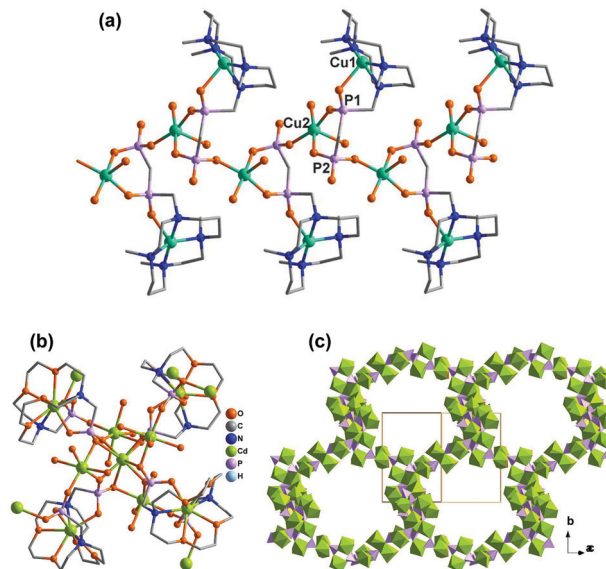


Fig. 8 (a) Chain structure of **44**. (b) The $Cd_5\{Cd_4(L^2)_4\}$ cluster and (c) packing diagram of structure **45**. The C, N and H atoms in (c) are omitted for clarity. Colour codes: $\{CdO_8\}$ and $\{CdO_6\}$ green, $\{PO_3C\}$ purple.

(Fig. 8b). The $Cd_5\{Cd_4(L^2)_4\}$ clusters are cross-linked *via* $\{Cd_2O_2\}$ dimers, forming a 3D framework structure (Fig. 8c). The channel sizes are *ca.* $4.2 \times 4.2 \text{ \AA}$ along the *c* axis or $9.3 \times 14.1 \text{ \AA}$ towards $(0\ 5\ -4)$ (van der Waals radii not accounted), respectively.

To summarize, metallo-polyazamacrocycle-phosphonates can be used as efficient metalloligands for the construction of MMPs with structures ranging from 1D chain, 2D layer to 3D framework. For metallo-polyazamacrocycle with four, three, or two phosphonate arms, they usually form 2D layer structures with second metal ions. Only in a few cases, 1D chain or 3D open framework structures are observed. In contrast, metallo-polyazamacrocycle with one phosphonate arm tend to form a 1D chain structure. This statement is not rigorous, however, as the given examples are still extremely limited. The formation of a particular structure is highly dependent on the reaction conditions, the coordination geometries of the second metal ions and the involvement of metal clusters as nodes. In addition, these metalloligands possess a unique feature of polarity that distinguishes them from many other phosphonate ligands. This feature together with their structural diversity and tuneable composition endows them with advantages in the design and synthesis of materials with multifunction.

3.3. Metalloporphyrin units as metalloligands

Porphyrins and their complexes have attracted intensive research interest across many fields due to their rich coordination chemistry and photophysical, redox and catalytic properties.⁴⁷ In recent years a number of metal-organic frameworks (MOFs) have been obtained using metalloporphyrins as building units.⁴⁸ Especially Zr-MOFs containing H_2TCPP^{4-} are highly porous and stable.⁴⁹ However, it is still a great challenge to construct MMPs incorporating metalloporphyrins, especially those based on metalloporphyrin-phosphonic acids, due to the difficulties in synthesizing

Table 3 Metal–metalloid phosphonates incorporating metalloporphyrins, cyclometalate or metal–polypyridine

Compound	Space group	N^a	Structure	Proton conductivity (S cm ⁻¹) and other properties ^a	Ref.
Cu ₂ (^t BuCOO) ₄ {Cu(DEPPP)} (46)	$P\bar{1}$	1D	Linear chain	n.a.	54
Cu ₂ (^t BuCOO) ₄ {Pd(DEPPP)} (47)	$P\bar{1}$	1D	Linear chain	n.a.	54
Cu ₂ (^t BuCOO) ₄ {Ni(DEPPP)} (48)	$C2/c$	1D	Zigzag chain	n.a.	54
MnNi(TPPPH ₆)(H ₂ O) (49)	$P\bar{1}$	3D	Framework isostructural to 51	N ₂ and H ₂ O sorption	60
CoNi(TPPPH ₆)(H ₂ O) (50)	$P\bar{1}$	3D	Framework isostructural to 51	N ₂ and H ₂ O sorption	60
NiNi(TPPPH ₆)(H ₂ O) (51)	$P\bar{1}$	3D	Open-framework with channels	5.6 × 10 ⁻⁶ , 90% RH, 80 °C	60
CdNi(TPPPH ₆)(H ₂ O) (52)	$P\bar{1}$	3D	Framework isostructural to 51	N ₂ and H ₂ O sorption	60
Zn ₃ [Ni(TPPPH ₃) ₃] ₃ ·9(CH ₃) ₂ NH ₂ ·3DMF·17H ₂ O (53)	$C2/c$	3D	Open-framework with channels	1.55 × 10 ⁻³ , 80% RH, 75 °C	61
[Co ₂ Ni(TPPPH ₄)] ₂ DABCO·6H ₂ O (54)	$P\bar{4}c2$	3D	Open-framework with channels	N ₂ sorption, BET: 700 m ² g ⁻¹	62
[Zr ₂ Ni(TPPPH ₂)(OH/F) ₂] (55)	$I4_1/acd$	3D	Open-framework with channels	BET: 1070 m ² g ⁻¹ , CV, UV-vis	63
[Hf ₂ Ni(TPPPH ₂)(OH/F) ₂] (56)	$I4_1/acd$	3D	Open-framework with channels	BET: 1030 m ² g ⁻¹ , UV-vis	63
[Na ₂ Cu(TPPPH ₄)]·(NH ₂ (CH ₃) ₂) ₂ (57)	$P\bar{1}$	3D	Open-framework with channels	BET: 697 m ² g ⁻¹	64
[EuIr ₆ (ppy) ₁₂ (bpp) ₂ (bph) ₄](CF ₃ SO ₃) _x ·xH ₂ O (58)	$P\bar{1}$	0D	Supramolecular framework	Phosphorescence	65
[TbIr ₆ (ppy) ₁₂ (bpp) ₂ (bph) ₄](CF ₃ SO ₃) _x ·xH ₂ O (59)	$P\bar{1}$	0D	Supramolecular framework	Phosphorescence	65
[DyIr ₆ (ppy) ₁₂ (bpp) ₂ (bph) ₄](CF ₃ SO ₃) _x ·xH ₂ O (60)	$P\bar{1}$	0D	Supramolecular framework	Phosphorescence, magnetic	65
[ErIr ₆ (ppy) ₁₂ (bpp) ₂ (bph) ₄](CF ₃ SO ₃) _x ·xH ₂ O (61)	$P\bar{1}$	0D	Isostructural to 58–60	Phosphorescence, magnetic	65
[YbIr ₆ (ppy) ₁₂ (bpp) ₂ (bph) ₄](CF ₃ SO ₃) _x ·xH ₂ O (62)	$P\bar{1}$	0D	Isostructural to 58–60	Phosphorescence, magnetic	65
[La ₃ (H ₂ O) ₉ {Ru(4,4'-dpppy) ₃ H _{5.5} } ₂] ₃ ·nH ₂ O (63)	$P2_1/n$	3D	Isostructural to 64	2.8 × 10 ⁻⁶ , 95% RH, 25 °C	66
Pr ₃ [Ru(DBBPY) ₃ H _{5.5}] (64)	$P2_1/n$	3D	Open-framework with channels	3.7 × 10 ⁻⁶ , 95% RH, 25 °C	66
[GdCo ^{III} (5,5'-dpppyH ₂) ₃ (H ₂ O) ₂] ₂ ·6.5H ₂ O (65)	$C2/c$	2D	Isostructural to 67	Magnetic	67
[DyCo ^{III} (5,5'-dpppyH ₂) ₃ (H ₂ O) ₂] ₂ ·6.5H ₂ O (66)	$C2/c$	2D	Isostructural to 67	Magnetic	67
[TbCo ^{III} (5,5'-dpppyH ₂) ₃ (H ₂ O) ₂] ₂ ·6.5H ₂ O (67)	$C2/c$	2D	Layer with (4,4) connectivity	Magnetic	67
[Gd{Ni ^{II} (5,5'-dpppy) ₃ H ₇ }(H ₂ O) ₃ NaCl·6H ₂ O] (68)	$C2$	3D	Open-framework with channels	Magnetic	67

^a N : dimension of the structure, n.a.: not available. UV-vis absorption spectra and cyclic voltammetry (CV) measurements were carried out in solution.

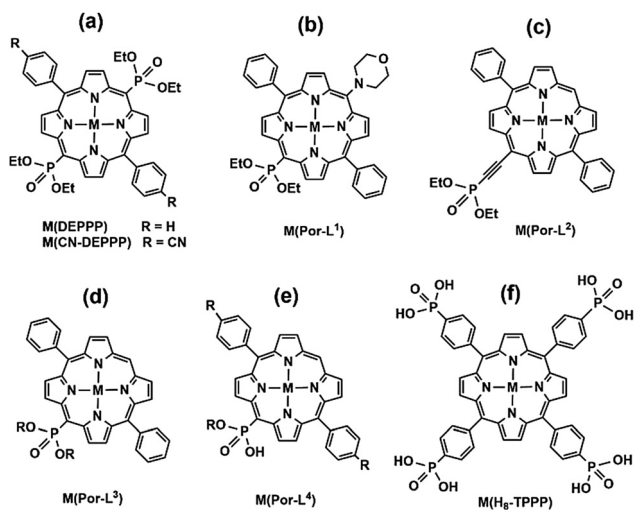
the metalloligands and in forming crystalline materials with sufficient size for structural determination. Table 3 gives the known MMPs based on the metalloporphyrins shown in Scheme 7.

