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# Facile synthesis of fluorescent hetero[8]circulene analogues with tunable solubilities and optical properties†

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Hetero[8]circulenes are an interesting class of polycyclic heteroaromatic molecules having rigid and planar structures, which are promising in light of their potential applications for OLEDs, OFETs and so forth. Although their synthetic methods have been developed in some specific cases, a facile synthetic protocol of novel hetero[8]circulenes with tunable properties is highly desirable. We herein report the unexpected formation of methoxy-substituted quasi-aza[8]circulene and its conversion into unprecedented triazaoxa[8]circulene. The structures and optical properties were comparatively studied. Remarkably, triazaoxa[8]circulene is highly soluble in THF, acetone and DMSO mainly because of effective hydrogen-bonding of the NH moieties to these solvents. Their highly soluble nature in various solvents enabled us to study the solvent effects of these molecules. In particular, triazaoxa[8]circulene displays a high fluorescence quantum yield of 0.72 in DMSO. Furthermore, enantiomeric separation of highly distorted quasi-aza[8]circulene was successfully achieved by chiral HPLC. Thus, these novel hetero[8]circulene derivatives are practically useful fluorescent nanographene-like molecules with intriguing optical properties.

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## Introduction

Hetero[8]circulenes belong to a class of polycyclic heteroaromatic molecules (PHAs) possessing eight alternately arranged condensed arenes and/or heteroles that surround a central eight-membered ring.<sup>1–3</sup> Unlike the parent all-benzene analogue (*i.e.* [8]circulene) adopting a saddle-shaped conformation,<sup>4</sup> most of the hetero[8]circulenes take planar structures and have attracted considerable attention as promising materials for organic electronics.<sup>5–10</sup> One of the old examples is tetraoxa[8]circulene **1** reported by Erdtman *et al.* in 1968,<sup>5</sup> which was synthesized by acid-catalyzed condensation of  $\alpha$ -naphthoquinone although its formation was only identified by mass analysis due to its extremely poor solubility (Chart 1). Several tetraoxa[8]circulene derivatives were reported later on,<sup>6</sup> but it was not until 2013 that hetero[8]circulenes including units other than furan units were prepared. Azatrioxa[8]circulene **2** and diazadioxo[8]circulene **3** were synthesized by Pittelkow *et al.* by using hydroxy-substituted carbazole precursors under acidic conditions.<sup>7</sup> Recently, we reported the first synthesis of tetraaza

[8]circulene **5** by a fold-in type oxidative fusion reaction of *ortho*-phenylene-bridged cyclic tetrapyrrole **4**.<sup>8a</sup> From octabromotetraphenylene **6**, Wong *et al.* synthesized tetrathia- and tetraselena[8]circulenes **7** and **8** by Li-mediated nucleophilic reactions followed by thermal annealing.<sup>9a</sup> Miyake and Shinokubo utilized tetraboryltetrathienylene **9** as a precursor of tetrathia[8]circulenes **10**.<sup>10</sup> Interestingly, multiply alkylated tetrathia[8]circulenes form aggregates in CDCl<sub>3</sub> depending on the number of peripheral substituents.<sup>10b</sup> Despite this progress, all these reactions required specific synthetic precursors or harsh reaction conditions, precluding the fine property tunability of [8]circulenes. A facile synthetic approach toward unprecedented hetero[8]circulenes is thus highly desirable.

Recently, we reported the synthesis of triaza-quasi-[8]circulene **12a**,<sup>11</sup> which lacks one heterole moiety from the hetero[8]circulene structure.<sup>12</sup> Due to the structural constraint, **12a** takes on a non-planar geometry. Interestingly, under milder conditions, triaza[7]helicene **13a** was exclusively obtained from the identical precursor **11a**. Dimethyl-substituted triaza[7]helicene **13b** was then synthesized to estimate a helicene racemization activation free energy of 23.8 kcal mol<sup>-1</sup> (at 298.15 K). On the other hand, the racemization rate of **12a** was computed to be too fast for experimental determination.

Herein, we report an unexpected formation of dimethoxy-substituted triaza-quasi-[8]circulene **12c** and its conversion to the hitherto unknown triazaoxa[8]circulene **14**. The planar and non-substituted triazaoxa[8]circulene **14** was surprisingly well soluble in polar solvents such as THF and acetone. The solvent-

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† Electronic supplementary information (ESI) available: Experimental and computational details and X-ray crystallographic data for **12c**, **13d**, **14**, **12a-Bu**, **12c-Bu** and **14-Bu** are available. CCDC 1953592–1953597. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c9sc05087f



