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## Anion-templated 2D frameworks from hexahydroxytriphenylene†

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Hexahydroxytriphenylene (HHTP) forms 2D frameworks through O–H⋯anion hydrogen bonds with a range of anions. In all cases, 1:1 co-crystals of HHTP and tetraalkylammonium-anion salts are obtained, which have a layered 2D structure. When HSO<sub>4</sub><sup>−</sup> was used as anion, an unprecedented methylation reaction was observed giving crystals containing the methylsulfate anion (MeOSO<sub>3</sub><sup>−</sup>).

The vast majority of research into the supramolecular chemistry of anions has focused on their solution phase behaviour spurred by the key role negatively-charged species play in a range of biological and environmental processes.<sup>1</sup> However recently, a handful of reports have demonstrated that anions can also be used to template the formation of self-assembled structures.<sup>2</sup>

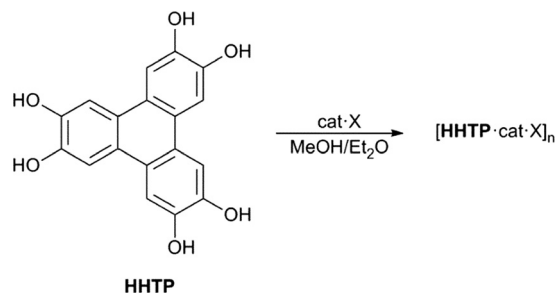
While the interaction of anions with O–H hydrogen bond donors has received comparatively little attention, the groups of Smith,<sup>3</sup> and Kass and Wang,<sup>4</sup> have demonstrated that receptors containing these groups can be potent anion hosts. Very recently, White and MacLachlan have demonstrated that O–H⋯anion hydrogen bonds can also be used to form extended structures including 1D polymers and nanotubes.<sup>5</sup>

Hexahydroxytriphenylene (HHTP, Scheme 1) is an important precursor for a range of supramolecular framework structures. When reacted with transition metals in the presence of base, HHTP gives two or three-dimensional metal organic frameworks (MOFs) containing the sextuply-deprotonated form of the ligand [*i.e.* triphenylene-hexakis(olate)].<sup>6</sup> Some of these MOFs show very high electrical or proton conductivity. Additionally, HHTP has been used to form highly stable covalent organic frameworks (COFs) with boronate derivatives, which show impressive gas storage properties.<sup>7</sup>

Herein, we show that O–H⋯anion hydrogen bonding from the six hydroxyl groups of HHTP can be used to prepare 2D layered structures. The complexes form sheets with one sheet comprised of anions and HHTP ligands, and the second made up of cations and solvent molecules.

We initially tried to prepare HHTP directly from catechol using ammonium persulfate as oxidant,<sup>8</sup> but recovered only unreacted catechol. Instead we found that high purity HHTP could be readily prepared by FeCl<sub>3</sub>-mediated oxidation of 1,2-dimethoxybenzene (veratrole),<sup>9</sup> followed by demethylation using BBr<sub>3</sub> (see ESI†).

We attempted to determine if HHTP binds anions in solution using <sup>1</sup>H NMR spectroscopy, but were unable to obtain quantitative data. We have found HHTP only shows solubility in acetone or polar solvents (alcohols, DMSO), however in polar solvents, solution binding is too weak to be observed, while addition of anions to HHTP in acetone causes precipitation. Previous studies have shown that the interactions between catechol<sup>3</sup> or catechol-containing<sup>5b</sup> molecules and anions are modest in acetonitrile, so we suggest that more competitive solvents such as DMSO or alcohols disrupt the interactions between HHTP and anions allowing them to remain soluble. Less competitive acetone presumably does not do this, and therefore some kind of oligomeric or polymeric HHTP-anion complex forms and precipitates from solution.



**Scheme 1** Structure of HHTP and its assembly into anion-templated 2D polymeric structures (cat = tetraalkylammonium cation, X = anion).

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which markedly different structures were obtained when the cation was changed from TBA to TEA.<sup>5a</sup> In the structure of  $[\text{HHTP}\cdot\text{TEA}\cdot\text{Cl}]_n$ , substantial solvent is present, located within the TEA cation layer, so it would appear that the crystal compensates for the smaller cation by filling the empty space with solvent molecules. To investigate this further, we also attempted to crystallise HHTP in the presence of TEA-Br and TEA-I, but did not obtain any singly crystalline material containing HHTP.<sup>11</sup> While we cannot be certain, we hypothesise that a range of factors determines whether framework crystallisation is favourable, and that the comparatively strong O-H $\cdots$ Cl<sup>-</sup> interactions are sufficiently strong to overcome the otherwise unfavourable crystallisation with TEA cations and associated solvents. Conversely, the O-H $\cdots$ anion interactions are weaker with the less basic anions, Br<sup>-</sup> and I<sup>-</sup> and so the framework structures do not form.

In all cases, we were able to prepare the framework materials in bulk, generally in good yields (67–83%); in the case of the framework prepared from HSO<sub>4</sub><sup>-</sup>, the yield was lower (41%). These materials were investigated by thermogravimetric analysis (TGA), and IR and NMR spectroscopies, which revealed that all organic solvent had been removed from the structures on drying (see ESI<sup>†</sup>).