Gorbunova and co-workers synthesized a meso-phosphorylporphyrin, e.g. 5,15-bis(diethoxyphosphoryl)-10,20-diphenylporphyrin (DEPPP, Scheme 7a), through Pd-catalyzed coupling reactions. The solution reactions of H₂-DEPPP with metal salts resulted in compounds M(DEPPP) (M = Zn,⁵⁰ Cu,⁵¹ Cd⁵²) with similar layer structures, where DEPPP serves as both donor and acceptor. For M(DEPPP) (M = Pt, Pd, Ni), layered coordination polymers cannot be obtained because the central

metal ions favour a four-coordinated square-planar geometry.⁵² Another related 2D polymeric MMP is Zn(CN-DEPPP) based on 5,15-bis(diethoxyphosphoryl)-10,20-bis(4'-cyanophenyl)porphyrin (CN-DEPPP, Scheme 7a), described by Kadish and co-workers.⁵³ The electrochemical properties were studied for all these compounds.

Direct reactions of Cu₂(^tBuCOO)₄(CH₃CN)₂ and M(DEPPP) (M = Cu, Pd, Ni) in chloroform resulted in the formation of compounds Cu₂(^tBuCOO)₄{M(DEPPP)} [M = Cu (46), Pd (47), Ni (48)].⁵⁴ These compounds display similar zigzag chain structures, where the paddlewheel dimers of Cu₂(^tBuCOO)₄ are connected by the M(DEPPP) metalloligands (Fig. 9a). In 46, there exist two kinds of Cu atoms. One locates at the centre of the porphyrin cavities adopting a square-planar CuN₄ environment. The other incorporates in the paddlewheel unit showing a square-pyramidal CuO₅ geometry with the apical site occupied by phosphoryl O atom. It is noted that the geometry of the porphyrin macrocycle can be rather different depending on the central metal ion. A strong nonplanarity of the porphyrin core is observed only for Ni(DEPPP). This leads to the formation of a zigzag chain in 48, instead of a linear chain as in 46 and 47. The packing of the zigzag chains results in regular pores in the crystal with dimensions of ca. 10 × 12 Å (van der Waals radii not accounted) (Fig. 9b).

Recently, Ermakova *et al.* developed synthetic routes to accessing mono(diethoxyphosphorylporphyrins) with different substituents. Compounds Zn(Por-L¹)⁵⁵ (Scheme 7b) and Zn(Por-L²)⁵⁶ (Scheme 7c) show chain structures, where the porphyrin macrocycle units are connected to each other through coordination of zinc with phosphoryl O atoms. Apart from these polymers, discrete dimeric or tetrameric compounds were also reported



Scheme 7 Molecular structures of metalloporphyrin–phosphonate ligands.

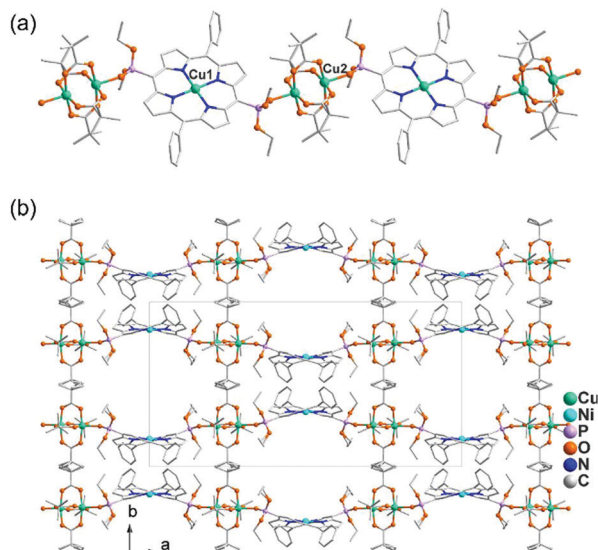


Fig. 9 (a) Linear chain structure of **46**. (b) Zigzag chain structure of **48**. Adapted from ref. 50 and 54.

based on zinc complexes of 10-(dialkoxyphosphoryl)-5,15-diarylporphyrinates or β -dialkoxyphosphoryl-5,10,15,20-tetraphenylporphyrins (Por-L³, Scheme 7d),⁵⁷ and gallium(III) and indium(III) complexes of *meso*-mono(diethoxyphosphoryl)porphyrins (Por-L⁴, Scheme 7e).⁵⁸

Noting that the diethoxyphosphoryl group can provide only one oxygen donor for coordination and the resulted M–O(P) bond is weak, it is highly desired to prepare metal phosphonates using porphyrin-phosphonic acids as ligands. A typical ligand is tetra(4-phosphonophenyl)porphyrin (H₁₀TPPP). Although it was known early in 1992,⁵⁹ coordination polymers containing H_xTPPP-based metalloligands were reported only recently.

Stock and co-workers employed a high-throughput method and synthesized a series of metal–organic frameworks using Ni(H_xTPPP) metalloporphyrins (Scheme 7f) as the precursor. Compounds M(Ni-H₆TPPP)(H₂O) [M = Mn (**49**), Co (**50**), Ni (**51**), Cd (**52**)]⁶⁰ are isostructural, and the structure of Ni(Ni-H₆TPPP)(H₂O)·11H₂O (**51**) is shown in Fig. 10. Ni-H₆TPPP²⁻ acts as a pentadentate metalloligand, binding five equivalent Ni1 atoms *via* four phosphonate oxygen atoms, each from a different phosphonate group (Fig. 10a). The phosphonate oxygen O1 serves as μ_3 -O bridging two Ni1 atoms forming an edge-sharing {Ni₁₂O₂} dimer (Fig. 10b). Each dimer is connected to eight Ni-H₆TPPP²⁻ metalloligands and each Ni-H₆TPPP²⁻ links to four {Ni₁₂O₂} dimers, therefore constructing a 3D porous structure containing two kinds of rhombic channels with dimensions 2.8 × 5.0 and 3.8 × 3.8 Å (van der Waals radii accounted) or 5.5 × 8.0 and 7.0 × 7.0 Å (van der Waals radii not accounted) that host water molecules (Fig. 10c). Sorption measurements revealed that **51** shows the highest water uptake of 181 mg g⁻¹. Moreover, it is stable in the pH range between 1 and 11.

A dinuclear building unit is also found in compound Zn₃(Ni-H₃TPPP)₃·9(CH₃)₂NH₂·3DMF·17H₂O (**53**).⁶¹ Here the metalloligand Ni-H₃TPPP⁵⁻ acts as a tetra-dentate ligand, binding to four Zn atoms (Fig. 10d). The Zn atoms have tetrahedral geometries, unlike the octahedral geometry of the Ni(II) ion in **51**. The Zn atoms are doubly bridged by O–P–O units forming {Zn₂(OPO)₂} dimers (Fig. 10e). Each dimer is connected to six Ni-H₃TPPP⁵⁻, while each Ni-H₃TPPP⁵⁻ is connected to three {Zn₂(OPO)₂} dimers. Although the connecting topology in **53** is quite different from that in **51**, a 3D open framework containing rhombic channels is again constructed in **51** (Fig. 10f). The channel size is *ca.* 4.5 × 10 Å (van der Waals radii not accounted). The anionic framework is charge-balanced by (CH₃)₂NH₂⁺ which, together with the solvent molecules, occupies the channel space.



Fig. 10 (a–c) The structure of compound **51**: (a) coordination mode of Ni-H₆TPPP²⁻, (b) dimeric building unit, (c) packing diagram. (d–f) The structure of compound **53**: (d) coordination mode of Ni-H₆TPPP²⁻, (e) dimeric building unit, (f) packing diagram. Adapted from ref. 60 and 61.

Another example containing the Ni-H_xTPPP metalloligand is [Co₂(Ni-H₄TPPP)]·2DABCO·6H₂O (**54** or Co-CAU-36) (DABCO = 1,4-diazabicyclo[2.2.2]octane),⁶² the structure of which was determined by single-crystal electron diffraction using continuous rotation electron diffraction (cRED) data collected at low temperature (96 K). In this compound, Ni-H₄TPPP⁴⁻ is octadentate, coordinating to eight Co atoms using eight of its twelve phosphonate oxygen atoms (Fig. 11a). The Co atom has a distorted tetrahedral geometry. The equivalent Co atoms are doubly bridged by phosphonate groups through corner-sharing of {CoO₄} and {PO₃C} forming an inorganic chain (Fig. 11b). The chains are cross-linked by the Ni-H₄TPPP⁴⁻ moieties to form a 3D open framework with 1D rectangular channels generated along the *c*-axis (Fig. 11c). The channel has a pore diameter of *ca.* 9 Å (van der Waals radii accounted) when taking the DABCO guest molecules into account, and a diameter of *ca.* 12.5 Å (van der Waals radii accounted) when the pores are empty.

Interestingly, a similar framework structure with rectangular channels is also observed for tetravalent metal compounds [M₂(Ni-H₂TPPP)(OH/F)₂·H₂O [M = Zr (**55**), Hf (**56**); (M-CAU-30)].⁶³ **55** and **56** are isostructural. Like that in **54**, the metalloligand Ni-H₂TPPP⁶⁻ acts as an octa-dentate ligand, binding to eight Zr^{IV} (or Hf^{IV}) ions (Fig. 11d). A significant difference between **55** and **54** is the coordination geometry of second metal ions, which is tetrahedral for Co(II) in **54** but octahedral for Zr(IV) in **55**. The adjacent Zr atoms are bridged not only by two phosphonate groups but also by OH⁻/F⁻ anions, hence leading to an infinite chain containing corner-sharing {ZrO₆} octahedra (Fig. 11e). The chains are cross-linked by the Ni-H₂TPPP⁶⁻ ligands to form a 3D framework (Fig. 11f). The channel diameters are 1.3 × 2.0 nm (van der Waals radii accounted). M-CAU-30 show the highest specific surface area

among metal phosphonates of 1070 and 1030 m² g⁻¹ for **55** and **56**, respectively, derived from the experimental data for the corresponding compounds contaminated with ZrO₂ or HfO₂ impurities. They are chemically stable in the pH range 0–12 and thermally stable up to 420 °C in air. The remarkable stability of M-CAU-30 makes them promising candidates for many applications such as gas separation and storage, catalysis and proton conduction.

