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Permanently porous hydrogen-bonded organic frameworks comprising rod-like molecules with two MIDA boronate termini have been prepared. We show that MIDA boronates self-assemble through multiple hydrogen-bonding interactions. Thiophene-containing frameworks are fluorescent and have a 6.6% absolute quantum yield. The approach appears to be general and introduces new design rules for constructing hydrogen-bonded organic frameworks.

Porous materials such as metal organic frameworks,^{1–3} covalent organic frameworks (COFs),^{4–6} and porous aromatic frameworks^{7,8} have emerged as promising candidates for a range of uses. Porous organic materials include intrinsic and extrinsic porous structures. Organic cage compounds are a typical class of intrinsic porous crystalline solids.^{9–14} Extrinsic porous materials are formed by the assembly of nonporous small molecules including COFs and the recent revival of hydrogen-bonded organic frameworks (HOFs). HOFs are connected by non-covalent interactions such as hydrogen bonds and various π - π interactions. HOFs that incorporate solvent guest molecules typically lose their porosity after the solvent is removed. As a result relatively few HOFs exhibit permanent porosity.^{15–22} Permanently porous materials are important for a range of uses,^{23–26} yet designing building blocks for permanently porous HOFs is challenging.¹⁵ The most frequently used building blocks contain the 2,4-diaminotriazinyl group (DAT) and related nitrogen-containing heterocycles (Scheme 1).¹⁶ There are a few alternative motifs such as benzimidazolones¹⁸ and pyrazoles,¹⁹ however, the design rules for HOFs are still in their infancy and very important for the continued development of porous materials.

Herein we report that hydrogen-bonded porous frameworks can be constructed from rod-like molecules that are capped

Permanently porous hydrogen-bonded frameworks of rod-like thiophenes, selenophenes, and tellurophenes capped with MIDA boronates†

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Scheme 1 (a) Previously reported functional groups for constructing HOFs. (b) The synthesis of diphenylheterocycle MIDA boronates in this report.

with *N*-methyliminodiacetic acid (MIDA) boronate termini. The rod-like molecules contain heterocycles that are typical of optoelectronically active materials, including thiophene, selenophene, and tellurophene.^{27–31} The MIDA boronate group has been applied to a range of cross-coupling synthetic methodologies.^{32–34} MIDA esters are easily handled, stable under air, and contain both hydrogen bond donors and acceptors. Here we demonstrate that this synthon is useful for the self-assembly of permanently porous materials.

MIDA-containing thiophene, selenophene, and tellurophene rod-like molecules (hereafter denoted **DPT-MIDA**, **DPSe-MIDA**, and **DPTe-MIDA**, respectively) were synthesized by a Pd(PPh₃)₄ and CuI catalyzed Stille-coupling reaction between

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† Electronic supplementary information (ESI) available: Detailed experimental information and characterization data. CCDC 1434463 (**DPT-MIDA**) and 1434461 (**DPSe-MIDA**). For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5dt04960a

4-iodophenylboronic acid MIDA ester (3.5 eq.) and the corresponding 2,5-bis(trimethyltin)heterocycle^{35,36} (1.0 eq.) (Scheme 1). This reaction afforded yellow solids of **DPT-MIDA**, **DPSe-MIDA**, and **DPTe-MIDA**, respectively in high yields (84%–90%). Pure and crystalline products were obtained by recrystallization from acetonitrile. The MIDA boronates were characterized by FT-IR, NMR, and HR-MS (see the ESI†).

The single crystal structures of **DPT-MIDA** and **DPSe-MIDA** reveal that two molecules of **DPT-MIDA** or **DPSe-MIDA** are contained within each unit cell and the MIDA end groups are in the *trans*-configuration. **DPT-MIDA** crystallize through multiple C–H...O hydrogen bonding and C–H... π interactions (Fig. 1a). The C–H...O hydrogen bonding interactions between the oxygen atom in the carbonyl groups and hydrogen in the methylene groups of adjacent MIDA functionalities vary between 2.20 to 2.65 Å. There are two kinds of C–H... π interactions, one is a 2.85 Å interaction between the thiophene and the phenyl groups of adjacent molecules, the other is a 2.71 Å interaction between the phenyl group and the hydrogen in the methyl group of adjacent MIDA functional groups. **DPSe-MIDA** assembled in a similar fashion (see S4, ESI†). Since the only difference is the chalcogen atom, both of the microporous frameworks have a $7 \text{ \AA} \times 10 \text{ \AA}$ pore size along the shortest distance across the channel. Solvent molecules are incorporated in the channel of the frameworks. Three molecules of acetonitrile are incorporated within the **DPT-MIDA** or **DPSe-MIDA** unit cells. In **DPT-MIDA**, one of the acetonitrile molecule is disordered (see S6, ESI†). In both frameworks, guest acetonitrile molecules are bonded with the wall of the channel through weak C–H...N hydrogen bonding and C–H... π interactions.

