


Cite this: *Nanoscale*, 2023, **15**, 13102

Control over product formation and thermodynamic stability of thiolate-protected gold nanoclusters through tuning of surface protecting ligands†

Manju P. Maman,^a Eyyakkandy Nida Nahan,^a Greeshma Suresh,^a Arunendu Das,^b Akhil S. Nair,^b Biswarup Pathak^b and Sukhendu Mandal^{*,a}

Surface-protecting ligands can regulate the structure of a cluster's core either through electronic or steric effects. However, the influence of the steric effect along with the electronic effect over controlling the structure during ligand exchange reactions remains elusive. To understand this, we have carried out ligand exchange on $[\text{Au}_{23}(\text{CHT})_{16}]^-$ (CHT: cyclohexane thiol) using aromatic thiolates where we have tuned the bulkiness at the *para* position of the thiolate group on the incoming ligands. The outcome of the experiments reveals that each of the ligands in the chosen series is precisely selective towards the parent cluster transformation through specific intermediates. The ligand with more steric crowding directed the reaction pathway to have Au_{28} nanocluster as the major product while Au_{36} was the final product obtained with the gradual decrease of bulkiness over the ligand. The combined experimental and theoretical results elucidated the mechanism of the reaction pathways, product formation, and their stability. Indeed, this study with the series of ligands will add up to the ligand library, where we can decide on the ligand to obtain our desired cluster for specific applications through the ligand exchange reaction.

Received 3rd June 2023,
Accepted 5th July 2023

DOI: 10.1039/d3nr02617e

rsc.li/nanoscale

Introduction

Surface-protecting organic ligands in inorganic nanomaterials, like metal nanoparticles or metal nanoclusters (NCs), provide stability to the clusters from aggregation and thereby control the interparticle interactions.^{1–4} The role of ligands becomes more crucial and emanant as the size of the metal NCs decreases. For instance, in the case of metal NCs with a size less than 2 nm, the presence of a ligand influences the structure because of the quantum confinement effect, and it also controls the high surface energy of bare NCs. This affects the physico-chemical properties and reactivity of the metal NCs as they are very sensitive to their structures/size.^{5–9} The ligand exchange reaction (LER) is one of the best strategies to understand the ligand effect in changing the structure of metal nanoclusters.^{10–15} Particularly, LER has gained much attention in thiolated gold NCs because of their fascinating structural and physico-chemical properties that vary upon surface

modification.^{16–19} For example, the optical properties of $\text{Au}_{25}(\text{SR})_{18}$ were tuned through ligand exchange, and this happened due to the structural deformation of the parent cluster.²⁰ The LER strategy is not only used to introduce new properties, but it also has been used as a probe to obtain mechanistic insights into structural changes.^{21–23} The knowledge of ligand parameters that influence the transformation of the parent cluster during the ligand exchange reaction will help us provide an insight into its mechanism, which will be beneficial for designing a new cluster with desired properties. Currently, there are some progress on the mechanistic aspect of LERs.^{24–26} However, in view of the fact that surface-protecting ligands can exchange between the formed clusters, the material nature is quite dynamic.²⁷ So, a detailed elucidation of the ligand-exchange reaction in an atom-precise manner with control in position and the number of incoming ligands remains challenging.

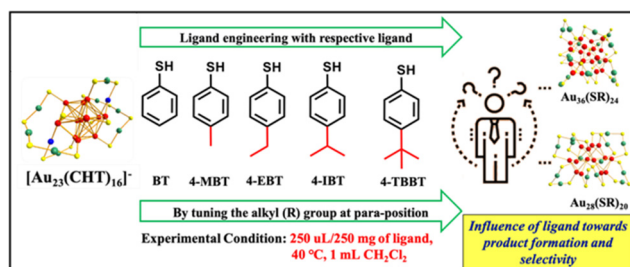
Single crystal structure analysis revealed that thiolate-protected gold-nanoclusters can be represented as $\text{Au}_n(\text{SR})_m$ (where n and m represent the number of gold and thiolate ligands, respectively) and contain the $\text{Au}(0)$ core that is protected by the $\text{Au}(\text{I})$ -thiolate staple motif.^{28,29} The variability in the staple motif is crucial for the ligand exchange as the reactivity depends on the binding energy of the respective staple motif.³⁰ For example, the $[\text{Au}_{23}(\text{SR})_{16}]^-$ nanocluster consists of

