

Spectroscopy of corannulene cations in helium nanodroplets[†]

Michael Gatchell,^a Paul Martini,^a Felix Laimer,^a Marcelo Goulart,^a Florent Calvo^c and Paul Scheier^a

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Helium tagging in action spectroscopy is an efficient method for measuring the absorption spectra of complex molecular ions with minimal perturbations to the gas phase spectra. We have used superfluid helium nanodroplets doped with corannulene to prepare cations of these molecules complexed with different numbers of He atoms. In total we identify 13 different absorption bands from corannulene cations between 5500 Å and 6000 Å. The He atoms cause a small, chemically induced redshift of the band positions of the corannulene ion. By studying this effect as a function of the number of solvating atoms we are able to identify the formation of solvation structures that are not visible in the mass spectrum. The solvation features detected using action spectroscopy agree very well with the results of atomistic modeling based on path-integral molecular dynamics simulations. By additionally doping our He droplets with D₂, we produce protonated corannulene ions. The absorption spectrum of these ions differs significantly from the case of the radical cations as the numerous narrow bands are replaced by a broad absorption feature that spans nearly 2000 Å in width.

Introduction

Messenger spectroscopy¹ with helium atoms as tags has in recent years paved the way for determining the previously unobtainable gas phase absorption spectra of complex molecules with the accuracies required for identifying the origin of astronomical absorption features. This was demonstrated particularly well in 2015 when C₆₀⁺ was shown^{2–4} to be the first identified carrier of at least five of the several hundred Diffuse Interstellar Bands (DIBs), the origin of which have been a mystery for nearly a century.⁵ Maier and coworkers achieved this result by

^aInstitut für Ionenphysik und Angewandte Physik, Universität Innsbruck, Technikerstr. 25, A-6020 Innsbruck, Austria. E-mail: michael.gatchell@uibk.ac.at

^bDepartment of Physics, Stockholm University, 106 91 Stockholm, Sweden

^cUniversité Grenoble Alpes, CNRS, LIPhy, Grenoble, France

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storing the C_{60}^+ ions in an ion trap cooled to 4 K. The trapped ions interacted with He gas so that complexes of $He_nC_{60}^+$ ($1 \leq n \leq 3$) were formed. These complexes were irradiated with a tunable laser and when a photon was absorbed by the fullerene ion, the excess energy would cause any attached He to boil off. By measuring the depletion of $He_nC_{60}^+$ using a mass spectrometer as a function of wavelength and cluster size, the gas phase spectrum of bare fullerene ions could be deduced through interpolation down to $n = 0$.^{2,6–8} The key benefit of using helium for spectroscopy in this way is the extremely weak interaction between the He atoms and the molecule being studied that gives a relatively small shift compared to other methods such as matrix isolation in solid neon and argon.^{9,10}

One of the difficulties of performing the experiments that led to the identification of C_{60}^+ as a DIBs carrier was attaching the He atoms to the fullerene ions. Helium is the most chemically inert element and the binding energy of a He atom to C_{60}^+ is only about 10 meV.¹¹ However, droplets of He, containing thousands or millions of atoms, can efficiently capture and swallow gas phase dopants, cooling them to the 0.37 K equilibrium temperature of the surrounding He.^{12,13} By impacting these doped droplets with energetic electrons, both positively and negatively charged products can be formed.¹³ This leads to charged dopants being expelled from the droplets, often complexed with some He atoms still attached, so that they can be studied using mass spectrometry.^{12,13} Compared to measuring the spectra of molecular systems fully solvated in He droplets,¹⁴ this method of producing ions tagged with a relatively small number of He atoms allows for the distortions to the electronic transitions to be minimized and the effects to be quantified as a function of the level of solvation.

Kuhn *et al.*¹⁵ used helium nanodroplets to produce C_{60}^+ decorated with up to about 100 He atoms, on which spectroscopy measurements were performed. These showed that the absorption lines of $He_nC_{60}^+$ were systematically redshifted by $0.72(1) \text{ \AA}$ per He atom for $5 \leq n \leq 32$,¹⁵ which is in good agreement with the findings of Campbell *et al.*³ For $32 < n \leq 60$ a slight blueshift per He atom was observed before the lines again were redshifted towards the positions for C_{60}^+ solvated in bulk He. These observations were explained by the first 32 He atoms occupying essentially localized positions above the 20 hexagonal and 12 pentagonal faces of the fullerene cages, each providing a nearly equally strong influence to the electronic structure of C_{60}^+ .^{11,15} After that, the subsequent He atoms continue to fill up the first monolayer of He adsorbed to the fullerene, displacing other atoms from their preferred positions above the pentagonal rings until this layer is filled at $n = 60$.^{11,15} This leads to a general weakening of the interaction between the individual He atoms and the fullerene ion, giving a slight blueshift with each additional He. For $n > 60$ a second solvation shell of He begins to form, which has a much weaker interaction with the C_{60}^+ cage.^{11,15} This picture, together with the statistical sampling of different isomers of the $He_nC_{60}^+$ complexes from path-integral molecular dynamics (PIMD) simulations, was also used to explain observed non-monotonic variations in the width of the 9635 Å line as a function of n .¹⁶

Recently, Hardy *et al.*¹⁷ presented measurements of the absorption spectra of coronene ($C_{24}H_{12}$) and corannulene ($C_{20}H_{10}$) cations, two species of polycyclic aromatic hydrocarbons (PAHs) considered as potential DIBs carriers, using the He tag technique. While that study produced a negative result—the measured absorption bands do not match any known DIBs—they observed that $C_{20}H_{10}^+$ ions decorated with one and two He atoms had resonance energies at the same



position, *i.e.* no blue- or redshift was observed, and concluded that the red-shift induced by the He tagging was approximately 0.1 \AA ,¹⁷ significantly smaller than the shift observed for C_{60}^+ .

