



Cite this: *Dalton Trans.*, 2018, **47**, 11572

Received 13th June 2018,
Accepted 26th July 2018

DOI: 10.1039/c8dt02432d
rsc.li/dalton

Design of oxophilic metalloporphyrins: an experimental and DFT study of methanol binding†

Sandra Olsson,  Christian Dahlstrand  and Adolf Gogoll *

By systematic measurements we have evaluated a series of tetraphenyl metalloporphyrins and halogenated tetraphenyl metalloporphyrin derivatives for binding to ligands with oxygen containing functional groups, using methanol, acetic acid and acetone as examples. Experimental binding constants identified three metalloporphyrins with good binding to all three ligands: MgTPFP, MgTPPBr₈ and ZnTPPBr₈ as well as a range of porphyrins binding to select ligands. Based on these results the optimal porphyrins can be selected for the desired binding interactions. We also show how to use DFT calculations to evaluate the potential binding between a metalloporphyrin and a ligand, which is deduced from free energies of binding ΔG , charge transfer ΔQ , and change of metal spin state. Computations on unsubstituted porphyrins in lieu of tetraphenyl porphyrin systems yield reliable predictions of binding interactions with good correlation to the corresponding experimental data. The calculations have also yielded interesting insights into the effect of halogenation in the β -position on the binding to ligands with oxygen containing functional groups.

Introduction

Metalloporphyrins are well known for their capability to bind ligands at the axial positions of the metal center.¹ This results in a plethora of functionality, based on varying electronic properties of the metalloporphyrin, increased reactivity of the axial ligands, and the possible formation of supramolecular entities.^{2–8} Compounds with several metalloporphyrin units have attracted particular interest, since they allow for a stronger host–guest interaction due to cooperativity or preorganization. Amongst these, bisporphyrin molecular tweezers and clips with two metalloporphyrin units connected by a linker are particularly well known for their capacity of ditopic binding to guest molecules.^{9,10} This is the foundation of Exciton-Coupled Circular Dichroism (ECCD), aimed at the determination of the absolute stereochemistry of guest molecules.¹¹ Recently, our group has extended the scope of such bisporphyrin tweezers to determine the relative stereochemistry of small flexible molecules using NMR spectroscopy *via* conformational restriction of the bound guest.^{12–14} In these studies, as in most ECCD investigations, zinc is the metal of choice. While zinc binds favorably to guests with nitrogen-containing functional groups, its affinity to oxygen-containing functional groups is much lower.¹⁵

Most molecules of interest for these analytical methods do not possess two nitrogen-containing functional groups, the presence of at least two oxygen-containing functional groups is more likely.¹⁶ Low affinity between host and guest may be compensated for in ECCD by a huge excess of the guest, since the signal originating from the bisporphyrin, and not the guest is detected. In our NMR-based method, instead we observe the signal originating from the guest molecule. Thus, a huge excess of guest is not tolerable, since it would result in NMR signals that are dominated by the contribution of unbound guest molecules.

Our aim therefore is to identify metalloporphyrins with increased affinity to oxygen-containing functional groups. On one hand it has been shown that this affinity may be increased by a derivatization of the porphyrin that reduces the electron density at the metal center.¹⁵ On the other hand, choice of suitable “oxophilic” metals may result in stronger binding.¹¹ Also, it should be possible to combine both parameters. Although both approaches have been reported, we found that a systematic approach is lacking. A more complete knowledge of possible binding interactions between metalloporphyrins and common oxygen-containing functional groups also should make it possible to choose a metal based on the desired binding behavior. For example, differences in affinity might be exploited to design bismetalloporphyrins that selectively bind components of a mixture. To facilitate both ECCD and NMR studies our aim is to find porphyrins binding strongly enough so that interactions between host and guests are easily detectable without requiring guest excess.

Department of Chemistry-BMC, Uppsala University, Box 576, Uppsala, S-75123, Sweden. E-mail: adolfgogoll@kemi.uu.se

†Electronic supplementary information (ESI) available. See DOI: 10.1039/c8dt02432d



To achieve this, we investigate here a series of metallated porphyrins, including, based on previous results with electron poor metal centers a number of halogenated metalloporphyrins.^{15,17}

To facilitate NMR studies most of the investigated metals are diamagnetic. We have also performed DFT calculations in order to devise a method for accurate prediction of whether a metalloporphyrin will bind strongly enough to oxygen containing functional groups to be detectable even at low guest concentrations. For practical reasons, some of the calculations have been carried out for simplified systems, while others used the full molecular structures used in the experimental study.

Results and discussion

Metalloporphyrins

The porphyrins used in this study are *meso*-tetraphenylporphyrins (TPP) and derivatives thereof. This porphyrin is widely used in host guest chemistry, molecular electronics and molecular materials due to the relatively good solubility in organic solvents.^{2–8,18} The TPP derivatives chosen are the tetra(pentafluorophenyl)porphyrin (TPFPF) and the octabromo tetraphenylporphyrin (TPPBr₈). Metals were chosen based on their prospective oxophilicity. For example, on the oxophilicity scale calculated by Kepp Mg²⁺ is more oxophilic than Zn²⁺.¹⁹ The insertion of metals into the porphyrins followed standard methods.^{20,21} All metalloporphyrins are shown in Scheme 1.

Ligands

The target of this study is to achieve reasonably strong binding of the common oxygen containing functional groups –OH, –COOH and –C=O to metalloporphyrins. To keep the studied systems as simple as possible the ligands chosen were methanol, acetic acid and acetone. The nitrogen containing ligand

pyridine that is very well studied as a ligand to porphyrins was included as a reference.^{22–24}

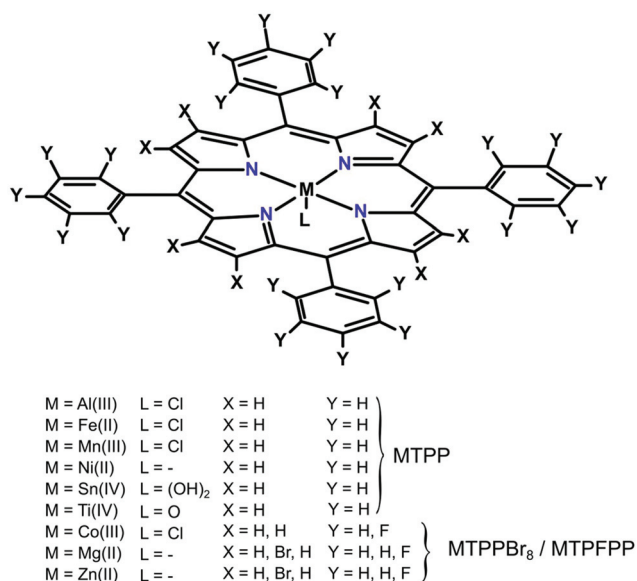
Determination of binding constants

While the standard way to measure binding constants involving porphyrins is UV-vis spectroscopy due to simplicity and high sensitivity, we opted for NMR spectroscopy instead. This allows for direct observation of ligand : metalloporphyrin stoichiometry. Furthermore, possible interference of water, the presence of which is difficult to avoid completely, may be indicated in the spectra. Binding of water during titration of porphyrins may result in ligand exchange rather than binding. In UV-vis titrations, this situation would result in a possibly very small and difficult to detect red-shift. For the NMR titrations, it needs to be considered that the chemical shifts of ligand protons are affected by two counteracting effects. Binding to the Lewis-acidic metal is expected to result in an increase of chemical shift ($\Delta\delta > 0$), whereas the anisotropy effect of the porphyrin ring system results in decreased chemical shift ($\Delta\delta < 0$). In practice, normally the anisotropy effect dominates. The gradual change in chemical shift is recorded and the binding constants are then obtained by non-linear fitting (Fig. 1). To probe if the binding constant was directly related to the magnitude of the chemical shift change $\Delta\delta$, this parameter also was recorded at the same $[G]/[H] = 1$ ratio for various metalloporphyrin complexes with G = ligand and H = metalloporphyrin (Table 2).

