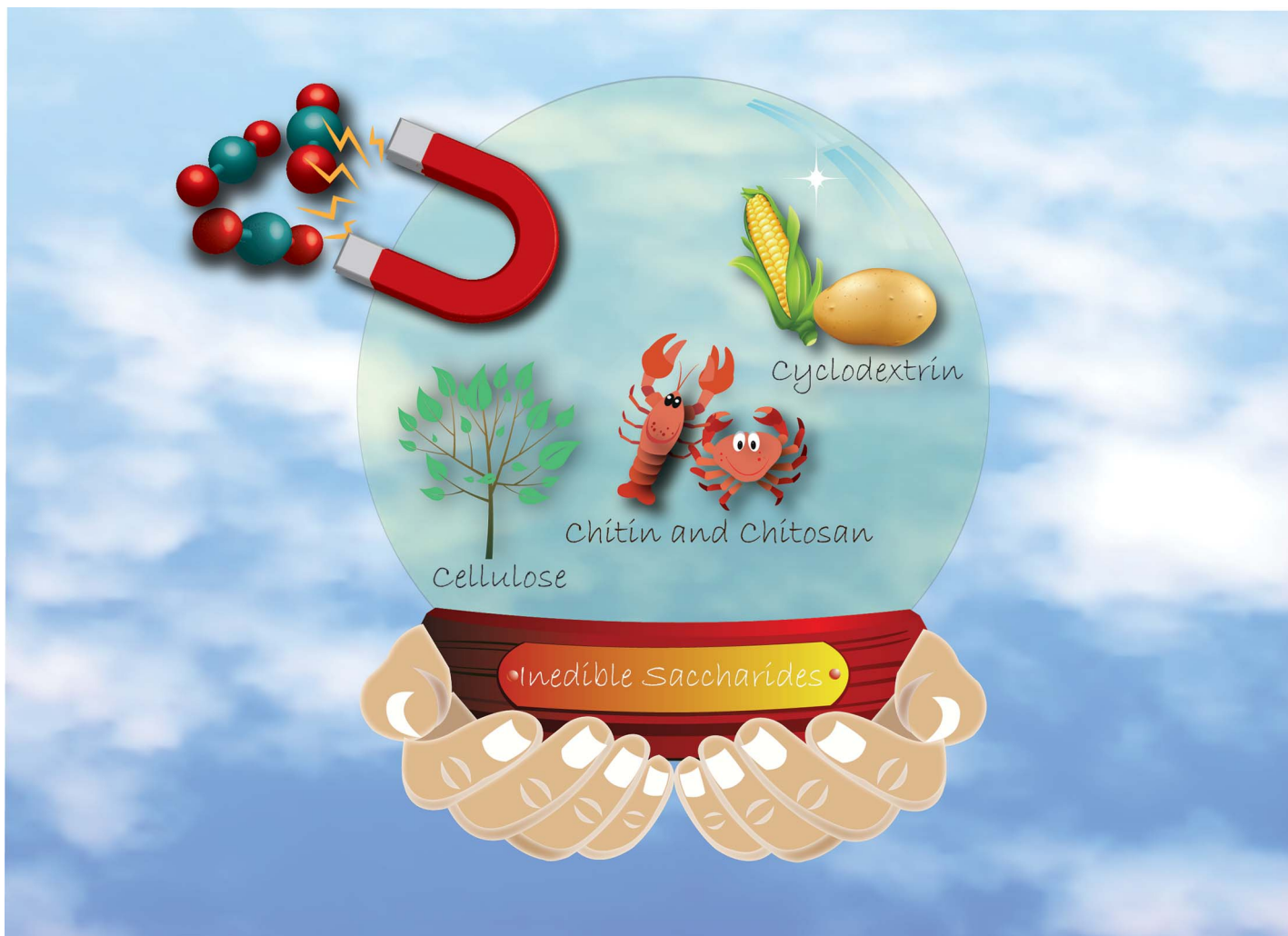




Showcasing collaborative research by the Jordanian CO₂ team (JCO₂T), represented by A. K. Qaroush (The University of Jordan, Jordan), K. I. Assaf (Jacobs University Bremen, Germany) and A. F. Eftaiha (The Hashemite University, Jordan)

Inedible saccharides: a platform for CO₂ capturing

Abundant, cost-effective, inedible saccharides represent potential candidates that can be exploited as green sorbents for CO₂ capturing. The efforts of the scientific community on using a wide spectrum of saccharides including cyclodextrins, chitin, chitosan and cellulose to capture CO₂ through physi- and chemisorption via dry and wet scrubbing applications are highlighted. The usage of oligomers, functionalisation of polymers, passing of “supramolecular chemisorption” convention and introduction of a Green tax are also suggested to mitigate CO₂.

As featured in:



See Abdussalam K. Qaroush, Khaleel I. Assaf, Ala'a F. Eftaiha *et al.*, *Chem. Sci.*, 2018, 9, 1088.