Here we present measurements of the absorption spectra of corannulene cations, produced from doped He nanodroplets, using the He messenger technique for systems containing up to 60 He atoms. In this way we have identified 13 absorption bands in the range of 5500 to 6000 \AA . Compared to the spectrum of C_{60}^+ decorated with He, where lines are redshifted linearly as a function of the number of solvation atoms, we find a much more complex behavior with corannulene with a varying magnitude of chemical shift being induced by each additional He atom. These features are well explained by atomistic modeling using path-integral molecular dynamics simulations, which accurately sample the delocalized nature of the weakly interacting He atoms. Spectroscopic measurement of He-tagged molecular ions proves to be a highly sensitive method for probing the structures of weakly bound complexes and it allows us to identify changes in the solvation structures that are not accessible by conventional mass spectrometry alone.

Experimental setup

Neutral helium nanodroplets are formed in the supersonic expansion of compressed He gas (2.5 MPa, 99.9999% purity) through a $5 \mu\text{m}$ nozzle that is cooled to 9.5 K. This results in superfluid He droplets with an equilibrium temperature of 0.37 K that have a typical size of a few million atoms.¹⁸ The central part of the beam passes through a 0.8 mm skimmer located 8 mm downstream. The droplets then transverse a series of pickup chambers where the droplets are doped with corannulene (evaporated from an oven at $70 \text{ }^\circ\text{C}$) and D_2 gas (99.8% purity, fed from a bottle through a gas line in the pickup chamber in experiments where D_2 is used). The doped droplets are ionized by the impact of 46 eV electrons, which mainly forms He^+ near the surface of the droplet. The charge migrates through the droplet by resonant hole hopping,¹⁹ attracted by the dopant that is more polarizable than the surrounding He. The large difference in ionization energy between He and the dopant leads to the latter being ionized and expelled from the droplet, often together with some additional neutral He atoms. This leaves a distribution of dopant ions with up to several tens of He atoms (or D_2 molecules) attached, upon which spectroscopy is performed. The ionized systems are then deflected 90° towards the entrance of the time-of-flight mass spectrometer (Tofwerk AG model HTOF) that is mounted perpendicularly to the neutral beam line. The deflection of the charged beam ensures that the laser does not overlap with the neutral molecular beam.

The mass spectrometer operates on a 10 kHz cycle and before every tenth extraction pulse, a tunable OPO (Optical Parametric Oscillator, EKSLPA model NT242 with up to $450 \mu\text{J}$ per pulse) laser with a 1 kHz repetition rate is used to probe the decorated corannulene ions in the region between the ion source and the mass spectrometer. During the remaining nine out of ten cycles of the mass spectrometer, a background spectrum is accumulated with the laser switched off. This background spectrum is then used to flatten the measured absorption spectrum to correct for the small variations in the current of ions over time.



The wavelength of the laser is scanned in steps of 0.1 \AA (5 cm^{-1} linewidth) over the measurement window and mass spectra are recorded with a $m/\Delta m$ resolution of approximately 2000 for each step. The absorption spectra of the $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ (or $\text{D}_n\text{C}_{20}\text{H}_{10}^+$) complexes are obtained by measuring the depletion of ions of a given mass as a function of wavelength as they boil off He atoms (D_2 molecules) following the absorption of photons.

Simulation details

Classical and quantum structures of $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ clusters were theoretically determined following an already established computational framework.²⁰ Under the experimental conditions, the corannulene cation is essentially frozen in its rovibrational ground state and its geometry was obtained from a standard electronic structure calculation employing the B3LYP functional and the 6-311++G** basis set. In our modeling, helium atoms interact with the corannulene cation through pairwise repulsion–dispersion forces of the Buckingham type, and a vectorial polarization contribution is included as well. Here the electric field on individual helium atoms results from the partial charges distributed on the molecular cation, which were obtained using a RESP (Restrained ElectroStatic Potential) procedure at the optimized geometry. The geometry and partial charges employed in the calculations are provided in the ESI.†

The stable structures of the $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ clusters were first explored using basin-hopping global optimization, which relies on a Monte Carlo exploration of locally minimized geometries.²¹ Here, in addition to conventional (large amplitude) Monte Carlo moves, we also implemented a specific mirror displacement move in which a specific atom was moved to the other side of the corannulene with a fixed probability (5%). Such random displacements are especially useful at small sizes, where successfully moving across the large hydrocarbon is unlikely. Ten series of 10^5 collective Monte Carlo moves were thus performed and generated putative classical global minima for n ranging from 1 to 60.