The stoichiometry of the formed complexes are mostly 1 : 1 [eqn (1)], but the zinc porphyrins form 1 : 2 complexes with pyridine [eqn (2)]. As mentioned above, the stoichiometry can be directly observed in NMR as the integrals of the peaks correlates to the number of protons with that chemical shift.

$$K_1 = \frac{[HG]}{[H][G]} \quad (1)$$

$$K_2 = \frac{[HG_2]}{[H][G]^2} \quad (2)$$



Scheme 1 Metalloporphyrins and the naming used in this study.

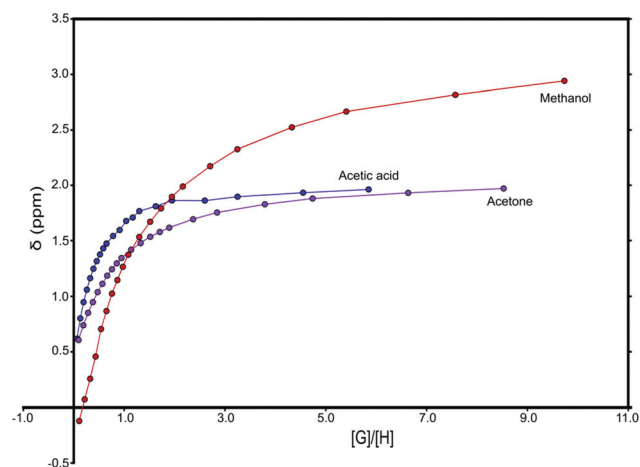


Fig. 1 NMR titration curves for MgTPFPF with oxygen containing ligands.



The binding constants can then be calculated from the measured NMR data utilizing an iterative fitting program, using eqn (3), for 1:1 stoichiometry and eqn (4) for 1:2 stoichiometry.²⁵

$$[HG] = \frac{1}{2} \left\{ \left([G]_0 + [H]_0 + \frac{1}{K_a} \right) - \sqrt{\left([G]_0 + [H]_0 + \frac{1}{K_a} \right)^2 - 4[H]_0[G]_0} \right\} \quad (3)$$

$$[G]^3(A) + [G^2](B) + [G](C) - [G]_0 = 0 \quad (4)$$

with: $A = (K_1K_2)$, $B = \{K_1(2K_2[H]_0 - K_2[G]_0 + 1)\}$, $C = \{K_1([H]_0 - [G]_0) + 1\}$.

The measured binding constants are presented in Table 1. In Fig. 2 the binding constants to different metallotetraphenylporphyrin-(MTPP) complexes are visualized.

As can be clearly seen the metals binding to MeOH are Al(III), Co(III), Mg(II) and Zn(II). All these also bind to AcOH, with the exception of Mg(II) due to demetallation following porphyrin protonation.²¹ For Mg(II) we also observe binding to acetone. There is only one case where we observe binding to an oxygen containing functional group but not to methanol. This is in the case of Ti(IV) which binds readily to AcOH but to no other ligand. For Co(III)TPPCL the binding constant to pyridine is too high to be determined by this method. Here, pyridine binds to the metal without any exchange with free pyridine in the solution. This results in one set of signals for the bound pyridine and another set of signals for free pyridine at $[G]/[H] > 1$.

Somewhat surprisingly, Fe(II)TPP, Mn(III)TPPCL and Ni(II)TPP show no binding, as well as the Sn(OH)₂TPP complex with two anionic hydroxide ligands on the metal which apparently bind more strongly than uncharged methanol ligands.^{22,26,27}

Introduction of halogen substituents, either in the *β*-positions (*i.e.* TPPBr₈) or as tetrakis(pentafluorophenyl) sub-

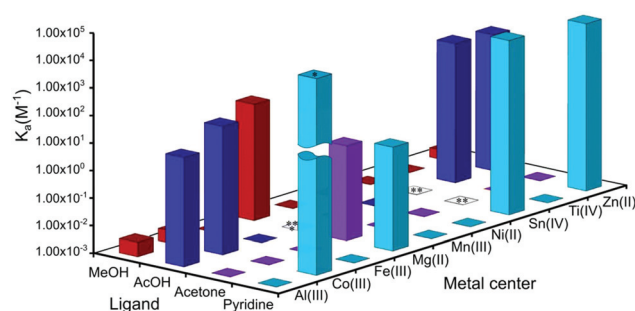


Fig. 2 Binding constants K_a for metallotetraphenylporphyrins (MTPP). * K_a too high to measure. **Not measured. ***Demetallation. Data from Table 1.

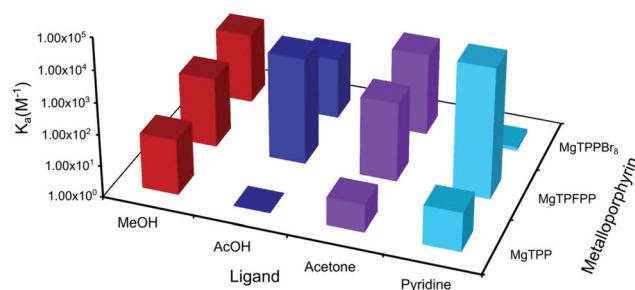


Fig. 3 Binding constants K_a for various ligands binding to MgTPP, MgTPFP and MgTPPBr₈. Data from Table 1.

stituents (*i.e.* TPFP) results in enhanced binding to the oxygen containing ligands (Fig. 3–5).

When comparing MgTPFP and MgTPPBr₈ to MgTPP the binding to MeOH is slightly increased (factor 2.6 and 2.8, respectively), but the binding to acetone more substantially so (factor 42 and 51, respectively). It is also interesting to note that the well-known demetallation of MgTPP under acidic con-

Table 1 Binding constants K_a for various metalloporphyrins from NMR titrations (1:1 stoichiometry, CDCl₃ solutions). Standard errors are given in brackets

Porphyrin	K_a (M ⁻¹), with standard error in brackets							
	Methanol		Acetic acid		Acetone		Pyridine	
Al(TPP)Cl	3.8×10^{-3}	$(2.2 \times 10^{-2})^a$	2.8×10^1	(2.6×10^{-4})	nb	—	nb	—
Co(TPP)Cl	3.2×10^{-3}	(1.5×10^{-4})	1.7×10^2	(4.9×10^{-3})	nb	—	cc	—
Co(TPFP)Cl	5.4×10^1	(8.9×10^{-3})	1.8×10^4	(3.6×10^{-4})	nb	—	cc	—
Fe(TPP)	nb	—	nb	—	nb	—	nb	—
Mg(TPP)	5.7×10^1	(3.0×10^{-2})	D	—	7.4	(1.4×10^{-3})	1.7×10^1	(0.2)
Mg(TPFP)	1.5×10^2	(8.6×10^{-2})	2.3×10^3	(8.0×10^{-2})	3.1×10^2	(2.9×10^{-2})	1.3×10^4	(1.9×10^{-2})
Mg(TPP)Br ₈	1.6×10^2	(4.6×10^{-2})	8.2×10^2	(1.0×10^{-1})	3.8×10^2	(2.2×10^{-2})	1.3	(9.3×10^{-2})
Mn(TPP)Cl	nb	—	nb	—	nb	—	nb	—
Ni(TPP)	nb	—	nb	—	nb	—	nb	—
Sn(TPP)OH ₂	nb	—	nm	—	nm	—	1.3×10^4	(6.6×10^{-3})
O=Ti(TPP)	nb	—	4.8×10^2	(8.1×10^{-4})	nb	—	nb	—
Zn(TPP)	2.4×10^{-3}	$(3.8 \times 10^{-3})^a$	3.6×10^2	(1.0×10^{-3})	nb	—	6.8×10^3	(0.2) ^b
Zn(TPFP)	1.5×10^1	(1.0×10^{-2})	1.2×10^3	(1.1×10^{-3})	1.1×10^{-2}	(1.4×10^{-3})	2.7×10^4	(0.5) ^b
Zn(TPP)Br ₈	2.0×10^1	(1.1×10^{-2})	2.5×10^2	(3.1×10^{-3})	6.0×10^{-3}	(1.4×10^{-3})	3.2×10^2	$(6.0 \times 10^{-2})^b$

nb = no binding; nm = not measured; cc = cannot calculate, binding too strong; D = demetallation. ^a Very broad signals. ^b Stoichiometry 1:2.