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MINIREVIEW

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[View Journal](#) | [View Issue](#)Cite this: *Chem. Sci.*, 2018, 9, 1088Inedible saccharides: a platform for CO₂ capturingAbdussalam K. Qaroush,^a Hiba S. Alshamaly,^b Shrouq S. Alazzeq,^a
Ream H. Abeskhron,^a Khaleel I. Assaf^c and Ala'a F. Eftaiha^b

The economic viability of eco-friendly and renewable materials promotes the development of an alternative technology for climate change mitigation. Investigations reported over the past few years have allowed understanding the mechanism of action for a wide spectrum of saccharides toward carbon dioxide (CO₂), in terms of reactivity, reversibility, stability and uptake. Exploiting bio-renewables, viz., inedible saccharides, to reduce the anthropogenic carbon footprint upon providing a sustainable and promising technology that is of interest to different groups of scientists, to overcome demerits associated with the current state-of-the-art aqueous amine scrubbing agents, following a "green chemistry guideline", by employing materials with properties relevant to the environment toward sustainable development. The interdisciplinary nature of research in this area provides a large body of literature that would meet the interest of the broad readership of different multidisciplinary fields. Although many reports emphasize the use of biomass in various industrial products ranging from pharmaceuticals, medical preparations, soaps, textiles, cosmetics, household cleaners, and so on, to our knowledge there is no focused article that addresses the application of saccharides for CO₂ sequestration. In this review, we highlight the recent advances on the use of oligo-, poly- and cyclic saccharides to achieve a reversible binding of CO₂. The future research directions are discussed to provide insight toward achieving sustainable development through implementing bio-renewables.

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Introduction

The on-going rise of the anthropogenic greenhouse gas (GHG) footprint has a substantial impact on ecological systems,^{1–3} particularly in the absence of stringent international policies for climate change mitigation. Although the ultimate goal of the intended nationally determined contributions, which are

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Abdussalam K. Qaroush received both BSc and MSc degrees from the Hashemite University, Jordan. While performing his PhD, he joined the group of Professor Bernhard Rieger (TUM) developing the concept "Green Sorbents for CO₂ Activation". In 2015, he was called by the Higher Council for Science and Technology in Jordan as a Jordanian Scientist and Technologist Abroad (JoSTA) for technology transfer regarding

CO₂ activation where he established the Jordanian CO₂ team (JCO₂T). In 2017, Dr Qaroush started his tenure track at the University of Jordan, working on CO₂ mitigation (storage, utilisation, recycling) upon applying green technology, polymer chemistry, and organic synthesis.

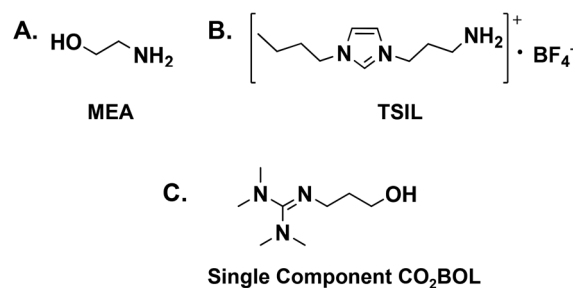


Hiba S. Alshamaly received her BSc in chemistry from the University of Jordan in 2009. Currently, Hiba is pursuing her MSc degree at the Hashemite University upon a joint project under the supervision of both Dr Eftaiha and Dr Qaroush. Her MSc thesis involves CO₂ activation utilising commercially available materials as green sorbents for CO₂ capturing.



adopted in the Paris agreement, is to maintain the average temperature of the globe at two degrees Celsius above that of the pre-industrial revolution era,^{4,5} preliminary indicators demonstrate that advanced economies have failed to meet their commitments to depress GHG emissions,⁶ which keeps the high concentration of CO₂ (the most dominant GHG) as one of the main environmental concerns that should be addressed worldwide through boosting carbon offset strategies^{7,8} toward a concrete international climate policy for the benefit of future generations.⁹

One of the most efficient approaches for climate change mitigation is carbon capture and storage (CCS), which allows the reduction of CO₂ emissions while facing the sustained increase in the global energy demand.^{10–12} The limitations associated with the state-of-the-art monoethanolamine (MEA, Scheme 1A) solvents such as corrosiveness, volatility and the energy penalty required for solvent regeneration, together with other related CCS based technologies, make implementing the principles of green chemistry,^{13–15} which were introduced by Paul Anastas and John Warner, of particular importance to



Scheme 1 The chemical structures of (A) monoethanolamine (MEA); (B) 1-*n*-propylamine-3-butyliimidazolium tetrafluoroborate, an example of TSIL prepared by the Jr Davis group;¹⁶ and (C) 2-(3-hydroxypropyl)-1,1,3,3-tetramethylguanidine (a single component CO₂BOL).¹⁷

design environmentally benign materials and innovative routes to ultimately achieve sustainable development by eliminating hazardous chemicals and energy consuming processes.



Shrouq S. Alazzeah is a bachelor student at the University of Jordan who is interested in green and analytical chemistry. She is keen to participate in scientific events, conferences and workshops to improve her skills in order to help her continue in the advancement of her studies after graduating from her BSc degree. She always has the curiosity to discover unknown chemical reactions which motivated her to start improving her scientific skillfulness at a young age. Shrouq is looking forward to being a distinctive researcher.



Khaleel I. Assaf was born in Amman, Jordan, in 1985. He received his BSc and MSc degrees in chemistry at the Hashemite University, Jordan. In 2015, he graduated with a PhD degree from Jacobs University Bremen, Germany, under the supervision of Prof. Dr Werner Nau, and has subsequently proceeded as a post-doctoral fellow. His research interests lie in the areas of computational and supramolecular chemistry.



Ream H. Abeskchron was born and raised in Amman, Jordan. Currently, she is a third-year bachelor student at the University of Jordan. She has developed a passion for chemistry from a very young age, especially in exploring the mechanisms of chemical reactions. Now, she is a research fellow who is passionate to work with green chemistry. Ream is looking forward to getting her MSc and PhD degrees in organic chemistry.