Yücesan, Zorlu and co-workers reported an alkali phosphonate MOF containing the Cu-H₄TPPP⁴⁻ metalloligand, formulated as [Na₂(Cu-H₄TPPA)]·(NH₂(CH₃)₂)₂ (**57**).⁶⁴ **57** crystallizes in a triclinic space group *P* $\bar{1}$. It also shows a 3D open framework structure with rectangular shaped channels (7.6 × 7.6 Å, van der Waals radii not accounted), composed of Na–O–P–O inorganic chains and Cu-H₄TPPP⁴⁻ linkages. Each Cu-H₄TPPP⁴⁻ binds to eight Na^I ions. The Na^I ions are either four-coordinated (for Na1) or five-coordinated (for Na2), each of which is edge-shared with its equivalent ones forming {Na₁O₂} or {Na₂O₂} dimers. The dimers are connected *via* phosphonate oxygen and O–P–O units in an alternative manner, forming an infinite chain. These chains are cross-linked by Cu-H₄TPPP⁴⁻ metalloligands into a 3D framework (Fig. 12). After activation, the compound shows a surface area of 698 m² g⁻¹ which is dramatically larger than that of the other Na-MOFs, but close to that of [Co₂(Ni-H₄TPPP)]·2DABCO·6H₂O (700 m² g⁻¹).

To summarize, metalloporphyrins with one or two esterified phosphonate groups form 1D chains with second metal ions, while those with four phosphonic acid groups form 3D open framework structures with different channel sizes. These MMPs show very high thermal and chemical stability and a large surface area, especially those containing tetravalent Zr^{IV} and Hf^{IV}

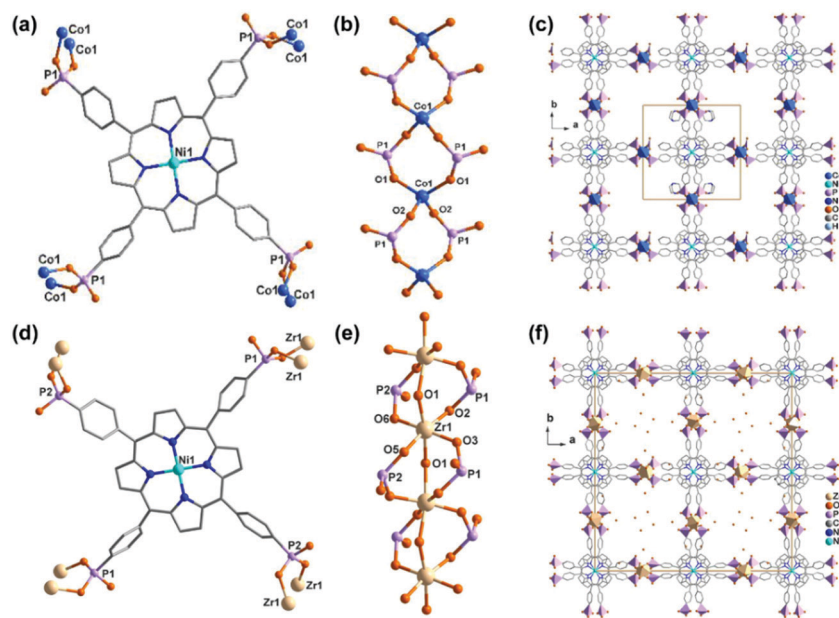


Fig. 11 (a–c) The structure of compound **54**: (a) coordination mode of Ni-H₄TPPP²⁻, (b) dimeric building unit, (c) packing diagram. (d–f) The structure of compound **55**: (d) coordination mode of Ni-H₆TPPP²⁻, (e) dimeric building unit, (f) packing diagram. Adapted from ref. 62 and 63.

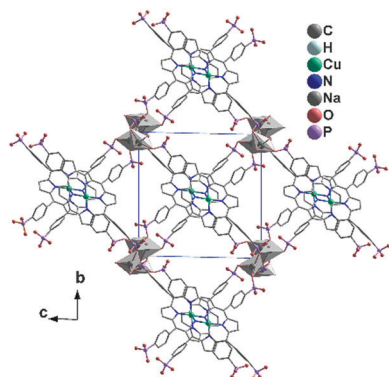


Fig. 12 Packing diagram of structure **57**. Adapted from ref. 64.

ions. By retaining chemically transformable groups like esters in some of these porphyrin-based MMPs, there exist further opportunities to diversify their structures and functions *via* crystal to crystal transformations. All these features are remarkable for the development of functional MOFs for future applications.

3.4. Cyclometalate or metal-polypyridine complexes as metalloligands

Metalloligands such as cyclometalated Ir(III) complexes and Ru(II) polypyridine complexes are very attractive for constructing functional coordination polymers because they can act not only as a chromophore or luminophore but also as a reversible redox reaction centre. By using Ir(ppy)₂(bppH) (ppy = 2-phenylpyridine, bppH₂ = 2-pyridylphosphonic acid) as the metalloligand, we isolated isostructural heptanuclear LnIr₆ clusters [LnIr₆(ppy)₁₂(bpp)₂(bppH)₄](CF₃SO₃)₃·xH₂O [Ln = Eu (**58**), Tb (**59**), Dy (**60**), Er (**61**), Yb (**62**); ppy = 2-phenylpyridine; bpp = 2-pyridylphosphonate].⁶⁵ The clusters are packed in the lattice, forming a supramolecular network containing channels in the *a*- and *c*-directions, where the lattice water molecules and the CF₃SO₃⁻ anions reside (Fig. 13). The two edge distances of the quadrilateral window in the *ab* plane are 15.877 and 16.958 Å, respectively. However, coordination polymers incorporating Ir(III) cyclometalate-phosphonate metalloligands have not been documented.

MMPs containing Ru(II)-polypyridine metalloligands are also rare. The only examples are Ln₃(H₂O)₉[Ru(4,4'-dppby)₃H_{5.5}]₂·nH₂O [Ln = La^{III} (**63**), Pr^{III} (**64**)],⁶⁶ where dppbyH₄ is 2,2'-bipyridinyl-4,4'-diphosphonic acid (Scheme 8). The two compounds

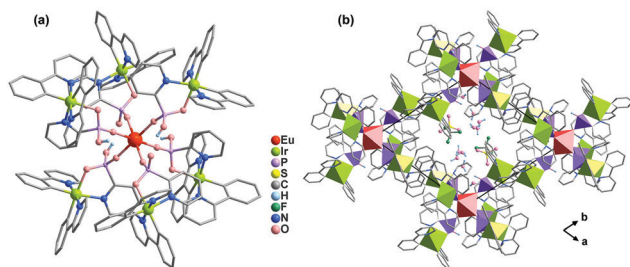
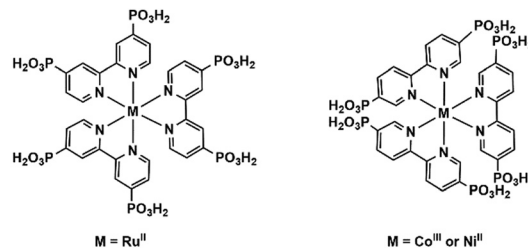


Fig. 13 (a) The heptanuclear cluster of **58**. (b) Packing diagram of **58** viewed along the *c*-axis. Color codes: EuO₆, red; IrC₂N₃O, yellow green; PO₃C, purple. Adapted from ref. 65.



Scheme 8 Molecular structures of metalloligands Ru^{II}(4,4'-dppbyH₄)₃²⁺ and [M(5,5'-dppbyH₄)₃]ⁿ⁺.



Scheme 9 Coordination modes of metalloligands Ru^{II}(4,4'-dppbyH₄)₃²⁺ and [M(5,5'-dppbyH₄)₃]ⁿ⁺.

are isostructural. Single crystal structural analysis of **64** reveals that it crystallizes in monoclinic space group *P2₁/n*. There are two kinds of Pr atoms in the structure. The Pr1 atom is coordinated by five phosphonate oxygen atoms and three water molecules to form a square antiprismatic geometry. The disordered Pr2 atom (half/half) is surrounded by five phosphonate oxygen atoms and three water molecules. Ru^{II}(4,4'-dppby)₃H_{5.5}^{1.5-} serves as an octa-dentate metalloligand, binding to eight Pr atoms *via* eight of its eighteen phosphonate oxygen atoms (Scheme 9a). Two of the six phosphonate groups are involved in bridging the Pr1 atoms into {Pr₁₂(OPO)₂} dimers. The remaining four phosphonate groups are involved in bridging the Pr1 atoms into {Pr₁₂(OPO)₂} dimers. The remaining four phosphonate groups are involved in bridging the Pr1 atoms into {Pr₁₂(OPO)₂} dimers. As a result, the Pr atom and {Pr₁₂(OPO)₂} dimers are cross-linked by the [Ru(4,4'-dppby)₃H_{5.5}]^{1.5-} metalloligands forming a three-dimensional open framework (Fig. 14). The framework contains two kinds of narrow porous channels along the *b*-axis with dimensions of *ca.* 4.4 × 6.0 Å and 3.2 × 4.6 Å (van der Waals radii not accounted).