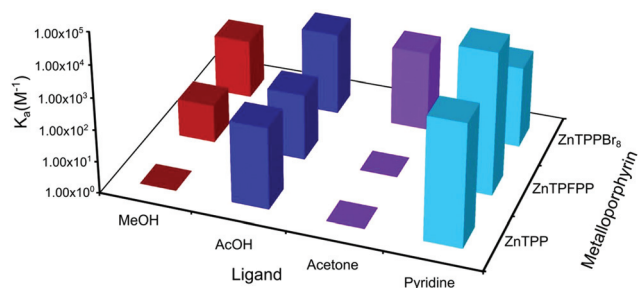


Fig. 4 Binding constants K_a for various ligands binding to ZnTPP, ZnTPFP and ZnTPPBr₈. Data from Table 1.

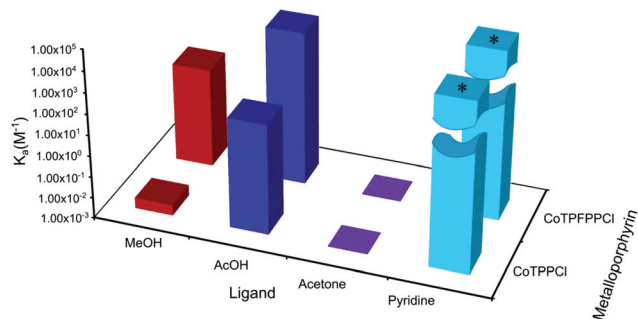


Fig. 5 Binding constants K_a for various ligands binding to cobalt-metalloporphyrins CoTPP and CoTPFP. Data from Table 1.

ditions is not observed for MgTPFP and MgTPPBr₈ allowing binding to all three oxygen containing functional groups.²¹ Another interesting feature is the low binding affinity of MgTPPBr₈ to pyridine.

For the zinc porphyrins the binding to methanol increases for the halogenated MTPFP and MTPPBr₈ (factors 625 and 8333, respectively) but the binding to acetic acid shows only slight changes (factors 3.3 and 0.7, respectively). The most interesting point is that ZnTPPBr₈ binds to acetone, unlike the ZnTPP and ZnTPFP. This makes ZnTPPBr₈ one of the most versatile metalloporphyrins investigated in this study as it binds to all four ligands with similar strength.

In the cobalt series the CoTPPBr₈ was excluded as the reported synthetic procedures were unsuccessful in our hands. For the CoTPFP porphyrin the binding to both methanol and acetic acid was increased as compared to CoTPP. None of the cobalt porphyrins binds to acetone.

The ¹H chemical shift changes for the target ligands' methyl groups are shown in Table 2. The largest changes for each ligand are indicated in bold face. The most apparent result is that all Mg(II)-metalloporphyrins produce a comparably large chemical shift change. The MgTPFP has one of the highest values for all tested ligands. However, we cannot see a clear correlation to the binding constant, indicating that $\Delta\delta$ is influenced by additional parameters, *e.g.* the distance between the methyl protons and the porphyrin ring. ZnTPPBr₈ has good binding constants to all the target ligands, but the change in chemical shift is less impressive. The same can be

Table 2 Chemical shift changes $\delta = \delta_{\text{bound}} - \delta_{\text{free}}$ for oxygen containing ligands (CDCl₃ solution)

Porphyrin	¹ H chemical shift change $\Delta\delta$ (ppm) for ligand protons at 1 : 1 host : guest ratios					
	Methanol		Acetic acid		Acetone	
	δCH_3	$\Delta\delta$	δCH_3	$\Delta\delta$	δCH_3	$\Delta\delta$
Al(TPP)Cl	3.01	−0.48	2.09	−0.01	nb	—
Co(TPP)Cl	3.47 ^a	−0.02	2.07 ^a	−0.03	nb	—
Co(TFPFP)Cl	1.15	−2.34	2.09	−0.01	nb	—
Mg(TPP)	2.25	−1.24	D	—	2.02	−0.15
Mg(TFPFP)	1.38	−2.11	1.68	−0.42	1.35	−0.82
Mg(TPP)Br ₈	1.76	−1.73	1.64	−0.46	1.56	−0.61
O=Ti(TPP)	nb	—	2.09	−0.01	nb	—
Zn(TPP)	3.33	−0.16	2.09	−0.01	nb	—
Zn(TFPFP)	2.79	−0.70	2.04	−0.06	2.13	−0.04
Zn(TPP)Br ₈	3.12	−0.37	1.99	−0.11	2.12	−0.05

^aThis metalloporphyrin could not be obtained free from water, thus the binding to oxygen containing functional groups here includes competition with water. nb = no binding.

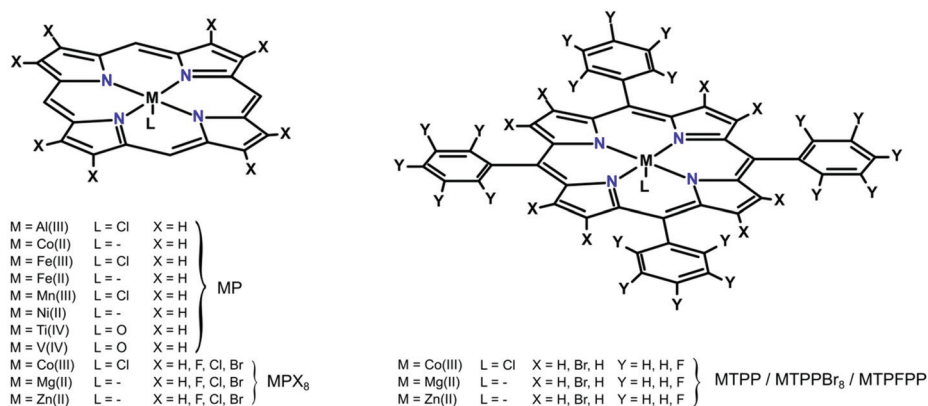
said for the cobalt porphyrins. With the exception of the CoTFPFP binding to methanol, the chemical shift changes are small whereas the binding constants are good (Table 1). Thus, $\Delta\delta$ seems more dependent on the metal of the metalloporphyrin than the binding strength. This might be a result of the two counteracting effects, *i.e.* an electronic effect from the metal and an anisotropy effect from the porphyrin ring (and possibly the phenyl rings).

Computational studies

To rationalize the experimental results a DFT study was carried out. As almost all porphyrins binding to R–OH also bind to R–COOH the smaller hydroxyl group was chosen as functional group. Most of the calculations were carried out on simplified metalloporphyrin systems (MP) to reduce computational costs (*vide infra*). Calculations were also performed on β -halogenated porphyrins to investigate the effect of halogenation on the binding constant as well as the change in charge distribution. To investigate if calculations on the rudimentary metalloporphyrins (MP) gave results indicative of the full systems (*i.e.*, MTPP), we performed a limited number of calculations on MTPP as well as halogenated MTPP derivatives (MTPFP and MTPPBr₈). All investigated systems are shown in Scheme 2.