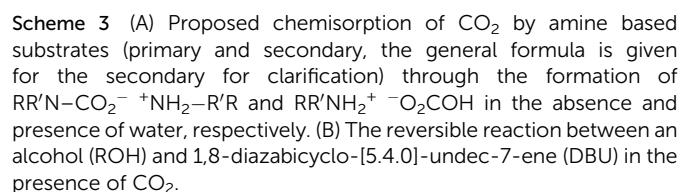


Ala'a Eftaiha has been an assistant professor of physical chemistry at the Hashemite University, Jordan since 2014. He obtained his BSc and MSc in chemistry from the Hashemite University. Prior to doctoral studies, he was a researcher at the Jordanian Pharmaceutical Manufacturing Company. In 2013, Ala'a earned his PhD at the University of Saskatchewan under the supervision of Prof. Dr

Matthew F. Paige, and subsequently worked at Dalhousie University as an NSERC-DREAMS postdoctoral fellow. In 2016, he was a Fulbright visiting scholar at the University of California Santa Barbara. His research interests include CO₂ capturing, emerging photovoltaic technologies and surfactant monolayers.



The second pathway results in the formation of an inorganic or organic carbonate ($M_2(CO_3)_x$, M^{+x} : a metal cation, and $ROCO_2^-$, respectively) following a 1 : 1 reaction mechanism. Aside from $M_2(CO_3)_x$, the formation of $ROCO_2^-$ by activating alcohols to react with CO_2 was reported in the literature by two



research groups. Firstly, Jessop and co-workers^{39,40} devised CO₂ binding organic liquids (CO₂BOLs) which comprised of anhydrous alcohol and a superbase (SB). An example reaction is shown in Scheme 3B. The high volatility associated with CO₂-BOLs and the hardship of drying/handling, together with the high corrosive character, paved the way toward single component CO₂BOLs¹⁷ (Scheme 1C) that unfortunately exhibited high viscosity upon chemisorption. Afterwards, Sir J. F. Stoddart's group^{41–44} reported that the hydroxyl groups located at the rim of γ -cyclodextrin-based metal-organic frameworks (CD-MOFs) are capable of reacting with CO₂ reversibly to form a metal-stabilized organic alkyl-carbonate (a detailed discussion is provided, *vide infra*).

Most of the literature reports focus on the capturing process highlighting the high-energy penalty associated with the regeneration of conventional aqueous amine solvents in comparison with alcohol/SB binary mixtures that suffer from high volatility and corrosiveness, which makes the use of aprotic solvents a reasonable alternative to increase the efficiency of carbon dioxide desorption as they have a lower heat capacity than aqueous based sorbents and a higher vapor pressure as introduced by Park and colleagues,⁴⁵ and followed-up by our research group.^{28,46,47} As a matter of fact, the solvation of ions is of great importance when forming CO₂ adducts from neutral solutions. Accordingly, cations are well solvated in polar aprotic solvents, while anions are poorly solvated, therefore, this might favour a shift in the equilibrium towards the neutral species and, consequently, pK_a values greater than that formed while working in aqueous solutions.⁴⁵

In this minireview, recent advances on the sorption of CO₂ by inedible saccharides is discussed as follows: (1) a review of the reversible reaction between cellulose and CO₂, and the subsequent dissolution and regeneration; (2) the capture and storage of CO₂ by cyclodextrins; (3) the chemisorption of CO₂ by chitin and chitosan using ILs, oligomers and nanoparticles; and (4) CO₂ capturing using functionalised cellulosic materials. We conclude with the challenges and future outlooks of exploiting bio-renewables to be implemented in the area of CCS research.

Cellulose

Cellulose, a low-value agricultural by-product, is the most naturally occurring biopolymer, with a production capacity of *ca.* 180 billion tons per annum by plants.^{48,49} It is a linear polymer consisting of β -(1,4)-linked glucopyranose units (Scheme 2A). It is characterized by strong inter- and intra-molecular H-bonding which play a major role in its structural stability, chemical properties and resistance to enzymatic hydrolysis. One major obstacle that hinders the effective utilisation of cellulose is its insolubility in conventional solvents due to the semicrystalline structure.²¹ Over the past few years, several solvent systems for the dissolution of cellulose have been reported^{50–54} that suffered from serious drawbacks such as corrosiveness, high cost, and reduced dissolution capacity in the presence of polar solvents as well as recycling and purification difficulties which required alternative eco-friendly

approaches for cellulose processing such as the use of pressurised CO₂ to precipitate cellulose from ILs.^{55–59}

Following the concept of switchable solvents, Jérôme and co-workers⁶⁰ reported on the dissolution of microcrystalline cellulose (MCC) suspended in a mixture of dimethyl sulfoxide (DMSO) and 1,5-diazabicyclo[4.3.0]non-5-ene (DBN) upon bubbling CO₂. In the presence of a SB, the primary hydroxyl groups along the cellulose backbone reacted with CO₂ resulting in the formation of carbonate species. The process could be reversed by heating the solution at 30 °C under vacuum, by which the cellulose was precipitated (Fig. 1). After exposing this precipitate to CO₂, the cellulose was dissolved again which highlighted the reversibility of the system. It is noteworthy that the use of weak bases was not efficient to dissolve cellulose under the investigated conditions, as they were incapable of activating the hydroxyl groups.

