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Selective dual-purpose photocatalysis for simultaneous H₂ evolution and mineralization of organic compounds enabled by a Cr₂O₃ barrier layer coated on Rh/SrTiO₃†

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Dual-functional photocatalysis for H_2 evolution with the simultaneous mineralization of 4-chlorophenol was achieved under de-aerated conditions using a $Cr_2O_3/Rh/SrTiO_3$ photocatalyst which has Rh nanoparticles covered with a thin Cr_2O_3 barrier layer to selectively control and maximize the dual-functional photocatalytic activity.

Hydrogen is considered as an ideal energy storage medium and a promising energy carrier since it can be obtained from abundant natural resources such as water and biomass instead of fossil fuels. 1-4 In particular, photocatalytic water splitting is widely studied as a promising technology to produce hydrogen using solar light.⁵⁻⁷ Another important application of photocatalysis is the degradation of organic compounds for the remediation of polluted water and air.8-12 In photocatalytic H₂ production, organic electron donors are commonly used as sacrificial agents to scavenge photogenerated holes. 13-17 However, the intentional addition of organic electron donors (e.g., alcohols, organic acids, amines) for H₂ production is not practically acceptable since the electron donors themselves are another energy resource (often more expensive than H₂). Therefore, a more desirable strategy is to use organic waste and pollutants in water as in situ electron donors. This is the concept of dual-functional photocatalysis which produces H2 along with the simultaneous degradation of organic pollutants. 18-20 Recently, Kim et al. demonstrated that the simultaneous H₂ production with the anoxic photocatalytic degradation of organic compounds can be achieved by using TiO₂ modified by both surface fluorination and Pt deposition (F-TiO₂/Pt). 19,20

In this study, a selective dual-purpose photocatalysis that achieves H_2 production and TOC removal simultaneously under anoxic conditions is reported. Instead of using the combination of TiO_2 (as a base photocatalyst) and Pt (as a cocatalyst for H_2 evolution), $SrTiO_3$ (base photocatalyst) and $Rh@Cr_2O_3$ core–shell nanostructure (cocatalyst) were employed in this work. The photocatalytic activities of the composite materials of $Cr_2O_3/Pt/SrTiO_3$ and other $Rh@Cr_2O_3$ (core–shell) loaded metal oxides have been previously demonstrated for the overall water splitting. ^{21,22} In this work, we demonstrate that $Cr_2O_3/Rh/SrTiO_3$ can be a promising dual-purpose photocatalyst that can produce H_2 along with TOC removal (mineralization) of aromatic pollutants in a de-aerated aqueous suspension which is considered an inappropriate condition to achieve the mineralization of organic pollutants.

Cr₂O₃/Rh/SrTiO₃ photocatalyst was prepared by step-wise photo-deposition of Rh nanoparticles on the SrTiO₃ surface as a core and then the Cr₂O₃ nano-shell on the Rh core.²² The loading amount of Rh and Cr₂O₃ was 0.5 wt% and 0.75 wt%, respectively. Fig. S1 (ESI†) shows the HRTEM images of Cr₂O₃/ Rh/SrTiO₃ and Rh/SrTiO₃ photocatalysts. Rh nanoparticles of 2–4 nm diameter were observed to be deposited on the surface of both photocatalysts and the Cr2O3 shell was seen around the Rh core of the Cr₂O₃/Rh/SrTiO₃ catalyst. Rh/SrTiO₃ and Cr₂O₃/Rh/SrTiO₃ were compared for the photocatalytic degradation of 4-CP under de-aerated conditions as shown in Fig. 1. Cr₂O₃/Rh/SrTiO₃ exhibited a much higher activity than Rh/SrTiO₃ in both the removal of 4-CP and the concurrent production of chloride (Fig. 1a). The TOC removal was also highly enhanced with Cr₂O₃/Rh/SrTiO₃ (Fig. 1b). These results clearly indicate that the removal of 4-CP in the suspension of Cr2O3/Rh/SrTiO3 proceeded along with the mineralization