Calculations on the metalloporphyrins were performed on the DFT level using the B3LYP-D3 method as described by Grimme.^{28–32} This method has previously been shown to predict the geometry, vibrations and electronic structure of porphyrin systems with reasonable accuracy.^{33,34} No solvent model was used in these calculations since their primary purpose was a comparative binding study of various metalloporphyrins independent of choice of solvent. Choice of solvent has a substantial impact on complexations and therefore without a solvent model the calculated binding constants will be considerably higher than the experimentally measured values, but should show the same trends. As it is well known





Scheme 2 Structures and designations of metalloporphyrins included in the calculations.

that metals can adopt several different spin states the three lowest spin states were investigated for each metalloporphyrin and its corresponding metalloporphyrin-methanol complex.³⁵ For metals with an even number of electrons the states with total spin $S = 0, 1, 2$ and multiplicity $M = 1, 3, 5$ were calculated and for metals with an odd number of electrons the states with a total spin $S = 1/2, 3/2, 5/2$ and multiplicity $M = 2, 4, 6$ were calculated. The lowest states were used in the calculations of binding energies and theoretical binding constants. The energies of the chosen metalloporphyrins were calculated both with and without methanol as ligand. To investigate the change in charge distribution a natural population analysis (NPA) was carried out on the free methanol ligand and on its complexes.^{36–38}

Geometries and symmetries

The simple MPs have planar structures whereas the substituted TPP, TPFPP and TPPBr₈ porphyrins have more complex geometries. The phenyl rings are a source of steric repulsion and depending on their position will affect the symmetry and the energy of the system.³⁹ Different rotamers with respect to the phenyl rings were used as starting geometries in order to find the most stable geometry. Our calculations show that for the free TPP metalloporphyrins (CoTPPCL, MgTPP, ZnTPP) the most stable positioning of the phenyl groups is with rings in opposing *meso*-positions in parallel orientation ($\parallel$, angle between ring planes $\approx 0^\circ$), but when binding to methanol both

MgTPP and ZnTPP have the opposite rings in an angular orientation ($\times$, angle between ring planes $\approx 40^\circ$) (Fig. 6).

For the TPFPP metalloporphyrins the ($\times$) configuration is favoured in both the free porphyrins and the complexes. The only exceptions are MgTPFPF where the ($\parallel$)-configuration is favoured for the free porphyrin, and ZnTPFPF where the ($\parallel$)-configuration is favoured in the methanol complex.

In the MTPPBr₈ the *meso*-phenyl rings are restricted into adopting one position due to repulsion from the bromo substituents in the β -positions. These bromo substituents are also the cause of substantial distortions of the porphyrin ring planarity, resulting in a saddle shaped structure (Fig. 7).⁴⁰

We have also observed a distortion of the porphyrin ring in the methanol complexes of Co(III)TPPCL, Co(III)TPFPFCL and to

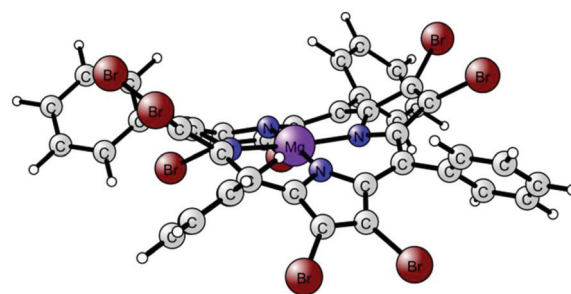


Fig. 7 The structure of MgTPPBr₈.

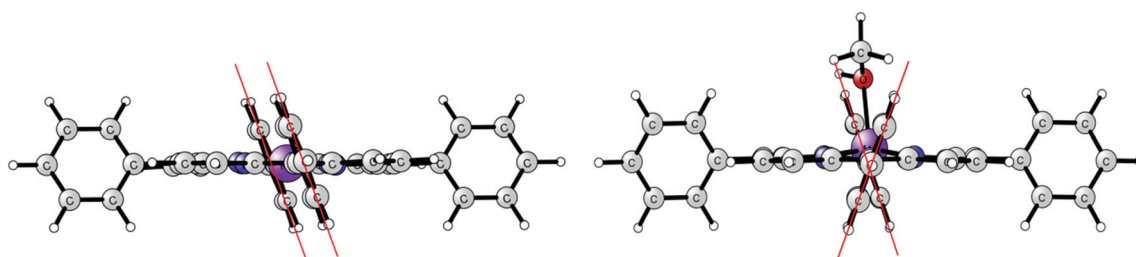


Fig. 6 MgTPP, position of phenyl rings when free and when complexing methanol.



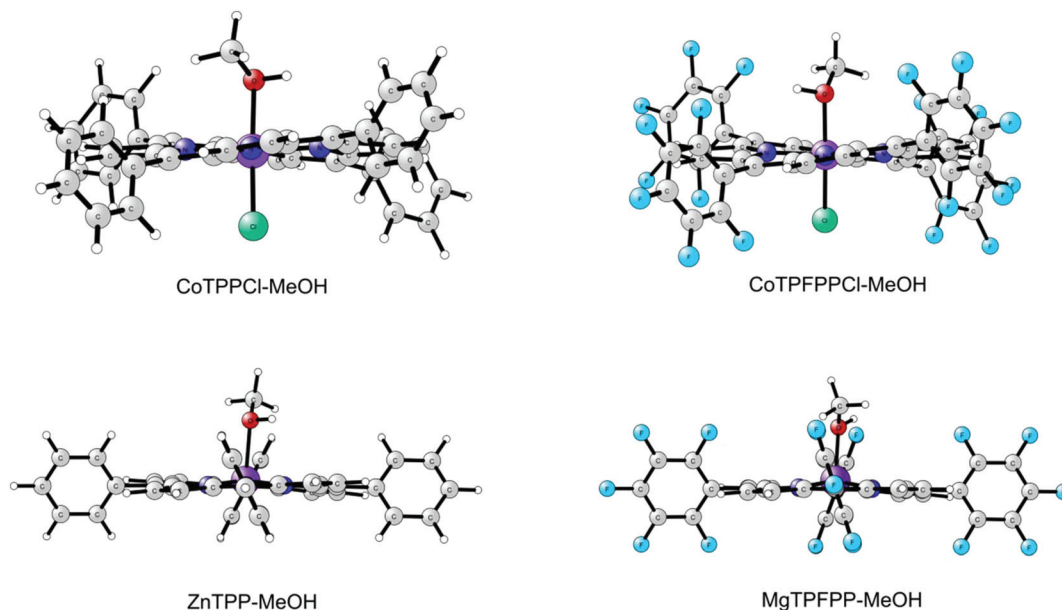


Fig. 8 Distortions of the porphyrin ring planarity in the methanol complexes of CoTPPCL, CoTPFPPL, ZnTPP and MgTPFPPL.

a lesser extent in the complexes of Zn(II)TPP and MgTPFPPL, Fig. 8. This kind of distortion was previously observed for Co(II)TPP by de Melo *et al.*⁴¹

Free energies of binding

The lowest calculated energies of the free porphyrins and their methanol complexes were used to determine the free energies of binding ΔG . From these energies the theoretical binding constants were calculated using (eqn (5)).

$$K_a = e^{-\frac{\Delta G}{RT}} \quad (5)$$

where K_a is the binding constant, ΔG the free energy of binding, R the universal gas constant and T the absolute temperature. The calculated free energies of binding and the theoretical binding constants are shown in Table 3.

The results are presented graphically in Fig. 9. As can be seen, MgP shows the strongest binding interaction followed by Co(III)P. For the other metallocporphyrins ΔG are at the same level or weaker than for Zn(II)P. As it is known from the literature that Zn(II)P binds weakly to alcohols the porphyrins of most interest are the ones that show stronger interactions, *i.e.* (Mg(II)P and Co(III)P).¹⁵

The calculations on the β -halogenated porphyrins showed a stronger binding to methanol, Table 4 and Fig. 10. This effect seems independent of which halogen is used. As it has been observed that halogenation on the phenyl rings of TPP increases the binding constant this was an expected result.^{15,17} Worth noting is that for Co(III) and Mg(II) there is a clear trend with increased ΔG in the series F < Cl < Br, whereas for Zn(II) the brominated porphyrin shows the weakest increase in free energy of binding.