Another example of a reaction between cellulose and CO₂ in the presence of a SB (namely 1,8-diazabicyclo[5.4.0]-undec-7-ene, DBU), yielding [DBUH]⁺[cellulose-OCO₂][–], was reported by Yang *et al.*⁶¹ The reaction allowed a successful dissolution of the cellulose pulp, and subsequently allowed its modification into cellulose acetate using acetic anhydride (Scheme 4A). Mu and co-workers⁶² monitored the formation of the aforementioned adduct *via in situ* infrared spectroscopy. The formed carbonated cellulose underwent a chemical reaction with a silane-based reagent (TEOS: tetraethyl orthosilicate) which provided a green method for the preparation of a super-hydrophobic nanocoated material (Scheme 4B). Moreover, a cellulose-graft-poly(L-lactide) copolymer was synthesized, exploiting the reversible character of the aforementioned reaction. DBU had a dual action in cellulose dissolution with its role as an organocatalyst for the subsequent ring-opening polymerization process (Scheme 4C).⁶³

Once again, Mu's research group⁶⁴ reported a detailed study of the dissolution process to understand the mechanism for the formation of a MCC carbonated/SB adduct in DMSO upon bubbling CO₂. It was found that DMSO helped in the dissociation of the adduct into free ions by balancing the concentration of the ions and the number of hydrogen bonds. In this context, density functional theory (DFT) calculations revealed multiple interactions between the protonated base and the carbonated cellulose. Very recently, the chemisorption of CO₂ onto cellulose in aqueous NaOH solution in the absence of a SB has been

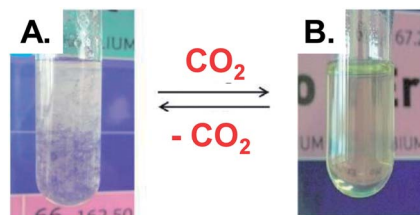


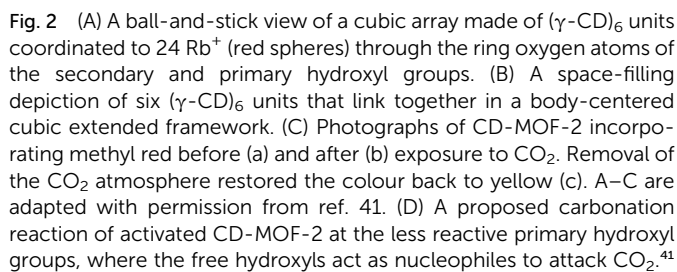
Fig. 1 Photographs of (A) a suspension of MCC in DMSO containing 9 wt% of DBN and (B) a transparent, viscous solution with a maximum concentration of 6 wt% cellulose, obtained after exposing the suspension at 40 °C to 2 bar of CO₂ for 1 hour. Reprinted with permission from ref. 60.





The use of 1,1,3,3-tetramethylguanidine (TMG), in combination with an alcohol dissolved in DMSO revealed a new switchable solvent system that was capable of dissolving cellulose after exposure to CO₂.⁶⁷ The authors attributed the dissolution phenomenon to the formation of a guanidinium alkylcarbonate salt, as identified by both ¹³C NMR (the emergence of a new peak at 159.2 ppm) and *in situ* FTIR (newly formed asymmetric and symmetric carbonyl stretching bands at 1663 and 1410 cm⁻¹, respectively). The dissolution of cellulose was alcohol dependent. In the case of ethylene glycol, a fully transparent cellulose solution was observed, allowing for a maximum de-crystallization up to 10 wt% concentration with high stability. In accordance with a previous report,⁶⁰ no dissolution of cellulose was observed when weak bases ($pK_a < 12$) were utilised.⁶⁷ In the same context, Nanta *et al.*⁶⁸ applied a similar switchable solution consisting of DBU and ethylene glycol in DMSO under high-pressure carbon dioxide to facilitate the dissolution and modification of cellulose derived from cassava pulp waste.

The chemisorption of CO₂ by CD-MOF-2 was proved using cross-polarization magic-angle-spinning (CP/MAS) ¹³C NMR spectroscopy through the emergence of a new peak at 158 ppm that was attributed to the formation of the alkyl carbonic acid based CD-MOF, which was further supported calorimetrically. As shown in Fig. 2C, the addition of methyl red to CD-MOF-2 crystals resulted in a yellow colour due to the basic environment within the framework. Upon exposure to CO₂, the yellow



The challenge in synthesizing CD-MOFs with cations other than K^+ , Rb^+ and Cs^+ , is the inability to obtain porous, crystalline frameworks. Very recently, Patel *et al.*⁴⁴ demonstrated the partial substitution of K^+ with Li^+ through co-crystallizing KOH and LiOH with γ -CD. As shown in Table 1, the most stable Li/K-CD-MOF with the highest $\text{Li}^+:\text{K}^+$ ratio (0.61 : 1.18) had a comparable Brunauer–Emmett–Teller (BET) surface area and a characteristic better sorption in comparison with those obtained for other iso-structural CD-MOFs.

windows (small triangular-shape voids of diameter 4.2 Å, Fig. 3) of the $[(\text{RB}^+)_4(\gamma\text{-CD})]_6$ cubic cages after the carbonation reaction, where the reactive primary hydroxyl groups are located, thus reducing the mass transfer of intermediate molecules. The sensor reusability was tested by measuring the change in conductivity of a CD-MOF-2 pellet following sequential CO_2 -sorption and desorption (by heating at $\sim 80^\circ\text{C}$ for several cycles). It was noteworthy that the conductivity change was higher at low CO_2 concentration (less than 20% in the presence of N_2 gas).⁴⁰