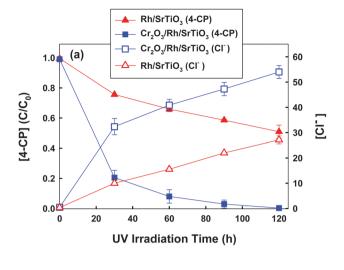
However, the total organic carbon (TOC) removal efficiency was negligible and TOC remained almost unchanged during the photocatalytic degradation of 4-chlorophenol (4-CP) on F-TiO₂/Pt, because dioxygen was needed for mineralization. Meeting the optimal condition for dual-purpose photocatalysis is contradictory because the H₂ evolution requires anoxic conditions whereas the mineralization of organic compounds needs dioxygen.

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 $[\]dagger$ Electronic supplementary information (ESI) available: Experimental details of the synthesis of Rh/SrTiO3, $Cr_2O_3/Rh/SrTiO_3$ and F-TiO2/Pt, experimental details of photocatalytic reaction and photoelectrochemical tests, and Fig. S1–S7. See DOI: 10.1039/c6cc04260k

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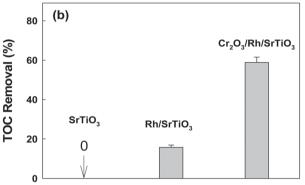
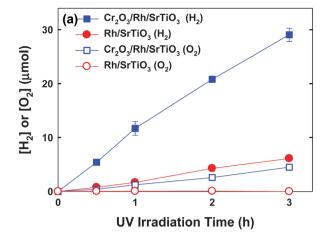


Fig. 1 (a) Photocatalytic degradation of 4-CP and the concurrent production of chloride in a de-aerated catalyst suspension. (b) Comparison of TOC removal efficiencies after 2 h photocatalytic reaction. Experimental conditions: [catalyst] = 0.5 g L $^{-1}$, [4-CP] $_0$ = 100 μ M, pH $_0$ = 7, λ > 320 nm, air-tight, and initially Ar-purged for 1 h before UV irradiation.

(*i.e.*, TOC removal) even under the de-aerated conditions, where the initial dissolved O_2 was measured to be less than 0.1 ppm using a dissolved O_2 meter. In the absence of O_2 that serves as a main electron scavenger, the photocatalytic oxidation and mineralization should be inhibited on a bare semiconductor but the loading of Rh and Cr_2O_3 changes the photocatalytic reaction mechanism.

 H_2 evolution was also monitored during the anoxic degradation of 4-CP. As seen in Fig. 2a, markedly increased H_2 production was observed with the $Cr_2O_3/Rh/SrTiO_3$ photocatalyst compared with $Rh/SrTiO_3$. The initial rate of H_2 production on $Cr_2O_3/Rh/SrTiO_3$ was around 12 µmol h^{-1} . The apparent photonic efficiency of H_2 evolution (with 300 µM 4-CP) was separately measured under the irradiation centered around $\lambda = 330 \pm 10$ nm and determined to be 0.7%. Concurrent O_2 evolution was also observed in the case of $Cr_2O_3/Rh/SrTiO_3$, whereas no O_2 production was observed with $Rh/SrTiO_3$. Such a difference can be ascribed to the fact that the Cr_2O_3 layer over Rh can suppress the back reaction of H_2 with O_2 to produce $H_2O.^{23}$ The fact that O_2 was evolved along with the degradation of 4-CP on $Cr_2O_3/Rh/SrTiO_3$ indicates that holes react with both 4-CP and water molecules. The *in situ* generated O_2 should be subsequently consumed during the



