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# Enhanced corrosion resistance of zinc phosphate coatings on mild steel through incorporation of nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>–CuO nanocomposite†

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Herein, 5–10 nm-sized rod-like nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>–CuO nanocomposite were synthesized and incorporated into zinc phosphate coatings as corrosion inhibition additives. The formulated coatings were studied thoroughly on mild steel in 3.5% NaCl solution and compared with unmodified coating. The concentration of the additive was varied from 0.3 g L<sup>−1</sup> to 0.9 g L<sup>−1</sup> with three step-ups. X-ray and surface analysis results revealed that both additives favoured the formation of more compact coatings with densely packed zinc phosphate crystals and reduced iron content. Compared with the unmodified coating, the metal oxide incorporated coatings showed low corrosion current density (*I*<sub>corr</sub>) and corrosion rate and high polarisation resistance (*R*<sub>p</sub>) and charge transfer resistance (*R*<sub>ct</sub>), confirming that these nanomaterials act as corrosion inhibitors. Among the three concentrations used, nanocrystalline CeO<sub>2</sub> performed well at 0.6 g L<sup>−1</sup> and CeO<sub>2</sub>–CuO nanocomposite performed well even at low concentration of 0.3 g L<sup>−1</sup>, with a more than fifty-one times lower corrosion rate and twenty times higher polarisation resistance (*R*<sub>p</sub>) compared with unmodified coating. Nanocrystalline CeO<sub>2</sub> increased the *R*<sub>ct</sub> from 114 Ω cm<sup>2</sup> to 427.5 Ω cm<sup>2</sup> at 0.6 g L<sup>−1</sup>, whereas the CeO<sub>2</sub>–CuO nanocomposite increased this value up to 3645 Ω cm<sup>2</sup> at 0.3 g L<sup>−1</sup>. Enhanced corrosion resistance of these coatings was confirmed with an accelerated salt-spray test.

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## 1. Introduction

Low carbon steel, commonly known as mild steel in which the carbon content normally varies from 0.05 to 0.3 weight percentage, is known for its high ductility, machinability and weldability. In addition to these properties, its cost-effectiveness makes it a highly sought-after material for many large and small fabrications, notably automobile parts such as body panels, camshaft, shaft axles, engine cylinders and bolts, gears, brackets, pins and washers. This alloy is more susceptible to rusting when exposed to moisture, which deteriorates the metal's performance, leading to random failures.<sup>1</sup> Despite this drawback, mild steel remains an ideal choice for applications where high tensile strength and durability are required.

Preventive measures such as using corrosion inhibitors and applying protective coatings can help mitigate these failures.<sup>2–5</sup> Coatings are one of the smart ways to control corrosion. Phosphate conversion coatings—particularly zinc phosphate

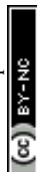
conversion coatings—are more advantageous over other coating methods.<sup>6–8</sup> This coating protects the surface of the steel against corrosion by forming a thin, non-conducting layer of insoluble phosphates.<sup>9</sup> Zinc phosphate coatings can be applied to a variety of materials, including high carbon steel, medium carbon steel,<sup>10</sup> low carbon steel,<sup>11</sup> galvanized steel,<sup>12</sup> magnesium alloy,<sup>13</sup> and aluminum alloy.<sup>14</sup> Owing to the ease of application, efficiency and compatibility with various substrates, this coating emerged as a preferred choice for many industrial applications.<sup>15</sup>

However, the corrosion resistance of the phosphate coatings in adverse environments is often limited due to open pores and cracks through which the electrolytes diffuse and react with the underlying steel surface and induce the conversion of iron into iron oxides and hydroxides.<sup>16,17</sup> Many additives improve the coating performance. Careful selection and infusion of these additives are vital for achieving desired coating characteristics, such as improved corrosion resistance, adhesion, hardness, and overall durability of the coated metal surface.

In the recent past, the use of nanomaterials as corrosion inhibitors gained significant attention. The presence of nanomaterials in the coating composition control the electrochemical process that occurs at the interface of the coating and the metal surface. Metal oxides like SiO<sub>2</sub>,<sup>18</sup> ZrO<sub>2</sub>,<sup>19</sup> TiO<sub>2</sub>,<sup>20</sup> and

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Al<sub>2</sub>O<sub>3</sub> (ref. 21) in the nanoscale, nanocomposites like Al<sub>2</sub>O<sub>3</sub>/Mo<sup>22</sup> and two-dimensional nanosheets like graphene oxide<sup>23,24</sup> and boron nitride<sup>25</sup> were studied as additives for zinc phosphate coatings and the results were promising. Their unique properties such as high surface area, water repellent nature, antioxidant properties, and barrier functions against aggressive agents collectively improve the performance of the coatings.<sup>26</sup> They also speed-up the phosphate treatment process at room temperature, and thus reduce energy consumption.

Bagal and co-workers found that the corrosion rate significantly reduced from 8 mpy to 3.5 mpy, and the coating porosity decreased by 50% upon incorporation of TiO<sub>2</sub> in the phosphate coating.<sup>20</sup> In another report, Young Zhou and co-workers explored the effect of the addition of nano CeO<sub>2</sub> into zinc phosphate coatings applied on magnesium alloy AZ91D and showed improved adhesion and micro-hardness of the coating at a concentration of 2 g L<sup>-1</sup>.<sup>27</sup> Riyas and co-workers reported Al<sub>2</sub>O<sub>3</sub>/Mo composite incorporated zinc phosphate coatings for galvanised steel.<sup>22</sup> They found that the Al<sub>2</sub>O<sub>3</sub>/Mo particles acted as nucleating agents, forming highly compact and dense phosphate crystals. Moreover, the resulting coating showed high charge-transfer resistance and improved corrosion resistance. Deepa and co-workers formulated zinc coatings with BiVO<sub>4</sub>/TiO<sub>2</sub> composites.<sup>28</sup> This coating achieved a water contact angle of 120.2°, which improved its self-cleaning properties. Additionally, it demonstrated a high inhibition efficiency of 91.68% and great corrosion resistance. These recent findings recognise metal oxides as anti-corrosion additives for coatings.

The present study focuses on abundant rare earth metal oxide nanoparticle, *i.e.*, nano CeO<sub>2</sub> and its nanocomposite with CuO as additives for zinc phosphate coatings. Generally, this class of metal oxides enhances microstructure formation, refines grain boundaries, and improves overall durability of the coating.<sup>29</sup> Rare earth elements inhibit corrosion by forming an insoluble film of oxides and hydroxides, which replaces the natural oxide film on metal surfaces.<sup>30</sup> Among the various rare earth elements, cerium with high oxidation state (Ce<sup>3+</sup> and Ce<sup>4+</sup>) forms stable oxides that provide barrier protection against corrosion.<sup>31</sup> As far as copper is concerned, it has good thermal stability, mechanical workability and acts as like a noble metal in stringent environmental conditions. A nano copper oxide coating on the same metal formed by anodization was found to increase the corrosion protection efficiency by 86.2% in 3.5% NaCl and 74.5% in 2 mg per L NH<sub>3</sub>.<sup>32</sup> The insoluble nature of copper in water increases the hydrophobicity of the coating.<sup>33</sup> We aimed to develop new zinc phosphate conversion coatings using nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite as corrosion inhibition additives. Accordingly, some zinc phosphate conversion coatings were formulated and applied on mild steel. Herein, the synthesis of nanoadditives was reported and the anti-corrosion performance of the coatings was impressive.

## 2. Materials and methods

### 2.1 Materials

Analytical grade cerous nitrate hexahydrate (Ce(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O, CAS number 10294-41-4), cupric nitrate trihydrate

(Cu(NO<sub>3</sub>)<sub>2</sub>·3H<sub>2</sub>O, CAS 10031-43-3), zinc oxide (ZnO, CAS number 1314-13-2), orthophosphoric acid (85% phosphoric acid (H<sub>3</sub>PO<sub>4</sub>), CAS number 7664-38-2), sodium hydroxide, sodium carbonate, sodium nitrite and sodium chloride were purchased from Sisco Research Laboratories Pvt. Ltd and used without further purification. Mild steel of grade IS 513/2008 GR4 was purchased from the local market and its chemical composition was given in Table S1 (ESI).† All the experiments were carried out in distilled water.

### 2.2 Synthesis of CeO<sub>2</sub> nanoparticle

CeO<sub>2</sub> in the form of nanoparticle was prepared from cerium nitrate hexahydrate (Ce(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O) and sodium carbonate (Na<sub>2</sub>CO<sub>3</sub>) by the precipitation method. A 0.2 M solution of cerium nitrate hexahydrate was prepared by dissolving 4.34 g of Ce(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O in 50 mL of distilled water and the solution was subjected to continuous stirring. To this, a 1 M sodium carbonate solution was added dropwise until the pH of the solution reached 10. The whole mixture was stirred for a further 60 minutes, resulting in a pale, yellow coloured liquid-solid suspension. This suspension was centrifuged at 2000 rpm, filtered and washed with deionized water several times to remove the impurities. The obtained precipitate was first dried in a vacuum oven maintained at 80 °C, milled to obtain a fine powder and subsequently calcined at 450 °C in the presence of air for 3 h to obtain CeO<sub>2</sub> as a nanopowder.

### 2.3 Preparation of CeO<sub>2</sub>-CuO nanocomposite

CeO<sub>2</sub>-CuO nanocomposite was prepared *via* the co-precipitation method. Twenty-five mL of 0.2 M solutions of Ce(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O and Cu(NO<sub>3</sub>)<sub>2</sub>·3H<sub>2</sub>O were mixed together and stirred for 1 h. Next, a 1 M sodium hydroxide solution was added dropwise to the solution until the pH of the solution attained 10. Then, the whole mixture was stirred for 60 min to obtain a pale blue coloured liquid-solid suspension. The rest of the procedure employed to obtain CeO<sub>2</sub>-CuO nanocomposite was similar to that adopted for the preceding synthesis.

### 2.4 Characterizations of CeO<sub>2</sub> nanoparticle and CeO<sub>2</sub>-CuO nanocomposite

The functional groups of the synthesised metal oxide and composite were analysed using a Thermo iS50 Fourier-transform infra-red spectrophotometer (FT-IR) in the mid IR range of 400–4000 cm<sup>-1</sup>; samples were prepared using the KBr pellet method. The morphology and elemental distribution of the prepared CeO<sub>2</sub> nanoparticle and CeO<sub>2</sub>-CuO nanocomposite were examined by field emission scanning electron microscopy (FE-SEM) combined with energy dispersive X-ray spectroscopy (EDAX) (TESCAN VEGA3 model and Carl Zeiss Sigma 300 model instruments). The phase and crystalline characteristics of the prepared CeO<sub>2</sub> and CeO<sub>2</sub>-CuO composite were studied using a powder X-ray diffraction (XRD) technique with an Empyrean Series III Diffractometer operated with a monochromatic Cu Kα radiation source (λ = 0.15418 nm). The samples were examined between 2θ of 20° and 90° with a 0.05° step-up and 0.02 s per

step. The crystallite size was calculated using the Debye-Scherrer formula as given in eqn (1).

$$D = K\lambda/\beta \cos \theta \quad (1)$$

where  $D$  represents the average crystallite size,  $K$  is a dimensionless shape factor (0.98),  $\lambda$  is the X-ray wavelength,  $\theta$  is the diffraction angle and  $\beta$  is the full width at half maximum of the diffraction peak in radians. The elemental compositions of the prepared nanomaterials were analysed using X-ray photoelectron spectroscopy (XPS) (Thermo Fisher, model K-Alpha-KAN9954133) using monochromatic Al-K $\alpha$  radiation at 150 W with an X-ray power gun at room temperature. The thermal stability of the nanoparticle and nanocomposite were determined by thermogravimetry (TG) and differential thermal analyses (DTA). These experiments were performed with a simultaneous thermal analyser (Shimadzu, DTG 60 Series) under a nitrogen atmosphere at a 100 mL per min gas flow rate. The temperature was varied from 30 °C to 700 °C and the heating rate was set at 10 °C min<sup>-1</sup>.

