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Controlled synthesis of copper sulfide-based catalysts for electrochemical reduction of CO₂ to formic acid and beyond: a review

Anirban Mukherjee, ^a Maryam Abdinejad, ^{*b} Susanta Sinha Mahapatra ^c and Bidhan Chandra Ruidas ^{*a}

Converting carbon dioxide (CO₂) into value-added chemicals is considered as a promising strategy to mitigate climate change. Among the various CO₂ reduction techniques, electrochemical CO₂ reduction (ECO₂R) using renewable energy sources holds significant potential. Consequently, the design and development of electrocatalysts capable of offering both high performance and cost-effectiveness hold the potential to expedite reaction kinetics and facilitate widespread industrial adoption. In recent years, abundant copper sulfide (Cu/S)-based nanomaterials among various metal–chalcogenides have attracted extensive research interest due to their semiconductivity and low toxicity, enabling them to be used in a wide range of applications in the ECO₂R field. This review highlights the progress in engineered Cu/S-based nanomaterials for ECO₂R reactions and elaborates on the correlations between engineering strategies, catalytic activity, and reaction pathways. This paper also summarises the controllable synthesis methods for fabricating various state-of-the-art Cu/S-based structures and outlines their possible implementation as electrocatalysts for CO₂ reduction. Finally, challenges and prospects are presented for the future development and practical applications of Cu/S-based catalysts for ECO₂R to value-added chemicals.

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1. Introduction

Since industrialization, the utilization of non-renewable energy sources, which encompass fossil fuels, *i.e.*, coal, crude oil, and natural gas, has tremendously increased.¹ Fossil fuels, as non-renewable energy sources, have two disadvantages: (1) limited supplies cannot meet increasing demands for energy, resulting

^a Department of Chemical Engineering, Birla Institute of Technology, Mesra, Ranchi 835215, India. E-mail: bidhanruidas@gmail.com^b Department of Chemical Engineering, Massachusetts Institute of Technology, 77 Massachusetts Avenue, Cambridge, Massachusetts 02139, USA^c Department of Chemistry, Birla Institute of Technology, Mesra, Ranchi 835215, India

Anirban Mukherjee

Mr Anirban Mukherjee completed his Master of Engineering at Jadavpur University, Kolkata. Currently, he is pursuing his PhD as an institute research fellow at the Birla Institute of Technology, Mesra (BIT Mesra), Ranchi 835215, India. His research mainly focuses on synthesizing and characterizing high-performance electrocatalysts for the electrochemical reduction of CO₂ into value-added products.



Maryam Abdinejad

Dr Maryam Abdinejad completed her PhD at the University of Toronto, followed by a postdoc at the Delft University of Technology. In 2022 she continued her research as an associate postdoc researcher at the Massachusetts University of Technology (MIT). Her research mainly focuses on the rational design and development of novel high-performance homogeneous and heterogeneous catalysts for the capture and electrochemical reduction of CO₂ to value-added materials.

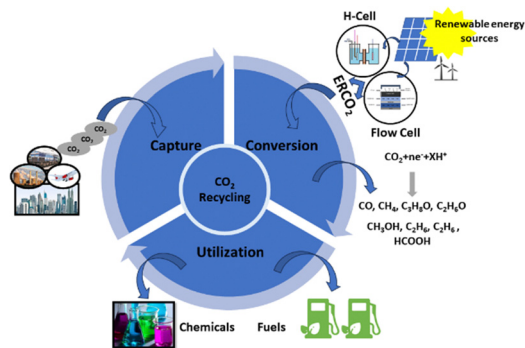


Fig. 1 A graphical illustration of CO₂ capture, utilization, and storage (CCUS).

in a severe energy crisis and (2) the consumption of fossil fuels for energy generation releases a substantial amount of CO₂ into the earth's atmosphere, contributing to global ecological issues such as global warming, global sea level rise, land degradation, and many more.^{1–3} To achieve a sustainable future, reducing CO₂ concentrations in the atmosphere is crucial.⁴ In this context, the scientific community has made numerous efforts to achieve this goal by using the CO₂ capture, utilization, and storage (CCUS) approach, as depicted in Fig. 1.^{5,6} The conversion of CO₂ into C₁ and highly dense C₂ products through chemical routes has been considered a promising technology for reutilizing CO₂.^{7,8} It can be achieved by several methods, such as photochemical,⁹ electrochemical,^{10,11} photo-electrochemical,¹² and biochemical methods.¹³ Among these methods, electrochemical CO₂ reduction (ECO₂R) has gained a lot of attention owing to its (a) controllable process parameters (*i.e.*, potential and temperature), (b) feasibility with the reaction environment (*e.g.*, organic and aqueous electrolytes), and (c) ability to scale up.^{14,15} Additionally, ECO₂R can be conducted using renewable energy resources such as solar and wind as power sources, allowing for renewable energy storage and redistribution.¹⁶ Therefore, ECO₂R has emerged as a significant research area with industrial prospects, and, in the

recent past, much progress has been made in this prospering domain.^{17,18}

As reported in the scholarly literature and with the progress of research outcomes, the ECO₂R catalysts are classified into metal and metal-free catalysts based on their design, synthesis, and product selectivity.¹⁹ Nonmetal electrocatalysts are mostly made of carbon-based nanoparticles.²⁰ On the other hand, metal electrocatalysts include bare metals and their hybrid complexes, such as metal alloys,^{19,21,22} metal oxides,^{23,24} metal sulfides,²⁵ and metal complexes.^{26,27} The activity and selectivity for the ECO₂R vary with different metal catalysts owing to variations in the activity and adsorption strength of the intermediates.²⁸ Thus, metal catalysts are further classified into four categories based on their selectivity for final products: (a) Cd, Hg, Tl, Sn, In, Bi, and Pb produce the HCOO[−] intermediate, resulting in formate as their primary product;^{29–34} (b) Au, Ag, Zn, and Pd produce CO as their main reduction product because of the weak CO adsorption capacity;^{35–42} (c) Ni, Fe, Si, V, and Pt produce H₂ owing to the low HER potential;^{43–47} and (d) Cu produces up to sixteen carbon-based nanomaterials, comprising several highly energy-dense multi-carbon products, under different conditions.^{48,49} Hori *et al.*⁵⁰ conducted ECO₂R in 0.5 M KHCO₃ at 5 mA cm^{−2} for an hour on various polycrystalline metal electrodes. The outcomes showed that each metal electrode needs a significant overpotential to reach a particular current density (in this case, 5 mA cm^{−2}). Au requires the lowest overpotential (−0.6 V_{RHE}) to obtain CO (87%), but Pb needs the highest overpotential (−1.1 V_{RHE}) to produce formate (97%).⁵⁰

Although significant progress has been made, some challenges still exist with the industrial aspects of ECO₂R,^{51,52} such as (1) high overpotential for ECO₂R as CO₂ is a linear molecule that is thermodynamically stable and chemically inert;⁵³ (2) sluggish kinetics of CO₂ electroreduction because the complex proton-coupled electron transfer (PCET) steps and kinetics of electron transfer were slowed down by the insolubility of CO₂ in the aqueous medium;⁵⁴ (3) low exchange current densities; a



Susanta Sinha Mahapatra

electrochemical supercapacitors, electrochemical CO₂ reduction and hydrogen evolution reactions.

Dr Susanta Sinha Mahapatra is an Assistant professor at the Chemistry Department at the Birla Institute of Technology, Mesra, Ranchi, India. He received his MSc in Chemistry and MTech in Energy Science & Technology from Visva-Bharati, Santiniketan, India, and Jadavpur University, Kolkata, India, respectively, and his PhD from Indian Institute Engineering Science & Technology, Shibpur, India. His research mainly focuses on direct alcohol fuel cells,



Bidhan Chandra Ruidas

Dr Bidhan Chandra Ruidas completed his PhD in Chemical Engineering at the Indian Institute of Technology Kharagpur, India. Currently, he is working as an Assistant Professor in the Department of Chemical Engineering at the Birla Institute of Technology, Mesra Ranchi, India. His research mainly focuses on enhanced oil recovery and electrochemical reduction of CO₂ to value-added chemicals.



majority of CO₂ electrocatalysts reported thus far operate at less than 20 mA cm⁻²; however, this is significantly less than those of the commercial electrolyzers, which generally operate at over 70% efficiency at current densities exceeding 200 mA cm⁻²,⁵⁵ (4) unsatisfactory selectivity, suggesting expensive separation procedures;⁵⁶ (5) catalyst deactivation in less than 100 h, limiting practical application and industrialization of technology;⁵⁷ and (6) competitive hydrogen evolution reaction (HER); the HER surpasses the ECO₂R due to its favourable energy conditions at higher overpotentials, leading to a higher FE_{H₂} (faradaic efficiency) than FE for other ECO₂R products.⁵⁸ Recently, much research has concentrated on designing and synthesizing innovative, cost-effective, and robust electrocatalysts that can counter these bottlenecks and reduce CO₂ at high rates with minimal overpotential.^{57,58}

As pointed out in an earlier paragraph, Cu can electrochemically reduce CO₂ to highly energy-dense C₂₊ products, which has gained massive attention.⁵⁹ However, the critical path of the C–C coupling process required for the C₂₊ product formation introduces a high activation energy barrier, resulting in low activity and poor selectivity.^{60,61} In this context, synthesizing Cu-based electrocatalysts is crucial for overcoming the energy barrier in ECO₂R. Researchers showed that the CO₂ electroreduction performance (*i.e.*, activity and selectivity) of Cu catalysts is affected by multiple factors. For example, in their work, Hori *et al.*^{62,63} showed that Cu with several crystal facets could yield various ECO₂R products. The Cu(111) surface produces methane as a primary product, while C₂ products such as ethylene and ethanol are formed over Cu(100) surfaces. Reske and colleagues⁶⁴ also showed that faradaic efficiency and current density of Cu nanoparticles were significant compared to bulk Cu, while Chen *et al.*⁶⁵ demonstrated that the single-atom Cu showed better ECO₂R performance than other reported Cu-based materials, *i.e.*, bulk metal Cu.

Moreover, bimetallic Cu-based electrocatalysts have shown promise in enhancing copper's selectivity and overall catalytic performance. For example, Burdyny's group⁵⁸ could electrochemically reduce CO₂ to formate as the primary product using bimetallic Cu–Pd with a faradaic efficiency of 93% and a current density of 150 mA cm⁻² at a cell potential of $-2.9 V_{RHE}$ using a zero-gap flow cell, also known as the membrane electrode assembly cell. They successfully demonstrated how the engineering design of an electrochemical cell, coupled with catalyst structure, achieves the highest overall reported energy efficiency (EE) for formate production at 47%. These results show the benefit of using non-post-transition metals as primary catalysts for formate production. Another study shows that catalyst morphology is also essential: Chorkendorff *et al.*⁶⁶ found that metallic Cu with varying surface roughness exhibits distinct selective ECO₂R product formation. They discovered that the Cu nanoparticle-coated electrocatalysts have improved hydrocarbon selectivity compared to non-coated electrocatalysts. Motivated by the aforementioned work, several research groups synthesized different morphologies, such as nanowires,^{67,68} films,⁶⁹ microcubes,^{70,71} core-shell,^{72,73} and Cu-based bimetallic⁵⁸ catalysts and studied their catalytic activity towards ECO₂R. They

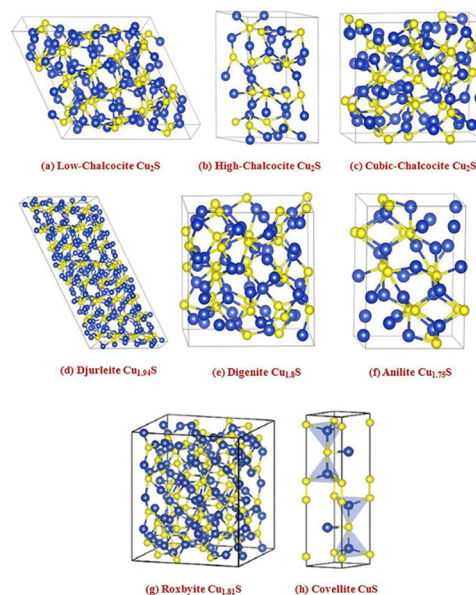
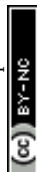


Fig. 2 Crystal structures of some representative copper sulfide phases. Blue spheres – copper atoms and yellow spheres – sulfur atoms, surrounded by black lines showing each unit cell boundary ((a)–(f) reproduced with permission from ref. 95, copyright American Institute of Physics 2012 and (g) and (h) reproduced with permission from ref. 77, copyright RSC 2019).

showed that the catalytic activity of Cu-based catalysts towards ECO₂R can be remarkably enhanced by tailoring the catalysts' structure and morphology.

In recent years, Cu/S-based nanomaterials have gained significant attention as electrocatalysts due to their p-type semi-conducting, earth-abundance and nontoxic characteristics. Various phases of copper sulfide are reported in the literature, namely chalcocite (Cu₂S), djurleite (Cu_{1.97}S), digenite (Cu_{1.80}S), anilite (Cu_{1.75}S), geerite (Cu_{1.60}S), spionkopite (Cu_{1.40}S), yarrowite (Cu_{1.12}S) and covellite (Cu_{1.00}S) as copper-rich systems, while villamanite (CuS₂) as a sulfur-rich system and covellite (CuS) as a 1/1 system, usually denoted as Cu_{2–x}S having minimal values of *x*.^{74–77} Based on the packing of sulfur atoms in the lattice, the aforementioned crystal structures have been grouped into three categories, as illustrated in Fig. 2, specifically cubic close packing (anilite and digenite), close hexagonal packing (djurleite and chalcocite), and a combination of close hexagonal packing and covalent bonding of the sulfur atoms (covellite).⁷⁷ But for the remaining forms, *i.e.*, yarrowite, spionkopite, and geerite, crystal structures remain unknown. It is to be noted that copper sulfides' electrical conductivity depends on their phases, decreasing from copper-deficient to copper-rich.^{78,79} For example, it has been observed that at 1.63 K, the naturally occurring covellite phase of CuS exhibits exceptional electrical conductivity.⁸⁰ Therefore, owing to their unique versatile properties, copper sulfides are fascinating nanomaterials for various applications, *i.e.*, optoelectronic devices,⁸¹ photocatalysis,⁸² photovoltaic cells,⁸³ sensors,⁸⁴ battery electrodes,⁸⁵ and biomedical applications.⁸⁶ Thus, numerous studies have focused on the engineering strategies of these materials for modifying their properties, including the electronic



structure modulation for ECO₂R. Notably, developing ECO₂R catalytic systems that can overcome bottlenecks is becoming an essential topic with the increasing use of this electrocatalytic technology. According to the reported literature, a variety of physical⁸⁷ and chemical⁷⁵ methods, such as hydrothermal and solvothermal,^{88,89} ball-milling,⁹⁰ electrodeposition,⁹¹ microwave irradiation,⁹² thermolysis,⁹³ and template-assisted^{94,95} approaches, have been widely used for constructing different nano-dimensional (*i.e.*, zero-dimensional, one-dimensional, two-dimensional, and three-dimensional) Cu/S-based nanomaterials. So far, several synthetic processes have yielded various shapes of Cu/S-based nanostructures, such as nanoparticles, nanoplates, hollow spheres, nanorods, nanowires, nanotubes, nanosheets, *etc.* Therefore, the significant connections within several nanomaterials and their distinct properties have prompted research into the controlled fabrication of valuable electrocatalysts with altered nanostructures.

Until recently, many outstanding review articles on Cu/S-based nanomaterials for several energy applications have been published,^{74,79,96} including the preparation, characterization methods, and the effects of modification strategies on catalytic activity. However, an in-depth and systematic investigation of synthesis strategies of Cu/S-based electrocatalysts and electron modulation effects on ECO₂R is lacking and is desired. This review summarizes the most recent advancements in Cu/S-based catalysts for ECO₂R reactions and outlines the connection between catalytic efficiency and several engineering approaches. The primary section discusses various modification strategies for Cu/S-based nanomaterials. After this, the effects of these strategies on electron modulation will be summarized. This fundamental understanding can help with the design of high-efficiency ECO₂R catalysts. Finally, this article highlights the research challenges and future prospects of Cu/S-based nanomaterials for ECO₂R.

2. Characteristics for improving the ECO₂R performance of Cu/S-based catalysts

This section discusses the essential features of Cu/S-based catalysts to understand their fundamental characteristics, which are mentioned below in each subsection.

2.1 Exposed facets for improving electrocatalytic performance

As reported by many literature studies, tuning the exposed facets of nano-electrocatalysts may alter atomic rearrangements, reaction intermediate affinity, and surface energy, influencing the electrochemical activity.^{97,98} Inspired by this, a research group presented a simple and effective technique for limiting product distribution using sulfur-modified Cu₂O electrocatalysts.⁹⁹ A wet chemical technique was used to synthesize distinct morphologies of Cu₂O effectively with varied exposed facets (Fig. 3a). Surprisingly, they observed that the faradaic efficiency of formate on sulfur-modified Cu₂O electrocatalysts was significantly sensitive to Cu₂O crystal facets, as shown in Fig. 3b–i, with selectivity in the order of

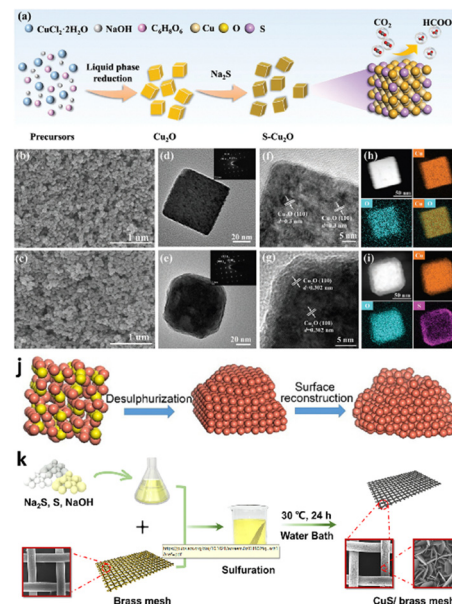


Fig. 3 Synthesis and morphological characterization. (a) Schematic illustration of the electrocatalyst preparation process. SEM images of (b) S0-Cu₂O-70 and (c) S3-Cu₂O-70. TEM images of (d) S0-Cu₂O-70 and (e) S3-Cu₂O-70 (the inset shows the selected area electron diffraction (SAED) patterns). HRTEM images of (f) S0-Cu₂O-70 and (g) S3-Cu₂O-70. The corresponding elemental mappings of (h) S0-Cu₂O-70 and (i) S3-Cu₂O-70 (reproduced with permission from ref. 99, copyright Wiley-VCH GmbH 2023). (j) Diagrammatic representation of the conversion from Cu_{2-x}S to Cu-s (reproduced with permission from ref. 100, copyright Springer Nature 2021). (k) Diagrammatic representation of the CuS NS arrays/BM synthesis method (reproduced with permission from ref. 101, copyright ACS 2021).

Cu₂O(100) > Cu₂O(100)/(111) > Cu₂O(111). Among all the sulfur-modified Cu₂O electrocatalysts prepared, the optimized S3-Cu₂O-70 demonstrated a FE of 90% at −0.9 V_{RHE} and an extended stability of more than 80 h in an H-type cell. Furthermore, a flow-associated cell system achieved a *j*_{formate} of 260 ± 16 mA cm^{−2}, outperforming most formate-producing Cu-based electrocatalysts. Sulfur can improve water activation for synthesizing unique H₂ species and reduce the activation energy of *OCHO intermediate formation on the surface of sulfur-modified Cu₂O, boosting formate selectivity in ECO₂R, according to experimental data and density functional theory (DFT) calculations. Similarly, He *et al.*¹⁰⁰ investigated the role of rich high-index facets of polycrystalline Cu (Cu-s) nanoparticles successfully derived from Cu_{2-x}S nanocrystals, as illustrated in Fig. 3j. They observed that the formation of high-index facets during surface reconstruction is beneficial for providing surface active sites for C–C coupling, thus boosting C₂H₄ generation. The Cu–S nanocatalysts exhibited high catalytic performance with an FE of 68.6% (a *j*_{C₂H₄} of 40.8 mA cm^{−2}) for C₂H₄ because of vicinal facet formation during surface engineering. Furthermore, *in situ* studies demonstrated that Cu–S electrocatalysts use the *COCHO intermediate route for producing C₂H₄ *via* ECO₂R. Similarly, Dou *et al.*¹⁰¹ investigated the facet-dependent selectivity and activity of CuS nanosheet arrays on brass mesh for ECO₂R prepared *via* a facile and green chemical bath

deposition approach (Fig. 3k). Meanwhile, connecting CuS with BM increased overall performance ($j = 75 \text{ mA cm}^{-2}$ at $-0.7 \text{ V}_{\text{RHE}}$ and $\text{FE}_{\text{HCOO}^-} = 67.8 \pm 1\%$) for ECO_2R . Instead, PTFE-coated CuS/BM achieved more enhanced CO_2 conversion to generate $\text{HCOOH}/\text{HCOO}^-$ ($\text{FE} = 70.2 \pm 1\%$ at $0.7 \text{ V}_{\text{RHE}}$) than CuS/BM. Their investigations reveal that the reconstruction of CuS/BM resulted in a uniform nanowire framework with abundant active surfaces during the ECO_2R , considerably increasing catalytic reactivity. They integrated DFT research with experimental findings. The work attributes the high selectivity for HCOO^- production to the reconstructed formation of the $\text{Cu}(111)/\text{CuS}(102)$ facets throughout the electrolysis process. According to the theoretical investigation, S under the CuO layer reduces the binding energies of HCOO^* and $^*\text{COOH}$ on $\text{Cu}(111)/\text{CuS}(102)$ compared to the $\text{Cu}(111)$ plane, allowing the development of HCOOH or HCOO^* .