^aSchool of Chemistry, Indian Institute of Science Education and Research Thiruvananthapuram, Kerala 695551, India. E-mail: sukhendu@iisertvm.ac.in

^bDepartment of Chemistry, Indian Institute of Technology, Indore, Madhya Pradesh 453552, India

†Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3nr02617e>


In this report, we have tried to understand the role of the bulky alkyl group (R) at the *para* position of the aromatic thiolate ligand in core selectivity. For this, we tuned the alkyl (R) group at the *para* position of the aromatic thiolate ligand and carried out ligand exchange in the $[\text{Au}_{23}(\text{CHT})_{16}]^-$ nanocluster with five different ligands (Scheme 1). Herein, we observed that the aromaticity (electronic effect) and a slight change in the bulkiness of the incoming ligand (steric effect) have a greater influence over the core selectivity and stability of the products.



Scheme 1 Role of the ligand towards size/structural evolution of $[\text{Au}_{23}(\text{CHT})_{16}]^-$ via ligand exchange where BT: benzenethiol, 4-MBT: 4-methylbenzenethiol, 4-EBT: 4-ethylbenzenethiol, 4-IBT: 4-iso-propylbenzenethiol, 4-TBBT: 4-*tert*-butylbenzenethiol.

Results and discussion

Transformation of the $[\text{Au}_{23}(\text{CHT})_{16}]^-$ nanocluster *via* ligand exchange was carried out with the given ligands (Scheme 1), where we tuned the bulkiness of the substituent on the *para* position of the benzenethiol ligand.

Case of 4-TBBT

At the initial stage of the time-dependent MALDI-MS, a single peak depicting the $[\text{Au}_{23}(\text{CHT})_{16}]^-$ nanocluster converted into a group of peaks with a consistent mass difference of 50 Da indicating successful ligand exchange over the Au_{23} nanocluster (Fig. 1). As the reaction proceeds further complete ligand exchange in the Au_{23} nanocluster was observed as depicted by an intense peak at 7171.4 m/z at 0.5 h which corresponds to $\text{Au}_{23}(4\text{-TBBT})_{16}$. Indicating that the 4-TBBT ligand prefers to transform the parent cluster into the $\text{Au}_{23}(4\text{-TBBT})_{16}$ nanocluster, majorly. Later (from 1 h to 24 h), $\text{Au}_{23}(4\text{-TBBT})_{16}$ was found to be transforming majorly into the $\text{Au}_{28}(4\text{-TBBT})_{19}$ nanocluster as a peak at 8656.3 m/z was noticed with the disappearance of the $\text{Au}_{23}(4\text{-TBBT})_{16}$ nanocluster peak. Ultimately after 48 h, we found two intense peaks at 8656.3 m/z and 4937 m/z that correspond to the $\text{Au}_{28}(4\text{-TBBT})_{19}$ nanocluster and the $\text{Au}_{15}(4\text{-TBBT})_{12}$ gold complex, respectively. During the reaction, we observed Au_{25} and Au_{20} nanoclusters were formed as side products while the specific core selection occurred *via* 4-TBBT ligand exchange. Here, the peak obtained for $\text{Au}_{15}(4\text{-TBBT})_{12}$ in the time-dependent MALDI-MS spectra is actually an in-source mass-analyzed adduct of $\text{Au}_{10}(4\text{-TBBT})_{10}$ indicating that $\text{Au}_{15}(4\text{-TBBT})_{12}$ is an $\text{Au}_{10}(4\text{-TBBT})_{10}$