Nuclear delocalization was then accounted for by performing PIMD simulations, employing as in our earlier investigation²⁰ a Trotter discretization number of $P = 64$ at temperature $T = 1 \text{ K}$, which is high enough for exchange effects to be safely ignored. The PIMD equations of motion were numerically solved using a velocity Verlet algorithm with a time step of 1 fs and the temperature being imposed using a massive Nosé–Hoover thermostat with mass $3n/k_{\text{B}}T$, which is the usually recommended value.²² The PIMD trajectories were integrated for 1.5 ns, with statistical averages accumulated after the first 500 ps equilibration period. From these simulations, the quantum energies were evaluated from the standard virial definition,²² and for ease of visualization we also recorded the three-dimensional atomic densities.

Results and discussion

Absorption spectrum of corannulene complexed with He

Fig. 1 shows an overview of the $\text{HeC}_{20}\text{H}_{10}^+$ absorption spectrum detected by the depletion of ions of mass 254 u as a function of laser wavelength in the range of 5550 to 6000 Å. In this range we have identified 13 absorption features that we can reliably fit with Gaussian functions for several different numbers of attached He



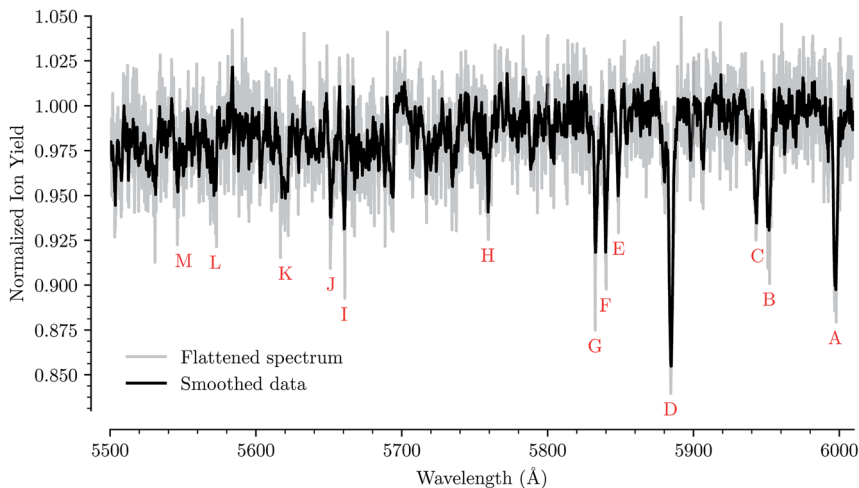


Fig. 1 Absorption spectrum in a range from 5500 Å to 6000 Å of $C_{20}H_{10}^+$ complexed with a single He atom. The flattened raw spectrum is shown in gray and is overlaid with a spectrum where the data has been smoothed (moving average over 10 nearest neighbors) for easier viewing. All of the analysis is performed on the unsmoothed dataset. Identified absorption lines are labeled and their positions are reported in Table 1.

atoms. These are labeled from A through M in Fig. 1. Other minima that appear to be visible in Fig. 1, but which are not found in the spectra of systems containing more He atoms, are not considered for further analysis. Bands A, B, C, D, F, and G were reported by Hardy *et al.*¹⁷ as two series of bands originating from the Jahn–Teller distorted electronic ground state of $C_{20}H_{10}^+$, which they assigned to $S_3(^2A') \leftarrow S_0(^2A'')$ and $S_3(^2A') \leftarrow S_0(^2A')$ transitions, respectively. We do not conclusively assign the other seven bands that we identify, but expect that vibrational excitations of the S_3 state contribute to them. Calculations show that transitions to the $S_4(^2A')/S_4(^2A'')$ states may also be responsible for some of these bands.²³

For each of the bands labeled in Fig. 1 we have performed a least square fit of a Gaussian line profile. This was then repeated for the corresponding spectra with more He atoms attached to the $C_{20}H_{10}^+$ ions. Fig. 2 shows a few examples of such fits to band D near 5885 Å and the positions and relative intensities of all bands (for the $HeC_{20}H_{10}^+$ complex) are given in Table 1. From Fig. 2 it is clear that the band position is systematically redshifted by the addition of solvating He atoms. This trend continues up to $n = 32$ before the positions are slightly blueshifted again.