2.2 Edge engineering for improving electrocatalytic performance

Recent research has established that the exposed edges of transition metal chalcogenides are more catalytically active than planar surfaces in electrochemical reactions.^{102,103} The more exposed edges on the electrode/electrolyte interfaces can improve the electrochemical performance owing to more active sites. Owing to the sufficiently exposed-edge planes, the hierarchical hollow CuS microcubes (MCs) exhibited remarkable electrocatalytic properties for CO_2 reduction with a FE_{CO} of 32.7% at a lower onset potential of $-0.2 \text{ V}_{\text{RHE}}$. Shao *et al.*¹⁰⁴ reported a novel engineered surface and morphology-enhanced exposed edge sites, offering significantly higher electrocatalytic activities and selectivity toward ECO_2R . They demonstrated CO_2 electro-reduction with lower overpotential for hollow CuS microcubes (MCs). According to the morphological characterization, it has been shown that as-synthesized h-CuS MCs lead to a higher density of defective edge exposed sites, essential for enhanced conversion activity and selectivity.

2.3 Porosity effect for improving electrocatalytic performance

Owing to their high specific surface areas, nanoporous structures significantly accelerate surface reactions and facilitate mass transfer by enhancing contact between the electrolyte and the active site.¹⁰⁵ The adsorption and transformation of intermediates, the release and diffusion of gas products, and the lowering of electron transfer resistance are all significantly more accessible by the active sites of the nanoporous structures.^{106,107} Inspired by this, Li *et al.*¹⁰⁸ designed and synthesized hollow-ordered porous copper sulfide cuboctahedra (HOP CuS-CO) with regulated shell thicknesses, as shown in Fig. 4a–h, to selectively produce formate from ECO_2R . The hollow shells of HOP CuS-CO possess uniformly distributed and interconnected pores. With these advantages and benefits of porous cages, the HOP CuS-CO catalyst exhibited an exceptional $\text{FE}_{\text{formate}}$ of 70.3% with stability of up to 26 h at a potential of $-1.1 \text{ V}_{\text{RHE}}$. *In situ* Raman spectroscopy studies showed that the HCOO^* intermediate adsorption energy is favourable on the surfaces of HOP CuS-CO by a spatial confinement effect, resulting in highly effective ECO_2R for formate generation.

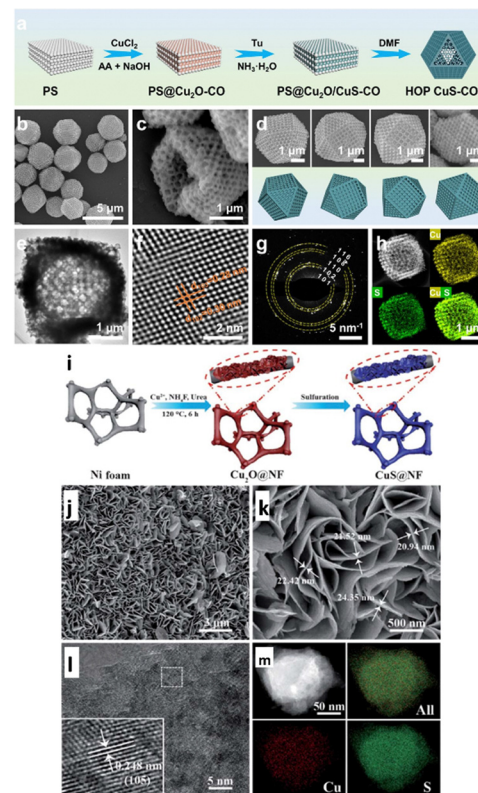


Fig. 4 (a) Schematic illustration of synthesis of HOP CuS-CO. (b) SEM image of HOP CuS-CO. (c) SEM images of one individual HOP CuS-CO crystal projected from four different directions. (d) SEM image of partially broken HOP CuS-CO particles. (e) TEM and (f) HRTEM images and (g) SAED patterns of HOP CuS-CO. (h) HAADF-STEM image and the corresponding elemental maps of Cu (yellow) and S (green) (reproduced with permission from ref. 108, copyright Elsevier 2024). (i) Schematic illustration of synthesis of CuS@NF samples. (j and k) SEM images of CuS@NF samples with lower and higher magnification. (l) HR-TEM and (m) HAADF STEM images and corresponding EDS elemental mapping of CuS@NF samples. (reproduced with permission from ref. 110, copyright RSC 2017).

The above studies provided insights into designing novel morphologies for more outstanding formate production using ECO_2R . Recently, Yabuki *et al.*¹⁰⁹ employed the thermal breakdown of a sulfur and copper-amine complex ink to create copper sulfide film electrodes. Furthermore, the XRD study confirmed that $\text{Cu}_{1.8}\text{S}$ and CuS nanoparticles were present in the film. The copper sulfide film possessed variations in the surface area caused by the film's micropores, resulting in ECO_2R to CO , CH_4 , and C_2H_4 , with a more significant percentage of C_2H_4 (C_2 product) than a copper electrode.

Zhu *et al.*¹¹¹ synthesized highly porous Cu_2S -decorated copper foam (Cu-foam), an active CO_2 reduction electrocatalyst, in an H-cell system. They used anodization followed by a heat treatment process to deposit Cu_2S nanoarrays on Cu foam. Anodization was placed in the first phase in an electrocatalytic cell, employing platinum foil as the cathode and Cu-foam as the anode in an aqueous Na_2S solution. After the anodization, the anodized Cu-foam was washed multiple times with distilled water, followed by thermal treatment. The 3D-shaped $\text{Cu}_2\text{S}/\text{Cu}$ -foam electrode produces much more HCOOH ($\text{FE}_{\text{CO}} = 85\%$ with



$j = 5.3 \text{ mA cm}^{-2}$ at $-2.0 \text{ V}_{\text{Ag/AgCl}}$) than the Cu-foam ($\text{FE}_{\text{CO}} = 38.9\%$ at $-1.8 \text{ V}_{\text{Ag/AgCl}}$). Using a simple hydrothermal procedure, Zhao *et al.*¹¹⁰ efficiently fabricated economical, robust and highly porous CuS nanosheet arrays successfully decorated on a porous nickel foam support, as illustrated in Fig. 4i, for ECO_2R activity. The CuS nanosheet is highly dense and is evenly dispersed over the highly porous Ni foam structure, forming a 3D organized foam CuS/NF, as shown in Fig. 4j–m. The thicknesses of the highly porous CuS/NF framework range from 20 to 25 nm. Their findings revealed that S concentration in the electrodes facilitates CO_2 adsorption and speeds up the rate-limiting step by converting CO_2 to $\text{CO}_2^{\bullet-}$ intermediates. Next, the $\text{CO}_2^{\bullet-}$ intermediate produces CH_4 by PCET reaction. Thus, CuS@NF achieved an extraordinary faradaic efficiency of 73.5% at $-1.1 \text{ V}_{\text{RHE}}$ for CH_4 formation and was stable up to 60 h.

3. Emerging regulations for improving the ECO_2R activity

3.1 Phase effect for improving electrocatalytic performance

This section discusses the ECO_2R activity of catalysts, which is determined by the final stage of the active metal's conversion, not the initial phase.^{112–114} Notably, Cu/S-based materials typically act as pre-catalysts for ECO_2R due to their conversion to oxide form during electrolysis. Due to their high oxidation potential, Cu/S-based materials could be almost totally/partially transformed into their oxide species.^{115,116} For example, Chen *et al.*¹¹⁷ observed that during the ECO_2R , copper sulfide nanoflowers (Fig. 5a and b) undergo restructuring to metallic Cu (Fig. 5c), and S^{2-} ions are released into the electrolyte and adsorbed on the surface of Cu catalysts. This phenomenon suppressed the formation of other hydrocarbon products, resulting in high selectivity towards HCOOH during ECO_2R . Copper sulfide pre-catalysts with high S content may absorb more S^{2-} ions on the Cu surface, resulting in higher FE_{HCOOH} . To assess this hypothesis, they used copper foil as the electrode and added varying amounts of K_2S to the KHCO_3 electrolyte. Later, they discovered that the FE_{HCOOH} of copper gradually increased with increasing K_2S concentrations. In this context, Phillips *et al.*¹¹⁸ used *in situ* electro-reduction to investigate selective formate production over copper sulfide-derived copper surfaces (S-derived copper) (Fig. 5d–i). As evidenced by surface-enhanced infrared absorption (SEIRA) spectroscopy, the aforementioned electrocatalyst reduces H_2 and CO formation while enhancing formate selectivity. The authors interpreted this increase in formate selectivity by describing a plausible reaction mechanism in which the active sites are occupied by CO_{ads} , preventing the adsorbed hydrogen molecules (H_{ads}) from combining and producing an H_2 molecule. As a result, H_{ads} could only generate H_2 in solution by making bonds with protons *via* PCET. Theoretical simulations suggest that H_{ads} might create HCOOH by reacting with a solution-phased- CO_2 molecule rather than an H^+ -containing solution. As a result, H_{ads} molecules adsorbed on S-derived Cu surfaces and generated HCOOH by interacting with solution-phase CO_2 *via* PCET.

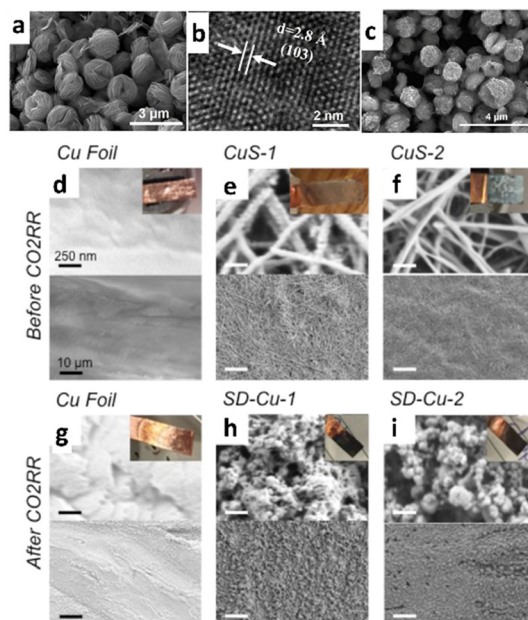


Fig. 5 (a) SEM image and (b) HR-TEM image of CuS nanoflowers before ECO_2R ; (c) SEM image of CuS nanoflowers after ECO_2R (reproduced with permission from ref. 117, copyright Elsevier 2024). SEM image of catalyst surfaces (d)–(f) before and (g)–(i) after ECO_2R . (d) and (g) Copper foil, (e) and (h) copper sulfide (CuS) electrodeposited for 30 min (CuS-1), and (f) and (i) CuS electrodeposited for 2 h (CuS-2). The CuS samples convert to SD-Cu after ECO_2R (top scale bars – 250 nm and bottom scale bars – 10 μm for each panel, reproduced with permission from ref. 118, copyright ACS 2018).

In a different work, through a straightforward two-step coupling procedure *via* a hydrothermal method followed by pyrolysis, as illustrated in Fig. 6a, Zhang *et al.*¹¹⁹ created a $\text{Cu}_{1.81}\text{S}$ catalyst supported by a multi-walled carbon nanotube (MWCNT). Due to the highly active sites of the uniformly dispersed $\text{Cu}_{1.81}\text{S}$ particles and the effective electron transport and active sites provided by the MWCNT, the $\text{Cu}_{1.81}\text{S}@ \text{MWCNT}-600$ composite catalyst (Fig. 6b–e) was able to achieve superior ECO_2R performance with 30 h stability during continuous operation. Later, they reported that $\text{Cu}_{1.81}\text{S}@ \text{MWCNT}-600$ ($\text{Cu}_{1.81}\text{S}@ \text{MWCNT}-600\text{-OD}$) with oxide modification showed improved catalytic activity and had a high $\text{FE}_{\text{Formate}}$ of 82%. According to the authors, copper oxide, which changed the phase into a needle-shaped structure during ECO_2R , provided more active sites and improved electrocatalytic activity (Fig. 6f–i). In another work, Oversteeg *et al.*¹²⁰ investigated the role of phase engineering *via* Cu_{2-x}S derived copper sulfide-supported carbon (CuS/C and $\text{Cu}_2\text{S}/\text{C}$) (Fig. 6j, m and n) ECO_2R catalysts synthesized using the liquid phase sulfidation of CuO/C nanoparticles (Fig. 6l and o). All the prepared phases are confirmed by XRD analysis, as shown in Fig. 6k. According to the electrochemical and *in situ* X-ray absorption (XAS) spectroscopy investigations, the metallic Cu reduction occurs in $\text{CuS}@ \text{C}$ and $\text{Cu}_2\text{S}/\text{C}$ nanoparticles during electrochemical CO_2 reduction (Fig. 6p and q). Later, their observation revealed that CuS/C - and $\text{Cu}_2\text{S}/\text{C}$ -derived catalysts had higher selectivity towards creating formate at low current densities than the CuO/C -derived electrocatalyst. Surprisingly, the catalyst only needed less

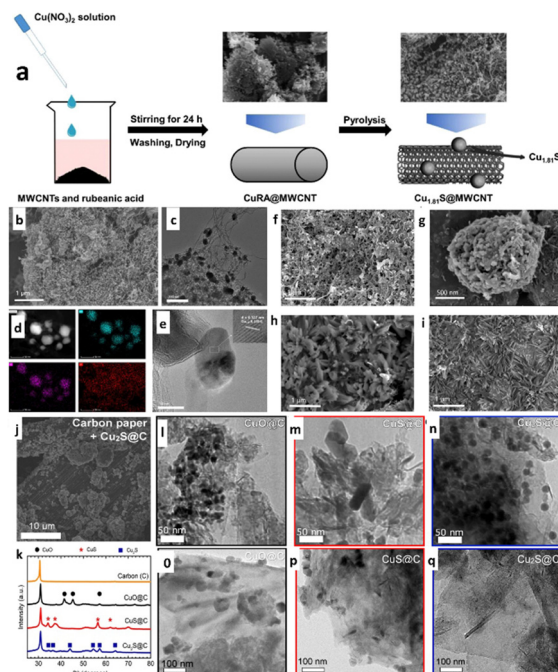


Fig. 6 (a) Schematic illustration of the preparation of $\text{Cu}_{1.81}\text{S}@ \text{MWCNT}$. (b) SEM and (c) TEM images of $\text{Cu}_{1.81}\text{S}@ \text{MWCNT}$ -600. (d) HAADF-STEM image and the corresponding EDS maps revealing the uniform distribution of Cu (blue) and S (purple) in the $\text{Cu}_{1.81}\text{S}$ particles. (e) HRTEM image of $\text{Cu}_{1.81}\text{S}@ \text{MWCNT}$ -600. SEM images of used $\text{Cu}_{1.81}\text{S}@ \text{MWCNT}$ -600-OD after (f) 10 min, (g) 20 min, (h) 1 h, and (i) 20 h of ECO_2R (reproduced with permission from ref. 119, copyright Elsevier 2020). (j) SEM images of the carbon paper substrate with $\text{Cu}_2\text{S}@ \text{C}$ deposited on the carbon fibres by spraying. (k) XRD patterns of the bare GNP-500 carbon support (orange) and of $\text{CuO}@ \text{C}$ (black), $\text{CuS}@ \text{C}$ (red) and $\text{Cu}_2\text{S}@ \text{C}$ (blue) nanoparticles on this carbon support. TEM images of (l) $\text{CuO}@ \text{C}$, (m) $\text{CuS}@ \text{C}$ and (n) $\text{Cu}_2\text{S}@ \text{C}$ nanoparticles before electrolysis; TEM images of the (o) $\text{CuO}@ \text{C}$ -, (p) $\text{CuS}@ \text{C}$ -, and (q) $\text{Cu}_2\text{S}@ \text{C}$ -derived catalysts after 5 h of ECO_2R (reproduced with permission from ref. 120, copyright Elsevier 2021).

carbon surface coverage ($<4\%$) for total formate selectivity to achieve the highest FE (12%). The efficiency of formate generation in ECO_2R can be increased using sulfur-derived copper with carbon catalysts.

An interesting approach has been reported recently, in which they created a series of CuS catalysts using different precursors and examined the relationship between phase restructuring and ECO_2R catalytic activity during electrolysis. First, Guo *et al.*¹²¹ employed hydrothermal synthesis to create several CuS-based catalysts using different sulfur precursors (*i.e.*, TU = thiourea, STS = sodium thiosulfate, TAA = thioacetamide, and SS = sodium sulfide). Surprisingly, in terms of ECO_2R activity, CuS-TU outperformed the other electrodes (*i.e.*, CuS-STS, CuS-TAA, and CuS-SS) in CO_2 -saturated 0.1 M KHCO_3 electrolytes, with an FE_{CO} of 72.67% ($-0.51 \text{ V}_{\text{RHE}}$) and a high CO selectivity. They hypothesized that the thiourea precursor's rapid S decomposition led to a higher concentration of dissolved S^{2-} in the electrolyte, enabling quicker nucleation and restructuring of a nanoflower like CuS-thiourea electrocatalyst for more enhanced mass transfer kinetics and favourable ECO_2R .

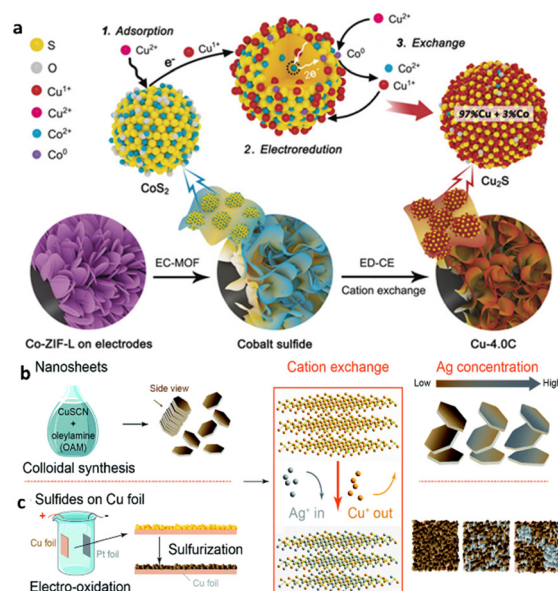


Fig. 7 (a) Schematic illustration of the experimental pathways and mechanisms for electrochemically driven cation exchange (ED-CE). Synthetic strategies for Ag/Cu sulfide catalysts (reproduced with permission from ref. 122, copyright Wiley-VCH 2020). (b) Cu sulfide nanosheets (C-nano-0, 100 nm lateral dimension, 14 nm thick) were obtained through colloidal synthesis with CuSCN in oleylamine (OAM). (c) Cu sulfides on Cu foil (C-foil-x) were obtained through electrooxidation in 1 M NaOH to produce an oxide layer of a few 10 s of microns thick, followed by sulfurization with 0.1 M Na_2S . After cation exchange where Ag^+ replaces the Cu^+ in the Cu sulfides, Ag/Cu sulfide nanosheets (CA-nano-x) remain in nanosheet structure with some distortion in shape as the Ag/Cu mass ratio ranges from 0.3 to 25, while for C-foil-x, Ag nucleates at higher Ag concentration, which impedes the uniform distribution of Ag and Cu (reproduced with permission from ref. 123, copyright RSC 2021).

