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Rational molecular design for realizing high performance sky-blue-emitting gold(III) complexes with monoaryl auxiliary ligands and their applications for both solution-processable and vacuum-deposited organic light-emitting devices

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A new class of sky-blue-emitting arylgold(III) complexes containing tridentate bis-cyclometalating ligands derived from 2,6-diphenylpyridine (C^N^C) has been successfully designed and synthesized. By systematically varying the electron-withdrawing groups from cyano, fluoro, and trifluoromethyl to trifluoromethoxy groups on the phenyl ring of the tridentate C^N^C ligands, the emission maxima of these complexes have been significantly blue-shifted from 492 nm to 466 nm in dichloromethane solution. In addition, the higher excited state distortion with respect to the ground state associated with the multiple fluoro substitutions at the tridentate ligand has been successfully reduced by the employment of trifluoromethyl and trifluoromethoxy groups, as revealed by the Huang-Rhys factor. Taking advantage of their high photoluminescence quantum yields of up to 43% in the solid-state MCP thin-films, high performance solution-processable and vacuum-deposited organic light-emitting devices with external quantum efficiencies of up to 5.3% and 11.3%, respectively, have been realized. This work represents the first demonstration of sky-blue-emitting gold(III) complexes with an x chromaticity coordinate of <0.2 .

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Introduction

With the successful launching of high resolution mobile phones and television displays on the market, recent developments have turned organic light-emitting devices (OLEDs) into a practical reality with unprecedented success.^{1–5} The advancement in the design of high-performance luminescent materials as well as smart device architecture has led to a leap-forward development for the display industry. To realize full-color displays, emitters with emission energies covering the whole visible region from saturated blue to saturated red are required. Among the RGB colors, the development of blue-emitting materials with emission in the blue region at *ca.* 400–480 nm is still one of the most challenging tasks. Rapid progress in device performance with extraordinary high external quantum efficiencies (EQEs) has been demonstrated for blue emission in recent years.^{6,7} Of particular interest are thermally activated delayed fluorescence (TADF) and phosphorescent materials, because of their capabilities to harvest all singlet and triplet excitons. Notably, Adachi and co-workers have reported high performance blue-emitting TADF materials with nearly 100%

internal quantum efficiencies.⁸ Efficient vacuum-deposited OLEDs with EQEs of up to 25.9% have also been realized although the relatively long excited lifetimes of the TADF materials have inevitably resulted in severe triplet-triplet annihilation at high brightness.⁸

Meanwhile, blue-emitting phosphorescent complexes, mostly based on iridium(III) and platinum(II) systems, with good color purities and outstanding electroluminescence (EL) performances with EQEs of up to 30% have been demonstrated.^{9–13} Typically, the incorporation of high triplet energy cyclometalating ligands or carbene-type ligands is the most effective way to blue-shift the emission.^{14–18} In particular, a number of blue-emitting phosphors based on iridium(III) and platinum(II) metal centers have been reported; for example, deep blue-emitting phosphorescent OLEDs (PHOLEDs) based on iridium(III) complexes with picolinic acid *N*-oxide as the auxiliary ligand were reported by Kim and co-workers.¹⁹ High EQEs of 23.3% and Commission Internationale de L'Eclairage (CIE) coordinates of (0.15, 0.20) have been realized *via* smart device engineering.¹⁹ Meanwhile, Li and co-workers reported a series of deep blue-emitting phosphorescent $C^C^C^N$ tetridentate-ligand containing platinum(II) complexes.²⁰ PHOLEDs based on these complexes exhibited a maximum EQE of 23.7% with CIE coordinates of (0.14, 0.15).

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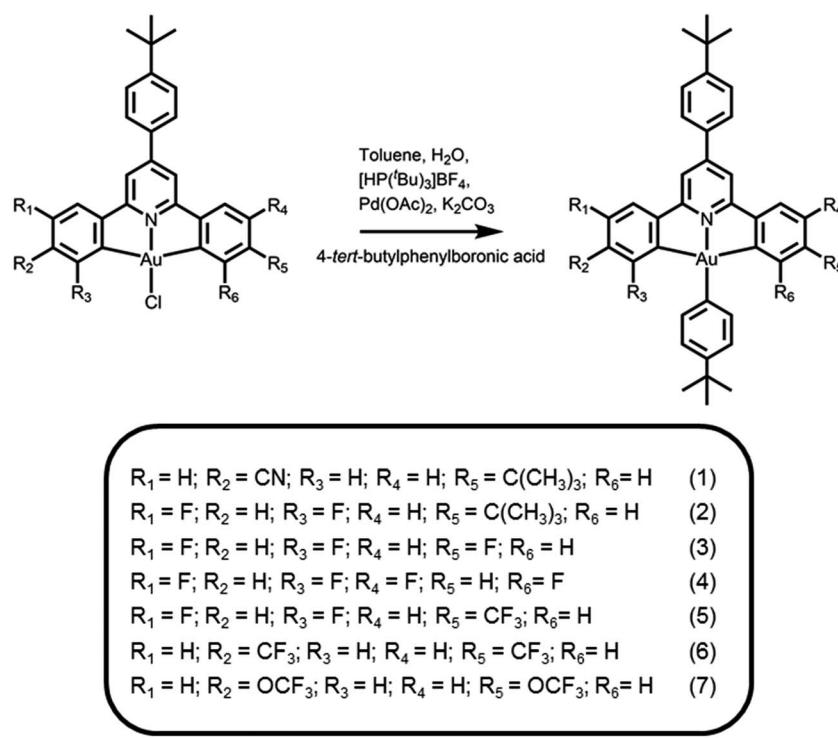
Compared to the isoelectronic platinum(II) analogues with a square-planar structure, there are only a handful of phosphorescent gold(III) complexes that have been reported to emit in the blue region.²¹ The occurrence of red-shifted excimeric emission, which is commonly observed in d⁸ square-planar metal complexes, may be one of the most important key factors for the synthetic challenges for blue-emitting gold(III) complexes.²² Besides, our recent demonstrations on the design of new classes of high performance cyclometalated gold(III) complexes with bidentate, tridentate and tetradentate ligands have provided inspiration for the strategy for further modification of this system.^{23–34} In particular, arylgold(III) complexes appear to be an ideal candidate for further development of a versatile design strategy for blue-emitting gold(III) complexes, where the steric interactions between the auxiliary aryl ring and the 2,6-diphenylpyridine (C^NC) ligand can effectively suppress the π -stacking.³⁵ To further blue-shift the emission energy, herein we report a novel class of sky-blue-emitting arylgold(III) complexes **1–7** by incorporating different tridentate ligands derived from the C^NC ligand (Scheme 1). Particularly, the 4,4'-positions of the phenyl ring of the tridentate ligand have been changed to an electron-withdrawing cyano and electron-donating *tert*-butyl groups in **1**, fluoro and *tert*-butyl groups in **2**, while complex **3** is 4-fluoro- and 3',5'-difluoro-substituted and complex **4** is tetrafluoro-substituted at 3,3',5,5'-positions. By varying their electron-withdrawing substituents as well as the positions of the phenyl ring of the tridentate ligand, the emission maxima of the complexes, *i.e.* **1–4**, are found to be progressively blue-shifted from 492 nm to 466 nm. By replacing the 4,4'-positions of the phenyl ring by

other stronger electron-withdrawing groups, *i.e.* trifluoromethyl in **6** and trifluoromethoxy in **7**, the CIE coordinates can be further blue-shifted from (0.22,0.43) in **4** to (0.22,0.42) in **5**, (0.21,0.42) in **6**, and (0.19,0.39) in **7** in the solution state. In addition to their high photoluminescence quantum yields (PLQYs) of up to 43% in doped *N,N'*-dicarbazolyl-3,5-benzene (MCP) thin films, this class of arylgold(III) complexes is highly soluble in common organic solvents and is thermally evaporable, allowing them to serve as phosphorescent dopants for both solution-processable and vacuum-deposited OLEDs. Remarkable EQEs of up to 11.3% with a peak wavelength of 476 nm and CIE coordinates of (0.19,0.43) have been achieved, demonstrating that the emission of arylgold(III) complexes can span over the entire visible region into the blue region and their broad color-tuning ability is comparable to that of state-of-the-art iridium(III) and platinum(II) counterpart systems.