Table 3 Calculated binding energies ΔG and theoretical binding constants $K_{a,theo}$ for the formation of MP-MeOH complexes

Complex formation (spin state of metal indicated)	Change in spin	ΔG (kJ mol ⁻¹)	$K_{a,theo}$
¹ Al(III)PCL + MeOH → [¹ Al(III)PCL-MeOH]	1 → 1	-20.4	3.70×10^3
⁵ Co(III)PCL ^a + MeOH → [¹ Co(III)PCL-MeOH]	5 → 1	-28.4	9.37×10^4
² Co(II)P + MeOH → [² Co(II)P-MeOH]	2 → 2	-22.0	7.06×10^3
⁶ Fe(III)PCL + MeOH → [⁴ Fe(III)PCL-MeOH]	6 → 4	7.0	—
³ Fe(II)P ^b + MeOH → [⁵ Fe(II)P-MeOH] ^c	3 → 5	-18.9	2.06×10^3
¹ Mg(II)P + MeOH → [¹ Mg(II)P-MeOH]	1 → 1	-48.1	2.65×10^8
⁵ Mn(III)PCL + MeOH → [⁵ Mn(III)PCL-MeOH]	5 → 5	-10.6	7.18×10^1
¹ Ni(II)P + MeOH → [³ Ni(II)P-MeOH]	1 → 3	-21.8	6.72×10^3
O= ¹ Ti(IV)P + MeOH → [O= ³ Ti(IV)P-MeOH]	1 → 3	235.5	—
O= ² V(IV)P + MeOH → [O= ² V(IV)P-MeOH]	2 → 2	-11.1	8.70×10^1
¹ Zn(II)P + MeOH → [¹ Zn(II)P-MeOH]	1 → 1	-21.3	5.41×10^3

^a The singlet state is only 2.62 kJ mol⁻¹ higher in energy. ^b The quintet state is only 0.58 kJ mol⁻¹ higher in energy. ^c The triplet state is only 1.99 kJ mol⁻¹ higher in energy.



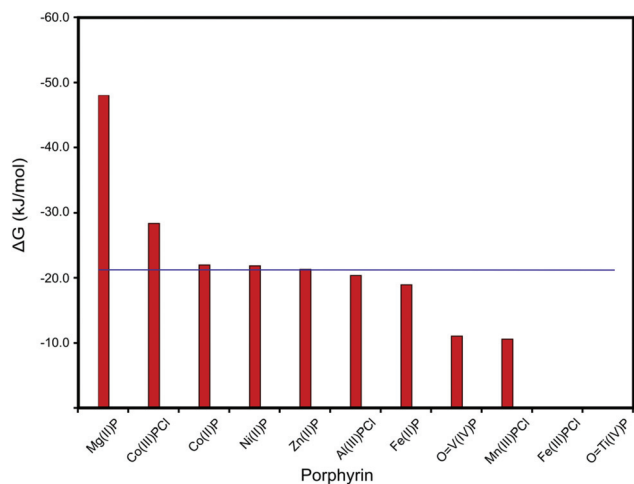


Fig. 9 The theoretical free energies of binding ΔG for the formation of a complex between the studied metalloporphyrins and methanol. The horizontal line is positioned at the value for Zn(II)P.

The free energies of binding for the complexation between the MTPP and methanol are shown in Table 5. It can be seen that for Co(III)TPP and Zn(II)TPP the binding constants are similar to those for their MP congeners. For Mg(II)TPP the difference is almost one order of magnitude when compared to Mg(II)P. This supports the use of the simpler porphyrin derivatives (*i.e.* without *meso*-phenyl substituents) in most of the calculations.

The halogenated MTPP derivatives TPFPP and TPPBr₈ show about as high or higher free energies of binding ΔG than the MTPP congener but, depending on the metal, different porphyrins present the best result (Table 6). With the distorted geometries of especially the MTPPBr₈ in mind, a direct comparison is difficult. The structures of both the free porphyrin and the complex might be significantly stabilized by solvent. Further computational studies on those systems with different solvent models and ligands would be required to estimate this effect.

The spin state of the metal

The investigated MPs have different spin states as their most stable state. In Table 3 the change in spin state when forming

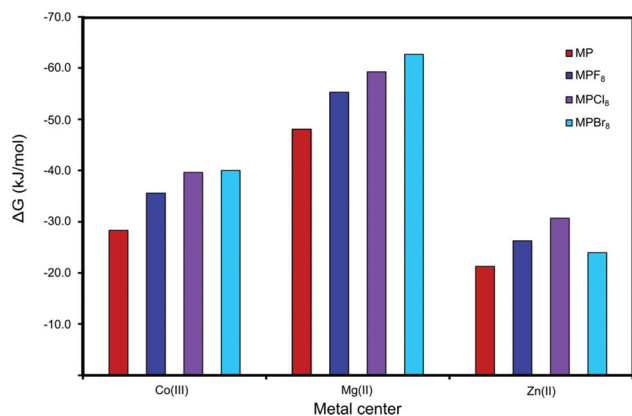


Fig. 10 Calculated free energies of binding ΔG for the formation of MPX₈-MeOH complexes.

Table 5 Calculated free energies of binding ΔG and theoretical binding constants $K_{a,theo}$ of methanol with MP and MTPP

Complex formation	ΔG (kJ mol ⁻¹)	$K_{a,theo}$
Co(III)PCL + MeOH → [Co(III)PCL-MeOH]	-28.4	9.37×10^4
Co(III)TPPCL + MeOH → [Co(III)TPPCL-MeOH]	-26.7	4.74×10^4
Mg(II)P + MeOH → [Mg(II)P-MeOH]	-48.1	2.65×10^8
Mg(II)TPP + MeOH → [Mg(II)TPP-MeOH]	-52.2	1.44×10^9
Zn(II)P + MeOH → [Zn(II)P-MeOH]	-21.3	5.41×10^3
Zn(II)TPP + MeOH → [Zn(II)TPP-MeOH]	-23.3	1.24×10^4

the MP-MeOH complex is shown. As it is known that DFT has limitations regarding the calculation of spin states we also indicate when another spin state is very close in energy.^{35,42} Considering the free energy of binding ΔG it is indicated that the Fe(III) and Ti(IV) porphyrins do not form stable complexes, as their ΔG are positive. For the Al(III), Co(II), Mg(II), Mn(III) and Zn(II) porphyrins there is no change in spin state during complex formation. Co(III) porphyrin goes from a quintet to a singlet when forming the complex and also shows the second highest energy of binding. Both Fe(II) and Ni(II) porphyrin complexes change to a higher spin state. In comparison with the experimental results (Table 1) we might expect detectable binding for any metalloporphyrins that have similar or higher free energy of binding than Zn(II)P. However, of these Fe(II) and

Table 4 Calculated free energies of binding ΔG and theoretical binding constants $K_{a,theo}$ for the formation of MPX₈-MeOH complexes

Complex formation halogen	ΔG (kJ mol ⁻¹)	$K_{a,theo}$
Co(III)PF ₈ Cl + MeOH → [Co(III)PF ₈ Cl-MeOH]	-32.6	5.18×10^5
Co(III)PCL ₈ Cl + MeOH → [Co(III)PCL ₈ Cl-MeOH]	-33.1	6.34×10^5
Co(III)PBr ₈ Cl + MeOH → [Co(III)PBr ₈ Cl-MeOH]	-32.3	4.65×10^5
MgPF ₈ + MeOH → [MgPF ₈ -MeOH]	-55.3	4.96×10^9
MgPCL ₈ + MeOH → [MgPCL ₈ -MeOH]	-59.3	2.43×10^{10}
MgPBr ₈ + MeOH → [MgPBr ₈ -MeOH]	-62.7	9.72×10^{10}
Zn(II)PF ₈ + MeOH → [Zn(II)PF ₈ -MeOH]	-26.3	4.07×10^4
Zn(II)PCL ₈ + MeOH → [Zn(II)PCL ₈ -MeOH]	-30.7	2.43×10^5
Zn(II)PBr ₈ + MeOH → [Zn(II)PBr ₈ -MeOH]	-24.0	1.60×10^4



Table 6 Calculated free energies of binding ΔG and theoretical binding constants $K_{a,theo}$ of methanol with MTPP and halogenated MTPP derivatives

