


 Cite this: *RSC Adv.*, 2022, 12, 32853

Copper oxide nanostructured thin films processed by SILAR for optoelectronic applications

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The lack of high-functioning p-type semiconductor oxide material is one of the critical challenges that face the widespread performance of transparent and flexible electronics. Cu_xO nanostructured thin films are potentially appealing materials for such applications because of their innate p-type semiconductivity, transparency, non-toxicity, abundant availability, and low-cost fabrication. This review summarizes current research on Cu_xO nanostructured thin films deposited by the SILAR technique. After a brief introduction to the advantages of Cu_xO semiconductor material, diverse approaches for depositing and growing such thin films are discussed. SILAR is one of the simplest deposition techniques in terms of better flexibility of the substrate choice, the capability of large-area fabrication, budget-friendly, deposition of stable and adherent film, low processing temperature for the film fabrication as well as reproducibility. In addition, various fabrication parameters such as types of copper salts, pH of precursors, number of cycles during immersion, annealing of as-deposited films, doping by diverse dopants, and growth temperature affect the rate of fabrication with the structural, electrical, and optical properties of Cu_xO nanostructured thin films, which led the technique unique to study extensively. This review will include the recent progress that has recently been made in different aspects of Cu_xO processed by the SILAR. It will describe the theory, mechanism, and factors affecting SILAR-deposited Cu_xO. Finally, conclusions and perspectives concerning the use of Cu_xO materials in optoelectronic devices will be visualized.

 Received 7th October 2022
 Accepted 9th November 2022

DOI: 10.1039/d2ra06303d

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1. Introduction

1.1 Background

Copper (Cu) and copper oxide (Cu_xO) thin films have been studied extensively due to their potential application in semiconductor technology long before the Ge and Si era started, and researchers have faced much more difficult to work with this oldest material ever. The n-type window layer semiconductors such as ZnO, ITO, FTO, and GaN with large bandgap energies have already achieved outstanding optical as well as electronic transport properties. Consequently, the effort of detecting new, prospective p-type absorber layers for optoelectronics devices has led to intensive research.

Cu_xO semiconductors are very attractive and have been broadly studied in both theoretical analysis and investigations into applied executions of nano or optoelectronic devices due to

their chemically stable nature, nontoxicity, relative abundance, potential particle size effects, excellent performance as a catalyst, and fulfill all the requirements for low-cost manufacturing at ambient conditions, which have high potential usage in energy storage, conversion, and next-generation rechargeable lithium-ion batteries.^{1–6} Furthermore, Cu_xO nanostructures are extensively used in other diverse applications, including photovoltaics,⁷ photodetectors,⁸ nanofluid,^{9,10} energetic materials,¹¹ field emissions,¹² supercapacitors,^{13,14} biosensors,^{15,16} gas sensors,^{17,18} photocatalysis,^{19,20} removal of inorganic pollutants,^{21,22} and magnetic storage media.^{23,24}

Both the Cu₂O and CuO show direct transition nature with a direct band gap of around 2.1 and 1.5 respectively, having a high extension coefficient of above 10⁵ cm⁻¹. Since the theoretical limit of the energy conversion efficiency of Cu₂O and CuO is as high as 20 and 29%, respectively under air mass (AM) 1.5 solar illumination, numerous efforts were done to increase the efficiency of Cu_xO solar cells, but the performance remains very poor.²⁵ In the case of Cu₂O solar cells, it is not more than 8.1%,⁷ whereas in the case of CuO it is lower and still about 3%.²⁶ Toward the large area fabrication, it is crucial to establish the thin film growth technique for Cu_xO. Thus, the research of Cu_xO thin films has both high-tech and scientific consequences.

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Cu_xO nanostructured thin films have been synthesized by various approaches like electrodeposition,²⁷ electron beam evaporation,²⁸ magnetron sputtering,^{29–31} molecular beam epitaxy,³² sol–gel,³³ solution growth,³⁴ spin coating,³⁵ successive ionic layer adsorption and reaction (SILAR),^{36,37} thermal evaporation,³⁸ and vapor deposition.³⁹ Among all the deposition methods, SILAR is one of the simplest methods in terms of better flexibility on substrate choice, the capability of large area fabrication and deposition of the stable and adherent film, low processing temperature for film fabrication as well as reproducibility.⁴⁰ This technique is very budget friendly since it does not require any sophisticated equipment. Moreover, various fabrication parameters such as pH, annealing temperature and time, doping elements, the concentration of precursor solutions, and temperature of the precursor solutions affect the rate of fabrication as well as the structural, optical, and electrical properties of the fabricated thin films led the technique unique to study in an extensive manner.

More than a few reviews of different aspects of Cu_xO-based optoelectronics have been published based on the fabrication technique but still no such report for the SILAR technique. This paper concerns the progress that has recently been made in diverse aspects of Cu_xO-based thin films processed by the SILAR method, following the introduction in section one, several deposition techniques are reviewed in section two. The third section of this paper describes the theory and mechanism of Cu_xO-based thin films fabricated by the SILAR method. The fourth section, which incorporated the core focus of this review, leads to the factors that affect SILAR-based Cu_xO deposition which is followed by the application of Cu_xO in section five. Finally, conclusions and perspectives concerning the use of Cu_xO in optoelectronic devices are presented.

1.2 Properties of copper oxides (Cu_xO)

Cu₂O exists as a simple cubic Bravais lattice^{8–10} with a space group of (*Pn3m*) or (*O_h*). Each unit cell consists of six atoms, the four Cu atoms are in a face-centered cubic lattice while the two O atoms are at the tetrahedral positions creating a body-centered cubic sublattice. Thus, O atoms are fourfold coordinated with Cu atoms as closest neighbors, and Cu atoms are linearly coordinated with two O atoms as closest neighbors as shown in Table 1. On the other hand, the unit cell of CuO fits into a monoclinic structure with the space group *C2/c* and the lattice parameters are represented in the table (PDF No. 89-5898). In each CuO unit, there exist four Cu–O bonds. As demonstrated in the table, in a unit, each Cu atom is surrounded by the four closest coplanar O atoms. The four O atoms are positioned at the angles of an almost rectangular parallelogram, which then unites another two O atoms to shape a highly distorted octahedron. The O atom is enclosed by the four closest Cu atoms positioned at the angle of a tetrahedron.

1.3 Band-structure calculation

Ab initio calculations are mandatory to understand the optical and electronic properties of the Cu_xO systems. But there is a challenge for standard *ab initio* investigations based on DFT for both Cu₂O and CuO. The exchange–correlation function is the crucial ingredient in the theoretical description. Fig. 1 and 2 represent the band structures, density of states (DOS), and partial density of states (PDOS) of the Cu₂O and CuO compounds. The results were simulated for both Cu₂O and CuO unit cells using CASTEP software within the LDA + *U* and the calculated bandgaps were found as 1.647 and 1.52 eV respectively.⁴²

Table 1 Crystallographic properties of Cu₂O and CuO⁴¹

Parameters	Cu ₂ O	CuO
Structure: Cu(I)-yellowish, Cu(II)-greyish black, O-red		
	Cubic	Monoclinic
Unit cell	$a = b = c = 4.26 \text{ \AA}$	$a = 4.6837 \text{ \AA}, b = 3.4226 \text{ \AA}, c = 5.1288 \text{ \AA}$
Space group	$\alpha = \beta = \gamma = 90^\circ$ <i>Pn3m</i> (224)	$\alpha, \gamma = 90^\circ, \beta = 99.54^\circ$ <i>C2/c</i> (15)
Bond length, \AA	Cu–O O–O Cu–Cu	
	1.849 3.68 3.012	1.96 2.62 2.90
Cell volume, \AA ³	77.83	81.08
Formula weight	143.14	79.57
Density, g cm ^{–3}	5.749–6.140	6.515
Melting point, °C	1235	1201



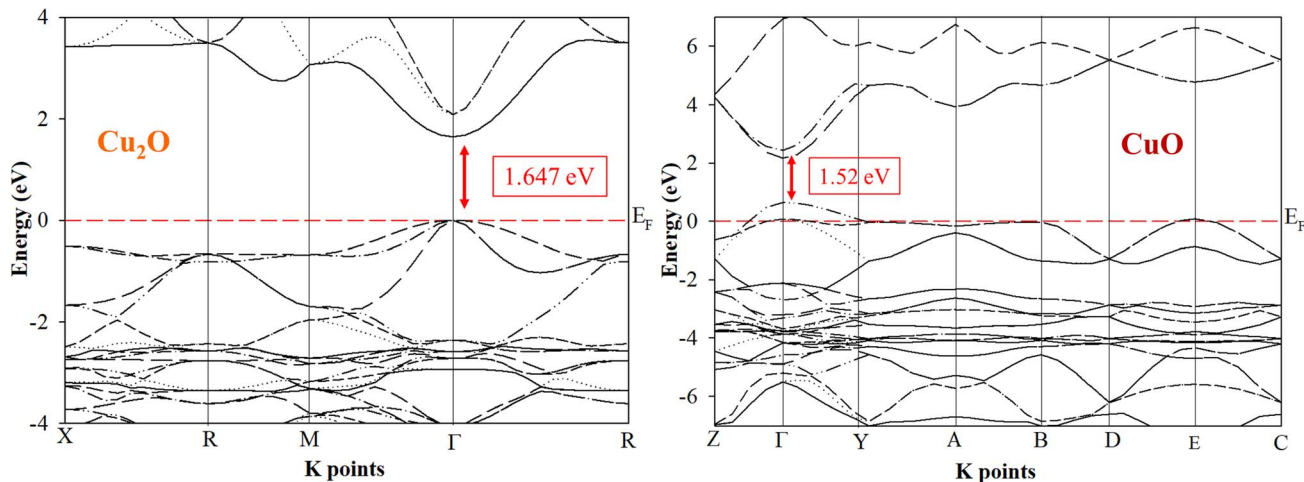


Fig. 1 Band structures of Cu_2O and CuO unit cells are drawn by CASTEP using LDA + U .

2. Thin film deposition process

2.1 Physical deposition methods

The physical deposition methods have diverse techniques to attain thin films with good quality. It can be summarized with the raw materials, deposition conditions as well as cost of production as shown in Table 2.

2.2 Chemical deposition methods

Likewise, diverse chemical deposition techniques with the deposition condition, raw materials, cost of production, the

usual use of substrate *etc.* are discussed in Table 3 as shown below:

2.3 Advantages and disadvantages of deposition techniques

Till now, a lot of deposition techniques are available to fabricate high-quality thin films having diverse applications. For a better understanding, the advantages, and disadvantages of some of the chemical deposition techniques such as chemical bath deposition (CBD), atomic layer deposition (ALD) as well as spin coating are summarized to understand the potentiality of the SILAR method in Table 4.



Fig. 2 DOS and PDOS of Cu_2O and CuO unit cells are drawn by CASTEP using LDA + U .



Table 2 Physical deposition techniques with the requirements during deposition

Physical deposition					
Techniques	Raw materials	Conditions	Substrate	Film quality	Budget Ref.
Thermal evaporation	Red Cu ₂ O powder	Base pressure: 5×10^{-4} Pa Temperature: 300 °C Annealing temperature: 500 °C	Glass tantalum SiO ₂	—	High 43
Electron beam evaporation	Cu ₂ O pellets	Deposition time: 5–16 min Substrate temperature: 200 °C Accelerating voltage: 2, 4 and 6 kV Evaporation pressure: 3×10^{-2} Pa Ultimate pressure: 4×10^{-4} Pa Filament current: 30 mA	Glass	Moderate	High 44
Pulsed laser deposition	Cu, O ₂	Substrate temperature: 25–400 °C p_{O_2} : 0–10 mTorr Vacuum chamber pressure: Torr	Quartz ITO NaCl	Excellent	High 45
Molecular beam epitaxy	Cu, O ₂	Incident O ⁺ beam energy: 50 eV Substrate temperature: 100–400 °C Cu flux: 2.5×10^{13} to 1.6×10^{14} atoms per cm ² s Base pressure: 3×10^{-10} torr Total pressure: 3×10^{-9} to 2×10^{-8} O ⁺ flux: 2.7×10^{14} atoms per cm ² s	MgO	Excellent	High 46
Ion plating evaporation	Cu, O ₂	Base pressure: 10^{-4} Pa O ₂ flow rate: 15–45 sccm N ₂ flow rate: 0–6 sccm RF power: 300 W Annealing temperature: 300 °C Substrate temperature: 25 °C	Glass	Excellent	High 47
Direct current (DC) sputtering	Cu, O ₂	Deposition pressure: 6.3×10^{-3} torr Base pressure: 6×10^{-6} torr Sputtering power: 60 W Ar gas pressure: 20 sccm p_{O_2} : 8.0×10^{-4} to 1.8×10^{-3} torr	Glass	Excellent	High 48 and 49
	Cu, O ₂	Sputtering power: 10–40 W Ar flow rate: 15 sccm O ₂ flow rate: 10 sccm Substrate temperature: 300 °C Ambient gas pressure: 0.045 Pa	Glass stainless steel		
Radio frequency (RF) sputtering	Cu, O ₂	Substrate temperature: 300 °C Ar flow rate: 10 sccm O ₂ flow rate: 0–2 sccm Deposition time: 60 min Background pressure: $<3 \times 10^{-4}$ Pa Working pressure: 1.7–1.8 Pa	Glass	Excellent	High 50