Results and discussion

Synthesis of arylgold(III) complexes

The synthetic route and molecular structures of complexes **1–7** are shown in Scheme 1. Complexes **1–7** are prepared *via* modification of the literature procedure for bis-cyclometalated arylgold(III) complexes.^{36,37} The identities of all the complexes have been confirmed by ¹H and ¹³C NMR spectroscopy, FAB mass spectrometry and elemental analyses. These complexes have been isolated as thermally stable white to pale yellow solids with high decomposition temperatures (T_d) of >290 °C (Table 1). The thermogravimetric traces of selected complexes have been shown in Fig. 1.



Scheme 1 Synthetic route and molecular structures of cyclometalated arylgold(III) complexes.



Table 1 Thermal properties of complexes 1–7

Complex	$T_d^a/^\circ\text{C}$
1	344
2	310
3	296
4	325
5	304
6	293
7	315

^a T_d is defined as the temperature at which the material showed a 5% weight loss.

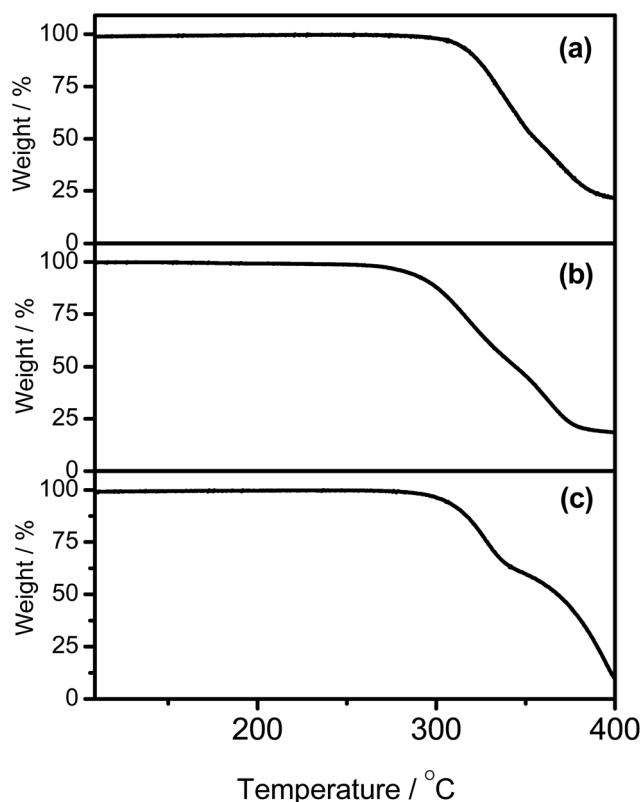


Fig. 1 Thermogravimetric analysis (TGA) curves of (a) 1, (b) 2, and (c) 5.

Photophysical properties

All the complexes show intense absorption bands at *ca.* 280–320 nm with a moderately intense vibronic-structured band at *ca.* 360–410 nm with extinction coefficients (ε) on the order of $10^4 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$. Table 2 summarizes the photophysical data of 1–7. The UV-vis absorption and emission spectra of complexes 1–7 in dichloromethane at 298 K are shown in Fig. 2 and 3, respectively. The low-energy absorption shoulder is found to be sensitive to the substituent groups on the phenyl moiety of the C^NC ligand, and has been tentatively assigned as the metal-perturbed intraligand (IL) $\pi \rightarrow \pi^*$ transition of the C^NC ligand with IL charge transfer (ILCT) character from the phenyl ring to the pyridine unit, possibly with mixing of some ligand-to-ligand charge transfer (LLCT) transition from the aryl

ligand to the pyridine unit of the C^NC ligand.^{35–37} A progressive blue shift of the low-lying absorption bands at *ca.* 384 and 404 nm in 1 to *ca.* 366 and 386 nm in 4 is observed, attributed to the stronger electron-withdrawing effect of the fluorine substituents than that of the cyano group. In addition, on going from 4 to 7, it is revealed that the incorporation of trifluoromethyl or trifluoromethoxy at the 4-position of the phenyl ring of the tridentate ligand could exert a similar electronic effect to 3,5-difluoro-substitution. It is worth noting that such modification of the substituent groups on the phenyl ring can also affect the emission energies of this class of gold(III) complexes. Upon excitation at $\lambda \geq 350 \text{ nm}$ in degassed dichloromethane solution at 298 K, a vibronic-structured band with an emission maximum at *ca.* 466–492 nm has been observed for 1–7. The emission energies are found to be dependent on the nature of the tridentate ligand and have been assigned as originating from the metal-perturbed ³IL [$\pi \rightarrow \pi^*(\text{C}^N\text{C})$] state, with some aryl-to-pyridine ³ILCT character.^{35–37} In parallel to the UV-vis studies, the emission wavelength maximum has been blue-shifted with increasing number of the fluorine atoms on the phenyl moiety of the C^NC ligand, *i.e.* 1 (492 nm) > 2 (474 nm) > 3 (470 nm) > 4 (466 nm) (Fig. 3a). Unfortunately, the emission band shape has been found to be sensitive to multiple substitutions of fluorine atoms on the C^NC ligand, which perturbs the relative transition probability of the vibronic bands, as governed by the vibrational wavefunction overlap between the excited state and the ground state, commonly known as the Franck–Condon overlap integral, which is associated with the Franck–Condon factor.^{38–40} Specifically, the Huang–Rhys factor (S), which is defined as the ratio of the vibronic (1,0) and (0,0) intensities, serves to quantify the structural distortion ΔQ of the excited states with respect to the ground state.³⁹ Tetrafluoro substitution on the phenyl ring of the C^NC ligand has resulted in a larger excited state distortion of 4, as revealed by a Huang–Rhys factor of >1 (Fig. 3a). Interestingly, upon replacing the fluorine atoms by either one or two stronger electron-withdrawing trifluoromethyl and trifluoromethoxy groups (*i.e.* 5–7) according to the Hammett constant,⁴¹ the ratio of the vibronic (1,0) and (0,0) intensities of 5 and 6 becomes closer and even comparable in the case of 7 (Fig. 3b), indicating that the Huang–Rhys factor is approaching 1. More importantly, the higher emission color purity can be reflected by their emissions in MCP thin films. For instance, all the complexes feature vibronic-structured emission bands with peak maxima at *ca.* 474–498 nm; the thin film emission spectra of 4 doped with different concentrations are illustrated in Fig. 4. The emission energies are found to be independent of the dopant concentration in MCP thin films, without the presence of any undesirable excimeric emission arising from the π – π stacking of the C^NC ligand, which is in contrast to those commonly observed in square-planar d⁸ metal complexes.^{23,24,26}

Electrochemistry

The electrochemical data of 1–7 in dichloromethane solution (0.1 M ${}^n\text{Bu}_4\text{NPF}_6$) have been measured by cyclic voltammetry



Table 2 Photophysical properties of complexes 1–7

Complex	Medium (T/K)	Absorption λ_{\max} /nm (ϵ_{\max} /dm ³ mol ⁻¹ cm ⁻¹)	Emission λ_{\max} /nm ($\tau_0/\mu\text{s}$)	Φ_{sol}^a	Φ_{film}^b
1	CH ₂ Cl ₂ (298)	300 (42 055), 365 (3815), 384 (5635), 404 (4340)	492, 530, 570 (40.4)	0.13	
	Glass (77) ^c		500, 540, 582, 630 (168)		
	Thin film (298)				
	5% in MCP				0.30
2	CH ₂ Cl ₂ (298)	298 (32 990), 358 (1980), 378 (3155), 398 (2290)	498, 536, 578 (93.7) 474, 508, 546 (12.3)	0.06	
	Glass (77) ^c		472, 510, 546, 592 (261)		
	Thin film (298)				
	5% in MCP				0.37
3	CH ₂ Cl ₂ (298)	290 (31 225), 304 (34 390), 354 (2070), 374 (3530), 394 (2300)	478, 512, 548 (86.0) 470, 504, 540 (14.0)	0.10	
	Glass (77) ^c		466, 502, 540, 584 (216)		
	Thin film (298)				
	5% in MCP				0.30
4	CH ₂ Cl ₂ (298)	290 (33 580), 308 (38 120), 330 (15 790), 366 (3735), 386 (2430)	478, 512, 548 (62.9) 466, 500, 538 (15.2)	0.07	
	Glass (77) ^c		464, 498, 534, 576 (191)		
	Thin film (298)				
	5% in MCP				0.26
	10% in MCP				0.23
	15% in MCP				0.23
	20% in MCP				0.22
5	CH ₂ Cl ₂ (298)	286 (32 210), 308 (39 285), 366 (4280), 386 (2845)	466, 500, 538 (20.1)	0.11	
	Glass (77) ^c		464, 500, 534, 576 (192)		
	Thin film (298)				
	5% in MCP				0.34
6	CH ₂ Cl ₂ (298)	284 (8745), 306 (13 040), 366 (1425), 386 (960)	476, 508, 544 (66.3) 466, 500, 538 (17.5)	0.10	
	Glass (77) ^c		466, 500, 536, 578 (160)		
	Thin film (298)				
	5% in MCP				0.30
7	CH ₂ Cl ₂ (298)	290 (28 285), 314 (21 350), 360 (1915), 372 (2450), 392 (1795)	476, 504, 544 (61.9) 466, 500, 538 (15.7)	0.11	
	Glass (77) ^c		462, 496, 534, 576 (203)		
	Thin film (298)				
	5% in MCP				0.20
	10% in MCP				0.31
	15% in MCP				0.28
	20% in MCP				0.27
	5% in MCP ^d				0.43

