



Cite this: *CrystEngComm*, 2016, 18, 4438

Received 3rd March 2016,
Accepted 13th April 2016

DOI: 10.1039/c6ce00496b

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Janus head type lone pair– π –lone pair and S \cdots F \cdots S interactions in retaining hexafluorobenzene†

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A series of eight tris-arylthiotriazines were synthesized to study the lone pair- π interaction between the triazine ring centroid of these molecules and halogenated solvents. All the eight compounds were characterized using ^1H and ^{13}C NMR spectroscopy and single crystal X-ray diffraction techniques. All these compounds show interesting structural properties in the solid state. Unprecedented Janus head type $\text{lp}\cdots\pi\cdots\text{lp}$ and $\text{S}\cdots\text{F}\cdots\text{S}$ interactions were observed between one of the tris-arylthiotriazines and hexafluorobenzene.

1. Introduction

It is apparent from the recent literature that lone pair- π ($\text{lp}\cdots\pi$) interactions are weak yet important for the stabilization of three-dimensional structures of biomolecules¹ in addition to other well-established weak interactions such as hydrogen bonding, anion- π interactions, and cation- π interactions. Some experimental evidence has been reported very recently to establish the applications of anion- π interactions²⁻⁶ and cation- π interactions⁷⁻¹⁰ while the $\text{lp}\cdots\pi$ interactions are yet to be proved useful in small molecular systems. To the best of our knowledge, all the reported synthetic molecular systems that exhibit $\text{lp}\cdots\pi$ interactions were discovered serendipitously. We hypothesised to devise molecular systems that will have lone pair- π interactions with small molecules of environmental and industrial interest (CFCs, CO_2 , SO_2) as these interactions may be strong enough to sequester the small molecules and weak enough to release them with minimal energy input such as slight elevation of temperature and/or depression in pressure. This prompted us to investigate host molecules comprising electron deficient ring systems in order to demonstrate the functional relevance of $\text{lp}\cdots\pi$ interactions. With this motivation, we have synthesized a number of tris-arylthiotriazine molecules (TArTTz), as the valley of electron deficiency present at the centre of the C_3N_3 ring of these molecules^{6,11-13} is known to attract lone pairs of electronegative atoms of neighbouring molecules.¹⁴⁻¹⁶ We have anticipated that this will allow us to park small molecules in between C_3N_3 rings of adjacent molecules through $\text{lp}\cdots\pi$ interactions. For this, we

wished to take advantage of the well-known “Piedfort Diad” formation (Chart 1).^{17–23} According to literature reports^{17,18} *sym*-triaryloxytriazines (TArOTz) have an interesting structural feature of peripheral aryl rings oriented almost perpendicular to the central C₃N₃ ring and form stacked diads known as Piedfort units²¹ which eventually are the synthons for higher level supramolecular assemblies.^{17–19}

One of the inherent disadvantages of the known ‘Piedfort Diads’ is that the distance $D_{T_2 \cdots T_2}$ between the C_3N_3 rings are in the range of 3 and 4 Å whereas to park a small molecule in between two *sym*-triaryloxytriazines through $lp \cdots \pi$ interaction(s), a minimum of ~ 8 Å distance might be required.^{17–19} To achieve this we have intended to utilize arylthio substitutions instead of aryloxy units where the aryl rings comprise varying substitutions in order to tune the distance between adjacent C_3N_3 rings and the type of interactions between neighbouring molecules.

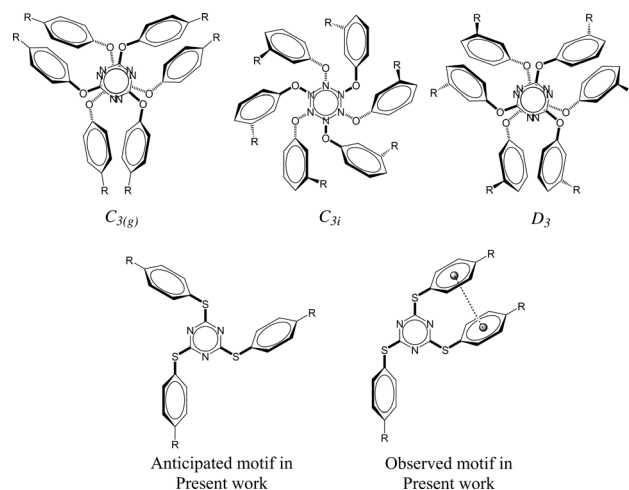


Chart 1 Piedfort units of tris-aryloxotriazines (TAROTz) with 3-fold symmetry ($C_{3(g)}$, C_{3i} , and D_3)^{17–21,23–25} and the motif anticipated and observed in the present work for tris-arylthiotriazines (TARTTz).