3. Theory and mechanism of SILAR process

SILAR is an extensively applied technique to fabricate high-quality metal oxide or halide thin films.^{84,85} During deposition, successive ionic layer adsorption and reaction of the ions take place at the solid–solution interface of the substrate. Thus, the thin film of the compound, A_xB_y, is deposited onto the substrate surface by dint of the adsorbed cations, xA^{p+} and anions, yB^{q-} due to the following heterogeneous chemical reaction:



where, x, p, q, y and p^+, x^-, y^+, q^- are the number and charges of the corresponding ions A (metal ions), P (cationic precursor), Q (anionic precursor) and B (anions) respectively.^{85,86} Sometimes, the ligands, L_n are a necessity to complete the reaction.^{87–90}

In the case of Cu_xO film deposition mechanism, salts of Cu²⁺ are used to deposit copper oxide thin films. In most of the research on Cu₂O, firstly copper(I) thiosulfate complex is formed by the redox reaction between Cu²⁺ and S₂O₃²⁻ ions which results in a colorless solution. The corresponding reactions are:

Oxidation half-reaction:



Reduction half-reaction:



Table 3 Chemical deposition techniques with the requirements during deposition

Techniques	Raw materials	Chemical deposition			Film quality	Budget	Ref.
		Conditions	Substrate				
Sol-gel	CuCl ₂ ·2H ₂ O, CH ₃ OH, (CH ₂ CH ₂ OH)NH, gluconone, ethylene glycol	Rotating speed: 2000 rpm	TiO ₂	Good	Low	51 and 52	
		Annealing temperature: 200–400 °C	FTO Si	Excellent	Low	53 and 54	
Chemical bath deposition	Cu(NO ₃) ₂ ·3H ₂ O, triethanolamine, hydrazine hydrate	Rotating speed: 1000 and 1500 rpm	Glass	Excellent	Low	53 and 54	
		Annealing temperature: 350, 500 °C	Cu	Excellent	Low	36 and 37	
SILAR	Cu(NO ₃) ₂ ·3H ₂ O, triethanolamine, hydrazine hydrate	Temperature: 30 °C	Glass	Excellent	Low	53 and 54	
		Molar ratio: HNO ₃ : HF-10 : 1 to 135 : 1	Glass FTO	Excellent	Low	36 and 37	
Spray pyrolysis	Cu, HNO ₃ , HF, C ₂ H ₅ OH, Na ₂ SO ₄ , CH ₃ COCH ₃	HNO ₃ conc: 0–1.2 mmol L ⁻¹	Glass FTO	Excellent	Low	36 and 37	
		Water bath time: 2–6 days	Glass FTO	Excellent	Low	36 and 37	
Electrodeposition	CuSO ₄ ·5H ₂ O, NaOH, Na ₂ S ₂ O ₃ ·5H ₂ O	Water-bath temperature: 10–45 °C	Glass	Excellent	Low	55 and 56	
		Temperature: 70 °C	Glass	Excellent	Low	55 and 56	
Electrodeposition	Cu(CH ₃ COO) ₂ ·H ₂ O, (CH ₃) ₂ CHOH, C ₆ H ₁₂ O ₆	pH: 2.35–7.33	Glass	Excellent	Low	55 and 56	
		Temperature: 70 °C	Glass	Excellent	Low	55 and 56	
Electrodeposition	Cu(CH ₃ COO) ₂ ·H ₂ O, (CH ₃) ₂ CHOH	No. of cycles: 20–80	Glass	Excellent	Low	55 and 56	
		Annealing temperature: 75–350 °C	Glass	Excellent	Low	55 and 56	
Electrodeposition	CuSO ₄ ·5H ₂ O, lactic acid, NaOH	Temperature: 200–350 °C	Ti	Excellent	Low	57–60	
		Deposition time: 45 min	FTO	Excellent	Low	57–60	
Electrodeposition	Cu(CH ₃ COO) ₂ ·H ₂ O, (CH ₃) ₂ CHOH	No. of cycles: 450	ITO	Excellent	Low	57–60	
		Temperature: 200–350 °C	Ti ITO Pt	Excellent	Low	57–60	
Chemical vapor deposition (CVD)	CuSO ₄ ·5H ₂ O, tri-sodium citrate dehydrate, C ₆ H ₅ Na ₃ O ₇ ; KOH	Glucose conc: 0–0.08 M	Borosilicate glass	Excellent	High	61–63	
		Isopropanol volume (%): 0–100	Borosilicate glass	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cu(CH ₃ COO) ₂ ·H ₂ O, CH ₃ COONa·3H ₂ O	Cu salt conc: 0–0.08 M	Si wafer	Excellent	High	61–63	
		Temperature: 30 and 60 °C	Glassy carbon	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cupric acetate, sodium acetate, CH ₃ COOH, NaOH	pH: 9, 12	SiO ₂ glass	Excellent	High	61–63	
		Applied potential: –150 to –800 mV	Sapphire	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cu(CH ₃ COO) ₂ ·H ₂ O, CH ₃ COONa·3H ₂ O	Time: 20–60 min	MgO	Excellent	High	61–63	
		Applied potential: –0.4 V	MgO	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cu(CH ₃ COO) ₂ ·H ₂ O, CH ₃ COONa·3H ₂ O	Temperature: 30 °C	MgO	Excellent	High	61–63	
		Temperature: 20–80 °C	MgO	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cu(CH ₃ COO) ₂ ·H ₂ O, CH ₃ COONa·3H ₂ O	Time: 2–80 min	MgO	Excellent	High	61–63	
		NaCl conc: 1–10 mM	MgO	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cupric acetate, sodium acetate, CH ₃ COOH, NaOH	Potential: –0.1 to –0.4 V	MgO	Excellent	High	61–63	
		Temperature: 20, 55 °C	MgO	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cu(CH ₃ COO) ₂ ·H ₂ O, CH ₃ COONa·3H ₂ O	Deposition time: 45 min	MgO	Excellent	High	61–63	
		pH: 5.4–7.49	MgO	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cu(CH ₃ COO) ₂ ·H ₂ O, CH ₃ COONa·3H ₂ O	Applied potential: –200 to –400 mV	MgO	Excellent	High	61–63	
		Cu(II) acetate con.: 0–16 mM	MgO	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cu(CH ₃ COO) ₂ ·H ₂ O, CH ₃ COONa·3H ₂ O	O ₂ flow rate: 1 and 300 cm ³ min ⁻¹	MgO	Excellent	High	61–63	
		Temperature: 300 and 500 °C	MgO	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cu(CH ₃ COO) ₂ ·H ₂ O, CH ₃ COONa·3H ₂ O	p _{O₂} : 1.689 × 10 ² and 5.07 × 10 ⁴ Pa	MgO	Excellent	High	61–63	
		Substrate temperature: 125–225 °C	MgO	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cu(CH ₃ COO) ₂ ·H ₂ O, CH ₃ COONa·3H ₂ O	Process pressure: 1–10 torr	MgO	Excellent	High	61–63	
		Vapor flow rate: 5 sccm	MgO	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cu(CH ₃ COO) ₂ ·H ₂ O, CH ₃ COONa·3H ₂ O	N ₂ flow rate: 100 sccm	MgO	Excellent	High	61–63	
		Pressure: 5–200 sccm	MgO	Excellent	High	61–63	
Chemical vapor deposition (CVD)	Cu(CH ₃ COO) ₂ ·H ₂ O, CH ₃ COONa·3H ₂ O	Temperature: 350–500 °C	MgO	Excellent	High	61–63	
		Temperature: 350–500 °C	MgO	Excellent	High	61–63	



Table 3 (Contd.)

Techniques	Raw materials	Chemical deposition			
		Conditions	Substrate	Film quality	Budget Ref.
Atomic layer deposition (ALD)	Cu(CH ₃ COO) ₂ · H ₂ O, Cu(CH ₃ COO) ₂ , H ₂ O vapor	Temperature: 180–220 °C Reactor pressure: 10 mbar N ₂ , H ₂ O and O ₂ flow rate: 400 sccm Deposition cycles: 500–7000	Glass Si	Excellent	High 64 and 65
	Cu(II)-bis-(dimethylamino-2-propoxide), O ₃	p _{O₂} : 34 Pa Substrate temperature: 112–165 °C Deposition cycles: 500–10000	SiO ₂ /Si		

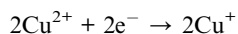
Table 4 Involved advantages and disadvantages of the deposition techniques

Techniques	Merits	Demerits
CBD	Simple and cost effective Stoichiometric deposition ⁶⁶ Low fabrication temperature Capability of depositing large area films (~10 cm ²) ⁶⁷ Various types of substrates are used ⁶⁸ Tuning film qualities by controlling growth parameters Deposition of ternary and quaternary compounds ⁶⁹	Precipitation occurs in the bath causing serious problems Materials are lost ⁷⁰ Films are badly cohered onto the substrate Produces powdery films
SILAR	Facile and economical Potentiality to grow large-surface films (~10 cm ²) ⁷¹ Reproducibility Any kind of substrate can be used No need to use sophisticated instrument or vacuum pump No precipitation occurs in the bath Synthesis of doped, ternary, and quaternary compounds Does not need premium quality target Controlling on film thickness Avoid unnecessary heating Minimization of dislocation density by controlling deposition parameters ⁷² Fabricated stable and sticky films ³⁶	Failed to control film thickness Deposited films are contaminated though organic additives Opposite ions present in the reaction bath Perfect adsorption of ions requires on the substrate surface Substrate surface must be balanced completely through precursor solution ⁷³
ALD	Films thicknesses are under controlled Layer by layer film deposition Deposition can be performed at relatively low temperature Soft substrates can be used	Sluggish deposition process Highly refined substrate is required Instrument and substrate are highly priced
Spin coating	Ability to use thermally unstable precursors due to slow degradation ^{74,75} Dominated over film thicknesses Easy and effortless Affordable Fabricated stable cohered films Deficit of coupled variables Reproducibility ^{78,79}	Several trials required to set optimize film growth condition Process is restricted to non-volatile compounds Unfavorable for heat sensitive biological substrates ^{76,77} Hindered to use multiple substrates at a time Restricted to utilize big substrate Low material productivity Inexpensive with respect to photoresist and substrate size 95–98% materials are wasted ⁸⁰ Less effective in nanotechnology due to quick drying ⁸¹



Table 4 (Contd.)

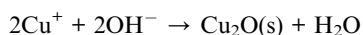
Techniques	Merits	Demerits
Solvothermal	Dominant over crystallinity of the deposited films Produces intermediate to specific quality films Potentiality to synthesize solid-state materials	Requirement of expensive autoclave Safety issues Impossible to observe <i>in situ</i> reaction process ⁸² Less dominance over particle size ⁸³



Overall reaction:



In the above reactions, $[\text{Cu}(\text{S}_2\text{O}_3)]^{-}$ the complex solution is regarded as the cationic precursor solution (cold solution) while NaOH is the anionic precursor solution, which is being kept at 70 °C (hot solution).⁹¹ When the substrate is immersed in the hot solution, OH^{-} ions are adsorbed onto the substrate and subsequently dipping into the cold solution results in the adsorption of Cu^{+} ions. Thus, one SILAR cycle is completed and Cu_2O thin film is formed due to the reaction between Cu^{+} and OH^{-} ions. Rinsing is carried out after every immersion to exclude loosely adhered particles. The number of cycles as well as dipping time varies based on required film thicknesses. Corresponding reactions are given below⁹² and the growth mechanism is schematically represented in Chart 1.



Therefore, a basic SILAR cycle comprises four different steps. The steps are represented in the following chart:

Consequently, a SILAR cycle covers four diverse steps on the surface, associating alternative immersion of the substrate into cationic and anionic precursor solution followed by rinsing in each immersion cycle to remove loosely adhered particles as demonstrated in Chart 2 and discussed below:

3.1 Adsorption

First SILAR stage forms the Helmholtz double layer owing to the initial adsorption of the cationic precursor such as Cu^{+} on the substrate surface. This layer is generally composed of two charged layers, the positively charged, Cu^{+} , inner layer and the negatively charged, $(\text{S}_2\text{O}_3)^{2-}$, outer layers.

3.2 Rinsing I

In the second stage, extra adsorbed ions, Cu^{+} and $(\text{S}_2\text{O}_3)^{2-}$, are rinsed away from the diffusion layer towards the bulk solution and a hypothetical monolayer is formed, resulting in a saturated electrical double layer.