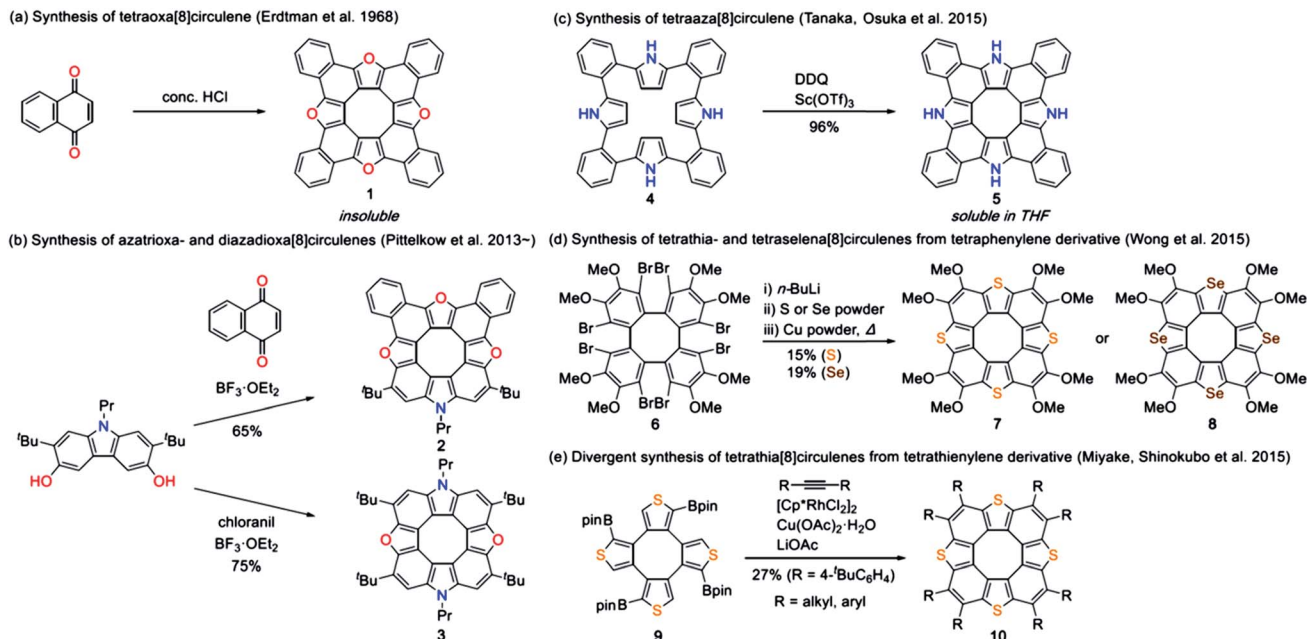


Chart 1 (a) Synthesis of tetraoxa[8]circulene from  $\alpha$ -naphthoquinone. (b) Synthesis of azatrioxa[8]circulene and diazadioxo[8]circulene from a dihydroxycarbazole precursor. (c) Synthesis of tetraaza[8]circulene from *o*-phenylene-bridged tetrapyrrole. (d) Synthesis of tetrathia- and tetraselena[8]circulenes from an octabromotetraphenylene precursor. (e) Synthesis of tetra[8]circulenes from a tetraboryltetrathienylene precursor (DDQ = 2,3-dichloro-5,6-dicyano-*p*-benzoquinone, chloranil = 2,3,5,6-tetrachloro-*p*-benzoquinone).

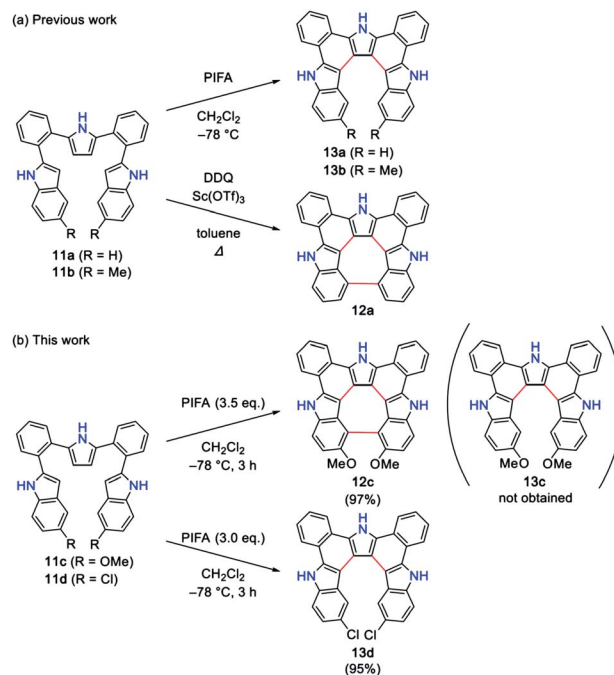
dependent optical properties will be discussed in conjunction with further *N*-substituted derivatives.

intermolecular  $\pi$ - $\pi$  stacking is not favorable. In the crystal lattice, two or three THF molecules are encapsulated and form hydrogen bonds with pyrrolic NH sites.

## Results and discussion

### Synthesis

The synthesis of triaza-quasi-[8]circulene is straightforward; the coupling reaction of indole derivatives with 2,5-bis(2-bromophenyl)pyrrole followed by oxidation under appropriate conditions.<sup>12</sup> This time, we synthesized a dimethoxy-substituted precursor **11c** as a potential cyclization substrate leading to dimethoxy-substituted triaza[7]helicene **13c**. However, even under milder oxidative reaction conditions (PIFA at  $-78$  °C), quasi-[8]circulene **12c** was produced in 97% yield and **13c** was not detected (Scheme 1). This result has been ascribed to the higher spin density at the 4-position than at the 3-position in the cation radical of **11c** (see Fig. S5-1, ESI†). Therefore, intramolecular 4,4'-oxidative coupling in **11c** proceeded preferably in spite of the steric repulsion imposed by the two methoxy groups. In contrast, the spin density distribution is reversed in the case of the dichloro-substituted precursor **11d**, which preferred oxidative coupling between the central pyrrole and indole 3-position to give triaza[7]helicene **13d** in 95% yield. The structures of **12c** and **13d** were unambiguously revealed by X-ray diffraction analysis (Fig. 1). Both compounds displayed highly twisted structures, the degree of which can be quantified by mean-plane deviation (MPD) values to be 0.45 Å for **12c** and 0.55 Å for **13d** (35 atoms). The value of **12c** is distinctly larger than that of **12a** (0.31 Å) due to the two methoxy groups. Because of the twisted structures,



Scheme 1 (a) Selective synthesis of triaza-quasi-[8]circulene **12a** and triaza[7]helicenes **13a,b**. (b) Unexpected formation of triaza-quasi-[8]circulene **12c** under mild conditions, and the formation of triaza[7]helicene **13d** under the same conditions (PIFA = bis(trifluoroacetoxy)iodobenzene).