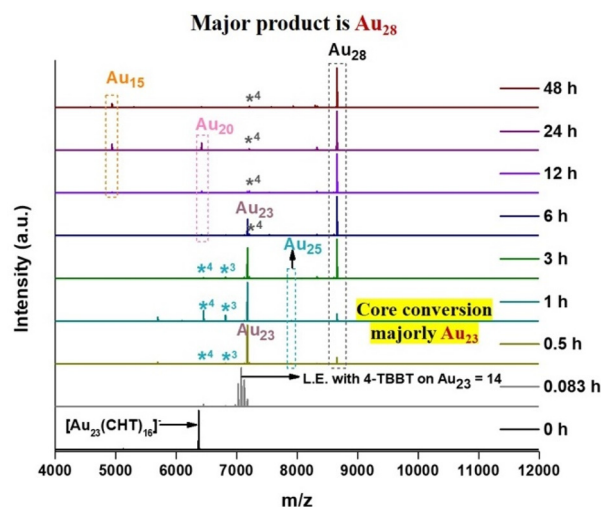


Fig. 1 Time-dependent MALDI-MS spectra showing the transformation process of the $[\text{Au}_{23}(\text{CHT})_{16}]^{-}$ nanocluster during the ligand exchange with the 4-TBBT ligand. Note: magenta: $\text{Au}_{23}(\text{SR})_{16}$ nanocluster, cyan: $\text{Au}_{25}(\text{SR})_{18}$ and its fragment, black: $\text{Au}_{28}(\text{SR})_{19}$ nanocluster and its fragment, pink: $\text{Au}_{20}(\text{SR})_{15}$ nanocluster and its fragment, orange: Au_{15} nanocluster and its fragment. SR denotes the 4-TBBT ligand. (*) Asterisks represent the loss of AuL fragments, where the number with the asterisk represents how many AuL fragments are lost from the respective nanoclusters.

gold–thiol complex.³⁶ This is further supported by UV-visible spectroscopy and electrospray ionization-mass spectrometry (ESI-MS) data (Fig. S1†). Similarly, the peak of $\text{Au}_{28}(\text{SR})_{19}$ corresponds to that of the $\text{Au}_{28}(\text{SR})_{20}$ cluster, which observed with a loss of one ligand in MALDI-MS is caused by the labile nature of the aromatic ligand.^{33,37}

Case of BT

At the initial stage, 0 to 1 h in time-dependent MALDI-MS, an intense peak at 5665 m/z that corresponds to the loss of an Au_4L_4 fragment from the $\text{Au}_{25}(\text{BT})_{18}$ nanocluster was observed (Fig. 2). It indicates that the parent nanocluster, *i.e.*, $[\text{Au}_{23}(\text{CHT})_{16}]^-$ was found to be successfully transformed into the $\text{Au}_{25}(\text{BT})_{18}$ nanocluster. As the reaction proceeded further from 1 h to 3 h, an intense peak at 4263 m/z was observed that corresponds to the $\text{Au}_{15}(\text{BT})_{12}$ nanocluster. At the same time, we observed the disappearance of all the peaks corresponding to the Au_{25} nanocluster. This indicates either the Au_{25} nanocluster is transformed into another nanocluster or degraded to form the gold–thiol complex. Later, after 48 h, only one peak was observed at 8380 m/z which corresponds to the mass of $\text{Au}_{32}(\text{BT})_{19}$ which is a fragment of the Au_{36} cluster formed due to the loss of the Au_4L_4 fragment from the single ligand lost $\text{Au}_{36}(\text{BT})_{24}$ nanocluster. During the process, very less intense peaks of Au_{28} , Au_{20} , and Au_{15} nanoclusters were found, those are the side products during the specific core selection. For better understanding, the same transformation process was monitored in different mass ranges of the mass spectrometer (Fig. S2†). From 3 h to 48 h, we observed the disappearance of the Au_{25} nanocluster peak, and at the same time, we observed

the peak emergence at 8380 m/z that corresponds to $\text{Au}_{32}(\text{BT})_{19}$, formed due to the fragmentation of the Au_{36} cluster as same as the previous spectra (Fig. 2).

This clearly indicates that the Au_{25} nanocluster is involved in the formation of the Au_{36} nanocluster. Ultimately after 48 hours, only one peak was observed at 8380 m/z which corresponded to the loss of the Au_4L_4 fragment from the single ligand lost $\text{Au}_{36}(\text{BT})_{24}$ nanocluster.