Complex formation	ΔG (kJ mol ⁻¹)	$K_{a,theo}$
Co(III)TPPCL + MeOH $\rightarrow$ [Co(III)TPPCL-MeOH]	-26.7	4.74×10^4
Co(III)TPFPCL + MeOH $\rightarrow$ [Co(III)TPFPCL-MeOH]	-43.8	4.80×10^7
Co(III)TPPBr ₈ Cl + MeOH $\rightarrow$ [Co(III)TPPBr ₈ Cl-MeOH]	-26.2	3.89×10^4
Mg(II)TPP + MeOH $\rightarrow$ [Mg(II)P-MeOH]	-52.2	1.44×10^9
Mg(II)TPFPP + MeOH $\rightarrow$ [Mg(II)TPFPP-MeOH]	-44.8	7.07×10^7
Mg(II)TPPBr ₈ + MeOH $\rightarrow$ [Mg(II)TPPBr ₈ -MeOH]	-71.7	3.70×10^{12}
Zn(II)TPP + MeOH $\rightarrow$ [Zn(II)TPP-MeOH]	-23.3	1.24×10^4
Zn(II)TPFPP + MeOH $\rightarrow$ [Zn(II)TPFPP-MeOH]	-31.8	3.71×10^5
Zn(II)TPPBr ₈ + MeOH $\rightarrow$ [Zn(II)TPPBr ₈ -MeOH]	-40.5	1.25×10^7

Ni(II) porphyrins did not show any binding in the experiments. Experimental binding is observed for those metalloporphyrins (Al(III), Co(III), Mg(II) and Zn(II)) that preserve or lower their spin state when forming the complex. The non-binding Fe(II) and Ni(II) porphyrins both would increase their spin state to form a stable complex according to the computations. Apparently, here an increase of spin state indicates non-binding although a negative ΔG indicates binding. It would be interesting to extend this study to probe whether this is a general trend in metalloporphyrins.

Bond length

We also investigated the ligand–metal bond length, Fig. 11. We can see a clear correlation between the binding energies and the metal–ligand bond length. A stronger binding corresponds to a shorter bond. For the linear fitting the period 3 metals (Mg and Al) were omitted as they are smaller and differ in orbital composition as compared to the period 4 transition and post-transition metals.

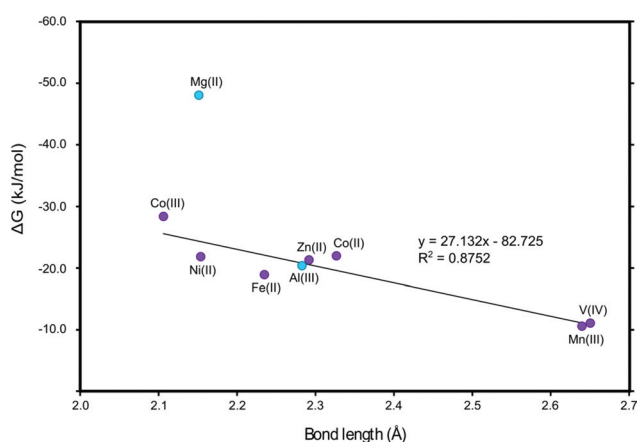


Fig. 11 Calculated metal–ligand bond length vs. free energy of binding ΔG for MP-MeOH complexes. Mg and Al are included in the figure for comparison but not included in the data fitting. The reason for this is that they belong to a different period and have a substantially smaller ion radius, and thus are not directly comparable to the other metals in the study.

The calculated O–M bond lengths for the complexes of the full MTPP and halogenated MTPP derivative systems can be related to the change in chemical shifts observed on methanol during the experimental determination of the binding constants (Fig. 12). This effect is caused mainly by the anisotropy effect from the porphyrin ring current and is therefore directly related to the distance between the measured methyl group protons and the porphyrin ring.

Charge distribution

A natural population analysis (NPA) was performed on the optimized structures of the free metalloporphyrins and the methanol-metalloporphyrin complexes. To probe how the halogenation in β -position affected the charge distribution the change in natural charge for the metal and nitrogen atoms were investigated (Table 7). Interestingly the metals of Mg(II)PX₈ and of Zn(II)PX₈ become more electron depleted, as expected when introducing electron withdrawing groups, whereas the metal of Co(III)PX₈Cl gets more electron rich. The nitrogens of Co(III)PX₈Cl show mixed behaviour which suggest that looking on only the metal and nitrogens might be an oversimplification in the case of the Co(III)PX₈Cl porphyrins. Further study of this, including comparison of orbital overlap and more metals would be interesting. For all the studied porphyrins we also observe an increased free energy of binding upon β -halogenation (Fig. 10).

Investigating the change in charge distribution ΔQ in the MTPP and halogenated MTPP derivatives we can see a similar trend in the Mg(II) and Zn(II) metalloporphyrins (Table 8). For the Co(III) metalloporphyrins there is an increased charge on the nitrogen atoms in Co(III)TPFPCL, but very little change on the metal. In Co(III)TPPBr₈Cl however there is a large increase in negative charge on the metal and a decrease in negative charge on the nitrogen atoms.

The natural charges for free and complexed methanol were extracted (Table 9). A charge transfer can be observed in the porphyrin as well, but with the charge spread out over the system the change is more perspicuous for the methanol ligand. All of Al(III)PCL, Co(III)PCL, Mg(II)P and Zn(II)P show a negative change in charge when forming a complex with methanol. Those porphyrins also have calculated free energies of binding on the same level as Zn(II)P or higher. Co(II)P as



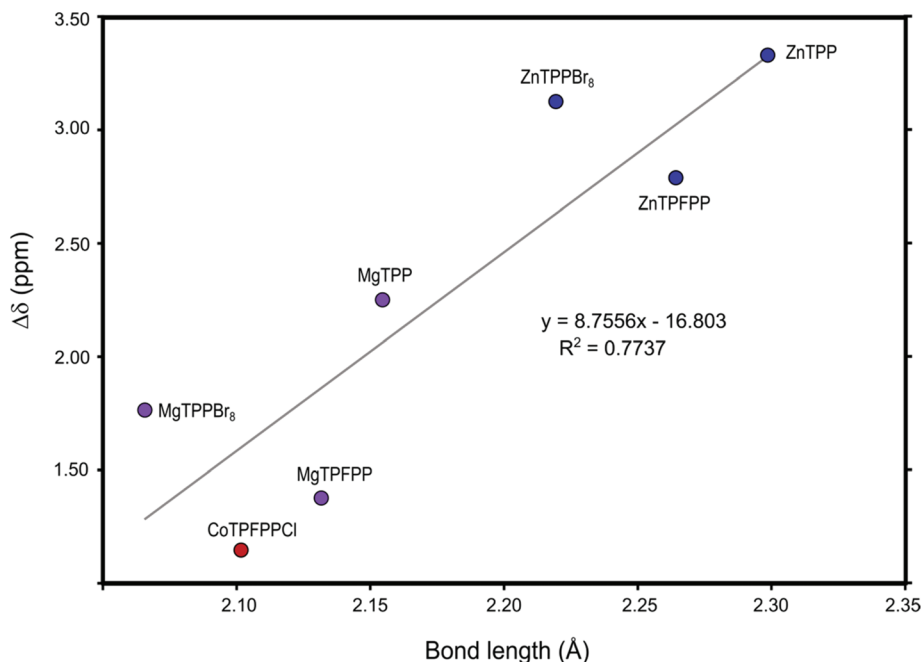


Fig. 12 Calculated M–O bond lengths compared to the experimentally determined changes in chemical shift $\Delta\delta$ for the methyl group on methanol in ^1H NMR. The CoTPFPFPPCl is not included in this figure as competition with water was observed during the experiments and this affects the binding interaction with the porphyrin (*vide supra*).