Fig. 3 shows how the central positions of two absorption bands (D and H) vary as a function of the number of complexing He atoms and the data for the remaining bands are shown in the ESI.† For the first three He atoms, there are only small changes in the resonance wavelengths of the $He_nC_{20}H_{10}^+$ complex, with less than 0.2 Å separating them. After this there is strong redshifting in each step between $n = 3$ and $n = 8$, where the absorption bands are shifted by, on average, 0.7 Å per He atom (however the trend is not entirely linear). A plateau is then observed in the range between $n = 8$ and $n = 13$ where the wavelengths remain essentially constant. This is then followed by the $13 \leq n \leq 32$ region where the



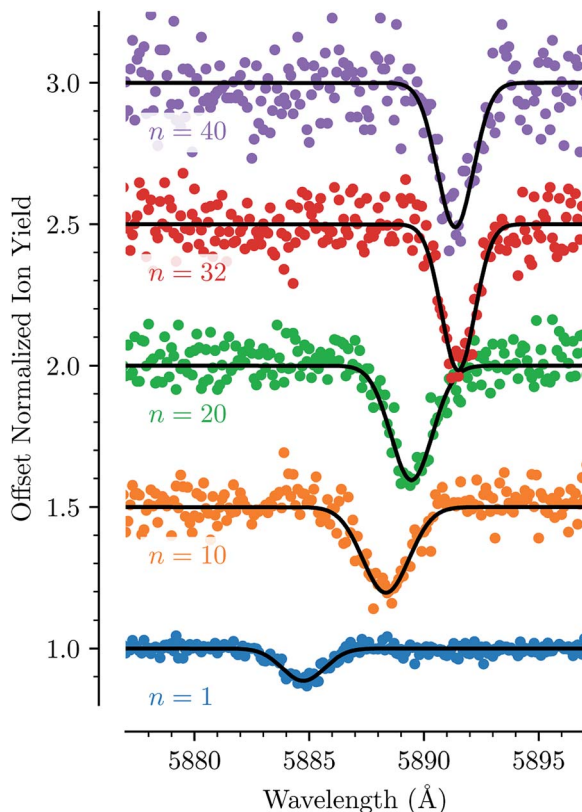


Fig. 2 The absorption feature of $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ near 5885 Å (line D) shown for different numbers of n . The line is redshifted in a non-linear fashion up to $n = 32$, after which the position is very slightly blue-shifted again. Gaussian fits are shown as solid black lines.

Table 1 Observed absorption bands for $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ complexes with measured values reported for $n = 1$. The given uncertainties are the statistical errors from the least square fits of Gaussian profiles to the absorption features

Band	Center wavelength (Å)	Relative strength
A	5997.06 ± 0.08	0.86 ± 0.06
B	5951.86 ± 0.18	0.34 ± 0.06
C	5943.75 ± 0.17	0.28 ± 0.06
D	5884.21 ± 0.08	1.00 ± 0.07
E	5848.60 ± 0.44	0.32 ± 0.10
F	5839.68 ± 0.21	0.27 ± 0.07
G	5832.51 ± 0.16	0.42 ± 0.07
H	5758.97 ± 0.08	0.76 ± 0.04
I	5660.43 ± 0.14	0.27 ± 0.05
J	5651.43 ± 0.14	0.36 ± 0.05
K	5620.23 ± 0.17	0.48 ± 0.06
L	5572.13 ± 0.34	0.18 ± 0.05
M	5550.43 ± 0.83	0.19 ± 0.07



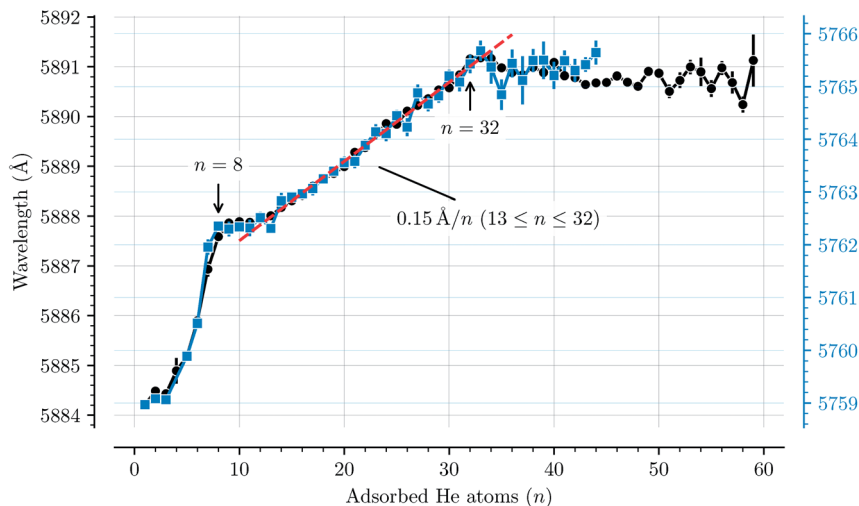


Fig. 3 Absorption wavelengths of bands D (black circles, left scale) and H (blue squares, right scale) as a function of the number of He atoms (n) adsorbed on the corannulene cations. Both bands show the same behavior. A linear redshift trend is observed in the $13 \leq n \leq 32$ range together with a fit (red dashed line). The $n = 4$ data point in the blue data set is missing due to the large uncertainty caused by corannulene complexed with water, which has a similar mass. All of the remaining bands that are studied show the same behavior and are presented in the ESI.†

wavelength is redshifted following a stable linear trend of 0.15 \AA per He atom. For $n > 32$ the wavelengths remain nearly constant with a slight blueshift relative to the $n = 32$ band positions. All 13 of the studied absorption features show the same basic behavior and no systematic variations in the band widths are observed as a function of the number of He atoms for any line.