## 2.5 Preparation of phosphating bath

First, the phosphating bath solution was partially prepared by mixing 7 g of zinc oxide (ZnO) in 14 mL of 85% ortho phosphoric acid (H<sub>3</sub>PO<sub>4</sub>). Stirring the obtained paste like mass with 50 mL of distilled water in a magnetic stirrer for 30 min resulted in a clear solution. Then, this solution was diluted to 1 litre with distilled water. The quantity of zinc oxide and orthophosphoric acid was fixed after several trial experiments. Appropriate quantities of the remaining ingredients were added to this partial solution immediately prior to experiments to make a complete formulation. The chemical composition of the zinc phosphating bath is given in Table 1.

The conventional zinc phosphate coating composition was modified by the addition of synthesised CeO<sub>2</sub> nanoparticle and CeO<sub>2</sub>-CuO nanocomposite into the phosphating bath. Based on the quantity of the additive added (0.3 g, 0.6 g or 0.9 g) into the zinc phosphating bath, the obtained coatings were named as (i) ZnP-0.3CO, ZnP-0.6CO, ZnP-0.9CO, respectively for cerium oxide and (ii) ZnP-0.3COCO, ZnP-0.6COCO, ZnP-0.9COCO, respectively for cerium oxide-copper oxide composite. The zinc phosphate coating formulated without the synthesised

nanomaterials as additive is simply represented as ZnP in this study.

## 2.6 Determination of free acid and total acid

To obtain quality coatings, bath parameters such as total acid value (TA), free acid value (FA) and their ratio should be maintained at appropriate levels. These values were calculated by titration of the bath solution against 0.1 N NaOH. The total acid value was determined by titrating 10 mL of the bath solution against 0.1 N NaOH using phenolphthalein as indicator. The end-point was colour change of the bath solution from colourless to pink (1 mL = 1 point). The free acid value was determined by following the same procedure using methyl orange as indicator. Here, the end-point was a colour change of the bath solution from orange to yellow. In general, the total acid value is maintained between 10 and 25 (points) and the free acid value is maintained between 0.5 and 3 (points) to obtain optimal coatings. The ratio of free acid to total acid gives the acid coefficient of the bath; the higher the acid co-efficient value, the better the coating produced. The total acid value, free acid value and the acid coefficient of the phosphating bath are given in Table 1.

## 2.7 Preparation of substrate and phosphate treatment

Steel panels were fabricated to a dimension of 20 mm × 20 mm × 0.8 mm. The residual rust and scales on the steel surface were removed by abrading them with SiC papers differing in grit number and washed with distilled water. Contaminants like dirt, oil and grease were removed by soaking the panel in 10% NaOH solution at room temperature for 30 min and washed with excess distilled water followed by rinsing with acetone. The cleaned steel substrates were immersed in a zinc phosphating bath at room temperature for 30 min with the composition given in Table 1. After this surface treatment, the panels were dried in air and kept in a desiccator for further studies.

## 2.8 Characterizations of zinc phosphate coatings

**2.8.1 Evaluation of corrosion resistance of CeO<sub>2</sub> nanoparticle and CeO<sub>2</sub>-CuO nanocomposite incorporated zinc phosphate coating.** The corrosion protection properties of the phosphate coated steel panels were studied in 3.5 wt% NaCl aqueous solution employing an electrochemical work station (CH Instruments Model 1130A). The electrochemical analyses were performed with a three-electrode cell system; the developed coating with an exposed area of 1 cm<sup>2</sup> was used as a working electrode, calomel electrode was used as reference electrode and platinum electrode was used as a counter electrode. The coatings were immersed in NaCl solution for 30 min before commencing the analysis to achieve a stable open circuit potential. The potentiodynamic polarisation curves were plotted to determine the corrosion current density ( $I_{\text{corr}}$ ) and corrosion potential ( $E_{\text{corr}}$ ). The polarisation resistance ( $R_p$ ) was calculated using the Stern-Geary eqn (2).

$$R_p = \frac{\beta_a \times \beta_c}{2.3i_{\text{corr}}(\beta_a + \beta_c)} \quad (2)$$

Table 1 Chemical composition of zinc phosphating bath

Ingredient	Quantity
ZnO	7 g L <sup>-1</sup>
Phosphoric acid	14 mL L <sup>-1</sup>
Sodium nitrite	1 g L <sup>-1</sup>
Nanocrystalline CeO <sub>2</sub> <sup>a</sup>	0.3 g L <sup>-1</sup> (or), 0.6 g L <sup>-1</sup> (or), 0.9 g L <sup>-1</sup>
CeO <sub>2</sub> -CuO nanocomposite <sup>a</sup>	0.3 g L <sup>-1</sup> (or), 0.6 g L <sup>-1</sup> (or), 0.9 g L <sup>-1</sup>
Total acid	28.2 (Points)
Free acid	2.9 (Points)
TA : FA	9.7 : 1
pH	2.8

<sup>a</sup> Either CeO<sub>2</sub> or CeO<sub>2</sub>-CuO was used for a formulation.



where  $\beta_a$ ,  $\beta_c$  and  $I_{\text{corr}}$  are the anodic Tafel slope, cathodic Tafel slope and corrosion current density, respectively. The corrosion rate in mm per year was calculated using eqn (3).

$$\text{CR} = K \frac{i_{\text{corr}}}{\rho} \text{EW} \quad (3)$$

where  $K$  ( $3.27 \times 10^{-3}$ ) is the unit conversion factor,  $I_{\text{corr}}$  is the corrosion current density in  $\mu\text{A cm}^{-2}$ ,  $\rho$  is the density of the metal in  $\text{g cm}^{-3}$  and EW is the equivalent weight of the metal (considered dimensionless).

The data obtained from the electrochemical impedance spectroscopy were presented as Nyquist and Bode plots. Bode

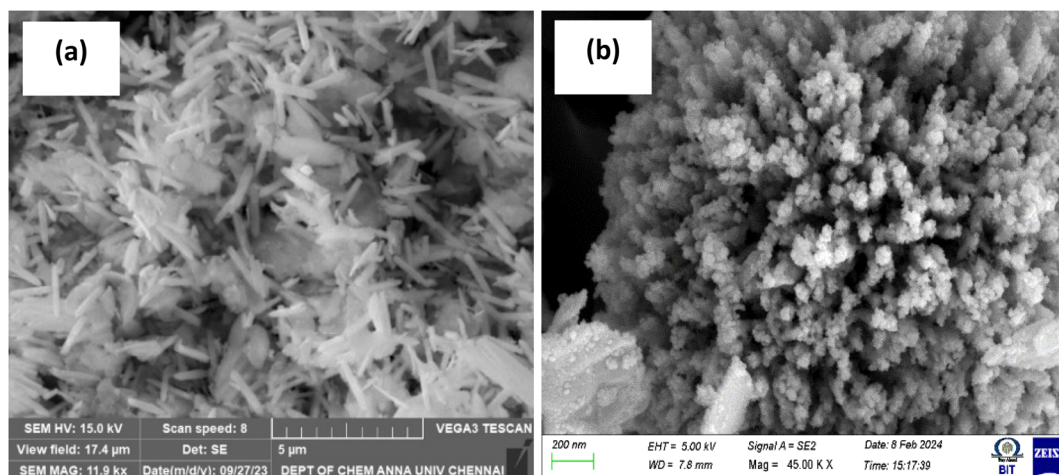


Fig. 1 SEM micrograph of (a) nanocrystalline CeO<sub>2</sub> (scale bar = 5 μm) and (b) CeO<sub>2</sub>-CuO nanocomposite (scale bar = 200 nm).

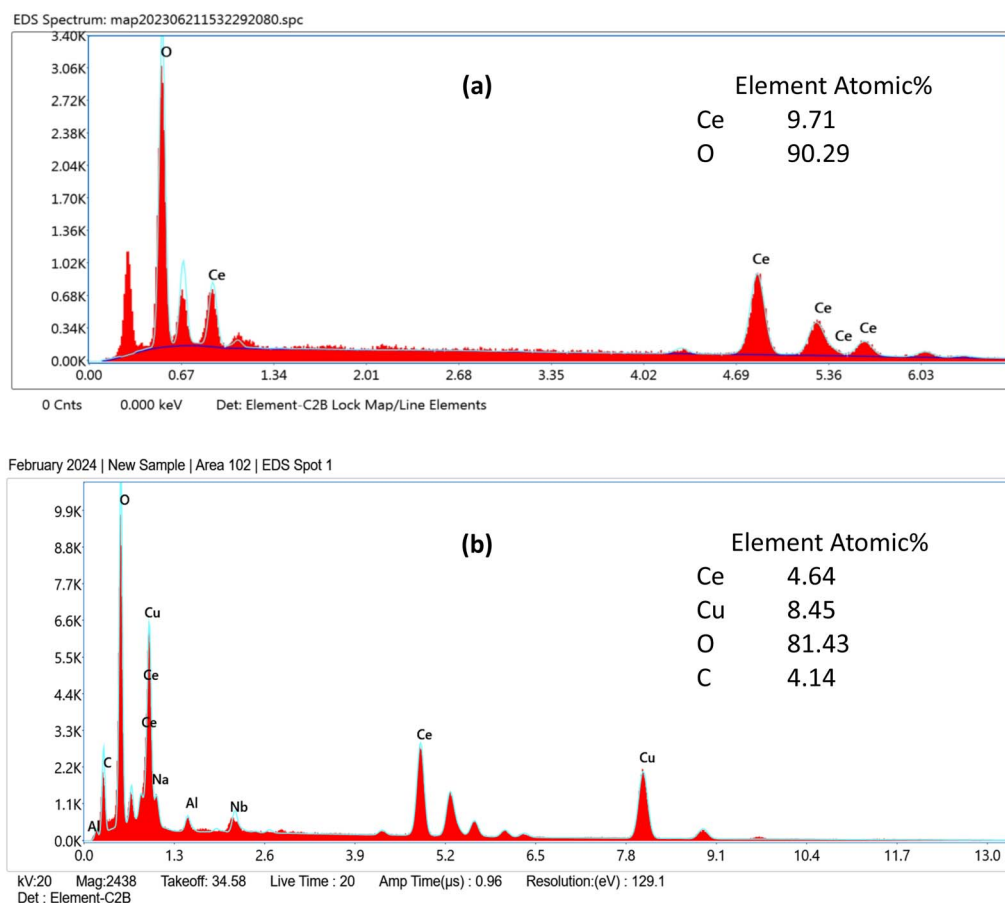


Fig. 2 EDAX spectrum of (a) nanocrystalline CeO<sub>2</sub> and (b) CeO<sub>2</sub>-CuO nanocomposite.



impedance and the phase angle were plotted as  $\log|Z|$  vs.  $\log(f)$  and  $-\text{phase angle}$  vs.  $\log(f)$ , where  $|Z|$  represents the absolute impedance and  $f$  indicates the frequency. The obtained data were fitted using ZSimpWin software (Princeton Applied Research, USA) to obtain an appropriate equivalent circuit.

**2.8.2 Salt-spray test.** The coated samples (panels) were subjected to a neutral salt spray test. These tests were performed in a closed chamber, where a 5% sodium chloride solution was atomized into a fine mist and sprayed over the coatings. This salt fog simulates a corrosive environment in accordance with the ASTM B117 standard. After completion of 48 hours of spray time, the corrosion resistance of the coatings was recorded through photographs of rust formed and evaluated. From the coating's weight noted before and after the test, the corrosion rate was calculated using eqn (4).

$$\text{Corrosion rate} = 534W/DAT \quad (4)$$

where  $W$  = weight loss in grams;  $D$  = density of the metal in  $\text{g cm}^{-3}$ ;  $A$  = exposed area of the sample in  $\text{cm}^2$ ;  $T$  = time of exposure in hours and the numerical value; 534 is a conversion factor for  $\text{g cm}^{-2}$  to mpy units.

The percentage inhibition efficiency ' $P$ ' of the additives was calculated from the following relationship:

$$P = 100[\text{CR}_{\text{uninhibited}} - \text{CR}_{\text{inhibited}}/\text{CR}_{\text{uninhibited}}] \quad (5)$$

where  $\text{CR}_{\text{uninhibited}}$  and  $\text{CR}_{\text{inhibited}}$  are the corrosion rates of the coating without and with additive (*i.e.*,  $\text{CeO}_2$  nanoparticle or  $\text{CeO}_2$ -CuO nanocomposite), respectively.