^a The relative luminescence quantum yield in solution was measured at room temperature using quinine sulfate in 0.5 M H₂SO₄ as the reference (excitation wavelength = 365 nm, $\Phi_{\text{lum}} = 0.546$). ^b Φ_{film} of the gold(III) compound doped into MCP thin film and excited at a wavelength of 320 nm.

^c Measured in EtOH–MeOH–CH₂Cl₂ (40 : 10 : 1, v/v). ^d Prepared by vacuum deposition.

and are depicted in Table 3, in which their highest occupied molecular orbital (HOMO) and lowest unoccupied molecular orbital (LUMO) energy levels are estimated. Selected cyclic voltammograms are shown in Fig. 5. In general, a quasi-reversible reduction couple at -1.39 V to -1.53 V vs. the saturated calomel electrode (SCE) and an irreversible oxidation wave at $+1.75$ V to $+1.89$ V vs. SCE are found for **1**–**7**. The first reduction is attributed to the ligand-centered reduction of the cyclometalating ligand.³⁷ Notably, the occurrence of the first reduction couple of **2** (-1.53 V vs. SCE) at a more negative value than those of **3** (-1.50 V vs. SCE) and **5** (-1.43 V vs. SCE) is attributed to the positive inductive effect exerted by the *tert*-butyl group on the

phenyl moiety that causes a destabilization of the π^* orbital. The first oxidation wave of all the complexes has been assigned as the mixing of ligand-centered oxidation of the cyclometalating ligand and the auxiliary aryl ligand. It is found that the oxidation potential is sensitive to the nature of the substituent groups on the phenyl moiety of the C^NC ligand. The more positive potentials for **6** ($+1.86$ V vs. SCE) and **7** ($+1.89$ V vs. SCE) than that for **1** ($+1.80$ V vs. SCE) are attributed to the more negative inductive effect of the trifluoromethyl and trifluoromethoxy groups exerted on the phenyl moiety that causes a stabilization of the filled π orbital. This assignment can be further rationalized by the computational studies in the



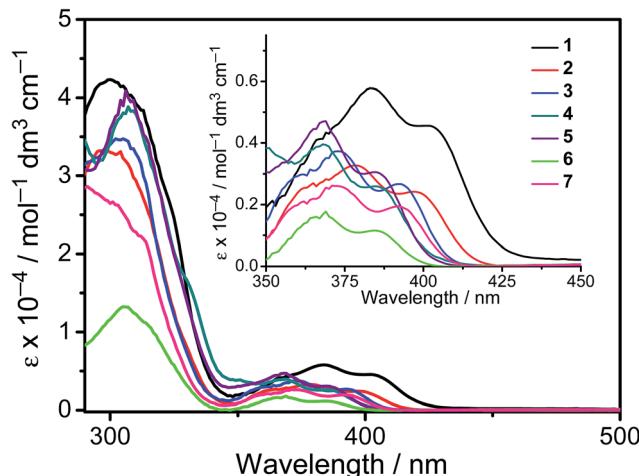


Fig. 2 UV-Vis absorption spectra of 1–7 in dichloromethane solution at 298 K.

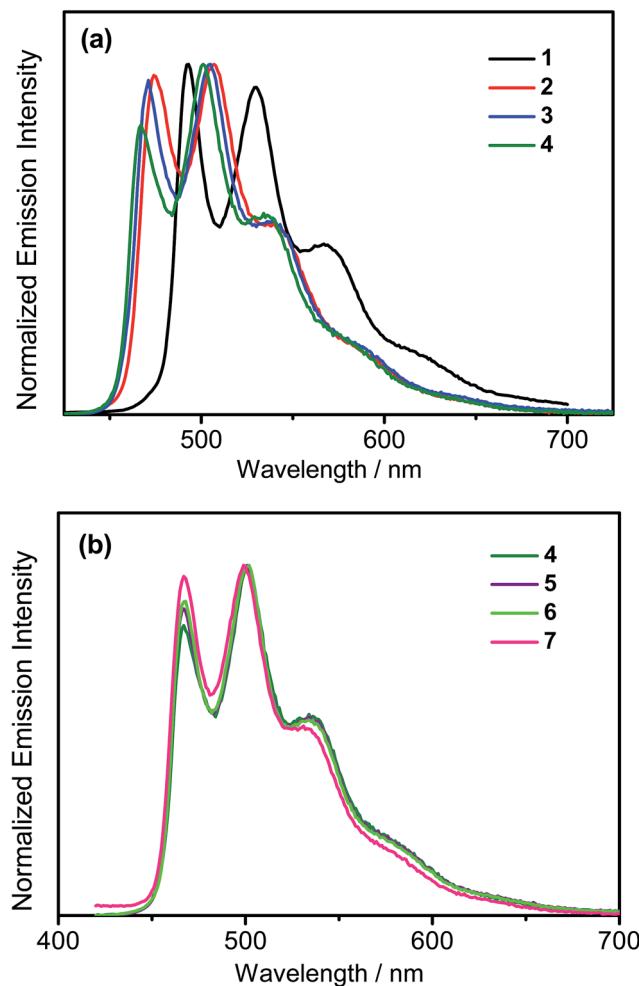


Fig. 3 (a) Normalized emission spectra of (a) 1–4 and (b) 4–7 in dichloromethane solution at 298 K.

structurally related cyclometalated arylgold(III) system reported by our group recently.³⁵ In the computational studies, the energy difference between the HOMO and HOMO–1 for this

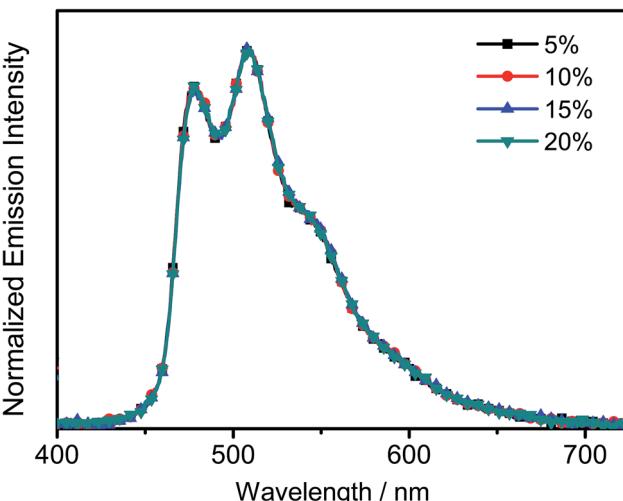


Fig. 4 Normalized PL spectra of thin films of 4 doped into MCP at different concentrations (wt%) at 298 K.

system is found to be *ca.* 0.15 eV, indicating that these two orbitals are close-lying in energy. Moreover, the HOMO–1 is predominantly from the π orbital of the C^N^C ligand mixed with a metal $d\pi$ orbital, and their HOMO is predominantly from the π orbital of the auxiliary ligand mixed with a metal $d\pi$ orbital. This further confirms that the first oxidation waves of these complexes can be fine-tuned by the choice of substituent groups.