Here, in the case of $\text{Au}_{25}(\text{SR})_{18}$, its Au_4L_4 fragment is found to be more pronounced in MALDI-MS spectra during transformation due to the stability of the fragment.³⁸ Similarly, in the case of the $\text{Au}_{36}(\text{SR})_{24}$ nanocluster, the peak observed in MALDI-MS is recognized by the fragment peak corresponding to $\text{Au}_{32}(\text{SR})_{19}$.³⁹

It was observed from the above two cases that the core conversion of $[\text{Au}_{23}(\text{CHT})_{16}]^-$ occurred either through the formation of Au_{23} or Au_{25} , which is completely determined by ligand parameters. The preference for core conversion towards $\text{Au}_{23}(\text{SR})_{16}$ NC is more likely to occur when there is a bulky alkyl group on the benzenethiol, whereas when the bulkiness reduces, the core conversion towards the Au_{25} NC is more preferred. For 4-IBT, 4-EBT, and 4-MBT, both Au_{23} and Au_{25} NCs were observed, and their relative amount depends on the bulkiness of the respective ligand (Fig. 3).

The reaction mixture for each case was quenched, dried and further purified through size exclusion chromatography (SEC) and the bands obtained in each case were collected separately for quantification and further studies (for more information, see the ESI†). Qualitative analysis of each band was conducted using UV-visible spectroscopy and mass spectrometry (Fig. S3–S7†). In the case of 4-TBBT, 4-IBT, and 4-EBT, we observed three bands in SEC, which corresponds to $\text{Au}_{36}(\text{SR})_{24}$ (band-1), $\text{Au}_{28}(\text{SR})_{20}$ (band-2), and the $\text{Au}_{10}(\text{SR})_{10}$ gold thiol complex (band-3) respectively (Fig. S3–S5†). While in the case of 4-MBT, we observed two bands, in which band-1 stands for $\text{Au}_{36}(\text{SR})_{24}$ and band-2 stands for $\text{Au}_{10}(\text{SR})_{10}$ (Fig. S6†). In the case of BT, we observed only one band which corresponds to the $\text{Au}_{36}(\text{SR})_{24}$ nanocluster (Fig. S7†). Each band was carefully collected separately and the solvent was evaporated to dryness. The obtained products were quantified for respective cases (Table 1).

The quantification of products obtained in each case reveals that as the bulkiness of the alkyl group (R) at the *para*-position of the aromatic thiolate ligand increases, the formation of the Au_{28} nanocluster was found to be more pronounced. At the same time, a reverse trend was observed in the case of the Au_{36} nanocluster formation. In addition to this, we also observed two final products, namely Au_{28} and Au_{36} in the case of 4-TBBT, 4-IBT, and 4-EBT. To confirm any further transformation, the reaction was continued for 11 days by maintaining similar conditions for these three cases, and the quantification of the obtained product was done similarly. In the case of 4-IBT and 4-EBT, we could obtain only a single product which was confirmed to be Au_{36} , and in the case of 4-TBBT, it gave rise to three products as we observed at 48 hours (Table 2).

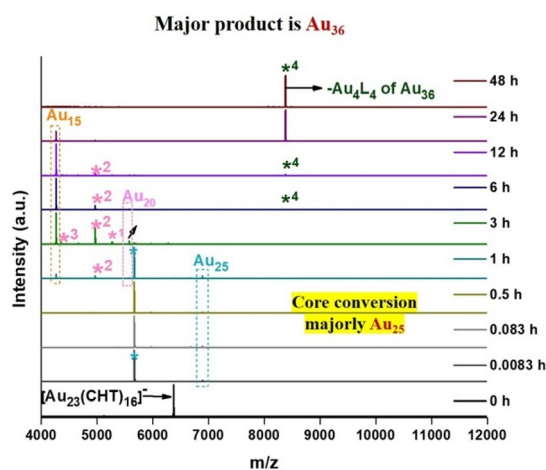


Fig. 2 Time-dependent MALDI-MS spectra showing the transformation process of the $[\text{Au}_{23}(\text{CHT})_{16}]^-$ nanocluster during the ligand exchange with the BT ligand. Note: dark green: Au_{36} nanocluster and its fragment, cyan: $\text{Au}_{25}(\text{SR})_{18}$ and its fragment, black: $\text{Au}_{28}(\text{SR})_{19}$ nanocluster and its fragment, pink: $\text{Au}_{20}(\text{SR})_{15}$ nanocluster and its fragment, orange: Au_{15} nanocluster and its fragment where SR denotes the BT ligand. (*) Asterisks represent the loss of the AuL fragment, where the number with the asterisk represents how many AuL fragments are lost from the respective nanoclusters.