The phase engineering of a target nanocrystal (NC) can be systematically varied by carefully substituting metal cations in a prefabricated NC template using an emerging electrochemically driven cation exchange (ED-CE) approach. He and coworkers¹²² recently designed a Cu_2S catalyst for ECO_2R from the CoS_2 template *via* the ED-CE approach (Fig. 7a). Employing the ion exchange process, Cu almost entirely replaces Co cations in CoS_2 , and the Cu/S atomic ratio is ~ 2.4 . The Cu_2S catalyst produced by converting the predesigned template retains the initial morphology of CoS_2 by preserving its high grain-boundary density, improving CO_2 adsorption. Also, the electronic structure of the nearby Cu sites has changed due to electronegative S heteroatoms, creating positively charged Cu^+ sites. The primary formate intermediate, $^*\text{OCHO}$ species, is formed when partially positive-charged Cu sites adsorb CO_2 molecules with more electronegative oxygen. The as-prepared 3D-shaped Cu_2S catalysts showed an FE of 87% with a current density of 19 mA cm^{-2} at $-1.9 \text{ V}_{\text{RHE}}$ in a 0.1 M NaHCO_3 medium for CO_2 conversion to HCOOH . Li *et al.*¹²³ utilized copper sulfide nanosheets as a template to develop Ag/Cu electrocatalysts through a straightforward ED-CE process (Fig. 7b and c). When the Ag^+ concentration in the exchange solution increased, the crystal structure of Cu_{2-x}S nanosheets

with lateral dimensions of 100 nm and a thickness of 14 nm progressively changed from Cu_7S_4 to Ag_2S (CA-nano- x , cation-exchanged nanosheets, where x indicates a higher concentration of Ag). The Ag/Cu mass ratio varied between 0.3 and 25. Hence, at an average overpotential ($-0.2 V_{\text{RHE}}$), both C-nano-0 and CA-nano- x exhibited outstanding FE_{HCOOH} . Surprisingly, when the Ag content increases, formate-producing C-nano-0 can generate C_{2+} products at $-1.0 V_{\text{RHE}}$. This observation concludes that the nanosheets show shape distortion as the Ag content rises while maintaining their original morphology after the cation exchange process.

In 2024, Goh *et al.*¹²⁴ investigated phase-engineered sulfide-derived Cu-Sb electrodes for electrochemical CO_2 conversion in a gas diffusion electrode (GDE) based cell. They synthesized several distinct Cu-Sb-S phases, skinnerite (SK; Cu_3SbS_3), tetrahedrite (TH; $\text{Cu}_{12}\text{Sb}_4\text{S}_{13}$), and chalcostibite (CS; CuSbS_2) using a heat-up colloidal nanoparticle route, as illustrated in Fig. 8a–f, and each showed a different selectivity for ECO_2R with CO as the main product, which contrasts with the individual CuS_x and SbS_x control samples, which show a preference for the formate product. They also demonstrated that the different elemental compositions caused the different selectivity patterns when the parent phases were reduced using fundamental composition characterization after reduction. Interestingly, lower Cu concentrations reduce phase segregation into harmful S-doped Cu that converts CO_2 to HCOO^- and H_2 , whereas higher sulfur concentrations disrupt crystallinity and promote CO formation. This outstanding performance is attributed to the tetrahedral Cu-Sb-S sample, which has the highest residual sulfur, with a CO_{FE} of about 80.5% at $-1.0 V_{\text{RHE}}$ and a $j_{\text{geometrical}}$ of 37.6 mA cm^{-2} . Post-electrocatalysis characterization combined with DFT calculations demonstrated that adding sulfur to Sb sites enhances $^*\text{COOH}$ binding compared to $^*\text{CO}$, rupturing scaling relations and aiding in CO(g) formation afterwards.

Similarly, in 2023, Mai *et al.*¹²⁵ used a facile solvothermal technique to create several cuprous sulfide nanoparticle-modified copper hydroxide nanowire array (S-CNW) pre-catalysts for elucidating the reaction mechanism of ECO_2R (Fig. 8g–n). They explored the effect of cuprous sulfide nanoparticle modification on formate generation during CO_2 reduction and observed sulfur modification changes in the intermediate during CO_2 reduction, leading to improved formate selectivity (60% of FE_{CO} with $j_{\text{HCOO}^-} = 10 \text{ mA cm}^{-2}$ at $-0.58 V_{\text{RHE}}$). Therefore, the role of trace sulfur alteration of copper surfaces in selective formate production is investigated using DFT. According to the study, sulfur modification in copper, as compared to a pure copper surface, can accelerate the synthesis of $^*\text{OCHO}$, a critical step along the formate pathway.

3.2 Size effect for improving electrocatalytic performance

Researchers discovered that the surface of tiny nanoparticles is rich in unsaturated atoms and defective sites compared to the bulk, exposing the number of active sites to improve catalytic activity.¹²⁶ Reducing nanoparticles even further to nanoclusters (NCs), which are made up of a few hundred or fewer atoms and have an average size of about 2 nm, reveals intriguing electrochemical performance



Fig. 8 SEM characterization of the samples before reduction. (a)–(c) SEM images of the samples SK (a), TH (b) and CS (c) sprayed on carbon paper. SEM characterization of the samples after reduction. (d)–(f) SEM images of the samples SK (d), TH (e) and CS (f) sprayed on carbon paper (reproduced with permission from ref. 124, copyright RSC 2024). Morphological and structural characterization; (g) schematic illustration of S-CNW preparation, (h)–(k) SEM images of S-CNWs with different loadings of TAA at 75°C for 0.75 h, (50, 75, 100, and 200 mg per 80 mL) for S-CNWs-1, S-CNWs-2, S-CNWs-3, and S-CNWs-4, respectively; (l) XRD patterns of CNWs and S-CNWs with different degrees of vulcanization, (m) Raman spectra of CNWs and S-CNWs with different degrees of vulcanization, (n) HAADF-STEM and the corresponding EDS images of S-CNWs (reproduced with permission from ref. 125, copyright ACS 2023).

with significant quantum size effects.¹²⁷ Shinagawa *et al.*¹²⁸ used a wet chemistry approach to create carbon-supported CuS nanocatalysts by controlling the size. The experimental investigation demonstrated that during ECO_2R the aforementioned nanometric-sized CuS was restructured and reformed to S-modified copper (Cu-S). As a result, at considerable overpotential ($-0.8 V$ vs. RHE), Cu-S catalysts produced formate with an FE of $>60\%$ and negligible quantities of CO as a byproduct. They discovered that as the particle size dimension became 3 nm to 20 nm, there was a slight increase in HCOOH production, highlighting the catalytic size-activity relationship. Following this, in a similar article, it was demonstrated that solvothermally prepared submicron-sized CuS electrodes had better FE_{HCOOH} (80%) for ECO_2R to HCOOH than nanometric CuS ($\text{FE}_{\text{HCOOH}} > 60\%$), indicating the important link between the particle size and electrochemical activity. Later, Lim *et al.*¹²⁹ studied the impact of size engineering by fabricating a size-controlled CuS_x electrocatalyst in an aqueous medium using Cu foil dipped in industrial CO_2 that contains H_2S (Fig. 9a–c). As per their observations, the Cu foil and the sulfur interacted appropriately when the concentration of sulfur in the solution increased. The simultaneous interactions raised the average particle size and surface sulfur density of CuS_x nanoparticles (NPs) to $133.2 \pm 33.1 \text{ nm}$ and $86.2 \pm 3.3\%$, respectively, as shown in energy dispersive X-ray spectroscopy (EDX) images (Fig. 9c). When the sulfur percentage and sizes of the CuS_x nanoparticles increased steadily, the $\text{FE}_{\text{Formate}}$ improved from 22.7 to 72.0% at $-0.6 V_{\text{RHE}}$ (Fig. 9d–g). Although the CuS_x nanocatalysts had a lower current

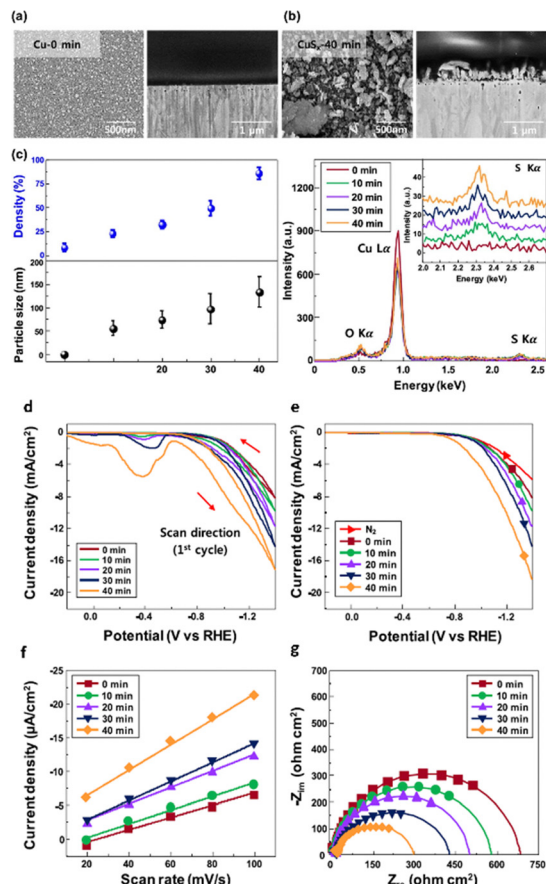


Fig. 9 Top-view and cross-sectional SEM images of (a) Cu-0 min without sulfur species and (b) CuS_x-40 min catalysts. (c) The average particle size and surface density of CuS_x NPs were calculated using a computing-based image analyzer and EDX spectra of Cu and CuS_x catalysts. (d) CV (1st scan) of CuS_x catalysts with different amounts of sulfur species. (e) LSV curves in N₂-purged and industrial CO₂-purged 0.1 M KHCO₃ electrolytes at a scan rate of 50 mV s⁻¹. (f) Current density plots at various CV scan rates. The current densities were obtained from the double-layer charge/discharge curves at 0.40 V_{RHE} in industrial CO₂-purged 0.1 M KHCO₃ electrolyte. (g) Nyquist impedance plots of catalysts in industrial CO₂-purged 0.1 M KHCO₃ electrolyte from 1 MHz to 0.1 Hz with 30 mV amplitude at -0.6 V_{RHE} (reproduced with permission from ref. 129, copyright ACS 2020).

density, the 72-hour stability must be addressed for industrial CO₂ conversion.

3.3 Grain boundary effect for improving electrocatalytic performance

Grain boundaries (GBs) can be efficient active sites for catalytic processes because they provide an electrodynamically beneficial surface.^{130,131} Also, it has been shown that catalytically energized surfaces can be stabilized using grain boundaries (GBs).^{132,133} For example, Yang *et al.*¹³⁴ used electrochemical reduction techniques to synthesize S-Cu₂O/Cu hybrid catalysts derived from Cu₇S₄/Cu nanoflowers (Fig. 10a–g). They showed that GB surface defects in Cu₂O/Cu interfaces were more energized than grain surfaces for ECO₂R. They observed that Cu₇S₄/Cu is dynamically restructured *in situ* to provide an S-Cu₂O/Cu hybrid catalyst for efficient ECO₂R to formate with

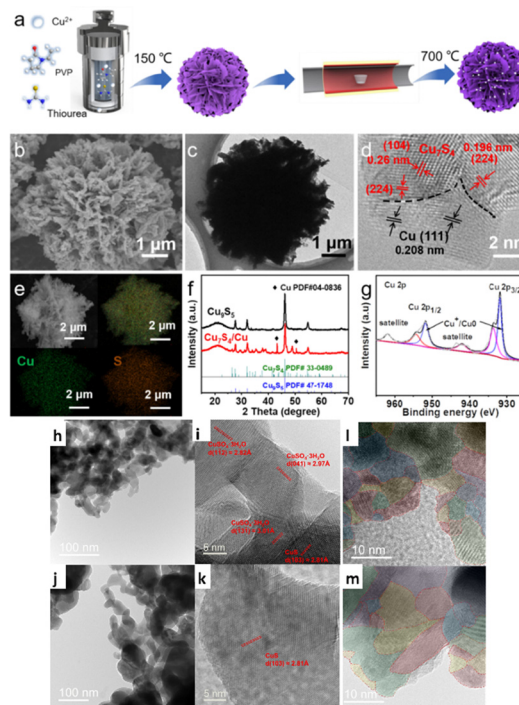


Fig. 10 (a) Schematic illustration of the synthetic process, (b) SEM image, (c) TEM image, (d) HRTEM image, (e) EDS elemental mapping, and (f) XPS pattern of Cu₇S₄/Cu. (g) XRD patterns of Cu₇S₄/Cu and Cu₉S₅ (reproduced with permission from ref. 134, copyright Wiley-VCH 2024). Morphological and structural characterization of CuS 811 and CuS MKL before ECO₂R by electron imaging. (h) and (i) Aberration-corrected TEM images of (h) CuS 811 and (i) CuS MKL. The crystal facets with heterogeneous species are highlighted in (j) and (k). Structural characterisation of CuS 811 and CuS MKL post ECO₂R by electron imaging. (l) and (m) HRTEM images of (l) CuS 811 and (m) CuS MKL. Red dash-dotted lines highlight the boundaries of induced grains, and the grains with heterogeneous lattice fringes are marked in various colours (l) and (m). The marking colours are only used to distinguish every grain, and there is no correlation between the colour and phase (reproduced with permission from ref. 135, copyright Wiley-VCH 2024).

a FE_{HCOOH} of up to 70% at -1.0 V_{RHE} and a partial current density of 5 mA cm⁻²), thus outperforming Cu₂O/Cu and Cu₇S₄. The authors attributed this performance to (i) thermodynamic and experimental investigations suggesting that the optimized adsorption of the HCOO* intermediate on the S-Cu₂O/Cu surface is modified, and S-doping suppresses the H₂ route (surface H), (ii) GBs at the Cu₂O/Cu interfaces reduce the adsorption energy favoured by S-doping and increased formate efficiency by inhibiting the HER route and CO₂-to-CO conversion. In another work, Wang *et al.*¹³⁵ fabricated S-doped Cu₂O derived from CuS 811 (consisting of CuS and CuSO₄·3H₂O) using the electrolysis method to study the relationship between the catalytic activity and GBs. This catalyst performed exceptionally well for formate production in ECO₂R, reaching an optimal FE of 92% in an H-type cell and an excellent *j*_{formate} of 321 mA cm⁻² in a flow cell while retaining an FE of more than 80%. The authors explained this outstanding performance from two viewpoints. (1) Structural studies showed that CuSO₄·3H₂O inhibits CuS growth and *vice versa*, leading to

lower grain sizes and more primary GBs (Fig. 10h–k). After electrolysis and electrochemical reduction, these grains break down into the smaller ones, resulting in more dense grain boundaries suitable for electrocatalysis (Fig. 10l and m). (2) The kinetics of S shedding during electrochemical ECO₂R are comparatively slow because of the initial mixture of multi-crystalline phases; this leads to a higher S content and a relatively complete CuS crystalline phase in CuS 811 post ECO₂R. They did DFT calculations and showed that retaining sulfur-sulfur bonds from covellite may reduce the binding energy, mainly by weakening the binding energies with several reaction intermediates, thus decreasing the energy barriers and facilitating the desorption steps and increasing formate generation activity. The catalytic impact of the GB was comparable to its dislocation strain field, establishing a way for a more comprehensive application of the GB effects in heterogeneous catalysts. More extensive research is required to understand the fundamental principle of GBs and apply it to developing advanced catalytic schemes for the ECO₂R into valuable products.

3.4 Defect/vacancy creation effect for improving electrocatalytic performance

Defect engineering, such as vacancy creation, exposing edge sites, and heteroatom doping, is essential for electrocatalytic activities because defects are usually rich in active sites.^{136,137} Furthermore, defects can modulate the electronic properties of active sites, increasing the density of catalytic active sites. As a result, defect engineering has become a strategy mostly used to increase active sites and improve charge transfer ability to fine-tune electrocatalytic activity.^{138,139} Therefore, considering the connection between diverse defects and the specific catalytic properties of various materials is critical for developing advanced catalysts.^{140,141}

For example, it is difficult to obtain the *n*-propanol product during electrolysis because of the complicated C₃ creation mechanism that requires the stabilization of *C₂ intermediates and subsequent C₁–C₂ coupling. Zhuang and colleagues¹⁴² synthesized a bifunctional core–shell nanostructure (Fig. 11a) showing that adding sulfur atoms to the catalyst core and copper vacancies to the shell named core–shell vacancy engineered Cu (CSVE-Cu) results in excellent ECO₂R activity towards *n*-propanol formation. The CSVE-Cu electrocatalyst exhibited satisfactory reduction performance by generating highly energy-dense C₂₊ alcohols (*i.e.*, C₃H₇OH and CH₃CH₂OH) with a FE of 32% ± 1%. The alcohol-to-ethylene ratio increased sixfold compared to simple copper nanoparticles, indicating an alternative route for producing alcohols instead of alkenes. According to DFT modelling, the incorporated vacancy, as shown by morphology analysis from Fig. 11b–g, on a bifunctional core–shell catalyst raises the activation energy of the C₂H₄ route (1.148 eV). Still, it does not affect the CH₃CH₂OH path (0.427 eV). Peng *et al.*¹⁴³ created a double-sulfur vacancy (DSV) engineering structure to achieve enhanced ECO₂R performance. The mechanistic study showed that the DSV-engineered CuS(100) planes facilitated the stabilization of both CO* and OCCO*, a *C₂ dimer which undergoes

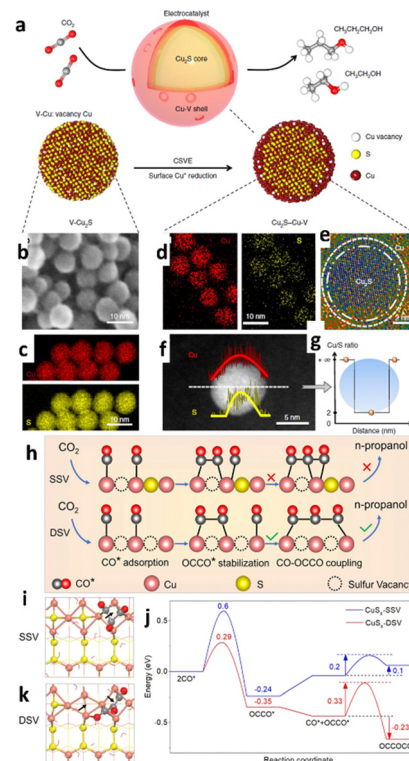
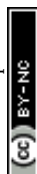


Fig. 11 Catalyst design and structural characterization. (a) Schematic illustration of Cu₂S–Cu–V CSVE electrocatalyst design for production of multi-carbon alcohols by CO₂ reduction. (b) and (c) TEM (b) and EDS (c) mapping of the original V–Cu₂S nanoparticles, showing the uniform size and the homogeneous distribution of Cu and S. (d)–(g) EDS mapping (d), high-resolution TEM (e), EDS line scan (f) and the ratio of Cu/S concentration (g) of the reduced CSVE nanocatalysts after electrochemical reduction, showing the removal of S from the nanoparticle surface. V–Cu indicates Cu with surface vacancies (reproduced with permission from ref. 142, copyright Nature 2018). (h) The mechanism of *n*-propanol formation on adjacent Cu_xS–SSV shows the dimerization of CO–CO followed by CO–OCCO coupling. (i) and (k) Top views of the optimized OCCOCO* intermediate configurations on the (100) surface of (i) Cu_xS–SSV and (k) Cu_xS–DSV. The arrows indicate the positions of sulfur vacancies. (j) The corresponding energy diagrams of Cu_xS–SSV (blue curve) and Cu_xS–DSV (red curve) at 0 V vs. RHE. The pink, yellow, grey, and red spheres and red wireframe in (h)–(k) represent copper, sulfur, carbon, and oxygen atoms and water molecules, respectively (reproduced with permission from ref. 143, copyright Nature 2021).

the subsequent interaction with a third *CO *via* CO–OCCO coupling (Fig. 11h–k). The DSV-engineered Cu_xS exhibited an improved FE_{*n*-PrOH} of ~15.4% –1.05 V_{RHE} for *n*-propanol formation in 0.1 M KHCO₃ medium in a H-cell set up, but in flow cells, *j*_{*n*-PrOH} was increased to 9.9 mA cm^{–2}. This study provides an appealing strategy for using the lithium electrochemical tuning method to create an array of novel frameworks with ion vacancies as active sites for electrochemical reactions.