Other characterizations of the coatings were described in the ESI section,<sup>†</sup> including the analysis of functional groups, surface morphology, elemental and phase composition, measurements of surface wetting, coating weight, hardness and adhesion tests.

### 3. Results and discussion

#### 3.1 Synthesis of nano $\text{CeO}_2$ and $\text{CeO}_2$ -CuO nanocomposite

Cerium oxide nanoparticles and its composite with copper oxide were synthesized *via* an economically simple-cum-attractive precipitation method as described in Fig. S1 and S2, respectively (ESI).<sup>†</sup> For both the synthesis, the respective hydrous forms of metal nitrates were used as precursors with sodium carbonate and sodium hydroxide used as precipitating agents. The reactions were carried out in water medium at pH = 10 and subsequent physical processes of fine grinding and calcination carried out at 450 °C yielded pale yellow  $\text{CeO}_2$  and greyish black  $\text{CeO}_2$ -CuO nanoparticles. The yield of the compound was quantitative, *i.e.*, >80%. It is worth mentioning that the amount of obtained  $\text{CeO}_2$  was considerably higher than the expected theoretical value, indicating that this material was in hydrated form.<sup>34,35</sup>

#### 3.2 Characterisation of prepared nanoadditives

**3.2.1 Formation, morphology and composition of nano  $\text{CeO}_2$  and  $\text{CeO}_2$ -CuO nanocomposite.** The FT-IR spectra of nano

$\text{CeO}_2$  and  $\text{CeO}_2$ -CuO nanocomposite were essentially identical and were shown in Fig. S3 (ESI).<sup>†</sup> They exhibited a peak for fundamental stretching of metal-oxygen bonds in the expected finger-print region at around  $500 \text{ cm}^{-1}$ , confirming the formation of metal oxides from the metal nitrate precursors. The broad and strong absorption peak at  $3433 \text{ cm}^{-1}$  and a strong absorption band at  $1384 \text{ cm}^{-1}$  was due to moisture adsorbed on the surface of the compounds. This suggests that these nano-materials are porous in nature. The peaks at  $1572 \text{ cm}^{-1}$  and  $1384 \text{ cm}^{-1}$  are respectively due to the anti and sym stretching vibration of the N-O group of residual nitrate reagents present in the compounds. A weak band at  $1059 \text{ cm}^{-1}$  is due to atmospheric carbon dioxide absorbed by the compounds during the storage and analysis.<sup>35-38</sup>

The surface morphology and nano-scale structure of  $\text{CeO}_2$  and  $\text{CeO}_2$ -CuO composite were examined through FE-SEM and the vivid pictures are shown in Fig. 1. The SEM images reveal that the  $\text{CeO}_2$  crystals possess small rod-like structures that grew randomly in all directions with an even surface.<sup>39</sup> On the other hand, the  $\text{CeO}_2$ -CuO composite grew perpendicularly with an uneven surface and showed a bamboo stem like structure. This well-defined surface morphology is an important

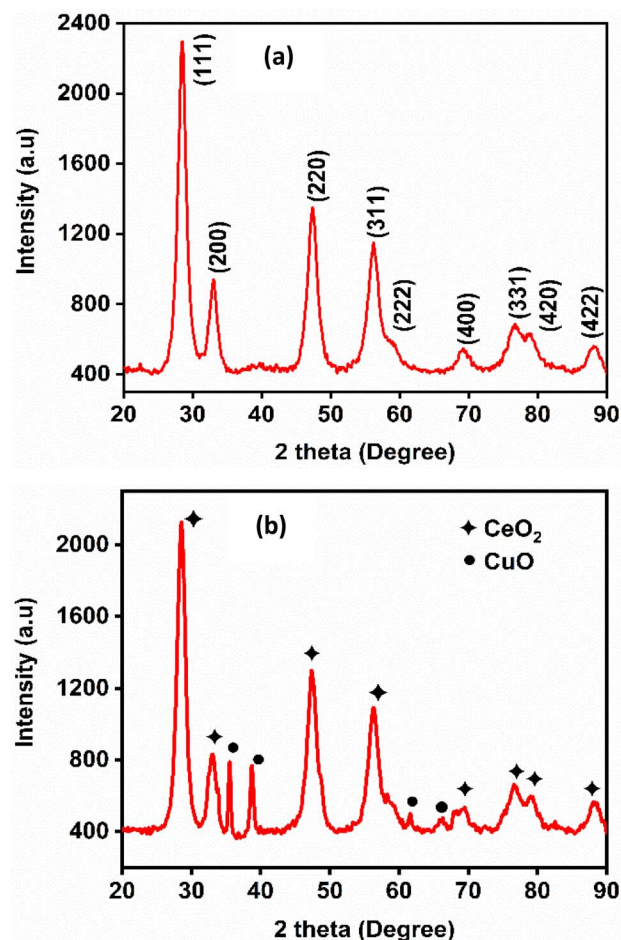


Fig. 3 XRD patterns of (a) nanocrystalline  $\text{CeO}_2$  and (b)  $\text{CeO}_2$ -CuO nanocomposite.



requirement that will enhance the electrochemical performance of the nanomaterials due to its high surface-to-volume ratio and large number of active sites.<sup>40</sup> This intrinsic characteristic of nanomaterials is particularly beneficial in the context of corrosion protection, where a large surface area leads to improved interactions between the protective layer and the corrosive environment. The EDAX spectra of both nano CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite along with elemental composition are shown in Fig. 2. The characteristic peaks corresponding to cerium, copper and oxygen appearing in the spectra with the expected ratios confirm the successful synthesis of these materials.

**3.2.2 Crystallinity and phase structure.** The XRD spectrum of CeO<sub>2</sub> displays reasonably sharp diffraction peaks at  $2\theta$  values of 28.53°, 32.92°, 47.35°, 56.33°, 58.80°, 69.17°, 76.56°, 78.69° and 88.20° (Fig. 3). These peaks correspond to Miller indices (111), (200), (220), (311), (222), (400), (331), (420) and (422), respectively,<sup>39</sup> indicating a defective face-centred cubic phase of CeO<sub>2</sub> and are consistent with JCPDS no. 34-0394. In the XRD spectrum of the CeO<sub>2</sub>-CuO composite (Fig. 3b), the peaks at 35.44°, 38.85°, 61.61° and 66.18° are assigned to the monoclinic phase of CuO corresponding to the (−111), (111), (−113) and (−311) planes, respectively in accordance with JCPDS no. 089-

2530. As this XRD pattern includes all characteristic peaks arising from both the defective face-centred cubic phase of CeO<sub>2</sub> and the monoclinic phase of CuO, it confirms the successful formation of CuO within the CeO<sub>2</sub> matrix.<sup>41</sup> The approximate crystallite sizes of nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite were found to be  $5.2 \pm 0.4$  nm and  $10.6 \pm 0.4$  nm, respectively, calculated using the Debye-Scherrer eqn (1).

The chemical state and elemental composition of nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite were investigated using XPS and the results were illustrated in Fig. 4 and 5, respectively. The survey scan spectrum of CeO<sub>2</sub> identified the presence of C 1s, O 1s, and Ce 3d (Fig. 4a). The XPS spectrum of CeO<sub>2</sub>-CuO composite was similar to that of CeO<sub>2</sub>; it showed the presence of Cu 3d in addition to all other elements observed for CeO<sub>2</sub> (Fig. 5a). The elemental compositions derived from the XPS survey scans for both CeO<sub>2</sub> and CeO<sub>2</sub>-CuO are shown in the respective figures: the data are consistent with the size of the elements and composition of the materials. In the high-resolution Ce 3d spectrum of both CeO<sub>2</sub> and CeO<sub>2</sub>-CuO composite (Fig. 4b and 5b, respectively), eight peaks were observed corresponding to the spin-orbit splitting of Ce 3d<sub>5/2</sub> and Ce 3d<sub>3/2</sub>. This indicates that CeO<sub>2</sub> contains both Ce<sup>3+</sup> and

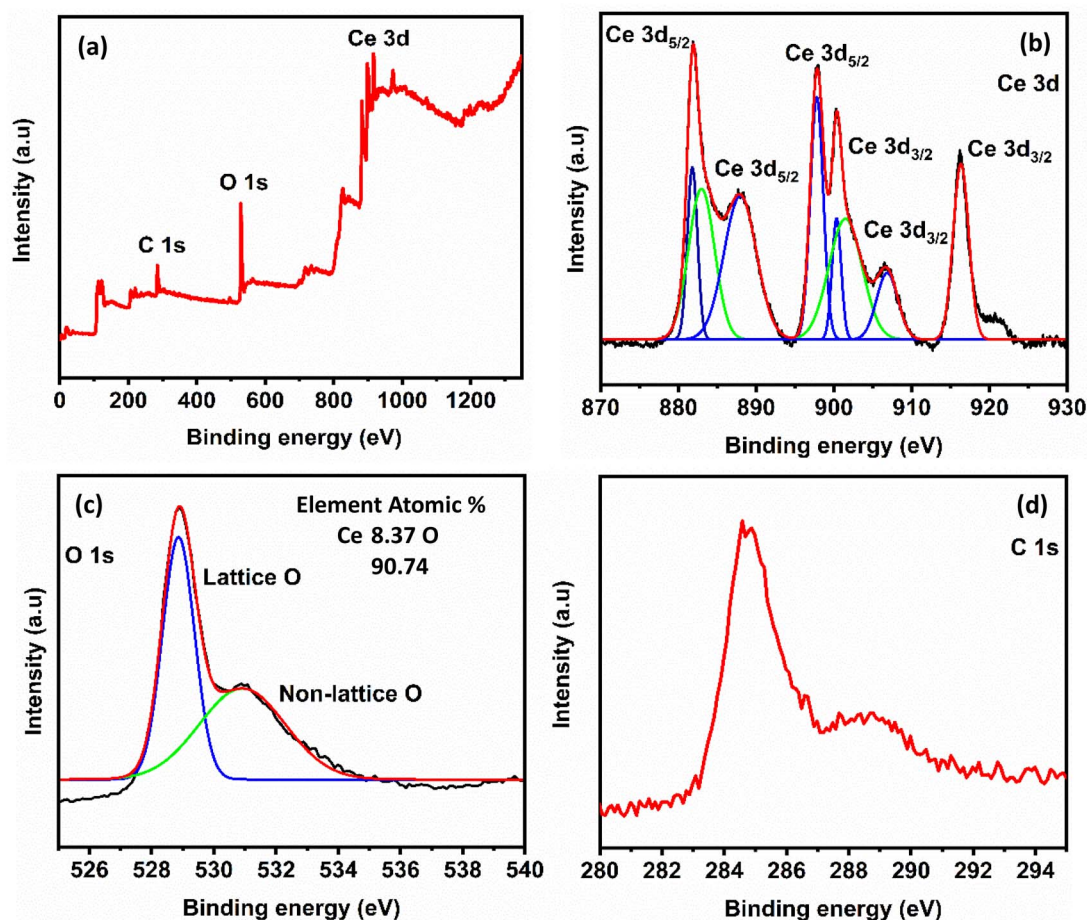


Fig. 4 (a) XPS spectrum of nanocrystalline CeO<sub>2</sub>, (b) high resolution spectrum of Ce 3d, (c) high resolution spectrum of O 1s and (d) high resolution spectrum of C 1s.