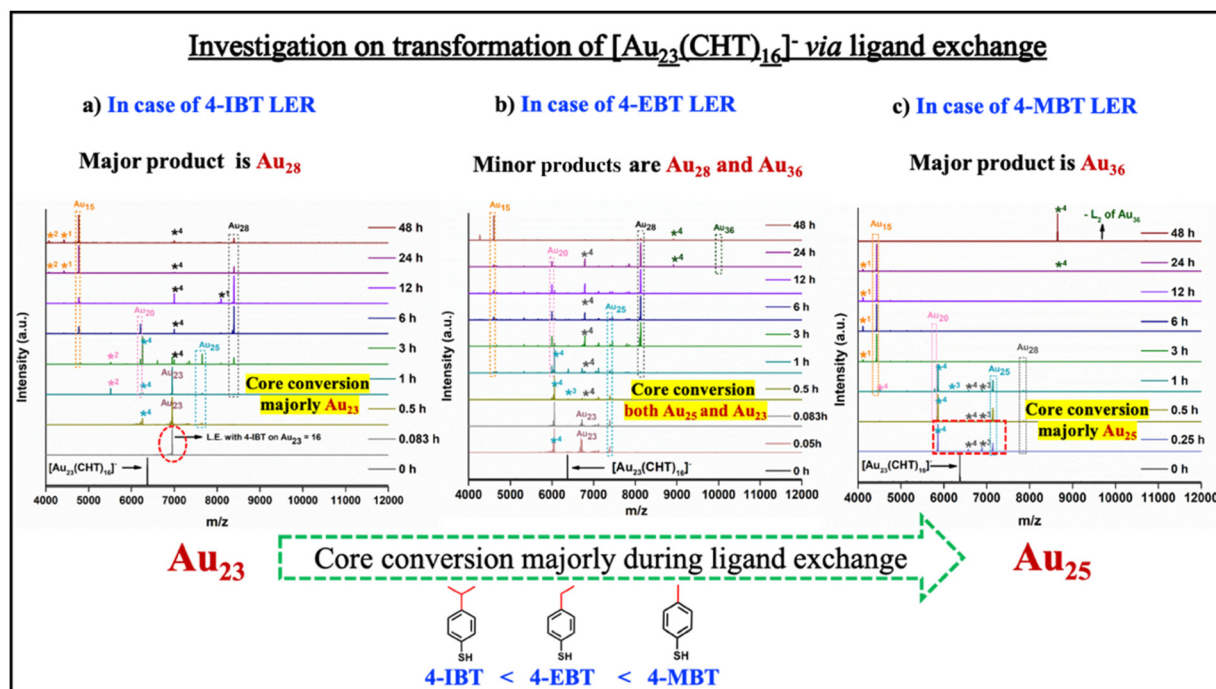


Fig. 3 Time-dependent MALDI-MS spectra showing the transformation process of the $[\text{Au}_{23}(\text{CTHT})_{16}]^-$ nanocluster during (a) 4-IBT ligand exchange, (b) 4-EBT ligand exchange, (c) 4-MBT ligand exchange. Note: dark green: Au_{36} nanocluster and its fragment, cyan: $\text{Au}_{25}(\text{SR})_{18}$ and its fragment, pink: $\text{Au}_{20}(\text{SR})_{15}$ nanocluster and its fragment, orange: Au_{15} nanocluster and its fragment. SR denotes 4-IBT, 4-EBT, and 4-MBT in the respective spectra. (*) Asterisks represent loss of the AuL fragment where the number with the asterisk represents how many AuL fragments are lost from the respective nanoclusters.