3.5 Heterostructure effect for improving electrocatalytic performance

A heterostructure commonly comprises more than one component that accomplishes different roles in the electrocatalytic reactions.¹⁴⁴ Modification of physical and chemical properties will enhance each component's combined advantages. As a



result, the various components will produce a synergistic effect, contributing to ECO₂R kinetics.^{145,146} Unlike single-component catalysts, the synergistic effects caused by heterogeneous interfaces can significantly improve catalytic activity.¹⁴⁷ Inspired by this, recently, Tao *et al.*¹⁴⁸ developed a novel Cu₂S/SnO₂@C nanocomposite by solvothermal heating, where SnO₂@C is confined on the snowflake-like Cu₂S surface and has a combined interfacial effect at the Cu₂S sites (Fig. 12a–e). Specifically, the Cu₂S snowflake improves the CO₂ concentration near the surface, and carbon spheres increase the surface electron transport capacity. Afterwards, the as-synthesized Cu₂S/SnO₂@C nanocomposites are used as a working electrode to investigate ECO₂R to HCOOH. They varied the loading amount to examine the impact of the SnO₂@C content on the ECO₂R. The prepared Cu₂S/1%SnO₂@C catalyst shows good selectivity and activity for ECO₂R to form liquid HCOOH products, thus outperforming all the existing Cu-based electrocatalysts. The electrochemical study demonstrated a high ECSA, exceptional CO₂ adsorption capacities and a fast electron transport rate on the surface of Cu₂S/1%SnO₂@C. Additionally, DFT calculations revealed the plausible reaction pathways for the enhanced HCOOH production by Cu₂S/SnO₂@C. Notably, through heterostructure construction, this work provides an easy method to fabricate effective Cu-based catalysts that can improve HCOOH selectivity and activity in the ECO₂R process.

Another work used a redox process at ambient temperature to create a heterostructure of CuS and S-doped SnO₂ (CuS/SnO₂-S) for ECO₂R, as shown in Scheme 1 in Fig. 13.¹⁴⁹ Structure analysis methods revealed the structural regeneration phase that had occurred during the first electrolysis (Fig. 13a–i). The unique restructuring of the CuS/SnO₂-S heterostructure to

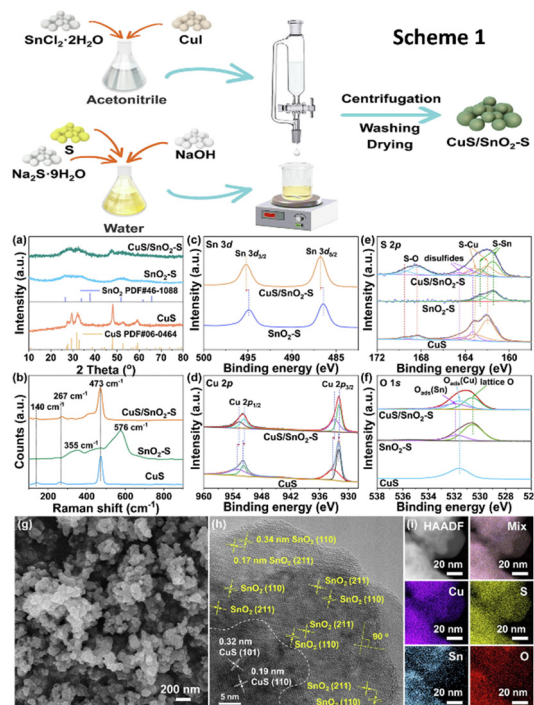


Fig. 13 Scheme 1: the synthetic process of CuS/SnO₂-S. (a) XRD patterns, (b) Raman spectra and (c)–(f) XPS analysis of Sn 3d, Cu 2p, S 2p and O 1s of CuS/SnO₂-S RCu/Sn high, SnO₂-S and CuS. (g) SEM images and (h) HRTEM and (i) HAADF image and EDS mapping of as-prepared CuS/SnO₂-S RCu/Sn high (reproduced with permission from ref. 149, copyright Elsevier 2023).

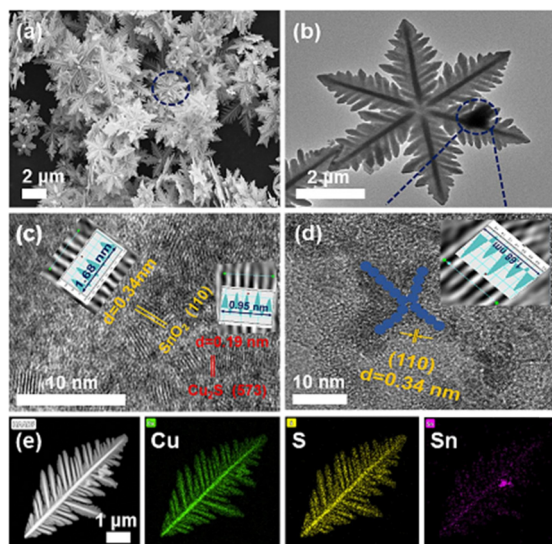


Fig. 12 SEM image (a) and TEM image (b) of Cu₂S/SnO₂@C. HRTEM images of Cu₂S/SnO₂@C (c) and SnO₂@C (d). The HAADF-STEM image and the corresponding elemental mapping of Cu₂S/SnO₂@C (e) reveal the distribution of Cu (green), S (yellow), and Sn (rosy), respectively (for interpretation of the references to colour in this figure legend, the reader is referred to the web version of this article; reproduced with permission from ref. 148, copyright Elsevier 2023).

Cu/Sn/Cu_{6.26}Sn₅ nanowires reduces CO₂ adsorption energy while increasing *H adsorption and reducing the competing HER. During ECO₂R, at $-0.8 V_{RHE}$, it achieves a formate conversion with a FE of 84.9% and a yield of 8860 $\mu\text{mol h}^{-1} \text{cm}^{-2}$ in an H-cell at a j_{formate} of 18.8 mA cm^{-2} . This research focused on the structural development of CuSn sulfides from precursor materials' early state as well as the process of formate production.

Wang *et al.*¹⁵⁰ created an innovative framework of 0D/2D composites of SnO₂ nanoparticles dispersed on CuS nanosheets (SnO₂/CuS) for selective syngas generation (a CO/H₂ ratio of 0.11–3.86), as illustrated in Scheme 1 in Fig. 14. The electrocatalytic system was highly efficient for syngas selectivity, with a faradaic efficiency of nearly 85%, a turnover frequency (TOF) of 96.12 h^{-1} , and stability of 24 hours. They explained the increased catalytic activity based on two factors: (a) the most active sites are provided by the uniformly distributed ultrasmall SnO₂ nanoparticles on ultrathin CuS nanosheets, enabling a faster electron transfer rate (Fig. 14a–c) and (b) the interfaces between SnO₂ and CuS lower the activation energy of reaction intermediates, enhancing ECO₂R performance to generate high-ratio tunable syngas. Both the SnO₂(110) and CuS(001) facet surfaces, as shown in Fig. 14d and e, favoured HCOOH creation. In contrast, the SnO₂/CuS contact considerably lowered the free energy of COOH* intermediate synthesis by 0.52 eV and encouraged CO formation.

In 2023, Liu and group¹⁵¹ reported the production of monoclinic-phase colloidal Cu₂SnS₃ nanoplates with precise



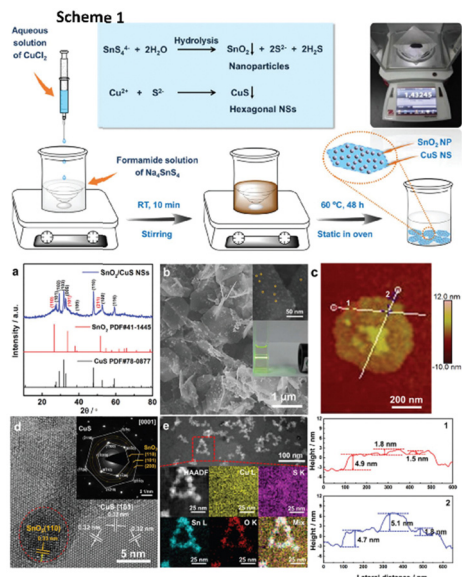


Fig. 14 Scheme 1: schematic illustration and the proposed reaction mechanism for the one-pot scalable synthesis of hierarchical SnO_2/CuS NSs. Structural characterization of SnO_2/CuS NSs. (a) XRD patterns. (b) SEM and (inset) magnified SEM image of the surface and photograph of the dispersion of SnO_2/CuS NSs in EtOH, demonstrating the Tyndall effect with a laser pointer. (c) AFM image and the corresponding height profiles of the as-obtained ultrathin SnO_2/CuS NSs marked by lines in different colours. (d) HRTEM image and the corresponding SAED pattern (inset). (e) HAADF-STEM images and the corresponding elemental mapping images of SnO_2/CuS NSs (reproduced with permission from ref. 150, copyright RSC 2020).

surfaces (Fig. 15a–h). Their findings revealed that thiocyanate-capped plate-shaped Cu_2SnS_3 nanoparticles exhibit outstanding formate selectivity over a broad spectrum of potentials, observing a formate production with a FE_{max} of 92% with a j_{formate} as high as 181 mA cm^{-2} in a GDA-based flow cell. Compared to prior studies of mono metal- and bimetal-based Cu- and Sn-sulfide nanoparticles that typically experienced phase separation or the creation of metal-based domains, Cu_2SnS_3 demonstrated outstanding structural robustness, as demonstrated by the concurrent retention of nanoplate morphology and the crystal phase during ECO_2R . *In situ* and DFT studies, as shown in Fig. 15i–m, have demonstrated that the Sn site basal-planes are the multi-active sites for favourable HCOO^* adsorption to produce formate by ECO_2R . According to DFT studies, thiocyanates also inhibit Cu sites on the surface, and the Sn site's electronic structure modulation is observed, favouring the activation energy barrier of ECO_2R to formate.

Also, Xiong *et al.*¹⁵² used a homogenous mixing approach to create bimetallic CuInS_2 hollow nanoparticles. According to their observations, the synergy between metal centres and hollow-shaped nanostructures accelerates the electron transfer kinetics. Consequently, the bimetallic catalyst had a FE_{CO} of 82.3% at $-1.0 \text{ V}_{\text{RHE}}$ and a FE_{HCOOH} of 72.8% at $-0.7 \text{ V}_{\text{RHE}}$. *In situ* studies showed faster conversion of CO_2 to $\text{CO}_2^{\bullet-}$ radicals as the rate-limiting step, and afterwards, electron

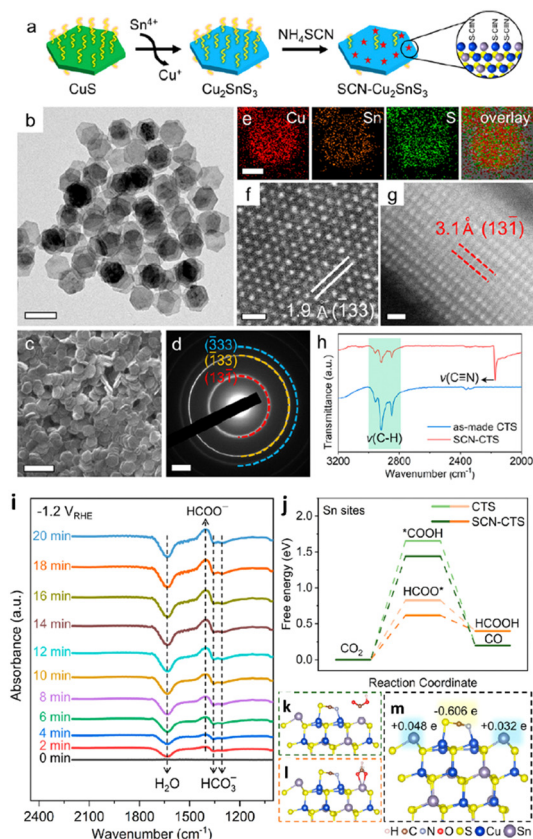


Fig. 15 Synthesis and characterization of Cu_2SnS_3 nanoplates. (a) Synthetic scheme of Cu_2SnS_3 nanoplates and ligand exchange with NH_4SCN . (b) TEM and (c) SEM images, (d) SAED pattern, (e) STEM-EDS elemental maps, (f) top-view and (g) side-view ACSTEM images of Cu_2SnS_3 nanoplates. (h) FTIR spectra of Cu_2SnS_3 nanoplates (basal plane edge length: $41.2 \pm 2.4 \text{ nm}$) before and after ligand exchange with NH_4SCN . Scale bars: (b) 100 nm, (c) 200 nm, (d) 2 nm^{-1} , (e) 30 nm, (f) and (g) 0.5 nm. (i) *In situ* FTIR spectra recorded at different times for CTS-3 at -1.2 V vs. RHE. (j) Free energy diagrams of ECO_2R into formate and CO on Cu_2SnS_3 nanoplates. (k)–(m) Optimized geometric structures of $^*\text{COOH}$ (j), HCOO^* (k), and SCN-modified Cu_2SnS_3 (001) surface (m). Also shown in (m) are charge transfers obtained using Bader charge analysis where e is the elementary charge (reproduced with permission from ref. 151, copyright ACS 2023).

redistribution happened at different potentials, leading to a product distribution shift (CO to HCOOH). Furthermore, Nyquist plots showed that hollow-like CuInS_2 nanocomposites have a substantially greater interfacial charge-transfer rate during electrocatalysis than Cu_2In because the interfacial charge-transfer impedance (R_{ct}) of Cu_2In is lower than that of CuInS_2 .

Graphdiyne (GDY) is a novel 2D all-carbon structure where alkyne bonds (sp -hybridized C) bind each benzene ring (sp^2 -hybridized C). GDY's particular sp/sp^2 -hybridized architecture provides numerous distinctive and intriguing qualities that are exceptional to standard carbon materials, *i.e.*, abundant carbon chemical bonds, massive conjugated pi structures, a favourable band gap, *etc.*¹⁵³ More enticingly, the exceptionally variable distribution of surface charges and incomplete charge transfer between GDY and metal centres can offer more active sites,

higher intrinsic activity, and efficient control of reaction intermediates' adsorption and desorption behaviour on functional site surfaces.¹⁵³ Owing to these advantages, in 2023, Li and colleagues¹⁵³ discovered a novel graphdiyne/copper sulfide (GDY/CuS_x) heterostructured electrocatalyst (Fig. 16a) controlling *in situ* development of GDY over the surface of CuS_x, as illustrated in Fig. 16b–x. The authors showed that the imperfect charge transfer between GDY and atomic Cu increased catalytic conductivity, providing additional active sites and enhancing the conversion performance. Thus, the heterostructure accomplished an FE of 70% and a j_{total} of 65.6 mA cm^{−2} at −0.9 V_{RHE} for ECO₂R to formic acid.

Heterostructure engineering modulates interfacial charge distribution and promotes CO₂ adsorption. Wen *et al.*¹⁵⁴ used a local sulfur doping strategy (SHKUST-1) to appropriately develop an isolated Cu@S motif on the HKUST-1 pre-catalyst (Fig. 17a). The *in situ* reconstruction of S-HKUST-1 results in a Cu(S) array and active biphased copper/copper sulfide (Cu/Cu_xS_y) interfaces (Fig. 17b), achieving highly selective ethylene (C₂H₄) formation in the H-cell with a FE_{max} of 60.0%. In a flow

cell configuration, ECO₂R occurred at a high current density of 400 mA cm^{−2} with a FE_{C₂H₄} of approximately 57% and FE_{C₂} (FE of C₂H₄, C₂H₅OH and CH₃COOH) of 88.4%. The S motif is stable before and after the ECO₂R, as evidenced by the systematic characterization (Fig. 17c–h). The researchers explain S-HKUST-1's superior selectivity as follows: the approximate distance between Cu⁰ and Cu^{δ+} favours the *CO dimerization step at the interface of Cu/Cu₂S. Also, a high S-concentrate electrocatalyst (*i.e.*, Cu₂S) showed a significant Cu–Cu distance that practically restricted *CO dimerization. Compared to a pure metallic Cu catalyst, the Cu/Cu₂S interface decreased to *CO species binding energy at the surface and fastened *CO dimerization (Fig. 17i–m). Similarly, in 2023, Yu *et al.*¹⁵⁵ constructed a Cu₂S nanocrystal on Cu nanosheets (Cu–Cu₂S), as shown in Fig. 18a, for ECO₂R to C₂H₅OH. The author's design focuses on attaining three discrete features: (i) the nanocomposites produce a positive charge locally on Cu (Cu^{δ+}) to offer multi-active sites during ECO₂R; (ii) the evenly dispersed tiny Cu₂S on Cu generate interfaces, as illustrated in Fig. 18b–g, between Cu^{δ+} and Cu in the zero-valence state Cu⁰; and (iii) the uneven and stepped Cu–Cu₂S surface offers a spatially advantageous arrangement for C₂H₅OH production. Due to these structural features, a total FE of 90% for C₂₊ products (a FE of 6% for C₂H₅OH and a FE of 15% for C₂H₄) with a partial current density of 45 mA cm^{−2} at −1.2 V_{RHE} was attained in a H-cell configuration. According to the *in situ* spectroscopy and DFT investigations, as illustrated in Fig. 18h–l, they showed that the three characteristics of Cu–Cu₂S work together to (1) improve CO₂ adsorption by enabling high electronic conductivity, (2) facilitate the adsorption of the *CO intermediate on the Cu–Cu₂S surface, (3) lower the energy barrier for forming the *COCO intermediate, and (4) make the reaction path more thermodynamically viable for C₂H₅OH over C₂H₄. This study emphasizes the potential for commercializing alcohol and related product formation from CO₂ by proposing an effective method and the underlying mechanism for the significant increase of ECO₂R to C₂H₅OH conversion.

Mosali *et al.*¹⁵⁶ developed an array of Cu_xZn_yS nanoparticles (Fig. 19a–l) with varying Cu:Zn compositions and that are electrochemically stabilized, resulting in SD-Cu_xZn_yS catalysts. The catalysts showed composition-dependent product selectivity during ECO₂R. Operating in an H-cell configuration, it was discovered that SD-Cu_xZn_yS catalysts with a higher copper content exhibited a significantly higher FE_{CH₄} than the Zn-rich catalysts. In contrast, Zn-rich counterparts produced CO as the primary ECO₂R product. The best composition of 1:1 Cu:Zn resulted in highly selective CH₄ as the main product, with an FE of 76 ± 3% at −0.98 V_{RHE}. The current density increased after switching to a flow cell arrangement for ECO₂R in 1.0 M KHCO₃ or 1.0 M KOH electrolyte. However, the selectivity changed to produce syngas, with molar ratios of 2:3 to 3:2 for the formation of H₂/CO. This tendency holds irrespective of the applied potential, particularly in the 1.0 M KOH electrolyte. An *ex situ* investigation revealed a significant reduction in CuS in electrocatalysts with higher copper content. In addition, pre-reduction of CuZnS catalysts and time-dependent observations showed the importance of copper's higher oxidation state and the interaction of partially reduced

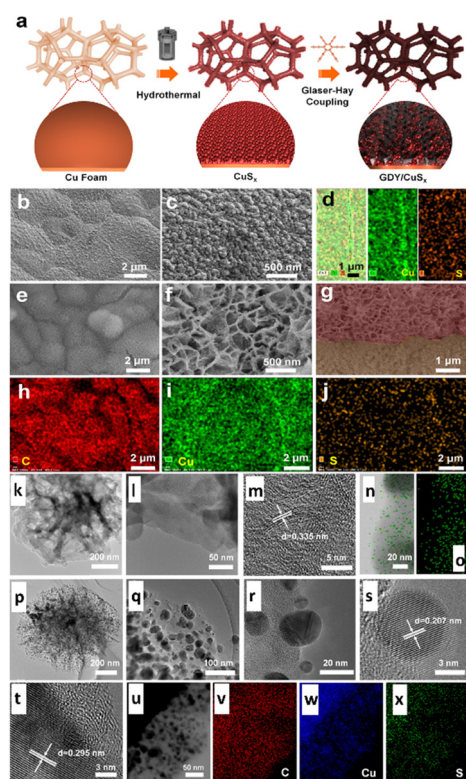
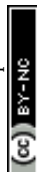


Fig. 16 (a) Schematic representation of the synthesis route of GDY/CuS_x; and (b) low- and (c) high-magnification SEM images of CuS_x. (d) Energy dispersive spectroscopy mapping of CuS_x. (e) Low- and (f) high-magnification SEM images of GDY/CuS_x. (g) Cross-sectional SEM image of GDY/CuS_x; and (h)–(j) energy dispersive spectroscopy mapping of GDY/CuS_x. (k) TEM and (l) and (m) HRTEM images of the GDY nanosheet. (n) STEM image (left) and (o) the corresponding elemental mapping (right) of the elemental carbon of GDY. (p)–(r) TEM and (s) and (t) HRTEM images of GDY/CuS_x. (u) STEM image and the corresponding elemental mapping of (v) C, (w) Cu, and (x) S elements of the GDY/CuS_x electrode (reproduced with permission from ref. 153, copyright RSC 2023).