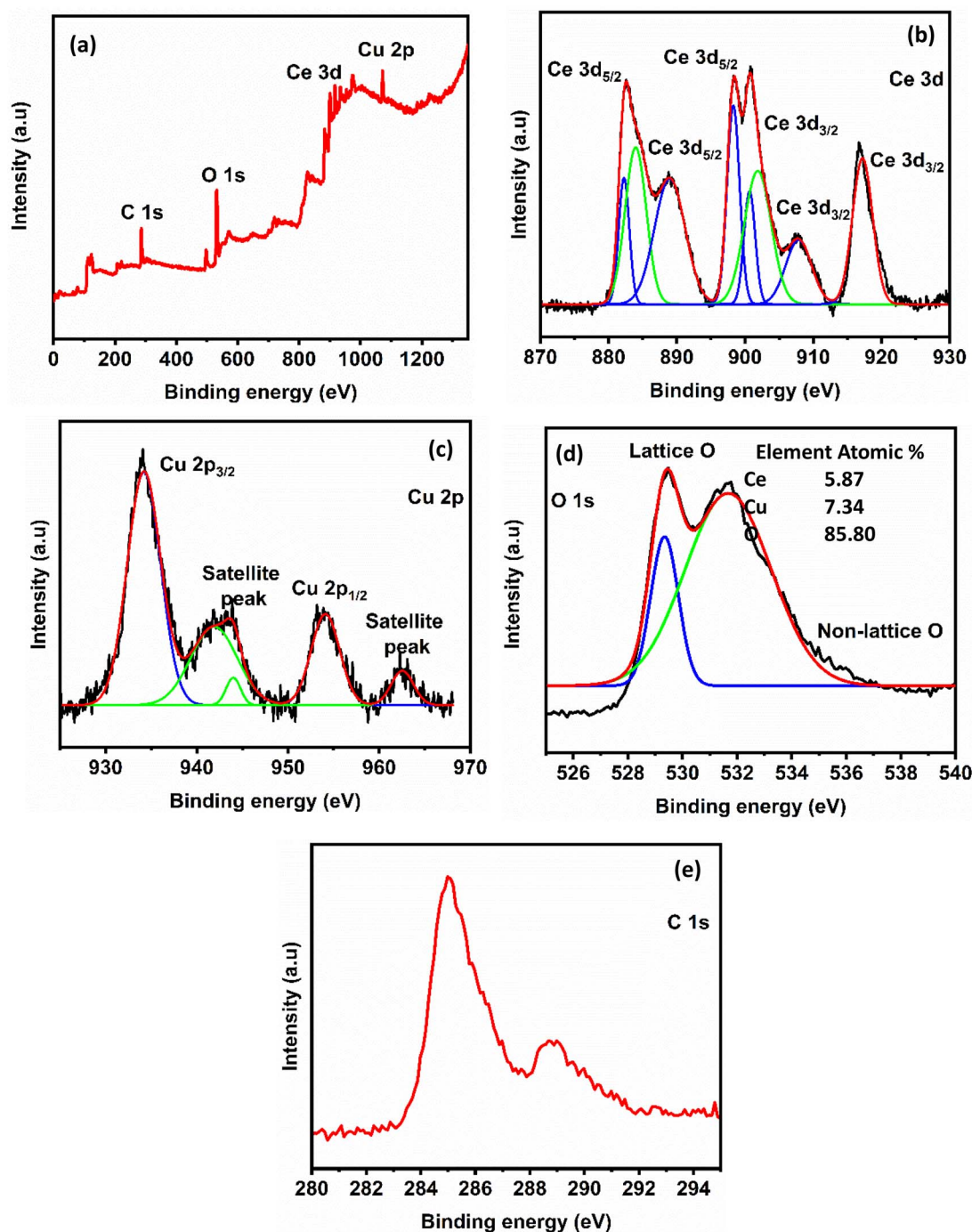
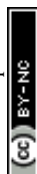


Fig. 5 (a) XPS spectrum of  $\text{CeO}_2\text{-CuO}$  nanocomposite, (b) high resolution spectrum of Ce 3d, (c) high resolution spectrum of Cu 2p, (d) high resolution spectrum of O 1s and (e) high resolution spectrum of C 1s.

$\text{Ce}^{4+}$  oxidation states. The six distinct peaks with binding energy values at 881.8, 888.0, 898.0, 900.4, 906.4, and 916.2 eV are associated with the  $\text{Ce}^{4+}$  3d state, confirming the predominant valence state of Ce as +4. The two weak XPS peaks at the binding energy values of 883.1 and 901.9 eV correspond to the  $\text{Ce}^{3+}$  oxidation state.<sup>42,43</sup> The observed yellowish colour of  $\text{CeO}_2$  supports the presence of both  $\text{Ce}^{4+}$  and  $\text{Ce}^{3+}$  ions in the structure.<sup>44</sup> In the  $\text{CeO}_2\text{-CuO}$  composite, two significant peaks appeared at 933.9 eV and 953.7 eV: these values are the binding

energies of  $\text{Cu } 2p_{3/2}$  and  $\text{Cu } 2p_{1/2}$ , respectively (Fig. 5c). The three satellite peaks in the XPS spectrum of  $\text{Cu}^{2+}$  at binding energies of 941.0 eV, 943.6 eV and 962.3 eV indicate a partially filled d-block configuration ( $3d^9$ ) for  $\text{Cu}^{2+}$  in the ground state.<sup>45,46</sup> The high-resolution XPS spectra of O 1s of both  $\text{CeO}_2$  and  $\text{CeO}_2\text{-CuO}$  composite revealed two peaks: the first peak around 528.8 eV originated from lattice oxygen and the peak of higher binding energy (531.1 eV) is attributed to physically adsorbed oxygen and/or a hydroxyl group. In both of the cases,





the C 1s peaks correspond to the carbonate impurities present in the compounds.

**3.2.3 Thermal properties.** From the TGA curves shown in Fig. S4 (ESI),<sup>†</sup> it can be seen that both of the compounds undergo single-stage weight loss up to 700 °C. An initial weight loss of 5% in the case of CeO<sub>2</sub> and 8% in the case of CeO<sub>2</sub>-CuO composites observed below 100 °C is primarily due to the evaporation of trapped moisture.<sup>35</sup> The indistinguishably observed meagre and gradual weight loss of 4% in CeO<sub>2</sub> and 3% in the CeO<sub>2</sub>-CuO composite at temperatures ranging from 100 °C to 500 °C is attributed to the release of carbon dioxide adsorbed from the atmosphere and decomposition of the residual nitrate present in the compounds. FT-IR spectral analyses confirm this observation. The endothermic peak detected below 100 °C in the DTA curves of both of the compounds is indicative of water evaporation. In the absence of any other transitions in the DTA curves, these results together with the TGA results confirm that these compounds are stable up to the studied temperature of 700 °C.

### 3.3 Characterization of the coatings

**3.3.1 Bonding characteristics, crystallinity and phase composition.** FT-IR spectra of the unmodified zinc phosphate coating and representative spectrum of coatings modified with nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite are given in Fig. S5 (ESI).<sup>†</sup> They showed characteristic absorption bands for different functional groups present in the coatings. The bands appearing at around 3165 cm<sup>-1</sup> and 1620 cm<sup>-1</sup> are respectively attributed to O-H stretching and H-O-H bending vibrations of water molecules. These broad bands are indicative of the hydrated nature of phosphate crystals formed. The bands observed at 1126 cm<sup>-1</sup> and 1029 cm<sup>-1</sup> are associated with the asymmetric stretching vibrations of HPO<sub>4</sub><sup>2-</sup> and PO<sub>4</sub><sup>3-</sup> groups, respectively. The peak at 922 cm<sup>-1</sup> arises from ν<sub>1</sub> P-O symmetric stretching of HPO<sub>4</sub><sup>2-</sup>. Lastly, the absorption at 570 cm<sup>-1</sup> corresponds to O-P-O bending vibrations. These observations are consistent with previous work and confirm the formation of a protective layer of phosphate on the metal surface.<sup>24,47</sup>

The crystallinity and phase composition of the coatings were studied using XRD and the obtained diffraction patterns are shown in Fig. 6. There are two crystal phases in all the zinc phosphate coatings: the first one is hopeite (Zn<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>·4H<sub>2</sub>O, JCPDS file 37-0465), which is a major phase and the other one is phosphophyllite (Zn<sub>2</sub>Fe(PO<sub>4</sub>)<sub>2</sub>·4H<sub>2</sub>O, JCPDS file 29-1427), which is a minor phase.<sup>5</sup> After comparing the XRD patterns of unmodified zinc phosphate coating and those modified with nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite, it was evident that the addition of both of the additives resulted in stronger (high intensity) reflections for the hopeite phase and this was more significant in the case of coatings formulated with CeO<sub>2</sub>-CuO nanocomposite. The strong reflection observed at 19.5° correspond to the (020) plane of hopeite crystals.<sup>22</sup> This increase in intensity indicates that the additives play a vital role in the formation of the phosphate coating layer. The diffraction peak for the phosphophyllite phase appeared at 31.5°, but was weak in the case of coating containing CeO<sub>2</sub> and very weak in

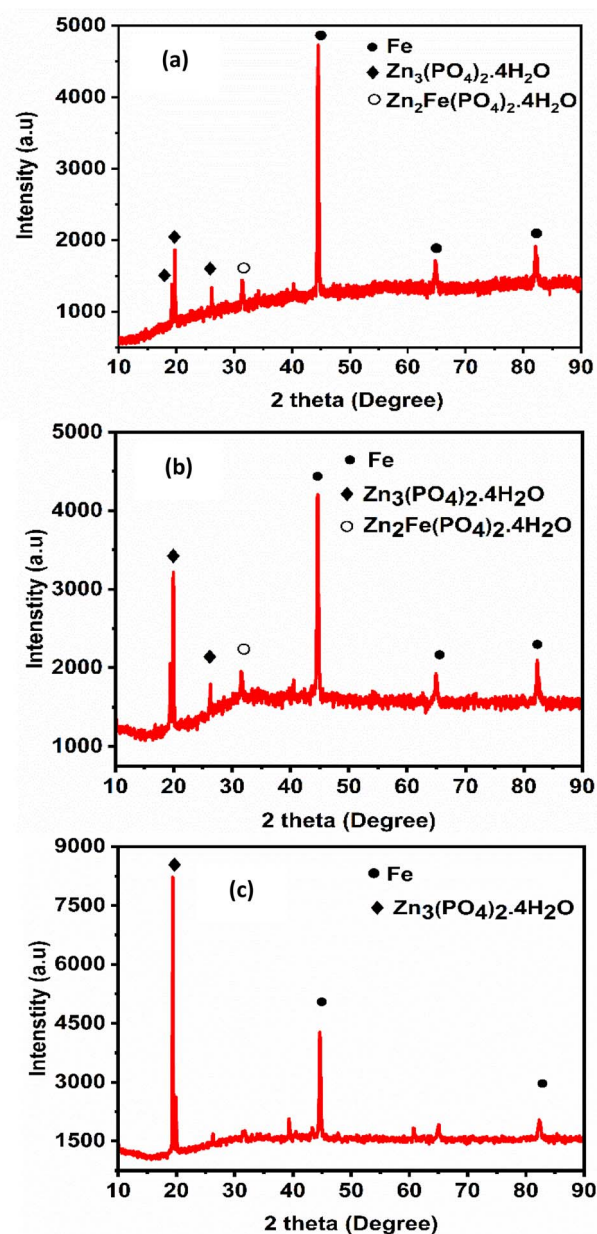


Fig. 6 XRD pattern of zinc phosphate coating corresponding to (a) ZnP, (b) ZnP-0.6CO and (c) ZnP-0.3COCO.

the CeO<sub>2</sub>-CuO modified coatings. Further, the XRD patterns reveal a considerable reduction in the intensity of iron upon the addition of nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite, suggesting improved surface coverage and enhanced growth of hopeite.<sup>48</sup> The absence of shifts in the 2 theta values of the major components confirms that the addition of these additives did not affect the phase composition of phosphate crystals while influencing its growth. The sizes of the phosphate crystallites were 32.9 ± 0.4 nm, 19.3 ± 0.8 nm and 18.1 ± 0.4 nm for ZnP, ZnP-0.6CO and ZnP-0.3COCO, respectively. These results together with the SEM micrograph showed that both of the nanometal oxides enhanced the nucleation sites, reduced the crystal size and resulted in coatings with more phosphate crystals.



