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# *In situ* reconstruction of vegetable sponge-like Bi<sub>2</sub>O<sub>3</sub> for efficient CO<sub>2</sub> electroreduction to formate†

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The electrochemical reduction reaction of CO<sub>2</sub> provides a renewable method to close the carbon cycle and alleviate the global energy issue. Bi-based electrocatalysts present huge prospects for catalyzing formate-selective CO<sub>2</sub> reduction. Herein, we fabricated a porous vegetable sponge-like bismuth oxide (VS-Bi<sub>2</sub>O<sub>3</sub>) for the selective electroreduction of CO<sub>2</sub> to formate, which underwent *in situ* reconstruction to form 2D nanosheets containing metallic Bi and Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub>. We propose that the unique porous morphology and low crystallinity of VS-Bi<sub>2</sub>O<sub>3</sub> are beneficial for the *in situ* generation of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub>, which might play a significant role in enhancing the CO<sub>2</sub> electrocatalysis performance. The catalyst delivers a 93.7% faradaic efficiency of formate at 400 mA cm<sup>-2</sup> and shows stability for over 18 h at 100 mA cm<sup>-2</sup>.

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## Introduction

The electrochemical CO<sub>2</sub> reduction reaction (CO<sub>2</sub>RR) using renewable electricity sources is considered to be an ideal strategy for storing energy and relieving the global environment crisis caused by carbon emissions.<sup>1–3</sup> Through an electrocatalytic reaction, CO<sub>2</sub> could be converted into valuable chemicals and fuels such as formic acid, CO, methanol, ethanol, ethylene, and acetate.<sup>4,5</sup> Formic acid or formate, with a high energy density and versatile chemical features, can be applied in plenty of areas including pharmaceuticals, the leather industry, hydrogen storage, fuel cells, and metallurgy.<sup>6,7</sup> Its critical potential has been presented in industrial applications for reducing CO<sub>2</sub> to formate *via* electrocatalysis.<sup>8,9</sup> However, some obstacles, such as the sluggish reaction kinetics and high electricity consumption, need to be overcome before large-scale application. CO<sub>2</sub> is a linear molecule with the highest oxidation state, which means that it has a high thermodynamic stability (*i.e.*, 806 kJ mol<sup>-1</sup> energy is needed to break the C=O bond).<sup>4,10–14</sup> A large reorganization energy in the first electron transfer step to form CO<sub>2</sub><sup>•-</sup> from CO<sub>2</sub>, -1.90 V *versus* the standard hydrogen electrode, imposes a high overpotential

and a slow reaction rate.<sup>15</sup> Furthermore, the hydrogen evolution reaction (HER) competes with the desired CO<sub>2</sub>RR, leading to a selectivity loss for the generation of formate.<sup>3,11</sup> Therefore, developing efficient electrocatalysts with low overpotentials and high selectivity becomes a key challenge for exclusively reducing CO<sub>2</sub> to formate.

Bismuth-based (Bi-based) materials, as efficient electrocatalysts, have aroused great interest in producing formic acid or formate from CO<sub>2</sub>, due to their low cost, low toxicity and environmentally benign traits.<sup>16–18</sup> Due to the low carbon monoxide adsorption energy and strong stabilization of intermediates, Bi-based materials are thermodynamically favorable for yielding formate instead of competitive CO or H<sub>2</sub>.<sup>19,20</sup> To implement the reaction, a high overpotential of over 300 mV must generally be applied along with a low current density (usually less than 100 mA cm<sup>-2</sup>, especially in H cell system).<sup>21,22</sup> Structural engineering methods, such as morphology, component and defect engineering, have been reported to promote the catalytic activities of catalytic materials.<sup>23,24</sup> Earlier studies regarding Bi-based catalysts with different morphologies, including dendrites, nanowires, nanoflakes, bismuthene nanosheets as well as nanoparticles, have been reported to deliver a high formate selectivity of over 90%.<sup>7,25</sup> However, once they interact with the surrounding reactants or products under the reduction conditions, most of the catalysts are inclined to undergo structural self-reconstruction, which can change their morphology and structure, which thus further alters the activity and selectivity of the catalysts.<sup>22,26</sup> For instance, Yao *et al.* proposed that the KHCO<sub>3</sub> electrolyte could moderate the dissociation and conversion of the

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Bi-based metal–organic framework (Bi-MOF) into  $\text{Bi}_2\text{O}_2\text{CO}_3$ , and the applied cathodic potential further helped to reduce the  $\text{Bi}_2\text{O}_2\text{CO}_3$  to metallic Bi, of which the unsaturated surface Bi atoms served as active sites.<sup>22</sup> The work of Ma *et al.* reported bismuthene (Bi-ene) nanosheets derived from monoclinic scheelite  $\text{BiVO}_4$  under working conditions, which significantly enhanced the  $\text{CO}_2$  reduction performance.<sup>27</sup> We anticipate that designing new precatalysts with specific structures will play a significant role in developing highly efficient Bi-based catalysts.