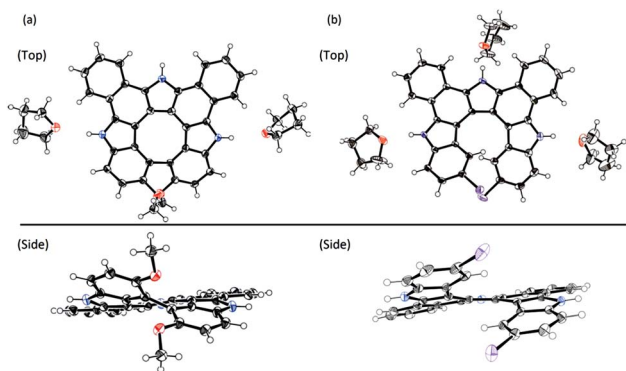
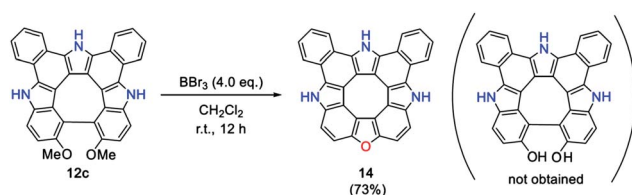


Fig. 1 X-ray crystal structures of (a) **12c** and (b) **13d**. The thermal ellipsoids were scaled to 50% probability level. Solvent molecules were omitted in the side views.

It is well known that biphenylene-2,2'-diol derivatives undergo thermal dehydration reaction to afford the corresponding dibenzofuran derivatives.<sup>13</sup> We attempted demethylation of **12c** with  $\text{BBr}_3$  in  $\text{CH}_2\text{Cl}_2$  at room temperature, which gave triazaoxa[8]circulene **14** as a yellow solid in 73% yield (Scheme 2). Matrix-assisted-laser-desorption-ionization time-of-flight mass spectrometry (MALDI-TOF-MS) revealed the molecular ion peak of **14** at  $m/z = 457.1218$  (calcd for  $\text{C}_{32}\text{H}_{15}\text{N}_3\text{O} = 457.1210$   $[\text{M}]^+$ ). The  $^1\text{H}$  NMR spectrum in  $\text{DMSO}-d_6$  showed two NH peaks at 12.93 and 12.64 ppm and six peaks at 9.09, 8.82, 7.98, 7.96, 7.83 and 7.77 ppm. Even at low temperature, the corresponding diol intermediate was not isolated, presumably because of the fast dehydration reaction in the diol intermediate driven by strain release. It is noteworthy that this is the first example of hetero[8]circulene consisting of three pyrrole units and one furan unit. All the reaction steps are high yielding and the reaction products are easily separable, thus allowing us to obtain **14** in more than 300 mg quantity. In spite of the lack of solubilizing substituents, **14** is quite soluble in acetone and THF but less soluble in  $\text{CH}_2\text{Cl}_2$ . The single crystal structure of **14** is shown in Fig. 2. The molecule is almost planar with a MPD value of  $0.014 \text{ \AA}$  (35 atoms). Three acetone molecules adopt hydrogen bonding with the three NH sites, and the layer of three acetone molecules prevents the planar core molecule from intermolecular  $\pi$ - $\pi$  contact. The contribution of the central formal  $8\pi$  cyclooctatetraene moiety is another point to discuss. The constituent eight C-C bond lengths are in the range of  $1.404(2)$ – $1.431(2) \text{ \AA}$  (see Fig. S6-5, ESI<sup>†</sup>). The NICS(0) and NICS(1)<sub>zz</sub> values at the center of the eight-membered ring were calculated by the GIAO method to be +7.00 and +18.05 ppm,



Scheme 2 Synthesis of triazaoxa[8]circulene **14**.

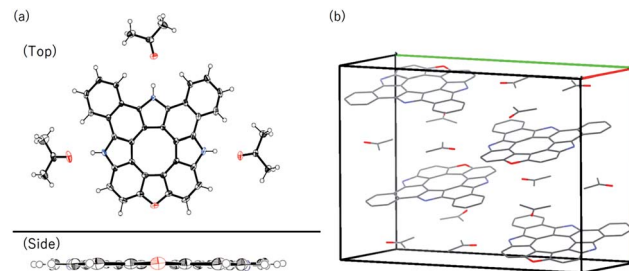


Fig. 2 (a) X-ray crystal structures of **14**. The thermal ellipsoids were scaled to 50% probability level. Solvent molecules in the side view were omitted for clarity. (b) Packing diagram.