Aside from CD-MOFs, it has been reported that the inclusion complex of 1 : 1 β -CD and aniline (CD-aniline) in association with residual water molecules inside the CD cavity, chemisorbed CO_2 with a sorption efficiency of 0.69 mmol of CO_2 per g sorbent through the formation of bicarbonate according to the following reaction: $\text{C}_6\text{H}_5\text{NH}_2 + \text{CO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{C}_6\text{H}_5\text{NH}_3^+ \text{HCO}_3^-$. The sorption capacity was decreased by 37% when the sorbent was heated to 100 °C under vacuum, which emphasized the importance of the water molecules to the sequestration process.⁷⁵ Although the performance characteristics of the CD-aniline complex were much lower than the corresponding CD-MOFs, it sorbed 0.85 mol CO_2 per mol nitrogen (N) in comparison with the amine efficiency of 0.5 mol CO_2 per mol N following a 2 : 1 sorption mechanism under dry conditions, which was unlikely to take place in the case of the explored inclusion complex. Unlike CD-MOFs (*vide supra*), the sorption capacity of the amorphous CD-aniline complex (obtained upon grinding) was almost similar to the as-synthesised crystalline material.

Han and co-workers⁷⁶ reported the facile and elegant synthesis of CD-based microporous carbon (CDPC) using acid catalysed solvothermal carbonization at 180 °C. The obtained materials were regular spheres with diameters between 300 and 800 nm, BET surface areas ranging from 600 to 700 m² g⁻¹ and high content of oxygen-containing functional groups, such as carboxyl, carbonyl, and hydroxyl as inferred from FTIR and ¹³C CP/MAS spectra. As shown in Table 1, the CO₂ gravimetric uptake of α -, β - and γ -CDPC measured at 0 °C and 1.0 bar were between 2.8 and 3.5 mmol CO₂ per g sorbent. Similarly, it has been reported that grafting β -CD with sulfonic acid functional groups (-SO₃H) led to the formation of a porous structure (CD-SO₃H) with a low oxygen content due to dehydration, which was able to physisorb CO₂ 30 times more than β -CD at the same investigated conditions (1.2 of the former against 0.04 mmol CO₂ per g sorbent of β -CD) with high selectivity over O₂, N₂ and CH₄, which makes it a potential candidate for CO₂ separation from flue gas.⁷⁷

Turning to wet scrubbing applications, the hydroxyl groups of the CD macrocycle should be activated to become susceptible to nucleophilic attack prior the exposure to CO₂. It has been reported that a mixture of β-CD and an organic SB, namely DBU, was able to fix CO₂ chemically as an amidinium alkylcarbonate. Due to viscosity issues, the ratio (in equivalents) of β-CD : DBU was kept at 0.5 : 1 throughout the work. The carbonation reaction was followed by ¹H and ¹³C NMR together with FTIR spectroscopy, following the peaks associated with the quaternary carbon of DBU. The relatively low uptake of CO₂ (1.8 wt%)

Table 1 Surface area, mechanism of sorption and sorption capacity of different solid CD-based sorbents

Sorbent	BET surface area (m ² g ⁻¹ sorbent)	Sorption mechanism	Sorption capacity (mmol CO ₂ per g sorbent)	
			0 °C	25 °C
CD-MOF-1	1145 (ref. 44) (1320 (ref. 71))	Chemi & physisorption	4.2 (ref. 44)	3.0 (ref. 44)
CD-MOF-2	1110 (ref. 71)		1.7 (ref. 44)	1.4 (ref. 44)
Li/K-CD-MOF	1205 (ref. 44)		4.5 (ref. 44)	3.1 (ref. 44)
CD-aniline	NA	Chemisorption		0.69 ^c (ref. 75)
α-CDPC	703 (ref. 76)	Physisorption	3.3 (ref. 76)	NA
β-CDPC	699 (ref. 76)		2.9 (ref. 76)	NA
γ-CDPC	678 (ref. 76)		2.8 (ref. 76)	NA
CD-SO ₃ H	211 (ref. 77)		NA	1.2 ^a (ref. 77) 1.7 ^b (ref. 77)

^a The uptake was measured at 1 bar. ^b The uptake was measured at 3 bar. ^c The CO₂ adsorption experiment was carried out at 30 °C.

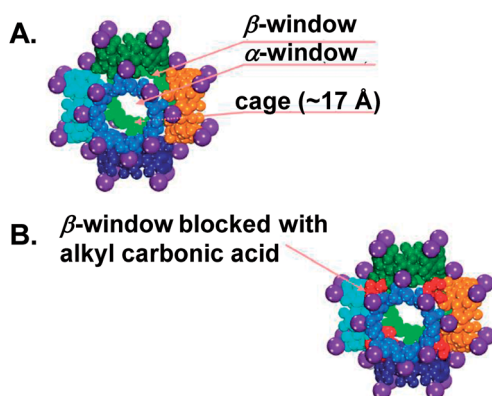


Fig. 3 A space-filling representation of [(Rb⁺)₄(γ-CD)]₆ cubic cages. Purple spheres indicate Rb⁺; the sides of the cube (six-CD rings) are depicted in different colours. Adapted with permission from ref. 43.

was explained by the excessive use of the SB together with the inefficient stirring during the synthesis of the sequestered species. Notably, the sorption capacity was 7.9 wt% in the case of glucose : DBU mixture.⁷⁸