$\text{Bi}_2\text{O}_3$ , which has both high physical and chemical stability, is convenient for synthesis reactions without being much affected by the surrounding environment. This can help to rule out unnecessary influences and focus on the research priorities. Besides,  $\text{Bi}_2\text{O}_3$  possesses  $\text{CO}_2\text{RR}$  advantages like other Bi-based materials. For example, the Bi–O structure of  $\text{Bi}_2\text{O}_3$  has been reported to be conducive to  $\text{CO}_2\text{RR}$  *via* enhancing the  $\text{CO}_2$  adsorption capacity and improving the stability of the  $\text{CO}_2^{\bullet-}$  intermediate.<sup>28,29</sup> Many Bi-based catalysts with the Bi–O structure, such as  $\text{Bi}_2\text{O}_3$  nanosheet/nitrogen-doped graphene quantum dots ( $\text{Bi}_2\text{O}_3\text{-NGQDs}$ ),<sup>30</sup>  $\beta\text{-Bi}_2\text{O}_3$  fractals<sup>31</sup> and  $\text{Bi}_2\text{O}_3$  nanosheets grown on a conductive multi-channel carbon matrix ( $\text{Bi}_2\text{O}_3\text{NS@MCCM}$ ),<sup>32</sup> exhibit a high  $\text{CO}_2\text{RR}$  performance. Our aim is to prepare a highly active  $\text{Bi}_2\text{O}_3$  material and study its *in situ* reconstruction process in electrolytic  $\text{CO}_2$  reduction.

Herein, a porous vegetable sponge-like bismuth oxide (VS- $\text{Bi}_2\text{O}_3$ ) was synthesized using a microwave ultrasonic synthesis method and exhibited an excellent performance for catalyzing  $\text{CO}_2$  to formate with a faradaic efficiency (FE) of around 93% under a potential ranging from  $-0.53$  to  $-1.29$  V (*versus* the reversible hydrogen electrode (RHE)), where all potentials are referenced to RHE unless mentioned otherwise) and a current density up to  $400$   $\text{mA cm}^{-2}$  in the flow cell system. *In situ* reconstruction of VS- $\text{Bi}_2\text{O}_3$  took place under  $\text{CO}_2\text{RR}$  conditions, with a fine nanosheet structure containing metallic Bi and  $\text{Bi}_2\text{O}_2\text{CO}_3$  formed. The porous structure of VS- $\text{Bi}_2\text{O}_3$  might be favorable for reconstruction during the  $\text{CO}_2\text{RR}$ , and the generated  $\text{Bi}_2\text{O}_2\text{CO}_3$  could maintain the superior catalytic performance of formate generation, with scarce attenuation occurring under a current density of  $100$   $\text{mA cm}^{-2}$  over 18 hours.

## Experimental section

### Synthesis of materials

The VS- $\text{Bi}_2\text{O}_3$  was fabricated using a featured microwave ultrasonic synthesis method based on a previous report with modifications.<sup>33</sup> Firstly,  $489.8$  mg  $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$  and  $505.0$  mg dicyandiamide were ultrasonically dissolved in  $50$  mL ethylene glycol (EG) in sequence to form a clear solution using a special three-neck flask. Then the flask was assembled with the corresponding condensation glassware using a microwave/ultrasonic/UV combined catalytic synthesizer (XH-300UL-2+, Beijing Xianghu Science and Technology Development Co., Ltd). A constant-temperature heating mode was chosen to synthesize the VS- $\text{Bi}_2\text{O}_3$  sample. Typically, the microwave heating power was

limited to the maximum of  $300$  W and the assisting ultrasonication was applied using a constant power of  $500$  W. The temperature was increased to  $150$  °C within  $10$  min and held there for another  $10$  min by the automatic program. The obtained white precipitate was then centrifugated and washed thoroughly using deionized water as well as ethanol in sequence. The white VS- $\text{Bi}_2\text{O}_3$  was obtained after drying at  $60$  °C for  $12$  h using a vacuum oven.