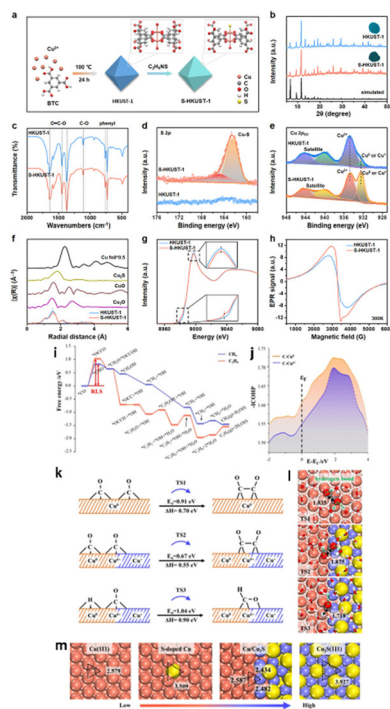


Fig. 17 (a) Schematic diagram of the preparation of S-HKUST-1. (b) XRD patterns of the prepared S-HKUST-1 and HKUST-1 precatalysts are well indexed to the simulated HKUST-1. The insets are digital photos of S-HKUST-1 (dark green) and HKUST-1 (blue). (c) FTIR spectra of S-HKUST-1 and HKUST-1 precatalysts indicate negligible differences. The XRD and FTIR results prove that the long-range ordered structure is not destroyed after S incorporation. (d) and (e) High-resolution XPS spectra of S-HKUST-1 and HKUST-1 precatalysts in the (d) S 2p region, showing the characteristic Cu₂S bonds in the S-HKUST-1 precatalyst and (e) Cu 2p region, showing the increased content of Cu^{δ+} or Cu⁰ species. (f) FT of the EXAFS spectra and (g) Cu K-edge XANES spectra of HKUST-1 and S-HKUST-1 precatalysts. The inset in (g) is the magnified image. The XAFS results in (f) and (g) prove the successful incorporation of local heteroatoms, which might be bonded to Cu atoms in MOFs. (h) EPR spectra of the samples measured at 300 K. (i) Calculated free energy profiles for ECO₂R to CH₄ and C₂H₄ over a pure Cu(111) surface, indicating that the initial *CO hydrogenation and dimerization steps determine the CH₄ and C₂H₄ product distribution. (j) Integrated crystal orbital Hamilton population (ICOHP) curves of Cu^{δ+}-CO and Cu⁰-CO bonds. (k) The reaction barriers together with enthalpies and (l) the corresponding transition state configurations for *CO dimerization and hydrogenation over Cu(111) and Cu/Cu₂S surfaces, respectively. Yellow, red, grey, white, orange and blue balls refer to S, O, C, H, Cu⁰ and Cu^{δ+} atoms, respectively. (m) Surface configuration (top view) of Cu-based structures with different degrees of vulcanization. The distances between two neighbouring Cu atoms on different surfaces are given in b (reproduced with permission from ref. 154, copyright Wiley-VCH GmbH 2021).

CuS and ZnS present in the catalysts in attaining methane selectivity. These findings demonstrate that copper's oxidation state is crucial in determining selective product formation using copper-based catalysts. Likewise, the same research group created S-derived copper-cadmium (SD-Cu_xCd_y) electrocatalysts earlier, where *x* and *y* denote the molar ratio of Cu/Cd to assist in generating vital intermediates.¹⁵⁷ In 0.1 M KHCO₃ solution, the SDCuCd₂ catalyst observed 32% ethanol selectivity at a low overpotential of -0.89 V_{RHE}. They exhibited selective ethanol production at lower overpotentials with the

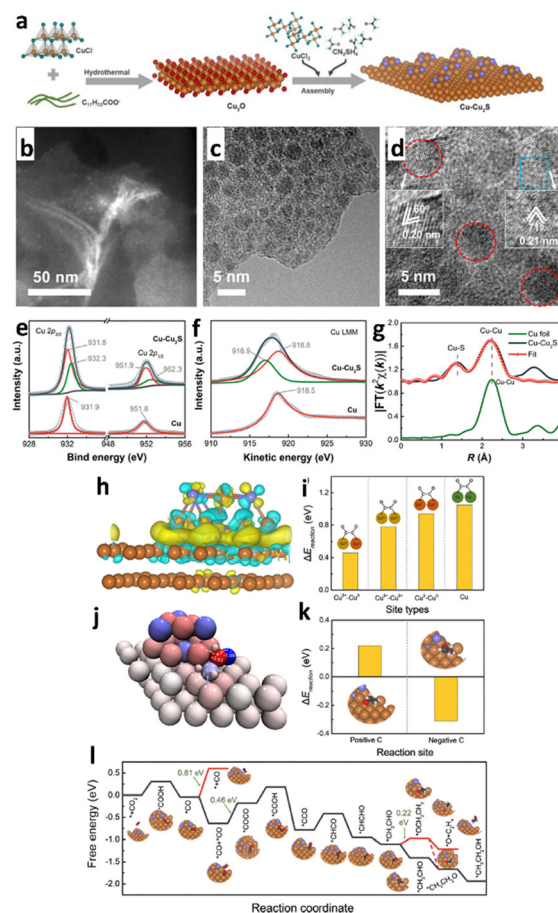
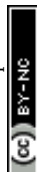


Fig. 18 (a) Schematic illustration of the synthesis of Cu-Cu₂S. Characterization of the Cu-Cu₂S sample. (b) and (c) TEM, (d) high-resolution TEM (HRTEM), (e) high-resolution Cu 2p XP spectra, (f) Cu LMM Auger spectra, and (g) Cu K-edge *K*₂-weighted $\chi(k)$ function of Fourier transform extended X-ray absorption fine structure (EXAFS) spectra of the sample, where pristine Cu and Cu foil are used as controls in panels (e)–(g), revealing that both Cu(0) and Cu(I) states are involved, respectively. Calculated ECO₂R on Cu-Cu₂S. (h) DFT-calculated charge density difference of Cu-Cu₂S (the Cu and S atoms are in brown and violet, respectively; the yellow and cyan colours represent charge accumulation and depletion, respectively; the isosurface value is 0.002 e Å^{−3}). (i) Energy barrier ($\Delta E_{\text{reaction}}$) for two *CO forming one *OCCO via C-C coupling at Cu^{δ+}-Cu^{δ+}, Cu^{δ+}-Cu^{δ+}, and Cu^{δ+}-Cu⁰ sites of Cu-Cu₂S and at the pristine Cu surface, where Cu^{δ+} and Cu⁰ of Cu-Cu₂S and Cu of pristine are in brown, red, and green, respectively, showing that the energy barrier at the Cu^{δ+}-Cu⁰ site of Cu-Cu₂S is the lowest. (j) Bader charge of *CH₂CHO adsorbed on Cu-Cu₂S, where the colour ranging from blue to red indicates the negative to the positive charge of the atoms, respectively. (k) $\Delta E_{\text{reaction}}$ for hydrogenation on the positively charged C and negatively charged C of *CH₂CHO converting to CH₂CH₂O* and CH₃CHO*, respectively. (l) Calculated free energy for each step of the reaction pathway converting CO₂ to C₂H₅OH on Cu-Cu₂S (reproduced with permission from ref. 155, copyright ACS 2023).

best-performing SD-CuCd₂ electrocatalyst of Cu₃Cd₁₀alloy/Cu₂S/CdS phased composites, as illustrated in Fig. 19m and n.

Xi and colleagues¹⁵⁸ recently developed and built CeO₂/CuS (Fig. 20a–k) with a -0.5 V_{RHE} overpotential for ECO₂R to ethanol and a FE_{C₂₊} of 75% at -0.8 V_{RHE}. During ECO₂R, they investigated catalysts' structural development, interactions, and activity sources, as illustrated in transmission electron



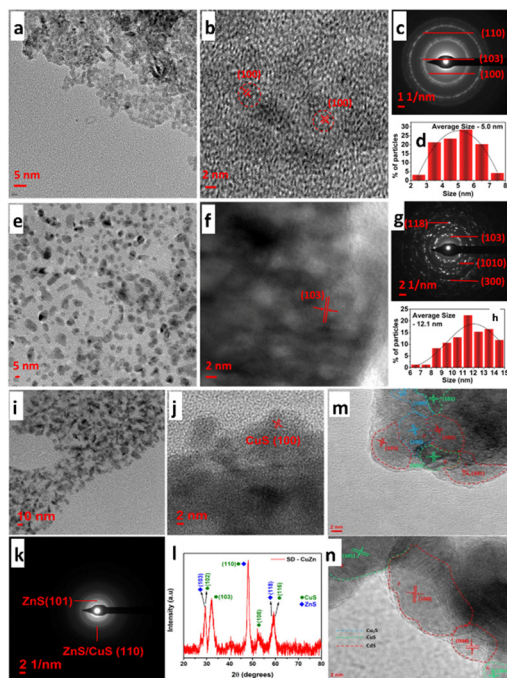


Fig. 19 (a) and (e) TEM images, (b) and (f) HR-TEM images, (c) and (g) SAED pattern and (d) and (h) particle size distribution of (a)–(d) CuS and (e)–(h) ZnS nanoparticles. (i) TEM images, (j) HR-TEM images, (k) SAED pattern, and (l) XRD powder pattern of SD-CuZn nanoparticles obtained upon pre-reduction at -0.80 V vs. RHE in 0.10 M KHCO_3 solution saturated with CO_2 (blue diamonds (♦) represent the peaks corresponding to ZnS, and green circles (●) represent the peaks corresponding to CuS; reproduced with permission from ref. 156, copyright Elsevier 2024). HR-TEM images of (m) CuCd_2S ; (n) Cu_2CdS showing boundary distributions of CdS, CuS and Cu_2S (reproduced with permission from ref. 157, copyright Wiley-VCH GmbH 2023).

microscopy (TEM) and high-resolution transmission electron microscopy (HRTEM) images (Fig. 20b–d). The fast electron movement route provided by CeO_2 regions in catalysts prevents electron aggregation around Cu^+ sites, preserving Cu^+ sites during ECO_2R . The experimental *in situ* studies and DFT calculations reveal that altering CeO_2 on CuS thermodynamically reduces the production energy for $^*\text{COCHO}$ compared to CuS nanoplates. The fast water molecule activation near CeO_2 speeds up the synthesis of $^*\text{COCHO}$. As a result, the C–C coupling is accelerated *via* the $^*\text{CHO}$ route, providing CeO_2/CuS catalysts with exceptional electrocatalytic efficacy towards C_{2+} products. Wang *et al.*¹⁵⁹ proved that heterointerfaces of Bi/ CeO_2/CuS nanohybrids (Fig. 20l–s) can be advantageously employed as a highly efficient and selective catalyst for ECO_2R to produce formate over a wide potential range. The nanohybrid demonstrated excellent activity at -0.9 V_{RHE}, with a formate efficiency of 88% and a current density of -17 mA cm⁻². The morphological analysis showed no change in the structure after a one-hour reaction at -0.9 V_{RHE}, and the FE of formate remained stable at that potential. The Bi/ CeO_2/CuS heterostructure substantially decreases the formation energy barrier of OCHO^* intermediates because, during electrolysis, Ce^{4+} rapidly suffers Ce^{3+} reduction, forming a conductive network of

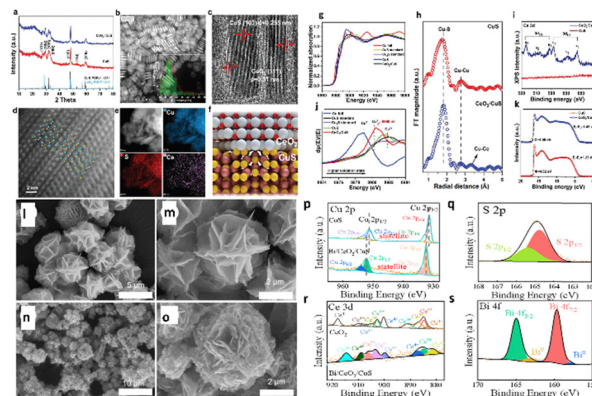
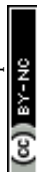


Fig. 20 Crystal structures and composition of CuS and CeO_2/CuS . (a) XRD patterns of CuS and CeO_2/CuS . (b) TEM image of CeO_2/CuS (inset: size distribution). (c) HRTEM image of CeO_2/CuS . (d) The HAADF-STEM image of CeO_2/CuS (inset: CuS 9 000 523. cif). (e) The HAADF-STEM image and corresponding EDS element mapping of CeO_2/CuS . (f) Structural representation of CeO_2/CuS , white: Ce atoms; red: O atoms; yellow: S atoms; gold: Cu atoms. Electronic structures of CuS and CeO_2/CuS . (g)–(i) X-ray absorption spectra of CeO_2/CuS , CuS catalyst, standard Cu_2S , and Cu foil. (j) Normalized Cu K-edge XANES. (k) The first derivative $\mu(E)/E$. (l) Fourier-transformed $k^2\chi(k)$ of CeO_2/CuS (lower panel) and CuS (upper panel). (m) Ce 3d XPS spectra. (n) UPS spectra (reproduced with permission from ref. 158, copyright Wiley-VCH GmbH 2023). SEM images of (l) and (m) CuS and (n) and (o) Bi/ CeO_2/CuS . XPS patterns of Bi/ CeO_2/CuS : (p) Cu 2p, (q) S 2p, (r) Ce 3d, and (s) Bi 4f (reproduced with permission from ref. 159, copyright MDPI 2024).

$\text{Ce}^{4+}/\text{Ce}^{3+}$, resulting in high activity and selectivity of ECO_2R to formate. Overall, the system increased electron mobilization, stabilized Cu^+ species, and improved CO_2 adsorption and activation over the catalytic surface. Furthermore, sulfur boosts the transformation of OCHO^* to formate.

Han *et al.*¹⁶⁰ employed a simple solution-phase method to synthesize a ternary metal–metal sulfide Bi– Cu_2S heterostructure electrocatalyst (Fig. 21a and b). Due to the high synergistic and interfacial effects between Bi and Cu_2S , the heterostructure showed a lower overpotential (240 mV) than Bi with an outstanding FE_{formate} of >98% and 2.4 and 5.2 times higher partial current density than bare Cu_2S and Bi at -1.0 V_{RHE}. According to the theoretical study, the HCOOH generation was aided by stabilizing the $^*\text{OCHO}$ intermediates over $^*\text{COOH}$ and $^*\text{H}$ due to the higher electron transfer rate between the Bi and Cu_2S interfaces (Fig. 21c–i).

Prasanna *et al.*¹⁶¹ designed a rationally created novel heterostructure of CuS decorated NH_4^+ ion incorporated stable 1T- WS_2/WO_3 using a simple hydrothermal method followed by a reflux technique (Fig. 22a). The nanohybrid favoured reaction mechanism through the $^*\text{OCHO}$ route, which reduces (by obtaining $\text{H}^+ + \text{e}^-$) to HCOO^- (HCOO^- pathway) as the main product. Thus, $\text{CuS}@1\text{T-N-WS}_2/\text{WO}_3$ nanohybrids yielded $55.6\% \pm 0.5$ at -1.3 V_{RHE} and a j_{geo} of -125.05 mA cm⁻². Interestingly, from their analysis, the intercalation of NH_4^+ ions with effective surfaces that donate WO_3 and accept protons from WS_2 stabilized the metallic 1T phase. Besides, an effective hydrogen spillover mechanism in the stable heterointerface of 1T-N- WS_2/WO_3 may provide kinetic support for CuS active



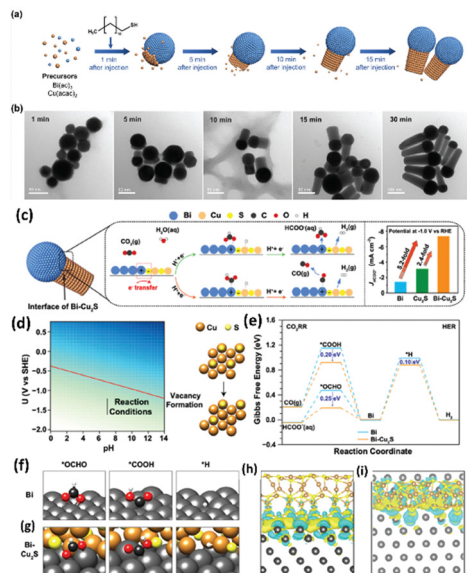


Fig. 21 (a) Schematic illustration of the nucleation and growth of the heterostructured Bi–Cu₂S nanocrystals. (b) TEM images of Bi–Cu₂S at different time intervals after introducing 1-dodecanethiol at 220 °C. (c) Proposed CO₂ reduction mechanism on the Bi–Cu₂S interfacial system. (d) Pourbaix diagram showing the sulfur vacancy formation. (e) Free energy diagrams of the ECO₂R and HER on Bi(001) and Bi–Cu₂S model systems. Optimized geometry structures of key intermediates (*OCHO, *COOH, and *H) on Bi(001) and Bi–Cu₂S systems are shown in (f) and (g), respectively (dark grey, brown, yellow, black, red, and white spheres denote Bi, Cu, S, C, O, and H atoms, respectively). (h) and (i) are the top and front views of the charge density difference for the Bi–Cu₂S interfacial surface, respectively. Cyan corresponds to an isosurface of $-0.001 \text{ e Bohr}^{-3}$ and yellow corresponds to an isosurface of $+0.001 \text{ e Bohr}^{-3}$ (reproduced with permission from ref. 160, copyright RSC 2022).

centres, leading to notable improvements in product efficiency and selectivity (Fig. 22b).

Guo *et al.*¹⁶² synthesized transition metal sulfide (TMS)-supported CuS catalysts with microflower-shaped frameworks *via* a facile hydrothermal approach for ECO₂R to CO (Fig. 22c). They investigated the effect of doping TMSs (*i.e.*, ZnS, Bi₂S₃, and MoS₂) with CuS on ECO₂R performance. All the heterostructured catalysts, *i.e.*, ZnS–CuS, Bi₂S₃–CuS and MoS₂–CuS, exhibited flower-shaped morphology, with the doped TMSs (*i.e.*, ZnS, Bi₂S₃, and MoS₂) attached as microcrystals on the surface of petals. On the other hand, strong Cd²⁺ and Cu²⁺ interactions with S^{2−} caused CdS and CuS to aggregate and expand disorderly, forming incomplete sphere-like cage-shaped microstructures. Electrochemical experiments revealed that TMS-supported CuS catalysts outperformed pristine CuS in terms of ECO₂R activity, but their CO₂ conversion rate was reduced. Decorating CuS with MoS₂ resulted in a flower-shaped nanomorphology (Fig. 22d and e), enhancing the catalyst's CO selectivity from 72.67% to 83.20% at $-0.6 \text{ V}_{\text{RHE}}$. With the increase in applied potential from -0.2 to $-1.0 \text{ V}_{\text{RHE}}$, CO selectivity initially improved and then decreased, but the CO₂ conversion rate increased considerably from 0.2% to 21.9%. It has been observed that, during a 300-minute electrolysis at $-0.6 \text{ V}_{\text{RHE}}$, the target MoS₂–CuS catalyst exhibited constant ECO₂R activity, with CO selectivity

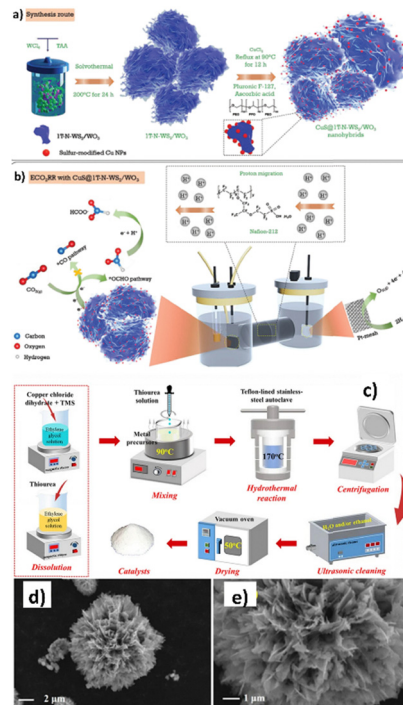


Fig. 22 (a) Schematic illustration of the synthetic process of 1T-N-WS₂/WO₃ and CuS@1T-N-WS₂/WO₃ nano hybrids using a simple solvothermal process and reflux method. (b) Proposed schematic illustration of the ECO₂R in an H-type cell (reproduced with permission from ref. 161, copyright Wiley-VCH GmbH 2023). (c) Schematic diagram of the synthesis procedures and formation of the pristine CuS and TMS-decorated CuS microflower-like structures; SEM images of the MoS₂–CuS catalysts (d) at a low magnification and (e) at a high magnification (reproduced with permission from ref. 162, copyright Springer Nature 2024).

being improved over a limited range of 73.6–88.5% and j_{CO} remaining stable at about 3.6 mA cm^{-2} . The authors attributed the increased selective CO formation linked to the synergistic effect of the generated GBs, undercoordinated S-vacancies, and edge-exposed Mo sites in boosting CO₂ activation, stabilizing *COOH adsorption, facilitating *CO desorption, and reducing the energy barrier of the potential-limiting step.

