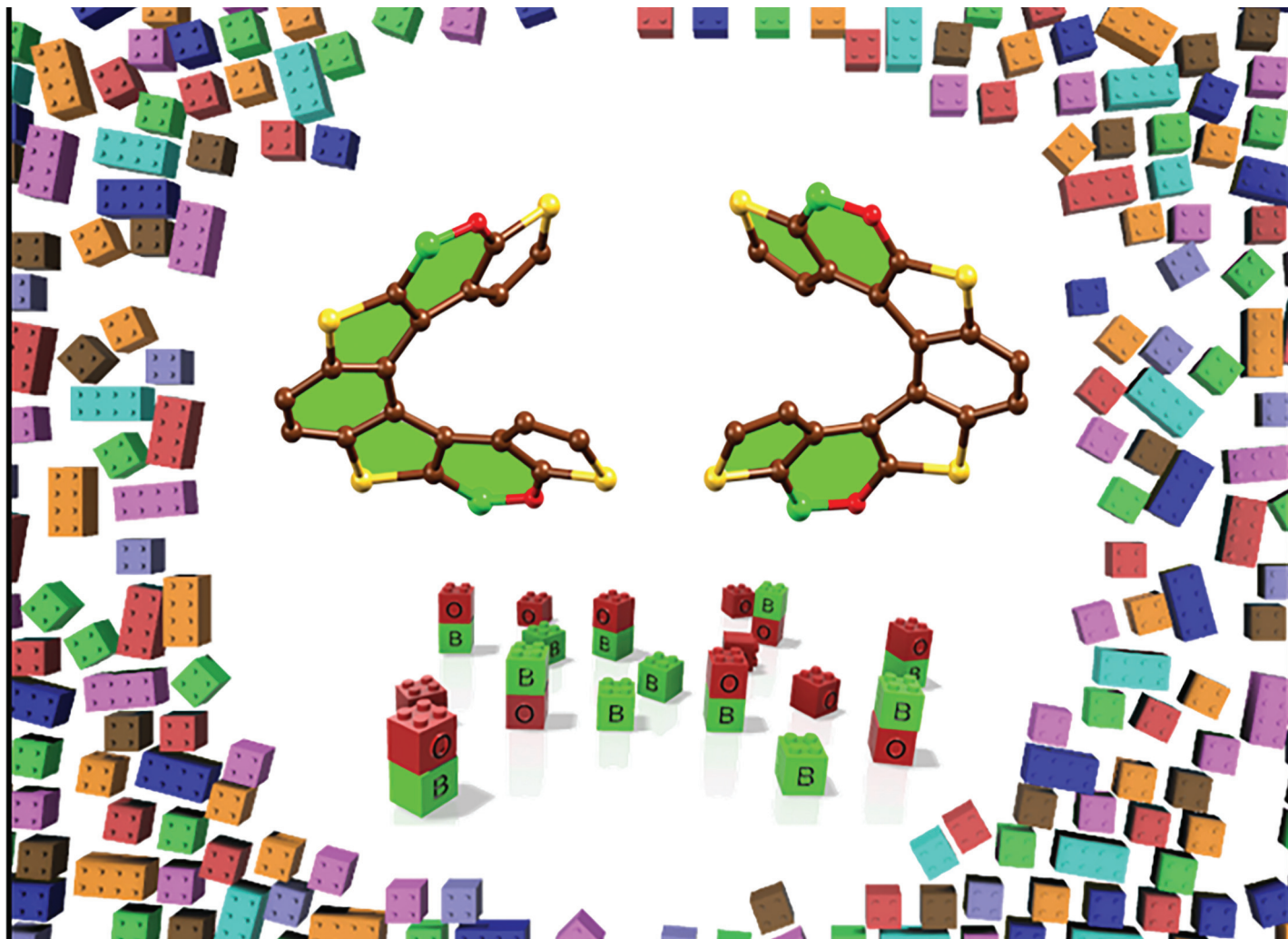




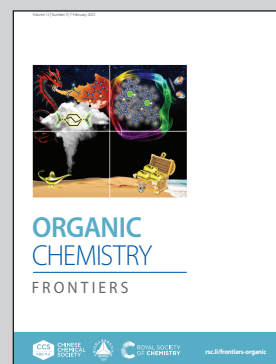
Showcasing research from the laboratories of Dr Menduti, Department of Chemistry, University of Milan, Italy and Prof. Dr Wagner, Institut für Anorganische Chemie, Goethe-Universität Frankfurt, Germany.

(BO)₂-doped tetrathia[7]helicenes: synthesis and property-change induced by “BO bond inversion”

Make a wish: revert BO bonds and get improved optoelectronic properties. Synthesis of a new (BO)₂-doped tetrathia[7]helicene and comparison with its isomer featuring reverted BO bonds.

Image designed and realized by Dr Simone Manetto, Dr Sara Grecchi and Dr Luigi Menduti

As featured in:



See Luigi Menduti, Matthias Wagner *et al.*, *Org. Chem. Front.*, 2025, 12, 725.

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Cite this: *Org. Chem. Front.*, 2025, **12**, 725**(BO)₂-doped tetrathia[7]helicenes: synthesis and property-change induced by “BO bond inversion”†**Luigi Menduti,^{id} *^{a,b} Clara Baldoli,^{id} ^c Simone Manetto,^d Claudio Villani,^{id} ^d Marta Penconi,^{e,f} Sara Grecchi,^a Serena Arnaboldi,^{id} ^a Giuseppe Mazzeo,^g Giovanna Longhi,^{id} ^g Michael Bolte,^b Alexander Virovets,^{id} ^b Hans-Wolfram Lerner,^{id} ^b Matthias Wagner^{id} *^b and Emanuela Licandro^{id} ^a

Helical distortion of polyaromatic hydrocarbons gives rise to a special class of π -conjugated systems, namely helicenes. Owing to their configurational stability and easily tunable optoelectronic properties (*via* heteroatom-doping), helicenes have recently come to the fore as building blocks for applications in materials science (CP-OLEDs, chiroptical switches). In this context, boron-doped helicenes are particularly promising. Herein, we report the synthesis of the new (BO)₂-doped tetrathia[7]helicene **2**, derived by the formal inversion of (Mes)B–O moieties in the (previously reported) isomer **1**. Theoretical characterization of **2**, and comparison with **1**, revealed that the inversion of the BO vectors promotes the extension of the LUMO *via* the central thiophene–benzene–thiophene fragment (and not *via* the terminal thiophene rings, as in **1**), resulting in a considerable lowering of the LUMO energy ($E_{\text{LUMO}}(\mathbf{2}) = -2.22$ eV vs. $E_{\text{LUMO}}(\mathbf{1}) = -1.65$ eV). Spectroscopic studies revealed that the “BO bond inversion” also contributes to the narrowing of the energy gap ($E_{\text{g}}^{\text{opt}}(\mathbf{2}) = 2.90$ eV vs. $E_{\text{g}}^{\text{opt}}(\mathbf{1}) = 3.16$ eV) and causes a significant red-shift of the absorption/emission bands (≈ 40 nm). Interestingly, besides low fluorescence quantum yield ($\Phi_{\text{PL}}(\mathbf{2}) = 7\%$), **2** shows detectable circularly polarized luminescence ($g_{\text{lum}} = 0.8 \times 10^{-3}$) and pronounced phosphorescence at low temperature (77 K). (*P*)-/(*M*)-enantiomers of **2** were successfully separated by CSP-UHPLC and proved to be stable ($\Delta G_{\text{enanti}}^{\ddagger} = 29.4 \pm 0.1$ kcal mol⁻¹ at 353 K). Racemization studies combined with theoretical calculations confirmed that BO-doping is an extremely perturbative tool for tuning the mechanical rigidity of tetrathia[7]helicenes ($\Delta G_{\text{enanti}}^{\ddagger}(\mathbf{2})$ is 10 kcal mol⁻¹ lower than $\Delta G_{\text{enanti}}^{\ddagger}(\mathbf{7TH})$).