The produced VS- $\text{Bi}_2\text{O}_3$  was then calcined using a muffle furnace to fabricate the contrast bulk  $\text{Bi}_2\text{O}_3$  (B- $\text{Bi}_2\text{O}_3$ ). The temperature was set as  $500$  °C and the furnace was supposed to reach that temperature in  $100$  minutes after which it remained at the same temperature for  $1$  h.

### Characterization

X-Ray diffraction (XRD, D/max2550V) was carried out in order to interpret the crystal structure of the products. Scanning electron microscopy (SEM; Hitachi S4800) and transmission electron microscopy (TEM) were employed for the morphology change. Scanning transmission electron microscopy (STEM) and TEM characterization were performed using a Thermo Fisher Talos F200X instrument. Energy dispersive X-ray spectroscopy (EDS) was carried out using 4 in-column Super-X detectors. X-ray photoelectron spectroscopy (XPS; Thermo Escalab 250Xi) was carried out using an Al  $K\alpha$  X-ray beam ( $1486.6$  eV) and was used to obtain more detailed information on the chemical compositions. In the meantime, the C 1s peak centered at  $284.8$  eV was set as the reference for calibration of all of the binding energies. Raman analysis was performed using a Leica DMLM microscope (Renishaw). The excitation wavelength and laser power were set as  $532$  nm and  $3$  mW, respectively.

### Electrochemical measurements

The experiments were carried out using a home-made flow cell system with three poly(methyl methacrylate) (PMMA) plates to divide the anolyte, catholyte and  $\text{CO}_2$  gas chambers, respectively. The anion exchange membrane (Fumasep FAB-PK-130, Fuel Cell Store) was placed between the anolyte and catholyte chambers for separation of the electrolyte and the exchange of anions. Nickel foam of size  $3 \times 3$   $\text{cm}^2$  was used as the anode, counter electrode. The Ag/AgCl ( $3.5$  M KCl) reference electrode was put in the cathode compartment. A total of  $11.5$  mg VS- $\text{Bi}_2\text{O}_3$  was mixed with  $25$   $\mu\text{L}$  Sustainion™ XA-9 solution and  $1$  mL dispersion liquid (isopropanol: $\text{H}_2\text{O}$  =  $3:1$ , v/v), followed by ultrasonication for  $30$  min to form an ink, which was then sprayed onto a piece of commercial Sigracet gas diffusion layer (GDS 28BC, Fuel Cell store), whose area was set as  $3 \times 3$   $\text{cm}^2$ . After drying, a piece of  $1.5 \times 1.5$   $\text{cm}^2$  gas diffusion electrode (GDE) was tested using a flow cell reactor as the working electrode. The  $\text{CO}_2$  gas feed rate was set as  $20$  standard cubic centimeters per minute (sccm). The electrolyte was  $1$  M KOH, which flows at a rate of  $10$  mL per min. All potentials were calibrated to the reversible hydrogen electrode (RHE) with  $iR_{\text{cell}}$  compensation ( $i$  serves as the applied current and  $R_{\text{cell}}$  denotes the cell resistance) according to the

equation  $E_{\text{RHE}} = E_{\text{Ag/AgCl}} + 0.205 + 0.059 \times \text{pH} + i \times R_{\text{cell}}$ . Gas products were analyzed *via* gas chromatography (RAMIN, GC2060), using a thermal conductivity detector (TCD) for detecting  $\text{H}_2$  and a flame ionization detector (FID) for the detection of CO and other gaseous hydrocarbons. The liquid product of  $\text{HCOO}^-$  was quantified using  $^1\text{H}$  nuclear magnetic resonance (NMR). For analysis, 500  $\mu\text{L}$  electrolyte with the liquid product in it was taken out to mix with 100  $\mu\text{L}$   $\text{D}_2\text{O}$  (with TMSP as the internal standard) to form the mixture quantified later using NMR.

## Results and discussion

XRD measurements were performed in order to figure out the phase of the product. In Fig. 1a, the XRD pattern of the obtained VS- $\text{Bi}_2\text{O}_3$  matches with cubic  $\text{Bi}_2\text{O}_3$  (PDF#27-0052), revealing its cubic phase. Besides, the wide peaks indicate its weak crystallinity with peaks at  $27.9^\circ$ ,  $46.4^\circ$  and  $55.1^\circ$  corresponding to the (111), (220) and (311) facets, respectively.