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† Electronic supplementary information (ESI) available: NMR spectra of all the compounds. X-ray crystallographic data in CIF format have been deposited with the Cambridge Structural Database. CCDC 1403546–1403550 (1–5), 1436791–1436793 (6–0.5C₆F₆, 6, 7) and 1452878 (8) contain the supplementary crystallographic data for this paper. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c6ce00496b

2. Results and discussion

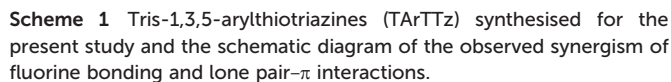
An equivalent of cyanuric chloride was reacted with three equivalents of thiophenol, 4-methylthiophenol, 4-chlorothiophenol, 3-nitrobenzenethiol, 2-methyl-5-nitrobenzenethiol, 2,6-dimethylthiophenol, 2,3,4,5,6-pentamethylthiophenol and 2,3,5,6-tetrafluorothiophenol in the presence of a suitable base to obtain 2,4,6-tris(phenylthio)-1,3,5-triazine (1), 2,4,6-tris(p-tolylthio)-1,3,5-triazine (2),²⁶ 2,4,6-tris((4-chlorophenyl)thio)-1,3,5-triazine (3), 2,4,6-tris((3-nitrophenyl)thio)-1,3,5-triazine (4), 2,4,6-tris((2-methyl-5-nitrophenyl)thio)-1,3,5-triazine (5), 2,4,6-tris(2,6-dimethylphenylthio)-1,3,5-triazine (6), 2,4,6-tris(2,6-dimethylphenylthio)-1,3,5-triazine (7) and 2,4,6-tris((2,3,5,6-tetrafluorophenyl)thio)-1,3,5-triazine (8), respectively (Scheme 1). All the eight compounds were characterized using ¹H and ¹³C NMR spectroscopy and the single-crystal X-ray diffraction technique. The thermal stability of these molecules was studied by thermogravimetric analysis.

The NMR spectra of all the eight TarTTz compounds indicate that these molecules are C_3 symmetric in solution at room temperature in NMR time scale though it is not the case in the solid state as seen in the SCXRD structures of these TarTTz compounds. As ^1H NMR spectroscopy does not differentiate the mono and disubstituted triazines from TarTTz, ^{13}C NMR spectroscopy was used as the finite preliminary analysis where all the TarTTz exhibit a signal around 180 ppm indicating the presence of three similar tertiary carbons in the triazine ring system.

Single crystals of TArTTz 1–8 suitable for SCXRD studies were grown by slow evaporation or diffusion methods as reported in the experimental section in detail. Unfortunately, none of the TArTTz molecules crystallised in any of the three required symmetries ($C_{3(g)}$, C_{3i} or D_3),^{17–19} which would have allowed the formation of the anticipated Piedfort units. To the best of our knowledge, many known free tris-aryloxotriazines form Piedfort units in their solid states.^{17–23} However, TAROTz are also known to attain lower symmetry when they are coordinated to metal ions.²⁷ One possible reason for the lowering of symmetry on going from TAROTz to TArTTz could be the possibility of intramolecular $\pi\cdots\pi$ interaction in TArTTZ molecules owing to the bigger size of sulphur. All the TArTTz molecules exhibit weak intramolecular $\pi\cdots\pi$ interactions of distances between 3.94 and 4.38 Å, which were not retained in solution at room temperature as observed from their NMR spectra.