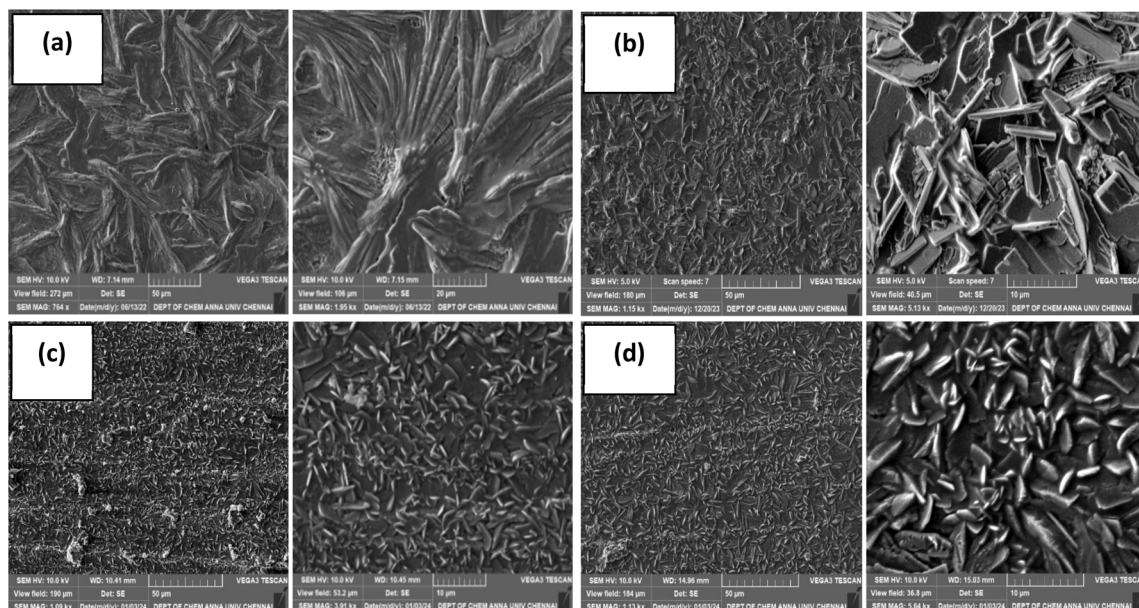


Fig. 7 Surface morphology of zinc phosphate coating formulated with different amounts of nanocrystalline CeO<sub>2</sub>. (a) ZnP; (b) ZnP-0.3CO; (c) ZnP-0.6CO; and (d) ZnP-0.9CO (right side images are magnified, scale bar = 10 μm).

**3.3.2 Morphology and elemental compositions of the coatings.** Fig. 7 and 8 illustrates the surface morphology of the zinc phosphate coatings modified with different amounts (0.3–0.9 g) of nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>–CuO nanocomposite. FE-SEM images of the coating without these additives reveal that the coating is uneven and the formed hopeite crystals are oriented and large in size. The images also confirmed the presence of numerous cracks in the coatings, which were caused by the stress developed due to hydrogen evolution during the crystallisation process.<sup>22,49</sup> These cracks allow the

electrolyte to penetrate the coating and potentially lead to corrosion. In contrast to the unmodified coating, the zinc phosphate coating modified with nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>–CuO nanocomposites exhibits distinct changes in morphology and crystal structure; the formed crystals are dense, well defined and randomly oriented across the surface. The hopeite crystals formed in the coating formulated with CeO<sub>2</sub> show plate-like structures, whereas some of these crystals formed with the addition of CeO<sub>2</sub>–CuO nanocomposites have a cubic-prismatic morphology, indicating unique crystallisation

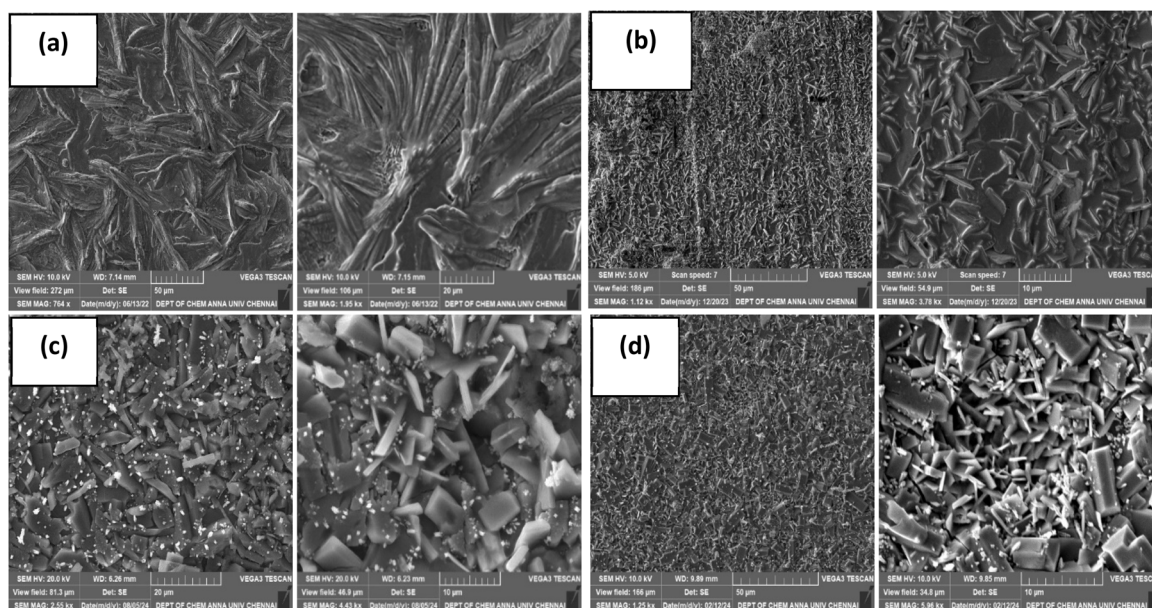


Fig. 8 Surface morphology of zinc phosphate coating formulated with different amounts of CeO<sub>2</sub>–CuO nanocomposite. (a) ZnP; (b) ZnP-0.3COCO; (c) ZnP-0.6COCO; and (d) ZnP-0.9COCO (right side images are magnified, scale bar = 10 μm).



**Table 2** EDAX-elemental composition of zinc phosphate coating formulated with different amounts of nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite

Coating ID	Nanoadditive added	Elemental composition (in wt%)					
		Zn	Fe	P	O	Ce	Cu
ZnP	Nil	21.30	15.40	24.16	38.00	—	—
ZnP-0.3CO	Nano CeO <sub>2</sub>	34.16	11.99	18.33	30.31	3.53	—
ZnP-0.6CO	Nano CeO <sub>2</sub>	36.29	3.18	20.87	31.63	5.61	—
ZnP-0.9CO	Nano CeO <sub>2</sub>	34.85	8.64	19.84	33.04	2.44	—
ZnP-0.3COCO	CeO <sub>2</sub> -CuO nanocomposite	39.75	1.39	18.37	24.70	1.85	12.53
ZnP-0.6COCO	CeO <sub>2</sub> -CuO nanocomposite	39.06	2.98	19.33	31.97	3.05	3.00
ZnP-0.9COCO	CeO <sub>2</sub> -CuO nanocomposite	39.40	1.64	19.69	28.33	4.76	6.21

behaviour. These crystals grow perpendicularly from the steel surface and differ in sizes.<sup>10</sup> Among the three concentration variations in two additives, ZnP-0.6CO and ZnP-0.3COCO produced uniform coatings with relatively smaller crystallites. It can be concluded from the explicit FE-SEM pictures that these nanometal oxide additives act as nucleation sites and produce more compact and dense coating.<sup>50</sup> The SEM micrographs also evidenced that both the additives reduced the cracks that were seen in the unmodified zinc phosphate coating.

The EDAX mapping spectra of all the zinc phosphate coatings are presented in Fig. S6 and S7 (ESI)<sup>†</sup> and the elemental composition data obtained thereof are given in Table 2. All the spectra confirm the presence of zinc, phosphorus, iron and oxygen present in all the coatings, whereas the spectra of coatings modified with CeO<sub>2</sub> and CeO<sub>2</sub>-CuO show the presence of cerium and copper in addition to all those elements. This indicates the successful incorporation of the additives into the zinc phosphate matrix which is essential for enhancing the coating's protective properties.

The zinc and iron content play a crucial role in determining the overall performance of the coating. As shown in Table 2, the zinc content is higher in the modified coatings than in the unmodified coatings, indicating a greater precipitation of zinc phosphate crystals. Another key observation is the low iron content in the modified coating; iron comes from the uncovered metal surface due to the porous nature of the coatings.<sup>25</sup> The low iron content of the modified zinc phosphate coating compared with the unmodified coating suggests that the nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite improves the surface coverage and integrity of the coatings.<sup>24</sup> Within the

variations in the additives and their concentrations, the increase in zinc content and decrease in iron content is more pronounced in the coatings formulated with 0.6 g L<sup>-1</sup> of CeO<sub>2</sub> and 0.3 g L<sup>-1</sup> of CeO<sub>2</sub>-CuO nanocomposites.

**3.3.3 Coating weight and hardness.** The additive effect of synthesized nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO composite on these properties of coatings was studied and the results were listed in Table 3. The data clearly indicate an increase in the coating weight against the addition of nanometal oxide additives. Upon adding the minimum quantity (0.3 g L<sup>-1</sup>) of nano CeO<sub>2</sub> to the phosphating bath, the coating weight increased from 4.8 ± 0.1 g m<sup>-2</sup> to 13.8 ± 0.2 g m<sup>-2</sup>, although the coating weight decreased upon further addition of nano CeO<sub>2</sub>. On the other hand, in the case of coatings based on CeO<sub>2</sub>-CuO nanocomposite, the coating weight successively increased up to 20.1 ± 0.8 g m<sup>-2</sup> for the 0.6 g L<sup>-1</sup> additive and then slightly decreased. This increase in coating weight against metal oxide additives can be attributed to a larger number of cathodic sites that facilitates the formation of phosphate crystals.<sup>18,51</sup> This can be further claimed with the formation of additional nucleation sites, which leads to a denser and more compact coating structure.<sup>11</sup>

The micro hardness of all the coatings was determined and the values were included in Table 3. In both of the series of coatings, the hardness increased with greater concentrations of nanoadditives. The highest hardness was achieved for the coating formulated with 0.9 g L<sup>-1</sup> of CeO<sub>2</sub>-CuO (173.8 ± 5.8 HV), which is 20.1 HV higher than that of unmodified coating. This improved hardness of the zinc phosphate coating modified with nano CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite indicate that

**Table 3** Weight and hardness of zinc phosphate coating formulated with different amounts of nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite

Coating ID	Nanoadditive added	Coating weight (g m <sup>-2</sup> )	Micro hardness (HV)
ZnP	Nil	4.8 ± 0.1	153.7 ± 2.9
ZnP-0.3CO	Nano CeO <sub>2</sub>	13.8 ± 0.2	159.9 ± 1.7
ZnP-0.6CO	Nano CeO <sub>2</sub>	12.0 ± 0.4	166.2 ± 5.9
ZnP-0.9CO	Nano CeO <sub>2</sub>	10.8 ± 0.2	169.0 ± 3.7
ZnP-0.3COCO	CeO <sub>2</sub> -CuO nanocomposite	17.4 ± 0.5	158.0 ± 3.1
ZnP-0.6COCO	CeO <sub>2</sub> -CuO nanocomposite	20.1 ± 0.8	166.6 ± 6.4
ZnP-0.9COCO	CeO <sub>2</sub> -CuO nanocomposite	19.4 ± 0.4	173.8 ± 5.8



significant changes occurred in the microstructure of hopeite crystals.

**3.3.4 Wettability and adhesion.** For a coating to combat corrosion effectively, it should repel water or any other corrosive agents. Therefore, it should be hydrophobic or neutral. The water contact angles of the coatings formulated for the present work are shown in Fig. S8 (ESI).† The zinc phosphate coating without metal oxide additives showed a hydrophilic character with a contact angle of 39.9°. The addition of nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite significantly improved the contact angle towards a neutral value: a maximum contact angle of 62.9° was obtained for the addition of 0.3 g L<sup>-1</sup> of CeO<sub>2</sub>. Upon increasing the concentration, the contact angle of the coating decreased. The opposite trend was observed in the case of CeO<sub>2</sub>-CuO based coatings.