3.6 Heteroatom doping effect for improving electrocatalytic performance

Doping is widely regarded as an effective method for modulating the electronic structure of a catalyst through charge redistribution.^{163,164} Doping different elements into the lattice of the pristine material can effectively modulate the electronic and physiochemical properties to fine-tune its electrocatalytic activity.^{165,166} Also, heteroatom doping has little impact on crystal structure disruption, so the catalyst's main composition remains unchanged. It has been observed that the doping approach can increase the adsorption/desorption energy of the reaction intermediates.^{167,168} The doping strategies are divided into metal and nonmetal doping, as described below.

3.6.1 Metal doping. Metal doping can efficiently catalyze the C–C coupling during ECO₂R. Very recently, Hu *et al.*¹⁶⁹ constructed Mo⁴⁺-doped CuS nanosheet-assembled hollow



spheres for ECO₂R to ethanol in a flow cell configuration. They revealed that doping with Mo⁴⁺ ions makes the movement of the electrons faster from Cu to S, optimizing the surrounding electrical conductivity and broadening the CuS binding sites, resulting in increased coverage of linear *CO_L on the Cu site away from the Mo site, producing bridge *CO_B on the Cu site near the Mo site and activating asymmetric C–C coupling of *CO_L–*CO_H for ethanol generation. Additionally, the highly electron-dense Cu site away from the Mo site favours the cleavage of Cu–O, promoting the hydrogenation of CH₃CH₂O* to ethanol. Moreover, the nanosheet-assembled hollow sphere structure and hydrophobic Mo⁴⁺-doped CuS were prepared by a GDE assisted flow-cell procedure. The best Mo_{9.0%}–CuS achieved an ethanol FE of 67.5% and a *j*_{ethanol} of 186.5 mA cm^{−2} at −0.6 V_{RHE} in a flow cell, with good stability over 26 h of continuous operation. In this work, high valence transition metal ions are doped into Cu-based sulfides to modify the coverage and configuration of related intermediates for ethanol production in a flow cell, thereby developing Cu-based materials for catalyzing ECO₂R to ethanol. In another work, a Cu-doped Bi (CDB) nanosheet was created by reconstructing a CuS–Bi₂S₃ heterojunction precursor *in situ* (Fig. 23a).¹⁷⁰ The *in situ* restructuring of the as-synthesized CuS–Bi₂S₃ heterojunction precursor, as shown in Fig. 23b–g, resulted in an enhanced CDB catalyst with an industrially acceptable ampere-level current density and remarkable stability for the formate production (Fig. 23h–m). The CDB surface with a high concentration of electrons resulting from an increase in the Fermi level not only helped to stabilize the *OCHO intermediate but also successfully lowered the rate-determining step's reaction kinetic barrier, leading to the exceptional electrocatalytic activity of the CDB catalyst. At the same time, the more excellent selectivity to formate can be due to the reduced Δ*G*_{OCHO*} on the CDB surface relative to Δ*G*_{COOH*} and Δ*G*_{H*}. Furthermore, the effective antioxidation capability of the CDB electrocatalyst allows it to exhibit outstanding stability beyond 100 hours at a high current density of 400 mA cm^{−2} in MEA cells (Fig. 22n).

3.6.2 Nonmetal doping. Many studies have also focused on using nonmetal doping to tune the electronic structure of the electrocatalysts. Nonmetals such as N and S have significantly increased ECO₂R catalytic activity. For example, Liang *et al.*¹⁷¹ conducted theoretical and experimental studies to investigate catalytic reactivity effects on N-doped Cu sulfides using the generalized morphology engineering principle of asymmetric metal pair sites, which promotes C–C coupling on electrode surfaces. First, a computational N-doped Cu₂S model layer featured electronically asymmetric Cu^{δ1+}–Cu^{δ2+} (0 < δ₁⁺ < δ₂⁺ < 1) metal pairs evidenced by considering their non-uniform charge distribution. The asymmetric electronic structure led to distinct CO adsorption energies and the associated self-adjusting structures, which significantly lowered C–C coupling energy barriers. Later, they experimentally validated the computational hypotheses using XPS of Cu–N moieties within N-doped Cu₂S layers. *In situ* FTIR spectroscopy confirmed linear *CO and *CO–CO adsorption configurations from the ≈2080 and 1920 cm^{−1} broad peaks, respectively. After N-doping, the

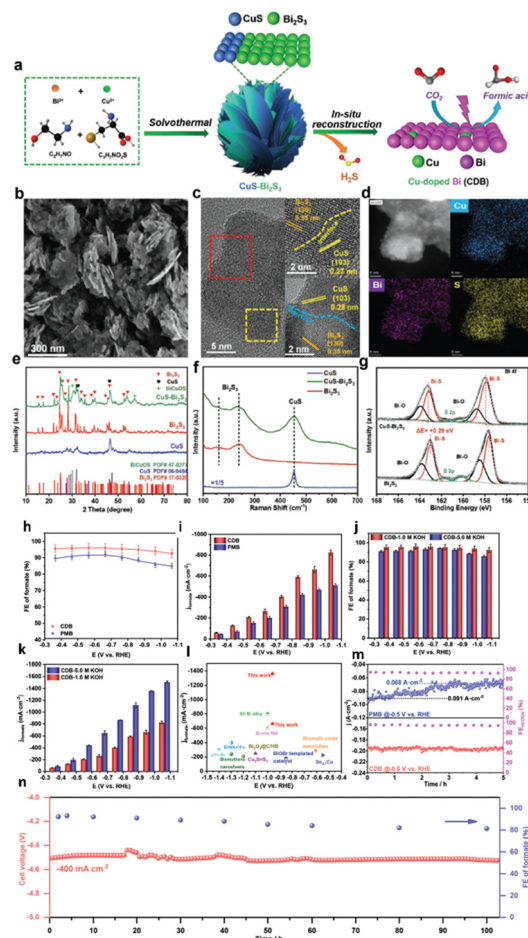
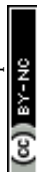


Fig. 23 (a) Synthetic strategy of a CuS–Bi₂S₃ nano-heterojunction precursor. (b) SEM images of a CuS–Bi₂S₃ nano-heterojunction precursor. (c) HR-TEM images of a CuS–Bi₂S₃ nano-heterojunction precursor. (d) Corresponding elemental mapping images of S, Cu, and Bi elements. (e) X-ray diffraction (XRD) patterns of CuS, Bi₂S₃, and CuS–Bi₂S₃. (f) Raman spectrum of CuS, Bi₂S₃, and CuS–Bi₂S₃. (g) The high-resolution XPS spectra of Bi 4f for Bi₂S₃ and CuS–Bi₂S₃. ECO₂R measurements in the lab-made flow cell. (h) FE of formate and (i) formate partial current densities at different electrolytic potentials in 1 M KOH electrolyte with CDB and PMB electrocatalysts. (j) FE of formate and (k) formate partial current densities of the CDB electrocatalyst at different electrolytic potentials in 1 M and 5 M KOH electrolytes. (l) Performance comparison of the CDB electrocatalyst with the reported electrocatalysts towards ECO₂R to formate in the flow cell. (m) Stability test of PMB and CDB electrocatalysts at the potential of −0.5 V_{RHE} for 5 h. (n) Continuous 100 h electrolysis in a 5 cm² MEA electrolyzer under a constant total current density of −400 mA cm^{−2}. The error bars in (f)–(h) represented the standard deviations of three independent measurements (reproduced with permission from ref. 170, copyright Wiley-VCH GmbH 2022).

catalytic C₂ faradaic efficiency can significantly be elevated to 14.72% due to the promotion of C–C coupling (Fig. 24a).

In another work, Li *et al.*¹⁷² demonstrated the effective modulation of ECO₂R pathways by designing and synthesizing three kinds of copper sulfides (Fig. 24b). Among all, they observed that SNC@Cu₂S (Cu₂S coated sulfur, nitrogen-doped carbon) without Cu vacancies displays a high FE for formate production. In contrast, the other two catalysts,



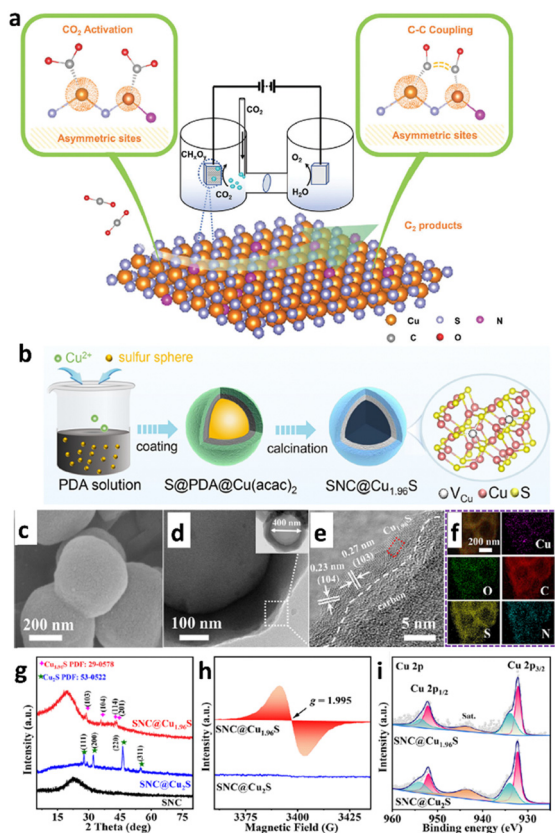


Fig. 24 (a) Illustration of the key hypothesis: the rate of C–C coupling by CO dimerization can be manipulated using structurally asymmetric metal^{δ1+}–metal^{δ2+} pair sites ($0 < \delta_1^+ \neq \delta_2^+$). A well-ordered CuS_x surface facet with surface Cu atoms adsorbing CO₂ molecules is shown. Dopant atoms introduce electronic asymmetries in Cu site pairs by adjusting electron-withdrawing/donating properties. Asymmetric CO chemisorption energies favor CO dimerization and hence ECO₂R to C₂ products. By contrast, sites with even electronic distributions feature strong dipole–dipole repulsion forces during CO₂ activation, which seriously hinders CO dimerization (reproduced with permission from ref. 171, copyright Wiley-VCH GmbH 2022). (b) Schematic showing the synthetic process of SNC@Cu_{1.96}S. Characterization of Cu vacancies in SNC@Cu_{1.96}S. (c) SEM image, (d) TEM image, (e) HR-TEM image, and (f) the elemental mapping images of SNC@Cu_{1.96}S. (g) XRD patterns of SNC@Cu_{1.96}S, SNC@Cu₂S, and SNC. (h) EPR spectra and (i) Cu 2p XPS spectra of SNC@Cu_{1.96}S and SNC@Cu₂S (reproduced with permission from ref. 172, copyright ACS 2022).

SNC@Cu_xS and SNC@Cu_{1.96}S, with Cu vacancies, generated CO as the primary product (Fig. 24c–i). Firstly, the Cu vacancies present in Cu_{1.96}S modify the S sites' electronic structures and significantly increase the energy barrier of H* formation. At the same time, the Cu vacancies generate the appropriate binding energy for the *COOH intermediate while attenuating the adsorbate–metal interaction. These advantageous characteristics, when combined, result in the favoured formation of CO over formate. Furthermore, the Cu vacancies lower charge transfer resistance, enrich the electronic structure of active sites and increase CO₂ adsorption capacity, enhancing the ECO₂R activity of SNC@Cu_{1.96}S.

Huang and his colleagues¹⁷³ synthesized Cu₂O-derived Cu catalysts doped with sulfur by immersing the Cu substrates in ammonium polysulfide solutions. Regarding cost, availability,

longevity, and catalytic efficacy, the catalyst outperforms alternative materials for ECO₂R to formate. Their XPS, ToF-SIMS, and μ XRF analyses revealed a positive correlation between catalyst sulfur content and formate production. According to their findings, among all the prepared electrocatalysts, Cu-5000S with 2.7% atomic sulfur has satisfactory catalytic activity with a faradaic efficiency of 75% and a current density of -13.9 mA cm^{-2} at $-0.9 \text{ V}_{\text{RHE}}$ for formate selectivity. Moreover, Cu-5000S exhibits outstanding stability for prolonged CO₂ reduction, as evidenced by its formate FE retained at approximately 75% for 12 hours. When sulfur doping was increased, formate production outpaced HER activity, as demonstrated by comparing the SRF-normalized j_{formate} and j_{H_2} . According to investigations of mechanisms, sulfur-doped Cu reduces the binding energy of *COOH intermediates for CO generation, facilitating the synthesis of HCOOH. This work demonstrates that sulfur doping is an effective strategy for enhancing the catalytic selectivity of Cu towards formate and expands the material choice for producing this commercially valuable fuel and chemical.

Li *et al.*¹⁷⁴ examined the influence of the doping strategy on the ECO₂R to ethylene using S-doped spherical coral-like CuO catalysts. With a current density of 15.5 mA cm^{-2} , the as-synthesized 5% S–CuO observed a good FE_{C₂H₄} of 48.4% at $-1.3 \text{ V}_{\text{RHE}}$ compared to pure CuO. The 5% S–CuO also retained a consistent FE for C₂H₄ upon long-term stability testing, with no appreciable drop in current density. The superior performance of 5% S–CuO was derived from the improved dynamic barriers of *CO intermediate dimerization at the Cu active site caused by S-doping, according to the experimental analysis and contact angle measurements. In addition, the hydrophobicity of the 5% S–CuO surface prevented accessible H₂O molecules from interacting, thereby preventing competitive HER.

Wang *et al.*¹⁷⁵ discovered various new morphologies of Cu₂S–X materials reduced to S–Cu₂O–X during ECO₂R in 2022. The researchers found that the optimized electronic structure, aided by the S dopant and microstructure reconstruction, resulted in a large surface area, critical in improving the ECO₂R performance and formate selectivity of the S–Cu₂O–X catalysts. In addition, as compared to Cu(OH)₂ and desulfurized Cu₂O materials, S–Cu₂O–X ($X = 6, 10, \text{ and } 14$) catalysts showed significantly higher catalytic activity in the formation of formate (60–70%). The S–Cu₂O–14 catalyst demonstrated a partial current density of 16 mA cm^{-2} at $1.0 \text{ V}_{\text{RHE}}$ and a FE of 67.2% for formate selectivity with a stability of 20 h among all evaluated electrocatalysts. DFT studies demonstrated that S-doped Cu(111) and S-vacancy (V_s) species promoted COOH* or OCHO* intermediates, which accelerated selectivity towards HCOOH pathways.

Recently, Zhang *et al.*¹⁷⁶ showed that S-doped Cu derived from hierarchical hollow-like CuS polyhedron (CuS-HP) nanostructures synthesised from a MOF significantly improves ECO₂R performance in neutral pH water environments (Fig. 25a–j). According to their findings, during electrolysis, the CuS-HP was gradually transformed into an S-doped Cu as the reduction process went on. The *in situ* formed electrocatalyst





Fig. 25 Structural characterization of CuS-HP. (a) Schematic illustration of the synthetic process. (b) XRD pattern. (c)–(f) TEM images. (g)–(j) HAADF-STEM image and the corresponding EDS elemental mapping images (reproduced with permission from ref. 176, copyright Chinese Chemical Society 2021). (k) SEM images, (l) TEM images, and (m) XPS spectra of (i) fresh CuS_x , (ii) AC- CuS_x , and (iii) AC- CuS_x samples after 40 min of ECO_2R in 0.1 M KHCO_3 at $-0.85 V_{\text{RHE}}$ and (iv) the DS- CuS_x sample (reproduced with permission from ref. 177, copyright ACS 2018).

exhibited a stability of 36 hours at a j_{formate} of 16 mA cm^{-2} at $-0.6 V_{\text{RHE}}$ with a FE of $>90\%$. The DFT study indicates that the Cu(111) facet in S-doped Cu lowers formate activation energy barriers while inhibiting the HER and the reaction intermediates remain unchanged. This work comprehensively explains the mechanisms for improving formate selectivity using CuS electrode materials. Deng *et al.*¹⁷⁷ have experimentally and theoretically demonstrated that sulfur-doped Cu catalyzes the ECO_2R to formate. They synthesized an active CuS_x (AC- CuS_x) catalyst, which exhibited a high FE of 75% and a partial current density of 9 mA cm^{-2} at $-0.9 V_{\text{RHE}}$ toward formate production (Fig. 25k–m). It was discovered that sulfur dopants were the key to the increased formate production on the AC- CuS_x surfaces. *Operando* Raman spectroscopy found that S dopants on the catalyst inhibited the CO intermediate formation during ECO_2R , with a lower FE for other products (*i.e.*, CO, hydrocarbons, and alcohols). DFT calculations validated the Raman band assignments. They observed that the adsorption strengths of adsorbed HCOO^* were modified by the presence of sulfur on the copper surface, which supported the formation of formate while inhibiting the formation of $^*\text{COOH}$, the CO intermediate.

3.7 Catalyst support materials for improving electrocatalytic performance

The supporting materials significantly impact the catalytic performance of electrocatalytic materials. In particular, TMD NSs show promise as supporting materials.^{178,179} Composite

catalysts have high catalytic performance due to the synergistic interaction of TMDs, supporting materials, and electrocatalysts.^{180,181} Kahsay *et al.*¹⁸² successfully deposited copper sulfide nanoparticles onto thermally synthesized copper oxide using a simple and facile SILAR method and examined its catalytic activity for ECO_2R (Scheme 1 in Fig. 26). The modified nanocomposites exhibited high selectivity for formate formation at low overpotential. Remarkably, a maximum faradaic efficiency of 84% and an enhanced partial current density of -20 mA cm^{-2} were obtained at an overpotential of $-0.7 V_{\text{RHE}}$. It was observed that copper sulfides undergo phase changes during ECO_2R , which can contribute to the enhanced electrocatalytic activity. As a result, together with copper oxides, a catalytic synergy is created in the composite, and more favourable adsorption sites are generated to facilitate the ECO_2R . This study paves the way for controlling the composition–selectivity relationship using a facile and scalable catalyst synthesis approach. Also, Li *et al.*¹⁸³ produced extremely porous $\text{Cu}_2\text{O}/\text{CuS}$ nanocomposites, as evidenced by X-ray diffraction (XRD), Raman spectroscopy and scanning electron microscopy (SEM) (Fig. 26a–i), showing a better formate FE of 67.6% at a j_{HCOOH} value of 15.3 mA cm^{-2} at $-0.5 V_{\text{RHE}}$ in ECO_2R . More significantly, with the same applied potential, the current density remained unchanged for at least 30 hours with an average FE of 62.9%. DFT modelling showed that $\text{CuS}(110)$ facets favour HCOOH over CO, with substitutional surface O_s or vacancy V_s species expected to result in lower onset potentials and increased catalytic activity, with both COOH^* - and OCHO^* -mediated ECO_2R routes predicted to be involved. As a result, this increases the yield of HCOOH with a better current density than the Cu, CuS, and Cu_2O electrocatalysts.