The TARTTz 1, where the phenyl rings do not have any substitution, possesses an intermolecular lone pair $\cdots\pi$ interaction between the lone pair of thioether sulphur and the triazine ring (S(lp) $\cdots\pi$) with a distance ($D_{\text{S(lp)}\cdots\pi}$) of 3.207 Å which is shorter than the sum of van der Waals radii (C + S = 3.5 Å; N + S = 3.35 Å).¹³ The S3 A \cdots ring distances are 3.471 Å (S3 A \cdots N1B), 3.472 Å (S3 A \cdots N2B), 3.527 Å (S3 A \cdots N3B), 3.391 Å (S3 A \cdots C1B), 3.453 Å (S3 A \cdots C2B), and 3.525 Å (S3 A \cdots C3B). This may be the first observed S(lp) $\cdots\pi$ involving a 1,3,5-triazine ring system whereas few other known S(lp) $\cdots\pi$ involve 1,2,3-triazine and 1,3,4-triazine rings (Fig. 1).¹³ The 4-methyl substituted TARTTz 2 does not have any significant lp $\cdots\pi$ interactions whereas the 4-chloro substituted TARTTz 3 exhibits an intermolecular Cl(lp) $\cdots\pi$ interaction ($D_{\text{Cl(lp)}\cdots\pi}$ = 3.542 Å). The sums of the van der Waals radii are 3.45 Å (C + Cl) and 3.30 Å (N + Cl);¹³ the Cl2 \cdots ring distances are 3.648 Å (Cl2 \cdots N1), 4.181 Å (Cl2 \cdots N2), 3.550 Å (Cl2 \cdots N3), 3.981 Å (Cl2 \cdots C1), 3.931 Å (Cl2 \cdots C2), and 3.353 Å (Cl2 \cdots C3). According to Mooibroek *et al.*, the observed mean $D_{\text{Cl(lp)}\cdots\pi}$ is 3.346 Å.¹² Hence, the Cl(lp) $\cdots\pi$ interaction observed in TARTTz 3 may be considered as a weak one.

The 3-nitro substituted TARITZ 4 has all three nitro groups facing the same side of the triazine ring while the 2-methyl-5-nitro substituted TARITZ 5 has two nitro groups in one side and the other one facing the opposite side of the triazine ring. This indicates that the 2-methyl-5-nitrophenyl rings have free rotation



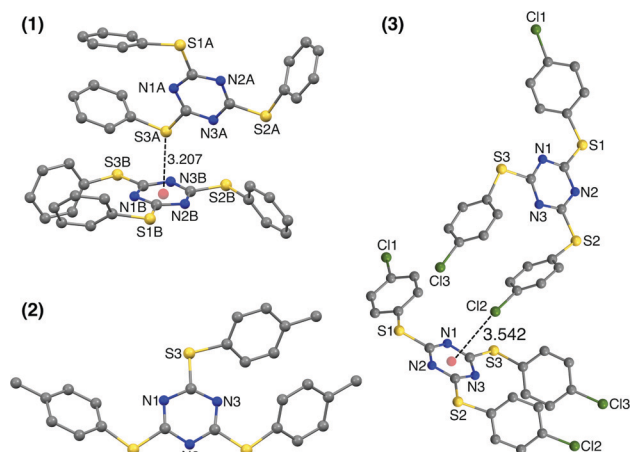


Fig. 1 X-ray structures of 2,4,6-tris(phenylthio)-1,3,5-triazine (1), 2,4,6-tris(p-tolylthio)-1,3,5-triazine (2), and 2,4,6-tris((4-chlorophenyl)thio)-1,3,5-triazine (3). The hydrogen atoms are not shown for clarity. Intramolecular $\pi\cdots\pi$ interactions: (1) 4.287 and 4.350; (2) 4.384; (3) 4.158 Å.

in solution as TARITZ 5 has C_3 symmetry in solution. Apart from the intramolecular $\pi\cdots\pi$ interactions, TARITZ 4 has an interesting intermolecular $\pi\cdots\pi$ interaction ($D_{Cg\cdots Cg} = 3.616$ Å) between the triazine ring of one molecule with the 3-nitrophenyl ring of another molecule. The TARITZ 5 possesses an intermolecular $\pi\cdots\pi$ interaction ($D_{Cg\cdots Cg} = 3.652$ Å), which presents a ladder of four aryl rings (Fig. 2). The TARITZ 6 and 7 also present the same ladder-type structure of four dimethylphenyl rings and four pentamethylphenyl rings with intermolecular ($D_{Cg\cdots Cg} = 3.868$ Å in 6; 3.913 Å in 7) and intramolecular $\pi\cdots\pi$ stacking ($D_{Cg\cdots Cg} = 3.965$ Å in 6; 4.114 Å in 7).