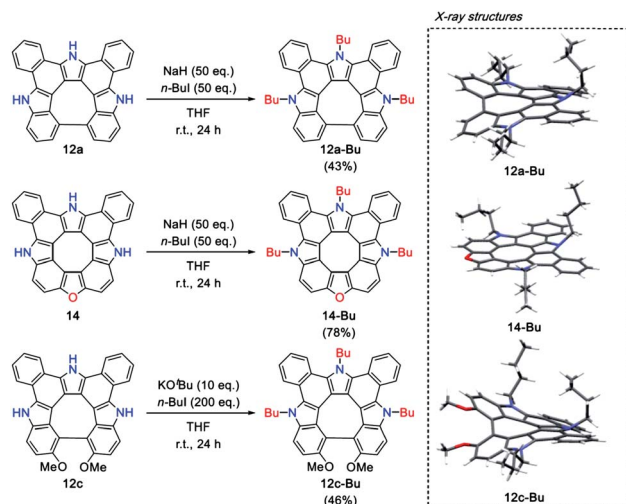
respectively (Fig. S5-12 and S5-13, ESI<sup>†</sup>). This situation is in analogy to other hetero[8]circulenes.<sup>7,8</sup> Although it may have decent antiaromatic contributions, we would say that the aromatic contributions of each benzene, pyrrole and furan moiety are predominant as indicated by HOMA and NICS(0) values at the centre points of each aromatic ring (see Fig. S5-14, ESI<sup>†</sup>).

### N-substitution reactions

An advantage of the new synthetic scheme for **14** against other [8]circulene syntheses may be the presence of the intact NH moieties which can be further decorated. Previously, we reported that sequential *N*-butylation of tetraaza[8]circulene **5** slightly altered its optical properties.<sup>14</sup> Along with this strategy, *N*-alkylation reactions were tested for **12a**, **12c** and **14**. *N*-Butylation was easily achieved for **12a** and **14** by utilizing the reported method (excess NaH, 1-iodobutane in THF) to give tributylated products **12a-Bu** and **14-Bu** in 43% and 78% yields, respectively.<sup>14,15</sup> On the other hand, the same reaction of **12c** resulted in a very low yield of *N*-butylated products. It is likely that the highly twisted structure of **12c** causes degradation under these conditions. Then, the reaction conditions were screened, and we found that the treatment of **12c** with 200 equivalents of 1-iodobutane in the presence of potassium *tert*-butoxide at room temperature for 24 hours afforded **12c-Bu** in 46% yield. All the solid-state structures have been revealed by X-ray diffraction analysis as shown in Scheme 3. The planarity and bond-length alternation do not change so much for **14-Bu**, while the degree of non-planarity of **12a-Bu** and **12c-Bu** increased compared with **12a** and **12c**.

With these in hand, we next investigated their aggregation behaviours in solution. Basically, all these hetero[8]circulene analogues are quite soluble in THF and the absorption spectra obey the Beer-Lambert law in the UV/vis concentration (*ca.*  $10^{-5}$ – $10^{-6} \text{ M}$ ) (see Fig. S7-1 and S7-2, ESI<sup>†</sup>). The tri-*N*-butylated products were more soluble in less polar solvents as compared with non-substituted ones. For instance, the solubilities of **14** and **14-Bu** were roughly estimated by a simple filtration experiment to be  $0.15 \text{ mg mL}^{-1}$  in  $\text{CH}_2\text{Cl}_2$  for **14** and  $26 \text{ mg mL}^{-1}$  in  $\text{CH}_2\text{Cl}_2$  for **14-Bu** at room temperature, while those in acetone were comparable ( $4.6 \text{ mg mL}^{-1}$  for **14** and  $5.1 \text{ mg mL}^{-1}$  for **14-Bu**) (see Table S9-1, ESI<sup>†</sup>). In addition, **12a-Bu**, **12c-Bu** and **14-Bu** are even soluble in *n*-hexane. On the other hand, **12a**, **12c** and **14**





Scheme 3 N-substitution reactions and the X-ray crystal structures of 12a-Bu, 12c-Bu and 14-Bu. Solvent molecules were omitted for clarity.

bearing intact NH sites are especially soluble in hydrogen-bond accepting solvents such as THF, acetone and DMSO. They are also soluble in a mixed solvent, up to 10% THF in *n*-hexane, without the observation of aggregation-induced fluorescence quenching (Fig. S7–9<sup>†</sup>). In the solid state, the emission spectra were completely quenched.

Nanographenes are, in many cases, only soluble in halogenated solvents, and the low solubility often hampers further functionalization and applications.<sup>16</sup> The high solubility can be a unique feature of hetero[8]circulenes possessing intact NH sites, and is useful for solution-processed electronics applications and is potentially applicable for biochemistry.<sup>17</sup>

### Optical properties

Hetero[8]circulenes are highly luminescent molecules upon photoirradiation because of their rigid structures that prevent thermal energy loss due to structural change in the excited state. In fact, tetraaza[8]circulene **5** showed a very small Stokes shift

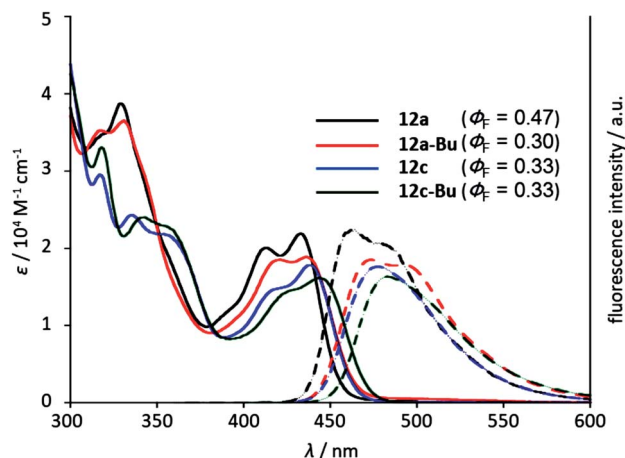


Fig. 3 UV/vis absorption and emission spectra of 12a, 12a-Bu, 12c and 12c-Bu in THF.