The central metal ions in Ru^{II}(bpy)₃-based metalloligands can be substituted by other metal ions such as Co(II) or Ni(II).

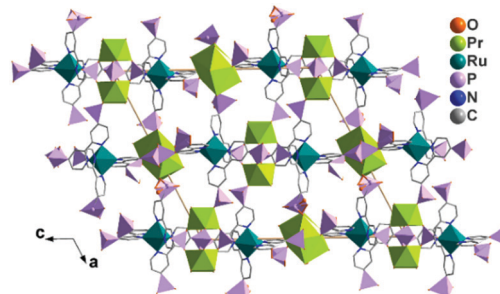


Fig. 14 Packing diagram of structure **64**. Adapted from ref. 66.

Rentschler and co-workers employed 2,2'-bipyridinyl-5,5'-diphosphonic acid (5,5'-dppbpyH₄) and obtained [LnCo^{III}(5,5'-dppbpyH₂)₃(H₂O)₂]₂·6.5H₂O [Ln = Gd^{III} (**65**), Dy^{III} (**66**) and Tb^{III} (**67**)] and [GdNi^{II}(5,5'-dppbpyH₂)₃(H₂O)₃]₂NaCl·6H₂O (**68**),⁶⁷ where M(5,5'-dppbpyH₄)₃ (M = Co^{III}, Ni^{II}) act as metalloligands (Scheme 8). Compounds **65**–**67** are isostructural, crystallizing in the space group *C2/c*. The lanthanide atom is seven-coordinated with a capped trigonal prismatic geometry, which is rare for lanthanide ions. The seven sites are occupied by five oxygen atoms from separate phosphonate groups of four Co(5,5'-dppbpyH₂)₃³⁻ and two water molecules. All phosphonates are mono protonated in the Co(5,5'-dppbpyH₂)₃³⁻ unit, which serves as a penta-dentate metalloligand and links four equivalent lanthanide atoms through its five phosphonate oxygen atoms (Scheme 9b). Therefore, each Ln(III) is connected by four Co(5,5'-dppbpyH₂)₃³⁻ units and *vice versa*, leading to a two-dimensional layer (Fig. 15a). The uncoordinated phosphonate groups are pendant on the two sides of the layer, and are involved in the interlayer hydrogen bonds.

Interestingly, different connectivity and dimensionality are observed in **68** when the metal center of Co^{III}(5,5'-dppbpyH₂)₃³⁻ is replaced by the divalent Ni ion. In this case, each Gd atom in **68** again has a capped trigonal prismatic coordination. But the seven coordination sites are filled with only four phosphonate oxygen atoms and three water molecules. Another significant difference for the metalloligand Ni(5,5'-dppbpyH₂)₃⁴⁻ is that it is tetra-dentate, binding to four Gd atoms *via* its four phosphonate groups. The two phosphonate groups of one 5,5'-dppbpyH₂ remain uncoordinated (Scheme 9c). As a result, each Gd is connected to four Ni(5,5'-dppbpyH₂)₃⁴⁻ and *vice versa*. Although this kind of connectivity is similar to that in **67**, the geometry of GdNi₄ in **68** differs from that of TbCo₄ in **67**. The four Ni atoms in GdNi₄ are arranged closer to a trigonal pyramid, unlike the

four Co atoms in TbCo₄ which are closer to a square plane. Thus a three-dimensional open framework is constructed in **68** which contains rhombic channels along the *c*-axis with dimensions of *ca.* 6.4 × 10.8 Å (van der Waals radii not accounted) (Fig. 15c). The sodium and chloride ions are occupying the space within the channels.

To summarize, the use of cyclometalate or metal–polypyridine complexes as metalloligands to construct MMPs brings not only internal functions such as luminescence and redox but also more diverse structures. The cyclometalated Ir(III) complex containing 2-pyridylphosphonate was used to assemble a series of LnIr₆ heptanuclear species. On the other hand, M(bpy)₃-based ligands were employed to build MMPs with 2D layer and 3D open framework structures. Moreover, the phosphonate groups in bipyridine can be placed in different substitution positions. The varied number of phosphonate groups in these metalloligands together with their different pre-organized conformation offers more coordination modes, leading to MMPs with different architectures. Although only a handful of examples have been given in this aspect, the result is encouraging in exploring new MMPs based on cyclometalate or metal–polypyridine complexes with versatile structures and interesting properties.

4. Properties and applications

Metal–metalloligand phosphonates offer several advantages over other metal phosphonates not only in structural complexity but also in functions. The presence of different metal ions could affect their physical properties and, in some cases, manipulate the property synergistically. The immobilization of the pre-organized functional metalloligand within metal phosphonate frameworks also provides opportunities in exploring new materials with high thermal and water stability, interesting magnetic, optical, and proton conductive properties and efficient heterogeneous catalytic performance.

4.1. Magnetic and optical properties

For metal–metalloligand phosphonates, the magnetic moment can originate from either the metalloligand, or the second metal ion, or both. Typical paramagnetic metalloligands are metal–metal bonded Ru₂(hedp)₂ⁿ⁻ species with paddlewheel structures. Early studies on diruthenium tetracarboxylates have demonstrated that the homovalent Ru₂^{II,II}(O₂CR)₄ and mixed-valent [Ru₂^{II,III}(O₂CR)₄][±] contain two (*S* = 1) and three unpaired electrons (*S* = 3/2) in electron configurations of σ²π⁴δ²(π*)²(δ*)² and σ²π⁴δ²π*²δ*¹.⁶⁸ Accordingly, the ground states for Ru₂^{II,II}(hedp)₂⁴⁻ and Ru₂^{II,III}(hedp)₂³⁻ are *S* = 1 and 3/2, respectively. As already described, a series of MMPs containing Ru₂⁵⁺ units have been obtained. They show large and positive zero-field splitting (ZFS) with the *D* parameters being 78.3–103.2 cm⁻¹ (Table 1). The interactions between the Ru₂ dimers are usually weak, which can be ferromagnetic as observed in layer compounds **1**–**2**, or antiferromagnetic as observed in Ln–Ru₂ frameworks. The introduction of the lanthanide ion into the Ru₂ system can

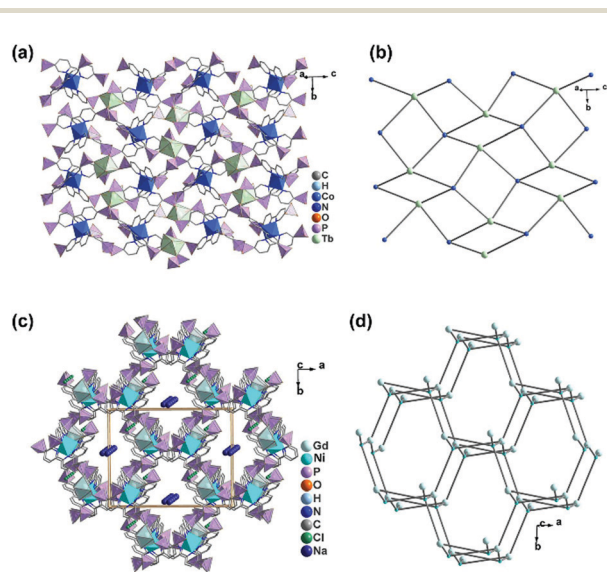


Fig. 15 (a) One layer and (b) layer topology of structure **67**. (c) Packing diagram and framework topology of structure **68**. The topologies in (b) and (d) are presented by the connections between the metalloligand centers and the lanthanide atoms. Adapted from ref. 67.

bring an additional magnetic source and multiple exchange pathways. For Dy–Ru₂ and Yb–Ru₂ compounds **17** and **18**, field-induced slow magnetic relaxation characteristic of single molecule magnet (SMM) behaviour was observed. But no peaks of in-phase and out-of-phase ac signals appear down to 2 K. Lanthanide complexes are well known to be promising systems for SMMs owing to their large single-ion magnetic anisotropy arising from the unquenched orbital angular momentum and strong spin–orbit coupling.⁶⁹ The observation of magnetization relaxation in **17** and **18** can be mainly attributed to the presence of Kramers ions Dy^{III} or Yb^{III}. Though the Ru₂⁵⁺ dimer also shows strong spin–orbit coupling, its contribution to magnetic relaxation remains unclear.

The magnetic properties of one metal ion could be affected by another metal ion in MMPs. An interesting example is [LaCo(notp)(H₂O)₄]_nH₂O (**22**)³⁰ which is diamagnetic with the octahedral Co^{III} ion in a low spin state. It has a layer structure where the metalloligand Co(notp)³⁻ is quadra-dentated binding to four La^{III} ions. Each La atom is eight-coordinated by four phosphonate oxygen atoms from four Co(notp)³⁻ and four water molecules. Thermal treatment up to 220 °C causes the release of both lattice and coordination water molecules. However, the coordination number of La remains the same as confirmed by EXAFS measurements. The vacant sites of La are filled with the phosphonate oxygen atoms from neighbouring Co(notp)³⁻, leading to the elongation of the Co–O bond length. The elongation of the Co–O length weakens the ligand field strength and stabilizes the high-spin Co^{II}. A thermally induced electron transfer occurs with the switching of the cobalt oxidation state, from diamagnetic low-spin Co^{III} to paramagnetic high-spin Co^{II} (Fig. 16). Moreover, the dehydrated H[hs-Co^{II}La^{III}(notp)] can turn back to the original [ls-Co^{III}La^{III}(notp)(H₂O)₄]_nH₂O upon rehydration in air. Compound **22** serves as a good example of metal–organic systems showing dehydration induced synergistic change of the coordination spheres of two kinds of metal ions linked by a polyazamacrocycle ligand and its impact on the magnetic properties.