OLED device fabrication and characterization

Compared to our previously reported alkynylgold(III) complexes, this new class of arylgold(III) complexes is both solution-processable and thermally evaporable, offering greater flexibility for the fabrication of OLEDs. To test the applicability, solution-processable OLEDs with the configuration of indium tin oxide (ITO)/poly(ethylenedioxythiophene):poly(styrene sulfonic acid) (PEDOT:PSS; 70 nm)/ x % gold(III) complex:MCP (60 nm)/tris(2,4,6-trimethyl-3-(pyridin-3-yl)phenyl)borane (3TPYMB; 5 nm)/1,3,5-tri[(3-pyridyl)phen-3-yl]benzene (TmPyPB; 30 nm)/LiF (0.8 nm)/Al (100 nm) have been prepared, in which the most bluish complexes 4 and 7 were used as phosphorescent dopants. Fig. 6 depicts the normalized EL spectra of solution-processable devices doped with 4 and 7 at different concentrations. All devices exhibit vibronic-structured emission and are independent of the dopant concentration, in agreement with the emission studies. The slight differences between the emission profile of the EL spectrum of 4 in solution-processable OLEDs and the PL spectrum of 4 measured in the doped film may be due to the different packing of the molecules that can perturb or modulate the extent of excited state distortion and hence the Huang-Rhys factor. Particularly, both x and y coordinates of all the devices remain unchanged with increasing dopant concentration from 5 wt% to 20 wt%, *i.e.* (0.20, 0.41) for 4 and (0.19, 0.40) for 7 for all of the different dopant concentrations in the range of 5 to 20 wt%. A similar phenomenon is also observed for the vacuum-deposited OLEDs with the structure of ITO/molybdenum

Table 3 Electrochemical data for 1–7^a

Complex	Oxidation [E_{pa} /V vs. SCE] ^b	Reduction $E_{1/2}$ /V vs. SCE ^c (ΔE_p /mV) ^d [E_{pc} /V vs. SCE] ^e	E_{HOMO}^f /eV	E_{LUMO}^f /eV
1	[+1.80]	-1.44 (60), [-1.93]	-6.60	-3.36
2	[+1.75]	-1.53 (60)	-6.55	-3.27
3	[+1.81], [+1.89]	-1.50 (63)	-6.61	-3.30
4	[+1.78], [+1.92]	-1.44 (53), [-1.97]	-6.58	-3.36
5	[+1.81]	-1.43 (62), [-1.92]	-6.61	-3.37
6	[+1.86], [+1.94]	-1.39 (71), [-1.91]	-6.66	-3.41
7	[+1.89], [+1.94]	-1.49 (65), [-2.04]	-6.69	-3.31

^a In CH_2Cl_2 solution with 0.1 M ${}^n\text{Bu}_4\text{NPF}_6$ as a supporting electrolyte at 298 K; working electrode, glassy carbon; scan rate = 100 mV s⁻¹. ^b E_{pa} refers to the anodic peak potential for the irreversible oxidation wave. ^c $E_{1/2} = (E_{pa} + E_{pc})/2$; E_{pa} and E_{pc} are the peak anodic and peak cathodic potentials, respectively. ^d $\Delta E_p = (E_{pa} - E_{pc})$. ^e E_{pc} refers to the cathodic peak potential for the irreversible reduction wave. ^f E_{HOMO} and E_{LUMO} levels were calculated from electrochemical potentials, i.e. $E_{HOMO} = -e(4.8 \text{ V} + E_{pa})$; $E_{LUMO} = -e(4.8 \text{ V} + E_{pc})$.

oxide (MoO_x ; 2 nm)/di-[4-(*N,N*-di-*p*-tolyl-amino)-phenyl] cyclohexane (TAPC; 40 nm)/x% 4 or 7:MCP (25 nm)/diphenyl-4-triphenylsilylphenyl-phosphine oxide (TSPO1; 5 nm)/1,3,5-tris(6-(3-pyridin-3-yl)phenyl)pyridin-2-yl)benzene (Tm3PyP26-PyB; 35 nm)/LiF (1 nm)/Al (100 nm), in which no undesired

emission from the adjacent carrier-transporting and/or host materials is observed (Fig. 7). This concentration independence is seldom observed for square-planar metal complexes, in which the π -stacking of the molecules usually leads to the formation of excimeric emission, generating a significant spectral shift and hence poor color purity of the OLEDs. It should be highlighted that the present results open up another effective means to realize sky-blue emission of gold(III) complexes. Unlike the alkynylgold(III) analogues, the steric repulsion between the aryl ring and the $\text{C}^{\text{N}}\text{C}$ ligand can effectively suppress the intermolecular interactions. By

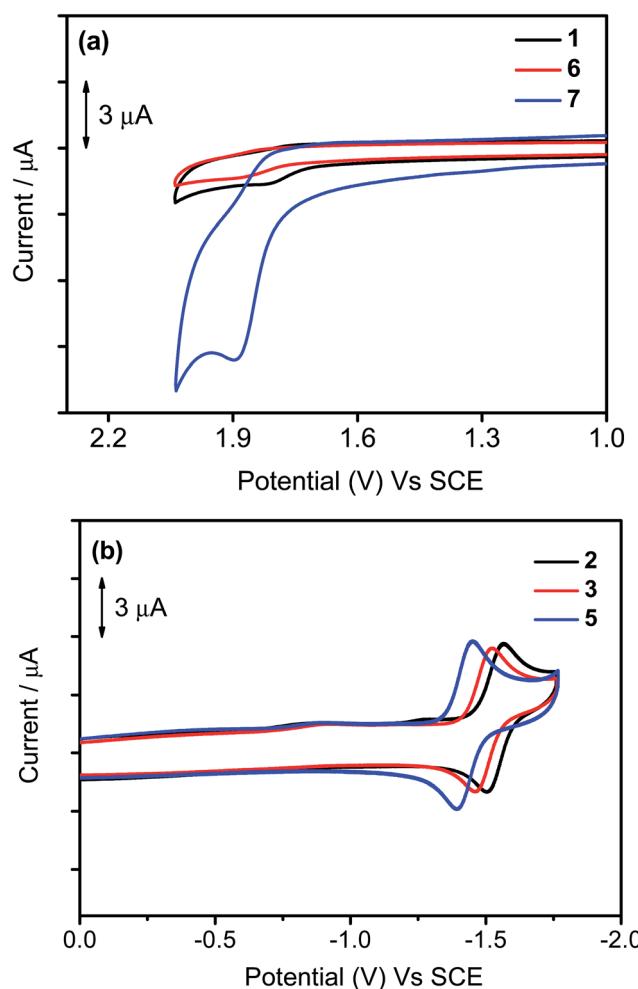


Fig. 5 Cyclic voltammograms for the (a) oxidation of 1, 6, and 7 and (b) reduction scans of 2, 3, and 5 in dichloromethane (0.1 M ${}^n\text{Bu}_4\text{NPF}_6$).

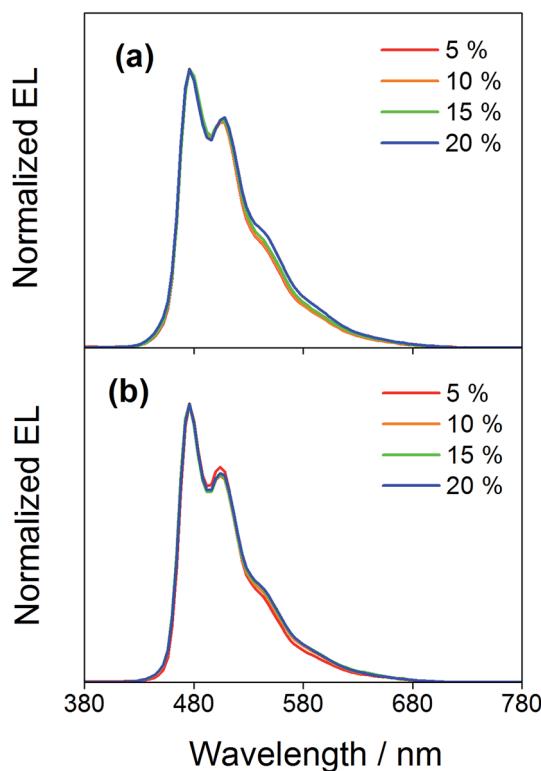


Fig. 6 Normalized EL spectra of solution-processable devices doped with 4 and 7 at different concentrations.