Chitin and chitosan

Among the various, spread polysaccharides, amine containing sugars, in which the -OH group at C-2 of the hexose monomeric unit is replaced by -NH₂, are the main component of chitin (Scheme 2C, *m* > 50%), which is mainly found in crustaceans, insects, molluscs, bacteria and fungi. The insolubility of chitin in common solvents represents a considerable barrier that confronts its employment in various fields. In the early eighties of the last century, Austin⁷⁹ explored the solubility of a chitin-lithium chloride (LiCl) complex in several amide solvents such as DMAc and *N*-methyl-2-pyrrolidone (NMP), followed by several attempts to dissolve chitin using salts and acids (for more details, see the review by Rinaudo²²). Deacetylation of chitin either enzymatically or under alkaline conditions⁸⁰ results in chitosan (Scheme 2C, *n* > 50%). The solubility of chitosan in aqueous acidic media is dependent on several structural parameters such as the degree of deacetylation, the distribution

of the acetyl groups along the polymer backbone and the molar mass.

In 2006, Zhang and co-workers⁸¹ reported on the use of ILs, namely 1-butyl-3-methyl-imidazolium chloride (Bmim)Cl, as an alternative solvent for chitin and chitosan due to its ability to break down the polymer crystallinity under mild conditions as inferred from wide-angle X-ray diffraction (WAXD) measurements. Interestingly, both chitin and chitosan were recovered from the IL solution using methanol or water. The sorption efficiency of 10 wt% solutions of chitosan/IL and chitin/IL were measured gravimetrically using a microbalance at 30 °C and 1 atm CO₂ pressure. Results indicated that the efficiency of the former was 8.1% compared to 3.8% for the latter. Although no spectroscopic evidence for chemisorption was provided, the authors justified the higher sorption capacity of the chitosan-based solution by the chemical fixation of CO₂ due to the presence of free amine groups which are either blocked or presented in a lower fraction in the case of chitin. Regarding regeneration, it was reported that heating the chitosan/IL solution at 100 °C for 30 minutes reduced the sorption efficiency to 1.4%, afterwards the IL solution can be used for subsequent capture/release cycles.

The sorption capacity measurements performed by Mu's research group⁸² using 6 wt% of chitosan or chitin dissolved in 1-butyl-3-methylimidazolium acetate (Bmim)OAc were consistent with those reported previously by Zhang's group.⁸¹ The values obtained using thermogravimetric analysis (TGA) were 6.7 and 3.5% for chitosan and chitin, respectively, whereas it was 1.3% for neat (Bmim)OAc. The almost doubled capacity of the former suggested that chemisorption was the main mechanism for CO₂ capturing through carbamate formation in the case of chitosan, while physical dissolution was responsible for the capturing in the case of chitin. The chemisorption was proved using ATR-FTIR spectroscopy through the disappearance of the peak at 1680 cm⁻¹ which belongs to the bending mode of N-H, the appearance of a new peak centred at 1652 cm⁻¹ ascribed to -N-C(O)- amide moiety of the carbamate ion. The sequestered species was regenerated under 1 atm N₂ at 80 °C. The ability of the compressed CO₂ to act as an anti-solvent^{56,83} was exploited to precipitate chitosan through the reaction between CO₂ and the acetate-based ILs to end up with



the formation of a carboxylate zwitterionic dimer that changes the hydrogen-bonding properties of the IL as inferred from the measurements of the solvatochromic parameters.

Recently, Wasserscheid and co-workers introduced a new type of immobilized IL by encapsulating IL⁸⁴ or ionogel droplets⁸⁵ (chitosan dissolved in IL and diluted with acetonitrile (ACN)) in fumed silica forming a free-flowing IL (ionogel)-in-air powder or inverse supported ionic liquid phase (SILP) material (Fig. 4) after evaporating the ACN. The key advantages of the inverse SILP is the high ratio of IL (ionogel) to the nanoporous material together with relatively short gas diffusion distances.

The gravimetric absorption capacity of the inverse SILP made of 40 wt% IL or ionogel (with 5 wt% of chitosan) and 60 wt% silica measured at 1 bar and 40 °C ranged between 0.1 and 0.8 mmol CO₂ per g sorbent. The results indicated that the CO₂ uptake was increased when an IL with a weakly basic anion such as chloride or 2-methoxyethyl methylphosphonate were used due to the extra carbamation reaction provided by chitosan together with the enhanced solubility of CO₂ in an ionogel compared with the corresponding IL. However, employing a more basic anion such as acetate obstructed the chemistry of sorption by chitosan due to an acidic side product obtained from the carbonation reaction of the IL itself. In comparison with neat chitosan powder, the absorption capacity calculated based on the number of amino groups present in chitosan was improved by 10 times in the inverse SILP. The low absorption efficiency of the neat chitosan was attributed to the restriction of the carbamation reaction at the interface of the powder.

Amine-functionalized porous materials represent an alternative technology for CO₂ capturing.¹² Chitosan represents a promising alternative for mesoporous silica in the quest to find eco-friendly and low-cost materials to be used as an amine support. Fujiki and Yogo^{86,87} reported on porous chitosan beads impregnated with polyethyleneimine (PEI) working under both dry and humid conditions with CO₂ adsorption capacity of 2.38 and 3.61 mmol per g sorbent, respectively, at 15 kPa of CO₂ and 66 °C through the formation of ammonium carbamate as confirmed by *in situ* diffuse reflectance infrared Fourier

transform spectroscopy (DRIFTS). The enhanced CO₂ adsorption under wet conditions was explained by the swelling of the chitosan framework, thus exposing buried amino groups to CO₂, which resulted in the formation of two adsorption states of CO₂ (whereas one state was recognised under dry conditions) as indicated by TGA measurements.