DPTe-MIDA was also recrystallized from acetonitrile, however only needle-like crystals could be obtained. In order to compare the structure of **DPTe-MIDA** with the other

samples, powder X-ray diffraction patterns (PXRDs) were collected on all three compounds. For reference the simulated PXRD patterns of **DPT-MIDA** and **DPSe-MIDA** were generated from their single crystal data. These closely match the experimental PXRD data (see S8–S9, ESI†), indicating that the structure in the powders are similar to the single crystals. The PXRD scattering pattern of **DPTe-MIDA** is weak compared to the other two frameworks, which indicates low crystallinity. None-the-less, the main diffraction peaks are close to the other two frameworks (see S10, ESI†), which implies all three frameworks are assembled into a similar structure.

Thermogravimetric analysis (TGA) was performed on each sample under a N_2 atmosphere. These experiments reveal that the incorporated acetonitrile molecules are removed at different temperatures. For S and Se, the guest-removal temperatures are around 150 °C; for Te, the temperature drops to 70 °C (see S11–S13, ESI†). The weight loss of **DPT-MIDA**, **DPSe-MIDA**, and **DPTe-MIDA** are 12%, 8.2%, and 6.1% respectively. This is close to the expected weight loss based on acetonitrile inclusion (10%, 9.4%, and 8.7% for **DPT-MIDA**, **DPSe-MIDA**, and **DPTe-MIDA**, respectively). This provides support that the degree of acetonitrile inclusion is similar in all frameworks. No additional weight loss was observed after the loss of acetonitrile until 350 °C, after which the frameworks began to decompose. The incorporation of acetonitrile in the frameworks was also observed in the ^1H NMR spectra at $\delta = 2.07$ ppm (see S14–S16, ESI†).

To investigate stability, variable-temperature PXRD patterns were collected on all three frameworks (Fig. 2 and see S20–S21, ESI†). Under air, structural changes were observed. **DPT-MIDA** is stable up to 100 °C. After that the structure begins to transform, the peaks at 8° and 8.8° become broad, and the peak at 12.2° shifts to 10.7° (Fig. 2). **DPSe-MIDA** is stable in air up to 125 °C. After that the structure begins to transform, the peaks at 5.4° , 8.2° , and 8.8° become broad, the intensity of peak

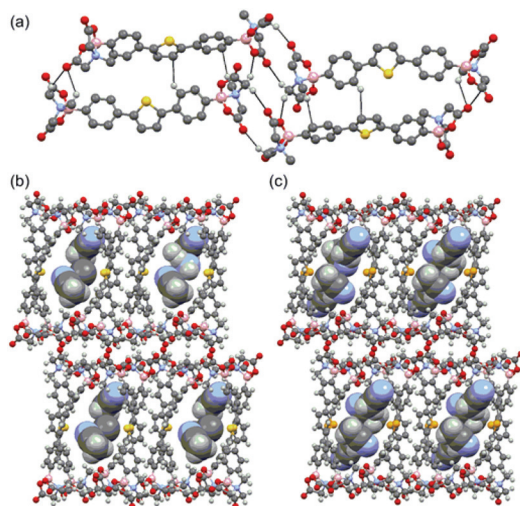


Fig. 1 (a) Hydrogen-bonding interactions in **DPT-MIDA**, (b) and (c) microporous structure of **DPT-MIDA** and **DPSe-MIDA** with acetonitrile solvent incorporated.

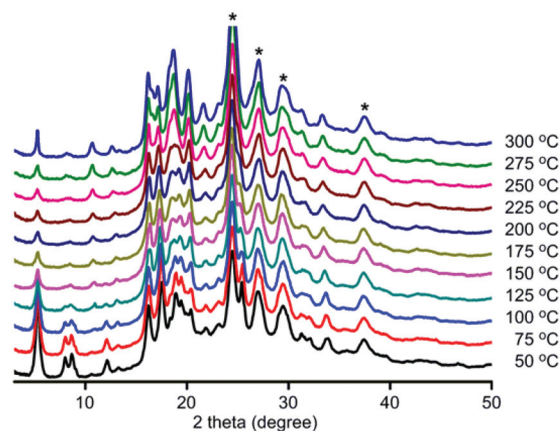


Fig. 2 Variable-temperature PXRD patterns for **DPT-MIDA**. The asterisks (*) denotes the background signal from a polyether ether ketone layer that protects the detector. The temperature was recorded in 25 °C increments.



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