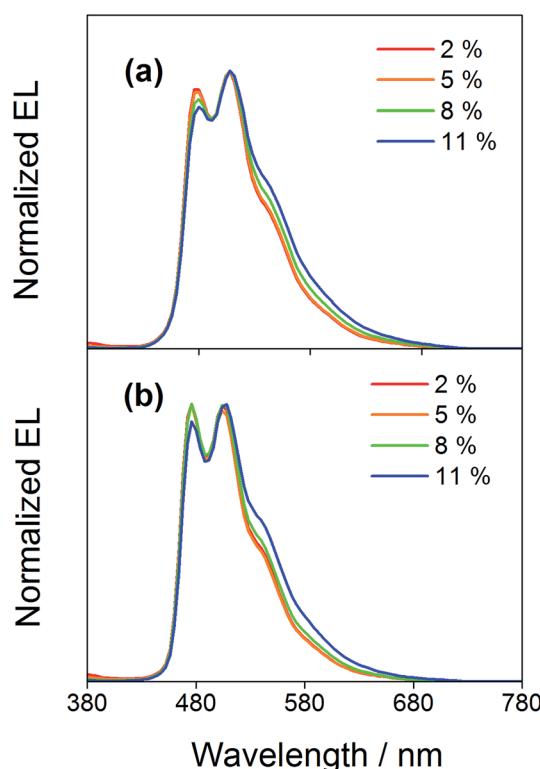


Fig. 7 Normalized EL spectra of vacuum-deposited OLEDs based on **4** and **7**.

modifying the substituents on the phenyl ring of the C^N^C ligand, ready fine-tuning of the energy of the arylgold(III) complexes into the blue region has been realized.

In addition to high color purity, satisfactory device performance has been obtained for both solution-processable and vacuum-deposited OLEDs. For solution-processable devices, the optimized devices doped with **4** and **7** show maximum EQEs of 2.3% and 5.3%, respectively. While both gold(III) complexes have similar PLQYs, the discrepancy in the EQEs may be due to the better solubility of **7**, yielding more uniform thin films. This solubility problem can be diminished by using thermal evaporation, in which both devices exhibit comparable EQEs of up to 10.7% and 11.3%, respectively (Tables 4 and 5). Fig. 8 and 9 depict the EQEs of solution-processable and vacuum-deposited devices made with **4** and **7**, respectively. The operational stability of the vacuum-deposited device based on 5 wt% complex **7** has been investigated and tested at a constant driving current density of 20 mA cm^{-2} . An operational half-life of ~ 48 hours at a luminance of 100 cd m^{-2} is estimated according to the accelerated lifetime testing method (Fig. 10). These findings clearly demonstrate that this new class of sky-blue-emitting arylgold(III) complexes shows promise as phosphorescent dopants for both high performance solution-processable and vacuum-deposited OLEDs. More importantly, this work represents the first successful demonstration of blue-emitting gold(III) complexes with CIE x coordinate < 0.2 .

Table 4 Key parameters of solution-processable devices based on **4** and **7**

Complex	Dopant concentration/wt%	Max. current efficiency/cd A $^{-1}$	Max. power efficiency/lm W $^{-1}$	Max. EQE/%	$\lambda_{\text{max}}^a/\text{nm}$
4	5	1.8	0.5	0.7	476 (0.20,0.41)
	10	3.8	1.2	1.6	476 (0.20,0.41)
	15	5.6	2.4	2.3	476 (0.20,0.41)
	20	3.7	1.4	1.5	476 (0.21,0.41)
7	5	5.8	2.1	2.5	476 (0.19,0.40)
	10	10.5	6.5	4.4	476 (0.19,0.40)
	15	11.4	7.2	4.8	476 (0.19,0.40)
	20	12.6	7.9	5.3	476 (0.19,0.40)

^a CIE coordinates in parentheses, measured at 100 cd m^{-2} .

Table 5 Key parameters of vacuum-deposited devices based on **4** and **7**

Complex	Dopant concentration/wt%	Max. current efficiency/cd A $^{-1}$	Max. power efficiency/lm W $^{-1}$	Max. EQE/%	$\lambda_{\text{max}}^a/\text{nm}$
4	2	12.5	9.8	4.8	508 (0.20,0.45)
	5	14.3	12.8	5.5	508 (0.20,0.46)
	8	30.1	31.6	10.7	508 (0.22,0.47)
	11	24.5	25.7	8.5	508 (0.23,0.48)
7	2	13.8	12.4	6.0	476 (0.19,0.43)
	5	21.1	18.9	11.3	476 (0.19,0.43)
	8	29.9	26.7	9.9	476 (0.20,0.44)
	11	35.2	31.6	11.2	508 (0.22,0.47)

^a CIE coordinates in parentheses, measured at 100 cd m^{-2} .

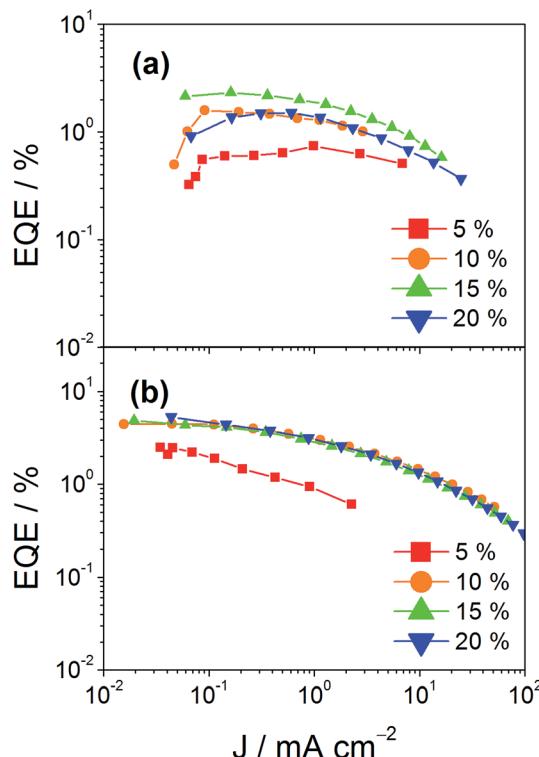


Fig. 8 EQEs of solution-processable devices made with (a) 4 and (b) 7 at different concentrations.

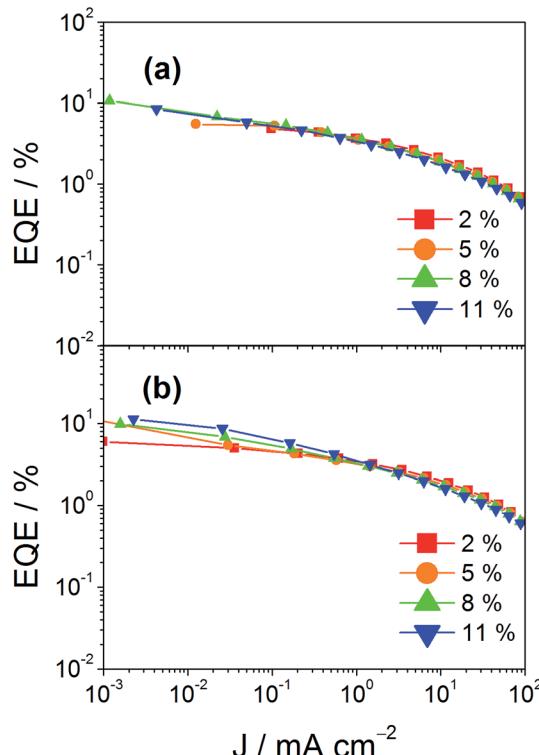


Fig. 9 EQEs of vacuum-deposited devices made with (a) 4 and (b) 7 at different concentrations.

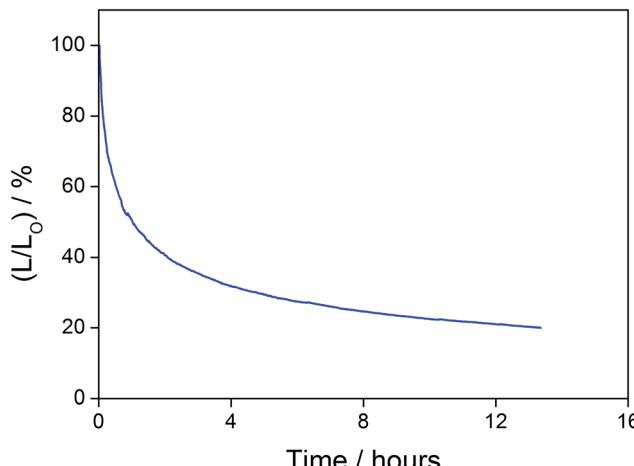


Fig. 10 Relative luminance of vacuum-deposited OLEDs based on 5 wt% complex 7 as a function of operation time at a constant driving current density of 20 mA cm^{-2} .