3.3 Reaction

In the reaction step, the anions, OH^{-} , from the anionic precursor solution are entered into the scheme. A solid

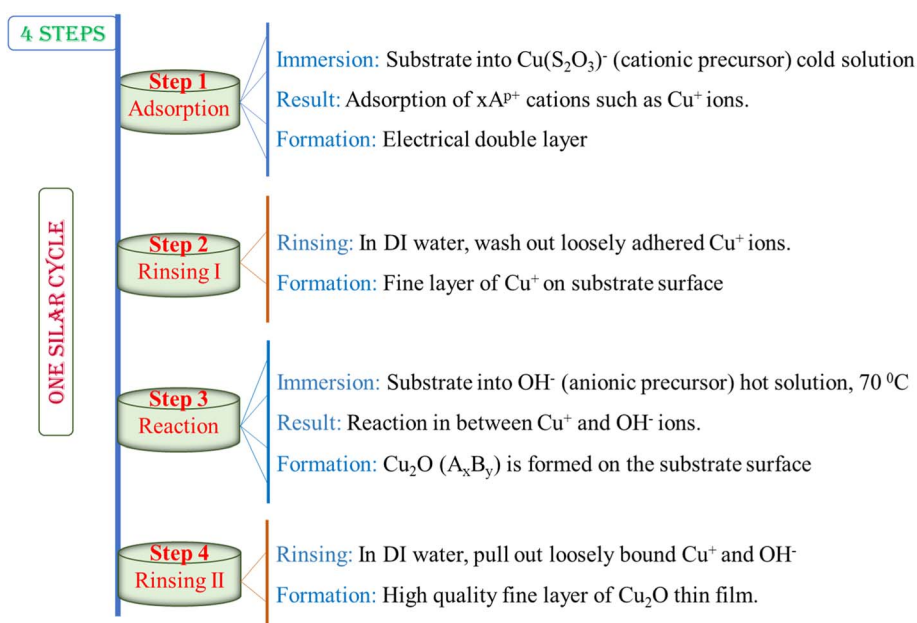


Chart 1 Representation of different steps during a SILAR cycle.



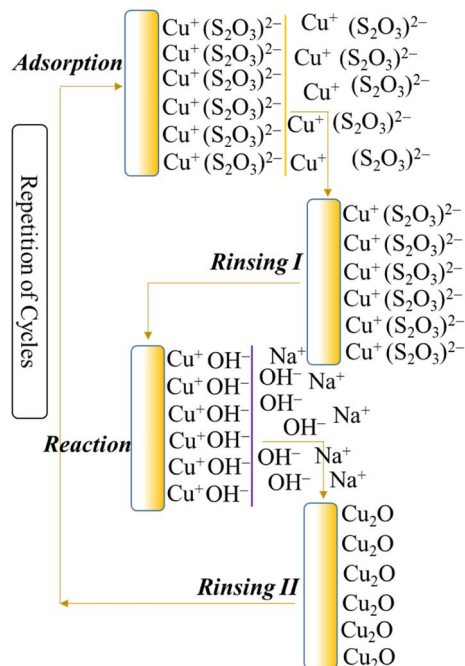


Chart 2 Schematic presentation of the deposited Cu_2O nanostructured films on the substrate surface during a SILAR cycle.

substance, Cu_2O , is synthesized on the interface due to the low stability of the material. This procedure pays the reaction of Cu^+ species with the anionic precursor such as OH^- .

3.4 Rinsing II

In the final SILAR cycle, the extra and unreacted species such as $(\text{S}_2\text{O}_3)^{2-}$, Na^+ as well as by-products of the reaction from the diffusion layer are removed leaving expected thin films.

The above deposition process involved alternate immersion of the substrate into cationic and anionic precursor solution followed by rinsing in every immersion cycle to eliminate loosely adhered particles.³⁶ Fig. 3 represents the synthesis of copper(i) oxide nanorod thin films in presence of NaCl using the SILAR deposition system.⁹³ Earlier to the film deposition, the colorless copper–thiosulfate complex was made ready by mixing 10 mL 1 M copper(II) sulfate and 40 mL 1 M sodium thiosulfate into a 100 mL volumetric flask. Then, in addition to DI water, the required amount of NaCl electrolyte was further added to the same flask and the produced complex solution was the cold solution. Meanwhile, 2 M NaOH solution was kept constant at 70 °C and treated as the hot solution. The substrate such as soda lime glass was then alternatively submerged in cold and hot solutions respectively for the required time interval and completed one SILAR cycle. To fabricate a thin film, this procedure was repeated for up to several immersion cycles.

The formation of Cu_2O nanorod thin films in presence of a NaCl electrolyte at various concentrations were discussed by using SEM micrographs as shown in Fig. 4. The film fabricated with no NaCl electrolyte demonstrated pencil-thin, and crack-free nanorod with an overgrown cluster in some areas on the substrate surface, as also detected in our earlier study.³⁷ When 2 mmol of NaCl of the electrolyte was introduced into the solution, the crowded nanorods were developed, and the formation of nanorods enhanced with the increase in the concentration of NaCl to 4 mmol, showing a larger size and shape as observed in the Fig. 4(c). Very rough, tiny and dense spherical grains as well as some overgrown clusters were seen with an additional increase in the concentration of NaCl to 6 mmol. Such an overgrown cluster was produced due to the coalescence of the particles.⁹⁴ Characteristically distributed, clear, and larger-sized spherical grains were revealed with further addition of NaCl electrolyte of 8 mmol. Thus, the NaCl

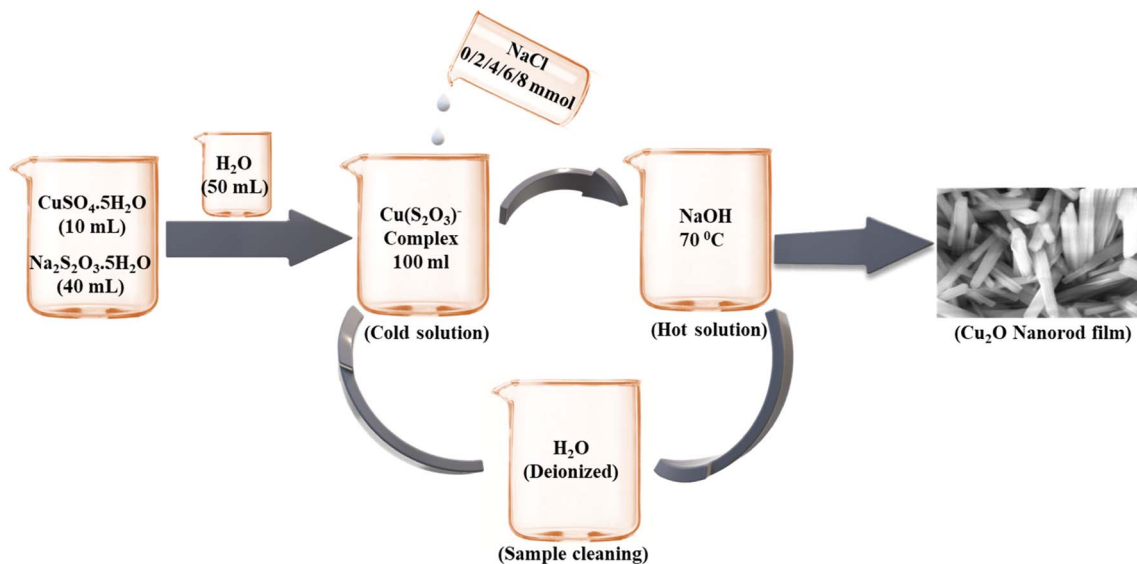


Fig. 3 Synthesis of the copper(i) oxide nanorod thin films.⁹³





Fig. 4 SEM images of the samples fabricated at (a) 0 mmol, (b) 2 mmol, (c) 4 mmol, (d) 6 mmol, and (e) 8 mmol of NaCl electrolyte.⁹³

electrolyte has the potential impact to change the surface morphologies from nanorods to spherical grains. The growth of Cu_2O nanorod thin films was sensitive to the concentration of salts added as also stated for CuO .⁹⁵ The growth of Cu_2O

nanostructures was increased gradually with the rise of NaCl concentration but until a limit. Such phenomena signify that NaCl concentration will consequence in similar morphology of the product and perform key roles in governing the size and

Table 5 Raw materials are used for the formation of Cu_xO thin films by the SILAR method

Formation	Cationic precursors	Complexing agents	Anionic precursors	Ref.
Cu_2O	$\text{Cu}(\text{CH}_3\text{COO})_2 \cdot \text{H}_2\text{O}$	$\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$	NaOH	96
	$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$	$\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$	NaOH	96
	$\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$	$\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$	NaOH	96
	$\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$	$\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$	NaOH	96
CuO	$\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$	NH_3	H_2O_2	97
	$\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$	NH_3	H_2O	98
	$\text{Cu}(\text{CH}_3\text{COO})_2 \cdot n\text{H}_2\text{O}$	$\text{NH}_4\text{CH}_3\text{COO}$	H_2O	99





Fig. 5 The FE-SEM images of Cu_2O nanostructured thin films deposited from four different Cu salts⁹⁶ (a) $(\text{CH}_3\text{COO})_2\text{Cu}\cdot\text{H}_2\text{O}$, (b) $\text{CuSO}_4\cdot 5\text{H}_2\text{O}$, (c) $\text{Cu}(\text{NO}_3)_2\cdot 3\text{H}_2\text{O}$, and (d) $\text{CuCl}_2\cdot 2\text{H}_2\text{O}$.

shape of the Cu_2O nanorods. Moreover, the steric hindrance caused by salt concentration might have affected the micelle aggregates, and these effects collected the assemblies of the products. More investigations are ongoing to elucidate the mechanisms for the growth process caused by the novel anticipated route.

4. Factors affecting SILAR deposition process

4.1 Types of copper salts

In the SILAR technique, Cu_xO thin films were studied and fabricated by using different copper salts as summarized in

Table 6 Properties of Cu_xO thin films deposited by varying solution pH applying SILAR method

pH	Precursors			Phase formed	Thickness (nm)	Crystallite (nm)	Band gap (eV)	Resistivity $\times 10^3$ (Ω cm)	Ref.
	Cationic	Anionic	To alter pH						
2.35	$[\text{Cu}(\text{S}_2\text{O}_3)]^-$	2 M NaOH ^a	H_2SO_4	Cu_2O	340	17	2.05	0.21	36
3.45		2 M NaOH ^a	CH_3COOH		729	21	2.10	0.18	
4.50		2 M NaOH	CH_3COOH		800	15–22	2.30	72	37
5.10		2 M NaOH	CH_3COOH		1000		2.28	103	
6.20		2 M NaOH	CH_3COOH		1800		2.43	742	
7.33		2 M NaOH ^a	—		1130	18	2.15	0.37	36
7.33		1 M NaOH	—		336	13	2.16	0.18	
7.95		2 M NaOH	—		1477	15–22	2.42	21.9	37
9.0	$[\text{Cu}(\text{NH}_3)_4]^{2+}$	H_2O	—	CuO	42	37.4	1.61	—	105
9.5		H_2O	—		67	22.4	1.49	—	
10.0		H_2O	—		85	22.9	1.49	—	
10.0		H_2O	H_2SO_4		520	14	2.17	6.5	94
10.5		H_2O	H_2SO_4		590	21	2.07	5.5	
11.0		H_2O	H_2SO_4		680	27	2.02	4.0	
11.5		H_2O	H_2SO_4		770	30	1.99	4.25	
12.0		H_2O	H_2SO_4		820	36	1.89	4.5	

^a OP = optimized precursor (2 M NaOH).





Fig. 6 SEM images representing $\{\text{Cu}_2\text{O}: (a)-(d)\}$ and $\text{CuO}: (e)-(g)\}$ thin films deposited by varying pH.^{36,105}

Table 5. Generally, most of the studies were done by using $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ as a basic salt for the formation of Cu_2O thin films, whereas $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ was used for the formation of CuO

thin films. The mechanism of both salts was discussed elsewhere in this article.

Altindemir *et al.* examined $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ with the other three different salts such as $(\text{CH}_3\text{COO})_2 \cdot \text{H}_2\text{O}$, $\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$ and





Fig. 7 (a) Variation of band gap with rinsing steps. (b) I - V characteristic curve for FTO/Cu₂O/ZnO heterojunction cell grown on FTO substrate at pH-7.95.¹⁰¹

CuCl₂·2H₂O to fabricate Cu₂O thin films. Field emission-scanning electron microscope (FE-SEM) photographs of the deposited Cu₂O thin films were demonstrated in Fig. 5(a)–(d). The deposited Cu₂O thin film using the (CH₃COO)₂Cu·H₂O salt demonstrated the cauliflower-like pattern having zero voids between the grains as seen from the images whereas, in the case of the CuSO₄·5H₂O salt, the grain size of the film showed the spherical form having no voids between the grains as in Fig. 5(a) and (b). The grain of the Cu₂O thin film deposited using the Cu(NO₃)₂·3H₂O salt revealed both cauliflower-like and rod shapes as in Fig. 5(c) whereas by means of the CuCl₂·2H₂O salt showed the cauliflower-like shape with discrete spaces between the grains as in Fig. 5(d).