Table 7 Change in natural charge ΔQ on the metal and nitrogen atoms of MP and MPX_8 . The change in charge is calculated relative to MP and given in au

Porphyrin	Charge of metal	Average charge of N1–N4	ΔQ_{metal}	ΔQ_{N}
Co(III)P	1.790	−0.160	—	—
Co(III)PF ₈ Cl	1.782	−0.162	−0.008	−0.002
Co(III)PCl ₈ Cl	1.781	−0.155	−0.009	0.004
Co(III)PBr ₈ Cl	1.780	−0.154	−0.010	0.005
Mg(II)P	1.429	−0.682	—	—
Mg(II)PF ₈	1.436	−0.680	0.007	0.001
Mg(II)PCl ₈	1.446	−0.669	0.018	0.012
Mg(II)PBr ₈ ^a	1.451	−0.663	0.022	0.019
Zn(II)P	1.236	−0.649	—	—
Zn(II)PF ₈	1.244	−0.648	0.009	0.001
Zn(II)PCl ₈	1.254	−0.637	0.018	0.013
Zn(II)PBr ₈	1.258	−0.634	0.023	0.015

^a The LanL2DZ basis set was used on all metals.

Table 8 Change in natural charge ΔQ on the metal and nitrogen atoms of MTPP, MTPFPFPP and MTPPBr₈. The change in charge is calculated relative to MTPP and given in au

Porphyrin	Charge of metal	Average charge of N1–N4	ΔQ_{metal}	ΔQ_{N}
Co(III)TPPCL	1.788	−0.158	—	—
Co(III)TPFPFPPCl	1.789	−0.162	0.001	−0.004
Co(III)TPPBr ₈ Cl	1.742	−0.151	−0.046	0.007
Mg(II)TPP	1.427	−0.677	—	—
Mg(II)TPFPFPP	1.439	−0.674	0.012	0.003
Mg(II)TPPBr ₈ ^a	1.453	−0.663	0.026	0.015
Zn(II)TPP	0.616	−0.322	—	—
Zn(II)TPFPFPP	0.623	−0.321	0.006	0.002
Zn(II)TPPBr ₈	0.628	−0.317	0.012	0.005

^a The LanL2DZ basis set was used on all metals.

well as the non-binding Fe(II)P and Ni(II)P also have calculated free energies of binding on the same level as Zn(II)P but show a positive change in charge.

The charge transfer when methanol forms complexes with the MTPP, MTPFPFPP and MTPPBr₈ shows exactly the same pattern as observed for the simplified systems (MP) (Table 10).

In Table 11 the data calculated for the simplified MP systems are compared to the experimental results for the corresponding MTPPs. The calculations of free energies of binding ΔG result in a number of false positives, *i.e.* for (Fe(II), Mn(III) and Ni(II)), for which experimentally no binding to

methanol was found. In contrast, the change of charge on the ligand correlates well to the experimental results, with negative values indicating binding and positive values indicating non-binding. From this we conclude that the change in charge on the ligand gives a good indication of whether binding can be expected. As these are relatively easy calculations we suggest this to be a useful tool for evaluation of different ligand–metal complexes before synthesising the compounds.

In Table 12 the calculated binding energies and charge transfer for the full systems (TPP, TPFPFPP and TPPBr₈) are compared to the experimentally determined binding constant K_a . They show similar results to those obtained for the simplified systems. The change in charge on the methanol ligand seems



Table 9 Natural charges for atoms in free and metalloporphyrin-bound methanol. The change in charge ΔQ upon complexation is shown in brackets and is given in au. Porphyrins not forming stable complexes in the calculations are omitted

Porphyrin	Natural charges for methanol in porphyrin complexes. Difference to free methanol in brackets.				
	H (ΔQ)	O (ΔQ)	C (ΔQ)	CH ₃ (ΔQ)	Total change
None	0.24	−0.37	−0.16	0.10	
Al(III)P	0.51 (0.28)	−0.75 (−0.38)	−0.31 (−0.16)	0.22 (0.12)	−0.14
Co(III)P	0.52 (0.28)	−0.66 (−0.29)	−0.31 (−0.15)	0.22 (0.12)	−0.04
Co(II)P	0.26 (0.02)	−0.35 (0.03)	−0.16 (0.001)	0.11 (0.01)	0.06
Fe(II)P	0.24 (0.01)	−0.35 (0.03)	−0.15 (0.003)	0.11 (0.01)	0.05
Mg(II)P	0.52 (0.28)	−0.78 (−0.41)	−0.31 (−0.16)	0.22 (0.12)	−0.16
Mn(III)P	0.25 (0.01)	−0.34 (0.03)	−0.15 (0.01)	0.10 (0.01)	0.05
Ni(II)P	0.26 (0.02)	−0.34 (0.03)	−0.16 (0.001)	0.11 (0.01)	0.07
V(IV)P	0.25 (0.01)	−0.36 (0.03)	−0.15 (0.002)	0.10 (0.01)	0.04
Zn(II)P	0.51 (0.27)	−0.75 (−0.38)	−0.31 (−0.15)	0.21 (0.11)	−0.14

Table 10 Natural charges for atoms in free and metalloporphyrin-bound methanol. The change in charge ΔQ upon complexation is shown in brackets and is given in au. Some MPs are included for reference

Porphyrin	Natural charges for methanol in porphyrin complexes. Difference to free methanol in brackets				
	H (ΔQ)	O (ΔQ)	C (ΔQ)	CH ₃ (ΔQ)	Total change
None	0.24	−0.37	−0.16	0.10	
Co(III)P	0.52 (0.28)	−0.66 (−0.29)	−0.31 (−0.15)	0.22 (0.12)	−0.04
Co(III)TPPCL	0.52 (0.28)	−0.66 (−0.29)	−0.31 (−0.15)	0.22 (0.12)	−0.04
Co(III)TPFPCL	0.52 (0.28)	−0.66 (−0.30)	−0.31 (−0.15)	0.22 (0.12)	−0.04
Co(III)TPPBr ₈ Cl	0.52 (0.28)	−0.67 (−0.29)	−0.31 (−0.15)	0.22 (0.12)	−0.04
Mg(II)P	0.52 (0.28)	−0.78 (−0.41)	−0.31 (−0.15)	0.22 (0.12)	−0.16
Mg(II)TPP	0.52 (0.28)	−0.78 (−0.41)	−0.31 (−0.15)	0.22 (0.12)	−0.16
Mg(II)TPFP	0.52 (0.28)	−0.79 (−0.42)	−0.31 (−0.15)	0.22 (0.12)	−0.16
Mg(II)TPPBr ₈ ^a	0.52 (0.28)	−0.80 (−0.43)	−0.31 (−0.15)	0.22 (0.12)	−0.16
Zn(II)P	0.51 (0.27)	−0.75 (−0.38)	−0.31 (−0.15)	0.21 (0.11)	−0.14
Zn(II)TPP	0.51 (0.27)	−0.75 (−0.38)	−0.31 (−0.15)	0.21 (0.11)	−0.15
Zn(II)TPFP	0.51 (0.27)	−0.76 (−0.39)	−0.31 (−0.15)	0.22 (0.12)	−0.15
Zn(II)TPPBr ₈	0.52 (0.28)	−0.77 (−0.40)	−0.31 (−0.15)	0.22 (0.12)	−0.14

^a The LanL2DZ basis set was used on all metals.**Table 11** Relation between free energies of binding ΔG , change in spin states and change in charge on complexed MeOH to experimentally determined binding constants for the same metal centres

Metal ^a	DFT calculations (MP)			Experimental (MTPP)
	ΔG (kJ mol ^{−1})	ΔQ on MeOH	Change in spin state	Binding constant K_a (M ^{−1})
Al(III)	−20.4	−0.14	1 → 1	3.8×10^{-3}
Co(III)	−28.4	−0.04	5 → 1	3.2×10^{-3}
Fe(II)	−18.9	0.05	3 → 5	nb
Mg(II)	−48.1	−0.16	1 → 1	5.7×10^1
Mn(III)	−10.6	0.05	5 → 5	nb
Ni(II)	−21.8	0.07	1 → 3	nb
Zn(II)	−21.3	−0.14	1 → 1	2.4×10^{-3}

^a Only including metals binding in DFT study and also present in experimental study. nb = no binding.

related to the metal centre rather than the porphyrin as all the Co(III), Mg(II) and Zn(II) porphyrins show almost identical results. When comparing the calculated free energies of binding to the experimentally determined binding constants we obtain a good linear correlation with a $R^2 = 0.9$ (Fig. 13).