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Introduction

Helicenes, which are non-planar, screw-shaped chiral structures composed of *ortho*-fused aromatic or heteroaromatic rings, represent a unique class of polycyclic aromatic hydrocarbons (PAHs).¹ Configurationally stable helicenes show peculiar chiroptical and optoelectronic features suitable for their application in the field of materials science.^{2,3}

Recently, the so-called heteroatom doping, based on the replacement of selected carbon atoms by other p-block elements,⁴ opened new perspectives in helicene chemistry, allowing the modulation of their electronic structures and optoelectronic properties. The most commonly used doping elements to date are Si, N, P, O and S, but recently several examples of B-atom introduction have been described.⁵ Only a limited number of purely B(sp²)-doped helicenes and chiral PAHs incorporating helicene subunits have been reported so far,⁶ while in most of the examples, C=C double bonds are replaced by B(sp²)-N/B(sp²)-O^{7,8} or B(sp³)-N bonds.⁹ Our groups have recently contributed to this field by establishing the synthesis and investigating the optoelectronic and chiroptical

^aDipartimento di Chimica, Università degli Studi di Milano, Via C. Golgi 19, 20133 Milano, Italy. E-mail: luigi.menduti@unimi.it^bInstitut für Anorganische Chemie, Goethe-Universität Frankfurt, Max-von-Laue-Strasse 7, 60438 Frankfurt am Main, Germany. E-mail: matthias.wagner@chemie.uni-frankfurt.de^cCNR Istituto di Scienze e Tecnologie Chimiche Giulio Natta, Via C. Golgi 19, 20133 Milano, Italy^dDipartimento di Chimica e Tecnologie del Farmaco, Sapienza Università di Roma, P.le A. Moro 5, 00185 Roma, Italy^eCNR Istituto di Scienze e Tecnologie Chimiche Giulio Natta, Via G. Fantoli 16/15, 20133 Milano, Italy^fSmartMatLab Center, via C. Golgi 19, 20133 Milano, Italy^gDipartimento di Medicina Molecolare e Traslazionale, Università di Brescia, Viale Europa 11, 25123 Brescia, Italy†Electronic supplementary information (ESI) available. CCDC 2384683 (for **2**), 2384684 (for **4**), 2384685 (for **5**), 2384686 (for **6**), 2384687 (for **7**), 2384688 (for **8**), 2384689 (for **9**), 2384690 (for **10**) and 2384691 (for **11**). For ESI and crystallographic data in CIF or other electronic format see DOI: <https://doi.org/10.1039/d4qo01897d>

tical properties of the first doubly BO-doped tetrathia[7]helicene **1**¹⁰ (Fig. 1).

Although **1** showed poor optical performance ($\Phi_{\text{PL}} = 6\%$, $E_{\text{g}}^{\text{opt}} = 3.16$ eV), our studies have demonstrated that the incorporation of two BO fragments into the tetrathiahelicene (**7TH**; Fig. 1) scaffold provides a powerful tool to manipulate the mechanical rigidity, the chiroptical properties, and the electron distribution across the helix.

With the aim of thoroughly investigating the effect of BO doping of the **7TH** framework to obtain helicenes with improved photophysical features, we designed the new (BO)₂-helicene **2**, an isomer of **1**, derived by the formal inversion of the BO bonds in the two oxaborine rings (Fig. 1).

(BO)₂-helicene **2** was tailored based on our previous results, which showed that the LUMO of **1** (highlighted in green in Fig. 1) is mainly localized on the B-substituted terminal thiophene rings,¹⁰ while the HOMO, as in **7TH**,¹¹ is spread over the entire helix (for DFT calculations, see Fig. 6 and Fig. S49†).

These results suggested that the position of the B atoms could be crucial for the extension of the LUMO across the helix, with only minor effects on the HOMO distribution/energy (as in the case of boron-doped planar polyaromatics).^{4c} The inverted positions of the BO bonds in **2** were therefore expected to promote the extension of the LUMO *via* the central thiophene–benzene–thiophene fragment rather than *via* the terminal thiophene rings only (Fig. 1). This could provide a helicene with an extended and energetically lowered LUMO, a narrowed HOMO–LUMO energy gap, and improved optoelectronic performance.

Herein, we report the synthesis of the (BO)₂-helicene **2**, along with its optoelectronic and chiroptical properties and a comparison with those of the previously reported (BO)₂-helicene isomer **1**.¹⁰ Moreover, we describe attempts to synthesize the (BO)-doped helicene **3** (Scheme 2a), formally derived by

the replacement of the central C=C bond in **7TH** with a (Mes) B–O fragment. Although the target BO-helicene **3** was not obtained, the investigation of two different synthetic approaches provided important information for the design of new derivatives and allowed us to isolate the penta-fused intermediate **11** (see Results and discussion).

Results and discussion

Synthesis, NMR spectra, and X-ray characterization of **2**

The (BO)₂-helicene **2** was prepared in four steps, starting from 4-bromothiophene-2-carbaldehyde¹² (Scheme 1), which was converted into the alkene **4** in 89% yield (*trans* : *cis* = 84 : 16) *via* a McMurry coupling. The *trans* isomer was isolated in 82% yield by washing the crude product with Et₂O.

The next step was performed on *trans*-**4** and involved Suzuki or Stille coupling with the corresponding partners 2-(2-methoxythiophen-3-yl)-1,3,2-dioxaborolane **S1** or (2-methoxythiophen-3-yl)trimethylstannane **S2** and furnished the alkene *trans*-**5** in 60% and 42% yields, respectively. Although Suzuki coupling gave a better yield, our preferred route for *trans*-**5** remains Stille coupling, due to the easier preparation and purification of the stannane **S2**. The subsequent two-fold borylation of *trans*-**5** was performed by demethylative borylation with BCl₃/[*n*Bu₄N]I/Et₃N¹³ in chlorobenzene at 135 °C, followed by the addition of MesMgBr. In this way, the alkene *trans*-**6** was obtained in 31% yield after purification.