4. Theoretical studies of Cu/S-based ECO_2R Catalysts

As knowledge of electrocatalysis grows, scientists are no longer satisfied with investigating macroscopic phenomena in experimentation; instead, they are exploring the complex microscopic world to gain insight into the mechanism underlying electrocatalytic reactions. This effort is significant for the design and evaluation of highly efficient electrocatalysts. The rapid development of supercomputers has resulted in significantly greater computational speed and performance, enhancing theoretical calculations. In addition, modelling and simulated environments are now closely consistent with the experimental surroundings, resulting in accurate results.^{184–186} Nowadays, computational models are used not only to shed light on experimental phenomena but also to direct research strategies and the development of efficient electrocatalysts.^{184–186}

Inspired by this, Fonseca *et al.* discussed theoretically how the tuning ratio of Cu/S on 2D CuS_x nanomaterials influenced the ECO_2R performance.¹⁸⁷ Their findings demonstrate that the adsorption modes and strength of the ECO_2R intermediate on CuS_x monolayers vary considerably based on the Cu/S ratio, resulting in different catalytic activities. For instance, $\text{CuS}_{0.5}$

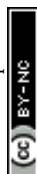




Fig. 26 Scheme 1: illustration of thermal oxidation and SILAR growth: (i) copper mesh before thermal oxidation, (ii) $\text{Cu}_2\text{O}/\text{CuO}$ grown on copper mesh, (iii) adsorption of Cu^{2+} and NO_3^- and the formation of an electrical double layer, (iv) rinsing (I) removes excess unadsorbed Cu^{2+} and NO_3^- , (v) reaction of S^{2-} with pre-adsorbed Cu^{2+} ions to form CuS , and (vi) rinsing (II) to remove excess and unreacted species and form the solid solution CuS on surface of the Cu mesh/ $\text{Cu}_2\text{O}/\text{CuO}$ (reproduced with permission from ref. 182, copyright Springer 2019). (a) Schematic illustration of the synthesis of $\text{Cu}_2\text{O}/\text{CuS}$ nanocomposites. (b) and (c) SEM images of $\text{Cu}_2\text{O}/\text{CuS}$ nanocomposites grown on a copper gauze collected at different magnifications. (d) HR-TEM image collected from a representative $\text{Cu}_2\text{O}/\text{CuS}$ particle. (e) The corresponding FFT image of (d). Blue and yellow dashed lines highlight the diffraction rings of polycrystalline Cu_2O and CuS particles, respectively. (f)–(i) HAADF-TEM image of $\text{Cu}_2\text{O}/\text{CuS}$ nanocomposites and the corresponding mapping images of elements O, Cu, and S (reproduced with permission from ref. 183, copyright ACS 2021).

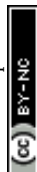
catalytic systems showed high onset potential values even though the system was slightly more favourable for reduction toward HCOOH than CO . The CuS system's minimal onset potential for CO or HCOOH suggests that the system could potentially act as a catalyst to form $2e^-$ products. In comparison, $\text{CuS}_{1.5}$ revealed minimal onset potential values for the generation of CO , HCOOH , CH_4 , and CH_3OH , which are, respectively, 0.19, 0.19, 0.53, and 0.53 (V_{RHE}) less than the values determined for $\text{Cu}(111)$. The simulation also revealed that, on $\text{CuS}_{1.5}$, the reaction toward $\text{CH}_4/\text{CH}_3\text{OH}$ was limited by the $\text{*CO} \rightarrow \text{*COH}$ formation, while on both the $\text{CuS}_{1.5}$ and CuS

surfaces, the $\text{CO}_2 \rightarrow \text{*COOH}$ generation was the potential determining step for CO and HCOOH formation. According to the calculation, the facile *CO desorption on CuS surfaces could restrict the formation of $\text{CH}_4/\text{CH}_3\text{OH}$. Therefore, methods for breaking the *CO and *COOH or $\text{*CHO}/\text{*COH}$ interactions' linear scaling could be investigated as well in order to enhance most of these systems, while the performance of CuS as a catalyst for the production of CH_4 and CH_3OH could be improved by simply increasing the CO adsorption strength. In summary, their DFT calculation discovered that increasing the sulfur content in 2D copper sulfide materials promotes CO and HCOOH formation at lower applied potentials and facilitates methane and methanol production. Although these outcomes seem promising from a thermodynamic standpoint, more research into the kinetic behaviour of processes using the most promising systems, considering defect creation under operating conditions, and more research examining additional reaction pathways are essential for strengthening the concept of using these systems as potential electrocatalysts.

Another group proposed experimentally supported theoretical simulation to investigate the mechanism of ECO_2R on S-modified Cu electrocatalysts for HCOOH selectivity.¹⁸⁸ They discovered that the surrounding environment and symmetry of the remaining sulfur atoms greatly influence their stability. Most sulfur exists in somewhat unstable forms that satisfy the strong CO^* surface-enhanced infrared absorption spectroscopy signal and the experimentally confirmed negative XPS shift. From the ECO_2R energy graphics, it was revealed that these types of S atoms cannot immediately promote formic acid generation; instead, they produce a highly dominant CO^* and have a large CO^* adsorption capacity. However, the study discovered that the sulfur atoms' strong CO^* adsorption improves the CO^* coverage to an almost four times higher level than on a bare Cu . With such a dense CO^* coverage, specific surface reactive sites are restricted, resulting in a solution-phase CO_2 reduction pathway leading to the formation of highly selective HCOOH . Also, it has been found that CO^* can considerably stabilize residual S, and the adsorption and associated electronic structure modulation studies have revealed the source of CO^* adsorption improvement. Lastly, the group proposed a synergy between residual S and CO^* dominating the HCOOH formation in ECO_2R using experimental findings from published research and their DFT calculations.^{118,177} These findings offer novel insights into the fundamental role of atomic sulfur in the selective formation of ECO_2R products and intermediates on metallic Cu electrodes. Advanced characterization techniques, *i.e.*, *in situ/operando* studies¹⁸⁹ and computational approaches,^{190,191} must be correlated to better understand the structure–stability–efficiency connections of electrocatalysts. This connection is also essential because of the high selectivity of these catalysts regarding particular catalytic processes and product generation.

5. Perspectives on Cu/S -based ECO_2R catalysts

The catalyst's nanostructure, porosity, morphology, stability, surface area, and particle size influence ECO_2R by affecting the



adsorption and desorption of CO₂ molecules. These changes in adsorption and desorption processes lead to different products. In addition to the several modification strategies, the performance of Cu/S-based nanomaterials towards ECO₂R is strongly connected with several other factors. This section highlighted the novelty aspects related to Cu/S-based nanomaterials.

Controlled morphology is critical for understanding the structure–activity relationships of Cu/S-based nanomaterials in ECO₂R catalysis. For example, Cu/S-based nanomaterials have been developed with various nanoscale morphologies, including nanoarrays,¹¹¹ cubes,¹⁰⁴ ultrathin nanosheets,¹⁵⁰ hollow polyhedrons,¹⁷⁶ and hierarchical structures.¹⁵⁰ Also, it is to be noted that the morphology and ECO₂R performance vary among Cu/S-based nanomaterials due to the precursor effect.¹²¹ The morphology-related characterization studies showed that different sulfur source precursor materials have different solubilities and rates of release during hydrothermal synthesis, resulting in diverse surface morphologies at the end.¹²¹ Several of these catalysts also have highly porous frameworks,¹⁰⁸ aiding electrolyte permeability. Moreover, a greater surface area might offer more multi-active sites, which is also advantageous for producing an optimal nanostructure for improved ECO₂R activity.^{151,153}

Notably, the phase engineering approach also has a vital role in ECO₂R. For example, phase changes in Cu/S-based nanomaterials are observed due to S leaching during ECO₂R. For instance, during the reduction process, XPS and HRTEM analyses showed that thermodynamically unstable CuS undergoes phase changes to metallic Cu, CuO, or Cu₂O.¹¹⁷ For example, Shinagawa *et al.* employed XPS with SAED studies for the post-electrolysis sample. The group observed that during ECO₂R, S-modified materials lose excess sulfur irrespective of initial sulfur concentration due to the cathodic environment, resulting in modification of particle size and the co-existence of Cu₂S and Cu metal, significantly interfering with the electrocatalyst's activity.¹²⁸ Zhao's group¹¹⁰ performed XPS and HRTEM to demonstrate that CuS nanosheets undergo partial reduction to metallic Cu, and the metallic Cu phase is partially oxidized to CuO after ECO₂R. Zhang *et al.*¹⁷⁶ also did XPS and revealed that CuS-HP synthesized from HKUST-1 was transformed to Cu(S) with a minimal S^{δ−} (0 ≤ δ ≤ 2) and metallic Cu. This study revealed that although Cu(S) has minimal S concentrations, the polyhedron structures offer high activity and stability during the ECO₂R. As a result, *in situ* reconstruction of CuS in ECO₂R promotes the ECO₂R activity and product selectivity.

Utilizing highly conducting substrates to load copper sulfides, such as copper oxide,¹⁸² carbon-based materials,¹¹⁹ and highly porous 3D foams,¹¹¹ improves electron conduction and generates more electrocatalytic active sites. Notably, forming copper sulfides–conducting substrate bonds improves catalyst stability, resulting in a longer lifetime.¹¹⁹

Other techniques for increasing catalytic activity include the compositional changes of bi-metallic sulfides in comparison to mono-metallic sulfides to generate a heterostructure capable of providing synergistic coupling,¹⁶⁰ interfacial phenomena,¹⁶¹ and defect-rich structures¹⁴² to maximize the benefits of metals

and sulfides. The number of exposed active sites can be increased, and ECO₂R performance can be improved through heterostructure engineering strategies by doping highly functional single-metal nanoparticles¹⁶⁹ and non-metals¹⁷¹ on metal sulfides. Also, constructing heterostructures can improve ECO₂R efficiency by adjusting the interfaces and electrical states and offering more active sites. So, lower onset potential, increased cathodic current, higher TOF, and lower Tafel slope values can be obtained during the ECO₂R process.¹⁸⁷

Furthermore, the ECO₂R performance varies across Cu/S-based nanomaterials based on their S content or S-vacancy concentration. Vacancy creation through defect engineering techniques, *i.e.*, S-vacancy¹⁶² and Cu-vacancy,¹⁷² can shift the product selectivity of ECO₂R to C₂₊ products by modifying the electronic structure of S and Cu sites in Cu/S-based nanomaterials.

6. Comparative analysis of Cu/S-based ECO₂R catalysts with other best-performing Cu-based and metal–sulfide-based ECO₂R catalysts

The performance data and recent research of earlier studies on different Cu/S-based nanomaterials are summarized in Table 1. Tables 2 and 3 present the best-performing Cu-based and metal–sulfide-based nanomaterials (MS-based nanomaterials) for ECO₂R (other than copper–sulfide-based) to compare the catalytic efficacy of Cu/S-based nanomaterials (summarized in Table 1) with an emphasis on partial current density and faradaic efficiency. Table 1 indicates that many Cu/S-based nanomaterials have demonstrated acceptable levels of stability during the ECO₂R as well as good partial current density and faradaic efficiency for mainly formate/HCOOH production. However, for C₁ and C₂₊ products, the performance of Cu/S-based nanomaterials is unsatisfactory in terms of faradaic efficiency and product selectivity. Other than HCOOH/formate, several Cu/S-based nanomaterials such as N-doped Cu₂S thin layers,¹⁷¹ polycrystalline Cu (Cu-s),¹⁰⁰ S-doped spherical coral-like CuO,¹⁷⁴ CSVE-Cu,¹⁴² DSV-engineered CuS,¹⁴³ S-HKUST-1,¹⁵⁴ CeO₂-modified CuS nanoplates,¹⁵⁸ and Cu₂S nanocrystal-decorated Cu nanosheets,¹⁵⁵ have yielded C₂₊ products. However, their selectivities are much lower compared to those of best-performing Cu-based catalysts, *i.e.*, dCu₂O/Ag2.3%,¹⁹² Ag–Cu₂O,¹⁹³ Cu₃N-derived Cu nanowires,¹⁹⁴ CuO-derived Cu,¹⁹⁵ Cu(salophen)-coated GDE,¹⁹⁶ Cu@Cu₂(OH)₃NO,¹⁹⁷ Cu₂O films,¹⁹⁸ oxygen-bearing copper,¹⁹⁹ and nano-defective Cu nanosheets.²⁰⁰ In conclusion, combining Cu/S-based nanomaterials with other nanomaterials, *i.e.*, metal oxides, single-atom metals, metal selenides and metal phosphides, through various engineering/modifications strategies could be effective for higher C₂₊ product selectivity. In addition, cadmium-based sulfide, molybdenum-based sulfide and zinc-based sulfide, as summarized in Table 3, electrocatalysts exhibited better current density and faradaic efficiency performance than Cu/S-based and Cu-based catalysts for CO production. Furthermore, for CH₄ production, the CuS/Ni foam¹¹⁰ electrocatalyst offered better

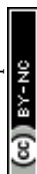


Table 1 A current overview of Cu/S-based catalysts for ECO_2R

Products	Electrocatalysts	Method of synthesis	Potential (V_{RHE})	Electrolyte used	Cell	Partial current density (mA cm^{-2})	Faradaic efficiency (%)	Stability (h)	Ref.
HCOOH	Sulfur-modified Cu_2O	Wet chemistry approach	−0.9	0.1 M KHCO_3	H-type cell	260 ± 16	≈90	80	99
C_2H_4	Cu-s	Electrochemical reduction	−1.2	0.1 M KHCO_3	H-type cell	40.8	68.6	8	100
HCOOH	PTFE-coated CuS/BM	One-step chemical bath process	−0.7	0.5 mol L^{-1} KHCO_3	H-type cell	50	70	1.5	101
CO	Hollow-shaped CuS microcubes	Galvanic replacement technique	−0.16	0.5 M KHCO_3	H-type cell	—	32.7	10	104
HCOOH	Hollow ordered porous copper sulfide cuboctahedra	Hard templating strategy	−1.1	0.1 M KHCO_3	H-type cell	7.8	70.3	26	108
C_2H_4	Copper sulfide film consisting of nanoparticles	Thermal decomposition	−0.13	1 M KHCO_3	H-type cell	≈0.08 (total)	—	—	109
HCOOH	Cu_2S over copper foam	Electrochemical deposition	−2.0 V (vs. Ag/Ag^+)	0.5 mol L^{-1} Bmim BF_4	H-type cell	5.3	85	6	111
CH_4	CuS/Ni foam	Hydrothermal method	−1.1	0.1 M KHCO_3	H-type cell	7.32 (total)	73.5 ± 5	60	110
HCOOH	CuS nanoflowers	Ethylene glycol solvothermal method	−0.8 to −1.0	0.1 M KHCO_3	H-type cell	≈6.5	≈52	2	117
HCOOH	$\text{Cu}_{1.81}\text{S}/\text{MWCNT}$ -600-OD	Two-step coupling approach	−0.67	0.5 M KHCO_3	H-type cell	3.75 (total)	82	20	119
HCOOH	$\text{Cu}_2\text{S}/\text{C}$	Liquid phase sulfidation	−0.78 ± 0.02	0.5 M KCl + 0.5 M KHCO_3	H-type cell	1.5 (total)	12	5	120
CO	CuS catalyst thiourea precursors	Hydrothermal synthesis	−0.51	0.1 M KHCO_3	H-type cell	—	72.67	5	121
HCOOH	Cu_2S nanocrystals	Electrochemically driven cation exchange method (ED-CE)	−0.9	0.1 M NaHCO_3	H-type cell	19	87.3	9	122
HCOOH	Cu_2S	ED-CE	−0.2	0.05 M K_2CO_3	H-type cell	—	≈70	16	123
CO	S-derived Cu–Sb	Heat-up colloidal nanoparticle route	−1.0	1 M KHCO_3	Flow cell	37.6	80.5	24	124
HCOOH	S-CNWs	Chemical oxidation followed by solvothermal reaction	−0.59	0.25 M KHCO_3	H-type cell	≈10	≈60	2	125
HCOOH	Submicron-sized Cu–S	Solvothermal approach	−0.8	0.1 M KHCO_3	H-type Cell	≈10 (total)	80	12	128
HCOOH	CuS_x nanoparticles	Dipping	−0.6	0.1 M KHCO_3	H-type cell	2.5	72	72	129
HCOOH	S- $\text{Cu}_2\text{O}/\text{Cu}$	Hydrothermal reduction followed by <i>in situ</i> reconstruction	−1.0	0.1 M KHCO_3	H-type cell	5	70	6	134
HCOOH	CuS 811 derived S-doped Cu_2O	Surfactant-free double-replacement reaction	—	0.1 M KHCO_3	Flow cell	321	80	33	135
Alcohols (ethanol + propanol)	CSVE-Cu	Solution-phase growth	−0.95	1.0 M KOH	Flow cell	126	31.6	16	142
<i>n</i> -Propanol	DSV-engineered CuS_x	Lithium tuning approach	−1.05	0.1 M KHCO_3	H-type cell	9.9	15.4 ± 1	10	143
HCOOH	SnO_2/C onto snowflake-like Cu_2S	Hydrothermal method	−1.0	0.1 M KHCO_3	H-type cell	15.6	88	50	148
HCOOH	CuS/SnO_2 -S derived Cu/Sn/ $\text{Cu}_{6.26}\text{Sn}_5$ catalysts	Redox reaction	−0.8	0.5 M KHCO_3	H-type cell	≈18.8	84.9	10	149
CO	SnO_2 confined on CuS nanosheets	One-pot precipitation synthesis	−1.0	0.1 M KHCO_3	H-type cell	15.24	>85	24	150
HCOOH	$\text{Cu}_2\text{S}/\text{SnS}_3$ nanoplates	Step-growth approach	−1.2	1 M KHCO_3	Flow cell	181	72.4	3	151
CO	CuInS_2 hollow nanostructures	Template-free approach	−0.7 & −1.0	0.1 M KHCO_3	H-type cell	≈13 (total)	82.3	4	152
HCOOH	Graphdiyne/copper sulfide	Hydrothermal followed by <i>in situ</i> growth	−0.9	0.1 M KHCO_3	H-type cell	65.6 (total)	70	4	153
C_2H_4	S-HKUST-1	Local sulfur doping strategy	−1.94	1.0 M KOH	Flow cell	228	57.2	8	154
Ethanol	Cu_2S nanocrystal-decorated Cu nanosheets	Hydrothermal followed by self-assembly	−1.2	0.1 M KHCO_3	H-type cell	$j_{\text{C}_2\text{H}_4\text{OH}} = 20.7$	46	20	155
CH_4	$\text{Cu}_x\text{Zn}_{1-x}$ nanoparticles	Seed-mediated growth method	−0.98	0.1 M KHCO_3	H-type cell	$j_{\text{methane}} = 4.58$	76	5	156
Ethanol	SD- CuCd_2	Co-Precipitation	−0.8	0.1 M KHCO_3	H-type cell	0.6	32	6	157
C_2H_4 products	CeO_2 -modified CuS nanoplates	Self-assembly method	−0.9	1 M KOH	Flow cell	$j_{\text{C}_2\text{H}_4} = 75.3$	75	180	158
HCOOH	Bi/ CeO_2 /CuS	Solvothermal process	−0.9	0.5 M KHCO_3	H-type cell	17 (total)	88	6	159
HCOOH	Heterostructure Bi- Cu_2S nanocrystals	One-pot solution-phase approach	−1.0	0.1 M KHCO_3	H-type cell	≈17	>98	10	160
HCOOH	$\text{CuS}@1\text{T-N-WS}_2/\text{WO}_3$	A facile solvothermal process followed by a reflux strategy	−1.3	0.5 M KHCO_3	H-type cell	125.05	≈55.6	2	161
CO	MoS_2 -CuS	Hydrothermal approach	−0.6	0.1 M KHCO_3	H-type cell	3.6	83.2	5.55	162



Table 1 (continued)

Products	Electrocatalysts	Method of synthesis	Potential (V _{RHE})	Electrolyte used	Cell	Partial current density (mA cm ⁻²)	Faradaic efficiency (%)	Stability (h)	Ref.
Ethanol	Mo ⁴⁺ -doped CuS nanosheet-assembled hollow spheres	One-step solution-phase approach	−0.6	1.0 M KOH	Flow cell	$J_{\text{ethanol}} = 186.5$	67.5	14	169
HCOOH	Cu-doped Bi derived from CuS-Bi ₂ S ₃	Solvothermal followed by <i>in situ</i> reconstruction	−0.86	1 M KOH	MEA cell	1132	90	100	170
C ₂ H ₄	N-doped Cu ₂ S thin layers	Annealing process	−0.98	0.1 M KHCO ₃	H-type cell	~2.6	14.72	0.25	171
CO	SNC@Cu _{1.96} S	Surface coating and calcination	−0.84	0.5 M KHCO ₃	H-type cell	37.2	85.2	13	172
HCOOH	Cu-5000S	Electrochemical deposition	−0.8	0.1 M KHCO ₃	H-type cell	13.9	~75	12	173
C ₂ H ₄	Sulfur-doped spherical coral-like CuO	Local sulfur doping strategy	−1.3	0.1 M KHCO ₃	H-type cell	15	48.4	20	174
HCOOH	S-Cu ₂ O-14	Hydrothermal sulfurization method followed by an electrochemical reduction process	−1.0	0.1 M KHCO ₃	H-type cell	16.3	67.2	20	175
HCOOH	CuS hollow polyhedron derived from a MOF	MOF self-sacrificial template method	−0.6	0.5 M K ₂ SO ₄ aqueous solution	H-type cell	~16	>90	36	176
HCOOH	AC-CuS _x	Pulsed electrochemical deposition	−0.9	0.1 M KHCO ₃	H-type cell	9	75	3	177
HCOOH	CuS deposited onto thermally prepared Cu ₂ O/CuO heterostructure	Thermal oxidation followed by SILAR	−0.7	0.1 M KHCO ₃	H-type cell	20	84	2.5	182
HCOOH	Cu ₂ S/Cu ₂ O nanocomposites	Hydrothermal approach and electrochemical treatment	−0.9	0.1 M KHCO ₃	H-type cell	12.3	67.6	30	183

performance than Cu/MoS₂²⁰¹ and Fe_{4.5}Ni_{4.5}S₈²⁰² regarding partial current density and faradaic efficiency.