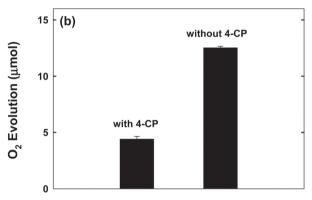


Fig. 2 (a) Time-profiled production of H_2 and O_2 in the suspension of $Cr_2O_3/Rh/SrTiO_3$ and $Rh/SrTiO_3$ with 4-CP (300 μ M) in the initially de-aerated suspension. (b) Comparison of photocatalytic O_2 evolution in the suspension of $Cr_2O_3/Rh/SrTiO_3$ (after 3 h reaction) in the presence and absence of 4-CP.

mineralization of 4-CP, 19,24,25 which explains why the degradation of 4-CP was possible under de-aerated conditions. As a result, the *in situ* O₂ evolution was significantly enhanced in the absence of 4-CP (Fig. 2b). The mineralization of 4-CP can be expressed by eqn (1).

$$C_6H_5OCl (6C_{org}) + 6.5O_2 \rightarrow 6CO_2 + 2H_2O + HCl$$
 (1)

During the initial stage (for 1 h) of photo-irradiation, the rates of $\rm H_2$ and $\rm O_2$ evolution were determined to be around 12 µmol h⁻¹ and 1.2 µmol h⁻¹, respectively. The $\rm O_2$ evolution rate is much lower than the expected stoichiometric rate (6 µmol h⁻¹) in the dual-purpose photocatalysis, which should be ascribed to the *in situ* consumption of $\rm O_2$ in the mineralization (as mentioned above). In this case, the average TOC removal rate was 5.0 µmol h⁻¹, which corresponds to 5.4 µmol h⁻¹ of the $\rm O_2$ consumption rate (according to eqn (1)). Therefore, the sum of the apparent $\rm O_2$ evolution (1.2 µmol h⁻¹) and the *in situ* consumption of $\rm O_2$ (5.4 µmol h⁻¹) is 6.6 µmol h⁻¹, which is close to the stoichiometric $\rm O_2$ evolution rate of 6.0 µmol h⁻¹. Incidentally, from a practical point of view, dual-functional photocatalysts working under air-saturated conditions would be desirable. Therefore, $\rm H_2$ evolution in an aerated suspension of $\rm Cr_2O_3/Rh/SrTiO_3$ was

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also tested. As shown in Fig. S2 (ESI†), H2 evolved even under air-saturated conditions although they were lower compared with Ar-saturated conditions. This result indicates that this composite photocatalyst can be effectively used for dualfunctional photocatalysis in both the presence and absence of dissolved O2.

The effects of dissolved O2 and the probe reagents (i.e., t-butyl alcohol (TBA) and EDTA) were further investigated to understand the photocatalytic mechanisms of Cr₂O₃/Rh/SrTiO₃ and Rh/SrTiO₃, as shown in Fig. S3 (ESI†). First, it is noted that the effects of dissolved O₂ are drastically different between Cr₂O₃/Rh/SrTiO₃ and Rh/SrTiO₃. The presence and absence of O₂ did not affect the 4-CP degradation on Cr₂O₃/Rh/SrTiO₃ at all while the degradation of 4-CP on Rh/SrTiO₃ was significant only in the presence of O₂. This observation is fully consistent with the previous reports that the Cr₂O₃ shell layer covering the Rh core blocks the contact of O_2 with the Rh core, but is still permeable to protons.^{22,23} As a result, the CB electron transfer to O2 on Cr2O3/Rh/SrTiO3 is insignificant but the CB electron transfer to protons (H⁺) is allowed with enabling the concurrent hole transfer to 4-CP. Fig. S3 (ESI†) also shows that the photocatalytic degradation of 4-CP remains unchanged in the presence of excessive TBA (as an OH radical scavenger) but is highly retarded in the presence of excessive EDTA (as a hole scavenger). 26 This supports that 4-CP degradation proceeds via a direct hole-transfer, and not an OH radical-meditated pathway.