The scanning electron microscopy (SEM) images (Fig. 1b and c, and Fig. S1 and S2, ESI<sup>†</sup>) show a vegetable sponge-like morphology, assembling multiple pieces of nanosheets into a hierarchical micron structure. Transmission electron microscopy (TEM) results (Fig. 1d and e and Fig. S3, ESI<sup>†</sup>) show lots of open pores in the VS- $\text{Bi}_2\text{O}_3$ . Moreover, the high-resolution TEM (HR-TEM; Fig. 1f) image displays no obvious lattice fringe. The selected area electron diffraction (SAED; inset in Fig. 1f) image exhibits a diffraction ring with a radius of 3.19 Å that corresponds to the (111) plane of  $\text{Bi}_2\text{O}_3$  (PDF#27-0052), confirming the weak crystallinity of VS- $\text{Bi}_2\text{O}_3$ . The B- $\text{Bi}_2\text{O}_3$  obtained from the VS- $\text{Bi}_2\text{O}_3$  *via* annealing in air has a high crystallinity, with the XRD pattern aligning well with monoclinic  $\text{Bi}_2\text{O}_3$  (PDF#41-1449) (Fig. S4, ESI<sup>†</sup>).

The  $\text{CO}_2\text{RR}$  processes for VS- $\text{Bi}_2\text{O}_3$  and B- $\text{Bi}_2\text{O}_3$  were carried out in 1 M KOH electrolyte (pH = 14) using a flow cell system (Fig. S5, ESI<sup>†</sup>). During the  $\text{CO}_2\text{RR}$ , VS- $\text{Bi}_2\text{O}_3$  underwent *in situ*

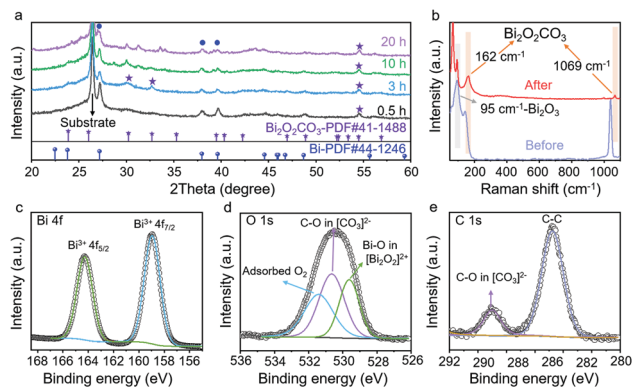


Fig. 2 (a) XRD patterns of VS- $\text{Bi}_2\text{O}_3$  electrolyzing the  $\text{CO}_2\text{RR}$  at  $100 \text{ mA cm}^{-2}$  after 0.5, 3, 10 and 20 h, respectively.  $\text{Bi}_2\text{O}_2\text{CO}_3$  and Bi were formed and maintained during the  $\text{CO}_2\text{RR}$ . (b) Raman spectra of VS- $\text{Bi}_2\text{O}_3$  before (blue line) and after (red line) the  $\text{CO}_2\text{RR}$  for 3 h.  $\text{Bi}_2\text{O}_2\text{CO}_3$  (orange shadow) is formed in the electrolytic process. (c–e) Bi 4f, O 1s and C 1s XPS spectra, respectively, of the post-electrolysis VS- $\text{Bi}_2\text{O}_3$  (at  $100 \text{ mA cm}^{-2}$  for 24 h), demonstrating the existence of  $\text{Bi}_2\text{O}_2\text{CO}_3$ .

reconstruction according to the XRD measurements for different electrolysis times. As shown in Fig. 2a, tetragonal  $\text{Bi}_2\text{O}_2\text{CO}_3$  (PDF#41-1488) appeared after 0.5 h of electrolysis at  $100 \text{ mA cm}^{-2}$ . As the catalysis process proceeds, the featured peaks of  $\text{Bi}_2\text{O}_2\text{CO}_3$  (marked with purple stars) become noticeable for the XRD patterns after 3, 10, and 20 h. Furthermore, the Raman spectrum of the VS- $\text{Bi}_2\text{O}_3$  sample after electrolysis also demonstrates the formed  $\text{Bi}_2\text{O}_2\text{CO}_3$  *via* the featured peaks appearing at 162 and  $1069 \text{ cm}^{-1}$  (in light orange shadow),<sup>22,34</sup> while the peak at  $95 \text{ cm}^{-1}$  is attributed to  $\text{Bi}_2\text{O}_3$  (in light grey shadow) (Fig. 2b).<sup>35</sup> As for B- $\text{Bi}_2\text{O}_3$ , scarce  $\text{Bi}_2\text{O}_2\text{CO}_3$  was detected after electrolysis (Fig. S6, ESI<sup>†</sup>).