The trend observed with $\text{C}_{20}\text{H}_{10}^+$ -He complexes differs significantly to that observed with C_{60}^+ -He complexes. In the latter a constant redshift of 0.7 \AA per He atom was observed for the first 32 He atoms.^{3,15} This persistent trend was explained by the spherical symmetry of the fullerene cages and their 32 faces, above which He atoms are preferably located,¹¹ and by utilizing a model to describe the polarization of the fullerene cage induced by the surrounding He atoms.¹⁵ Because of the high symmetry, all of the first 32 He atoms had nearly the same interaction energy with C_{60}^+ , thus inducing the same spectral shift. The linear trend also meant that the gas phase spectrum of the bare C_{60}^+ ion could be estimated using a linear extrapolation down to zero He atoms.² For corannulene, determining the unperturbed gas phase spectrum with the same accuracy as for C_{60}^+ presents a difficulty because of the more complex nature of how the He atoms influence the band positions.

Structures of corannulene complexed with He

We have used path-integral molecular dynamics simulations in order to investigate the solvation of the corannulene cation in He and explain the features seen in Fig. 3. Density plots for the locations of He atoms surrounding a corannulene cation for different degrees of solvation are shown in Fig. 4 and the corresponding



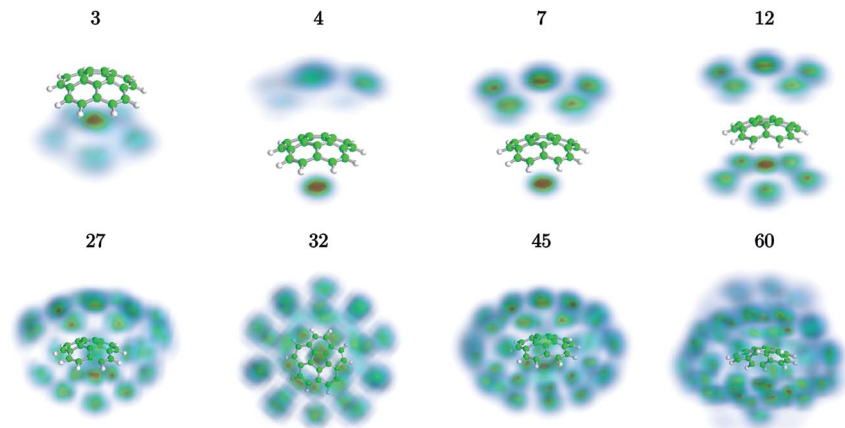


Fig. 4 Probability densities for the most stable geometries of He-solvated corannulene cations for different numbers of He atoms from our atomistic modeling. The first and most tightly bound He atom is preferentially positioned over the pentagonal face on the concave side of the corannulene ion.

binding energies are shown in Fig. 5. The simulations show that the first He atom that binds to $C_{20}H_{10}^+$ preferentially occupies the position above the central carbon pentagon of the concave side of the molecule. The binding energy at this site is found to be about 27 meV, which is considerably higher than for positions above the pentagonal and hexagonal faces on the convex side of the molecule that are equivalent to the outer surface of a C_{60}^+ ion. The next two additional He atoms also preferentially occupy the inside surface of the corannulene bowl, but due to the curvature and the interaction between the He atoms, they are not strongly localized and can easily move around the center He atom along the inner rim. This competition for space in the corannulene ion leads to the mean interaction energy per He sharply decreasing for the first three He adducts, as seen in Fig. 5.

For complexes with between four and seven He atoms, the most favorable structures all involve a single He atom near the center of the concave side, with the remaining He atoms now occupying the convex side. Like with $He_nC_{60}^+$ complexes, the He atoms on the convex outer surface of the corannulene cation are each preferentially located above the faces of the carbon rings. Once the six outer faces are occupied, the filling then continues on the concave side. From $n = 8$ to $n = 12$, the rim of the concave side is filled with the He atoms mainly occupying positions above the open HCCCH pockets. For $n \geq 13$, the additional He atoms fill the periphery of the corannulene ion, again starting with the convex side, until the first, essentially solid, solvation layer is entirely filled at $n = 32$. After this, additional He atoms may slot into the first solvation layer by displacing some of the localized He from their preferred positions. We do not identify a sharp closure of this sub-layer in the simulations, but it occurs close to $n = 45$. For sizes larger than this, we find that a second solvation layer begins to form, with the second shell's He atoms largely being shielded from the corannulene by the first layer of He. In the last structure of Fig. 4 ($n = 60$) we can see that a second, more delocalized layer of He has been formed with the outer layer of He mainly occupying the shallow dimples between He atoms in the first layer.