The structures and the *trans*-configurations of the double bonds in intermediates **4**, **5**, and **6** were confirmed by X-ray

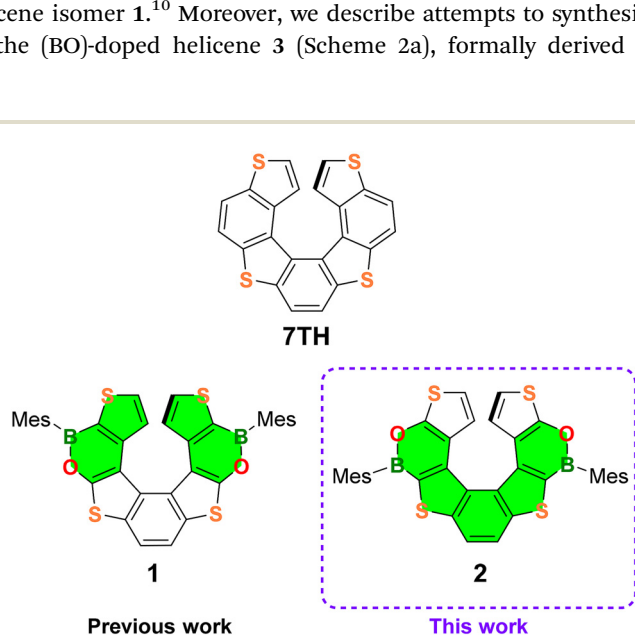
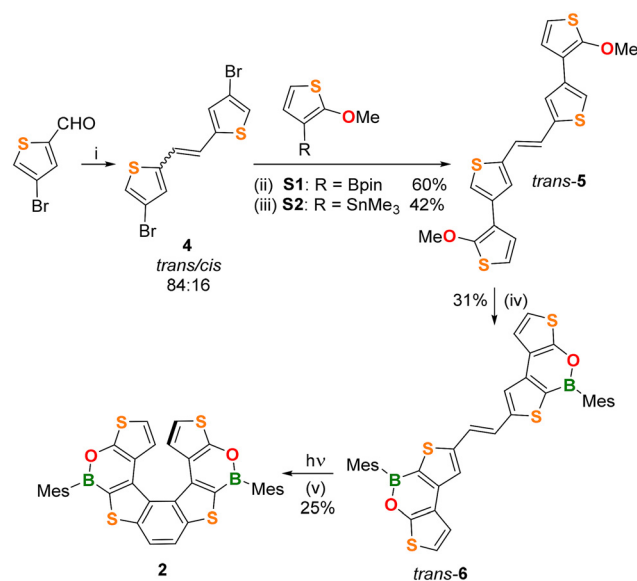


Fig. 1 Structures of the parent tetrathia[7]helicene (**7TH**), the previously reported **1** and the newly developed (BO)₂-helicene **2**. Mes = 2,4,6-trimethylphenyl (mesityl). The LUMO extension in the structures of **1** and **2** is highlighted in green.



Scheme 1 (i) 2.6 equiv. Zn, 1.2 equiv. TiCl₄, THF, 0–15 °C; (ii) 2.4 equiv. **S1** (R = Bpin), Pd(PPh₃)₄ 10 mol%, 4 equiv. NaOH (3 M), THF, 80 °C, 15 h; (iii) 2.1 equiv. **S2** (R = SnMe₃), Pd(Pt-Bu₃)₂ 10 mol%, toluene, 80 °C, 17 h; (iv) 2.2 equiv. BCl₃, 2 equiv. Et₃N, 2.4 equiv. [*n*Bu₄N]I, chlorobenzene, 135 °C, then THF, 6 equiv. MesMgBr; and (v) 405 nm LED, C₆D₆, r.t., 1.5 h. Mes = 2,4,6-trimethylphenyl (mesityl).



crystallography (Fig. S36–S38†). The last reaction to obtain the helicene structure was a Mallory photocyclization¹⁴ involving as the first step the isomerization of *trans*-6 to *cis*-6, which afterwards cyclized to give 2.

We first investigated the *trans* → *cis* isomerization. To this end, an NMR tube charged with a non-degassed C₆D₆ solution of *trans*-6 was irradiated with a 405 nm LED lamp and the reaction progress was monitored by ¹H NMR. Surprisingly, this test showed that not only does *trans*-6 smoothly isomerize to *cis*-6 but also that *cis*-6 undergoes photocyclization to give 2, even in the absence of an added oxidant (e.g. I₂; differently from what was required for the synthesis of 1).¹⁰

In light of these results, we then performed the synthesis of 2 under the same conditions: four NMR tubes charged with *trans*-6 and C₆D₆ (10 mg mL^{−1}) were simultaneously irradiated with the 405 nm LED lamp until almost complete disappearance of the alkene precursor (1.5 h; NMR-spectroscopic control). The NMR solutions were then combined and, after chromatographic purification, (BO)₂-helicene 2 was isolated in 25% yield.

NMR characterization of the new (BO)₂-helicene 2 was done along with a comparison with the chemical shifts of the isomeric helicene 1 (Fig. 2a and b). The inversion of the BO-bonds in the oxaborine rings causes a substantial upfield shift of the diagnostic ¹H and ¹³C signals of the terminal thiophene rings.

For instance, in 2, the protons at the α and β positions of the terminal thiophene rings substituted with an electron-donating O atom (O-thiophene) resonate at 6.93 and 6.52 ppm, respectively, whereas in the corresponding isomer 1, they are significantly deshielded due to the electron-withdrawing effect of the B atom (B-thiophene) and resonate at 7.37 and 7.19 ppm (cf. parent 7TH: 6.83, 6.67 ppm). The same trend is observed for the ¹³C shifts: 2 C(α) and C(β) resonate at 123.0 and 112.4 ppm, respectively, but are significantly deshielded in 1 (C(α) = 135.6 ppm, C(β) = 126.5 ppm).

A similar behavior is observed for the corresponding alkene precursors *trans*-6 and *trans*-6a for both the H(α)/C(α) and H(β)/C(β) resonances (Fig. 2c and d). Due to the proximity of the electron-withdrawing B atoms, the protons/carbons of the central benzene ring in 2 are deshielded (8.03/122.6 ppm) with respect to those at the same positions in 1 (7.81/119.0 ppm).

Interestingly, the α and β protons of the terminal thiophene rings of 2 are more shielded than those of its planar precursor *trans*-6 (6.93/6.52 ppm vs. 7.32/6.96 ppm). This effect can be attributed to the helical 3D-structure, in which each terminal thiophene ring falls into the shielding region of the corresponding ring in the opposite helicene wing. (BO)₂-helicene 1 and its precursor *trans*-6a show the same trend.

The crystal structure of 2 was determined by single-crystal X-ray crystallography using the high-flux DESY Petra III synchrotron source. 2 crystallizes as a racemic mixture of the (*P*)- and (*M*)-enantiomers in the monoclinic crystal system (Fig. 3a). 2 shows a dihedral angle between the two terminal thiophene rings of 58.36(6)° (Fig. 3b), significantly larger than those of 1 (50.26(9)°)¹⁰ and 7TH (48.6(1)°),¹⁵ while the associated intramolecular centroid...centroid distances have comparable values (4.299(4) Å for 2; 4.153(6) Å for 1).