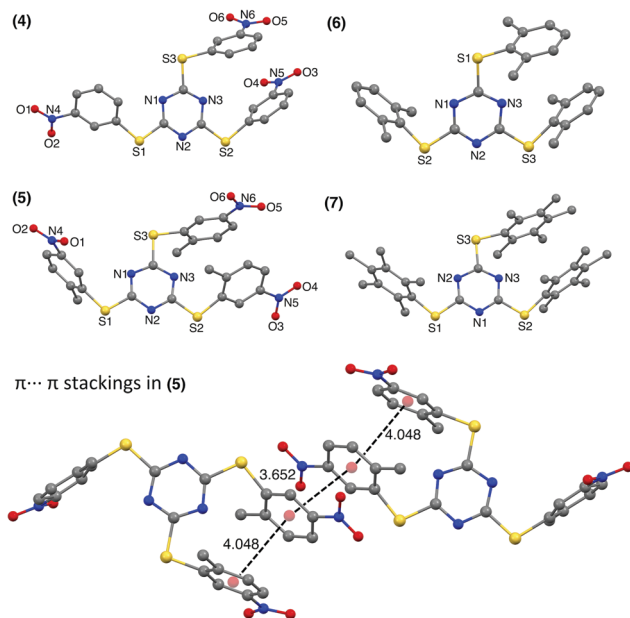


Fig. 2 X-ray structures of 2,4,6-tris((3-nitrophenyl)thio)-1,3,5-triazine (4), 2,4,6-tris((2-methyl-5-nitrophenyl)thio)-1,3,5-triazine (5), 2,4,6-tris(2,6-dimethylphenylthio)-1,3,5-triazine (6), and 2,4,6-tris(2,6-dimethylphenylthio)-1,3,5-triazine (7). The hydrogen atoms are not shown for clarity. Intramolecular $\pi\cdots\pi$ interactions: (4), 4.041; (5), 4.048; (6), 3.965; (7), 4.114 Å. Intermolecular $\pi\cdots\pi$ interactions: (4), 3.616; (5), 3.652; (6), 3.868; (7), 3.913 Å.

The SCXRD structure of TARITZ 8 (Fig. 3) presents an interesting dimeric motif exhibiting many intermolecular interactions such as $S\cdots F$ fluorine bonds ($S3\cdots F8$, 2.991 Å; $S1\cdots F1$, 3.251 Å; $S2\cdots F2$, 3.498 Å), an intermolecular $F(lp)\cdots\pi$ interaction ($D_{F(lp)\cdots\pi} = 2.967$ Å) and an intramolecular $\pi\cdots\pi$ interaction ($D_{Cg\cdots Cg} = 3.944$ Å). The $F(lp)\cdots\pi$ interaction observed in TARITZ 8 is a strong one as the $D_{F(lp)\cdots\pi}$ of 2.967 Å is well below the sums of the van der Waals radii ($C + F = 3.17$ Å; $N + F = 3.02$ Å).¹³ The $F4\cdots$ ring distances are 3.273 Å ($F4\cdots N1$), 3.111 Å ($F4\cdots N2$), 3.414 Å ($F4\cdots N3$), 3.404 Å ($F4\cdots C1$), 3.077 Å ($F4\cdots C2$), and 3.212 Å ($F4\cdots C3$) which are not too far from the sum of the van der Waals radii.

As we have observed the intermolecular $lp\cdots\pi$ interactions in TARITZ 1, 3 and 8, we set out to investigate the interactions with fluorinated molecules as solvents. The halogenated aromatic molecules, especially C_6F_6 , are known to be involved in $lp\cdots\pi$ interactions as they can participate through the halogen lone pair and through the electron deficient centre of the aromatic ring. We have observed a peculiar crystal structure where the C_6F_6 molecule involves in Janus head type of $lp\cdots\pi\cdots lp$ interaction on both sides of the ring ($D_{N(lp)\cdots\pi} = 3.51$ Å) where the nitrogen lone pairs of two TARITZ 6 molecules interact with the ring centre of C_6F_6 (Fig. 4). Furthermore, the fluorine atoms at the 1,4-positions of the C_6F_6 ring also take part in fluorine bonding^{28–31} with the thioether sulphur atoms of the TARITZ 6 molecules with interesting $S\cdots F\cdots S$ motifs ($D_{S\cdots F} = 3.365$ and 3.442 Å). Apart from the $lp\cdots\pi\cdots lp$ and $S\cdots F\cdots S$ interactions, the structure of $6\cdot 0.5C_6F_6$ possesses extensive inter and intramolecular $\pi\cdots\pi$ stacking networks.