To further analyse the surface valence states, X-ray photoelectron spectroscopy (XPS) was conducted. Post-electrolysis XPS inspections of VS- $\text{Bi}_2\text{O}_3$  (at  $100 \text{ mA cm}^{-2}$  for 24 h) are shown in Fig. 2c–e. The peaks at 159 and 164.3 eV in the Bi 4f spectrum (Fig. 2c) are assigned to  $\text{Bi}^{3+}$ , demonstrating a single oxidation state of  $\text{Bi}^{3+}$ .<sup>28,30</sup> Three obvious fitted peaks can be clearly identified in the O 1s spectrum (Fig. 2d). The peaks at 529.6, 530.7 and 531.5 eV correspond to the O atoms in the Bi–O bonds in the  $[\text{Bi}_2\text{O}_2]^{2+}$  layers, the O atoms in C–O in the  $[\text{CO}_3]^{2-}$  layers, and the oxygen of adsorbed  $\text{O}_2$ , respectively.<sup>7,28,34</sup> As for the C 1s XPS spectrum (Fig. 2e), the peak at 289.0 eV is indexed to the carbon atoms in the  $[\text{CO}_3]^{2-}$  layers.<sup>36</sup> XPS spectroscopy verifies the existence of  $\text{Bi}_2\text{O}_2\text{CO}_3$  after electroreduction, in line with the XRD and Raman results.

To gain further insight into the morphology changes, SEM, TEM and HR-TEM characterizations of the post-electrolysis sample of VS- $\text{Bi}_2\text{O}_3$  were carried out. As shown in Fig. 3a and b, vegetable sponge-like frames of VS- $\text{Bi}_2\text{O}_3$  evolve into a nanosheet morphology after  $\text{CO}_2\text{RR}$  for 10 h. Combined with more post-electrolysis SEM images for different timescales in Fig. S7 and S8 (ESI<sup>†</sup>), we speculate that the morphology of VS- $\text{Bi}_2\text{O}_3$  underwent several transformation steps under different cathodic potentials. The shape first transformed from the vegetable sponge-like to thick nanosheet-assembled plates, and

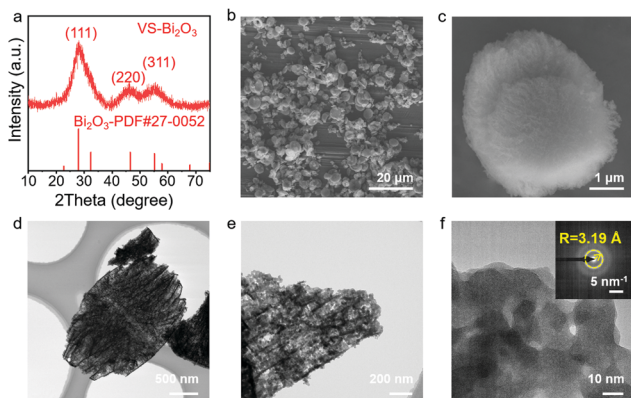


Fig. 1 (a) XRD pattern of VS- $\text{Bi}_2\text{O}_3$  in which the wide peaks indicate the weak crystallinity of VS- $\text{Bi}_2\text{O}_3$ . (b and c) SEM images of VS- $\text{Bi}_2\text{O}_3$  showing the vegetable sponge-like morphology. (d and e) TEM images of VS- $\text{Bi}_2\text{O}_3$  showing lots of open pores. (f) HR-TEM image of VS- $\text{Bi}_2\text{O}_3$  displaying no obvious lattice fringe, and (inset) SAED image exhibiting a classic diffraction ring of  $\text{Bi}_2\text{O}_3$  (PDF#27-0052).