Table 1 Quantification of products obtained in the LERs after 48 h with respect to the parent $[\text{Au}_{23}(\text{CTHT})_{16}]^-$ nanocluster

Ligand (-SR)	BT	4-MBT	4-EBT	4-IBT	4-TBBT
Yield (%) of $\text{Au}_{36}(\text{SR})_{24}$ w.r.t. precursor	25	17.8	12.8	8.3	~2
Yield (%) of $\text{Au}_{28}(\text{SR})_{20}$ w.r.t. precursor	0	0	14	21	38.4
Yield (%) of $\text{Au}_{10}(\text{SR})_{10}$ w.r.t. precursor	0	15.3	69.2	75.6	57.6

Table 2 Comparison of yield obtained at two different times as well as at the same time

Ligand	4-EBT	4-IBT	4-TBBT
Quantification of the products obtained after 48 hours			
Yield (%) of $\text{Au}_{36}(\text{SR})_{24}$ w.r.t. precursor	12.8	8.3	~2
Yield (%) of $\text{Au}_{28}(\text{SR})_{20}$ w.r.t. precursor	14	21	38.4
Yield (%) of $\text{Au}_{10}(\text{SR})_{10}$ w.r.t. precursor	69.2	75.6	57.6

vs.

Quantification of the products obtained after 11 days

Yield (%) of $\text{Au}_{36}(\text{SR})_{24}$ w.r.t. precursor	12.8	8.3	~2
Yield (%) of $\text{Au}_{28}(\text{SR})_{20}$ w.r.t. precursor	0	0	23.4
Yield (%) of $\text{Au}_{10}(\text{SR})_{10}$ w.r.t. precursor	0	0	71.7

The comparison of the percentage of the yield of the products obtained at 48 h with 11 days' reactions led to two main outcomes:

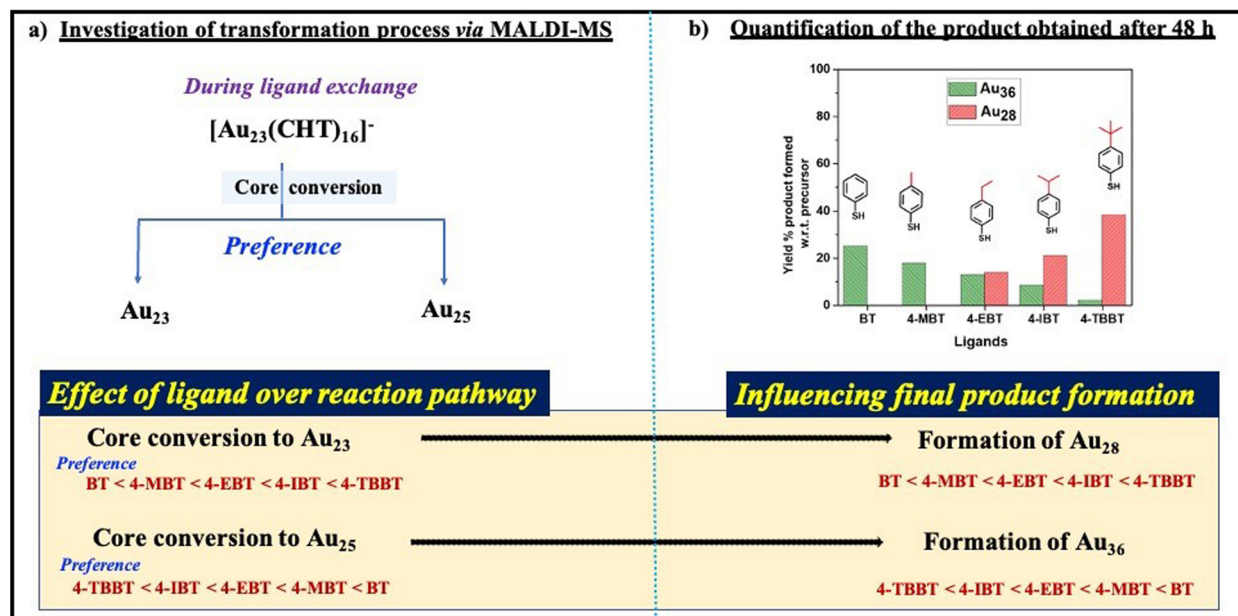
Point 1: No variation in the amount of the Au_{36} nanocluster even after 11 days indicated that the Au_{28} nanocluster was not involved in the formation of the Au_{36} nanocluster in any of the cases. This also suggested that the formation of Au_{28} and Au_{36} nanoclusters are from two different routes.