To investigate the role of the Cr₂O₃ shell on the Rh core and its effects on the interfacial electron transfer on Cr₂O₃/Rh/SrTiO₃ and Rh/SrTiO₃, the Fe^{3+/2+} redox couple-mediated photocurrent was collected (via reactions (2) and (3)) in the UV-irradiated suspension of each catalyst.27

$$Fe^{3+} + e_{cb}^{-} \rightarrow Fe^{2+}$$
 (on photocatalyst) (2)

$$Fe^{2+} \rightarrow Fe^{3+} + e^{-}$$
 (on Pt electrode) (3)

Fig. S4 (ESI†) shows that the time profile of photocurrent generation is completely suppressed by Cr₂O₃/Rh/SrTiO₃ in comparison with Rh/SrTiO3, which is the opposite to the photocatalytic activity of H₂ evolution. This result reconfirms that the Cr₂O₃ layer on the Rh core suppresses the interfacial electron transfer to Fe³⁺ ions as it hinders the electron transfer to O2 molecules. The presence of the Cr2O3 layer blocks the CB electron transfer to electron acceptors (e.g., O₂, Fe³⁺) except protons. On the other hand, holes generated on Cr₂O₃/Rh/SrTiO₃ are mostly consumed by water molecules with generation of O2, which enables the anoxic degradation of 4-CP. As a result, the photocatalytic production of H₂ on Cr₂O₃/Rh/SrTiO₃ under the de-aerated conditions depended little on the presence and kind of organic electron donors (see Fig. S5, ESI†). The photocatalytic H₂ production rates on Cr₂O₃/Rh/SrTiO₃ only moderately changed among the different conditions of water, 10 vol% MeOH, and 300 µM 4-CP whereas the electron donor effect on H2 production was very critical for Rh/SrTiO3. This implies that the photocatalytic production of H₂ with the simultaneous decomposition of organic pollutants can be achieved effectively in the Cr₂O₃/Rh/SrTiO₃ photocatalytic system regardless of the kind and concentration of organic pollutants.

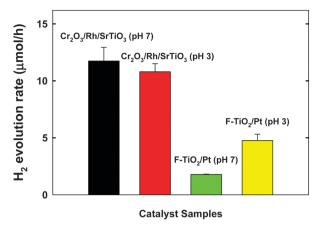
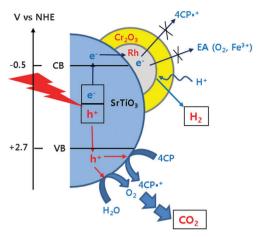


Fig. 3 Comparison of the initial photocatalytic H₂ production rate between Cr₂O₃/Rh/SrTiO₃ and F-TiO₂/Pt photocatalytic systems in the presence of 4-CP (300 μ M). Experimental conditions: [catalyst] = 0.5 g L⁻¹, $\lambda > 320$ nm, air-tight and initially Ar-purged for 1 h before UV irradiation.

Our recent studies demonstrated that the TiO₂ modified with both Pt and fluoride (F-TiO₂/Pt) exhibited a dual-functional photocatalytic activity for the simultaneous production of H₂ and degradation of 4-CP. 19,20 In Fig. 3, the H2 production on Cr₂O₃/Rh/SrTiO₃ was compared with F-TiO₂/Pt under neutral and acidic pH conditions. The photocatalytic activity of Cr2O3/ Rh/SrTiO₃ is higher than that of F-TiO₂/Pt and less affected by the pH change. It should be noted that the activity of F-TiO₂/Pt is markedly reduced at neutral pH whereas that of Cr₂O₃/Rh/SrTiO₃ was little influenced by pH. As a result, Cr2O3/Rh/SrTiO3 is a better dual-functional photocatalyst from a practical point of view. In terms of the charge transfer characteristics, the following two major features make Cr₂O₃/Rh/SrTiO₃ a practical dual-functional photocatalyst. For the electron transfer part, CB electrons are selectively consumed by protons only and their transfer to O2 and other electron acceptors (EA) is hindered because the Cr₂O₃ barrier layer is selectively permeable only to protons. On the other hand, VB holes are utilized to oxidize both H₂O (to O₂) and 4-CP (organic pollutants) simultaneously and the in situ generated O2 is immediately consumed for the mineralization of the organic pollutants. 19,24,28 The reaction mechanisms described above are schematically illustrated in Scheme 1. In the absence of the Cr₂O₃ layer, the CB electrons can be consumed by not only protons but also in situ generated O₂ and other reaction intermediates, which would reduce the overall dual-photocatalysis activity.