2.4. Thermal stability of TARITZ molecules

The thermal stability of TARITZ compounds 1–8 was studied by means of thermogravimetric analysis. Compounds 1–3, 6 and 7 were found to be stable until ~ 280 °C, 4 was stable until ~ 260 °C, whereas compounds 5 and 8 began to decompose at ~ 230 °C. Interestingly, adduct compound $6\cdot 0.5C_6F_6$ lost the hexafluorobenzene molecule at 140 °C (Fig. 5) which is 60 °C above the boiling point of C_6F_6 which indicates the synergism of the $lp\cdots\pi\cdots lp$ and $S\cdots F\cdots S$ interactions in retaining the C_6F_6 molecule in the lattice.

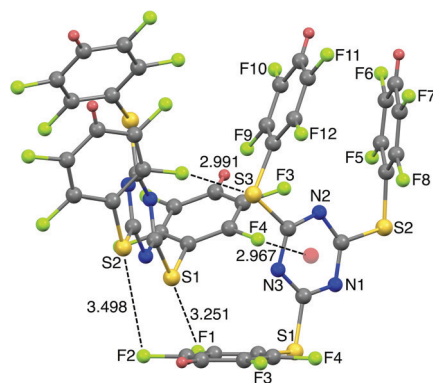


Fig. 3 X-ray structure of 2,4,6-tris((2,3,5,6-tetrafluorophenyl)thio)-1,3,5-triazine (8). Intramolecular $\pi\cdots\pi$ interaction distance: 3.944 Å.



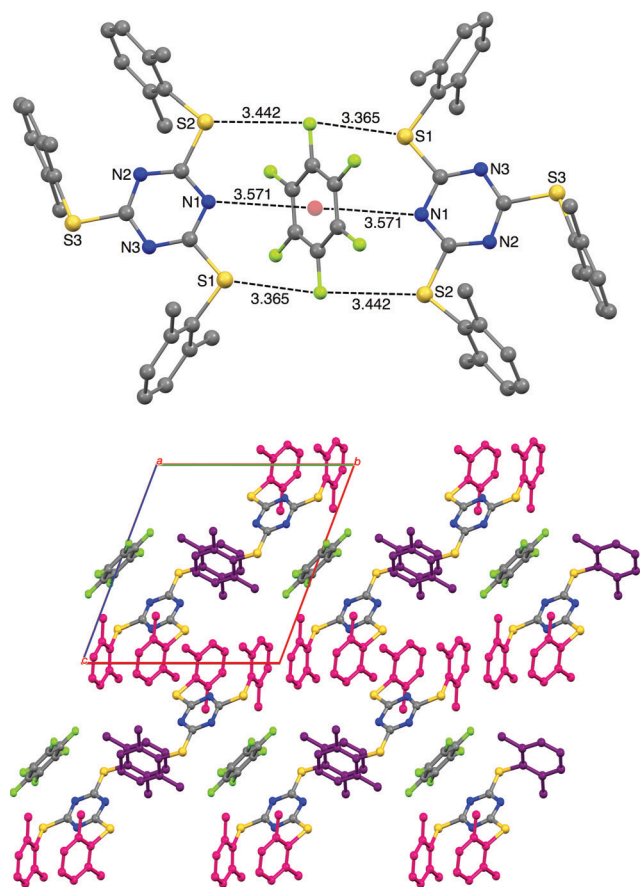


Fig. 4 X-ray structure of 2,4,6-tris(2,6-dimethylphenylthio)-1,3,5-triazine (**6**) crystallized in hexafluorobenzene (**6**·0.5C₆F₆). The hydrogen atoms are not shown for clarity. The packing diagram of **6**·0.5C₆F₆ shows the presence of extensive inter and intramolecular $\pi\cdots\pi$ stacking. Pink coloured dimethylphenyl rings are involved in both inter and intramolecular stacking and the purple coloured rings are involved in the intermolecular stacking. $\pi\cdots\pi$ interactions: 4.100 Å (intra, pink); 4.160 Å (inter, pink); 4.280 Å (inter, purple).