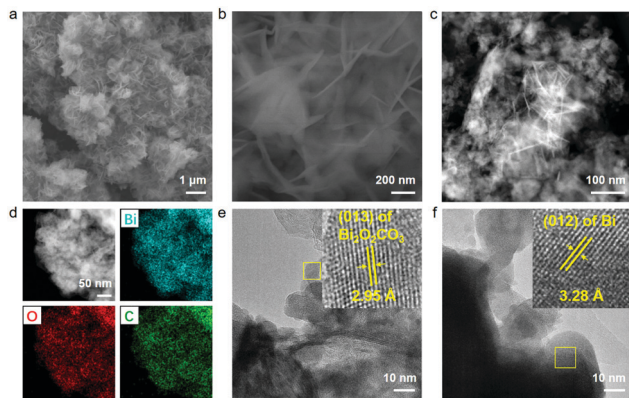


Fig. 3 (a and b) SEM, (c) TEM and (d) STEM-EDS elemental maps of VS-Bi<sub>2</sub>O<sub>3</sub> after the CO<sub>2</sub>RR for 10 h. (e and f) HR-TEM images of VS-Bi<sub>2</sub>O<sub>3</sub> after the CO<sub>2</sub>RR for 10 h in which Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> and metallic Bi are detected *via* their characteristic facets.

then gradually changed into thin nanosheets. The TEM image (Fig. 3c) confirms the nanosheet morphology in accordance with the SEM images. The SEM images after electrolysis of B-Bi<sub>2</sub>O<sub>3</sub> are also displayed in Fig. S9 and S10 (ESI<sup>†</sup>). Scanning transmission electron microscopy-energy dispersive spectroscopy (STEM-EDS; Fig. 3d) shows the uniformly distributed Bi, O and C elements. It is worth noting that the presence of the carbon element to some extent reveals the existence of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub>. The high-resolution TEM image in Fig. 3e proves this point as well, with the lattice spacing of 2.95 Å aligning well with the (013) facet of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub>. Fig. 3f shows a 3.28 Å lattice space in accordance with the (012) plane of metallic bismuth, which was derived from the *in situ* reduction of VS-Bi<sub>2</sub>O<sub>3</sub>.

The above measurements illustrate that the prepared VS-Bi<sub>2</sub>O<sub>3</sub> is prone to undergo reconstruction under the working conditions. The sample was changed into metallic bismuth and Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> after electrolysis, among which the Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> did not appear in the B-Bi<sub>2</sub>O<sub>3</sub> sample (Fig. S6, ESI<sup>†</sup>). Besides, the porous vegetable sponge-like morphology is inclined to be exposed to the operating environment since its open pores can be filled with electrolytes and reactants. This may lead to the reconstruction from VS-Bi<sub>2</sub>O<sub>3</sub> to Bi and Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> during the CO<sub>2</sub>RR process. Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> is a layered structure consisting of CO<sub>3</sub><sup>2-</sup> and [Bi<sub>2</sub>O<sub>2</sub>]<sup>2+</sup>, which has been reported to exhibit a lower overpotential compared with metallic bismuth.<sup>37</sup>

To figure out the phase transformation from Bi<sub>2</sub>O<sub>3</sub> to Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> and Bi, three XRD samples obtained under different conditions were prepared (Table S1 and Fig. S11, ESI<sup>†</sup>). The results indicate that VS-Bi<sub>2</sub>O<sub>3</sub> did not react directly with the KOH electrolyte in the absence of CO<sub>2</sub> under ambient conditions. However, when VS-Bi<sub>2</sub>O<sub>3</sub> was introduced into the CO<sub>2</sub> atmosphere, Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> was generated with no metallic Bi formed. Furthermore, the VS-Bi<sub>2</sub>O<sub>3</sub> changed into metallic Bi and Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> with the current density applied. Based on the above results, we speculated that the Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> might be formed *via* the intercalation of CO<sub>3</sub><sup>2-</sup> into the layered structure of the bismuth oxide compound. Previous reports have revealed that the presence of CO<sub>2</sub> during the CO<sub>2</sub>RR process plays a

significant role in maintaining the stable existence of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> thanks to the consumption of most cathodic electrons.<sup>37,38</sup> Besides, a high local pH was also reported to be beneficial for the formation of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub>, and the 1 M KOH electrolyte in this work can provide a large amount of OH<sup>-</sup>, which might promote the generation of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub>.<sup>7</sup>