On the other hand, CuCl₂·2H₂O salt was employed to fabricate CuO thin films using NH₃ as a complexing agent. But Chatterjee and co-workers utilized H₂O₂ as an oxidizing agent with CuCl₂·2H₂O and NH₃ solution to fabricate Cu₂O thin films instead of CuO thin films.⁹⁵ Similarly, Cu(CH₃COO)₂·nH₂O and NH₄CH₃COO solution could be the potential choice to fabricate CuO thin films.⁹⁹

4.2 pH of the precursor solution

Impact of solution pH on the properties of Cu_xO nanostructured thin films deposited by SILAR was studied in the pH scale range from 2.35 to 12 as shown in Table 6. The study was accomplished by controlling the pH of cationic and anionic precursor solutions by adjusting the additional acid and/or bases such as H₂SO₄, CH₃COOH, NaOH and NH₄OH.

To optimize the growth condition to fabricate the FTO/Cu₂O/ZnO heterojunction solar cell, Farhad and co-workers extensively studied the effect of pH in between 2.35 and 7.95.³⁶ During the study, the Cu₂O thin films have grown by slightly modifying the original SILAR method,¹⁰⁰ just by eliminating step 2 as shown in Chart 1 and named the technique as modified SILAR method. 10% H₂SO₄ and CH₃COOH were added dropwise into cationic precursor solution to adjust solution pH, as well as the concentration of anionic precursor (NaOH), was also varied (1–2 M). The addition of CH₃COOH into the optimized precursor solution (OP + CH₃COOH) improved the quality of the crystal having a larger crystallite size of 21 nm whereas H₂SO₄ played the opposite role. Strong H₂SO₄ etches

film thickness and it decreased with decreasing pH of the cationic precursor solution. From the SEM micrograph it is observed that the optimized solution (OP ~ 2 M NaOH) showed compacted and larger spherical grains ($D \sim 231$ – 259 nm) while the non-optimized solution (1 M NaOH) revealed irregular surface morphology with tiny grains ($D \sim 164$ nm) which is as shown in Fig. 6. Grain size and band gap decreased with decreasing pH of cationic precursor solution such as 259–164 nm and 2.16–2.05 eV respectively. The electrical resistivity varies in the range of 0.18–0.38 k Ω cm and among optimized solutions OP + CH₃COOH showed the lowest resistivity. The resistivity of the modified SILAR grown samples had (1–5) order magnitudes less than those deposited by Nair and Ristov's SILAR, and electrodeposition method.³⁶

To justify further the effects of CH₃COOH and rinsing steps, Farhad and co-workers again deposited thin films at pH 4.50–7.95 by adding CH₃COOH into cationic precursor solutions.^{37,101} It is seen that film deposited without any rinsing step showed the lowest band gap due to the high film thickness and *vice versa* which is shown in Fig. 7(a). pH 5.10 which was maintained by adding CH₃COOH exhibited a larger and densely packed grain size (~ 300 – 530 nm) compared to pH 7.95 where no use of

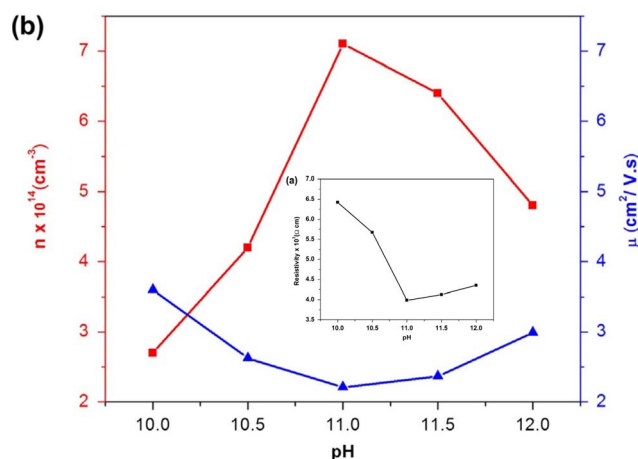


Fig. 8 (a) Electrical resistivity, (b) mobility and carrier concentration variation of SILAR-coated CuO thin films with pH.^{94,103}



CH₃COOH. Band gap and resistivity values were 2.41–2.30 eV and 742–72 kΩ cm respectively and decreased with decreasing pH of cationic precursor solution.³⁷ The activation energy was 0.004–0.19 and 0.01–0.68 eV respectively in the temperature range 40–90 and 100–250 °C. In these studies, FTO/Cu₂O/ZnO heterojunction shows diode-like characteristics¹⁰¹ at pH 7.95 as shown in Fig. 7(b) with maximum power (*P*) of ~45 mW cm⁻² under LED illumination with Knee voltage -0.5 V. These results indicate that CH₃COOH has a significant influence on the controlling of the properties of Cu₂O thin films.

Umeri and coworkers described the effect of pH and growth temperature during the deposition of Cu_xO thin films⁹⁴ at RT and 70 °C in the pH range from 8 to 11. At RT, the pure Cu₂O phase exists with pH 8, while mixed phases of Cu_xO appeared at pH 11. Whereas, at 70 °C, only the pure Cu₂O phase was deposited in the pH region of 8–11, which indicates that 70 °C is the optimum temperature for the growing phase of pure Cu₂O thin films as also supported by other reported results.¹⁰² At 70 °C, the optical band gap (*E_g*) rises from 1.85 to 2.0 eV with the rise of pH from 8 to 11, while the trend showed the opposite at RT and declines from 2.0 to 1.6 eV with the rise of pH. This might be because of the change in the composition from Cu₂O to CuO. From SEM micrographs, it is seen that at RT, the compact thin film was produced with pH 8 with an overgrown cluster in some spaces and when the pH increased to 11, overgrown cluster formation was diminished, and network-like nanofibers were observed. Conversely, at 70 °C and pH 9, a fiber-like nanostructure was formed, that looked like the morphology of films grown at RT with pH 11. Consequently, uniform, overgrown clusters free of close-packed and interconnected nanofibers of Cu₂O were observed at 70 °C and pH 11 with a band gap of 2.0 eV. Thus, temperature-dependent pH has a significant controlling overgrowth and properties of the deposited films.

Likewise, Cu₂O, the influence of pH on the physical properties of CuO thin films was investigated by Visalakshi *et al.*¹⁰³ The pH (~10–12) of the cationic precursor solution was maintained by adding concentrated NH₄OH. Film thickness, crystallite size, and texture coefficient rise with the rising pH of cationic precursor solution but dislocation density and strain decreases. SEM images concluded that pH 10 and 10.5 exhibited cluster-like surface morphologies due to the coalescence of the grains but when it reached pH 11, uniformly distributed spherical grains were observed. At pH > 11 the agglomeration of the grains occurs which outcomes in larger grain size. The optical transmittance and band gap (2.17–1.89 eV) reduces with increasing pH. The resistivity decreases initially from 6.5 × 10³ to 4.0 × 10³ Ω cm with increasing pH from 10 to 11, then further increases with increasing pH. As represented in Fig. 8(a), with the increase of solution pH from 11 to 12, the carrier concentration decreases from 7.1 × 10¹⁴ to 4.8 × 10¹⁴ cm⁻³ which is in good agreement with the obtained result by Saravanakannan *et al.*¹⁰⁴ Conversely, the mobility is first declined to pH 11; then, it is raised for further growth in pH. The decrease in mobility may be owing to the scattering formed at grain boundaries. The decrease in resistivity of the film synthesized at high pH may be attributable to the growth in film thickness without voids, whereas the rise in resistivity and decrease in carrier concentration and mobility detected at low pH may be owing to the existence of bulky voids. However, almost different properties were exhibited when the sample was annealed at 400 °C for 2 hours after deposition reported by Gençylmaz and coworkers.¹⁰⁵

4.3 Number of cycles during deposition

The film characteristic is closely related to the number of immersions of the substrates into the precursor solution. In

Table 7 Effect of the number of cycles on the fabrication of SILAR grown Cu_xO films

Product	Cycles	Thickness (nm)	Crystallite (nm)	Dislocation density ($\delta \times 10^{-3}$, nm ⁻²)	Strain ($\epsilon \times 10^{-3}$)	Bandgap (eV)	Ref.
Cu ₂ O	50	550	16	3.72	6.82	2.11	106
	75	830	19	2.90	6.01	2.06	
	100	1050	20	1.51	4.35	1.84	
	10	120	2.46	—	—	2.01	107
	30	350	5.23	—	—	2.48	
	40	520	7.15	—	—	2.53	
	20	654	—	—	—	2.48	
	40	1130	—	—	—	2.45	37
	60	1200	—	—	—	2.41	
	80	1477	—	—	—	2.38	
—	—	—	—	—	2.31		
CuO	30	500	11	8	3.12	1.56	108
	40	850	13	6	2.66	1.52	
	50	950	18	3	1.95	1.48	
	20	87	7	2.04	3.62	2.48	109
	30	179	8	1.56	3.73	2.41	
	40	298	9	1.23	3.83	2.37	
	50	415	11	0.83	4.87	2.31	
	30	—	10	10	3.43	1.92	110
	40	—	15	4.44	2.27	1.89	
	50	—	24	1.74	1.42	1.69	



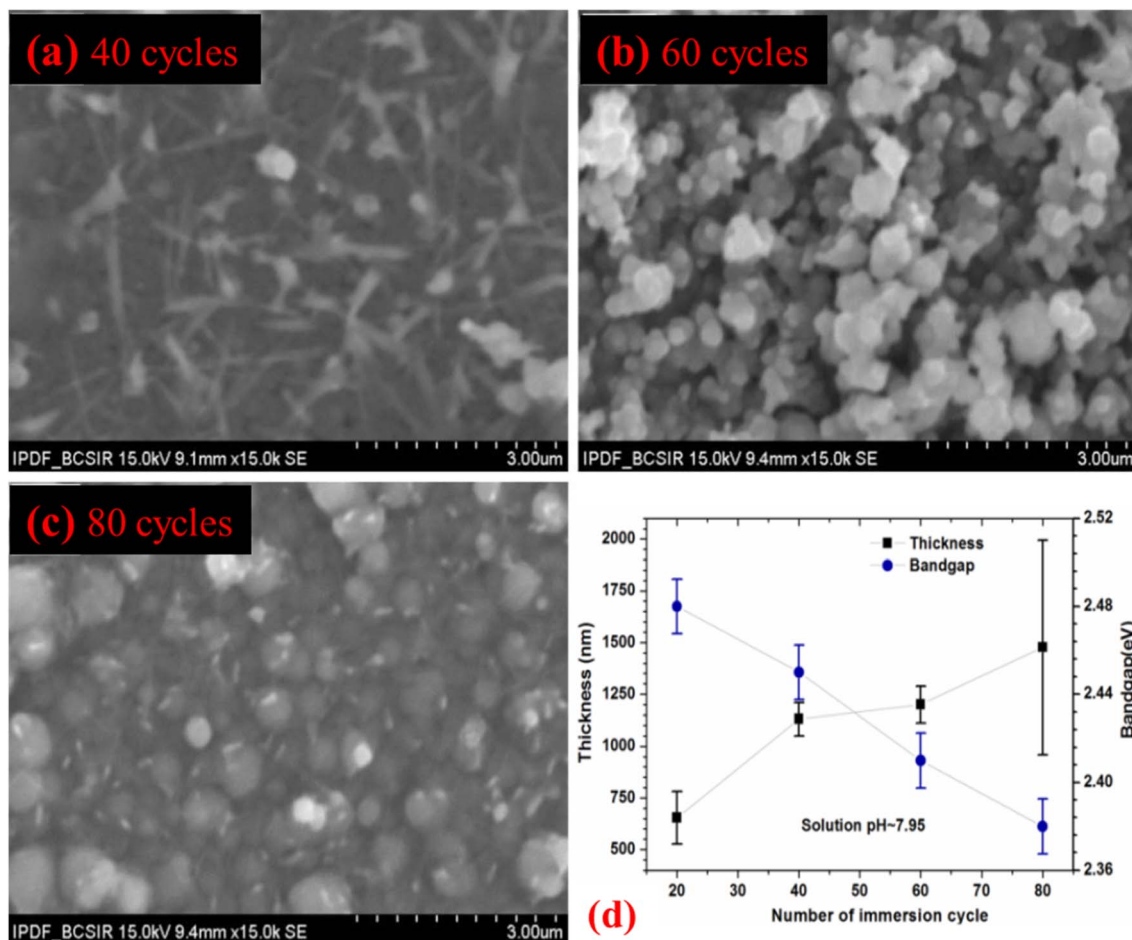


Fig. 9 SEM micrographs of the films grown with different immersion cycles at pH: 7.95 for (a) 40 cycles (b) 60 cycles (c) 80 cycles and (d) film thickness of Cu_2O thin films using different immersion cycles grown by m-SILAR.³⁷

almost all studies, the number of cycles is generally kept constant to understand the other properties of the Cu_xO films. There are few investigations where the effect of the number of cycles on the fabrication of Cu_xO films^{106–110} is discussed, as summarized in Table 7. It is obvious that with the increase of immersion cycles thickness of the films increase.