3. Conclusions

The tris-arylthiotriazines reported herein show that the lone pair– π interactions can indeed be engineered by design and be used for applications such as capturing or retaining small molecules which possess at least one lone pair (halocarbons, CO₂, or SO₂). The observed feature of the synergistic Janus head lp $\cdots\pi$ –lp and S \cdots F \cdots S motif is interesting in the context of fluorine bonding and warrants further experimental and computational studies as the fluorine bonding is a unique and emerging section of weak interactions which might be relevant to pharmacologically important fluorinated drug molecules.²⁹

4. Experimental

4.1. Materials

Cyanuric chloride (Avra Synthesis), thiophenol (SDFCL), 4-methylthiophenol (Avra Synthesis), 4-chlorothiophenol (Avra Synthesis), nitrobenzene (Merck), chlorosulfonic acid (Spectrochem), triphenylphosphine (Avra synthesis), 4-nitrotoluene

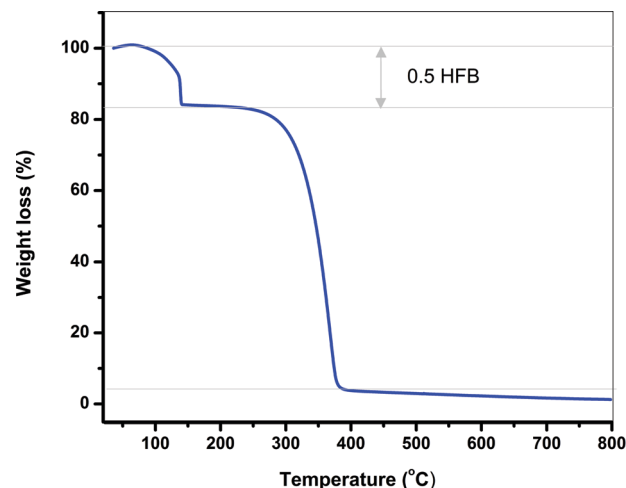


Fig. 5 The thermogravimetric diagram of **6**·0.5C₆F₆ showing the loss of the C₆F₆ solvate at 140 °C and the decomposition of **6** at 280 °C.

(Avra), 2,6-dimethylthiophenol (Alfa Aesar), 2,3,4,5,6-pentamethylbenzene-1-sulfonylchloride (Alfa Aesar) and 2,3,5,6-tetrafluorothiophenol (Alfa Aesar) were used as received from commercial sources. Solvents were distilled under dry nitrogen atmosphere using conventional methods. Heavy-walled high-pressure round bottom flasks were acquired from Ace Glass Inc.

4.2. Methods

Elemental analyses were carried out on a Perkin-Elmer CHNS/O analyser. The NMR spectra were recorded on JEOL 500 MHz and JEOL 400 MHz spectrometers. The temperature was kept constant using a variable temperature unit within the error limit of ± 1 K. The software MestReNova was used for the processing of the NMR spectra.³² Tetramethylsilane (TMS) or the deuterated solvent residual peaks were used for calibration. Mass spectrometry experiments were performed on a Waters-Q-ToF-Premier-HAB213 equipped with an electrospray interface. The spectra were collected by constant infusion of the sample dissolved in methanol or acetonitrile with 0.1% formic acid. Thermogravimetric analyses were carried out in dry nitrogen atmosphere at a heating rate of 10 °C min^{−1} on a Mettler Toledo TGA/DSC1 Star^c System.

Warning: Reactions using heavy-walled high-pressure round bottom flask with screw cap must be carried cautiously. The danger of explosion can be avoided by performing the reactions in smaller scales. The scales presented here are safe with a 250 ml heavy-walled high-pressure round bottom flask.

4.3. Synthetic procedures

4.3.1. Synthesis of 2,4,6-tris(phenylthio)-1,3,5-triazine (1). Thiophenol (3.3 g, 30 mmol) was added to a solution of cyanuric chloride (1.8 g, 10 mmol) in 80 ml acetone. The resulting solution was treated with an aqueous NaOH solution (1.2 g, 30 mmol in 40 ml H₂O) dropwise at 0 °C with continuous stirring for 1 h. Then the reaction mixture was stirred for 4 h at 35 °C. The resulting white precipitate was collected through filtration and washed with cold water



(10 ml), followed by cold ethanol (2 ml) and dried under high vacuum (3.4 g, 87%). X-ray quality crystals were grown by the slow evaporation of ethanol solution. ^1H NMR (500 MHz, CDCl_3): δ 7.36 (t, J = 6.9, 6H), 7.32 (t, J = 8.0, 6H), 7.23 (t, J = 8.0, 3H); ^{13}C NMR (500 MHz, CDCl_3): δ 180.43, 134.59, 129.22, 128.95, 127.17. Anal. calcd (%) for $(\text{C}_{21}\text{H}_{15}\text{N}_3\text{S}_3 \cdot 0.2\text{CH}_3\text{CH}_2\text{OH})$: C 61.97, H 3.94, N 10.13; found: C 61.86, H 3.61, N 10.62.