We then carried out the CO<sub>2</sub>RR to assess the electrocatalytic performance and identify the role of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub>. As shown in Fig. 4a and Fig. S12 (ESI<sup>†</sup>), VS-Bi<sub>2</sub>O<sub>3</sub> achieves a high current density of up to 100 mA cm<sup>-2</sup> at a small applied potential of only -0.53 V, with an accompanying high FE(HCOO<sup>-</sup>) of around 93.6%. The FE(HCOO<sup>-</sup>) remains at a high level of >93.0% over a wide potential range from -0.53 V to -1.29 V, with a potential window of 760 mV. The maximum FE(HCOO<sup>-</sup>) of 94.9% appears at -1.12 V delivering a current density of 300 mA cm<sup>-2</sup>. The overall current density of 400 mA cm<sup>-2</sup> is achieved at a potential of -1.29 V, in the meantime sustaining a high FE(HCOO<sup>-</sup>) of around 93.7%. As for the contrast sample, B-Bi<sub>2</sub>O<sub>3</sub> attains the same current density with a much more negative potential and a much lower selectivity for formate. Taking the current density of 200 mA cm<sup>-2</sup>, for example, B-Bi<sub>2</sub>O<sub>3</sub> needs an applied potential of -1.11 V to reach that goal, which is 380 mV more negative than VS-Bi<sub>2</sub>O<sub>3</sub> with a low FE(HCOO<sup>-</sup>) of 85.3% under the operational conditions (Fig. 4b, and Fig. S13, ESI<sup>†</sup>). Furthermore, the potential grows markedly for the larger current densities, and a potential of -1.58 V is needed for 400 mA cm<sup>-2</sup>, negatively shifted by 300 mV compared with VS-Bi<sub>2</sub>O<sub>3</sub>. The CO<sub>2</sub>RR performance measurements unveil the superb activity and exclusive selectivity of VS-Bi<sub>2</sub>O<sub>3</sub> for the electrochemical conversion of CO<sub>2</sub> into formate, indicating the significance of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> derived from VS-Bi<sub>2</sub>O<sub>3</sub>.

Since efficient charge transfer at the catalyst/electrolyte interface is beneficial to the catalytic process, we carried out electrochemical impedance spectroscopy (EIS) in order to measure the kinetics of charge transfer.<sup>35</sup> Fig. 4c shows the Nyquist plots of VS-Bi<sub>2</sub>O<sub>3</sub> and B-Bi<sub>2</sub>O<sub>3</sub> under open-circuit potential at a steady state, from which a smaller arc can be observed in VS-Bi<sub>2</sub>O<sub>3</sub> (red) compared with B-Bi<sub>2</sub>O<sub>3</sub> (blue). In the Nyquist plots, the charge transfer resistance (*R*<sub>ct</sub>) of the electrochemical reaction can be represented by the radius of the arc, and a smaller arc usually means a smaller resistance. Thus, VS-Bi<sub>2</sub>O<sub>3</sub> possesses a highly accelerated charge-transfer process thanks to its much shorter radius than the contrast bulk sample. We attributed this remarkable improvement to the appearance of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> derived from VS-Bi<sub>2</sub>O<sub>3</sub>. Hence durability testing was carried out at the current density of 100 mA cm<sup>-2</sup>, and the potential remained steady without any significant degradation for an electrolysis process of over 18 h, demonstrating its superior stability for converting CO<sub>2</sub> to formate (Fig. 4d). The performance comparison of VS-Bi<sub>2</sub>O<sub>3</sub> with previous Bi-based materials is listed in Table S2 (ESI<sup>†</sup>).

Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> is a layered structure consisting of CO<sub>3</sub><sup>2-</sup> and [Bi<sub>2</sub>O<sub>2</sub>]<sup>2+</sup>.<sup>37</sup> The presence of metastable oxides has previously been reported to be able to stabilize the reduced CO<sub>2</sub> intermediate, which plays a vital role in lowering the overpotential