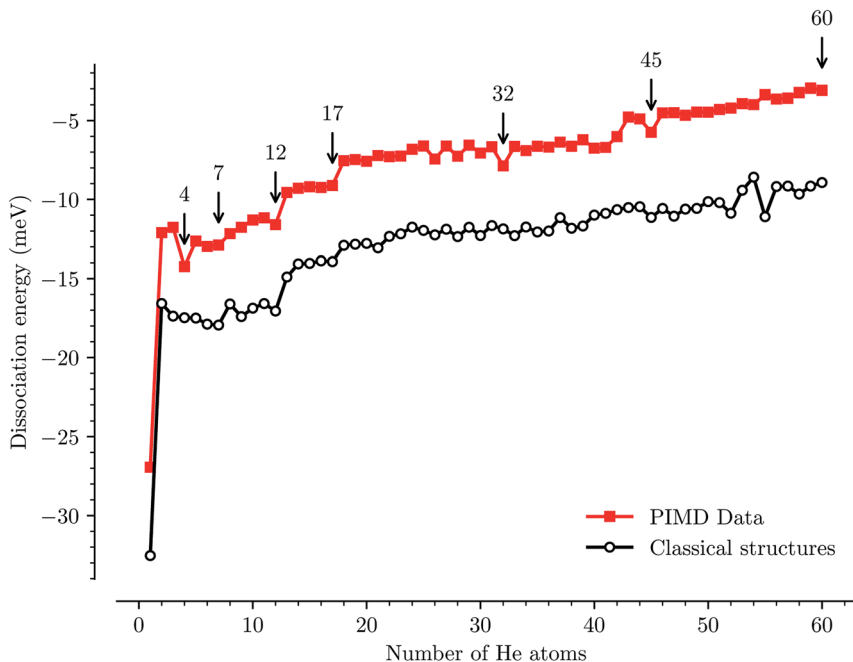


Fig. 5 Individual binding energies of He atoms physisorbed to the corannulene cation as a function of the number of atoms from our path-integral molecular dynamics simulations (red squares) and the equivalent classical structures (black circles). The PIMD simulations consider the delocalized quantum nature of the He atoms on the shallow potential energy surface and provide a more accurate description of the interactions than the classical calculations.

The changes in solvation structures determined by the simulations (Fig. 4 and 5) describe very accurately the features seen in the solvation-dependent band positions (Fig. 3). The small differences in band positions for $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ complexes with $1 \leq n \leq 3$ are consistent with the first three He atoms occupying the concave side of the corannulene ion and competing for space, so that the total interaction energy in the complex remains nearly constant as the second and third He atoms are added. Similarly, following the complete covering of the carbon faces on the convex side at $n = 7$, the plateau seen in Fig. 3 for $8 \leq n \leq 13$ is again well explained by the filling of the concave side of the corannulene ion in the simulations. Finally, the closure of the first solvation shell at $n = 32$ is clearly seen in both the spectroscopic data and the simulations. This value is consistent with results for the slightly larger coronene cation, which consists of seven hexagons in a planar structure, where the equivalent solvation shell forms after 38 He atoms.^{24,25} However, the fact that $\text{He}_n\text{C}_{60}^+$ also shows a form of shell closure (with an associated maximum in the redshift of the absorption bands) at the very same size of $n = 32$ is merely a coincidence.^{11,15}

Comparison with results from mass spectrometry

While detailed information about the solvation of corannulene in He can be deduced from the spectroscopy of $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ complexes, these features are for



the most part not visible in our mass spectra from these measurements. In Fig. 6 a mass spectrum obtained in these measurements to determine the depletion of $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ complexes as a function of laser wavelength is shown. The two strongest features in the mass spectrum are the peaks of the corannulene monomer and dimer. Between these peaks we detect the series of pure He_m^+ clusters ($63 \leq m \leq 124$), indicated by the blue trace in Fig. 6, and the $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ series for $1 \leq n \leq 60$ indicated in red. In addition, a number of contaminants litter the mass spectrum with the most abundant one coming from the corannulene cation complexed with a single water molecule (or O or OH fragments) near 268 u/e. These contaminants are mainly introduced *via* the corannulene sample being used, but at least some of the water could also be picked up from the residual gas in our device. The pure He_m^+ cluster series follows an exponential decrease in intensity as a function of size (hence the linear appearance on the logarithmic scale), with a few deviations being the result of overlapping peaks from contaminants (the largest one being caused by the aforementioned $\text{H}_2\text{OC}_{20}\text{H}_{10}^+$ complex). The $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ series follows a similar exponential decrease in intensity up to $n = 32$ (with deviations again mainly being caused by contaminants), and again up to $n = 45$, before clear drops in intensity are measured. These changes in intensity are typical of shell closures, in this case the filling of the first solvation shell in two steps, and similar features have been found for $\text{He}_n\text{C}_{60}^+$ at $n = 32$ and $n = 60$.^{11,13} Interestingly, the second kink in the exponential intensity curve visible after $n = 45$ is not observed in the absorption spectroscopy data (Fig. 3).