To check the photostability of Cr₂O₃/Rh/SrTiO₃, the photocatalytic H2 production was repeated up to four cycles in the same batch of catalyst by injecting 100 µM 4-CP every 2 h (Fig. S6, ESI†). The activity was not maintained and gradually decreased with repeated uses. To elucidate whether the gradual loss of activity was caused by the instability of the photocatalyst, the same experiment was performed without 4-CP injection. In this case, the photocatalytic activity was maintained without showing activity loss. The composite photocatalyst itself seems to be stable. Therefore, the gradual loss of photocatalytic activity observed in the presence of 4-CP might be ascribed to Communication ChemComm



Scheme 1 Schematic illustrations of photocatalytic reaction mechanisms occurring on the surface of Cr₂O₃/Rh/SrTiO₃.

the accumulation of organic degradation intermediates on the catalyst surface. 29,30 To further investigate the effect of organic compounds in this dual functional photocatalysis, the evolution of H2 and O2 was simultaneously measured with repeated photocatalysis cycles. In this case, 100 µM 4-CP was initially added but not replenished in the subsequent cycles. Fig. S7 (ESI†) shows that H₂ production was higher in the first cycle than in the subsequent cycles: the difference in H₂ production should be ascribed to the organic electron donor (i.e., 4-CP) effect. The extra holes scavenged by 4-CP make an equal number of electrons to be used for H₂ production. At the same time, O₂ evolved in the first cycle is immediately consumed for the mineralization of 4-CP. As a result, the ratio of H₂ to O₂ in the first cycle was significantly higher $(H_2/O_2(r) = 6.9)$ than the stoichiometric water splitting ratio (r = 2.0). The ratio progressively approached the stoichiometric ratio as the cycle was repeated $(r: 6.9 \rightarrow 2.7 \rightarrow 2.5 \rightarrow 2.2).$

In conclusion, this study demonstrated that the mineralization of organic pollutants can be achieved under the de-aerated conditions with the simultaneous H₂ production over a Cr₂O₃/ Rh/SrTiO₃ photocatalyst. The present study showed the highest H₂ evolution efficiency in dual-purpose photocatalysis to our knowledge. It is proposed that the Cr₂O₃ shell on the Rh nanoparticle core markedly enhances the H2 production and TOC removal of aromatic pollutants, which makes Cr₂O₃/Rh/SrTiO₃ an active dual-functional photocatalyst.