As a comparison, the Ln^{III}–Fe^{III} and Ca^{II}–Co^{III} compounds **24**–**28** based on metalloligands M(notp)³⁻ (M = Fe, Co) cannot undergo such valence change. This can be explained by the fact that the oxidation state of Fe^{III} is not so sensitive to the

ligand field, the coordination number of Ca^{II} is less than that of La^{III}, and the Ca–O bonds are weaker than La–O bonds. By incorporating paramagnetic IrCl₆²⁻ as a counter-anion, [CaCo₂(notpH₃)₂(H₂O)₂](IrCl₆)·4H₂O (**34**) exhibits slow magnetic relaxation originating from isolated Ir^{IV}Cl₆²⁻ species.³⁵

Paramagnetic materials can be potential contrast agents (CAs) for magnetic resonance imaging (MRI). However, most of the research work has been focused on mononuclear Gd^{III} complexes using polyazamacrocycle derivatives as ligands due to their high stability in aqueous solution. MasPOCH and co-workers immobilized Gd^{III} ions in the framework structure of [GdCu(dotpH₄)Cl]·4.5H₂O (**37**),³⁷ which is stable in physiological saline solution and cell culture media. This material can be made into colloidal nanoparticles and shows interesting relaxometric properties with *r*₁ at a high field (500 MHz) of 5 mM⁻¹ s⁻¹ and a maximum *r*₁ = 15 mM⁻¹ s⁻¹ at 40 MHz. Furthermore, it can be used in a broad pH range, and the performance increases with temperature.

Reports on optically active MMPs are quite limited. For those containing metalloporphyrins, their UV-visible absorption spectra usually show a strong Soret band located at 410–440 nm (S₀ → S₂ transition) and less intensive Q bands centered at 520–650 nm (S₀ → S₁ transition).^{50–52,56,57,63} Heptanuclear clusters [LnIr₆(ppy)₁₂(bpp)₂(bppH)₄](CF₃SO₃)_x·xH₂O [Ln = Dy (**60**), Er (**61**), Yb (**62**)]⁶⁵ are multifunctional materials showing both phosphorescence originating from the cyclometalated Ir(III) component and field-induced single molecule magnet behaviour arising from the anisotropic lanthanide ions. Compounds Ln₃(H₂O)₉[Ru(4,4'-dppbpy)₃H_{5,5}]₂·nH₂O [Ln = La^{III} (**56**), Pr^{III} (**57**)] provide rare examples of MMPs incorporating the luminescent Ru(II) polypyridine metalloligand.⁶⁶ Interestingly, they exhibit a two-step vapochromism triggered by humidity. The triplet metal-to-ligand charge-transfer (³MLCT) emission is blue-shifted on increasing the relative humidity (RH) in the low RH region, and red-shifted in the high RH region, resulting in the highest-energy ³MLCT emission at medium RH. The origin is ascribed to the water-adsorption triggered reconstruction of the porous structure and the proton release from the metalloligand to the water filled channels.

4.2. Proton conductive properties

Proton conductive materials play critical roles in many renewable energy and electronics technologies.⁷⁰ Metal phosphonates are promising candidates owing to their high thermal and water stability, their ability to provide phosphonic acid groups as the proton source within extensive H-bond networks and their insulating nature, which are essential for high performance proton conductors.¹³ Since the number of MMPs is quite limited, only a few of them have been subjected to proton conductivity measurements. As shown in Tables 2 and 3, most compounds show conductivity in the range of 10⁻⁷–10⁻⁴ S cm⁻¹ at 95% RH and 25 °C. The conductivity can increase with increasing temperature, reaching 10⁻³ S cm⁻¹ as observed for compound **53** at 90% RH and 80 °C.

It is interesting to observe the enhancement of proton conductivity of [CoLa(notpH)(H₂O)₆]ClO₄·5H₂O (**23**)³¹ via



Fig. 16 The scheme of the valence change between **22** and H[hs-Co^{II}La^{III}(notp)]. Adapted from ref. 30.

in situ solid state phase transition. This compound is closely related to $[\text{CoLa}(\text{notp})(\text{H}_2\text{O})_4]_n\text{H}_2\text{O}$ (**22**) except for one phosphonate group protonated and hence an extra perchloride anion for charge compensation. Proton conductivity measurements on pellets of the powder samples of **22** and **23** show identical conductivity at 95% RH and 25 °C, being 3.0×10^{-6} and $3.5 \times 10^{-6} \text{ S cm}^{-1}$, respectively. Remarkably, **23** undergoes a solid state phase transition at 45 °C and 93% RH, forming a new phase $(\text{H}_3\text{O})[\text{CoLa}(\text{notp})(\text{H}_2\text{O})_4]\text{ClO}_4 \cdot x\text{H}_2\text{O}$ which has the same layer structure as that of **22** but with additional hydronium ions and ClO_4^- anions between the neutral layers. Consequently, this new phase shows an enhanced proton conductivity of $4.2 \times 10^{-5} \text{ S cm}^{-1}$ at 25 °C and 95% RH, one order of magnitude higher than that of **22** ($3.00 \times 10^{-6} \text{ S cm}^{-1}$) and **23** ($3.50 \times 10^{-6} \text{ S cm}^{-1}$).

The single crystal proton conductivity measurement can offer a chance to understand the proton transport mechanism. We used a single crystal of $[\text{CoCa}(\text{notpH}_2)(\text{H}_2\text{O})_2]\text{ClO}_4 \cdot 4\text{H}_2\text{O}$ (**28**)³³ for AC impedance measurements. The result demonstrates that the [010] direction of H-bond extension is the preferred proton conduction pathway showing the highest conductivity of $1.00 \times 10^{-3} \text{ S cm}^{-1}$ at 25 °C and 95% RH. On the other hand, the [20-1] direction, which involves the phosphonate oxygen atoms in the H-bond network, shows the lowest conductivity of $4.35 \times 10^{-8} \text{ S cm}^{-1}$ at 25 °C and 95% RH. Moreover, compound **28** experiences a reversible humidity dependent SC-SC structural transformation forming $[\text{CoCa}(\text{notpH}_2)(\text{H}_2\text{O})_2]\text{ClO}_4 \cdot 2\text{H}_2\text{O}$ (**29**) below 70% RH at room temperature with the continuous H-bond network interrupted (Fig. 17a and b). The structural transformation leads to a drastic decrease in proton conductivity by ~5 orders of magnitude (Fig. 17c). The process is reversible and the response is quick (Fig. 17d), and thus the material can be used as an impedance humidity sensor for future applications.

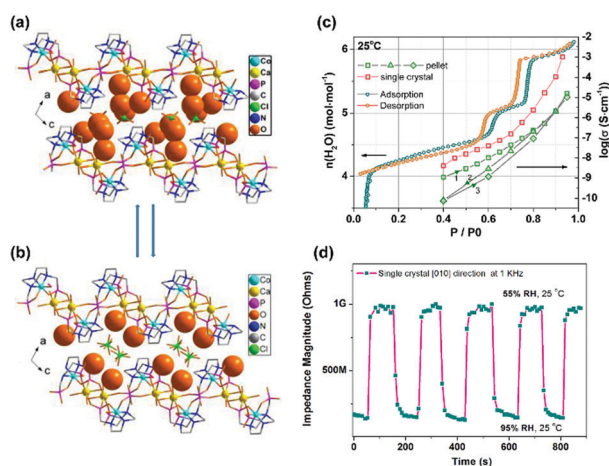


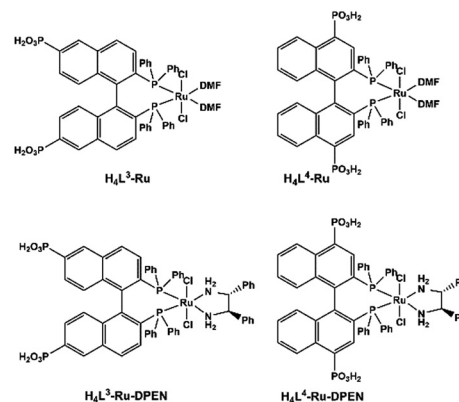
Fig. 17 (a and b) Packing diagrams of structures $[\text{CaCo}(\text{notpH}_2)(\text{H}_2\text{O})_2](\text{ClO}_4) \cdot 4\text{H}_2\text{O}$ (**28**) (a) and $[\text{CaCo}(\text{notpH}_2)(\text{H}_2\text{O})_2](\text{ClO}_4) \cdot 2\text{H}_2\text{O}$ (**29**) (b). (c) Proton conductivity and adsorption isotherms of **28**. (d) Impedance response and recovery time to RH variation between 55% and 95% for a single crystal of **28** along the [010] direction. Reprinted with permission from ref. 33. Copyright 2015, American Chemical Society.

To increase the proton conductivity of MMPs, one way is to increase the concentration of protons within the extensive H-bond network by providing more protonated phosphonate groups or acidic counterions. Another way is to produce composites by introducing salts into the framework or combine the MMP with suitable polymers. All these remain to be further explored.

