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Stabilisation of metastable polymorphs: the case of paracetamol form III†

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The design of a melt synthesis of the first air-stable formulation of the metastable form III of paracetamol is derived from thermo-spectroscopic and thermo-diffraction experiments. Melt crystallisation in the presence of β -1,4-saccharides produces form III selectively and the excipients appear to act as stabilising 'active' templates of the metastable polymorph.

Accessing and stabilising metastable polymorphic forms of active pharmaceutical ingredients remains a major challenge in pharmaceutical product innovation.^{1,2} While systemic approaches to polymorphic diversity in organic solids remain elusive, we report here studies into the polymorphism of paracetamol that indicates a promising role for the 'excipient' matrix to direct and stabilise metastable forms and the importance of *in situ* characterisation of polymorphism for formulation design.

Paracetamol exists in three polymorphic forms; a stable phase (I) and two metastable phases (II, III). Form I packs with a herringbone motif, while forms II and III display layered structures.^{3,4} The structure of form I results in poor compactibility, leading to poor tablet quality.³ However, difficulties in producing suitable quantities of either form II or III have limited research into their physical properties. While form III has been reported as highly unstable in air, even being referred to as 'elusive',⁴ recent reports indicate that surfaces such as glass⁵ and silica⁶ can promote its growth, and careful control of heating and cooling rates during melt crystallisation have also been demonstrated to lead to form III.^{7,8} Additionally, inclusion of hydroxypropyl methylcellulose (HPMC) into the paracetamol melts can

alter the kinetics of the crystallisation of the two metastable phases, which provides an initial insight into the use of polysaccharides as templating agents.^{9,10}

We report here the first air-stable formulations of paracetamol form III based on an apparently active interaction with an excipient matrix derived from saccharides. *In situ* spectroscopic and diffraction studies have offered unique insight into the melt synthesis. Furthermore, this approach is shown to be amenable to scale up.

The addition of molecular additives to alter the kinetics of the crystallisation process to obtain metastable polymorphs has been a frequent topic of study.^{11,12} The inhibition of the stable phase through preferential interactions between the additive and the surface available functional groups of the crystal may lead to slowing of crystal growth,^{13–17} while preferential interaction of the additive with the unstable phase may lead to stabilisation of the phase or alter the balance of the kinetic rates of the two phases.¹⁷ Design or selection of a suitable additive may then be used to control the crystallisation processes and tailor the creation of a desired polymorph.

Melt synthesis of air-stable form III utilises the sugar-based excipients lactose monohydrate and HPMC. The crystallisation process has been followed with simultaneous Raman-DSC and time-resolved X-ray diffraction-DSC experiments. Critically, computational modelling of the interfacial interactions at the predicted major faces of the respective paracetamol crystals and the excipient indicate significant drivers to the stabilisation of form III through hydrogen bonding.

Samples of paracetamol and an excipient (10% w/w) were subjected to heating and cooling cycles with the phase of the crystal established with *in situ* Raman spectroscopy throughout the heating cycles.‡ A standardised heating cycle was implemented: (i) ambient to 180 °C at 10 °C min⁻¹ and held for 2 minutes to melt the paracetamol and excipient mixture; (ii) quench cooling to 0 °C to generate an amorphous phase; (iii) reheating to 180 °C at 10 °C min⁻¹. This last cycle caused three major thermal events: an exotherm at *ca.* 90 °C corresponding to crystallisation [identified as form III with indicative Raman shifts of

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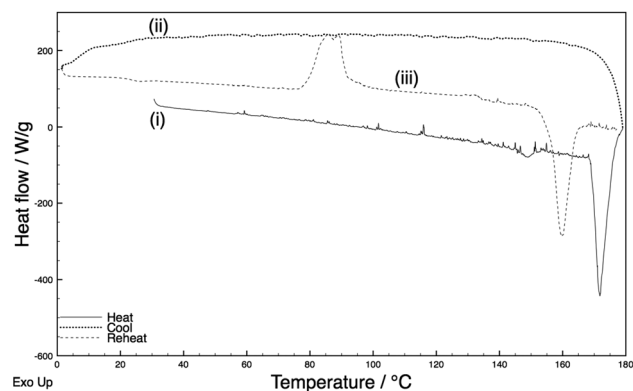
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† Electronic supplementary information (ESI) available: DSC and Raman spectra for all systems. Predicted morphologies for form I and III. See DOI: 10.1039/c6cc05006a





(a)



(b)

Fig. 1 (a) DSC thermograms and (b) Raman spectra recorded at (i) 30 °C before heating [form I], (ii) 30 °C after quench cooling [amorphous] and (iii) 110 °C [form III] after reheating for a 10% lactose containing sample of paracetamol (annotated with key spectral shifts observed from I to III).

phenyl C–C (*ca.* 1620 cm^{-1}), amide II (*ca.* 1560 cm^{-1}), and phenyl ring deformations (*ca.* 1320 and 1230 cm^{-1} bands),^{8,18} a second exotherm at 133 °C (conversion to form II), followed by an endotherm at 160 °C (melt of form II) (Fig. 1). The system with HPMC shows essentially identical behaviour (ESI,† Fig. S1 *cf.* Fig. S2). By comparison, an excipient free sample (ESI,† Fig. S3) melts on heating and cools to an amorphous phase. Upon reheating, recrystallisation to form II occurs at *ca.* 83 °C before the system melts at *ca.* 159 °C – importantly, no exotherm is apparent at 133 °C.