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References

- 1 G. Zhang, C. Ni, X. Huang, A. Welgamage, L. A. Lawton, P. K. J. Robertson and J. T. S. Irvine, Chem. Commun., 2016, 52, 1673.
- 2 M. Ni, D. Y. C. Leung and M. K. H. Leung, Int. J. Hydrogen Energy, 2007, 32, 3238-3247.
- 3 F. Sastre, M. Oteri, A. Corma and H. Garcia, Energy Environ. Sci., 2013, 6, 2211.
- 4 Y. Tachibana, L. Vayssieres and J. R. Durrant, Nat. Photonics, 2012, 6, 511.
- A. Kudo and Y. Miseki, Chem. Soc. Rev., 2009, 38, 253.
- 6 Y. Ma, X. Wang, Y. Jia, X. Chen, H. Han and C. Li, Chem. Rev., 2014,
- 7 T. Hisatomi, J. Kubota and K. Domen, Chem. Soc. Rev., 2014, 43, 7520.
- 8 H. Park, Y. Park, W. Kim and W. Choi, J. Photochem. Photobiol., C, 2013, 15, 1.
- 9 M. R. Hoffmann, S. T. Martin, W. Choi and D. W. Bahnemann, Chem. Rev., 1995, 95, 69.
- 10 Q. Xiang, J. Yu and M. Jaroniec, Chem. Soc. Rev., 2012, 41, 782.
- 11 J. Yang, R. Hu, W. Meng and Y. Du, Chem. Commun., 2016, 52,
- 12 H. Park, H.-i. Kim, G.-h. Moon and W. Choi, Energy Environ. Sci., 2016, 9, 411.
- 13 X.-Y. Zhang, H.-P. Li, X.-L. Cui and Y. Lin, J. Mater. Chem., 2010, 20, 2801.
- 14 M.-C. Wu, J. Hiltunen, A. Sápi, A. Avila, W. Larsson, H.-C. Liao, M. Huuhtanen, G. Tóth, A. Shchukarev, N. Laufer, A. Kukovecz, Z. Kónya, J.-P. Mikkola, R. Keiski, W.-F. Su, Y.-F. Chen, H. Jantunen, P. M. Ajayan, R. Vajtai and K. Kordás, ACS Nano, 2011, 5, 5025.
- 15 P. Gomathisankar, K. Hachisuka, H. Katsumata, T. Suzuki, K. Funasaka and S. Kaneco, Int. J. Hydrogen Energy, 2013, 38, 11840.
- 16 W. Zhang, Y. Wang, Z. Wang, Z. Zhong and R. Xu, Chem. Commun., 2010, 46, 7631.
- 17 M. Wen, Y. Kuwahara, K. Mori, D. Zhang, H. Li and H. Yamashita, J. Mater. Chem. A, 2015, 3, 14134.
- 18 Y.-J. Cho, H.-i. Kim, S. Lee and W. Choi, J. Catal., 2015, 330, 387.
- 19 J. Kim and W. Choi, Energy Environ. Sci., 2010, 3, 1042.
- 20 J. Kim, D. Monllor-Satoca and W. Choi, Energy Environ. Sci., 2012, 5, 7647.
- 21 K. Maeda, A. Xiong, T. Yoshinaga, T. Ikeda, N. Sakamoto, T. Hisatomi, M. Takashima, D. Lu, M. Kanehara, T. Setoyama, T. Teranishi and K. Domen, Angew. Chem., Int. Ed., 2010, 49, 4096.
- 22 K. Maeda, K. Teramura, D. Lu, N. Saito, Y. Inoue and K. Domen, Angew. Chem., Int. Ed., 2006, 45, 7806.
- 23 M. Yoshida, K. Takanabe, K. Maeda, A. Ishikawa, J. Kubota, Y. Sakata, Y. Ikezawa and K. Domen, J. Phys. Chem. C, 2009, 113, 10151.
- 24 D. Hufschmidt, D. Bahnemann, J. J. Testa, C. A. Emilio and M. I. Litter, J. Photochem. Photobiol., A, 2002, 148, 223.
- 25 J. Kim, J. Lee and W. Choi, Chem. Commun., 2008, 756.
- 26 C. Minero, G. Mariella, V. Maurino, D. Vione and E. Pelizzetti, Langmuir, 2000, 16, 8964.
- 27 H. Park and W. Choi, J. Phys. Chem. B, 2003, 107, 3885.
- 28 J. Theurich, M. Lindner and D. W. Bahnemann, Langmuir, 1996, 12, 6368.
- 29 M. I. Franch, J. Peral, X. Domenech and J. A. Ayllon, Chem. Commun., 2005, 1851,
- 30 S. Weon and W. Choi, Environ. Sci. Technol., 2016, 50, 2556.