4.3. Catalytic properties

MMPs can rationally immobilize catalytically active metal sites within the robust frameworks. Metalloporphyrins are prominent candidates for catalysis. Although the catalytic properties of MOFs containing metalloporphyrins have been well-explored,⁷¹ those of MMPs still remain rarely described due to the limited examples of MMPs. Manganese porphyrins $\text{Mn}(\text{porphyrin})\text{Cl}$ are active homogeneous catalysts for many oxidation reactions with H_2O_2 as the oxidizing agent in the presence of nitrogen base cocatalysts. By functionalizing manganese porphyrin with four phosphonate groups, Bujoli and co-workers reported the first examples of insoluble zinc phosphonate hybrid materials incorporating manganese porphyrins.⁷² They act as efficient heterogeneous oxidation catalysts in the epoxidation of cyclooctene with PhIO showing similar yields to their mononuclear precursor. Moreover, these catalysts show a different behaviour from their homogeneous counterparts in the competitive hydroxylation of alkane mixtures using PhIO as the oxidant, attributed to shape selectivity imposed by the phosphonate support.⁷³

The introduction of chirality into catalytically active MMPs can extend their application in enantioselective catalysis. Lin and co-workers designed and synthesized a series of chiral porous zirconium phosphonates embedding enantiopure and catalytically active Ru species (Scheme 10). Compounds $[\text{Zr}\{\text{Ru}(\text{L}^3)(\text{DMF})_2\text{Cl}_2\}] \cdot 2\text{MeOH}$ and $[\text{Zr}\{\text{Ru}(\text{L}^4)(\text{DMF})_2\text{Cl}_2\}] \cdot 2\text{MeOH}$ are found to be highly active catalysts for asymmetric hydrogenation of β -ketoesters.⁷⁴ $\text{Zr}[\text{Ru}(\text{L}^3)(\text{DPEN})\text{Cl}_2] \cdot 4\text{H}_2\text{O}$ and $\text{Zr}[\text{Ru}(\text{L}^4)(\text{DPEN})\text{Cl}_2] \cdot 4\text{H}_2\text{O}$ are excellent catalysts for enantioselective hydrogenation of unfunctionalized aromatic ketones.⁷⁵



Scheme 10 Molecular structures of Ru-L⁴ complexes of binaphthalene-bisphosphonic acids. Reprinted with permission from ref. 10f. Copyright 2020, Springer Nature.

5. Summary and outlook

In this feature article, we have presented the current status of metal–metalloid phosphonates (MMPs) from syntheses, structures to properties. It is clear that using the pre-formed metalloid as the reaction precursor is a more efficient and controllable approach to construct MMPs. Further, the metalloid can bring multiple phosphonate groups as coordination donors with different orientations, resulting in MMPs with new structure types and special functions such as polarity. Moreover, the immobilized metalloid with magnetic, optical, and catalytic properties can offer additional functions of MMPs for potential applications.

Given the limited types of known metallo-phosphonate ligands, the limited number of MMPs and the limited studies on their properties, the development of this field is still in its early stage.

Future challenges remain in several aspects. First, the exploration of new MMPs. This exploration can be based on the known phosphonate-based metalloids. For example, the assembly of paddlewheel diruthenium diphosphonate with transition metal ions or clusters has not been investigated. For metalloids based on polyazamacrocycles, the numbers of MMPs and the structural types are still very limited. Even less has been studied for metalloporphyrin and metallo-tris-bipyridine-phosphonate ligands. So there is plenty of room for these already known metalloids to construct new MMPs by optimizing reaction conditions and introducing different metal ions or clusters as nodes. The exploration of new MMPs can also be based on new phosphonate-based metalloids by pre-design and synthesis. In this regard, there are a great many choices considering the vast reservoir of organic chelating and bridging ligands. Second, the characterization of MMPs. Metal phosphonates are well recognized by their low solubility and easy precipitation. It is often difficult to grow single crystals with sufficiently large size for conventional single crystal determination. The situation is particularly true for MMPs because the metalloids usually possess more than one phosphonate groups which makes the solubility of MMPs even lower. Apart from the synthetic efforts to grow crystals with bigger size, a promising solution to solve the structures is by Rietveld refinement using powder X-ray diffraction data and, in some cases, combining with the single-crystal electron diffraction method. Third, the application of MMPs. The properties of MMPs originate from both the metalloids and the second metal ions or clusters. By designing and synthesizing specific metalloids, it is possible to construct MMPs with desired functions. For instance, the combination of paddlewheel diruthenium diphosphonate with paramagnetic transition metal ions may result in magnetic materials with long-range ordering, like that found for diruthenium phosphates.⁷⁶ The MMPs containing polar metallo-polyazamacrocycles may be designed and synthesized to possess polarity or even chirality for further applications. The incorporation of luminescent cyclometalate or metal–polypyridine complexes in MMPs brings optical functions for application in sensors and multifunctional materials. By taking

the advantages of metalloporphyrin, the related MMPs are expected to show interesting electrical and optical properties that can be utilized in photocatalysis and photodynamic therapy. Transformation of the channels may lead to switching in electrical and optical properties of these porphyrin-based materials. Some compounds exhibit conductance from combined mechanism such as proton conductivity together with inherent semiconducting properties, thus showing potential in multifunctional materials. Finally, the functional MMPs may be fabricated into nanoparticles, nanowires or nanosheets which may lead to applications in catalysis, molecular devices and nanotechnologies.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

Financial support from the National Natural Science Foundation of China (21731003 and 21671098) and the National Key R&D Program of China (2018YFA0306004) is acknowledged.

Notes and references

- (a) H. Furukawa, K. E. Cordova, M. O'Keeffe and O. M. Yaghi, *Science*, 2013, **341**, 1230444; (b) S. Horike, S. Shimomura and S. Kitagawa, *Nat. Chem.*, 2009, **1**, 695–704; (c) W.-X. Zhang, P.-Q. Liao, R.-B. Lin, Y.-S. Wei, M.-H. Zeng and X.-M. Chen, *Coord. Chem. Rev.*, 2015, **293–294**, 263–278; (d) T. D. Bennett, A. K. Cheetham, A. H. Fuchs and F.-X. Coudert, *Nat. Chem.*, 2017, **9**, 11–16; (e) A. Kirchon, L. Feng, H. F. Drake, E. A. Joseph and H.-C. Zhou, *Chem. Soc. Rev.*, 2018, **47**, 8611–8638.
- (a) B. B. Lin, S. C. Xiang, H. B. Xing, W. Zhou and B.-L. Chen, *Coord. Chem. Rev.*, 2019, **378**, 87–103; (b) S. Dissegna, K. Epp, W. R. Heinz, G. Kieslich and R. A. Fischer, *Adv. Mater.*, 2018, **30**, 1704501; (c) J.-G. Jia and L.-M. Zheng, *Coord. Chem. Rev.*, 2020, **403**, 213083; (d) Q. Wang and D. Astruc, *Chem. Rev.*, 2020, **120**, 1438–1511.
- (a) Y. Cui, B. Chen and G. Qian, *Coord. Chem. Rev.*, 2014, **273**, 76–86; (b) K. Chen and C. Wu, *Chin. Chem. Lett.*, 2018, **29**, 823–826; (c) L. Zhu, X.-Q. Liu, H.-L. Jiang and L.-B. Sun, *Chem. Rev.*, 2017, **117**, 8129–8176; (d) G. M. Espallargas and E. Coronado, *Chem. Soc. Rev.*, 2018, **47**, 533–557; (e) J. Calbo, M. J. Golomb and A. Walsh, *J. Mater. Chem. A*, 2019, **7**, 16571–16597; (f) D.-W. Lim and H. Kitagawa, *Chem. Rev.*, 2020, **120**, 8416–8467.
- (a) G. Kumar and R. Gupta, *Chem. Soc. Rev.*, 2013, **42**, 9403–9453; (b) E. Lee, S. Seo, S. S. Lee and L. F. Lindoy, *Coord. Chem. Rev.*, 2017, **348**, 121–170.
- (a) M. C. Das, S. Xiang, Z. Zhang and B. Chen, *Angew. Chem., Int. Ed.*, 2011, **50**, 10510–10520; (b) J. Liu, L. Chen, H. Cui, J. Zhang, L. Zhang and C.-Y. Su, *Chem. Soc. Rev.*, 2014, **43**, 6011–6061; (c) M. Y. Masoomi, A. Morsali, A. Dhakshinamoorthy and H. Garcia, *Angew. Chem., Int. Ed.*, 2019, **58**, 15188–15205.
- A. Clearfield, *Progress in Inorganic Chemistry*, Wiley, New York, 1998, vol. 47, pp. 371–510.
- (a) A. D. G. Firmino, F. Figueira, J. P. C. Tomé, F. A. Almeida Paz and J. Rocha, *Coord. Chem. Rev.*, 2018, **355**, 133–149; (b) J. M. Taylor, K. W. Dawson and G. K. H. Shimizu, *J. Am. Chem. Soc.*, 2013, **135**, 1193–1196; (c) Y. Wang, J. G. Alauzun and P. H. Mutin, *Chem. Mater.*, 2020, **32**, 2910–2918.
- (a) T. Zheng, Z. Yang, D. Gui, Z. Liu, X. Wang, X. Dai, S. Liu, L. Zhang, Y. Gao, L. Chen, D. Sheng, Y. Wang, J. Diwu, J. Wang, R. Zhou, Z. Chai, T. E. Albrecht-Schmitt and S. Wang, *Nat. Commun.*, 2017, **8**, 15369; (b) S.-B. Liu, S.-S. Bao and L.-M. Zheng, *Dalton Trans.*, 2020, **49**, 3758–3765.
- Metal Phosphonate Chemistry: From Synthesis to Applications*, ed. A. Clearfield and K. Demadis, The Royal Society of Chemistry, 2012.