As for 1, the central benzene ring of 2 is twisted (Fig. 3c), as evidenced by the torsion angle of −10.5° (9.7° for 1) between the (H)CC(H) vector and the opposite CC bond.

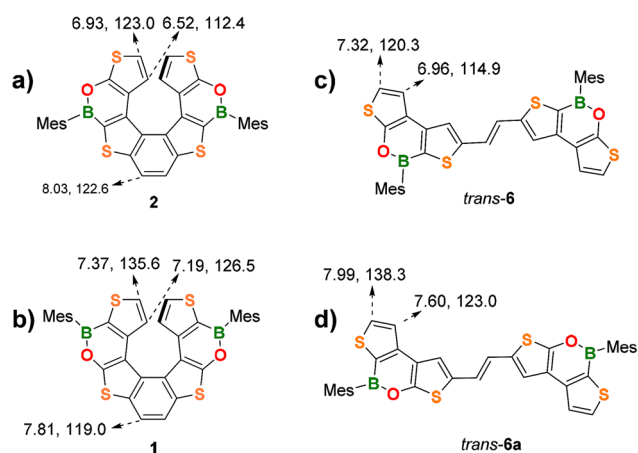


Fig. 2 Comparison of diagnostic ¹H and ¹³C NMR resonances (CDCl₃) of the two (BO)₂-helicenes (a) 2 and (b) 1 and of the *trans*-alkenes (c) 6 and (d) 6a.

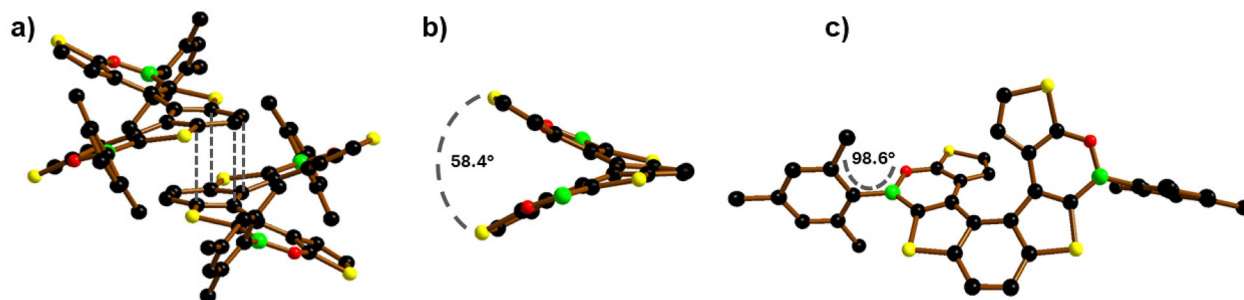


Fig. 3 (a) π -stacking interaction between a (*P*)/(*M*)-pair of 2 with H atoms omitted for clarity;¹⁸ grey dashed lines indicate the shortest intermolecular C...C contacts. (b) Dihedral angle between the terminal thiophene rings of (*P*)-2; Mes rings are omitted for clarity. (c) Front view of (*P*)-2. B: green, C: black, O: red, and S: yellow spheres.



Contrary to **1** and similar to most of the Mes-protected B-doped polycyclic aromatic hydrocarbons (PAHs),¹⁶ the Mes substituents of **2** are nearly orthogonal to the oxaborine rings (Fig. 3c), as evidenced by the Mes//O–B–C dihedral angle (98.6(2)° for **2**; 59.0(5)° and 58.6(6)° for **1**). In both **2** and **1**, the *ortho*-CH₃ groups of the Mes substituents are positioned at approximately 4 Å to the CH(β) atoms of the opposite thiophene rings, thus excluding that the different Mes orientations in the two (BO)₂-helicenes could be caused by steric factors (*e.g.* H/H clashes). The reason for the two different orientations apparently lies in the crystal packing of **1** and **2**.

The B–O bond lengths of **2** (1.393(3) Å) are slightly shorter than those of **1** (1.402(4) and 1.401(5) Å) but longer than those of the reference compound 9-mesityl-10,9-oxaboraphenanthrene (av. B–O = 1.378 Å; av. Mes//O–B–C = 104.1(3) and 84.8(3)° for two crystallographically independent molecules in the asymmetric unit).¹⁷ In the crystal, alternating molecules of (*P*)- and (*M*)-**2** form infinite rods by means of π -stacking interactions between the central benzene rings (Fig. 3a; similar to **7TH**).¹⁵ The corresponding intramolecular centroid...centroid distances are 3.666(4) Å, and the shortest intermolecular C...C contacts are 3.552(4) and 3.666(4) Å (grey lines in Fig. 3a). As for **1**, the shortest *intermolecular* S...S distances are smaller than the *intramolecular* ones (3.665(5) and 6.140(5) Å, respectively, for **2**; 3.9911(14) and 5.8156(2) Å, respectively, for **1**).

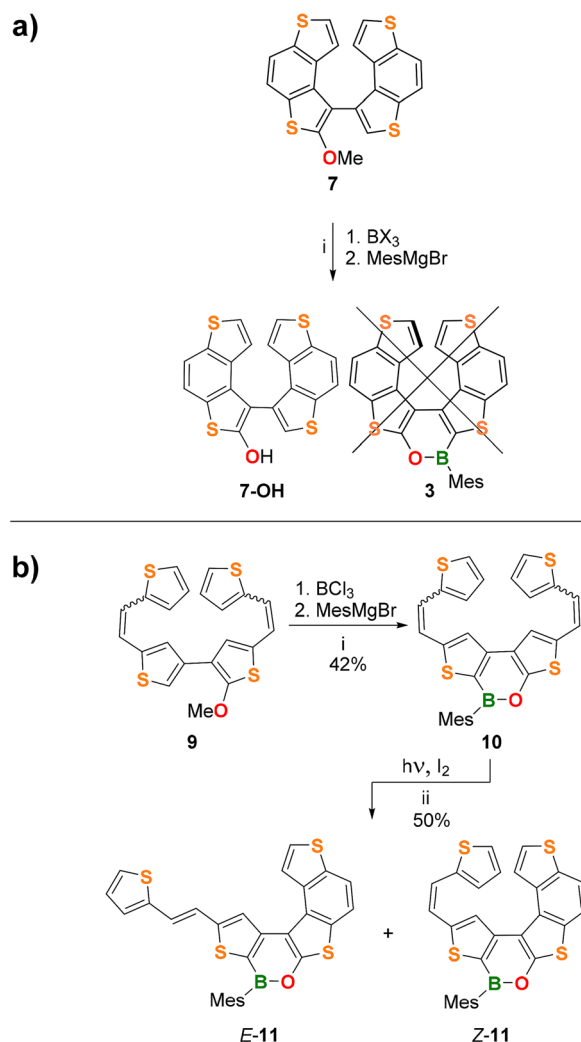
Synthesis attempts for **3**

Next, we attempted the synthesis of the (BO)-helicene **3** (Scheme 2a), formally derived by the replacement of the central C=C bond of **7TH** with a (Mes)B–O fragment.