The surface SEM morphologies of m-SILAR deposited Cu_2O films in the top of the FTO substrates using 40, 60, and 80 immersion cycles were shown in Fig. 9(a)–(c). Throughout the area investigated, the surface morphology of all films was seen to be compact as well as coherently carpets. However, the Cu_2O film grown with 40 immersion cycles demonstrated fiber-like microstructures with small grains of around 200 nm. Instead, thin films having 60 and 80 immersion cycles exhibited bigger spherically shaped grains of size around (200–550) and (350–650) nm respectively as shown in Fig. 9(b) and (c). This reflection recommends that grain size develops as the thickness increases with the increase of immersion cycles.³⁷ As can be seen from Table 7 and Fig. 9(d), the optical bandgap of the Cu_2O thin films deposited at pH \sim 7.95 using 20 cycles (film thickness \sim 654 nm), 40 cycles (film thickness \sim 1130 nm), 60 cycles (film thickness \sim 1200 nm) as well as 80

cycles (film thickness \sim 1477 nm) were calculated to be \sim 2.48, \sim 2.45, \sim 2.41 and \sim 2.38 eV respectively. Obviously, there is a decreasing trend of optical bandgap with increasing film as represented in Fig. 9(d), probably due to the bigger grains usually existing in the thicker films, which verifies the results stated by Nair *et al.*¹¹¹

Fig. 10(a)–(c) demonstrates the Cu_2O nanostructured films fabricated with different dipping cycles of the nanorods spread homogeneously on the substrate surface, showing a large number of grains with fine particle edges. As seen from morphological studies and Table 7, 50 cycles grown sample has a smaller crystallite size and higher dislocation density with good nanorod morphology. For light absorption, although it has an opportunity for a larger surface area to the photoelectrode, due to the presence of considerable grain boundaries, it creates recombination problems in the film. So, the electron trapping at the surface and in the intergrain boundaries lowered the efficiency value of the film grown through 50 cycles. The samples deposited by 75 and 100 cycles have comparatively better crystallite size and lower dislocation density, which leads to reduce grain boundaries.¹⁰⁶ Due to the drop in grain boundary resistance, the photogenerated





Fig. 10 SEM image of the Cu_2O nanostructured films deposited at (a) 50 (b) 75 and (c) 100 dipping cycles.¹⁰⁶

charge carriers can significantly reduce the recombination losses. The cell was fabricated as ITO/ZnO NRs/ Cu_2O /Al with varied efficiency mainly due to the number of cycles of the films shown in Fig. 11. Even though the attained efficiency of the ZnO/ Cu_2O heterojunction was lower, the efficiency was high in the samples deposited at high cycles such as 100. Hence, the effect of the film thickness on cell performance was evidenced by the enhancement of efficiency due to the substantial development of crystallinity and absorbance of Cu_2O films.

4.4 Effect of bath temperature

Fig. 12(a) illustrates the deposition of the thin films grown by varying bath temperatures of anionic precursors as a function of the immersion cycles. The figure demonstrated that the fabrication rate reduced as the immersion cycle proceeds characteristically at 10 nm per cycle. In the case of fabrication using the alkali solution at 90 °C, the production was faster, and the film thickness was $>0.3 \mu\text{m}$ with 20 cycles of immersions, whereas, at 70 °C, the film growth slightly falls after 20 immersions.¹¹¹ For the films fabricated with NaOH solution temperature of 50–90 °C, the photo response curves were given for a range of thicknesses as demonstrated in Fig. 12(b). Irrespective of the solution temperature, the dark current and the

photocurrent logged in the films were comparable for the films with thicknesses smaller than 0.1 mm. The values were higher in samples fabricated using NaOH solution at 70 °C having films of higher thickness. The measured electrical conductivity of a 0.15 μm film is about $5 \times 10^{-4} \Omega^{-1} \text{cm}^{-1}$. And it was found

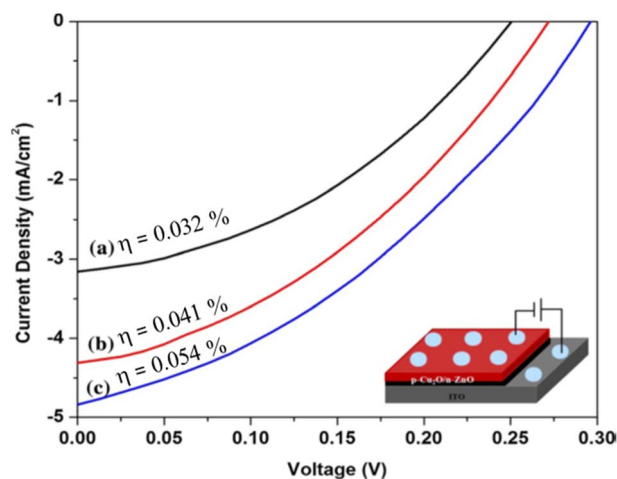


Fig. 11 Current–voltage curve of ITO/ZnO NRs/ Cu_2O /Al cells (a) 50 (b) 75 and (c) 100 dip Cu_2O films.¹⁰⁶



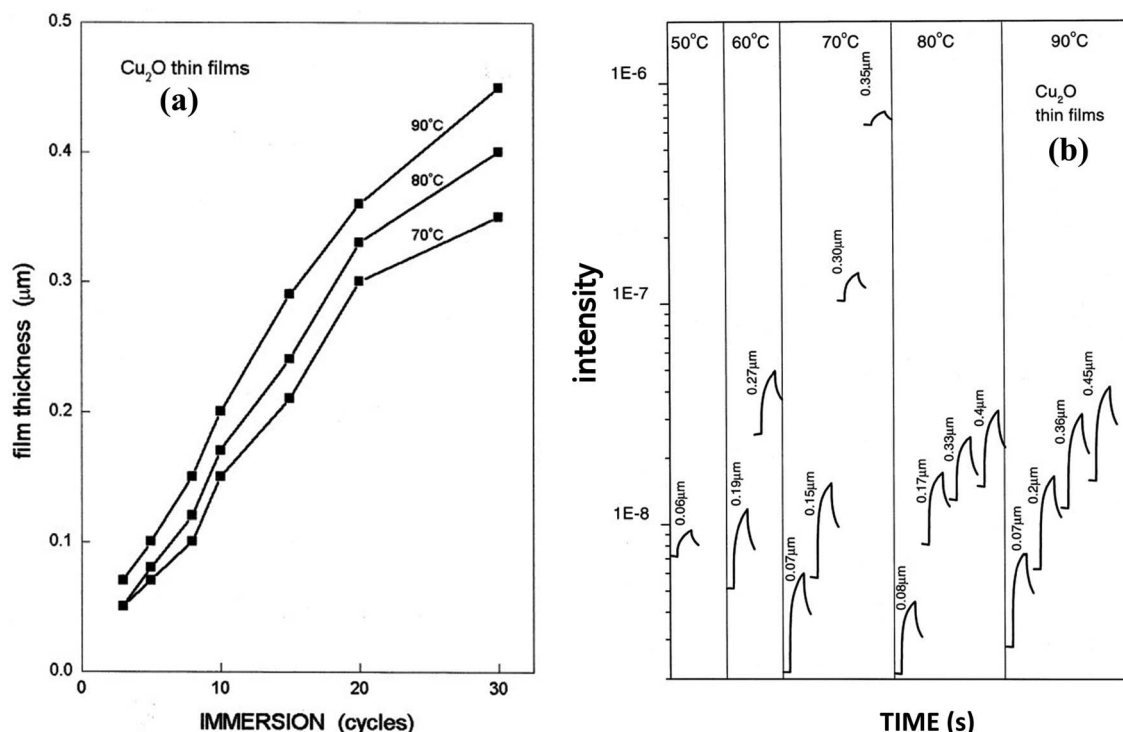


Fig. 12 Samples in NaOH solution (a) at different temperatures as a function of film thickness and immersions. (b) At the designated temperatures to record the photocurrent response of the Cu₂O thin films of different thicknesses.¹¹¹ [Light source: tungsten-halogen, intensity of illumination: 1 kW m⁻², time length: 60–180 s, bias: 1 V applied across Ag print electrodes 5 mm (long) × 5 mm (separation)].

that the increase of film thickness of two orders increases the conductivity by nearly two orders.

The structural parameters, elemental composition, and optical band gap for different bath temperature of Cu₂O films are given in Table 8, studied by Baig *et al.*¹¹² It is seen from the table that when the temperature climbed from 40 to 80 °C, the size of grain increased from 16.78 to 18.84 nm whereas strain in the crystal lattice was reduced. The fall in strain signifies that the imperfection in the crystal lattice with the rising temperature was decreased. The SEM images of Cu₂O thin films deposited on ITO substrate with different anionic bath temperatures are demonstrated in Fig. 13. From the figures, with the rise of anionic bath temperature the structure of the wire became compact compared to that at 40 °C. Likewise, the oxygen concentration was decreased with an increase in temperature as observed in Energy-dispersive X-ray spectroscopy (EDS) value of Cu₂O film in Table 8. Further, the photocatalytic activity for water splitting by the deposited Cu₂O thin

films at different temperatures was studied in a photochemical cell and the result revealed that film grown at 80 °C had a higher current ratio with respect to the other two samples and the photocurrent produced by that sample is relatively steady (figure in 5.2 section).

4.5 Addition of additives

The influence of the different additives on the surface morphological characteristics of CuO films was studied by using SEM. Fig. 14 illustrates the SEM images of the CuO thin films fabricated in the solution containing additives such as coumarin, saccharin, and sodium dodecyl sulfate (SDS) having different concentrations. In the first step without coumarin, Fig. 14(a) there were plate-like CuO nanostructures that homogeneously cover the entire surface. Then, the nanostructures start to change their shapes with the increase of coumarin concentration, form some clusters on the surface and lose their homogeneity. From Fig. 14(b), in the case of

Table 8 Structural parameters, elemental composition (EDS), and optical band gap for different bath temperatures of Cu₂O films

Bath temperature (°C)	D (nm)	Dislocation density (δ) × 10 ¹⁴ m ⁻²	Micro strain (ϵ)	Atomic (%)		Band gap (eV)
				Cu	O	
40	16.78	35.51	0.007025	50.75	49.25	2.25
60	17.10	34.19	0.006892	56.45	43.55	2.14
80	18.84	28.17	0.006193	61.76	38.24	2.07





Fig. 13 SEM images of Cu_2O nanostructured thin films deposited by SILAR at different bath temperatures (a) 40 °C, (b) 60 °C, and (c) 80 °C.¹¹²

saccharin, it was observed that all the samples have nearly the same morphology that is they are all composed of plate-like nanostructures as in ref. 113. On the other hand, the homogeneity of the films deteriorates with increasing saccharin concentration. In the case of SDS, Fig. 14(c) demonstrates that the fabricated CuO thin films were adhered and spread well onto the substrate surface without SDS. With the addition of SDS and its concentration, the surface morphology of the CuO film was changed dramatically. From this viewpoint, the discrepancy of SDS molar concentration impressed the morphology of the surface of all the fabricated thin films. This variation in morphology may be owing to the electrostatic interaction between Cu^{2+} and $\text{CH}_3(\text{CH}_2)_{11}\text{OSO}_3^-$. SDS may affect particle growth as well as morphology after nucleation. Thus, during the crystallization process, the SDS can affect nucleation.¹¹⁴

Through UV-Vis's spectrophotometer study, it was clear that both the optical band gap and the transmission spectra were affected by the additive concentration. The band gap, as well as spectral transmittance values of the films, were decreased for the higher content of both coumarin and saccharin,^{115,116} while showing the opposite tendencies in the case of SDS. The optical bandgap energy of both organic (coumarin and saccharin) additives decreased from around 1.50 to 1.27 eV, while increased from 1.32 to 1.49 eV for inorganic SDS, with the increasing concentration of the additives.¹¹⁷

4.6 Complexing agent

Cavusoglu and co-workers studied the role of the complexing agent such as triethanolamine (TEA) mediated fabrication of nanocrystalline CuO thin films *via* SILAR technique at room temperature and the results are summarized in Table 9. As a function of increasing TEA concentration, the optical band gap energy of the fabricated CuO thin films was increased from 1.33 to 2.00 eV while the average transmittance of all the films increased from 2.5 to 42.5%. A minimum resistivity of $3.74 \times 10^5 \Omega \text{ cm}$ was found with zero TEA in CuO thin films whereas, with a TEA concentration of 1.0 M%, the resistivity subsequently increased to $509 \times 10^5 \Omega \text{ cm}$. Surface morphology on the film surfaces demonstrated the homogeneous distribution of the nanostructured CuO as demonstrated in Fig. 15(a)–(d) whereas, the

figure of merit (FOM) was represented as a function of TEA concentration as shown in Fig. 15(e). TEA concentration of 0.25 M% in CuO thin film provided the high FOM values of $786 \times 10^{-12} \Omega^{-1}$ at distinct wavelengths of between 600 and 900 nm.¹¹⁸ Therefore, the range of optical and electrical properties developed by such a study having a different complexing agent is favorable for the applications of numerous optoelectronic devices.