- 10 (a) A. Clearfield, *Curr. Opin. Solid State Mater. Sci.*, 2002, **6**, 495–506; (b) K. J. Gagnon, H. P. Perry and A. Clearfield, *Chem. Rev.*, 2012, **112**, 1034–1054; (c) W. Yang, T. G. Parker and Z.-M. Sun, *Coord. Chem. Rev.*, 2015, **303**, 86–109; (d) A. D. G. Firmino, F. Figueira, J. P. C. Tomé, F. A. Almeida Paz and J. Rocha, *Coord. Chem. Rev.*, 2018, **355**, 133–149; (e) G. Yücesan, Y. Zorlu, M. Stricker and J. Beckmann, *Coord. Chem. Rev.*, 2018, **369**, 105–122; (f) G.-G. Weng and L.-M. Zheng, *Sci. China: Chem.*, 2020, **63**, 619–636; (g) P. Tholen, Y. Zorlu, J. Beckmann and G. Yücesan, *Eur. J. Inorg. Chem.*, 2020, 1542–1554.
- 11 J.-G. Mao, *Coord. Chem. Rev.*, 2007, **251**, 1493–1520.
- 12 S.-S. Bao and L.-M. Zheng, *Coord. Chem. Rev.*, 2016, **319**, 63–85.
- 13 S.-S. Bao, G. K. H. Shimizu and L.-M. Zheng, *Coord. Chem. Rev.*, 2019, **378**, 577–594.
- 14 (a) S. J. Shearan, N. Stock, F. Emmerling, J. Demel, P. A. Wright, K. D. Madis, M. Vassaki, F. Costantino, R. Vivani, S. Sallard, I. R. Salcedo, A. Cabeza and M. Taddei, *Crystals*, 2019, **9**, 270; (b) P. Bhanja, J. Na, T. Jing, J. Lin, T. Wakihara, A. Bhaumik and Y. Yamauchi, *Chem. Mater.*, 2019, **31**, 5343–5362.
- 15 (a) N. Hermer, H. Reinsch, P. Mayer and N. Stock, *CrystEngComm*, 2016, **18**, 8147–8150; (b) T. Willhammar, Y. Yun and X. Zou, *Adv. Funct. Mater.*, 2014, **24**, 173.
- 16 (a) F. A. Cotton, C. A. Murillo and R. A. Walton, *Multiple Bonds Between Metal Atoms*, Springer Science and Business Media, New York, 3rd edn, 2005; (b) K. P. Kornecki, J. F. Berry, D. C. Powers and T. Ritter, *Progress in Inorganic Chemistry*, John Wiley & Sons, Inc., 2014, vol. 58, pp. 225–302.
- 17 (a) H. Miyasaka, *Acc. Chem. Res.*, 2013, **46**, 248–257; (b) K. Uemura, N. Uesugi, A. Matsuyama, M. Ebihara, H. Yoshikawa and K. Awaga, *Inorg. Chem.*, 2016, **55**, 7003–7011.
- 18 J.-H. Yang, R.-M. Cheng, Y.-Y. Jia, J. Jin, B.-B. Yang, Z. Cao and B. Liu, *Dalton Trans.*, 2016, **45**, 2945–2954.
- 19 (a) F. A. Cotton, T. Datta, L. Labella and M. Shang, *Inorg. Chim. Acta*, 1993, **203**, 55–60; (b) B.-B. Yang, L.-N. Feng, X.-M. Fan, K.-X. Zhang, J.-H. Yang and B. Liu, *Inorg. Chem. Front.*, 2017, **4**, 1061–1065.
- 20 X.-Y. Yi, L.-M. Zheng, W. Xu and S.-H. Feng, *Inorg. Chem.*, 2003, **42**, 2827–2829.
- 21 B. Liu, Y.-Z. Li and L.-M. Zheng, *Solid State Sci.*, 2006, **8**, 1041–1045.
- 22 (a) X.-Y. Yi, B. Liu, R. Jiménez-Paricio, F. A. Urbanos, S. Gao, W. Xu, J.-S. Chen, Y. Song and L.-M. Zheng, *Inorg. Chem.*, 2005, **44**, 4309–4314; (b) D. Wang, H.-Q. Yang, G.-H. Wu, X.-F. Hou, J.-H. Yang and B. Liu, *Inorg. Chem. Commun.*, 2014, **46**, 241–243.
- 23 B. Liu, B.-L. Li, Y.-Z. Li, Y. Chen, S.-S. Bao and L.-M. Zheng, *Inorg. Chem.*, 2007, **46**, 8524–8532.
- 24 S. M. Elahi, Q.-H. Lai, M. Ren, S.-S. Bao, M. Kurmoo and L.-M. Zheng, *Inorg. Chem.*, 2019, **58**, 14034–14045.
- 25 M. Yu Antipin, A. P. Baranov, M. I. Kabachnik, T. Ya Medved, Yu. M. Polikarpov, Yu. T. Struchkov and B. K. Shcherbakov, *Dokl. Akad. Nauk SSSR*, 1986, **287**, 130.
- 26 M. I. Kabachnik, M. Yu Antipin, B. K. Shcherbakov, A. P. Baranov, Yu. T. Struchkov, T. YaMedved and Yu. M. Polikarpov, *Koord. Khim.*, 1988, **14**, 536.
- 27 S.-S. Bao, G.-S. Chen, Y. Wang, Y.-Z. Li, L.-M. Zheng and Q.-H. Luo, *Inorg. Chem.*, 2006, **45**, 1124–1129.
- 28 (a) S.-S. Bao, L.-F. Ma, Y. Wang, L. Fang, C.-J. Zhu, Y.-Z. Li and L.-M. Zheng, *Chem. – Eur. J.*, 2007, **13**, 2333–2343; (b) M. Ren, S.-S. Bao, N. Hoshino, T. Akutagawa, B. Wang, Y.-C. Ding, S. Wei and L.-M. Zheng, *Chem. – Eur. J.*, 2013, **19**, 9619–9628; (c) M. Ren, S.-S. Bao, R. A. S. Ferreira, L.-M. Zheng and L. D. Carlos, *Chem. Commun.*, 2014, **50**, 7621–7624; (d) M. Ren, C. D. S. Brites, S.-S. Bao, R. A. S. Ferreira and L.-M. Zheng, *J. Mater. Chem. C*, 2015, **3**, 8480–8484; (e) M. Ren, S.-S. Bao, B.-W. Wang, R. A. S. Ferreira, L.-M. Zheng and L. D. Carlos, *Inorg. Chem. Front.*, 2015, **2**, 558–566.
- 29 S.-B. Liu, S.-S. Bao and L.-M. Zheng, *Dalton Trans.*, 2020, **49**, 3758–3765.
- 30 S.-S. Bao, Y. Liao, Y.-H. Su, X. Liang, F.-C. Hu, Z. Sun, L.-M. Zheng, S. Wei, R. Alberto, Y.-Z. Li and J. Ma, *Angew. Chem., Int. Ed.*, 2011, **50**, 5504–5508.
- 31 S.-S. Bao, K. Otsubo, J. M. Taylor, Z. Jiang, L.-M. Zheng and H. Kitagawa, *J. Am. Chem. Soc.*, 2014, **136**, 9292–9295.
- 32 Y.-H. Su, S.-S. Bao and L.-M. Zheng, *Inorg. Chem.*, 2014, **53**, 6042–6047.
- 33 S.-S. Bao, N.-Z. Li, J. M. Taylor, Y. Shen, H. Kitagawa and L.-M. Zheng, *Chem. Mater.*, 2015, **27**, 8116–8125.
- 34 S.-S. Bao, Y.-X. Wu, N.-Z. Li and L.-M. Zheng, *Eur. J. Inorg. Chem.*, 2016, 4476–4482.
- 35 S.-B. Liu, X.-D. Huang, S.-S. Bao, M. Kurmoo and L.-M. Zheng, *Cryst. Growth Des.*, 2019, **19**, 4836–4843.
- 36 (a) M. P. C. Campello, S. Lacerda, I. C. Santos, G. A. Pereira, C. F. G. C. Geraldies, J. Kotek, P. Hermann, J. Vaněk, P. Lubal, V. Kubiček, É. Tóth and I. Santos, *Chem. – Eur. J.*, 2010, **16**, 8446–8465; (b) S. Hassfjell, K. O. Kongshaug and C. Rømming, *Dalton Trans.*, 2003, 1433–1437; (c) Z. Kotková, G. A. Pereira, K. Djanashvili, J. Kotek, J. Rudovský, P. Hermann, L. V. Elst, R. N. Muller, C. F. G. C. Geraldies, I. Lukeš and J. A. Peters, *Eur. J. Inorg. Chem.*, 2009, 119–136; (d) R. Ševčík, J. Vaněk, R. Michalíková, P. Lubal, P. Hermann, I. C. Santos, I. Santos and M. P. C. Campello, *Dalton Trans.*, 2016, **45**, 12723–12733.
- 37 A. Carné-Sánchez, C. S. Bonnet, I. Imaz, J. Lorenzo, É. Tóth and D. Maspocho, *J. Am. Chem. Soc.*, 2013, **135**, 17711–17714.
- 38 F. Avecilla, J. A. Peters and C. F. G. C. Geraldies, *Eur. J. Inorg. Chem.*, 2003, 4179–4186.
- 39 (a) T.-H. Yang, K. Zhou, S.-S. Bao, C.-J. Zhu and L.-M. Zheng, *Inorg. Chem. Commun.*, 2008, **11**, 1075–1078; (b) D. Kong, D. G. Medvedev and A. Clearfield, *Inorg. Chem.*, 2004, **43**, 7308–7314.
- 40 J. Arriñez-Soriano, J. Albalad, C. Vila-Parrondo, J. Pérez-Carvajal, S. Rodríguez-Hermida, A. Cabeza, J. Juanhuix, I. Imaz and D. Maspocho, *Inorg. Commun.*, 2016, **52**, 7229–7232.
- 41 G. Giambastiani, W. Oberhauser, C. Bianchini, F. Laschi, L. Sorace, P. Brueggeller, R. Gutmann, A. Orlandini and F. Vizza, *Eur. J. Inorg. Chem.*, 2005, 2027–2031.
- 42 J. Rudovský, P. Cígler, J. Kotek, P. Hermann, P. Vojtišek, I. Lukeš, J. A. Peters, L. V. Elst and R. N. Muller, *Chem. – Eur. J.*, 2005, **11**, 2373–2384.
- 43 P. Vojtišek, P. Cígler, J. Kotek, J. Rudovský, P. Hermann and I. Lukeš, *Inorg. Chem.*, 2005, **44**, 5591–5599.
- 44 G. Neri, M. Forster, J. J. Walsh, C. M. Robertson, T. J. Whittles, P. Farràs and A. J. Cowan, *Inorg. Commun.*, 2016, **52**, 14200–14203.
- 45 M. Paúrová, T. David, I. Císařová, P. Lubal, P. Hermann and J. Kotek, *New J. Chem.*, 2018, **42**, 11908–11929.
- 46 J.-G. Mao, Z. Wang and A. Clearfield, *Inorg. Chem.*, 2002, **41**, 3713–3720.
- 47 *Handbook of Porphyrin Science*, ed. K. M. Kadish, K. M. Smith and R. Guilard, World Scientific Press, Singapore, 2010–2014.
- 48 (a) W.-Y. Gao, M. Chrzanowski and S. Ma, *Chem. Soc. Rev.*, 2014, **43**, 5841–5866; (b) Z. Chen, J. Liu, L. Zhang and C. Su, *Acta Chim. Sin.*, 2019, **77**, 242–252; (c) J. Chen, Y. Zhu and S. Kaskel, *Angew. Chem., Int. Ed.*, 2020, **59**, DOI: 10.1002/anie.201909880.
- 49 (a) D. Feng, W.-C. Chung, Z. Wei, Z.-Y. Gu, H.-L. Jiang, Y.-P. Chen, D. J. Darensbourg and H.-C. Zhou, *J. Am. Chem. Soc.*, 2013, **135**, 17105–17110; (b) T.-F. Liu, D. Feng, Y.-P. Chen, L. Zou, M. Bosch, S. Yuan, Z. Wei, S. Fordham, K. Wang and H.-C. Zhou, *J. Am. Chem. Soc.*, 2015, **137**, 413–419.
- 50 Y. Y. Enakieva, A. G. Bessmertnykh, Y. G. Gorbunova, C. Stern, Y. Rousselin, A. Y. Tsvadze and R. Guilard, *Org. Lett.*, 2009, **11**, 3842–3845.
- 51 A. A. Sinelshchikova, S. E. Nefedov, Y. Yu. Enakieva, Y. G. Gorbunova, A. Yu. Tsvadze, K. M. Kadish, P. Chen, A. Bessmertnykh-Lemeune, C. Stern and R. Guilard, *Inorg. Chem.*, 2013, **52**, 999–1008.
- 52 R. I. Zubatyuk, A. A. Sinelshchikova, Y. Y. Enakieva, Y. G. Gorbunova, A. Y. Tsvadze, S. E. Nefedov, A. Bessmertnykh-Lemeune, R. Guilard and O. V. Shishkin, *CrystEngComm*, 2014, **16**, 10428–10438.
- 53 Y. Fang, X. Jiang, K. M. Kadish, S. E. Nefedov, G. A. Kirakosyan, Y. Y. Enakieva, Y. G. Gorbunova, A. Y. Tsvadze, C. Stern, A. Bessmertnykh-Lemeune and R. Guilard, *Inorg. Chem.*, 2019, **58**, 4665–4678.
- 54 M. A. Uvarova, A. A. Sinelshchikova, M. A. Golubnichaya, S. E. Nefedov, Y. Yu. Enakieva, Y. G. Gorbunova, A. Yu. Tsvadze, C. Stern, A. Bessmertnykh-Lemeune and R. Guilard, *Cryst. Growth Des.*, 2014, **14**, 5976–5984.
- 55 E. V. Ermakova, Y. Yu. Enakieva, S. E. Nefedov, V. V. Arslanov, Y. G. Gorbunova, A. Yu. Tsvadze, C. Stern and A. Bessmertnykh-Lemeune, *Eur. J. Org. Chem.*, 2019, 3146–3162.
- 56 A. Yu. Mitrofanov, Y. Rousselin, V. N. Khrustalev, A. V. Cheprakov, A. Bessmertnykh-Lemeune and I. P. Beletskaya, *Eur. J. Inorg. Chem.*, 2019, 1313–1328.
- 57 (a) S. E. Nefedov, K. P. Birin, A. Bessmertnykh-Lemeune, Y. Y. Enakieva, A. A. Sinelshchikova, Y. G. Gorbunova, A. Y. Tsvadze,