Comparing the data in Fig. 3 and 6 illustrates a practical side effect of performing action spectroscopy of ions solvated in He droplets: an increased

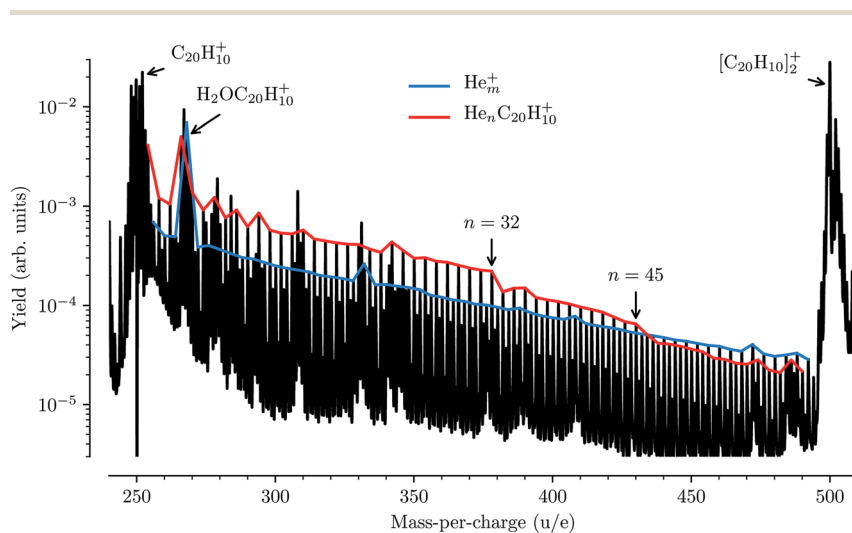


Fig. 6 Mass spectrum of corannulene cations formed by electron impact ionization of doped He nanodroplets. The series of He-complexed corannulene ions is indicated by the red line and the blue line shows the pure He_m^+ series. There are a number of contaminants that give additional features in the mass spectrum, the most prominent one being water giving a strong signal of $\text{H}_2\text{OC}_{20}\text{H}_{10}^+$ near mass 268 u. Note the clear intensity drops in the red $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ curve after $n = 32$ and $n = 45$.



sensitivity to the structures of loosely bound complexes over pure mass spectrometry and a decreased sensitivity to contaminants of a similar mass to the system being studied. The sensitivity of the spectroscopic measurements to small, chemically induced shifts has allowed us to identify features that give a detailed insight into the solvation of corannulene ions in He, whereas the mass spectrum alone only reveals the complete closure of the first solvation shell. This method could thus be applied for experimentally studying the structures of small nanoparticles, clusters, and molecules. As noted in the introduction, action spectroscopy using He tags is also a sensitive tool for performing spectroscopic studies on complex molecular systems at temperatures close to 0 K and with much smaller perturbations than with traditional matrix isolation techniques. Also, since it is unlikely that two very different systems have electronic transitions at exactly the same energy, the contaminants that make the mass spectrum difficult to analyze will generally be transparent to the photons that are at resonance with a transition in the studied $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ complexes. The main effect that the contaminants have on our measurements is that they increase the background noise level, reducing the signal-to-noise ratio in systems that overlap in mass with a contaminant. However, this is only an issue for very strong features such as the $\text{H}_2\text{OC}_{20}\text{H}_{10}^+$ complex which, as can be seen in Fig. 6, is an order of magnitude more intense than the $\text{He}_n\text{C}_{20}\text{H}_{10}^+$ series at that mass. Additionally, since we measure the depletion of all masses as a function of laser wavelength at the same time with high mass resolution, we can be sure that we are indeed only measuring the absorption from the desired He-complexed ions.

Absorption spectra of protonated corannulene complexes

By introducing D_2 into the pickup chamber of our experimental setup, we form He droplets doped with a mixture of deuterium and corannulene. Upon ionization, excess energy often leads to the breakup of at least one D_2 molecule and the formation of protonated $\text{C}_{20}\text{H}_{10}\text{D}^+$ systems. By using D_2 instead of H_2 we obtain a larger mass separation from the $\text{C}_{20}\text{H}_{10}^+$ radical, and avoid overlap with its $^{13}\text{C}^{12}\text{C}_{19}\text{H}_{10}^+$ isotopologue, which makes up about one out of every six corannulene molecules. Like He, intact D_2 molecules can form weakly bound solvation layers around molecules such as corannulene. Previous studies with coronene and C_{60} cations solvated in H_2/D_2 and He have shown that hydrogen molecules form closed solvation shells at the same numbers as helium.^{26,27}

Fig. 7 shows the absorption spectrum of $\text{C}_{20}\text{H}_{10}\text{D}^+$ measured by the depletion of ions with masses of 256, 300, and 348 u. Due to the similar masses of intact D_2 and He, we do not differentiate between the different mixtures of the two in the present measurement since both are weakly interacting and can be used as tags for action spectroscopy. The main feature in the absorption spectra is the very broad band between 4000 Å and 6000 Å. In addition, two narrow lines can be seen near 5280 Å and 5320 Å, the positions of which do not change as a function of the number of solvating molecules/atoms. These are due to an enhanced output of our OPO laser at the wavelength emitted by the pump laser (5325 Å is the second harmonic of the pump laser), and thus an enhancement of the already present depletion at these wavelengths, that we have not been able to fully correct for.