Adhesion of a coating with a metal surface plays a crucial role in the coating's protective performance. In corrosive environments, the penetration of corrosive agents into the metal-coating interface weakens the coating's adhesion. The cross-hatch adhesion test images of coated steel substrates were recorded (Fig. S9 and S10, ESI†) and the coating's adhesion strength was assessed according to the ASTM D3359-17

standard. The coating without the additive and the coatings modified with CeO<sub>2</sub> demonstrated good adhesion with no detachment or damage at the grid edges and no squares were eliminated in these coatings during removal of the adhesive tape. Thus, these coatings were rated 5B. On the other hand, the coatings formulated with CeO<sub>2</sub>-CuO composite showed trace removal of the coating within the square grids. However, the edges remained intact and it was rated 4B, indicating less than 5% removal. These adhesion test results suggest that the added nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite did not interfere with the interactions and bonding force acting between the coating and the steel substrate.

**3.3.5 Potentiodynamic polarisation studies.** The corrosion resistance of zinc phosphate coatings modified with nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite studied by analysing the respective potentiodynamic polarisation curves was depicted in Fig. 9. The corrosion parameters such as corrosion current density ( $I_{\text{corr}}$ ), corrosion potential ( $E_{\text{corr}}$ ), polarisation resistance ( $R_p$ ), cathodic ( $\beta_c$ ) and anodic ( $\beta_a$ ) Tafel slopes and the corrosion rate (CR) calculated using the Tafel extrapolation method were given in Table 4. It is evident from the data that the  $I_{\text{corr}}$  values of all the modified zinc phosphate

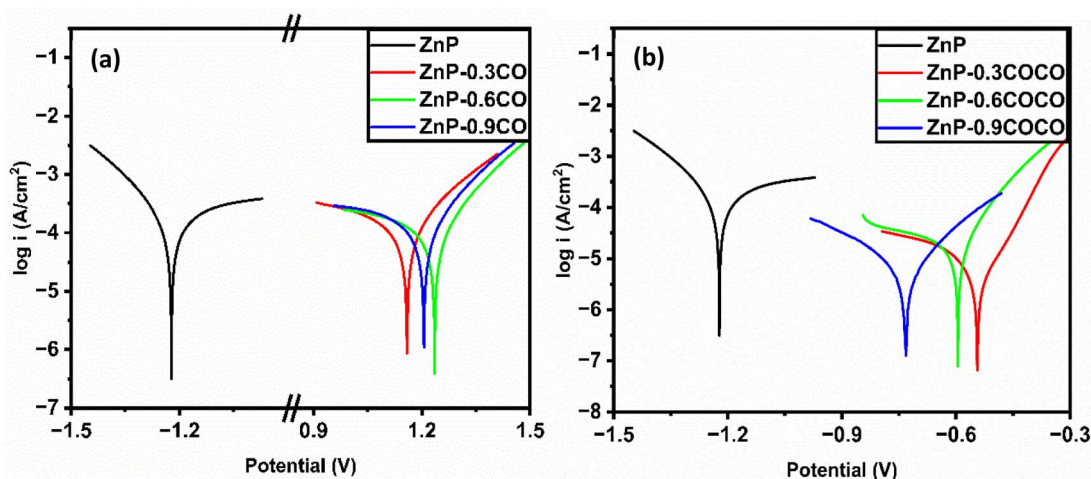


Fig. 9 Potentiodynamic polarisation curves of the uncoated and zinc phosphate coated steel substrates immersed in 3.5% NaCl. Coatings were formulated with different amounts of (a) nanocrystalline CeO<sub>2</sub> and (b) CeO<sub>2</sub>-CuO nanocomposite.

Table 4 Electrochemical parameters obtained by Tafel extrapolation of polarisation curves of zinc phosphate coatings formulated with different amounts of nanocrystalline CeO<sub>2</sub> and CeO<sub>2</sub>-CuO nanocomposite. Coated mild steel panels were immersed in 3.5% NaCl solution

Electrochemical parameters							
Coating ID	Nano additive added	$I_{\text{corr}}$ ( $\mu\text{A cm}^{-2}$ )	$E_{\text{corr}}$ (V)	$\beta_a$ (V dec <sup>-1</sup> )	$-\beta_c$ (V dec <sup>-1</sup> )	$R_p$ (ohm cm <sup>2</sup> )	Corrosion rate (CR) (mm per year)
ZnP	Nil	156.2	1.22	0.500	0.167	348	1.843
ZnP-0.3CO	Nano CeO <sub>2</sub>	87.44	1.16	0.155	0.333	525	1.032
ZnP-0.6CO	Nano CeO <sub>2</sub>	44.80	1.24	0.191	0.271	1086	0.528
ZnP-0.9CO	Nano CeO <sub>2</sub>	65.10	1.21	0.167	0.258	676	0.768
ZnP-0.3COCO	CeO <sub>2</sub> -CuO nanocomposite	3.025	-0.54	0.087	0.111	7003	0.036
ZnP-0.6COCO	CeO <sub>2</sub> -CuO nanocomposite	17.50	-0.59	0.106	0.500	2169	0.207
ZnP-0.9COCO	CeO <sub>2</sub> -CuO nanocomposite	3.98	-0.73	0.099	0.139	6321	0.047





coatings are low compared with that of unmodified coating. Furthermore, the proportional decrease in the corrosion rate in the modified coatings suggests better corrosion protection. The slightly high  $E_{\text{corr}}$  value of coating modified with  $0.6 \text{ g L}^{-1}$  of  $\text{CeO}_2$  compared with the unmodified coating indicates better corrosion resistance of this coating. Although the  $E_{\text{corr}}$  values of all the  $\text{CeO}_2$ -CuO composite modified coatings shifted to a negative direction, these parameters are best evaluated through  $R_p$  and CR. The  $R_p$  and CR of the coating without the additive are  $348 \Omega \text{ cm}^2$  and  $1.843 \text{ mm per year}$ , respectively. Upon the addition of nanocrystalline  $\text{CeO}_2$  and  $\text{CeO}_2$ -CuO nanocomposite to the coating recipe, both of these parameters improved significantly. In the case of coatings with  $\text{CeO}_2$ , higher  $R_p$  ( $1086 \Omega \text{ cm}^2$ ) and lower CR ( $0.528 \text{ mm per year}$ ) were observed at the additive concentration of  $0.6 \text{ g L}^{-1}$ . In comparison, coatings with the  $\text{CeO}_2$ -CuO composite showed more significant improvement even at a low additive concentration of  $0.3 \text{ g L}^{-1}$ . These results reveal that both nanocrystalline  $\text{CeO}_2$  and  $\text{CeO}_2$ -CuO nanocomposite increase the ability of the zinc phosphate coating to retard the electrochemical reaction occurring at the electrode-electrolyte interface and thereby increase the corrosion resistance, especially the  $\text{CeO}_2$ -CuO nanocomposite. This improved corrosion resistance of all the

modified coatings perceived from the polarisation studies is mainly due to formation of dense and compact coatings.

**3.3.6 Electrochemical impedance spectroscopy (EIS).** To further assess the protective properties of modified zinc phosphate coatings, EIS was employed and the Nyquist plots, Bode phase angles and Bode magnitude plots of the coatings formulated with nanocrystalline  $\text{CeO}_2$  and  $\text{CeO}_2$ -CuO nanocomposite are displayed in Fig. 10 and 11, respectively. The high frequency region of the Nyquist plot revealed the barrier properties of the coating, while the low-frequency regions corresponded to the diffusion-controlled corrosion reaction taking place at the coating and steel surface interface.<sup>52</sup> The Nyquist plot of the coating without metal oxide additive shows two capacitive loops, whereas that of each  $\text{CeO}_2$  modified coating has one additional inductive loop. The first is the large capacitive loop at high frequency, followed by a mid-frequency inductive loop and finally a small low frequency capacitive loop.<sup>53</sup> The shape of the Nyquist plots remains the same for all three concentrations of  $\text{CeO}_2$ . The high frequency capacitive loop is attributed to the zinc phosphate coating and the inductive loop is related with the relaxation of adsorbed species like  $\text{Cl}^-$  and  $\text{O}^{2-}$  on the surface of the coating. The capacitive loop at the low frequency region is due to the charge transfer reaction.<sup>31,54</sup> In the case of  $\text{CeO}_2$ -CuO modified coatings, they

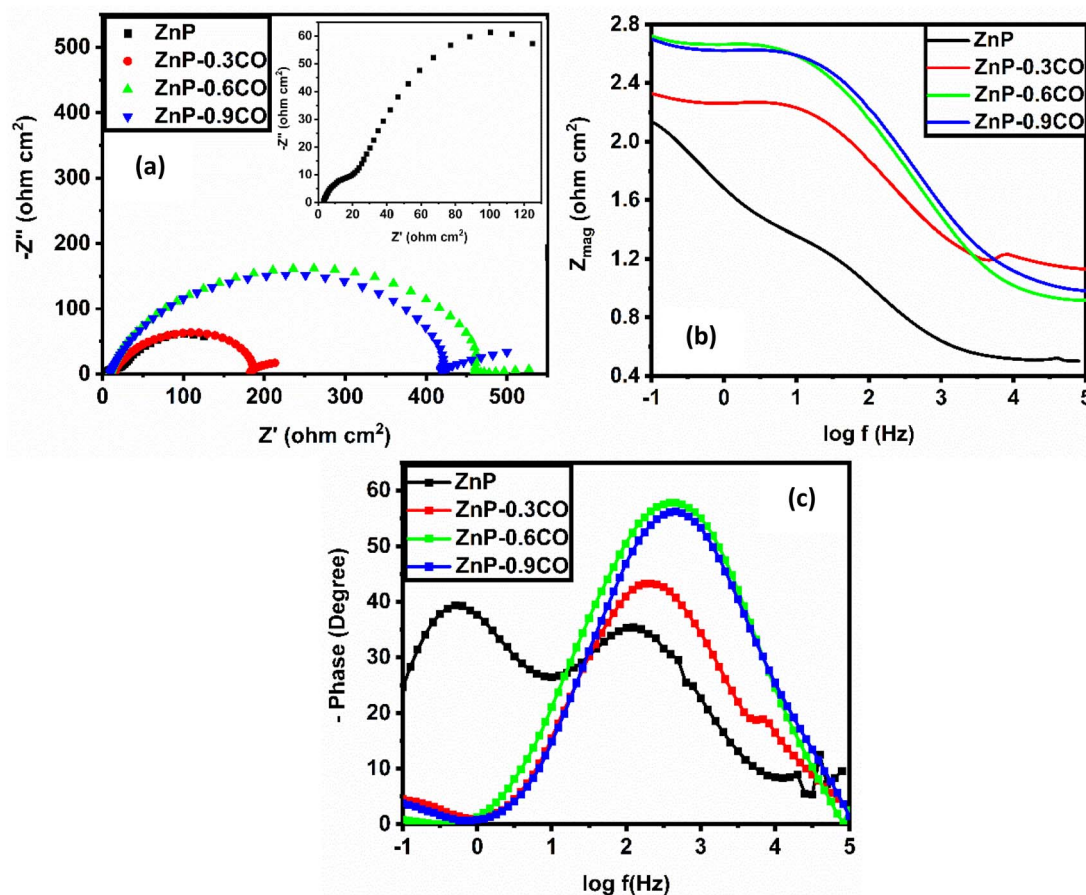


Fig. 10 (a) Nyquist plots, (b) Bode plots and (c) frequency dependence of phase angles vs. electrochemical impedance curves of the uncoated and zinc phosphate coated steel substrates immersed in 3.5% NaCl. Coatings were formulated with different amounts of nanocrystalline  $\text{CeO}_2$ .