4.3.2. Synthesis of 2,4,6-tris(p-tolylthio)-1,3,5-triazine (2). This compound has been reported in the literature and it was synthesised by the reaction of 1,3,5-triazine-2,4,6-trithiol with 4-bromotoluene catalysed by palladium nanoparticles.³³ 4-methylthiophenol (3.7 g, 30 mmol) was added to a solution of cyanuric chloride (1.8 g, 10 mmol) in 80 ml acetone. The resulting solution was treated with an aqueous NaOH solution (1.2 g, 30 mmol in 40 ml water) dropwise at 0 °C and stirred for an hour at the same temperature. Then the reaction mixture was stirred for 4 h at 35 °C. The resulting white precipitate was collected through filtration and washed with cold water (15 ml), followed by cold ethanol (2 ml) and dried under high vacuum (3.2 g, 74%). Needle-shaped crystals were collected through slow evaporation of methanol within two days. ^1H NMR (500 MHz, CDCl_3): δ 7.23 (d, J = 7.95, 6H), 7.04 (d, J = 7.95, 6H), 2.36 (s, 9H); ^{13}C NMR (500 MHz, CDCl_3): δ 180.60, 139.05, 134.92, 129.72, 123.86. Anal. calcd (%) for $(\text{C}_{24}\text{H}_{21}\text{N}_3\text{S}_3)$: C 64.39, H 4.73, N 9.39; found: C 64.27, H 4.61, N 9.83.

4.3.3. Synthesis of 2,4,6-tris((4-chlorophenyl)thio)-1,3,5-triazine (3). 4-chlorothiophenol (4.33 g, 30 mmol) was added to a solution of cyanuric chloride (1.8 g, 10 mmol) in 80 ml acetone. The resulting solution was treated with an aqueous NaOH solution (1.2 g, 30 mmol in 40 ml H_2O) dropwise at 0 °C and stirred for an hour at the same temperature. The resulting white precipitate was collected through filtration and washed with cold water (10 ml), followed by cold ethanol (2 ml) and dried under high vacuum (4.0 g, 81%). Needle-shaped crystals were collected from a methanolic solution through slow evaporation. ^1H NMR (500 MHz, CDCl_3): δ 7.27 (d, J = 9.1, 6H), 7.23 (d, J = 8.5, 6H); ^{13}C NMR (500 MHz, CDCl_3): δ 180.29, 136.47, 136.32, 129.41, 125.38. Anal. calcd (%) for $(\text{C}_{21}\text{H}_{12}\text{Cl}_3\text{N}_3\text{S}_3)$: C 49.56, H 2.38, N 8.26; found: C 49.63, H 2.31, N 8.13.

4.3.4. Synthesis of 3-nitrobenzene-1-sulfonyl chloride. A 250 ml heavy-walled high-pressure round bottom flask (Ace Glass, Inc.) was charged with nitrobenzene (3 ml, 29.23 mmol) and 5 ml of cold chlorosulfonic acid (77.22 mmol) was added slowly at room temperature with continuous stirring; the flask was sealed with a Teflon screw cap. The reaction mixture was heated at 40 °C for 30 min and then the temperature was increased by 20 °C every 30 minutes until it reached 140 °C. Then the reaction mixture was stirred at 140 °C for 40 hours. The resulting brown solution was cooled to room temperature before opening the cap and poured onto ~50 ml of crushed ice, stirred for 5 minutes, extracted with CHCl_3 (3×15 ml), then all the organic layers were combined, dried over MgSO_4 and the solvent was evaporated to obtain the crude batch of 3-nitrobenzene-1-sulfonyl chloride as a brown solid (63%)

which was purified by passing through a silica column as a solution in chloroform (6.5 g, 56%). ^1H NMR (500 MHz, CDCl_3): δ 8.88 (d, 1H), 8.61 (dd, 1H), 8.38 (dd, 1H), 7.90 (t, 1H); ^{13}C NMR (500 MHz, CDCl_3): δ 148.36, 145.58, 132.20, 131.52, 129.67, 122.31.