To this end, we pursued two different synthetic strategies (Scheme 2a and b). The first approach involved the late-stage borylation of **7** (Scheme 2a; see the ESI† for details of its synthesis). In the second approach, the oxaborine ring was constructed first (*via* borylative cyclization **9** → **10**; see the ESI† for the synthesis of **9**), and the obtained bis-olefinated oxaborine **10** was then subjected to Mallory photocyclization (Scheme 2b).

Unfortunately, the borylative cyclization **7** → **3** failed for BCl₃ and BBr₃; only the hydroxy derivative **7-OH** was isolated besides a small amount of unidentified products (Scheme 2a; see the ESI for details, Table S1†). The crystal structure of the precursor compound **7** (Fig. S39†) indicates that the final B–C bond-forming step must overcome adverse steric factors and is therefore unfavorable.

In contrast, borylation of the more flexible bis-olefin **9** (see the ESI† for details of its synthesis) gave the bis-olefinated oxaborine **10** in a satisfactory yield (42%) as a mixture of four *E/Z* isomers after chromatographic purification. Treatment of the mixture with *n*-pentane allowed the isolation of crystals of *E,E*-**10**, the configuration of which was confirmed by single-crystal X-ray diffraction (Fig. 4a and Fig. S42†). Photocyclization of **10** was carried out on the mixture of isomers in benzene solution by irradiation with a 385 nm LED lamp for 3.5 h at room temperature in the presence of 2 equiv. of I₂ and excess propylene oxide as an HI scavenger (Scheme 2b).¹⁹ In this way, the penta-



Scheme 2 (a) (i) 1.5 equiv. BCl₃ or BBr₃, 1 equiv. Et₃N, 1.2 equiv. [nBu₄N]I, chlorobenzene, 135 °C, then THF, 3 equiv. MesMgBr; (b) (i) 1.1 equiv. BCl₃, 1 equiv. Et₃N, 1.2 equiv. [nBu₄N]I, chlorobenzene, 135 °C, then THF, 3 equiv. MesMgBr; (ii) 385 nm LED, 2.1 equiv. I₂, 200 equiv. propylene oxide, C₆H₆, r.t., 3.5 h.

fused product *E/Z*-**11** was obtained in 50% yield as an almost equimolar mixture of the *E* and *Z*-isomers. We found no indication for the formation of **3**. ¹H NMR analysis of *E/Z*-**11** allowed the assignment of most of the proton resonances of both structures (see the ESI†). An X-ray diffraction study of single crystals of *E*-**11** confirmed that the photocyclization had occurred at the “O-thiophene” double bond (Fig. 4b and Fig. S43†).

Spectroscopic, electrochemical, theoretical, and chiroptical characterization of **2**

The electrochemical, spectroscopic, and chiroptical properties of the new (BO)₂-helicene **2** were investigated by cyclic voltammetry, UV/Vis and fluorescence spectroscopy, and chiral HPLC analysis and compared with the corresponding properties of isomer **1**.



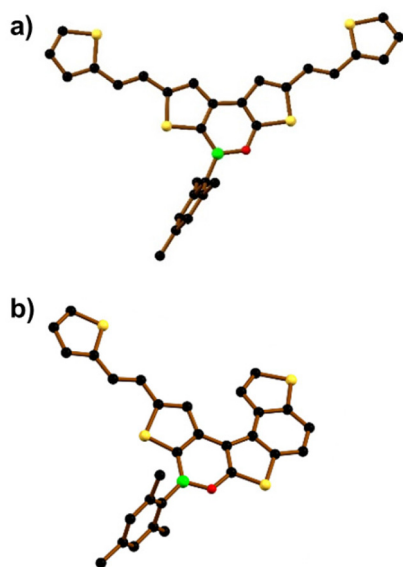


Fig. 4 Solid-state structures of (a) *E,E*-10 and (b) *E*-11. Hydrogen atoms are omitted for clarity. B: green, C: black, O: red, and S: yellow spheres.

Cyclic voltammetry measurements on **2** (THF, $[n\text{Bu}_4\text{N}][\text{PF}_6]$; referenced against the FcH/FcH^+ couple) showed two irreversible reduction processes with peak potentials of $E_{\text{pc}}^{\text{I}} = -1.66$ V and $E_{\text{pc}}^{\text{II}} = -2.36$ V in the cathodic scan and an irreversible oxidation peak in the anodic scan ($E_{\text{pa}} = 0.82$ V; Fig. S28, S29, and Table S3†). Helicene **2** is therefore easier to reduce than its isomer **1** ($E_{\text{pc}}^{\text{I}}(\mathbf{1}) = -2.83$ V). Consequently, the electrochemically determined LUMO energy of **2** ($E_{\text{LUMO}} = -3.14$ eV) is also significantly lower than that of **1** ($E_{\text{LUMO}} = -1.97$ eV), thus confirming that the B–O-bond inversion contributes markedly to the stabilization of the LUMO.

The molar absorptivity and emission spectra of **2** and **1** (*c*-hexane) are shown in Fig. 5a; the data are summarized in Table 1. Both compounds show an absorption band below 300 nm and a second, more structured band at lower energy, which is red-shifted in **2** and characterized by a lower absorption coefficient relative to **1**. Consequently, the energy gap between the frontier orbitals is narrowed ($E_{\text{g}}^{\text{opt}}(\mathbf{2}) = 2.90$ eV vs. $E_{\text{g}}^{\text{opt}}(\mathbf{1}) = 3.16$ eV). The emission spectrum of **2** shows a band with a well-resolved vibrational fine structure and a maximum at 446 nm which, similar to the absorption band, is bathochromically shifted with respect to the spectrum of **1**. The photoluminescence quantum efficiency of **2** remains unchanged, $\Phi_{\text{PL}} = 7\%$ (*c*-hexane; cf. **1**: $\lambda_{\text{em}} = 411$ nm, $\Phi_{\text{PL}} = 6\%$; **7TH**: $\lambda_{\text{em}} \approx 405$ nm, $\Phi_{\text{PL}} = 6\%$),²⁰ and the lifetime of the excited state is $\tau = 0.82$ ns.

To gain insight into the origin of the low quantum efficiency of **2**, low-temperature photoluminescence spectra (77 K) were measured. To avoid degradation of **2** in 2-MeTHF (which usually contains peroxides), measurements were carried out in a frozen toluene matrix; for comparison, the corresponding spectra of the helicene precursor *trans*-**6** and the mesityl dithienooxaborine (**MesDTOB**) subunit have also been included in Fig. 5b.