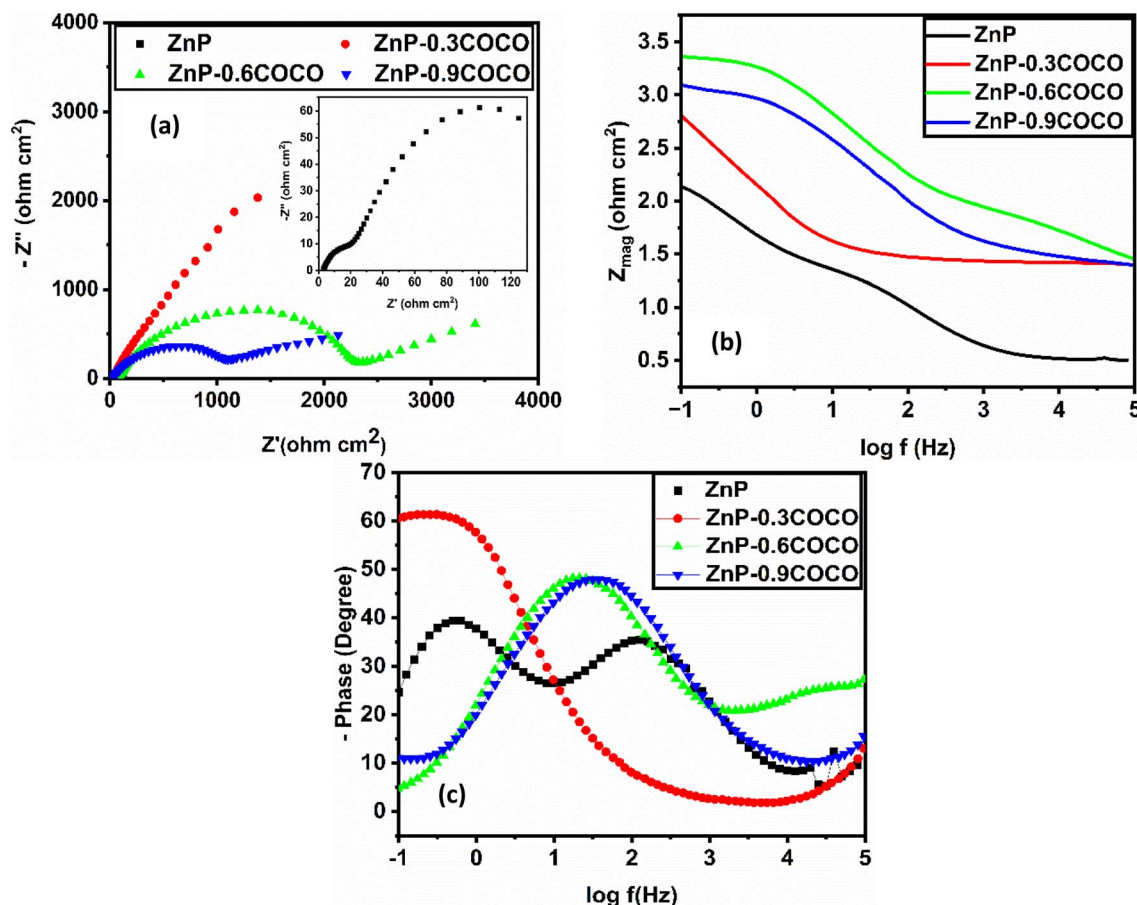


Fig. 11 (a) Nyquist plots, (b) Bode plots and (c) frequency dependence of phase angles vs. electrochemical impedance curves of the uncoated and zinc phosphate coated steel substrates immersed in 3.5% NaCl. Coatings were formulated with different amounts of  $\text{CeO}_2$ -CuO nanocomposite.

have a high frequency capacitive loop followed by low frequency Warburg impedance associated with the diffusion process.

In general, the large diameter of the capacitive loop indicates enhanced corrosion resistance as they reflect greater impedance to ion penetration.<sup>55</sup> For nanocrystalline  $\text{CeO}_2$  modified coatings, the diameter of the first capacitive loop increases successively from that without additive to the inclusion of  $0.3 \text{ g L}^{-1}$  and  $0.6 \text{ g L}^{-1}$  additive and then decreased. The Bode impedance at the low frequency region corresponds to the barrier properties of the coatings. The higher the impedance modulus, the greater the blocking of corrosive ions by the coating.<sup>22</sup> From Fig. 10b and 11b, it is clear that the impedance modulus

increased upon the addition of both  $\text{CeO}_2$  and  $\text{CeO}_2$ -CuO into the coatings. Similarly, these additives increased the phase angle of the coating up to  $60^\circ$ . This indicates that the stability of the modified coatings improved in this impedance measurement condition compared with the unmodified coating.

The impedance curves were fitted to equivalent electrical circuits using ZSimpWin software for further analysis and the generated circuits for unmodified and modified zinc phosphate coatings were given in Fig. 12. From the equivalent circuits, electrochemical parameters including  $R_s$  (solution resistance),  $R_c$  (coating resistance),  $R_{ct}$  (charge transfer resistance),  $C_c$  and  $Q_c$  (coating capacitance),  $C_{dl}$  and  $Q_{dl}$  (double layer capacitance)

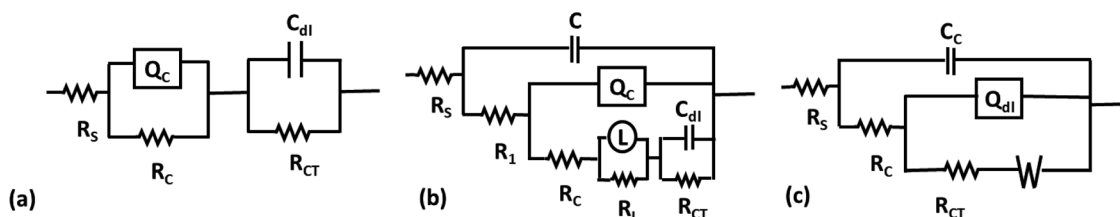


Fig. 12 Equivalent electrical circuit used to obtain the electrochemical parameters from the Nyquist plots of coatings in 3.5% NaCl. (a) ZnP; (b) ZnP-CO (common for three variations) and (c) ZnP-COCO (common for three variations). Fitted curves are given in Fig. S11, ESI.†



and  $W$  (Warburg impedance) were obtained and listed in Table 5. In the circuit,  $L$  and  $R_1$  are inductive elements. The  $R_s$  of all the modified coatings and  $R_c$  of all the coatings except the one prepared with  $0.9 \text{ g L}^{-1}$  of  $\text{CeO}_2\text{-CuO}$  nanocomposite are improved by the addition of metal oxide additives into the coating composition. The low  $Q_c$  value of coating containing nanocrystalline  $\text{CeO}_2$  indicates a restriction of the penetration of ions, water molecules and oxygen through the coating. This also means that this additive blocks the pores and reduces the porosity of the coating. Importantly, the  $R_{ct}$  component of all the modified coatings are higher than that of the unmodified coating, which indicates that the added nanocrystalline  $\text{CeO}_2$  and  $\text{CeO}_2\text{-CuO}$  nanocomposite successfully protects the substrate. In the case of  $\text{CeO}_2$  modified coatings, a high  $R_{ct}$  ( $427.5 \text{ } \Omega \text{ cm}^2$ ) was obtained for coating formulated with  $0.6 \text{ g L}^{-1}$  of this additive and a remarkable enhancement in this parameter was found in the case of coating added with  $0.3 \text{ g L}^{-1}$  of  $\text{CeO}_2\text{-CuO}$  nanocomposite for which the  $R_{ct}$  value ( $3645 \text{ } \Omega \text{ cm}^2$ ) was more than thirty-one times higher than the unmodified coating. This parameter represents the resistance towards charge transfer reactions occurring on the electrode. The high  $R_{ct}$  indicates that the defects in the coating are reduced due to refinement of hopeite crystals.<sup>25,56</sup> The presence of Warburg elements in the  $\text{CeO}_2\text{-CuO}$  nanocomposite added coatings shows that the corrosion reaction is diffusion controlled. From these EIS studies, it can be concluded that  $0.3 \text{ g L}^{-1}$  addition of  $\text{CeO}_2\text{-CuO}$  nanocomposite forms more corrosion resistant coating.

**3.3.7 Salt spray test.** A salt spray test is an accelerated test method used to assess the corrosion resistance of a coating in a simulated experimental condition. All the coatings modified with metal oxide additives were subjected to this standard test and the obtained results were compared with that of the

unmodified coating. The recorded photographs after the completion of the tests are displayed in Fig. 13 and the corrosion rate determined from the weight loss data are given in Table 6 along with additive's inhibition efficiency. The observed corrosion products (rust) that formed across the coated surface in the images indicate that the zinc phosphate coating prepared without metal oxide additive has poor corrosion resistance. The introduction of nanocrystalline  $\text{CeO}_2$  into the coating significantly reduced the degree of corrosion throughout the coating surface and lowered the corrosion rate at all concentrations used. Analysis of the additives concentration effect showed that  $\text{ZnP-0.6CO}$  had the lowest CR with the highest inhibition efficiency. This finding aligns with a previous study that describes the potential of using amorphous  $\text{CeO}_2$  of particle size  $\approx 60 \text{ nm}$  as a corrosion inhibitor.<sup>57</sup> When introducing  $\text{CeO}_2\text{-CuO}$  nanocomposite into the coating, a different effect was observed in the corrosion behaviour. At the minimum concentration ( $0.3 \text{ g L}^{-1}$ ), the coating was less corroded compared with the coating without the additive and this observation was consistent with the electrochemical characterisation results. Besides, this was supported by a previous study in which copper enhanced the corrosion protection of zinc phosphate coating on an aluminium surface.<sup>58</sup> The corrosion increased with greater concentrations of  $\text{CeO}_2\text{-CuO}$  nanocomposite. This observation contradicted with the potentiodynamic polarisation and electrochemical impedance results. The high corrosion observed in  $\text{ZnP-0.6COCO}$  and  $\text{ZnP-0.9COCO}$  indicates that structural changes occurred in these coatings after exposing them to a 5%  $\text{NaCl}$  solution for a longer time (compared with electrochemical characterisation) during the experiment. As previously discussed, the structural defects at the micro level allows the electrolyte to diffuse into the metal surface and initiates the corrosion. The reason for such structural defects may be due to

**Table 5** Fitting values of equivalent electrical circuits of zinc phosphate coatings formulated with different amounts of nanocrystalline  $\text{CeO}_2$  and  $\text{CeO}_2\text{-CuO}$  nanocomposite. Coated mild steel panels were immersed in 3.5%  $\text{NaCl}$  solution

Fitting values of equivalent electrical circuit								
Coating ID	Nano additive added	$R_s$ ( $\Omega \text{ cm}^2$ )	$R_c$ ( $\Omega \text{ cm}^2$ )	$R_{ct}$ ( $\Omega \text{ cm}^2$ )	$Q_c$		$C_{dl} \times 10^{-2}$ $\text{F cm}^{-2}$	$\chi^2 \times 10^{-3}$
					$Y \times 10^{-5}$ ( $\Omega^{-1} \text{ S}^n \text{ cm}^{-2}$ )	$n$		
ZnP	Nil	2.98	30.82	114.3	178.9	0.63	0.629	3.44
ZnP-0.3CO	Nano $\text{CeO}_2$	12.69	29.08	167.4	9.928	0.74	3.107	3.21
ZnP-0.6CO	Nano $\text{CeO}_2$	8.34	35.25	427.5	5.498	0.68	0.025	0.92
ZnP-0.9CO	Nano $\text{CeO}_2$	9.62	58.47	400.2	3.510	0.73	1.337	0.75

Fitting values of equivalent electrical circuit									
Coating ID	Nano additive added	$R_s$ ( $\Omega \text{ cm}^2$ )	$R_c$ ( $\Omega \text{ cm}^2$ )	$R_{ct}$ ( $\Omega \text{ cm}^2$ )	$Q_{dl}$			$W \times 10^{-3}$ ( $\Omega \text{ cm}^2$ )	$\chi^2 \times 10^{-3}$
					$C_c \times 10^{-5}$ $\text{F cm}^{-2}$	$Y \times 10^{-5}$ ( $\Omega^{-1} \text{ S}^n \text{ cm}^{-2}$ )	$n$		
ZnP-0.3COCO	$\text{CeO}_2\text{-CuO}$ nanocomposite	15.95	11.17	3645	0.10	204.8	0.73	0.53	1.40
ZnP-0.6COCO	$\text{CeO}_2\text{-CuO}$ nanocomposite	25.27	43.31	2424	0.23	7.38	0.69	6.44	6.19
ZnP-0.9COCO	$\text{CeO}_2\text{-CuO}$ nanocomposite	19.35	10.73	1179	0.22	12.75	0.69	4.47	3.67