4.7 Annealing of as-deposited films

Annealing is a vital parameter to control the phases of the deposited thin films. Both the phases of Cu_xO could be synthesized by changing the atmospheric condition (air, vacuum) and temperature of annealing as summarized in Table 10. Here, air or oxygen^{37,101,119–124} annealing of the SILAR-grown films has been studied more extensively than vacuum annealing.¹²⁵ Recently, SILAR deposited Cu_xO films are mainly studied between 75 to 500 °C^{37,101,119–124} in presence of air or N_2 . The study revealed that Cu_2O phase was stable until 250 °C,^{120,122} though Farhad and co-workers showed a mixed phase of both Cu_2O and CuO at 250 °C³⁷ due to the pH effect, while Ozaslan *et al.* showed a mixed phase even at 500 °C.¹²³ The CuO phase could be found at 300 °C¹²⁰ by annealing of Cu_2O or even could be deposited by using NH_3 solution (pH = 10) with the reaction of CuCl_2 at ambient temperature.¹²⁴

Amudhavalli and co-workers successfully showed the increasing trend of the crystallite size of copper oxides with the increase of annealing temperature while depositing the films at 0.5 M NaOH. Fig. 16 demonstrated the change of resistivity, mobility, and carrier concentration of copper oxide (Cu_2O and CuO) phases with annealing temperature as shown by Ozaslan *et al.*¹²³ It is found that the carrier concentration was decreased from 3.07×10^{17} to $6.61 \times 10^{15} \text{ cm}^{-3}$ with increasing annealing temperature from 70 to 500 °C respectively. The hole mobility of the films was increased from 4.20 to $31.87 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ with decreasing the carrier concentration, while the electrical resistivity of the films decreased with annealing temperature, inducing the increment in the conductivity of the films. Nair *et al.* observed the dark conductivity of the CuO film produced by air annealing of a 0.15 μm Cu_2O film at 400 °C is high about $7.2 \times 10^{-3} \Omega^{-1} \text{ cm}^{-1}$.





Fig. 14 SEM images of CuO nanostructured thin films as a function of different additive concentrations, such as (a) coumarin,¹¹⁵ (b) saccharin¹¹⁶ and (c) SDS.¹¹⁷

4.8 Doping by diverse dopants

Tuning of the structural, electrical, and optical properties of SILAR-deposited Cu_xO films through Fe, Eu, Zn, Co, B, Mg, Ni,

and Pb doping has been reported extensively by several authors. The summary of the effect of doping on SILAR-deposited Cu_xO films is represented in Table 11. Interestingly, in the case of

Table 9 Properties of CuO thin films as a function of TEA concentrations

TEA concentration M%	Crystallite (nm)	Thickness (nm)	Bandgap (eV)	Conductivity (σ) $\times 10^{-6} (\Omega \text{ cm})^{-1}$	Resistivity (ρ) $\times 10^5 \Omega \text{ cm}$	FOM \times $10^{-12} \Omega^{-1}$
0	19.95	797	1.33	2.67	3.74	149
0.25	19.80	387	1.57	1.05	9.50	786
0.50	18.92	199	1.67	0.07	149	37.2
1.00	17.47	101	2.00	0.02	509	1.70



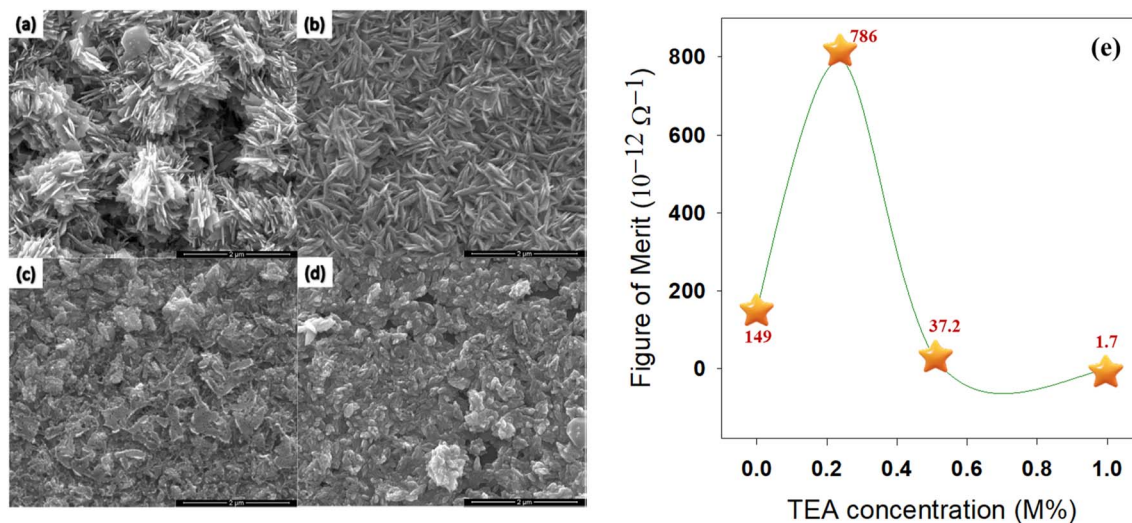


Fig. 15 SEM images of CuO thin films at (a) 0 M%, (b) 0.25 M%, (c) 0.50 M%, (d) 1.00 M% as well as (e) FOM with TEA at different molar concentrations.¹¹⁸

Cu₂O film fabrication, the reactants were the same except for the doping materials such as Fe, Eu, Zn, and Co.^{127–130} Similarly, during Co, B, Mg, and Pb doping into CuO films,^{131–134} the reactants were also the same except in the case of Ni doping.¹³⁵

In the case of Co or Fe doping, the crystallite size of the films of Cu₂O decreased between 62.83 and 28.44 nm when the concentration of the doped material increased gradually, whereas it showed an opposite trend in the case of Zn doping. Interestingly, Fe, Eu, Zn and Co doping into Cu₂O rises the bandgap of the material a little in every case. On the other hand, in the case of Co and B doping into CuO films, the bandgap decreased while it was increased for Mg and Pb doping.

The films prepared at high doped Cu₂O thin films such as 5% Eu showed a low resistivity value of $1 \times 10^3 \Omega \text{ cm}$ as shown in Fig. 17. The Hall mobility and carrier concentration values in such cases are $0.52 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ and $13.8 \times 10^{15} \text{ cm}^{-3}$, respectively.

Fig. 18(a) shows the current density–voltage (J – V) characteristics of the ZnO/Cu₂O heterojunction solar cells prepared using the Eu-doped Cu₂O thin films. The V_{oc} was increased with increasing Eu content from 265 mV (1% Eu) up to 332 mV (5% Eu). The conversion efficiency can be enhanced by dropping recombination centers avoiding lattice-mismatch defects, and by reducing the resistance of Cu₂O. The ionic radius of Eu³⁺ ion was 0.109 nm whereas, Cu⁺ is 0.077 nm. Therefore, Eu³⁺ ion could not be incorporated by substitution rather it was incorporated as an interstitial creating getter center. It overwhelms the recombination losses and thus advances current levels and improved Eu doping levels.¹³⁶

Fig. 18(b) illustrates the band structure as well as carrier transport of the deposited p–n junction. As there was much difference between conduction and valence band off-sets triggering effective separation of charge carriers, a built-in potential barrier was developed. When the light was absorbed onto the device photocarriers were generated and drifted to the respective electrodes depending upon the applied potential causing

current conduction. As an acceptor dopant, impurity levels of Eu were adjacent to the valence band edge. In the case of ZnO, the green luminescence at 535 nm could be produced by the diffused Cu ion and replacing Zn. The V_{O} center was atop the valence band whereas the Zn vacancy was (V_{Zn}) in an acceptor level, which occurred at 0.8 eV. Nevertheless, the ZnO coated over Eu: Cu₂O performed as a passivation layer improving the V_{oc} and declining the consequence of impurity center-mediated recombination loss.^{137,138}

Magnetic measurements were performed by employing a vibration sample magnetometer (VSM) at ambient temperature for both Fe and Co-doped Cu₂O. Undoped Cu₂O has a diamagnetic property.¹³⁹ The outcome agrees with Fig. 19(a) and (b) which demonstrate the change of magnetization against the applied magnetic field (M – H). In the case of Co-doped Cu₂O, undoped and minimum doped such as 1 and 2 wt% films showed diamagnetic (high magnetization) behavior whereas, at the maximum doped such as 10 wt%, the films showed ferromagnetic (low magnetization) properties.¹⁴⁰ The diamagnetic order was Cu_{2–x}Co_xO ($x = 0 > 1 > 2 > 5 > 10 \text{ wt}\%$). The ferromagnetic behavior was possibly due to the intrinsic coupling (Co–Co) between the atoms of doped material. Similarly, in the case of 1% Fe doped Cu₂O at 305 K, the film showed diamagnetic properties. An increase of the Fe-doping (2 wt%), slightly altered the diamagnetic property because of the increased hole concentrations and further doping of Fe ions (5 wt%), the film showed anti-ferromagnetic behavior. With the increase in the concentration of Fe, both the number of Fe³⁺–Cu²⁺ pairs and the hole concentrations increased and consequently, the crystallite size reduced.

Lobinsky and co-workers studied the cyclic voltammograms of the nickel foam electrode with Ni-doped CuO nanolayers in a potential space between 0 and 550 mV vs. Ag/AgCl electrode at the scanning rates of 5, 10, 15 and 20 mV s^{–1} as shown in Fig. 20. Two of the redox reactions on the anodic curve took place in the layer, including the Cu⁺ → Cu²⁺



Table 10 Annealing of the SILAR grown Cu_xO nanostructured thin films

Anionic salt, NaOH (M)	Time			Temperature		Crystal phase	Crystallite (nm)	Band gap (eV)	Ref.
	Cycles (s)	Dipping (s)	Annealing (min)	Growth ($^{\circ}\text{C}$)	Annealing ($^{\circ}\text{C}$)				
$\text{CuSO}_4 \cdot 5\text{H}_2\text{O} + \text{NaOH} + \text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$									
0.5	20	20	—	70	200–400	As deposited: Cu_2O 200 $^{\circ}\text{C}$: Cu_2O 300 $^{\circ}\text{C}$: $\text{Cu}_2\text{O} + \text{CuO}$ 400 $^{\circ}\text{C}$: CuO	27.76 49.95 40.88 62.32	—	119
1	10	20	60	70	200–350	As deposited: Cu_2O 200 $^{\circ}\text{C}$: Cu_2O 250 $^{\circ}\text{C}$: Cu_2O 300 $^{\circ}\text{C}$: CuO 350 $^{\circ}\text{C}$: CuO	14 14 14 14–26	2.20 2.20 2.20 1.35	120
1	30	20	—	70	200–400	As deposited: Cu_2O 200 $^{\circ}\text{C}$: Cu_2O 300 $^{\circ}\text{C}$: $\text{Cu}_2\text{O} + \text{CuO}$ 400 $^{\circ}\text{C}$: CuO	14–26 2.40 2.40 2.06	2.40 2.40 1.73	121
1	30	20	—	50–90	250–400 (air, N_2)	As deposited: Cu_2O 250 $^{\circ}\text{C}$: Cu_2O 300 $^{\circ}\text{C}$: $\text{Cu}_2\text{O} + \text{CuO}$ 350 $^{\circ}\text{C}$: CuO 400 $^{\circ}\text{C}$: CuO	~18 2.10 — 1.75 1.75	2.10 2.10 — 1.75	122
2	40	2	60	70	100–500	As deposited: Cu_2O 100 $^{\circ}\text{C}$: Cu_2O 300 $^{\circ}\text{C}$: Cu_2O (2.27%) + CuO (97.73%)	— 2.57 2.52	2.57 2.52 2.45	123
2	40–80	—	60–180	70	75–350	500 $^{\circ}\text{C}$: Cu_2O (1%) + CuO (99%) As deposited: Cu_2O 75 $^{\circ}\text{C}$: Cu_2O 150 $^{\circ}\text{C}$: Cu_2O 200 $^{\circ}\text{C}$: Cu_2O 250 $^{\circ}\text{C}$: $\text{Cu}_2\text{O} + \text{CuO}$ 350 $^{\circ}\text{C}$: $\text{Cu}_2\text{O} + \text{CuO}$ (1 h) 350 $^{\circ}\text{C}$: CuO (3 h)	15–22 2.42 2.02 1.98 1.94 1.62 — 1.44	1.91 2.42 2.02 1.98 1.94 1.62 — 1.44	37
2	60	5	60	70	350	As deposited: Cu_2O 350 $^{\circ}\text{C}$: $\text{Cu}_2\text{O} + \text{CuO}$ (1 h)	— 2.06–2.16 1.43–1.51	2.06–2.16 1.43–1.51	101
$\text{CuCl}_2 + \text{NH}_3\text{OH} + \text{H}_2\text{O}_2$									
—	2–10	30	30	RT	20–500	As deposited: Cu_2O 100 $^{\circ}\text{C}$: Cu_2O 150 $^{\circ}\text{C}$: Cu_2O 450 $^{\circ}\text{C}$: CuO 500 $^{\circ}\text{C}$: CuO	14 14 14 16 16	2.17 2.22 2.17 1.43 1.44	124
—	400	>30	60	RT	27–600 (air, vacuum)	As deposited: Cu_2O 100 $^{\circ}\text{C}$: Cu_2O 400 $^{\circ}\text{C}$: CuO 600 $^{\circ}\text{C}$: CuO	14–21 2.40 1.85 1.70	2.30 2.40 1.85 1.70	125
$\text{CuCl}_2 + \text{NH}_3$ solution (pH = 10)									
—	80	30	30	—	200–400	As deposited: CuO 200 $^{\circ}\text{C}$: CuO 300 $^{\circ}\text{C}$: CuO 400 $^{\circ}\text{C}$: CuO	11.09 12.05 13.86 14.88	1.17 1.29 1.30 1.36	126

transformation at 310 mV while the $\text{Ni}^{2+} \rightarrow \text{Ni}^{3+}$ at 390 mV at a scan rate of 5 mV s^{-1} . The proportionality of currents to scan rate delivers data that the film is sufficiently thick, and the charge transfer rate was restricted by the diffusion of charge carriers in the film.¹³⁴