Conclusion

A new class of sky-blue-emitting arylgold(III) complexes of 2,6-diphenylpyridine has been designed and synthesized by engineering the substituents on the phenyl ring, in which the emission energies have been effectively fine-tuned from 492 nm to 466 nm in dichloromethane solution. It is found that the incorporation of more fluorine atoms would cause a higher excited state distortion with respect to the ground state, while trifluoromethylation is an effective way to shift the emission to the blue region. More importantly, this class of arylgold(III) complexes is capable of serving as phosphorescent dopants with high EQEs of up to 5.3% and 11.3% for both solution-processable and vacuum-deposited OLEDs, respectively, opening up a versatile approach to realize blue emission in square-planar gold(III) complexes.

Experimental section

Materials

The tridentate ligands, $4\text{-CN-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C}^{\wedge}\text{Bu-4}$, $3,5\text{-F}_2\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C}^{\wedge}\text{Bu-4}$, $3,5\text{-F}_2\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C-F-4}$, $3,5\text{-F}_2\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C-F-3,5}$, $3,5\text{-F}_2\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C-CF}_3\text{-4}$, $4\text{-CF}_3\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C-CF}_3\text{-4}$, and $4\text{-OCF}_3\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C-CF}_3\text{-4}$, and the precursor gold(III) complexes, $[\text{Au}\{4\text{-CN-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C}^{\wedge}\text{Bu-4}\}\text{Cl}]$, $[\text{Au}\{3,5\text{-F}_2\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C}^{\wedge}\text{Bu-4}\}\text{Cl}]$, $[\text{Au}\{3,5\text{-F}_2\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C-F-4}\}\text{Cl}]$, $[\text{Au}\{3,5\text{-F}_2\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C-F-3,5}\}\text{Cl}]$, $[\text{Au}\{3,5\text{-F}_2\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C-CF}_3\text{-4}\}\text{Cl}]$, $[\text{Au}\{4\text{-CF}_3\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C-CF}_3\text{-4}\}\text{Cl}]$ and $[\text{Au}\{4\text{-OCF}_3\text{-C}^{\wedge}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C-CF}_3\text{-4}\}\text{Cl}]$, were prepared *via* modification of the procedures reported in the literature.^{42,43} Specifically, the $\text{C}^{\wedge}\text{N}^{\wedge}\text{C}$ pincer ligands were refluxed with mercuric(II) acetate in ethanol. Saturated lithium chloride in hot methanol was added to the reaction mixture to yield the mercury precursor complexes as a white precipitate. The mercury precursors were then refluxed with $\text{K}[\text{AuCl}_4]$ in acetonitrile. After removal of solvent, the crude product was purified by

column chromatography on silica gel to yield the chlorogold(III) precursor complexes as a pale yellow solid. 4-Acetylbenzonitrile, 3,5-difluoroacetophenone, 4-fluoroacetophenone, 4-*tert*-butylacetophenone, 4-(trifluoromethyl)acetophenone, 4-(trifluoromethoxy)acetophenone and 4-*tert*-butylphenylboronic acid were purchased from Sigma Aldrich or Matrix Scientific. All solvents were purified and distilled using standard procedures before use. All other reagents were of analytical grade and were used as received. Triethylamine was distilled over calcium hydride before use. Tetra-*n*-butylammonium hexafluorophosphate (Aldrich, 98%) was recrystallized for no less than three times from hot absolute ethanol prior to use. All reactions were performed under anaerobic and anhydrous conditions using standard Schlenk techniques under an inert atmosphere of nitrogen.

Physical measurements and instrumentation

The UV-vis absorption spectra were recorded on a Cary 60 UV/Vis (Agilent Technology) spectrophotometer equipped with a xenon flash lamp. ^1H and $^{13}\text{C}\{^1\text{H}\}$ NMR spectra were recorded on a Bruker Avance 600 (600 MHz for ^1H and 150 MHz for ^{13}C nuclei) Fourier-transform NMR spectrometer with chemical shifts reported relative to tetramethylsilane ($\delta = 0$ ppm) in deuterated chloroform. $^{19}\text{F}\{^1\text{H}\}$ NMR spectra were recorded on a Bruker Avance 400 (376 MHz for ^{19}F nucleus) Fourier-transform NMR spectrometer with chemical shifts reported relative to trifluoroacetic acid ($\delta = -76.55$ ppm) in deuterated chloroform. Positive FAB mass spectra were recorded on a Thermo Scientific DFS High Resolution Magnetic Sector Mass Spectrometer. IR spectra were recorded as KBr disks on a Bio-Rad FTS-7 FTIR spectrometer ($4000\text{--}400\text{ cm}^{-1}$). Elemental analyses were performed on a Carlo Erba 1106 elemental analyzer at the Institute of Chemistry, Chinese Academy of Sciences in Beijing. Steady-state excitation and emission spectra were recorded on a Horiba Scientific FluoroMax-4 fluorescence spectrophotometer equipped with an R928P PMT detector. Liquid nitrogen was placed into a quartz-walled optical Dewar flask for low temperature (77 K) photophysical measurements. Solid-state photophysical measurements were performed with solid samples loaded into a quartz tube inside the quartz-walled Dewar flask. Excited-state lifetimes of solution and glass samples were measured using a conventional laser system. The excitation source used was the 355 nm output (third harmonic, 8 ns) of a Spectra-Physics Quanta-Ray Q-switched GCR-150 pulsed Nd:YAG laser (10 Hz). Luminescence decay signals were detected with a Hamamatsu R928 photomultiplier tube, recorded on a Tektronix model TDS-620A (500 MHz, 2 GS s^{-1}) digital oscilloscope and analyzed with a program for exponential fits. PLQYs were measured by the optical dilute method reported by Demas and Crosby.⁴⁴ An aqueous solution of quinine sulfate in 0.5 M H_2SO_4 has been used as the reference ($\Phi_{\text{lum}} = 0.546$, excitation wavelength of 365 nm),⁴⁴ whereas thin film PLQYs were measured on a Hamamatsu C9920-03 Absolute PLQY Measurement System. Excited-state lifetimes of thin films were measured on a Hamamatsu C11367-34 Quantaurus-Tau Fluorescence lifetime spectrometer. Cyclic voltammetry was performed with a CH Instruments Model CHI620E (CH

Instruments, Inc.). All solutions for electrochemical measurements were purged with pre-purified argon gas prior to measurement. Thermal analyses were performed with a Q50 TGA (TA instruments), in which the decomposition temperature, T_d , is defined as the temperature at which the sample shows a 5% weight loss.

OLED fabrication and characterization

Solution-processable OLEDs were fabricated on patterned ITO glass substrates with a sheet resistance of $30\ \Omega$ per square. The substrates were cleaned with Decon 90, rinsed with deionized water, dried in an oven, and finally treated in an ultraviolet-ozone chamber. A 70 nm thick PEDOT:PSS layer was spin-coated onto the ITO coated glass substrates as a hole-transporting layer. After that, the emissive layer was formed by mixing the gold(III) complex with MCP to prepare a $10\ \text{mg cm}^{-3}$ solution in chloroform and spin-coating it onto the PEDOT:PSS layer to give uniform thin films of 60 nm thickness. Onto this, 5 nm thick 3TPYMB and a 30 nm thick TmPyPB were evaporated as a hole-blocking layer and an electron-transporting layer, respectively, while LiF/Al was used as the metal cathode. For vacuum-deposited OLEDs, sequential thermal evaporation of MoO_x , TAPC, the emissive layer, TSPO1, Tm3P26PyB, LiF and Al was done onto the ITO substrates. All organic and metal films were sequentially deposited at a rate of $0.1\text{--}0.2\ \text{nm s}^{-1}$ without any vacuum break. A shadow mask was used to define the cathode and to make four $0.1\ \text{cm}^2$ devices on each substrate. Current density–voltage–luminance characteristics and EL spectra were measured simultaneously with a programmable Keithley model 2420 power source and a Photoresearch PR-655 spectrometer. The operational stability of OLEDs was tested under accelerated lifetime testing conditions and the half-life was estimated using the equation $t_{1/2} = \text{constant}/(L_0)^m$, where L_0 is the initial luminescence, $t_{1/2}$ is the measured half-life, and m is an acceleration factor taken to be 1.7.