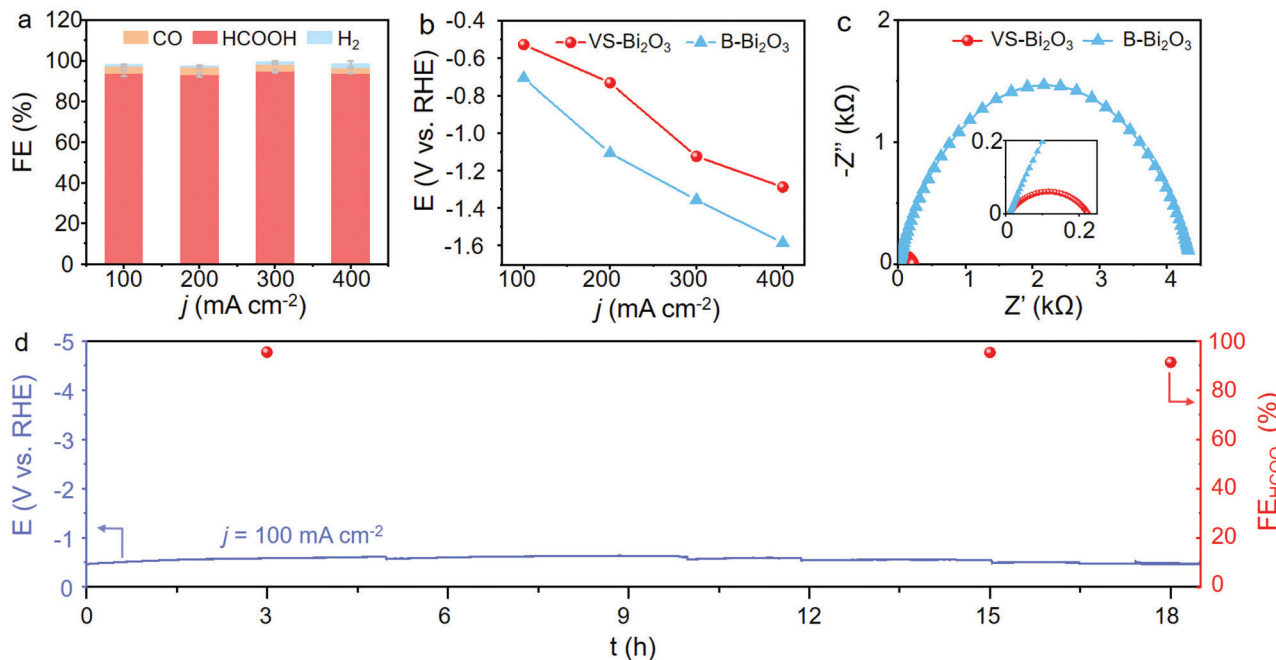


Fig. 4 (a) Faradaic efficiencies plot (with error bars) of VS-Bi<sub>2</sub>O<sub>3</sub> in 1 M KOH electrolyte using a flow cell system. (b) Current density dependence of potential values for VS-Bi<sub>2</sub>O<sub>3</sub> and B-Bi<sub>2</sub>O<sub>3</sub>. (c) Fitting results of Nyquist plots for VS-Bi<sub>2</sub>O<sub>3</sub> and B-Bi<sub>2</sub>O<sub>3</sub> under open-circuit potential at a steady state, where the inset shows an expanded version of the area near the origin. (d) Stability testing of VS-Bi<sub>2</sub>O<sub>3</sub> at 100 mA cm<sup>-2</sup>.

and accelerating the charge transfer required to reduce CO<sub>2</sub> into the target formate product.<sup>7,37–39</sup> The Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> was generated *in situ* from VS-Bi<sub>2</sub>O<sub>3</sub> during the CO<sub>2</sub>RR process, whose existence after electrolysis at 100 mA cm<sup>-2</sup> for 20 h revealed that the formed Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> was robust enough to survive under an applied cathodic bias. However, the contrast sample B-Bi<sub>2</sub>O<sub>3</sub> was reduced *in situ* to metallic Bi without Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> and was accompanied by an unsatisfactory performance. Therefore, the presence of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> might play a significant role in enhancement of the CO<sub>2</sub>RR performance.

## Conclusions

In summary, through the adoption of a facile microwave/ultrasonic synthesis method, we fabricated the porous vegetable sponge-like bismuth oxide (VS-Bi<sub>2</sub>O<sub>3</sub>), which underwent *in situ* reconstruction to form metallic Bi and Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> nanosheets, exhibiting excellent CO<sub>2</sub>RR activity with near-unity formate selectivity under a high current density of 400 mA cm<sup>-2</sup>. Besides, the durability at 100 mA cm<sup>-2</sup> could reach over 18 h, suggesting its remarkable long-term stability. A series of characterization measurements, such as post-electrolysis XRD, Raman, and TEM, indicate the existence of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub>, revealing its robust stability. We further demonstrate that the porous vegetable sponge-like configuration may play a critical role in the *in situ* reconstruction to form Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub>, and that the appearance of Bi<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> could enhance the CO<sub>2</sub>RR performance. This study emphasizes the significance of the precatalyst structure on its *in situ* reconstruction and CO<sub>2</sub>RR performance optimization.

## Author contributions

Hua Gui Yang, Fangxin Mao and Peng Fei Liu designed and guided the study; Yingli Shi and Chun Fang Wen performed the experiments and analyzed the experimental data; Xuefeng Wu carried out the TEM experiments; Jia Yue Zhao helped in the analysis. All the authors discussed and commented on the data and contributed to the manuscript.

## Conflicts of interest

There are no conflicts to declare.

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