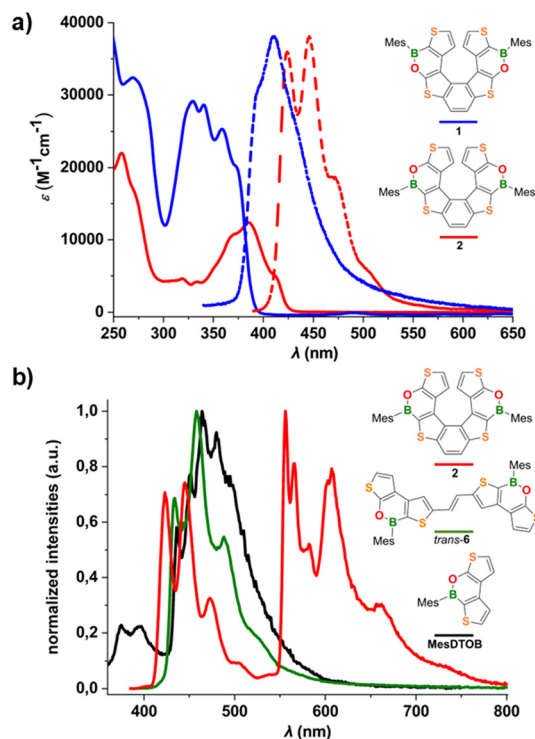


Fig. 5 (a) Molar absorptivity (solid lines) and emission (dashed lines) spectra of **2** (red) and **1** (blue) in *c*-hexane. (b) Low-temperature emission spectra of **2** (red), *trans*-**6** (green) and **MesDTOB** (black) in a frozen toluene matrix (77 K).

At 77 K, the fluorescence emission of **2** between 400 and 530 nm overlaps with the room temperature photoluminescence in *c*-hexane and the corresponding average lifetime of the singlet excited state is $\tau = 1.09$ ns. Interestingly, at low temperature, we observed long-wavelength photoluminescence, between 530 and 800 nm, not present at room temperature. This emission corresponds to phosphorescence from the triplet excited state with a typical lifetime of $\tau = 63.3$ ms. Both fluorescence and phosphorescence signals are highly structured and with spacings of about 1300 and 300–700 cm^{-1} , respectively, as observed for similar thiahelicene and carbohelicene structures reported in the literature.^{11,21} In particular, the low-frequency vibronic progression of the phosphorescence band is associated with breathing vibrational modes, so-called “helix breathing”.^{11,22} The observed phosphorescence of **2** results from the population of the triplet state by intersystem crossing from the singlet excited state and the subsequent relaxation of the triplet state by radiative processes. In this type of structure, these spin-forbidden singlet–triplet transitions become partially allowed by the enhanced spin–orbit coupling arising from two different contributions: the helically twisted molecular scaffold and the presence of heavy sulfur heteroatoms along the conjugated system. Notably, phosphorescence emission was also observed in our previous studies on α -dimesitylboryl benzodithiophene systems²³ (resembling the



Table 1 Photophysical data of **2**, *trans*-**6** and **MesDIOB** in *c*-hexane at room temperature and in toluene at 77 K

c-Hexane, r.t. ^a			Toluene, 77 K								
	λ_{abs} [nm] (ϵ [M ⁻¹ cm ⁻¹])	λ_{em} [nm]	τ [ns]	Φ_{PL}	k_{r} [10 ⁷]	k_{nr} [10 ⁹]	$\lambda_{\text{em,FL}}$ [nm]	τ_{FL} [ns]	$\lambda_{\text{em,PH}}$ [nm]	τ_{PH} [ms]	E_{T} [eV]
2	258 (21 900), 271 (sh), 319 (4600), 334 (4200), 368 (10 600), 385 (12 400), 413 (5000)	424, 446, 472, 506 (sh)	0.82	7	8.5	1.13	423, 445, 473, 505 (sh)	1.09	556, 566, 582, 603, 607, 660, 735 (sh)	63.3	2.23
	270 (26 500), 355 (30 700), 370 (30 000), 393 (sh), 414 (sh)	429, 452, 489 (sh)	0.31	6	19.3	3.03	433, 457, 489, 530 (sh)	0.48	—	—	—
MesDIOB	269 (32 400), 330 (29 100), 341 (28 600), 359 (25 300), 374 (sh)	366, 380	0.30	6	20.0	3.13	375, 395	0.70	436, 452, 464, 480	104	2.84

^a Quantum yields and lifetimes determined on N₂-saturated solutions. [FL] = fluorescence. [PH] = phosphorescence. sh = shoulder. Rate constants k_{r} and k_{nr} are calculated using the equations $k_{\text{r}} = \Phi_{\text{PL}}/\tau$ and $k_{\text{nr}} = (1 - \Phi_{\text{PL}})/\tau$. Absorption and emission spectra of *trans*-**6** at r.t. are given in the ESI.[†] For absorption and emission spectra of **MesDIOB** at r.t., see ref. 10.

B-thiophene–benzene–thiophene–B fragment of **2**) and in tetra-thia[7]helicene derivatives.¹¹

Under the same conditions, no phosphorescence was detected for the more flexible precursor *trans*-**6**, likely due to rotational deactivation processes. Interestingly, intense phosphorescence was also observed at 77 K for the dithieno-xaborine model compound **MesDIOB**, albeit at much shorter wavelengths than in the case of **2** (Fig. 5b). Overall, the photophysical characterization of **2** showed that the introduction of two BO moieties into the parent **7TH** results in a red-shift of the phosphorescence emission and an enhanced phosphorescence/fluorescence ratio (for low-temperature emission measurements on **7TH** (in 2-MeTHF), see ref. 11).

The HOMOs and LUMOs of the (BO)₂-helicenes **1** and **2** have been computed at the M06/6-311+G(d,p) level of theory (Fig. 6 and Fig. S49[†]). A comparison with the HOMO/LUMO of the parent **7TH** confirms that the introduction of the BO moieties substantially influences the LUMO delocalization and that the Mes rings make no contribution to the frontier orbitals, in agreement with reports on (BN)-doped carbocyclic helices.^{7e} The LUMO of **2** is ≈ 0.6 eV lower than that of **1** (−2.22 eV vs. −1.65 eV), in line with the trend observed for experimental values (−3.14 eV vs. −1.97 eV). As expected, the LUMO of **2** is mainly delocalized over the thiophene–benzene–thiophene central fragment rather than over the terminal thiophene rings (as in the case of **1**). Conversely, the HOMOs of **1**, **2**, and **7TH** present only minor differences, both in terms of energy and degree of delocalization.

Preliminary racemization tests showed that the enantiomers of **2** are stereochemically stable at room temperature. The (*P*)-/(*M*)-enantiomers of **2** were therefore separated by chiral stationary phase ultra-high performance liquid chromatography (CSP-UHPLC) on an (*S,S*)-Whelk-O 1 column, using *n*-hexane/CH₂Cl₂ (95 : 5) as the eluent (Fig. 7a).²⁴ Later, the analytical HPLC resolution was scaled up to the semipreparative level using the same eluent and a 10 mm internal diameter column. Racemization studies were conducted on the second eluted enantiomer showing an ee = 97 % ($R_{\text{t}} \approx 7$ min; Fig. 7a, red trace). The collected fractions containing the second eluted enantiomer were combined and evaporated at room temperature. The solid residue was dissolved in decalin and subjected to thermal racemization at 80 °C.

The changes in the enantiomeric excess of the sample were monitored by CSP-UHPLC (Fig. 7b). The racemization process was monitored by the decrease in the enantiomeric excess of the sample. Significant interconversion of the two enantiomers was visible after 12 min and the elution peak of the first enantiomer ($R_{\text{t}} \approx 6$ min) began to emerge (Fig. 7b).