Synthesis and characterization of gold(III) complexes

Complexes 1–7 were synthesized by reacting the respective chlorogold(III) precursors with 4-*tert*-butylphenylboronic acid in the presence of a catalytic amount of palladium(II) catalyst in an aqueous base and organic solvent. For instance, to a two-necked flask containing $[\text{Au}\{4\text{-CN-C}^{\text{H}}\text{N}(4\text{-}^t\text{BuC}_6\text{H}_4)\text{C}^{\text{H}}\text{Bu-4}\}\text{Cl}]$ (0.14 g, 0.23 mmol), K_2CO_3 (0.55 g, 0.43 mmol), $\text{Pd}(\text{OAc})_2$ (7.20 mg, 0.023 mmol), $[\text{HP}(^t\text{Bu})_3]\text{BF}_4$ (13.30 mg, 0.046 mmol) and 4-*tert*-butylphenylboronic acid (0.82 g, 0.46 mmol) were added into degassed toluene and H_2O (4 : 1, v/v) and the mixture was stirred at $80\text{ }^{\circ}\text{C}$ for 12 hours under a nitrogen atmosphere. After removing the solvent, the crude product was purified by column chromatography on silica gel (70–230 mesh) using hexane–dichloromethane (6 : 1, v/v) as the eluent. Subsequent recrystallization by diffusion of diethyl ether vapor into a concentrated dichloromethane solution gave the product as a pale yellow solid.

Complex 1. Yield: 53.6 mg, 34%. ^1H NMR (500 MHz, THF-d_8 , 298 K, δ/ppm): δ 8.09 (d, $J = 1.5\text{ Hz}$, 1H, -py of $\text{C}^{\text{H}}\text{N}^{\text{C}}$), 8.05–8.03 (m, 2H, -py and $-\text{C}_6\text{H}_3\text{'}\text{Bu-}$ of $\text{C}^{\text{H}}\text{N}^{\text{C}}$), 7.87 (d, $J = 8.5\text{ Hz}$, 2H, $-\text{C}_6\text{H}_4-$ of $\text{C}^{\text{H}}\text{N}^{\text{C}}$), 7.79 (d, $J = 8.0\text{ Hz}$, 1H, $-\text{C}_6\text{H}_3\text{CN-}$ of $\text{C}^{\text{H}}\text{N}^{\text{C}}$),



7.65 (d, 1H, $J = 2.0$ Hz, $-C_6H_3CN$ of C^N^C), 7.62 (d, $J = 8.5$ Hz, 2H, $-C_6H_4$ of C^N^C), 7.56 (dd, $J = 8.0$ and 2.0 Hz, 1H, $-C_6H_3CN$ of C^N^C), 7.52 (d, $J = 8.5$ Hz, 2H, $-C_6H_4^tBu-p$), 7.41 (d, $J = 2.0$ Hz, 1H, $-C_6H_3^tBu$ of C^N^C), 7.36 (d, $J = 8.5$ Hz, 2H, $-C_6H_4^tBu-p$), 7.27 (dd, $J = 8.5$ and 2.0 Hz, 1H, $-C_6H_3^tBu$ of C^N^C), 1.40 (s, 9H, $-tBu$, $-C_6H_4$ of C^N^C), 1.39 (s, 9H, $-tBu$, $-C_6H_4^tBu-p$), 1.23 (s, 9H, $-tBu$, $-C_6H_3^tBu$ of C^N^C). $^{13}C\{^1H\}$ NMR (150 MHz, $CDCl_3$, δ /ppm): δ 169.78, 166.25, 163.57, 161.07, 154.88, 154.77, 154.05, 153.82, 147.66, 146.61, 143.14, 138.82, 134.64, 133.33, 132.65, 130.02, 127.07, 126.40, 125.96, 124.98, 124.36, 123.60, 119.73, 115.62, 115.24, 113.75, 35.35, 34.92, 34.40, 31.52, 31.25, 31.06. Positive ESI-MS: m/z 773.3 [$M + H$]⁺. Elemental analyses: found (%): C 64.98, H 5.51, N 3.60. Calcd for $AuC_{42}H_{43}N_2$: C 65.28, H 5.61, N 3.63.

Complex 2. Yield: 50 mg, 44%. 1H NMR (500 MHz, $CDCl_3$, 298 K, relative to Me_4Si , δ /ppm): δ 7.67–7.65 (m, 3H, $-py$ and $-C_6H_4$ of C^N^C), 7.62 (s, 1H, $-py$ of C^N^C), 7.59–7.56 (m, 3H, $-C_6H_4$ and $-C_6H_3^tBu$ of C^N^C), 7.55–7.53 (d, $J = 8.0$ Hz, 2H, $-C_6H_4^tBu-p$), 7.29–7.23 (m, 4H, $-C_6H_3^tBu$ of C^N^C and $-C_6H_4^tBu-p$), 7.07–7.06 (m, 1H, $-F_2C_6H_2$ of C^N^C), 6.73–6.70 (m, 1H, $-F_2C_6H_2$ of C^N^C), 1.40 (s, 9H, $-tBu$, $-C_6H_4$ of C^N^C), 1.38 (s, 9H, $-tBu$, $-C_6H_4^tBu-p$), 1.17 (s, 9H, $-tBu$, $-C_6H_3^tBu$ of C^N^C). $^{13}C\{^1H\}$ NMR (150 MHz, $CDCl_3$, δ /ppm): δ 168.07, 167.99, 166.47, 166.39, 163.98, 163.96, 163.17, 163.08, 163.00, 161.63, 161.61, 161.46, 161.38, 154.83, 154.62, 153.95, 152.15, 152.09, 152.02, 151.97, 148.30, 148.28, 148.00, 147.98, 146.82, 146.10, 140.34, 140.32, 134.71, 133.11, 132.09, 127.06, 126.37, 125.10, 124.69, 123.59, 115.14, 114.85, 108.70, 108.68, 108.55, 108.53, 106.37, 106.21, 106.14, 35.21, 34.91, 34.25, 31.52, 31.25, 30.90. $^{19}F\{^1H\}$ NMR (470 MHz, $CDCl_3$, relative to CF_3COOH , δ /ppm): δ –85.70, –113.23. Positive ESI-MS: m/z 784.3 [$M + H$]⁺. Elemental analyses: found (%): C 62.54, H 5.37, N 1.74. Calcd for $AuC_{41}H_{42}NF_2$: C 62.83, H 5.40, N 1.79.

Complex 3. Yield: 200 mg, 58%. 1H NMR (500 MHz, CD_2Cl_2 , 298 K, δ /ppm): δ 7.76–7.68 (m, 5H, $-F_2C_6H_2$ and $-py$ of C^N^C , and $-C_6H_4^tBu-p$), 7.61 (d, $J = 8.5$ Hz, 2H, $-C_6H_4^tBu-p$), 7.45 (d, $J = 8.0$ Hz, 2H, $-C_6H_4$ of C^N^C), 7.30 (d, $J = 9.0$ Hz, 1H, $-FC_6H_3$ of C^N^C), 7.26 (d, $J = 8.0$ Hz, 2H, $-C_6H_4$ of C^N^C), 6.92–6.89 (m, 1H, $-F_2C_6H_2$ of C^N^C), 6.79–6.77 (m, 1H, $-FC_6H_3$ of C^N^C), 6.73–6.68 (m, 1H, $-FC_6H_3$ of C^N^C), 1.40 (s, 9H, $-tBu$, $-C_6H_4$ of C^N^C), 1.38 (s, 9H, $-tBu$, $-C_6H_4^tBu-p$). $^{13}C\{^1H\}$ NMR (150 MHz, $THF-d_8$, δ /ppm): δ 168.87, 168.77, 167.32, 167.25, 166.28, 164.59, 164.38, 164.31, 163.12, 162.77, 162.72, 162.70, 155.79, 154.81, 153.66, 153.60, 153.54, 153.48, 147.57, 147.43, 147.28, 146.49, 146.48, 140.50, 140.47, 135.32, 133.88, 128.44, 128.39, 128.24, 126.96, 125.98, 122.17, 122.04, 116.39, 116.25, 114.16, 144.00, 110.01, 109.99, 109.86, 109.84, 106.69, 106.53, 106.47, 106.31, 35.57, 34.90, 32.00, 31.58. $^{19}F\{^1H\}$ NMR (470 MHz, CD_2Cl_2 , 298 K, relative to CF_3COOH , δ /ppm): δ –86.87, –107.74, –113.42. Positive ESI-MS: m/z 746.2 [$M + H$]⁺. Elemental analyses: found (%): C 59.24, H 4.82, N 1.75. Calcd for $AuC_{37}H_{33}NF_3$: C 59.60, H 4.46, N 1.88.