- C. Stern, Y. Fang and K. M. Kadish, *Dalton Trans.*, 2019, **48**, 5372–5383; (b) E. V. Vinogradova, Y. Y. Enakieva, S. E. Nefedov, K. P. Birin, A. Y. Tsivadze, Y. G. Gorbunova, A. G. B. Lemeune, C. Stern and R. Guillard, *Chem. – Eur. J.*, 2012, **18**, 15092–15104.
- 58 Y. Y. Enakieva, M. V. Volostnykh, S. E. Nefedov, G. A. Kirakosyan, Y. G. Gorbunova, A. Y. Tsivadze, A. G. Bessmertnykh-Lemeune, C. Stern and R. Guillard, *Inorg. Chem.*, 2017, **56**, 3055–3070.
- 59 S. B. Ungashe, W. L. Wilson, H. E. Katz, G. R. Scheller and T. M. Putvinski, *J. Am. Chem. Soc.*, 1992, **114**, 8717–8719.
- 60 T. Rhauderwiek, K. Wolkersdörfer, S. Øien-Ødegaard, K.-P. Lillerud, M. Wark and N. Stock, *Chem. Commun.*, 2018, **54**, 389–392.
- 61 Y. Y. Enakieva, A. A. Sinelshchikova, M. S. Grigoriev, V. V. Chernyshev, K. A. Kovalenko, I. A. Stenina, A. B. Yaroslavtsev, Y. G. Gorbunova and A. Y. Tsivadze, *Chem. – Eur. J.*, 2019, **25**, 10552–10556.
- 62 B. Wang, T. Rhauderwiek, A. K. Inge, H. Xu, T. Yang, Z. Huang, N. Stock and X. Zou, *Chem. – Eur. J.*, 2018, **24**, 17429–17433.
- 63 T. Rhauderwiek, H. Zhao, P. Hirschele, M. Döblinger, B. Bueken, H. Reinsch, D. De Vos, S. Wuttke, U. Kolb and N. Stock, *Chem. Sci.*, 2018, **9**, 5467–5478.
- 64 M. Maares, M. M. Ayhan, K. B. Yu, A. O. Yazaydin, K. Harmandar, H. Haase, J. Beckmann, Y. Zorlu and G. Yücesan, *Chem. – Eur. J.*, 2019, **25**, 11214–11217.
- 65 (a) D. Zeng, M. Ren, S.-S. Bao, L. Li and L.-M. Zheng, *Chem. Commun.*, 2014, **50**, 8356–8359; (b) D. Zeng, K. Fan, L.-P. Wang, S.-S. Bao, M. Ren and L.-M. Zheng, *J. Magn. Magn. Mater.*, 2019, **484**, 139–145.
- 66 A. Kobayashi, K. Shimizu, A. Watanabe, Y. Nagao, N. Yoshimura, M. Yoshida and M. Kato, *Inorg. Chem.*, 2019, **58**, 2413–2421.
- 67 C. Köhler and E. Rentschler, *Dalton Trans.*, 2016, **45**, 12854–12861.
- 68 (a) F. A. Cotton, V. M. Miskowski and B. Zhong, *J. Am. Chem. Soc.*, 1989, **111**, 6177; (b) G. J. Norman, G. E. Renzoni and D. A. Case, *J. Am. Chem. Soc.*, 1979, **101**, 5256.
- 69 (a) D. N. Woodruff, R. E. P. Winpenny and R. A. Layfield, *Chem. Rev.*, 2013, **113**, 5110–5148; (b) Y.-S. Meng, S.-D. Jiang, B.-W. Wang and S. Gao, *Acc. Chem. Res.*, 2016, **49**, 2381–2389; (c) H. Tian and L.-M. Zheng, *Acta Chim. Sin.*, 2020, **78**, 34–55.
- 70 K. D. Kreuer, S. J. Paddison, E. Spohr and M. Schuster, *Chem. Rev.*, 2004, **104**, 4637–4678.
- 71 (a) M. M. Pereira, L. D. Dias and M. J. F. Calvete, *ACS Catal.*, 2018, **8**, 10784–10808; (b) C. F. Pereira, M. M. Q. Simões, J. P. C. Tomé and F. A. A. Paz, *Molecules*, 2016, **21**, 1348; (c) S. Nakagaki, G. K. B. Ferreira, G. M. Ucoski and K. A. D. de Freitas Castro, *Molecules*, 2013, **18**, 7279–7308.
- 72 D. Demand, B. Schollorn, D. Mansuy, J. Rouxel, P. Battioni and B. Bujoli, *Chem. Mater.*, 1995, **7**, 995–1000.
- 73 D. Deniaud, G. A. Spyroulias, J.-F. Bartoli, P. Battioni, D. Mansuy, C. Pinel, F. Odobel and B. Bujoli, *New J. Chem.*, 1998, **22**, 901–905.
- 74 A. Hu, H. L. Ngo and W. Lin, *Angew. Chem., Int. Ed.*, 2003, **115**, 6182–6185.
- 75 A. Hu, H. L. Ngo and W. Lin, *J. Am. Chem. Soc.*, 2003, **125**, 11490–11491.
- 76 B.-B. Yang, L.-N. Feng, X.-M. Fan, K.-X. Zhang, J.-H. Yang and B. Liu, *Inorg. Chem. Front.*, 2017, **4**, 1061–1065.