After approximately 8 h, the ee had dropped to 65% (Fig. 7b, purple trace). Data collected over this time (Fig. S46[†]) were used to determine the racemization constant $k_{\text{rac}} = 1.48 \times 10^{-5} \text{ s}^{-1}$.^{24,25} The enantiomerization constant was calculated using the relation $k_{\text{enant}} = 0.5 \times k_{\text{rac}}$. The free energies of activation, $\Delta G_{\text{rac}}^{\ddagger} = 28.6 \pm 0.1 \text{ kcal mol}^{-1}$ (at 353 K) and $\Delta G_{\text{enant}}^{\ddagger} = 29.4 \pm 0.1 \text{ kcal mol}^{-1}$ (at 353 K), were calculated by using the



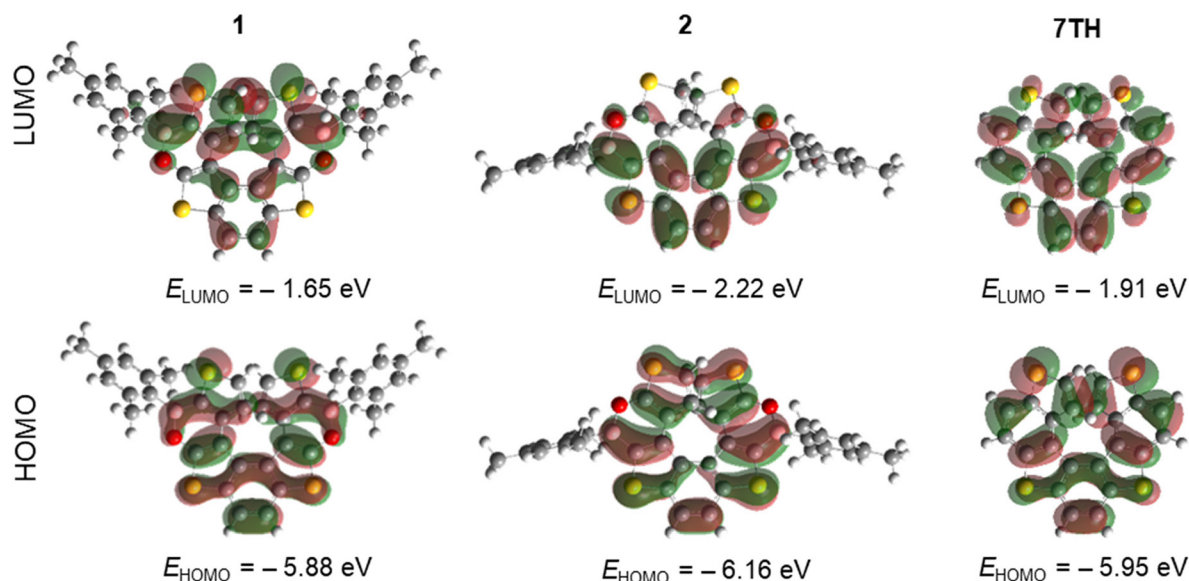


Fig. 6 Isovalue surfaces for the LUMO and HOMO of **1**, **2**, and **7TH**, calculated at the M06/6-311+G(d,p) level.

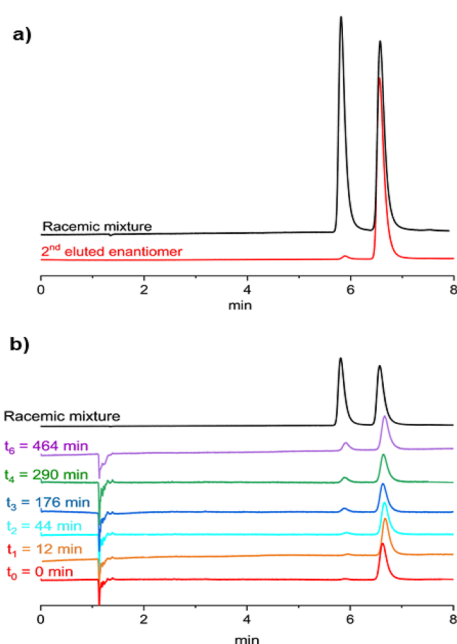


Fig. 7 (a) CSP-UHPLC resolution of **2** (stationary phase: (*S,S*)-Whelk-O **1**, mobile phase: *n*-hexane/CH₂Cl₂ (95 : 5), flow rate = 1.0 mL min⁻¹, 25 °C). (b) Thermal racemization study conducted on the second eluted enantiomer in decalin at 80 °C; solvent peak: *R*_t ≈ 1 min.

Eyring equation and setting a transmission coefficient equal to 0.5 (see section 8.1 of the ESI†).²⁶

Both $\Delta G_{\text{rac}}^{\ddagger}$ and $\Delta G_{\text{enant}}^{\ddagger}$ of **2** are close to the corresponding values of **1** ($\Delta G_{\text{rac}}^{\ddagger}$ (**1**) = 26.9 ± 0.1 kcal mol⁻¹, $\Delta G_{\text{enant}}^{\ddagger}$ (**1**) = 27.4 ± 0.1 kcal mol⁻¹ at 353 K), which reveals that the inversion of the B–O bonds does not reduce the configurational stability of the enantiomers.

We then evaluated the energy barrier through DFT calculations, using the same functional and basis set as for compound **1**.¹⁰ Optimization of the transition state has been performed at the PBE1-PBE/TZVP level with empirical dispersion corrections,²⁷ obtaining one imaginary frequency. This gave a barrier of 36 kcal mol⁻¹, significantly higher (≈ 8 kcal mol⁻¹) than the one obtained from the experimental data ($\Delta G_{\text{enant}}^{\ddagger}$ (**1**) = 27.4 ± 0.1 kcal mol⁻¹). The analysis of a relaxed scan to evaluate the racemization barrier of model compounds without Mes groups (**1*** and **2***; Fig. S50†) allows increasing the basis set level and comparing **1*** and **2*** with **carbo[6]helicene** and **7TH** to assess the influence of the heteroatoms.²⁸ However, similar values for the enantiomerization barrier were obtained (≈ 36 kcal mol⁻¹).

By contrast, a second scan carried out for **1**, **2**, **carbo[6]helicene**, and **7TH** with a semiempirical method (AM-1 level) produced results that were in better agreement with the experimental observations (27.8 kcal mol⁻¹ for **1** and 28.9 kcal mol⁻¹ for **2** (Fig. S51†)). In any case, the enantiomerization barrier is not much influenced by the presence of the Mes groups. As for **1**, a comparison of the minimum and transition-state structures of **2** and **7TH** suggests that the smaller double-bond character of the B=O pairs allows for greater flexibility of the six-membered heterocycles compared to benzene rings (Fig. S52†).¹⁰

Circular dichroism (CD) spectra recorded in *n*-hexane/CH₂Cl₂ (95 : 5) at 20 °C showed a perfect mirror-image relationship between the first and second eluted enantiomers of **2** (Fig. 8a); the optical rotation values are -264° (ee = 98%) and $+261^\circ$ (ee = 97%), respectively, which are in the same order of magnitude as those of **1** (-97.5° ; $+97.7^\circ$).