(ca. 180  $\text{cm}^{-1}$ ) in THF with a high fluorescence quantum yield ( $\Phi_F = 0.59$ ). The absorption and emission spectra of quasi-aza [8]circulene **12a** are red-shifted with a relatively larger Stokes shift (1450  $\text{cm}^{-1}$ ) probably due to their non-planar conformations. The absorption spectrum of **12c** is further shifted up to 474 nm compared with that of **12a** (433 nm) (Fig. 3 and Table 1). DFT calculations revealed that the larger non-planarity of **12c** raised its HOMO energy level, giving rise to the decreased HOMO–LUMO gap (see Fig. S5–8, ESI<sup>†</sup>). Instead, the emission quantum yield was decreased to  $\Phi_F = 0.33$ . Upon *N*-butylation, the absorption and emission spectra were red-shifted. While **12a-Bu** and **14-Bu** showed decreased  $\Phi_F$  values from non-butylated ones, **12c** and **12c-Bu** displayed the same  $\Phi_F$  values in THF.

Triaza[8]circulene **14** has a rigid and planar structure like **5**, but exhibited red-shifted absorption and emission bands (Fig. 4). The decrease of the molar extinction coefficient is ascribed to its lower symmetry, while the red-shift may be due to the lack of two benzo-fused segments. Azatrioxa[8]circulene **2**

Table 1 Summary of the properties

Compound	Solvent	$\lambda_{\text{abs}}$ (nm)	$\lambda_{\text{flu}}$ (nm)	Stokes shift ( $\text{cm}^{-1}$ )	$\Phi_F$	$\tau_F$ (ns)	$k_r$ ( $\times 10^8 \text{ s}^{-1}$ )	$k_{\text{nr}}$ ( $\times 10^8 \text{ s}^{-1}$ )
<b>5</b> <sup>a</sup>	THF	413	416	180	0.59	3.8	1.6	1.1
<b>12a</b>	THF	433	462	1450	0.47	5.6	0.83	0.94
<b>12a</b>	DMSO	436	490	2530	0.58	8.2	0.71	0.51
<b>12a-Bu</b>	THF	436	473	1790	0.30	5.1	0.59	1.4
<b>12a-Bu</b>	DMSO	435	496	2830	0.42	6.7	0.62	0.87
<b>12c</b>	THF	426	476	1680	0.33	3.6	0.92	1.8
<b>12c</b>	DMSO	444	485	1900	0.49	4.6	1.1	1.1
<b>12c-Bu</b>	THF	443	491	1870	0.33	4.0	0.82	1.7
<b>12c-Bu</b>	DMSO	447	494	2130	0.42	5.0	0.84	1.2
<b>14</b>	THF	426	437	590	0.52	4.5	1.2	1.1
<b>14</b>	DMSO	429	449	1040	0.72	5.7	1.3	0.50
<b>14-Bu</b>	THF	440	451	550	0.40	4.1	0.97	1.5
<b>14-Bu</b>	DMSO	442	460	890	0.70	4.8	1.5	0.62

<sup>a</sup> See ref. 8a.



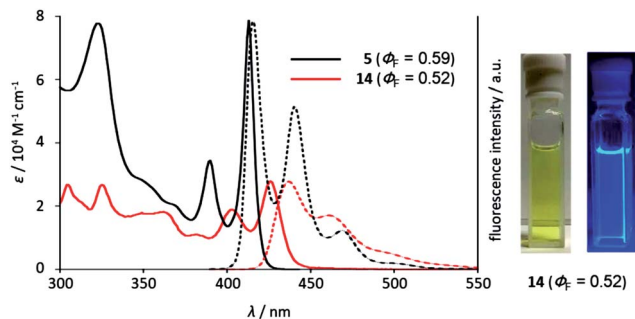


Fig. 4 UV/vis absorption and emission spectra of **5** and **14** in THF. The right figures show the photos (under room light and under UV light).

also showed a similar shift compared with a non-benzo-fused analogue.<sup>7a</sup> This tendency is consistent with theoretical calculations (see Fig. S5-10, ESI<sup>†</sup>).<sup>18</sup>

As described above, **14-Bu** is quite soluble both in polar and non-polar solvents without any aggregation effects. Thus, the absorption and emission spectra were recorded in various solvents as shown in Fig. 5. The absorption spectra are almost unchanged in various solvents with clear vibronic structures, while the emission spectra are slightly but gradually red-shifted as the polarity of the solvent increases. Interestingly, the emission is largely enhanced up to  $\Phi_F = 0.70$  in DMSO. This tendency is also observed for **14** (Table 1 and Fig. S7-5<sup>†</sup>). The fluorescence lifetimes ( $\tau_F$ ) were measured using the time-correlated single photon counting (TCSPC) technique. The longest lifetimes were observed in DMSO in all cases, but mostly the  $\tau_F$  values are in the range of 3.0–7.0 ns. From these optical parameters, the radiative rate constant ( $k_r$ ) and non-radiative rate constant ( $k_{nr}$ ) were estimated. In **14** and **14-Bu**, larger  $k_r$  and