7. Practical applications of Cu/S-based ECO₂R catalysts

The execution of ECO₂R on Cu/S-based catalysts on a practical level is still in its initial stages due to several difficulties. The primary challenges to the commercialization of Cu/S-based ECO₂R catalysts are (1) considerable CO₂ extraction and purification costs, as ECO₂R needs highly pure (99.999%) CO₂, (2) high energy consumption, (3) low yield of energy-dense C₂₊ products and (3) a restricted marketplace that is less attractive to investors.^{285,286} In addition, there are multiple technical obstacles related to ECO₂R, such as low catalytic performance, poor product selectivity, unsatisfactory catalytic stability, and non-optimization of cell design for use in practice.^{287,288} While Cu/S-based materials yielded several products during ECO₂R, producing formate/HCOOH was observed as the main product. However some catalysts can produce hydrocarbons, but their selectivities are much lower than those of copper nanoparticles or oxide-derived copper. This section discusses the possibilities of scaling up ECO₂R technology to the industrial level, focusing on the practical applications of formate/HCOOH production.

Indeed, scaling up an ECO₂R cell for formate/HCOOH production is complex because various factors and constraints must be investigated and carefully evaluated to maintain stable cell efficiency.²⁸⁹ A study revealed that the cell must maintain consistent operation for at least 8000 hours for commercial applications. Additionally, it should exhibit an overpotential below 1.0 V_{RHE}, a partial current density between 200 and 1000 mA cm⁻², and a formate/HCOOH faradaic efficiency of above 90% during stable operation.²⁹⁰ Another research stated that the electrolyzer should be stable for a minimum of 20 000 hours to be financially viable.²⁹¹ In this context, the proton exchange membrane and electrode framework also play a vital role in stabilizing the cell. Among the various membranes reported, bipolar membranes have received much attention for large-scale use because they can maintain pH gradient, decrease liquid crossover, and promote water separation at the membrane electrolyte interface.^{292,293} Operating variables such as electrolyte, flow rate, p^H, and CO₂ feed type can all be optimized to improve cell efficiency.²⁹⁴ Nowadays, there are only a few huge-scale CO₂ electrolyzers for methane, CO, C₁, C₂₊, and formic acid generation.^{295,296} Few academic institutions executed large-scale manufacture of formate/formic acid using ECO₂R. For example, in 2008, a pilot plant was constructed with a maximum output of 146 kg CO₂ per day, leading to approximately 110 kg formic acid per day at an operating pressure of around 10 bar.²⁹⁷ A different approach to the current scale-up investigation revealed a pre-pilot facility output of 55 kg CO₂ per day, resulting in approximately 12 kg of formate per day at ambient pressure.²⁹⁸

While the aforementioned studies demonstrated the practicality of this approach for scaling up, several obstacles must be addressed to maximize plant functionality. One of the primary



Table 2 An overview of other Cu-based ECO₂R catalysts

Electrocatalysts	Main products	Potential (V _{RHE})	Electrolyte used	Cell	Current density (mA cm ⁻²)	Faradaic efficiency (%)	Stability (h)	Ref.
Single Cu atom	Acetone	-0.76	0.1 M KHCO ₃	H-type cell	1.0	36.7	5	203
TS-Cu	Ethylene	3.5	1 M KOH	Flow cell	800	72	200	204
dCu ₂ O/Ag2.3%	Ethanol	-0.87	1 M KOH	H-type cell	326.4	40.8	12	192
Cu-NC	Ethanol	-1.01	0.1 M KHCO ₃	H-type cell	—	18.4	8	205
Cu ₂ O/NCS	Ethylene	-1.3	0.1 M KHCO ₃	H-type cell	—	24.7	4	206
CuN ₄	Ethanol	-1.2	0.1 M KHCO ₃	Flow cell	16.2	55	1	207
Cu ₄ Ni ₂ @N-MWCNT	Formate	-0.53	0.1 M KHCO ₃	H-type cell	3.0	90	10	208
Cu-doped carbon xerogel	Methane	—	0.1 M KHCO ₃	H-type cell	—	—	5	209
Cu-nanoparticles	Formate	-1.0	0.1 M KHCO ₃	H-type cell	6.53	54	12	210
GO-VB6-Cu-2	Ethanol	-0.250	0.1 M KHCO ₃	H-type cell	7.452	56.3	24	211
Cu-N-C	Carbon mono oxide	-0.7	0.1 M KHCO ₃	H-type cell	3.61	92	60	212
Cu ₂ O@Cu-MOF	Methane	-1.71	0.1 M KHCO ₃	H-type cell	8.4	79.4	1	213
PcCu-Cu-O	Ethylene	-1.2	0.1 M KHCO ₃	H-type cell	7.3	50	4	214
Cu ₉₀ In ₁₀ /C	Carbon mono oxide	-0.75	0.1 M KHCO ₃	H-type cell	9.85	85	4	215
CuO/CuTCPP	Formic acid	-1.65	0.5 M EMIMBF ₄	H-type cell	4.5	85.2	5	216
Cu[(111)@Cu-THQ	Ethylene	-1.4	0.1 M KHCO ₃	H-type cell	14.3	42	8	217
CuZn	C ₂ +	-1.0	0.1 M KHCO ₃	H-type cell	1.5	25	7	218
Cu@Cu _x O	C ₂ +	-1.78	0.1 M KHCO ₃	H-type cell	150	70	20	219
Cu-Sn	Carbon monoxide	-0.7	0.1 M KHCO ₃	H-type cell	4.5	90	12	220
Cu ₃ N-derived Cu nanowires	C ₂	-1.0	0.1 M KHCO ₃	H-type cell	-50.6	86	28	194
CuO-derived Cu	C ₂	-0.48	KH ₂ PO ₄	H-type cell	-210	9.6	1	195
Cu _v -Cu ₂ O	Ethylene	-0.76	0.1 M KHCO ₃	H-type cell	15.7	51	10	221
Cu-Zn	Carbon monoxide	-1.1	0.1 M KHCO ₃	H-type cell	4.3	74	4	222
Cu@Cu ₂ (OH) ₃ NO	Ethylene	-1.2	0.1 M KHCO ₃	H-type cell	80	31.8	20	197
Cu ₂ O-derived Cu/PdCl ₂	Ethylene	-1.0	0.1 M KHCO ₃	Flow cell	—	30.1	8	223
Cu ₂ O films	C ₂	-0.99	0.1 M KHCO ₃	H-type cell	-52	39	1	198
Cu ₂ O-derived Cu particles	Ethylene	-0.98	0.1 M KHCO ₃	H-type cell	-13.3	43	1	224
Branched CuO nanoparticles	Ethylene	-1.0	0.1 M KHCO ₃	H-type cell	-22.0	53	12	225
Cu-Pd alloy	Carbon monoxide	-0.9	0.1 M KHCO ₃	H-type cell	—	87	5	226
Cu-Pt alloy	Methane	-1.6	0.1 M KHCO ₃	H-type cell	0.598	21	0.5	227
CuMgAl LDH	Acetic acid	-0.4	0.1 M KHCO ₃	Flow cell	—	—	24	228
PD-CuO _x /C	Ethylene	-1.4	0.1 M KHCO ₃	MEA cell	26.0	45	48	229
CuCo ₂ Se ₄	Acetate	-0.25	0.3 M NaHCO ₃	H-type Cell	26	100	100	230
Cu(salophen)-coated GDE	C ₂ +	-1.2	1 M KOH	H-type cell	121	37	2	196
Ag-Cu ₂ O	Acetic acid	-0.9	1 M KOH	H-type cell	310	70	20	193
Hydroxo-bridged phenanthroline Cu(II) molecule	Ethylene	-1.25	0.1 M CsHCO ₃	H-type cell	-5.5	42	15	231
Cu NC spheres	Ethylene	-1.1	0.1 M KHCO ₃	H-type cell	-9	41	1	232
Copper electrodes	Methane	-1.2	0.1 M KHCO ₃	H-type cell	5	45	8	233
Oxygen-bearing copper	Ethylene	-0.95	0.3 M NaHCO ₃	H-type cell	44.7	45	1	199
Nano-defective Cu nanosheets	Ethylene	-1.0	0.1 M K ₂ SO ₄	Flow cell	66.5	83.2	1	200



Table 3 An overview of other metal–sulfide-based ECO₂R catalysts

Metal–sulfides	Electrocatalysts	Potential (V _{RHE})	Electrolyte used	Main products	Cell	Partial current density (mA cm ^{−2})	Faradaic efficiency (%)	Stability (h)	Ref.
Cadmium-based sulfide	Ag–CdS _{1–x}	−1.1	1 M KHCO ₃	CO	H-type cell	53.7	87.1	—	234
	P–CdS	−0.8	0.5 M KHCO ₃	CO	H-type cell	89.8	88	10	235
	Nanoneedle-shaped CdS	−1.2	0.1 M KHCO ₃	CO	Flow-cell	212	95.5 ± 4.0	24	236
	CdS nanorods	−1.2	0.1 M KHCO ₃	CO	H-type cell	~22	81	10	237
	CdS with S-vacancy	−0.8 to −1.1	0.5 M KHCO ₃	CO	H-type cell	~1.3	95	40	238
Tin-based sulfide	S vacancy engineered CdS nanorods	−1.1	0.5 M KHCO ₃	CO	H-type cell	20.5	100 ± 0.5	10	239
	CdS supported CNTs	−1.2	0.1 M KHCO ₃	CO	H-type cell	~13 (total)	~95	10	240
	CdS/MXene	−1.0	0.1 M KHCO ₃	CO	H-type cell	~6	94	8	241
	Ag ₂ S/CdS	−1.1	0.5 M KHCO ₃	CO	Flow-cell	10.6	95	20	242
	Hybridized composite of defective SnS ₂ nanosheets and Ag nanowires	−0.1	0.5 M KHCO ₃	HCOOH	H-type cell	23.3	60	10	243
Bismuth-based sulfide	5% Ni-doped SnS ₂ nanosheets	−0.9	0.1 M KHCO ₃	CO & HCOOH	H-type cell	19.6 (total)	93	8	244
	Sn(S)/Au	−0.75	0.1 M KHCO ₃	HCOOH	H-type cell	51	93.3	40	245
	Stannous sulfide modified with amino-functionalized carbon	−0.9	0.5 M KHCO ₃	HCOOH	H-type cell	41.1	92.6	15	246
	SnS ₂ nanosheets supported on rGO	−1.4	0.5 M NaHCO ₃	HCOOH	H-type cell	11.75	84.5	2	247
	Tin(IV) sulfide monolayers	−0.8	0.1 M KHCO ₃	HCOOH	H-type cell	~45 (total)	94 ± 5	80	248
Molybdenum-based sulfide	N–Sn(S) nanosheets	−0.7	0.1 M KHCO ₃	HCOOH	Flow-cell	25 (total)	93.3	20	249
	Semimetal phased 1H–SnS ₂	−0.8	0.1 M KHCO ₃	CO	H-type cell	10.9	98.2	15	250
	S doped-Bi ₂ O ₃ -CNT	−0.9	0.5 M KHCO ₃	HCOOH	H-type cell	48.64	97.06	10	251
	S-derived Bi	−0.75	0.5 M NaHCO ₃	HCOOH	H-type cell	5 (total)	84	24	252
	Bi ₂ S ₃ –Bi ₂ O ₃ /rGO	−0.9	0.1 M KHCO ₃	HCOOH	H-type cell	~3.8	90	24	253
Indium-based sulfide	Bi–Bi ₂ S ₃	−1.0	0.1 M KHCO ₃	HCOOH	H-type cell	~14	85	12	254
	S-doped two-dimensional (2D) bismuth subcarbonate (S–BiOC)	−0.9	0.5 M KHCO ₃	HCOOH	H-type cell	29	>90	20	255
	Bi ₂ S ₃ –PPy	−0.9	0.5 M KHCO ₃	HCOOH	H-type cell	56.95 (total)	91.18	>110	256
	2D defective Bi ₂ S ₃ NSs	−0.93	0.5 M KHCO ₃	HCOOH	H-type cell	~30	>90%	200	257
	Bulk MoS ₂	−0.764	4 mol% EMIM–BF ₄ solution	CO	H-type cell	65	98	10	258
Molybdenum-based sulfide	rGO–PEI–MoS _x	−0.65	0.5 M NaHCO ₃	CO	H-type cell	55 (total)	85.1	—	259
	N–MoS ₂ @NCDs-180	−0.9	6 mol% EMIM–BF ₄ solution	CO	H-type cell	36.2	90.2	10	260
	Cu–g-C ₃ N ₄ /MoS ₂	−0.67	0.5 M KHCO ₃	CH ₃ OH	H-type cell	78	19.7	30	261
	CuO–ZnO–MoS ₂	−0.6 & −0.9	0.5 M KHCO ₃	Methanol & ethanol	H-type cell	121 (total)	24.6 & 11.1	3	262
	NCMSH	−0.7	4 mol% EMIM–BF ₄ solution	CO	H-type cell	34.31	92.68	24	263
Indium-based sulfide	Cu/MoS ₂	−1.4 V _{SCE}	0.1 M KHCO ₃	CH ₄	H-type cell	~3.2	17.08	48	201
	VA–Mo _{0.95} Nb _{0.05} S ₂	−0.8	50 vol% EMIM–BF ₄	CO	H-type cell	237 (total)	80	—	264
	Cu ₂ O@MoS ₂ nanosheets	−1.3 and −1.1	0.5 M KHCO ₃	Methanol & ethanol	H-type cell	113 (total)	12.3 & 7.9	2	265
	H–E–MoS ₂ with decoration of fluorosilane	−0.9	6 mol% EMIM–BF ₄ solution	CO	H-type cell	—	81	10	266
	Monolayers of MoSeS alloys	−0.15	4 mol% EMIM–BF ₄ solution	CO	H-type cell	43	45.2	10	267
Indium-based sulfide	In nanoparticles on In ₂ S ₃ nanosheets	−1.0	1 M KHCO ₃	HCOOH	H-type cell	40.3	76	8	268
	Mn–In ₂ S ₃	−0.9	0.1 M KHCO ₃	HCOOH	H-type cell	20.1 (total)	90	8	269
	ZnIn ₂ S ₄	−1.18	1 M KHCO ₃	HCOOH	Flow cell	~298	99.3	60	270
	Flower-shaped In ₂ S ₃	−2.3	BrimPF ₆ /MeCN–H ₂ O solution	HCOOH & CO	H-type cell	25.6	86	10	271
	S–In	−0.98	0.1 M KHCO ₃	HCOOH	H-type cell	84 (total)	93	10	272



Table 3 (continued)

Metal-sulfides	Electrocatalysts	Potential (V _{RHE})	Electrolyte used	Main products	Cell	Partial current density (mA cm ⁻²)	Faradaic efficiency (%)	Stability (h)	Ref.
Lead-based sulfide	ZnO NW/CIGS/InS	-0.24	0.1 M KHCO ₃	HCOOH	H-type cell	0.35	77.2	10	273
	ZnO NF/CIGS/InS	-0.24	0.1 M KHCO ₃	HCOOH	H-type cell	—	68.1	—	273
	Wafer-structured sulfur-derived Pb	-1.08	0.1 M KHCO ₃	HCOOH	H-type cell	12	88	—	274
Zinc-based sulfide	PbS nanocrystals	-1.2	0.1 M KHCO ₃	HCOOH	H-type cell	—	97.6 ± 5.3	10	275
	ZnS@Zn	-2.4 V _{Fe/Fe+}	Propylene carbonate/ (CH ₃ CH ₂ CH ₂ CH ₂) ₄ N(ClO ₄)	CO	H-type cell	6.4	92	4	276
	S-Zn-S nanosheets	-0.8	0.1 M KHCO ₃	CO	H-type cell	~11	94.2	15	277
Titanium-based sulfide	ZnS@ZnO	-0.56	1 M KOH	CO	H-type cell	~109	97.2 ± 0.5	40	278
	TA-ZnS	-1.9 V _{Ag/AgCl}	0.1 M KHCO ₃	CO	H-type cell	~7	~83	20	279
	Semi-metallic titanium disulfide (TiS ₂)	-0.5	0.1 M NBu ₄ -PF ₆	CO	H-type cell	5	83	16	280
	Ni ₂ FeS ₄ nanosheets	-0.7	0.1 M KHCO ₃	CO	H-type cell	—	5.9	1	281
Other bimetallic-sulfide	FeS ₂ /NiS nanocomposite	-0.6	0.5 M KHCO ₃	CH ₃ OH	H-type cell	3.1 (total)	64	4	282
	Bulk Fe/Ni sulfides (Fe _{4.5} Ni _{4.5} S ₈ pentlandite)	-1.80 V _{NHE}	0.1 M TBAPF ₆	CO & methane	H-type cell	3 (total)	87 & 13	15	202
	Fe ₃ Ni ₆ S ₈	-1.80 V _{NHE}	0.1 M TBAPF ₆ /acetonitrile solution	CO	H-type cell	15 (total)	3.6	8	283
	Fe _{4.5} Ni _{4.5} S ₄ Se ₄	-1.80 V _{NHE}	Acetonitrile	CO	H-type cell	11	84	2	284

challenging tasks is maintaining a higher constant current density (recommended $>200 \text{ mA cm}^{-2}$) during the ECO_2R process. The higher current density value denotes a high electrochemical reaction rate, which raises the rate at which different products are produced. Notably, the electrode's properties and structure, as well as that of both the catalyst and the support material, significantly impact the current density. The lower energy efficiency presents another difficulty. The system demonstrated 28% and 33% energy efficiency in a lab-scale three-compartment set-up over a 1000-hour operation at current densities $>100 \text{ mA cm}^{-2}$.²⁹⁹ But, during 450 minutes of operation on a pilot scale, the energy efficiency exceeds 50%.²⁹⁷ However, data on the energy efficiency of various products formed on a large/pilot scale are scarce. Notably, the FE should be higher and the overpotential lower to maintain high cell energy efficiency. The support material on the working electrode, electrocatalytic properties, and anode materials primarily determine these parameters. In addition, the choice of the membrane is crucial because it may substantially decrease the IR drop between the compartments, *i.e.*, cathodic and anodic, lowering the cell overpotential and increasing energy efficiency. It is advised that the potential is between 2.5 and 3 V and that the FE is higher than 90% for industrial-scale implementation.^{300–302}

The cell construction for ECO_2R to formate/formic acid remains to be developed for industrial use. Every design has pros and cons of its own. However, the PEM cell layout appears to be a viable strategy for scaling up.³⁰³ The PEM design demonstrated a reasonably higher FE for formic acid generation for Cu/S-based electrocatalysts with a stable cell procedure of more than 100 hours.¹⁷⁰ However, multiple challenges restrict the commercial use of PEMs for ECO_2R to formate/formic acid and other products (*i.e.*, C_1 and C_{2+} products), including declination of membrane performance over time and GDE flooding. When the cell is operated at high pressure, the mechanical strength of the membrane becomes a significant issue, leading to a higher crossover rate of formic acid and membrane damage. Also, further research must be carried out on enhancing cell stability and efficiency. Although the PEM cell demonstrated a FE of 90% and steady performance for 100 hours,¹⁷⁰ industrial applications necessitate cell performance for a minimum of 8000 hours with a FE greater than 90%.²⁹⁰ Design aspects and operational variables should be considered when designing a cell assembly that can sustain long-term operation, encounter commercial operation demands, and be cost-effective. However, membranes are costly and need periodic replacement. According to researchers, membrane-less ECO_2R cells require more care to prevent frequent maintenance or membrane damage. Another significant problem is the overflow of GDE with electrolytes, which could harm the electrode structure and lower CO_2 conversion. Gas-phase ECO_2R cells may avoid such issues, so more research into these cells is needed to develop efficient, stable, and long-lasting ECO_2R cells.

8. Summary and outlook

This review discusses several strategies for modifying Cu/S-based nanomaterials, which are recognized as highly promising

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