To assign the absolute configuration, the (*M*)-structure of **2** was optimized at the M06/6-311+G(d,p) level of theory. Similar to compound **1**, only one stable conformer was found.

The comparison of calculated and experimental spectra is shown in Fig. 8b; a good correspondence is obtained between



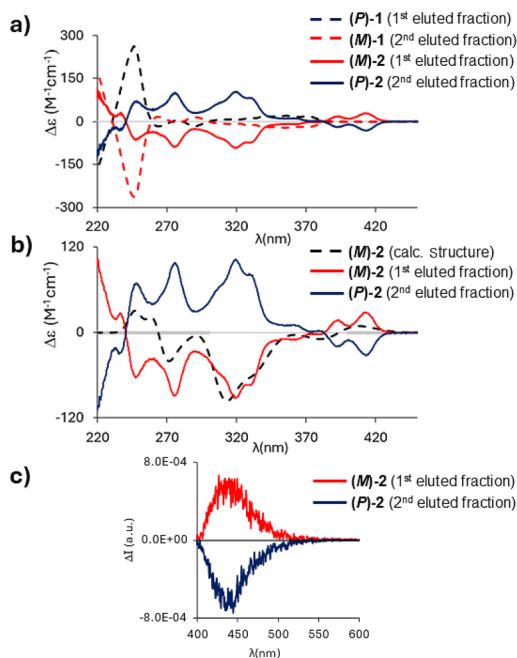


Fig. 8 (a) Circular dichroism (CD) spectra of the **1** (dashed lines) and **2** (solid lines) enantiomers in *n*-hexane/CH₂Cl₂ (95 : 5) at 20 °C; red traces correspond to the (*M*)-structures and blue traces to the (*P*)-structures. (b) CD spectra of (*M*)-**2** (red trace) and (*P*)-**2** (blue trace) in *n*-hexane/CH₂Cl₂ (95 : 5) at 20 °C; computed CD spectrum of the optimized (*M*)-**2** structure is shown for comparison (brown dashed trace; multiplied by a factor of 1.5). (c) Circularly polarized luminescence (CPL) spectra of (*M*)-**2** (red trace) and (*P*)-**2** (blue trace) in *n*-hexane/CH₂Cl₂ (95 : 5) at 20 °C.

the calculated spectrum and the spectrum of the first eluted enantiomer of **2** (opposite to the case of **1**, for which the (*M*)-structure matches the CD spectrum of the second eluted enantiomer). Fig. 8a illustrates how the inversion of the B–O bonds leads to remarkably different CD spectra: while **1** shows only weak spectral features above λ = 260 nm,¹⁰ **2** has a more intense and structured CD spectrum with the first feature recorded at 414 nm, in correspondence with the first shoulder observed in the absorption spectrum (Fig. 5a).

Circularly polarized luminescence (CPL) spectra of **2** showed a band at λ = 440 nm (Fig. 8c); the resolution of our CPL apparatus does not allow the reproduction of all the vibronic features observed in the emission spectrum of **2** (Fig. 5a). The dissymmetry ratio (defined as $g_{\text{lum}} = \Delta I/I$ and $g_{\text{abs}} = \Delta A/A$)²⁹ in emission is $g_{\text{lum}} = 0.8 \times 10^{-3}$ and is comparable to those of other organic molecules³⁰ and lower than g_{abs} of the first CD band ($g_{\text{abs}} = 2 \times 10^{-3}$; Fig. 8b). The first transition, responsible for the absorption/emission band (in the two optimized geometries), is associated with an electric and magnetic dipole transition moment parallel to the C₂-axis. Interestingly, the most intense CD band ($g_{\text{abs}} = 8.5 \times 10^{-3}$) is found at 320 nm (Fig. 8b); conversely, in the corresponding region of the absorption spectrum, a low-intensity band is observed (Fig. 5a). This indicates that electric and magnetic dipole moments have the prevalent component along the helical axis (see Table S15†).

Conclusions

We reported the synthesis and full characterization of the (BO)₂-helicene **2**, which was obtained through a four-step synthesis involving a Mallory photocyclization as the last step. Moreover, the study of the synthesis of the singly-doped BO-helicene (**3**) was also described.

A combined experimental and theoretical effort allowed us to evaluate (1) the property-switch produced by the BO bond inversion **1** → **2**, along with (2) the effect of BO-doping on **7TH**.

Although the effect on the quantum yield is negligible ($\Phi_{\text{PL}}(\mathbf{2}) = 7\%$ vs. $\Phi_{\text{PL}}(\mathbf{1}) = 6\%$), switching of BO vectors proved to be a useful tool to achieve a lowered LUMO, red-shifted absorption/emission bands (≈40 nm), and a smaller energy gap ($E_{\text{g}}^{\text{opt}}(\mathbf{2}) = 2.90$ eV vs. $E_{\text{g}}^{\text{opt}}(\mathbf{1}) = 3.16$ eV).

As previously observed for **1**, comparison of **2** with **7TH** highlighted that the incorporation of BO bonds does not significantly improve fluorescence performances ($\Phi_{\text{PL}}(\mathbf{2}) = 7\%$ vs. $\Phi_{\text{PL}}(\mathbf{7TH}) = 5\%$) but provides a detectable CPL signal ($g_{\text{lum}} = 0.8 \times 10^{-3}$).

At low temperature (77 K), **2** showed intense and structured phosphorescence (peaked at 556 nm; τ = 63.3 ms), reflecting noticeable vibronic components from “helix breathing” modes. This is important evidence that the incorporation of B–O bonds enhances the flexibility of the helical scaffold, as also highlighted by kinetics studies ($\Delta G_{\text{enant}}^{\ddagger}(\mathbf{2})$ is 10 kcal mol⁻¹ lower than $\Delta G_{\text{enant}}^{\ddagger}(\mathbf{7TH})$).

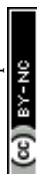
These findings now indicate how BO-doping is a means of influencing the rigidity of tetrathia[7]helicenes and demonstrate how it is possible to manipulate the electronic distribution across the helix by changing the position of the B centers within the scaffold. Overall, our results provide essential information for the synthesis and the structural design of new (BO)-helicene derivatives with tailored optoelectronic and mechanical properties.

Author contributions

L.M. and M.W. conceived the idea. L.M., M.W., E.L., and C.B. wrote the manuscript. All authors discussed and commented on the manuscript. In detail: L.M. and C.B. synthesized the compounds; L.M. and M.P. performed spectroscopic measurements; S.A. and S.G. performed the electrochemical measurements; S.M. and C.V. performed the enantiomer separation and measured CD spectra; G.L. and G.M. performed the theoretical calculations and measured CPL spectra; A.V. and M.B. performed single crystal X-ray diffraction studies; L.M., M.W., E.L. and C.B. directed and supervised the project.

Data availability

The data that support the findings of this study are available from the corresponding authors upon request.



Conflicts of interest

There are no conflicts to declare.

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