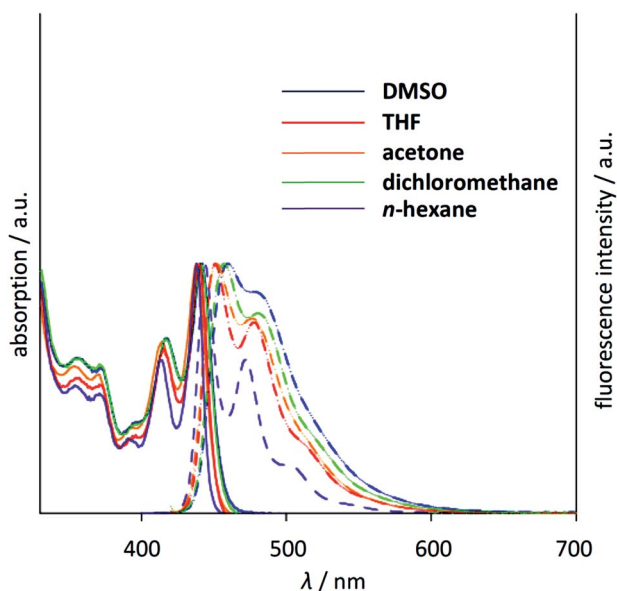


Fig. 5 Normalized UV/vis absorption (solid lines) and emission (dashed lines) spectra of **14-Bu** in *n*-hexane (purple), THF (red), CH<sub>2</sub>Cl<sub>2</sub> (green), acetone (orange) and DMSO (blue).

smaller  $k_{nr}$  both contributed to higher emission in DMSO.<sup>19</sup> For others, the larger  $\Phi_F$  values in polar solvents were mainly attributed to smaller  $k_{nr}$  values. Thus, deactivation processes in the excited state can be somehow suppressed in polar solvents.

### Chiral separation

The highly twisted structure of quasi-aza[8]circulene **12c** prompted us to examine its enantiomeric separation. Although the racemization activation barrier of **12a** was computed to be about 6.0 kcal mol<sup>-1</sup>,<sup>12</sup> that of **12c** was estimated to be more than 40 kcal mol<sup>-1</sup> (see Fig. S10-2, ESI<sup>†</sup>). Through screening of the chiral stationary phase and eluents, we found that **12c-Bu** was successfully separated by HPLC equipped with a CHIRALPAK IA-3 ( $\varphi = 4.6$  mm, Daicel Co.) with CH<sub>2</sub>Cl<sub>2</sub> as an eluent. The enantiomers were completely separated and the circular dichroism (CD) spectra showed a mirror-image pattern (Fig. 6). The first fraction which showed a negative Cotton effect at around 450 nm was assigned as the (*M*) twisted form by comparing with the calculated CD spectrum (see Fig. S10-3, ESI<sup>†</sup>). This CD spectral feature resembles that of quasi-[7]circulene derivatives,<sup>20</sup> but is reversed when compared with aza-helicene derivatives.<sup>12,21</sup> The enantiomers were quite stable and no racemization took place under heating in toluene. Since **12c-Bu** is soluble in various solvents, this may be useful in chiral recognition in suitable media. While Wong *et al.* reported the

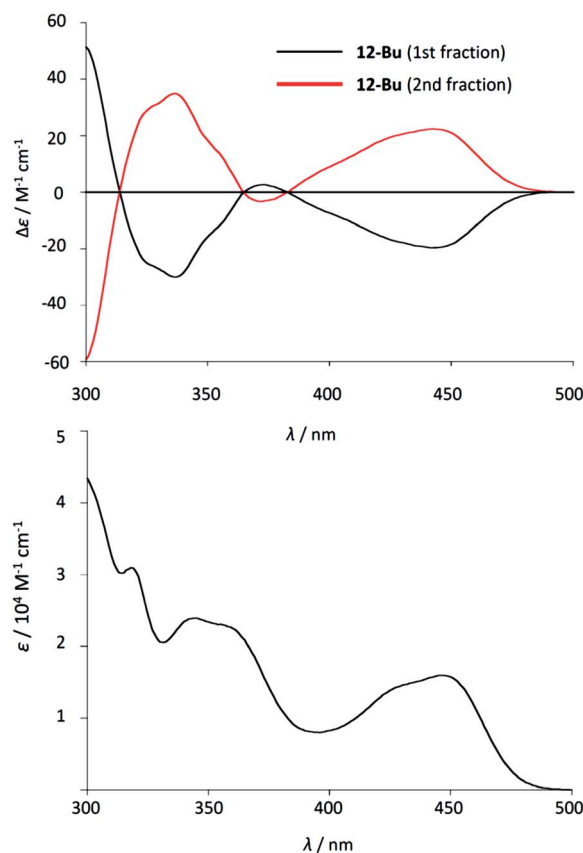


Fig. 6 (Top) CD and (bottom) UV/vis absorption spectra of the first fraction (black) and the second fraction (red) of **12c-Bu** in CH<sub>2</sub>Cl<sub>2</sub>.



chiral properties of tetraphenylene-based molecules,<sup>22</sup> enantiomeric separation of polycondensed nanographene-like molecules other than helicenes is quite rare.<sup>23</sup>

## Conclusions

In summary, non-planar quasi-aza[8]circulene derivatives and planar triazaoxa[8]circulene were prepared in good yields *via* a facile synthetic protocol. These hetero[8]circulenes are highly soluble in organic solvents without solubilizing substituents due to the hydrogen bonding interaction between the outer-pointing NH site and hydrogen-bond accepting solvent. Furthermore, *N*-alkylation reactions facilitated tuning of the solubility and optical properties. Taking advantage of the non-aggregation behaviours in various solvents, we studied the solvent effects on the absorption and emission spectra as well as emission lifetime. As an important finding, the fluorescence quantum yields were enhanced in DMSO mostly due to the suppression of non-radiative decay from the excited state. Among them, triazaoxa[8]circulenes **14** and **14-Bu** showed high fluorescence quantum yields of 0.72 and 0.70 in DMSO, respectively. Enantiomeric separation of **12c-Bu** was also successfully achieved using the chiral HPLC technique as the first chiral quasi-[8]circulene derivative. Further exploration of the synthesis and characterization of novel hetero[8]circulenes is in progress in our laboratory.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