PXRD patterns for the bulk materials containing TBA cations were generally consistent with those simulated from the single crystal structures, albeit with some loss of crystallinity upon drying (Fig. 3a and S19–S24, see ESI<sup>†</sup> for further details regarding PXRD experiments). In the TEA-containing structure  $[\text{HHTP}\cdot\text{TEA}\cdot\text{Cl}]_n$  there is a significant loss of crystallinity upon drying (Fig. 3b), which is perhaps not surprising given the large

amount of weakly-interacting solvent present in the lattice (this structure contains diethyl ether, whereas the TBA containing structures contain H-bonded methanol or water molecules).<sup>12</sup>

As mentioned above, crystallisation of HHTP and TBA-HSO<sub>4</sub> gives a crystalline product containing the MeOSO<sub>3</sub><sup>-</sup> anion. A search of the CSD<sup>13</sup> reveals 60 structures<sup>14</sup> containing this anion; of which the vast majority result from methylation of a nitrogen-containing species with dimethyl sulfate giving a cationic nitrogen centre as the MeOSO<sub>3</sub><sup>-</sup> salt. In a handful of cases, this anion was formed from the dehydration of methanol by sulfuric acid (*i.e.* H<sub>2</sub>SO<sub>4</sub> + MeOH → MeOSO<sub>3</sub>H + H<sub>2</sub>O, followed by deprotonation of MeOSO<sub>3</sub>H by a basic group within the molecule),<sup>15</sup> but we could not find any examples where the HSO<sub>4</sub><sup>-</sup> anion has been methylated to give MeOSO<sub>3</sub><sup>-</sup>. NMR analysis of bulk material obtained from crystallisation of HHTP with TBA-HSO<sub>4</sub> (Fig. S9 and S10<sup>†</sup>) reveals that approximately half of the anion is present as MeOSO<sub>3</sub><sup>-</sup>, with the remainder presumably unreacted HSO<sub>4</sub><sup>-</sup>. Interestingly, when HHTP and TBA-HSO<sub>4</sub> are dissolved in CD<sub>3</sub>OD, no methylsulfate formation is observed by <sup>13</sup>C NMR spectroscopy, even after seven days in solution,<sup>16</sup> suggesting that crystallisation may be important in the methylation process.

In conclusion, we have demonstrated that O-H $\cdots$ anion hydrogen bonds can be used to assemble the triphenylene-containing ligand HHTP into framework materials. The materials contain 2D sheets formed from HHTP and anions, which adopt different levels of corrugation dependent on the anion and cation size. Despite the fact that O-H $\cdots$ anion interactions are weak in solution (in polar solvents), they reproducibly give crystalline frameworks in good yields. These results demonstrate the importance of weak interactions in crystal engineering and we suggest that if anion-templated materials can be prepared using stronger anion $\cdots$ receptor interactions, these systems may prove useful for a range of materials applications.

## Acknowledgements

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## Notes and references

‡ For further details of SCXRD studies, see the ESI<sup>†</sup> CIF files for all structures are also provided as ESI<sup>†</sup> (CSD no.: 1451578–1451583).

§ This parameter allows useful comparison between hydrogen bonds to differently-sized anions. It is relatively imprecise, as the position of hydrogen atoms is poorly-determined by SCXRD and so values are given to only two significant figures. For a more detailed discussion, see the ESI<sup>†</sup> and N. G. White, C. J. Serpell and P. D. Beer, *Cryst. Growth Des.*, 2014, 14, 3472–3479.

- (a) P. D. Beer and P. A. Gale, *Angew. Chem., Int. Ed.*, 2001, 40, 486–516; (b) N. H. Evans and P. D. Beer, *Angew. Chem., Int. Ed.*, 2014, 53, 11716–11754; (c) N. Busschaert, C.

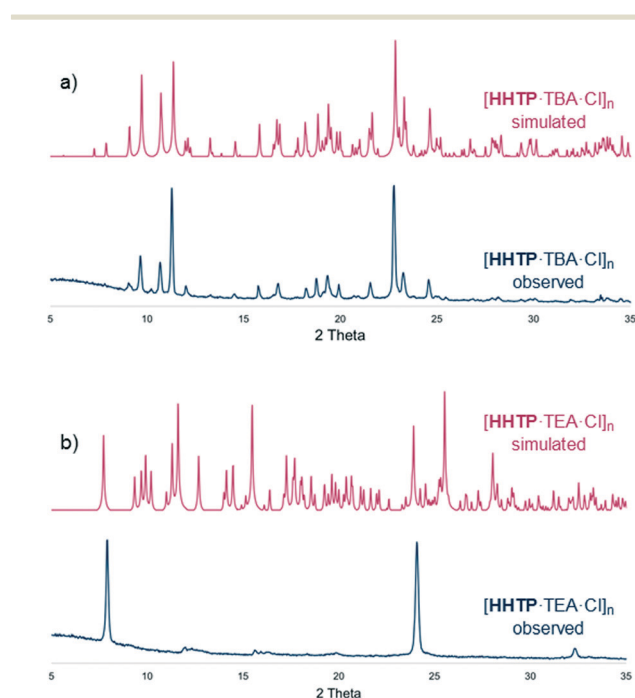


Fig. 3 PXRD traces of dried, bulk samples of a)  $[\text{HHTP}\cdot\text{TBA}\cdot\text{Cl}]_n$  and b)  $[\text{HHTP}\cdot\text{TEA}\cdot\text{Cl}]_n$  as well as traces simulated from SCXRD data (see ESI<sup>†</sup> for full details).