Although Zhang⁸¹ opened new horizons for exploiting chitin for CO₂ capturing through physical dissolution, our research group published the first report on the chemisorption of CO₂ by an oligomeric chitin acetate (CA)/DMSO binary mixture.^{88,89} The only functional groups along this oligomeric material available for the capturing process were the hydroxyl groups whereas the amino groups at C-2 of the amine pyranose rings were blocked either by protonation or acetylation. Herein, the use of DMSO was a prerequisite for the carbonation reaction to take place through the intermolecular hydrogen bonding between the polar aprotic solvent and primary hydroxyl group at C-6 of the pyranose ring which prompted the nucleophilic attack at CO₂ to form the organic alkyl-carbonate as follows: $\text{RCH}_2\text{OH} + \text{CO}_2 \rightleftharpoons \text{RCH}_2\text{OCO}_2^- + \text{H}^+$. The formation of the CO₂ sequestered species was confirmed using ¹³C NMR and *in situ* ATR-FTIR spectroscopy, through the appearance of the benchmark chemical shift at *ca.* 157 ppm together with a peak centred at 1555 cm⁻¹ which was ascribed to the vibrational mode of the carbonate. Notably, glucosamine hydrochloride (the monomeric unit of the oligomer) showed no chemisorption upon bubbling CO₂. The stability aspects of the organic carbonate adduct were understood using ¹H NMR spectroscopy and DFT calculations that indicated that the product was stabilized through non-bonding interactions, namely H-bonding and ionic interaction along the oligomer backbone. This indicated that the supramolecular stabilization was the driving force for the chemisorption, which was absent in the case of glucosamine hydrochloride. The CO₂ sorption uptake was measured volumetrically using *in situ* ATR-FTIR autoclave with a maximum capacity of 3.63 mmol CO₂ per g sorbent for a 10.0% (w/v) CA/DMSO solution. The decrease in the sorption capacity with increased concentration was attributed to the huge increase in viscosity at higher concentration of the oligosaccharide solution, which may have hindered the diffusion of CO₂ to the bulk of the solution.

To ultimately maximize the sorption capacity of CO₂, while using bio-renewable and inedible materials, we moved our attention over oligomeric chitosan due to the presence of two potential active sites to interact with CO₂, through the formation of carbonate and carbamate. In this context, our group reported on the use of a low molecular weight chitosan hydrochloride (CS·HCl) dissolved in DMSO for CO₂ sequestration.⁹⁰ Upon bubbling CO₂, the formation of the organic carbonate adduct was consistent with the previously affiliated system and ascribed to the supramolecular chemisorption. *Ex situ* ATR-FTIR spectroscopy of the CS·HCl/NaOH/DMSO solution confirmed the formation of the other sequestered functionality, *viz.* ammonium carbamate, after bubbling CO₂, through the appearance of the fingerprint frequency at 1709 cm⁻¹ which was further elucidated using ¹H-¹⁵N heteronuclear single quantum coherence (HSQC) NMR spectroscopy, with coupled cross-peaks

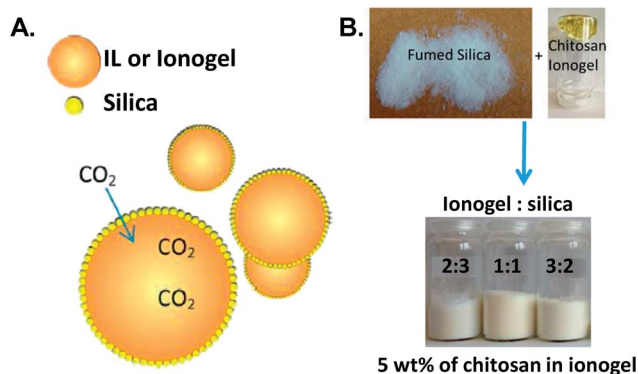
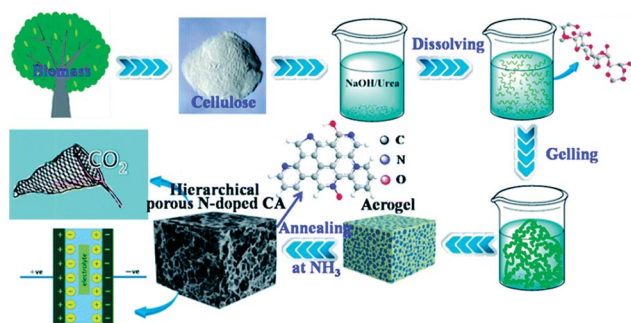


Fig. 4 (A) Depiction of an inverse SILP. (B) A photograph of inverse SILPs prepared from an ionogel comprised of 5 wt% chitosan dissolved in 1-ethyl-3-methylimidazolium acetate ([EMIM][OAc]) with different ionogel to silica ratios. Reprinted with permission from ref. 84 and 85.



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membranes *via* a casting methodology. The mechanical properties and thermal stability of the composites were a function of the plasticizing effect together with the gas affinity of the IL used. Its addition improved the permselectivities of the membranes but lowered the gas permeabilities. It is noteworthy that the use of 40% [E₁Py][NTf₂]/CAct showed a seven-fold increase in CO₂ permeability with CO₂/N₂ and CO₂/CH₄ permselectivities of 32 and 24, respectively.