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## Notes and references

- (a) *Carbon-Rich Compounds*, ed. M. M. Haley and R. R. Tykwinski, Wiley-VCH, Weinheim, 2006; (b) *Functional Organic Materials*, ed. T. J. J. Müller and U. H. F. Bunz, Wiley-VCH, Weinheim, 2007; (c) *Polycyclic Arenes and Heteroarenes*, ed. Q. Miao, Wiley-VCH, New York, 2015.
- (a) A. Narita, X.-Y. Wang, X. Feng and K. Müllen, *Chem. Soc. Rev.*, 2015, **44**, 6616; (b) M. Stępień, E. Gońka, M. yla and N. Sprutta, *Chem. Rev.*, 2017, **117**, 3479; (c) M. A. Majewski and M. Stępień, *Angew. Chem., Int. Ed.*, 2019, **58**, 86; (d) M. Grzybowski, B. Sadowski, H. Butenschön and D. T. Gryko, *Angew. Chem., Int. Ed.*, 2019, **58**, DOI: 10.1002/anie.201904934.
- (a) M. Stępień, *Synlett*, 2013, **24**, 1316; (b) C.-N. Feng, Y.-C. Hsieh and Y.-T. Wu, *Chem. Rec.*, 2015, **15**, 266; (c) T. Hensel, N. N. Andersen, M. Plesner and M. Pittelkow, *Synlett*, 2016, **27**, 498.
- (a) C.-N. Feng, M.-Y. Kuo and Y.-T. Wu, *Angew. Chem., Int. Ed.*, 2013, **52**, 7791; (b) Y. Sakamoto and T. Suzuki, *J. Am. Chem. Soc.*, 2013, **135**, 14074; (c) R. W. Miller, A. K. Duncan, S. T. Schneebeli, D. L. Gray and A. C. Whalley, *Chem.–Eur. J.*, 2014, **20**, 3705; (d) C.-N. Feng, W.-C. Hsu, J.-Y. Li, M.-Y. Kuo and Y.-T. Wu, *Chem.–Eur. J.*, 2016, **22**, 9198.
- H. Erdtman and H.-E. Högberg, The first synthetic attempt can be further traced back to 1881 when Knapp and Schultz reported that  $\alpha$ -naphthoquinone was reacted with concentrated hydrochloric acid to give insoluble material. Erdtman also synthesized the same material in 1933. However, the product had been mis-assigned and it was finally proved to be **1** in 1968, *Chem. Commun.*, 1968, 773.
- (a) H. Erdtman and H.-E. Högberg, *Tetrahedron Lett.*, 1970, **11**, 3389; (b) J.-E. Berg, H. Erdtman, H.-E. Högberg, B. Karlsson, A.-M. Pilotti and A.-C. Söderholm, *Tetrahedron Lett.*, 1977, **18**, 1831; (c) H. Erdtman and H.-E. Högberg, *Tetrahedron*, 1979, **35**, 535; (d) J. Eskildsen, T. Reenberg and J. B. Christensen, *Eur. J. Org. Chem.*, 2000, 1637; (e) R. Rathore and S. H. Abdelwahed, *Tetrahedron Lett.*, 2004, **45**, 5267; (f) C. B. Nielsen, T. Brock-Nannestad, T. K. Reenberg, P. Hammershøj, J. B. Christensen, J. W. Stouwdam and M. Pittelkow, *Chem.–Eur. J.*, 2010, **16**, 13030; (g) T. Brock-Nannestad, C. B. Nielsen, M. Schau-Magnussen, P. Hammershøj, T. K. Reenberg, A. B. Petersen, D. Trpceviski and M. Pittelkow, *Eur. J. Org. Chem.*, 2011, 6320; (h) S. L. Mejlsøe and J. B. Christensen, *J. Heterocycl. Chem.*, 2014, **51**, 1051.
- (a) C. B. Nielsen, T. Brock-Nannestad, P. Hammershøj, T. K. Reenberg, M. Schau-Magnussen, D. Trpceviski, T. Hensel, R. Salcedo, G. V. Baryshnikov, B. F. Minaev and M. Pittelkow, *Chem.–Eur. J.*, 2013, **19**, 3898; (b) T. Hensel, D. Trpceviski, C. Lind, R. Grosjean, P. Hammershøj, C. B. Nielsen, T. Brock-Nannestad, B. E. Nielsen, M. Schau-Magnussen, B. Minaev, G. V. Baryshnikov and M. Pittelkow, *Chem.–Eur. J.*, 2013, **19**, 17097; (c) M. Plesner, T. Hensel, B. E. Nielsen, F. S. Kamounah, T. Brock-Nannestad, C. B. Nielsen, C. G. Tortzen, O. Hammerich and M. Pittelkow, *Org. Biomol. Chem.*, 2015, **13**, 5937.
- (a) F. Chen, Y. S. Hong, S. Shimizu, D. Kim, T. Tanaka and A. Osuka, *Angew. Chem., Int. Ed.*, 2015, **54**, 10639; (b) Y. Nagata, S. Kato, Y. Miyake and H. Shinokubo, *Org. Lett.*, 2017, **19**, 2718.
- (a) X. Xiong, C.-L. Deng, B. F. Minaev, G. V. Baryshnikov, X.-S. Peng and H. N. C. Wong, *Chem.–Asian J.*, 2015, **10**, 969; (b) X. Xiong, C.-L. Deng, Z. Li, X.-S. Peng and H. N. C. Wong, *Org. Chem. Front.*, 2017, **4**, 682.
- (a) S. Kato, Y. Serizawa, D. Sakamaki, S. Seki, Y. Miyake and H. Shinokubo, *Chem. Commun.*, 2015, **51**, 16944; (b) S. Kato, S. Akahori, Y. Serizawa, X. Lin, M. Yamauchi, S. Yagai, T. Sakurai, W. Matsuda, S. Seki, H. Shinokubo and Y. Miyake, *J. Org. Chem.*, 2019, DOI: 10.1021/acs.joc.9b01655.
- A. Rajca, M. Miyasaka, S. Xiao, P. J. Boratynski, M. Pink and S. Rajca, Another expression of **13a** is dibenzohepta[8]circulene as previously reported,<sup>12,19</sup> but here we use the