Inset of Fig. 21 demonstrates the specific capacitance of the Ni-doped CuO nickel foam electrode, which was found from charge–discharge curves, and it was 154 mA h g^{-1} (1240 F g^{-1}) at the current densities of 1 A g^{-1} .¹³⁵ The high value of the specific capacitance of the sample can be explained based on the good conductivity of CuO and the substantial role of Ni atoms in





Fig. 16 Representation of the change of resistivity, mobility, and carrier concentration of copper oxide (Cu_2O and CuO) films with annealing temperature.¹²³

pseudo-capacity. The capacity retention of the Ni foam electrode with Ni-doped CuO nanolayers after 1000 charge–discharge cycles at a current density of 2 A g^{-1} was retained at 92%, showing good cycling stability of the material as presented in Fig. 21. High cycling stability could be described by the feature morphology of ultrathin nanocrystals of CuO which deliver fast diffusion of ions on the electrode surface and while not being ruined in the charge–discharge process.

5. Applications

The optoelectronic properties of SILAR synthesized thin films have shown outstanding performance in diverse applications, for instance, photovoltaics,¹⁴¹ supercapacitors,^{142,143} photoelectrochemical water splitting,¹⁴⁴ gas sensors^{143,145} and many more. The method appears to be easier and represents an efficient way to manufacture devices. Some of the potential applications such as antibacterial activities, supercapacitors, surface wettability and photoelectrochemical water splitting in presence of Cu_xO nanostructured thin films will be discussed in the following section.

5.1. Antibacterial activities

To control pathogens, nanoparticles are in great demand due to their huge applications in the health industries. Results achieved from nanocrystalline Cu_2O nano-thin films fabricated by Dhanabalan *et al.* possessed substantial antimicrobial activity against the experienced human pathogen at a maximum inhibition zone of 16 mm against Gram-positive *Staphylococcus aureus*.¹⁴⁶ The surface morphological studies exhibited that the needle-shaped grains which play a crucial role in the antibacterial activity of the fabricated Cu_2O films by SILAR technique as shown in Fig. 22(a) and (b). The synthesized Cu_2O thin film can exhibit antibacterial activity from 18 to 24 hours of incubation time. The bacterial growth will decrease with the increase in the concentrations of nanoparticles, which may be the cause of the

reduction of voids affording space for the growth of bacteria that remains resistant to the pathogenic bacterial strain.

5.2. Water splitting

Cu_xO was considered a good candidate for photoelectrochemical (PEC) water splitting due to its abundance, low price, and high stability in aqueous solution.^{147,148} Baig *et al.* fabricated Cu_2O at a high bath temperature of $80 \text{ }^\circ\text{C}$ by SILAR which showed high photocurrent and good stability¹¹¹ as discussed earlier. The photocatalytic activity for water splitting of the Cu_2O thin film was studied with the photochemical system containing Pt (counter), Ag/AgCl (reference), Cu_2O (working electrode) and KCl ($\text{pH} = 13.6$) as electrolytes. PCE data shown in Fig. 23 revealed that samples synthesized at $80 \text{ }^\circ\text{C}$ have a higher current ratio and produced a stable photocurrent compared to the other samples by using a 300 W Xenon lamp (PLSSXE300/300UV).

5.3. Surface wettability study

The surface wettability study of films determines its capability to interact with ions when immersed into electrolyte by measuring the contact angle with liquid electrolyte as shown in Fig. 24.¹⁴⁹ If the contact angle is $>90^\circ$, then the film surface is hydrophobic, while for $<90^\circ$, it is hydrophilic. For better interaction of electrolyte ions, the contact angle must be as low as possible with the electroactive site on the thin film surface.^{150,151} Fig. 24 (A1, A2, A3, and A4) signifies the image of the contact angle with the surface of the film. The observed angles of CuO thin films with 50, 60, 70 and 80 SILAR cycles were 65° , 58° , 50° , and 43° , respectively. The observed CuO films were hydrophilic in nature, as the contact angles for CuO decline with the rise in SILAR cycles, which will allow more interaction of electroactive sites of the CuO on the film surface.

5.4. Super capacitive behavior

The electrochemical impedance, as well as super capacitive properties, of SILAR synthesized CuO thin films are studied by Patil *et al.*¹⁴⁹ The synthesized CuO thin film showed the lowest charge transfer resistance of $41.45 \Omega \text{ cm}^{-2}$ with the highest specific capacitance of 184 F g^{-1} at the scan rate of 50 mV s^{-1} and demonstrated 83% capacitive retention after 5000 cycles. Super capacitive performance of the film was verified using cyclic voltammetry (CV) in 1 M KOH electrolytes in a three-electrode cell equipped with CuO (working electrode), Pt (counter electrode) and saturated Ag/AgCl (reference electrode). As shown in Fig. 25, the CVs were studied with a potential window of 0 to 0.6 V/ Ag/AgCl at several scan rates such as 10, 20, 50 and 100 mV s^{-1} .

To examine the charge–discharge properties of CuO , the chronoamperometry technique was applied. Fig. 26(a) demonstrates galvanostatic charge–discharge curves at various current densities for CuO and signify a good capacitive behavior of CuO electrode as ref. 152. In Fig. 26(b), the difference of specific capacitance with various scan rates was displayed, which enhanced exponentially with decreasing scan rate.¹⁵³ The electrochemical stability of CuO film electrode was examined by applying



Table 11 Cu_xO nanostructured materials grown in presence of different dopants with their properties

Product	Dopant		Cycle	Crystallite (nm)	Grain (nm)	Bandgap (eV)	Ref.	
	Material	Amount						
Reactants: CuSO₄·5H₂O + NaOH + Na₂S₂O₃·5H₂O								
Fe: Cu ₂ O	FeSO ₄ (wt%)	0	30	62.83	—	1.80	128	
		1		59.80		2.10		
		2		41.83		2.36		
		5		36.40		2.45		
Eu: Cu ₂ O	Eu(NO ₃) ₃ ·5H ₂ O (at%)	1	100	27	—	2.08	129	
		3		24		2.26		
		5		21		2.41		
		10		18		2.34		
Zn: Cu ₂ O	ZnSO ₄ (wt%)	0	50	18	—	2.34	130	
		1		30		2.35		
		2		39		2.37		
		3		52		2.38		
		5		69		2.41		
		10		44		2.39		
Co: Cu ₂ O	CoSO ₄ (wt%)	0	30	62.83	—	1.94	127	
		1		53.30		2.03		
		2		48.47		2.12		
		5		39.24		2.18		
		10		28.44		2.47		
Reactants: CuCl₂·2H₂O + H₂O + NH₃								
Co: CuO	CoCl ₂ ·6H ₂ O (at%)	0	10	22.7	70	1.53	132	
		0.5		15.7		44		1.47
		1		13.6		42		1.45
		2		13.1		36		1.41
		3		12.6		32		1.38
B: CuO	H ₃ BO ₄ (at%)	4	10	12.2	38	1.36	131	
		0		12.9		45		1.52
		1		13.1		42		1.48
		2		14.2		38		1.43
Mn: CuO	Mn(NO ₃) ₂ (at%)	3	10	15.9	30	1.39	133	
		0		—		9.94		1.42
		1		—		7.83		1.98
Pb: CuO	Pb(NO ₃) ₂ (at%)	3	10	—	9.94	8.21	135	
		5		—		9.76		2.20
		1		—		9.94		1.43
		2		—		17.22		1.80
		4		—		16.21		1.76
		8		—		15.79		
		4		—		15.79		1.72
		8		—		13.07		1.68
16	—	8.98	1.65					
Reactants: Cu(CH₃COO)₂·nH₂O + H₂O + NH₄CH₃COO								
Ni: CuO	Ni(CH ₃ COO) ₂ ·nH ₂ O		30	—	10–15	—	134	

CVs at a scan rate of 100 mV s⁻¹ for 5000 cycles. Fig. 26(c) demonstrated the cyclic voltammetry scan of CuO film electrode after the 1st to 5000th cycles and confirmed cyclic stability of 83% after 5000 cycles. By using the Ragone plot, the highest values of specific power and specific energy were measured as 3 and 14.1 W h kg⁻¹, respectively, using the GCD technique at a current density of 1 mA cm⁻² for CuO electrode attained in the potential range from 0 to 0.5 V as shown in Fig. 26(d).

Moreover, the electrochemical investigation of Ni-doped CuO nanolayers modified with Ni foam electrodes synthesized by Lobinsky *et al.* revealed the specific capacitance of 154 mA h g⁻¹ (1240 F g⁻¹) at a current density of 1 A g⁻¹,¹³⁵ as

already discussed in the doping section. Thus, SILAR-grown CuO material can be potential usage as an electroactive resource for alkaline batteries and pseudo-capacitors.

5.5. Photoelectrochemical characterization

The photo-responsive performance of m-SILAR grown Cu₂O/FTO electrodes was studied by Farhad and co-workers, through transient surface photovoltage under periodic illumination of a green LED by a HITACHI VG-4429 generator with ~0.1 Hz-square wave for '5 s ON and 5 s OFF' cycle.³⁶ The generated surface photovoltage of the Cu₂O/FTO electrode was observed by a Keithley SMU 2450 by employing Cu₂O/FTO



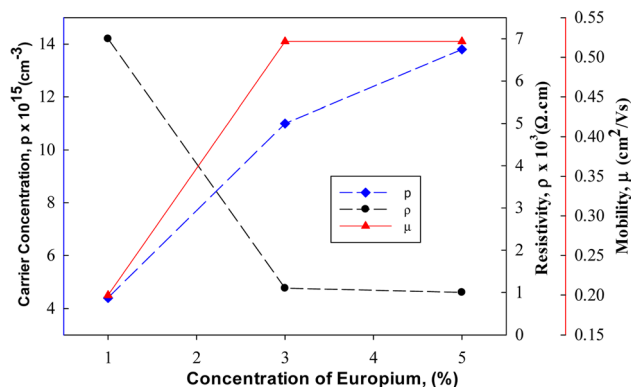


Fig. 17 Resistivity, carrier concentration and mobility for Eu doped Cu_2O films.¹²⁹

thin films as a working electrode, a graphite rod as a counter electrode and 0.1 M Na_2SO_4 aqueous solution as an electrolyte as established in Fig. 27(a) and (b). In the presence of an aqueous electrolyte, upon 2500 s LED exposure of the photocathode, the estimated V_{oc} for the samples grown with the non-optimized precursor, optimized precursor, and optimized precursor with CH_3COOH precursor solutions, were observed as $247 \pm 38 \mu\text{V}$, $36.0 \pm 2.0 \text{ mV}$ and $47 \pm 8 \mu\text{V}$ respectively. The large V_{oc} value projected for optimized precursor film revealed a better Schottky junction produced at $\text{Cu}_2\text{O}/\text{electrolyte}$ interface, consequently, advocating a better optoelectronic quality of Cu_2O thin film. The transient surface photovoltage and V_{oc} retention for 5000 s, advocating better stability of the SILAR fabricated Cu_2O thin films in aqueous electrolyte.