Surveying cellulose modified materials, Lin's research group¹⁰² doped cellulose triacetate (CTA, a polymer that is used industrially in membranes for the separation of CO₂/CH₄) with 1-ethyl-3-methylimidazolium tetrafluoroborate and 1-ethyl-3-methylimidazolium dicyanamide in order to enhance the cellulose affinity towards CO₂. Upon doping, both the crystallinity and glass transition temperature decreased. An increase in CO₂/CH₄ selectivity and CO₂/N₂ permeability was observed.

A new type of hydrophilic microfibrillated cellulose (MFC) was investigated by Baschetti and co-workers¹⁰³ who used pure MFC films and MFC/Lupamin® (a commercially-available polyvinylamine, BASF). The synthesized material showed good CO₂/N₂ and CO₂/CH₄ selectivity, in the order of 500 and 350 respectively, which was achieved by pure MFC films. The CO₂ permeability was very low and never exceeded 25 Barrer, even at 90% relative humidity. The permeability of a 1:1 MFC/Lupamin® nanocomposite (the fabrication process is shown in Fig. 6) was better with a six-fold enhancement over that of the

In 2017, Dassanayake *et al.*¹⁰⁴ reported an interesting study for a series of nanocrystalline cellulose mesoporous silica composite materials with tethered double-amidoxime functional groups (NCC-AO). The sorption capacities recorded for the NCC-AO were 3.30 and 5.54 mmol CO₂ per g sorbent at 25 and 120 °C, respectively. The authors claimed that the sorption capacity at such elevated temperatures and atmospheric conditions of CO₂ was the highest among other sorbents reported in the literature.

In the same year, Styring and co-workers¹⁰⁵ synthesized a cellulose supported ionic liquid (SoILs) which showed improved CO₂ uptake up to 39-fold in comparison with ILs. This means a reduction in the amount of IL used which can associate faster kinetics for sorption together with faster regeneration cycles. The optimal working ratio of cellulose to IL is 1 : 4, which might result in a reduction of operational costs and size of facilities.

Conclusions

Bio-renewable materials represented by the most abundant polysaccharides might be the next generation of sorbents that are introduced for the capturing of CO₂, which was present heavily in the atmosphere due to the severe anthropogenic effect. In our consensus, several research groups have introduced mild to harsh conditions for the activation of the title materials, and what matters the most is the efficiency/ commercial availability of the chosen sorbent. This is considered as one reason for the utilisation of water-immiscible polymers as solid supports for CO₂ capturing purposes. In addition, another good advantage of such materials is their naturally benign-by-design architecture that provides a wide platform to capture CO₂, ranging from physisorption to chemisorption based mechanisms, dry to wet scrubbing or even the introduction of hybrid materials that can work in parallel with the biopolymers to enhance the CO₂ storage. In our opinion, a strong recommendation for future work is maintained throughout the introduction of oligomers,²⁹ reported elsewhere as a first generation green sorbent, rather than polymers that can mimic the parent polysaccharides for wet scrubbing applications. “Supramolecular Chemisorption” is further facilitated by other hydrogen-bonding, future candidates of different functionalities. Furthermore, the implementation of a new fine named “green tax” (1% per successful fund taken from each accepted proposal submitted to any national/international funding agencies or world-wide universities to plant trees in arid areas and developing countries) might provide a good source of the title saccharides as well as cleansing the atmosphere from pollution while trying to save this green planet. As a matter of fact, what was damaged industrially and disrupted the carbon cycle can be fixed simply and in an eco-friendly manner by returning back to nature. Actually, Climeworks¹⁰⁶ have utilised filters that are primarily composed of cellulose upon grafting with amines to capture CO₂ from the air. Upon reaching the full capacity, the filter can be regenerated at 95 °C.

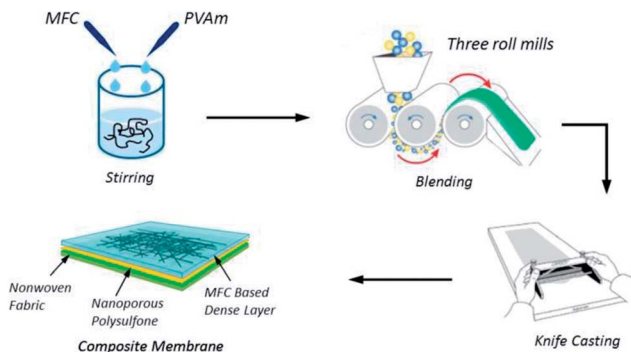


Fig. 6 MFC/Lupamin® nanocomposite preparation. Reprinted with permission from ref. 102.

A further recommendation is to find universal protocols to capture CO₂ in terms of gas uptake measuring conditions (volumetric *vs.* gravimetric), as this may hinder a fair comparison among sorbents.²⁹ Also, modification/functionalisation of the naturally-occurring polymers might provide better reactivities and performance characteristics.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- A further recommendation is to find universal protocols to capture CO₂ in terms of gas uptake measuring conditions (volumetric vs. gravimetric), as this may hinder a fair comparison among sorbents.²⁹ Also, modification/functionalisation of the naturally-occurring polymers might provide better reactivities and performance characteristics.
- ## Conflicts of interest
- There are no conflicts to declare.
- ## Acknowledgements
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