4.3.5. Synthesis of 3-nitrobenzenethiol. Triphenylphosphine (9.87 g, 37.63 mmol) was slowly added to a stirred solution of 3-nitrobenzene-1-sulfonyl chloride (2.78 g, 12.54 mmol) in toluene (30 ml) under nitrogen atmosphere, which resulted in a brown solution. To this, a solution of iodine in toluene (5.3 mg, 2 ml) was added dropwise and the resulting mixture was refluxed for 3 h. Addition of an iodine solution changed the colour of the reaction mixture from brown to yellow. The resulting yellow solution was cooled to room temperature and treated with 5 ml of water, which resulted in a turbid solution. The organic layer was extracted with 10% aqueous NaOH solution (3×15 ml). The aqueous alkali layer was acidified to pH 3 with HCl (6 N), extracted with dichloromethane (3×15 ml) and dried over MgSO_4 . The solvent was removed and dried under high vacuum to yield a yellow liquid (1.75 g, 89%). ^1H NMR (400 MHz, CDCl_3): δ 8.09 (s, 1H), 7.96 (d, J = 7.65, 1H), 7.54 (d, J = 7.65, 1H), 7.39 (t, J = 7.95, 1H), 3.72 (s, 1H); ^{13}C NMR (500 MHz, CDCl_3): δ 148.23, 134.51, 133.84, 129.66, 123.29, 120.28.

4.3.6. Synthesis of 2,4,6-tris((3-nitrophenyl)thio)-1,3,5-triazine (4). A solution of 3-nitrobenzenethiol (1.75 g, 11.2 mmol) in 20 ml of acetone was added to a stirred solution of cyanuric chloride (688 mg, 3.73 mmol) in 20 ml of acetone under nitrogen atmosphere. After that, NaOH (447 mg, 11.2 mmol) in water (30 ml) was added dropwise (during the addition of aqueous NaOH solution, a yellow precipitate was formed). This yellow reaction mixture was stirred at 0 °C for 1 h, at room temperature for 2 h and at 35 °C for 2 h. The resulting yellow precipitate was collected through filtration and dried under vacuum (1.1 g, 55%). Crystals were grown by the slow diffusion of diethyl ether into an acetonitrile solution. ^1H NMR (500 MHz, $\text{DMSO}-d_6$): δ 8.19 (m, 3H), 8.11 (dd, J = 8.2, 2.3, 3H), 7.86 (t, 3H), 7.54 (q, 1H); ^{13}C NMR (400 MHz, $\text{DMSO}-d_6$): δ 175.08, 147.79, 141.37, 132.45, 130.08, 129.42, 123.58. Anal. calcd (%) for $(\text{C}_{21}\text{H}_{12}\text{N}_6\text{O}_6\text{S}_3)$: C 46.66, H 2.24, N 15.55; found: C 46.64, H 2.06, N 15.49.

4.3.7. Synthesis of 2-methyl-5-nitrobenzene-1-sulfonyl chloride. A 250 ml heavy-walled high-pressure round bottom flask (Ace Glass, Inc.) was charged with 4-nitrotoluene (6.86 g, 50 mmol) and cold chlorosulfonic acid (8.3 ml, 125 mmol) was added dropwise, keeping the temperature below 5 °C. The reaction mixture was heated at 50 °C for 30 min and then the temperature was increased by 20 °C every 30 minutes until it reached 140 °C. Then the reaction mixture was stirred at 140 °C for 22 h. The dark brown solution was cooled to room temperature before opening the flask, poured onto crushed ice (~150 ml) with vigorous stirring for half an hour. The brown precipitate formed was filtered and dried under vacuum (9.5 g, 80.6%). ^1H NMR (500 MHz, $\text{DMSO}-d_6$): δ 8.49 (d, J = 2.85, 1H), 8.05 (dd, J = 8.27, 2.30, 1H), 7.43 (d, J = 8.60, 1H), 2.62 (s, 3H); ^{13}C NMR (400 MHz, $\text{DMSO}-d_6$): δ 147.55, 145.26, 144.61, 132.55, 123.78, 121.38, 20.53.