Point 2: In the case of 4-TBBT, the amount of the Au_{28} nanocluster was found to be decreased at the same time, we observed an increment in the amount of Au_{10} , indicating that Au_{28} degraded to form the gold-thiol complex. While in cases of 4-IBT and 4-EBT, we observed the decomposition of both the Au_{28} nanocluster and Au_{10} gold-thiol complex as the reaction proceeded. These results suggest that the bulkiness of the alkyl group favors Au_{28} nanocluster formation and provides stability.

The facts collected through the investigation of the transformation process and quantification of products after 48 hours (*i.e.*, from Schemes 2a and b, respectively), indicate that the formation of the Au_{28} nanocluster was caused by the transformation of the parent nanocluster through the $\text{Au}_{23}(\text{SR})_{16}$ intermediate during ligand exchange (Scheme 2). Whereas the formation of the Au_{36} nanocluster was caused by the transformation of the parent nanocluster to the $\text{Au}_{25}(\text{SR})_{18}$ intermediate (Scheme 2). This indicates that the selectivity and stability of a specific core during the ligand exchange are well governed by ligand parameters which is evident in this example.

The ligand effect over cluster stability

To understand the role of ligands on the stability of the obtained products, a stability test was done by heating the



Scheme 2 (a) Summary of the product formation based on time-dependent MALDI-MS spectra obtained for the LER process on the $[\text{Au}_{23}(\text{CHT})_{16}]^-$ nanocluster transformation with the series of ligands. (b) Summary obtained from the quantification of the product analysis for the $[\text{Au}_{23}(\text{CHT})_{16}]^-$ nanocluster transformation process with the series of ligands.

solution of Au_{28} and Au_{36} clusters at 40 °C. The solution was characterized using UV-visible spectroscopy conducted over a period. The results obtained for $\text{Au}_{28}(\text{SR})_{20}$ nanoclusters (where SR = 4-TBBT, 4-IBT, and 4-EBT respectively) is shown in Fig. 4.

In the case of 4-TBBT, the Au_{28} nanocluster was found to be stable even after one week. While in the case of 4-IBT, the UV-visible absorption spectra collected during the experiment changed after 72 hours, indicating less stability of $\text{Au}_{28}(4\text{-IBT})_{20}$ compared to $\text{Au}_{28}(4\text{-TBBT})_{20}$. For $\text{Au}_{28}(4\text{-EBT})_{20}$, the UV-visible spectra changed after 12 hours indicating the least stability of $\text{Au}_{28}(4\text{-EBT})_{20}$ among the three clusters. So, the trend obtained for the stability was found to be $\text{Au}_{28}(4\text{-TBBT})_{20} > \text{Au}_{28}(4\text{-IBT})_{20} > \text{Au}_{28}(4\text{-EBT})_{20}$. It indicates that the bulkiness

of the alkyl group at the *para* position of the aromatic thiolate ligand has a role in determining the stability of the Au_{28} cluster and it increases with the bulkiness.

The UV-Visible spectra showing the stability test of $\text{Au}_{36}(\text{SR})_{24}$ are given in Fig. S8.† Similar to the case of $\text{Au}_{28}(\text{SR})_{20}$ clusters, 4-TBBT-protected Au_{36} was the most stable among all the five clusters where the spectral features existed as such for more than 2 weeks. Whereas $\text{Au}_{36}(\text{BT})_{24}$ was the least stable and started to degrade within 6 hours. The order of stability was found to be $\text{Au}_{36}(4\text{-TBBT})_{24} > \text{Au}_{36}(4\text{-IBT})_{24} > \text{Au}_{36}(4\text{-EBT})_{24} > \text{Au}_{36}(4\text{-MBT})_{24} > \text{Au}_{36}(\text{BT})_{24}$. Here we could observe that even though the formation of the Au_{36} cluster is least preferred by the 4-TBBT ligand, $\text{Au}_{36}(4\text{-TBBT})_{24}$ is the most stable cluster among all the $\text{Au}_{36}(\text{SR})_{24}$ clusters. So, the