Fig. 18 (a) 1%, 3% and 5% Eu doped thin films having current–voltage characteristics of ITO/ZnO NRs/Eu: $\text{Cu}_2\text{O}/\text{Al}$ cell, (b) carrier transport and band structure of p–n junction.¹²⁹

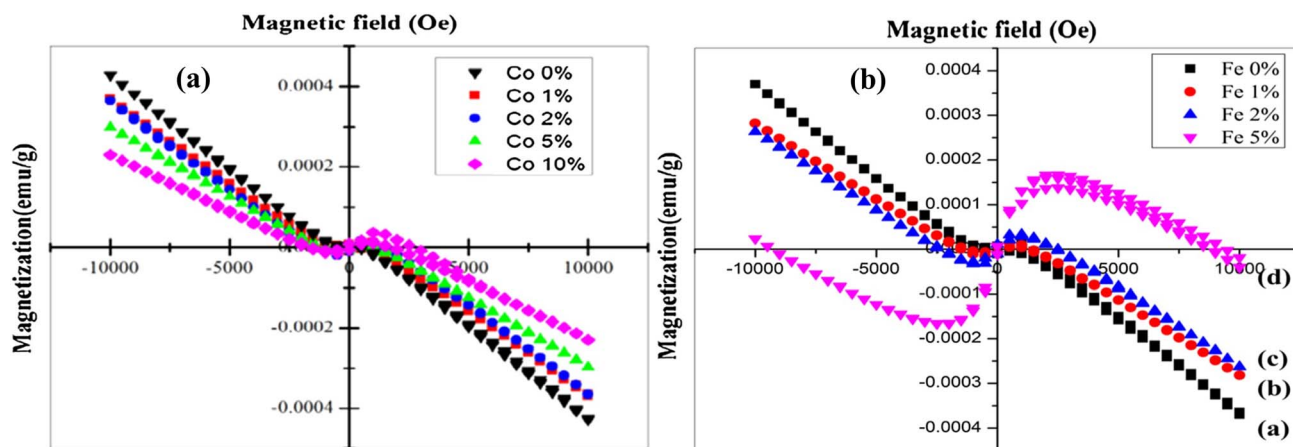


Fig. 19 Magnetic behavior of (a) Co-doped¹²⁷ (b) Fe-doped Cu_2O thin films.¹²⁸



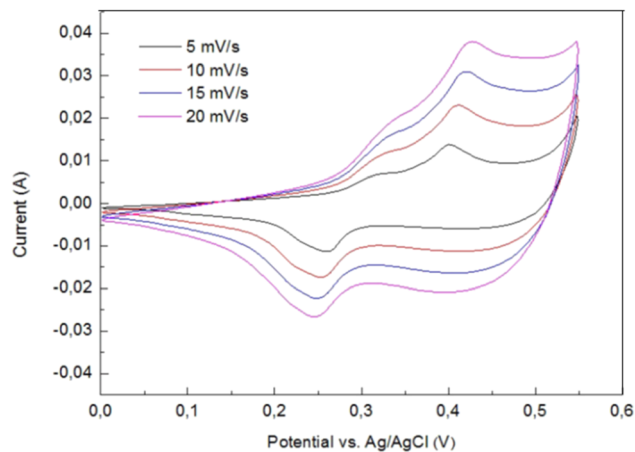


Fig. 20 CVA curves for nickel foam electrode with Ni-doped CuO nanolayers at a scan rate of 5, 10, 15 and 20 mV s^{-1} .¹³⁵

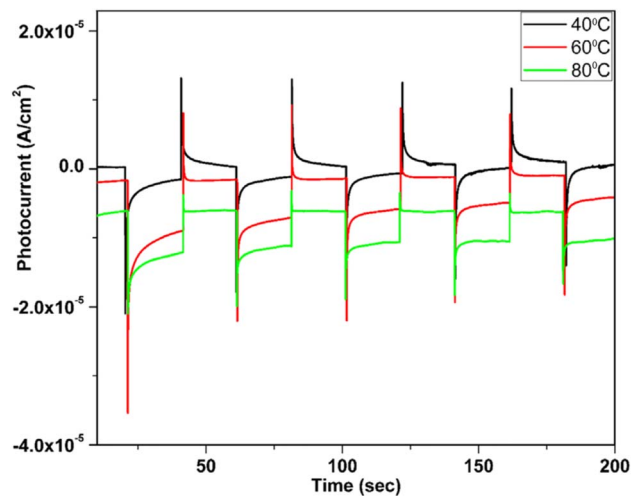


Fig. 23 PEC measurement of Cu₂O thin films at different bath temperatures.¹¹¹



Fig. 21 The cycling stability for the Ni Foam electrode with Ni-doped CuO nanolayers at 2 A g^{-1} whereas the inset graph represents galvanostatic charge-discharge curves of the electrode with Ni-doped CuO nanolayers.¹³⁵

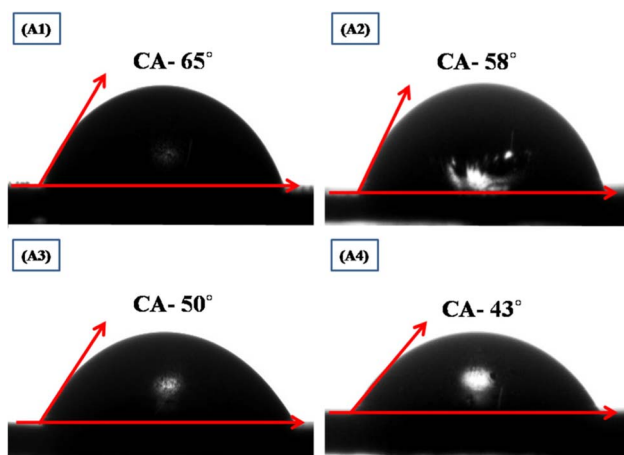


Fig. 24 Study of surface wettability of CuO thin films.¹⁴⁹ (A1) 50 SILAR cycles, (A2) 60 SILAR cycles, (A3) 70 SILAR cycles and (A4) 80 SILAR cycles.

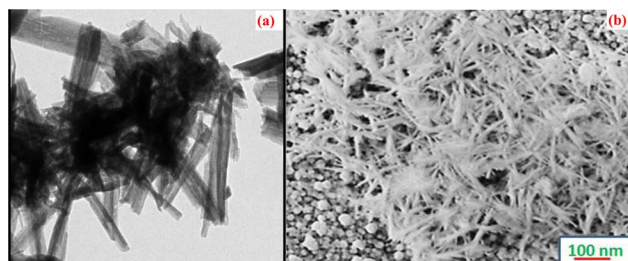


Fig. 22 (a) TEM (b) SEM images of the Cu₂O grains showing needle-shaped uniformity on the surface.¹⁴⁶

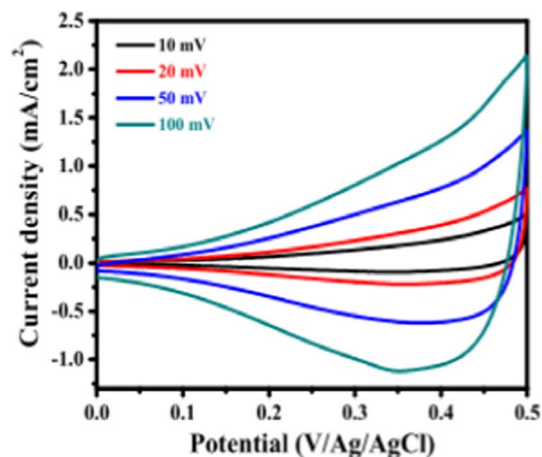


Fig. 25 Cyclic voltammograms of CuO with various scan rates of 1 M KOH electrolyte.¹⁴⁹



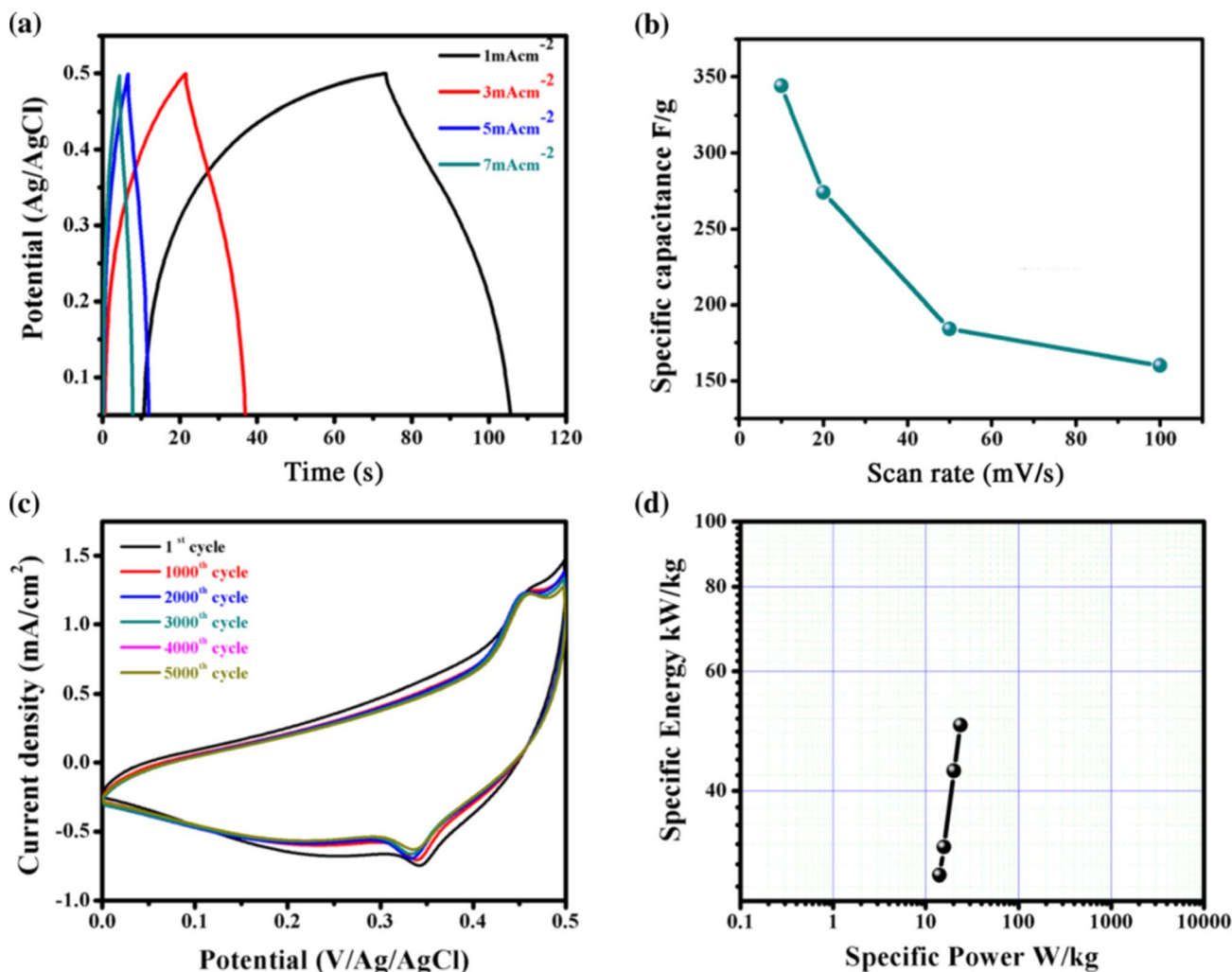


Fig. 26 (a) Representation of galvanostatic charge–discharge curves of CuO at diverse current densities. (b) Change of specific capacitance with scan rate (c) study of stability of CuO thin films (d) specific energy versus specific power of Ragone plot.¹⁴⁹



Fig. 27 (a) A representation of the transient surface photovoltage measurement system; (b) measurement of a typical Cu₂O/FTO electrode in a PEC with 0.1 M Na₂SO₄ electrolyte followed by real-time measurement in SMU 2450 demonstrating the shape of the LED modulated transient surface photovoltage.³⁶



6. Conclusion

In summary, Cu_xO thin films have been extensively studied and are receiving profound attention because of their fascinating properties and promising uses in a variety of fields. In this article, an inclusive review of the state-of-the-art research activities of diverse Cu_xO thin films was represented based on the SILAR method. This technique has fascinated substantial attention because of its simplicity and low cost, demands less time, and is fit for the large-scale growth of Cu_xO . The morphology, as well as diverse properties of Cu_xO , can be monitored by altering the number of SILAR cycles, the pH of precursor solutions, types of salt, bath temperature, annealing, doping, and the dipping time allowed for reactions. However, the technique does not yet allow for precise control of Cu_xO particle sizes, which can affect the power conversion efficiency in optoelectronic devices. The main limitation of this technique is the high rate of surface roughness as well as less study of the defects in the deposited sample which is very important to control the optical as well as electrical properties in optoelectronics. Having the optimum amount of the deposited Cu_xO is a very significant factor in improving optoelectronic performance. Thus, the inclusion of ligands, complexing agents and surfactants in the precursor solution employed during the SILAR growth could advance the stability of Cu_xO . Precise control of Cu_xO fabrication could accelerate multiple exciton generation effects, leading to a development of overall efficiency.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work is funded by Comilla University and University Grants Commission, Bangladesh. The authors like to acknowledge the assistance and scientific contribution from the Department of Chemistry, Comilla University, Cumilla 3506, Bangladesh, as well as the collaborative support from the Department of Electrical and Electronics Engineering, Saga University, Japan and Center for Nanotechnology, the Department of Natural Sciences, Coppin State University, Baltimore, MD, USA.

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