Complex 4. Yield: 80 mg, 42%. 1H NMR (500 MHz, acetone- d_6 , 298 K, δ /ppm): δ 8.38 (s, 2H, $-py$ of C^N^C), 8.03 (d, $J = 8.5$ Hz, 2H, $-C_6H_4$ of C^N^C), 7.81–7.79 (m, 2H, $-F_2C_6H_2$ of C^N^C), 7.67 (d, $J = 8.5$ Hz, 2H, $-C_6H_4$ of C^N^C), 7.44 (d, $J = 8.0$ Hz, 2H, $-C_6H_4^tBu-p$), 7.16 (d, $J = 8.0$ Hz, 2H, $-C_6H_4^tBu-p$),

6.82–6.65 (m, 2H, $-F_2C_6H_2$ of C^N^C), 1.41 (s, 9H, $-tBu$, $-C_6H_4$ of C^N^C), 1.35 (s, 9H, $-tBu$, $-C_6H_4^tBu-p$). $^{13}C\{^1H\}$ NMR (150 MHz, $THF-d_8$, δ /ppm): δ 168.12, 168.02, 166.15, 166.05, 163.60, 163.51, 161.79, 161.57, 155.03, 154.12, 152.33, 152.25, 152.20, 152.12, 145.82, 143.25, 134.16, 132.86, 127.41, 126.10, 124.33, 116.36, 109.13, 108.97, 106.33, 106.14, 106.07, 105.88, 34.69, 33.87, 31.17, 30.65. $^{19}F\{^1H\}$ NMR (470 MHz, $CDCl_3$, 298 K, relative to CF_3COOH , δ /ppm): δ –87.42, –112.24. Positive ESI-MS: m/z 764.2 [$M + H$]⁺. Elemental analyses: found (%): C 58.17, H 4.42, N 1.83. Calcd for $AuC_{37}H_{32}NF_4$: C 58.20, H 4.22, N 1.83.

Complex 5. Yield: 400 mg, 47%. 1H NMR (500 MHz, $CDCl_3$, 298 K, relative to Me_4Si , δ /ppm): δ 7.76 (d, $J = 1.0$ Hz, 1H, $-py$ of C^N^C), 7.73–7.70 (m, 2H, $-py$ and $-C_6H_3CF_3$ of C^N^C), 7.64 (d, $J = 6.5$ Hz, 2H, $-C_6H_4^tBu-p$), 7.59 (d, $J = 6.5$ Hz, 2H, $-C_6H_4^tBu-p$), 7.46–7.44 (m, 3H, $-C_6H_3CF_3$ and $-C_6H_4$ of C^N^C), 7.35 (s, 1H, $-C_6H_3CF_3$ of C^N^C), 7.27 (d, 2H, $J = 8.0$ Hz, $-C_6H_4$ of C^N^C), 7.23–7.21 (m, 1H, $-F_2C_6H_2$ of C^N^C), 6.71–6.68 (m, 1H, $-F_2C_6H_2$ of C^N^C), 1.41 (s, 9H, $-tBu$, $-C_6H_4^tBu-p$). $^{13}C\{^1H\}$ NMR (125 MHz, $CDCl_3$, δ /ppm): δ 168.20, 168.11, 166.26, 166.17, 165.14, 165.11, 163.36, 163.27, 161.93, 161.90, 161.58, 161.41, 161.31, 155.18, 154.43, 152.19, 151.71, 151.64, 151.56, 151.50, 147.37, 146.71, 146.36, 138.62, 138.59, 134.19, 132.76, 131.34, 131.32, 127.09, 126.51, 125.49, 124.63, 123.61, 123.58, 116.12, 116.06, 108.99, 108.97, 108.82, 108.79, 106.90, 106.72, 106.64, 106.46, 34.97, 34.29, 31.51, 31.24. $^{19}F\{^1H\}$ NMR (470 MHz, $CDCl_3$, relative to CF_3COOH , δ /ppm): δ –63.02, –85.75, –112.50. Positive ESI-MS: m/z 796.2 [$M + H$]⁺. Elemental analyses: found (%): C 57.57, H 4.37, N 1.59. Calcd for $AuC_{38}H_{33}NF_5$: C 57.36, H 4.18, N 1.76.

Complex 6. Yield: 110 mg, 44%. 1H NMR (500 MHz, $CDCl_3$, 298 K, relative to Me_4Si , δ /ppm): δ 7.76 (s, 2H, $-py$ of C^N^C), 7.72 (d, $J = 8.0$ Hz, 2H, $-C_6H_3CF_3$ of C^N^C), 7.67–7.66 (m, 2H, $-C_6H_3CF_3$ of C^N^C), 7.62–7.57 (m, 4H, $-C_6H_4$ of C^N^C), 7.47–7.45 (m, 4H, $-C_6H_3CF_3$ of C^N^C and $-C_6H_4^tBu-p$), 7.35 (d, $J = 8.5$ Hz, 2H, $-C_6H_4^tBu-p$), 1.42 (s, 18H, $-tBu$, $-C_6H_4$ of C^N^C and $-C_6H_4^tBu-p$). $^{13}C\{^1H\}$ NMR (150 MHz, $CDCl_3$, δ /ppm): δ 168.19, 161.84, 155.19, 154.37, 152.76, 147.96, 142.36, 134.22, 133.04, 132.18, 131.97, 131.94, 127.10, 126.46, 126.25, 125.53, 125.22, 124.67, 123.46, 123.44, 123.41, 116.14, 34.95, 34.41, 31.49, 31.23. $^{19}F\{^1H\}$ NMR (470 MHz, $CDCl_3$, 298 K, relative to CF_3COOH , δ /ppm): δ –62.81. Positive ESI-MS: m/z 828.2 [$M + H$]⁺. Elemental analyses: found (%): C 56.63, H 4.18, N 1.52. Calcd for $AuC_{39}H_{34}NF_6$: C 56.60, H 4.14, N 1.69.

Complex 7. Yield: 320 mg, 48%. 1H NMR (500 MHz, $CDCl_3$, 298 K, relative to Me_4Si , δ /ppm): δ 7.58 (d, $J = 8.5$ Hz, 2H, $-C_6H_3OCF_3$ of C^N^C), 7.55–7.50 (m, 6H, $-py$ and $-C_6H_4$ of C^N^C), 7.38 (d, $J = 8.5$ Hz, 2H, $-C_6H_4^tBu-p$), 7.31 (d, $J = 8.5$ Hz, 2H, $-C_6H_4^tBu-p$), 7.18–7.17 (m, 2H, $-C_6H_3OCF_3$ of C^N^C), 6.98–6.96 (m, 2H, $-C_6H_3OCF_3$ of C^N^C), 1.40 (s, 9H, $-tBu$, $-C_6H_4$ of C^N^C), 1.37 (s, 9H, $-tBu$, $-C_6H_4^tBu-p$). $^{13}C\{^1H\}$ NMR (125 MHz, $CDCl_3$, δ /ppm): δ 170.19, 161.86, 154.87, 154.24, 151.32, 147.91, 147.83, 142.67, 134.49, 133.23, 127.29, 127.15, 126.46, 126.24, 121.60, 119.55, 118.00, 115.10, 35.03, 34.50, 31.63, 31.35. $^{19}F\{^1H\}$ NMR (470 MHz, $CDCl_3$, relative to CF_3COOH , δ /ppm): δ –57.16. Positive FAB-MS: m/z 860.2 [$M + H$]⁺. Elemental analyses: found (%): C 54.47, H 4.15, N 1.47. Calcd for $AuC_{39}H_{34}NO_2F_6$: C 54.49, H 3.99, N 1.63.



Conflicts of interest

There are no conflicts to declare.

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