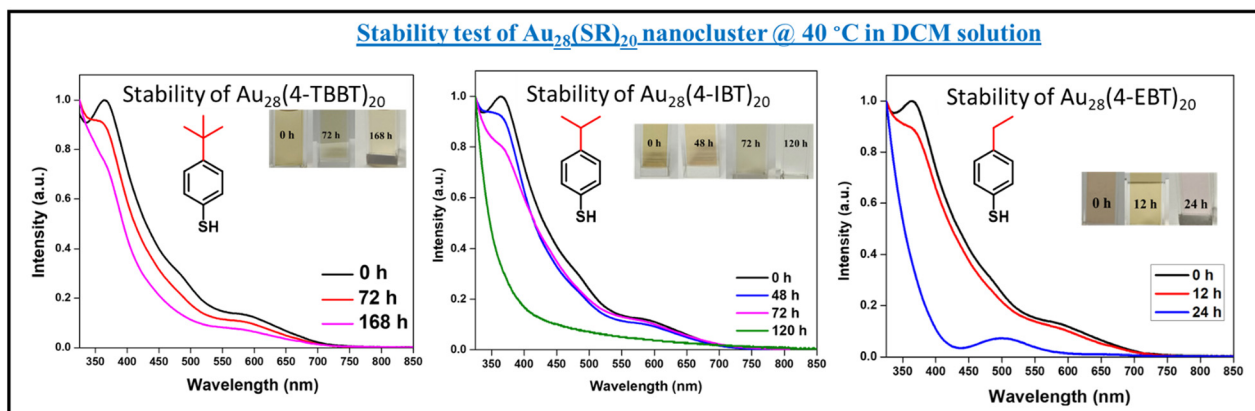
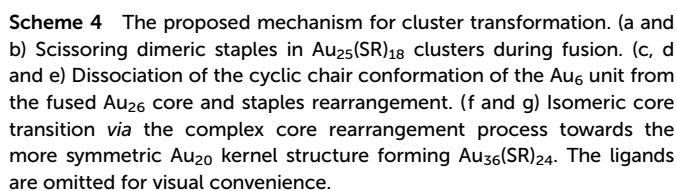
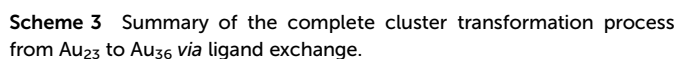


Fig. 4 UV-visible absorption spectra showing the stability test of $\text{Au}_{28}(\text{SR})_{20}$.



It has been observed in the experiment that the $\text{Au}_{36}(\text{BT})_{24}$ nanocluster is formed from the parent cluster $[\text{Au}_{23}(\text{CHT})_{16}]^-$ through the immediate formation of $[\text{Au}_{25}(\text{BT})_{18}]^-$ without any signature of $\text{Au}_{28}(\text{BT})_{20}$ *via* ligand exchange. This led us to

Scheme 3 Summary of the complete cluster transformation process from Au₂₃ to Au₃₆ via ligand exchange.



to each other, the feasibility of the isomeric symmetry assisted core transition followed by dissociation of the cyclic Au₆ unit is expected to be the reliable pathway for the ligand exchange process studied in our work.

In summary, we have shown that the electronic effect and a variation in the steric effect of the incoming ligands have a crucial role in the size selectivity and preference of the pathway of transformation during the ligand exchange reaction. We observed that when the ligands have a bulkier substituent, $[\text{Au}_{23}(\text{CHT})_{16}]^-$ preferred the transformation through the $\text{Au}_{23}(\text{SR})_{16}$ intermediate which led to producing $\text{Au}_{28}(\text{SR})_{20}$ as the major product. Whereas when the bulkiness gradually decreased, the reaction went through the pathway that favored the formation of the $\text{Au}_{36}(\text{SR})_{24}$ cluster through the $\text{Au}_{25}(\text{SR})_{18}$ intermediate. The influence of ligands in choosing the transformation pathway was evident from the gradual trend observed in the yield of the final products obtained. Steric crowding of the ligands not only dictates the pathway for the transformation, it also alters the stability of the clusters. A plausible mechanism based on symmetry-assisted isomeric core rearrangement is also proposed for Au_{25} to Au_{36} cluster

Nanoscale, 2023, 15, 13102–13109 | 13109