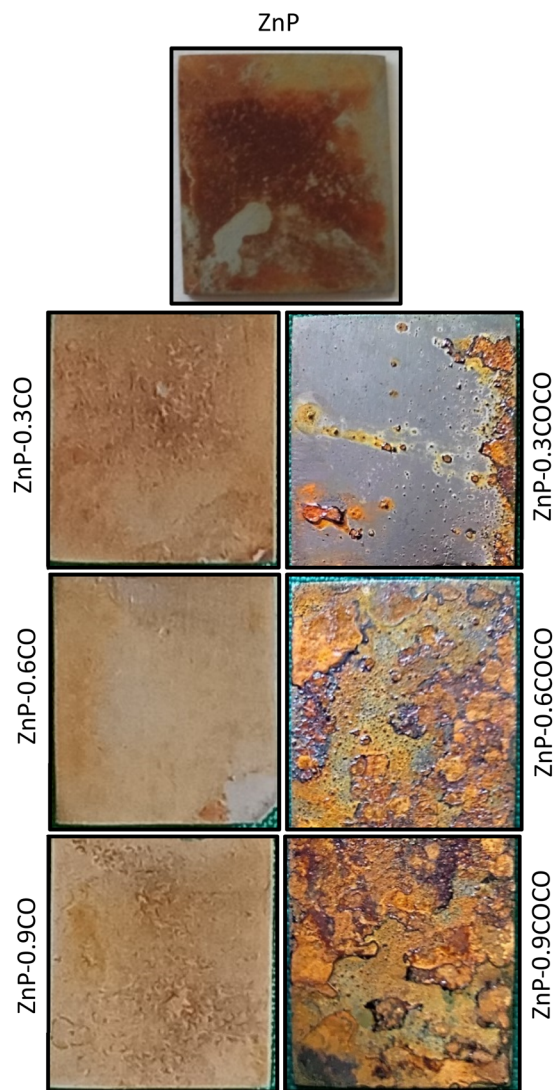
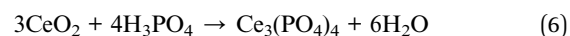


Fig. 13 Salt spray test images of coatings formulated without nano additive (top), with nanocrystalline  $\text{CeO}_2$  (left) and with  $\text{CeO}_2$ -CuO nanocomposite (right).

the formation of an association complex between more cathodic copper (compared with cerium) and chloride ions at high concentration, with subsequent dissolution of copper complex

into the salt solution. To verify this, ZnP-0.6COCO was subjected to a salt-spray test for 20 min and 20 h along with its counterpart ZnP-0.6CO and their surfaces were analysed using FE-SEM (Fig. 14). It can be clearly seen that the coating layer of ZnP-0.6CO shows less cracks, whereas that of ZnP-0.6COCO was not affected for 20 min and supports the obtained electrochemical characterisation results. However, the phosphate crystals are solvated in ZnP-0.6COCO exposed to 5% NaCl solution for 20 h, leaving more cracks in the coating surface and confirming the postulated reason.

**3.3.8 Mechanism of enhanced corrosion resistance of zinc phosphate coatings in the presence of nanocrystalline  $\text{CeO}_2$  and  $\text{CeO}_2$ -CuO nanocomposite.** The size of the formed phosphate crystallites decreased from  $32.9 \pm 0.4$  nm in the unmodified coating to  $19.3 \pm 0.8$  nm and  $18.1 \pm 0.4$  nm when adding nanocrystalline  $\text{CeO}_2$  and  $\text{CeO}_2$ -CuO nanocomposite at concentrations of  $0.6 \text{ g L}^{-1}$  and  $0.3 \text{ g L}^{-1}$ , respectively to the phosphate coating bath. At the same time, the intense reflections of crystalline  $\text{CeO}_2$  and CuO appearing at  $2\theta = 29.1$  and  $2\theta = 35.49$ , respectively (Fig. 3) were not observed in the XRD patterns of the coatings. This confirms the dissolution of both nano  $\text{CeO}_2$  and  $\text{CeO}_2$ -CuO nanocomposite into the excess phosphoric acid, forming their respective phosphates as shown below:



The characteristic reflections of these phosphates were not seen in the XRD patterns of the coat. The reason may be (i) the metal oxide additives used were in very low quantity ( $\approx 150$ – $600$  ppm) and (ii) these phosphates may be in the amorphous phase. Therefore, the cerium and copper phosphates that formed from their oxides have played the role of agents that nucleate the crystallisation and interfere with the crystallization of zinc phosphate into lengthy crystallites by suppressing the diffusion of nutrients for the growth of crystals. This can be supported with FE-SEM results of the coatings in which the formation of zinc phosphate as tiny crystals were evidently seen in large quantities compared with the coating without additive. Obviously, the formation of a dense, compact and uniform crystal layer enhanced the corrosion resistance of the coatings.

**Table 6** Corrosion rate and corrosion inhibition efficiency of the zinc phosphate coating formulated with different amounts of nanocrystalline  $\text{CeO}_2$  and  $\text{CeO}_2$ -CuO nanocomposite (data obtained from the salt spray test)

Coating ID	Nanoadditive added	Coating weight loss after 48 h ( $\times 10^{-3} \text{ g cm}^{-2}$ )	Corrosion rate (CR) ( $\times 10^{-3} \text{ mpy}$ )	Inhibition efficiency (%)
ZnP	Nil	1.32	1.86	No additive
ZnP-0.3CO	Nano $\text{CeO}_2$	1.20	1.70	8.6
ZnP-0.6CO	Nano $\text{CeO}_2$	0.81	1.14	38.7
ZnP-0.9CO	Nano $\text{CeO}_2$	1.02	1.44	22.5
ZnP-0.3COCO	$\text{CeO}_2$ -CuO nanocomposite	1.19	1.68	9.6
ZnP-0.6COCO	$\text{CeO}_2$ -CuO nanocomposite	2.36	3.33	Not calculated <sup>a</sup>
ZnP-0.9COCO	$\text{CeO}_2$ -CuO nanocomposite	2.58	3.67	

<sup>a</sup> Not calculated, since the corrosion rate of the inhibited systems were higher than that of uninhibited systems.





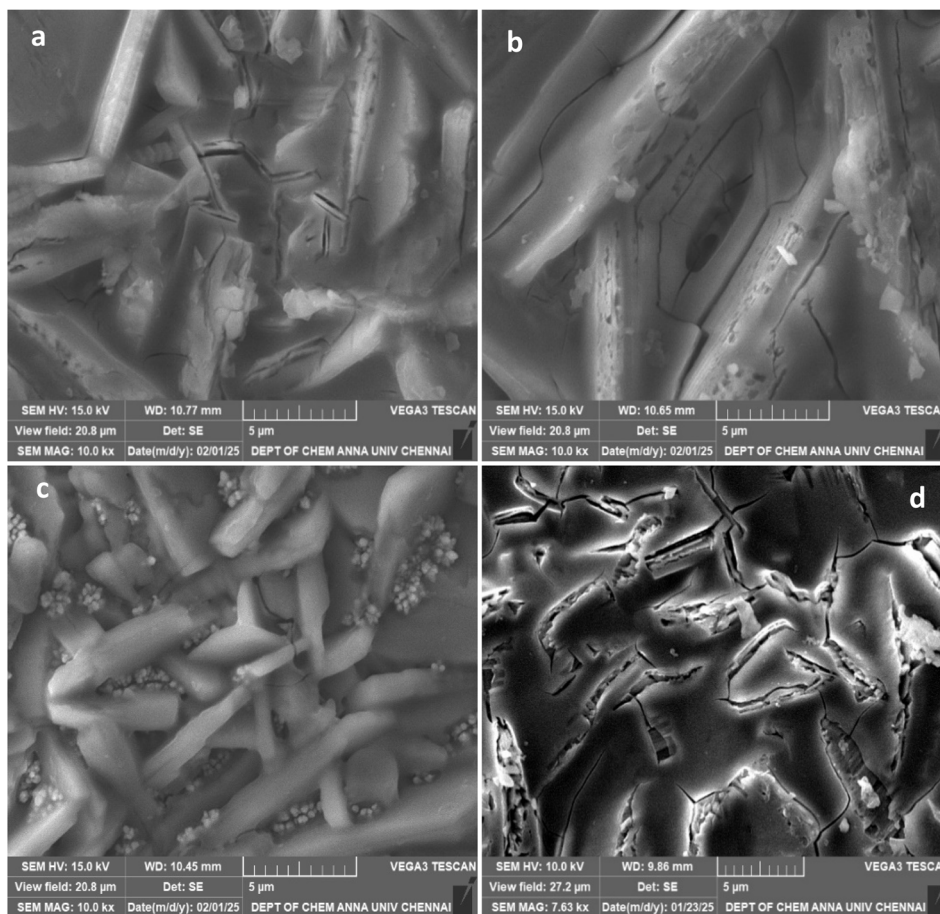


Fig. 14 FE-SEM micrographs of coatings exposed to the salt-spray test. (a and b) Are ZnP-0.6CO exposed to the salt spray test for 20 min and 20 h; (c and d) are ZnP-0.6COCO exposed to the salt spray test for 20 min and 20 h (scale bar for all images = 5 µm).

## 4. Conclusions

In this systematic study, nanocrystalline  $\text{CeO}_2$  and  $\text{CeO}_2\text{-CuO}$  nanocomposite were synthesized and their function as additives to enhance corrosion resistance of zinc phosphate coatings on mild steel substrates in 3.5% NaCl solution was evaluated. The XRD patterns of the coatings revealed that the addition of these compounds decreased the iron content (improved the surface coverage) and enhanced the growth of phosphate crystals without affecting the hopeite-phosphophyllite phase composition in the coatings. The XRD pattern and FE-SEM results revealed that these nanometal oxide additives act as nucleation sites and produced compact coating with more phosphate crystals. The SEM micrographs also revealed that both of these additives reduced the cracks that were seen in the coating prepared without the additive. Among the two metal oxides used in three different concentrations,  $0.6 \text{ g L}^{-1}$  and  $0.3 \text{ g L}^{-1}$  was found to be optimum for nanocrystalline  $\text{CeO}_2$  and  $\text{CeO}_2\text{-CuO}$  nanocomposite, respectively. It was found that (i) the iron content in the coating decreased from 15.4 wt% (ZnP) to 3.18 wt% upon the addition of  $\text{CeO}_2$  and this value was lower following the addition of  $\text{CeO}_2\text{-CuO}$  (1.39 wt%); (ii) the zinc content of the coating increased from 21.3 wt% (ZnP) to 36.29 wt% and 39.75 wt%, respectively following the addition of

$\text{CeO}_2$  and  $\text{CeO}_2\text{-CuO}$ ; and (iii) the corrosion rate decreased from 1.843 mm per year (ZnP) to 0.528 mm per year and 0.036 mm per year following the addition of nanocrystalline  $\text{CeO}_2$  and  $\text{CeO}_2\text{-CuO}$  nanocomposite, respectively. Based on the cross-hatch adhesion test, the coating without the additive and the coatings modified with  $\text{CeO}_2$  were rated as 5B and the coatings formulated with  $\text{CeO}_2\text{-CuO}$  composite were rated as 4B. From the results of all the coating characterizations including the accelerated salt-spray test, it was concluded that both of the nanocompounds act as potential corrosion inhibitors, except for the results of the salt-spray test carried out for coatings formulated with higher amounts of  $\text{CeO}_2\text{-CuO}$  nanocomposite, and all the results confirmed that the  $\text{CeO}_2\text{-CuO}$  nanocomposite showed more potential than nanocrystalline  $\text{CeO}_2$ .

## Data availability

All the data related to this research work have been included in this article and in the ESI section.†

## Author contributions

S. Ayesha Barsana: conceptualization, methodology, investigation, visualization, analysis and writing – original draft. M. J.



Umapathy: supervision, resources, investigation, methodology, data curation, reviewing and editing. A. Sultan Nasar: methodology, data curation and writing – original draft.

## Conflicts of interest

The authors declare no conflicts of interest.

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