- term, quasi-[8]circulene to emphasize the resemblance as referred to the Rajca's report, *J. Org. Chem.*, 2009, **74**, 9105.
- 12 F. Chen, T. Tanaka, T. Mori and A. Osuka, *Chem.–Eur. J.*, 2018, **24**, 7489.
- 13 (a) K. Nakanishi, D. Fukatsu, K. Takaishi, T. Tsuji, K. Uenaka, K. Kuramochi, T. Kawabata and K. Tsubaki, *J. Am. Chem. Soc.*, 2014, **136**, 7101; (b) M. S. Sundar and A. V. Bedekar, *Org. Lett.*, 2015, **17**, 5808; (c) C. Mitsui, M. Yamagishi, R. Shikata, H. Ishii, T. Matsushita, K. Nakahara, M. Yano, H. Sato, A. Yamano, J. Takeya and T. Okamoto, *Bull. Chem. Soc. Jpn.*, 2017, **90**, 931; (d) T. Okamoto, H. Dosei, M. Mitani, Y. Murata, H. Ishii, K.-i. Nakamura, M. Yamagishi, M. Yano and J. Takeya, *Asian J. Org. Chem.*, 2018, **7**, 2309.
- 14 (a) F. Chen, Y. S. Hong, D. Kim, T. Tanaka and A. Osuka, *ChemPlusChem*, 2017, **82**, 1048; (b) G. V. Baryshnikov, R. R. Valiev, B. F. Minaev and H. Ågren, *New J. Chem.*, 2017, **41**, 7621.
- 15 (a) F. Chen, T. Tanaka, Y. S. Hong, W. Kim, D. Kim and A. Osuka, *Chem.–Eur. J.*, 2016, **22**, 10597; (b) F. Chen, T. Tanaka and A. Osuka, *Chem. Commun.*, 2017, **53**, 2705.
- 16 J. C. Fetzer, *Large Polycyclic Aromatic Molecules – Chemistry and Analysis*, Wiley Interscience, New York, 2000.
- 17 Water-soluble nanographenes and their biological applications: (a) H. A. Lin, Y. Sato, Y. Segawa, T. Nishihara, N. Sugimoto, L. T. Scott, T. Higashiyama and K. Itami, *Angew. Chem., Int. Ed.*, 2018, **57**, 2874; (b) C. Liu, S. Zhang, J. Li, J. Wei, K. Müllen and M. A. Yin, *Angew. Chem., Int. Ed.*, 2019, **58**, 1638.
- 18 N. N. Karaush, G. V. Baryshnikov, V. A. Minaeva, H. Ågren and B. F. Minaev, *Mol. Phys.*, 2017, **115**, 2218.
- 19 From the Strickler–Berg relationship, larger  $k_r$  values can be attributed to a larger absorption coefficient at the lowest-energy band. Indeed, **12** and **12-Bu** displayed absorption bands at 429 nm ( $\epsilon = 38000 \text{ M}^{-1} \text{ cm}^{-1}$ ) and 441 nm ( $\epsilon = 39000 \text{ M}^{-1} \text{ cm}^{-1}$ ) in DMSO, which are larger than those in THF (Fig. S7–11, ESI†).
- 20 K. Yamamoto, T. Harada, Y. Okamoto, H. Chikamatsu, M. Nakazaki, Y. Kai, T. Nakao, M. Tanaka, S. Harada and N. Kasai, *J. Am. Chem. Soc.*, 1988, **110**, 3579.
- 21 I. Pischel, S. Grimme, S. Kotila, M. Nieger and F. Vögtle, *Tetrahedron: Asymmetry*, 1996, **7**, 109.
- 22 (a) J.-W. Han, X. Li and H. N. C. Wong, *Chem. Rec.*, 2015, **15**, 107; (b) P. Rashidi-Ranjbar, Y. M. Man, J. Sandstroem and H. N. C. Wong, *J. Org. Chem.*, 1989, **54**, 4888; (c) H.-Y. Peng, C.-K. Lam, T. C. W. Mak, Z. Cai, W.-T. Ma, Y.-X. Li and H. N. C. Wong, *J. Am. Chem. Soc.*, 2005, **127**, 9603.
- 23 Selected reviews of chiral nanographenes: (a) E. E. Maroto, M. Izquierdo, S. Reboredo, J. Marco-Martínez, S. Filippone and N. Martín, *Acc. Chem. Res.*, 2014, **47**, 2660; (b) I. R. Márquez, S. Castro-Fernández, A. Millán and A. G. Campaña, *Chem. Commun.*, 2018, **54**, 6705; (c) M. Rickhaus, M. Mayor and M. Jurićek, *Chem. Soc. Rev.*, 2017, **46**, 1643.