The protonated corannulene ion has a dramatically different absorption spectrum to the radical corannulene cation. Gone are the numerous narrow



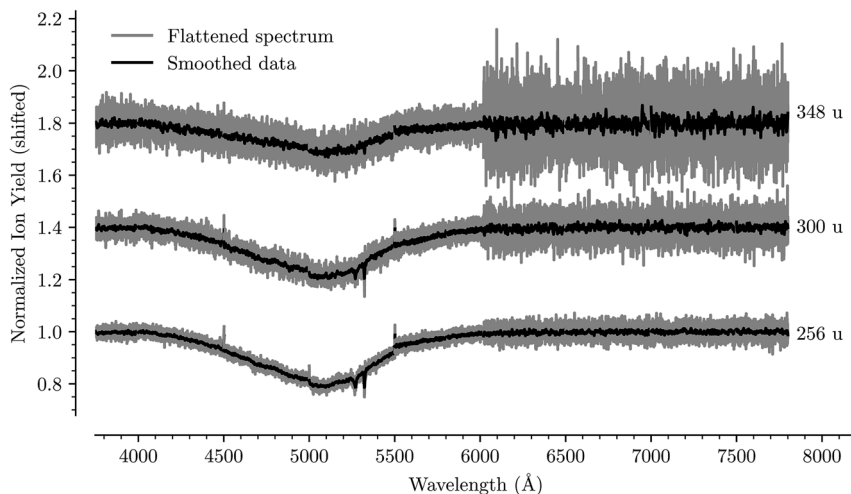


Fig. 7 Absorption spectrum in the range from 3750 Å to 7750 Å of $C_{20}H_{10}D^+$ complexed with 1, 12, and 24 He/D_2 atoms/molecules, respectively. The flattened raw spectrum is shown in gray and is overlaid with a spectrum where the data has been smoothed (moving average over 10 nearest neighbors) for easier viewing. A broad absorption feature is observed between 4000 Å to 6000 Å together with two narrow lines near 5280 Å and 5320 Å. The positions of the latter do not shift as a function of the solvation and are caused by an increased output from our laser that we have not been able to fully correct for. The measurements at wavelengths above 6000 Å were performed with shorter dwell times and have a higher signal-to-noise ratio than those at shorter wavelengths. A small discontinuity is also visible where two datasets were spliced together at 5500 Å.

absorption bands and we only detect a single broad feature. Since this absorption feature is so broad, we are not able to identify any shift in its position as a function of the number of solvating He/D_2 units. The position of this band is in agreement with the $S_3(^1A) \leftarrow S_0(^1A)$ transition of protonated corannulene that was measured using a resonant multiphoton fragmentation technique in an ion trap as a broad peak near 5150 Å.²³ The width of this peak, $FWHM \approx 800$ Å, can be explained by a short-lived $S_3(^1A)$ state with a lifetime of ≈ 2 fs, though unresolved vibrational excitations could contribute to the width making this lifetime a lower limit estimate. With no narrow bands originating from protonated corannulene in our measurements, it is unlikely that this system is responsible for any DIBs in this wavelength range. However, broad features like this could contribute to features in interstellar extinction curves like those that have been proposed to originate from hydrogenated fullerene molecules.²⁸

Summary and conclusions

We have used helium nanodroplets to produce corannulene cations solvated by He that we have measured the absorption spectra of. This method is highly efficient in producing cold molecular samples with weakly interacting tags for action spectroscopy whereby the absorption spectrum of the tagged molecule is measured indirectly from the depletion of He-complexed systems as a function of the laser wavelength. From our measurements we identify 13 absorption bands



between 5500 Å and 6000 Å and we have accurately studied the perturbation to the band positions as a function of the number of solvating He atoms. The behavior is found to be much more erratic than the linear behavior seen with $\text{He}_n\text{C}_{60}^+$ complexes up to $n = 32$. The different features seen with corannulene ions solvated in He are a sensitive tool for studying the structures of the weakly bound complexes and the results agree very well with our computational modeling, revealing features not seen when using mass spectrometry alone. By introducing D_2 to our He droplets before ionization, we form protonated corannulene complexed with He and intact D_2 . With the protonated system we do not detect any narrow absorption bands that could be potential DIBs candidates, but instead observe a broad absorption feature between 4000 Å and 6000 Å.

A disadvantage of the types of measurement performed here is that they are not directly accessing the gas phase band positions. This is a general problem with all forms of spectroscopy where a matrix or tags are used. However, He interacts only very weakly with other systems, minimizing these effects. So for systems with only one He atom attached, the chemical shift of the positions of the absorption bands should be no more than a few Ångströms. Given the relatively strong binding energy of the first He atom to the corannulene cation (about 27 meV) compared to the C_{60} cation (about 10 meV (ref. 11)), the chemical shift induced to the electronic transitions induced by the first He atom that binds to corannulene is likely larger than the 0.7 Å seen with $\text{C}_{60}^{+,3,15}$ perhaps greater than 1 Å. Although small for most applications, this shift is important to consider when comparing the experimental data with accurate astronomical measurements.

Finally, the ability to produce a wide range of cold mixed complexes in He nanodroplets and to perform action spectroscopy on them is a powerful tool for studying the structures of molecular ions and small nanoparticles. We foresee this being a useful method not only for investigating potential carriers of the diffuse interstellar bands, but for many other applications, including investigating systems where gas phase spectroscopy is difficult to perform, and as a highly sensitive tool for probing the structures of nano-solvated complexes.

Conflicts of interest

There are no conflicts to declare.

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