A parallel study of the melt synthesis of form III with lactose was undertaken using hyphenated *in situ* X-ray diffraction/DSC, using similar protocols to those in DSC-Raman studies.§ Due to a difference in the quenching rate, this sample results in a mixture of amorphous/form II paracetamol at room temperature as shown by the presence of Bragg reflections in the sample (Fig. 2, top). Upon reheating of this sample, the DSC shows a broad exotherm with a maximum at *ca.* 85 °C, followed by a second exotherm at 130 °C and an endotherm at just below 160 °C. The latter corresponds to the melt of form II, and XRD patterns collected above this temperature show no diffraction features. The two exotherms correspond to distinct changes in the XRD patterns. The first exotherm coincident with the

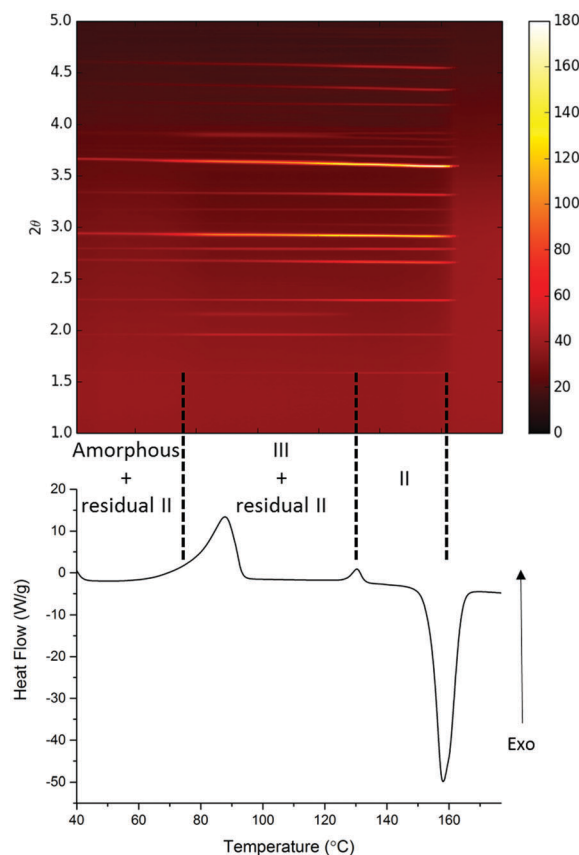


Fig. 2 Simultaneous XRD (top)-DSC (bottom) data collected on a sample of paracetamol/lactose. The sample was heated to 180 °C, held at this temperature for 2 min, and quenched before a second heating cycle performed at 10 °C min^{-1} . The data shown are for this second heat.

crystallisation of form III from the amorphous material present. The second exotherm corresponds to the conversion of form III present into form II, resulting in a pure form II above 130 °C which persists until the mixture melts. Each pattern was analysed using the Rietveld method (selected datasets are given in the ESI,† Fig. S4 and Table S1). The presence of form II throughout the experiment up to the melting point is clear, as is the fact that form III grows in during the first exotherm and is converted to form II at the second. The area under the diffraction pattern from each phase was calculated (ESI,† Fig. S5). It is clear that the amount of form III present is largely constant between the two exotherms, despite the fact that form II crystals are also present. The amount of form II present increases throughout the experiment until the second exotherm, presumably a result of seeding of crystallisation from the amorphous material (in contrast to the pure phase III inferred from the Raman-DSC experiment in this region of the thermogram). Thus the lactose additive is still able to stabilise form III even in the presence of form II crystals.

Interactions between the sugar excipients (modelled as lactose) and paracetamol were investigated computationally,¶ through the minimisation of the intermolecular interactions between crystal surfaces of the differing paracetamol polymorphs and a lactose molecule. For forms I and III, the dominant crystal



software, and XRD data using in-house routines. Rietveld refinements were performed using the GSAS and Topas Academic programmes.²⁰

¶ The molecular modelling was carried out using the in-house program DEX.^{21–23} The crystal structures of paracetamol polymorphs (form I (HXACAN30),²⁴ form II (HXACAN31)²⁴ and form III (HXACAN29),⁴) were extracted from the CSD,²⁵ energy minimised using the downhill simplex routine with in-house code using the potential described by No *et al.*²⁶ using atomic point charges fitted to the electrostatic potential calculated for the isolated molecules in the program orca (mpW2PLYP-D3/def2-TZVPPD).^{27–29} The interaction energy between the crystal surface and a probe molecule was calculated in the same way and minimised using the differential evolution (DE) optimisation algorithm^{30,31} by alteration of the position and orientation of the probe molecule on the selected crystal surface. The parameters were constrained to lie in the range -5 to 5 Å for translations and -180 to 180° for the rotations. The control parameters for the DE algorithm (K , F , G_{\max} , N_p) were set to 0.75, 0.75, 2000, 60. The underlying crystal surface was generated by slicing a $3 \times 3 \times 3$ crystal block through the selected surface removing a 3×3 layer.

|| Morphology predictions were carried out using BFDH and attachment energy methods. Dreiding and COMPASSII force fields were used with the relevant optimised crystal structure. Atomic point charges for the Dreiding force field were taken from the electrostatic fit carried out for the subsequent modelling. Resulting morphologies are presented in ESI.†

** Production of form III at 100 mg scale was achieved by preparing a physical mix of form I paracetamol and lactose monohydrate (10% w/w) on a glass coverslip. This was subjected to a standard heating cycle (ambient to 180°C at $10^\circ\text{C min}^{-1}$ (2 minute isotherm), followed by quench cooling to 0°C and reheating to 110°C at $10^\circ\text{C min}^{-1}$) in the cell in DSC.

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