When performing calculations on the full systems instead of MP we also gain some additional insights into the relative strengths of the binding constants and the bond lengths (see ESI†), which can be related to chemical shift changes observed in the ¹H NMR spectra. For the purpose of predicting if a



Table 12 Comparison of free energies of binding ΔG and change in charge for complexed MeOH to experimentally determined binding constants K_a

Porphyrin	DFT calculations		Experimental
	ΔG (kJ mol ⁻¹)	ΔQ on MeOH	Binding constant K_a (M ⁻¹)
Co(III)TPPCL	-26.7	-0.04	3.20×10^{-3}
Co(III)TPFPFPP	-43.8	-0.04	5.36×10^1
Mg(II)TPP	-52.2	-0.16	5.65×10^1
Mg(II)TPFPFPP	-44.8	-0.16	1.54×10^2
Mg(II)TPPBr ₈	-71.7	-0.16	1.63×10^2
Zn(II)TPP	-23.3	-0.15	2.40×10^{-3}
Zn(II)TPFPFPP	-31.8	-0.15	1.53×10^1
Zn(II)TPPBr ₈	-40.5	-0.14	6.61×10^1

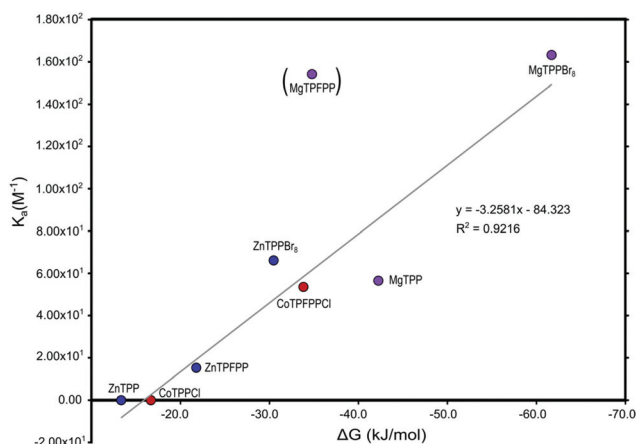


Fig. 13 Calculated free energies of binding ΔG plotted against the experimentally determined binding constant K_a , with MgTPFPFPP treated as an outlier. The free energy of binding for MgTPFPFPP is clearly underestimated in the DFT calculations. The R^2 -value with MgTPFPFPP included is 0.7.

certain metalloporphyrin will bind to methanol calculations on a simplified MP system are sufficient.

Conclusion

For the purpose of binding to oxygen containing ligands Mg(II)-porphyrins and to a lesser extent Co(III)- and Zn(II)-porphyrins are the prime choice. In particular, the porphyrin derivatives TPFPP and TPPBr₈ support binding to various ligands, *i.e.* MeOH, AcOH, and acetone. The MgTPFPFPP is the most versatile porphyrin derivative in this study. A weaker binding is observed for ZnTPPBr₈. Both the CoTPPCL and the ZnTPFPFPP bind to all ligands except acetone with reasonable strength. Also worth noting is that unlike most of the other porphyrins the MgTPPBr₈ binds stronger to the oxygen containing ligands than to pyridine. This could potentially be very useful for ligand recognition and separation of mixtures. Another interesting porphyrin is the O=TiTPP that only binds to the carboxylic acid but binds with a high binding constant.

This could also be used to discriminate between ligands or functional groups on the same ligand.

We also investigated the change in chemical shift of bound ligands caused by the porphyrin ring current anisotropy effect. With respect to maximization of this effect the TPFPP porphyrin stands out. The TPPBr₈ also gives better results than the TPP. The best choice to get both binding to all tested ligands and a large change in chemical shift is the MgTPFPFPP metalloporphyrin.

The DFT calculations using B3LYP-D3 show that the metalloporphyrins with both a high free energy of binding and a negative change in charge on the methanol ligand are those that experimentally have been found to bind methanol. Thus we propose to use such calculations as a relatively fast and easy method to predict binding between metalloporphyrins and prospective ligands. The calculations on the full MTPP systems showed correlations to the experimental binding constants and the observed changes in chemical shifts, indicating that they are a good representation of the studied systems. Furthermore, the calculations also indicated that formation of a stable complex cannot be expected if a change to a higher spin state of the metal is required. It would be interesting to confirm this preliminary observation by more extended computations involving further metalloporphyrins.

Experimental

Commercially available compounds were used without purification. *meso*-Tetraphenylporphyrin (TPP) was purchased from Porphyrin Systems GbR, Germany. Octabromo *meso*-tetraphenylporphyrin (TPPBr₈) was purchased from Frontier Scientific. *meso*-(Tetrapentafluorophenyl)porphyrin (TPFPFPP) was purchased from PorphyChem. The free porphyrins were metallated using known methods.^{21,43–50} Mn(III) *meso*-tetraphenylporphyrin chloride (Mn(TPP)Cl), Zn(II) *meso*-(tetrapentafluorophenyl)porphyrin (ZnTPFPFPP) and Sn(IV) *meso*-tetraphenylporphyrin dichloride (Sn(TPP)Cl₂) were purchased from Porphyrin Systems GbR, Germany. Ni(II) *meso*-tetraphenylporphyrin (Ni(TPP)) was purchased from Frontier Scientific. Analytical TLC was performed using Merck precoated silica gel 60 F₂₅₄ plates and for column chromatography Matrex silica gel (60 Å, 35–70 μm) or Merck basic aluminum oxide (60–200 mm) was used.

Computational details

The calculations were performed with the B3LYP functional as implemented in the Gaussian 09 program package.^{28,30} Geometries were optimized using the 6-31G(d,p) basis set for the H-, C-, O-, F-, Cl-atoms and SDD basis set and pseudo potential for all metals.^{30,31} Frequency calculations were performed at the same level to confirm that a minimum had been reached and to extract free energy corrections which were evaluated at 298.15 K. The energies were also corrected for dispersion effects according to the B3LYP-D3 method by Grimme.³² When optimizing the geometries of the MTPP and



MTPFPP structures different rotations of the phenyl rings where used as starting points in order to find the most stable geometry.

As metals can adopt several different spin states the three lowest spin states were investigated for each species.³⁵ For species with an even number of electrons the states with total spin $S = 0, 1, 2$ ($M = 1, 3, 5$) were calculated and species with an odd number of electrons the states with total spin $S = 1/2, 3/2, 5/2$ ($M = 2, 4, 6$) were calculated. A stability analysis was performed to ensure that a stable wave-function was attained for all species. As the goal of the computational study was to compare the most stable free metalloporphyrin and methanol-metalloporphyrin complex, the states which did not converge after repeated tries were deemed unstable and not relevant to this study.

Natural charges were calculated with Natural Population Analysis (NPA) using Natural Bond Orbital analysis (NBO) version 3 with the B3LYP functional and the 6-31G(d,p) basis set for the H-, C-, O-, F-, Cl-atoms and SDD basis set and pseudo potential for all metals.³⁶ For the MgPBr₈ and MgTPPBr₈ the LanL2DZ basis set was used for all metals.^{37,38}

UV-Vis spectra were recorded on a Varian Cary 3 Bio spectrophotometer using 5 mm quartz cuvettes.

¹H and ¹³C NMR spectra were recorded on Varian Mercury Plus (¹H at 300.03 MHz, ¹³C at 75.45 MHz), Agilent 400-MR DD2 (¹H at 399.98 MHz, ¹³C at 100.58 MHz), or Varian Unity Inova (¹H at 499.94 MHz, ¹³C at 125.7 MHz) spectrometers at 25 °C unless noted otherwise. Chemical shifts are reported referenced to tetramethylsilane *via* the residual solvent signal (CDCl₃, ¹H at 7.26 and ¹³C at 77 ppm).

Mass spectra were recorded on a Voyager-DE PRO MALDI-TOF spectrometer in positive or negative ion-mode without matrix.

For NMR titrations, aliquots of freshly prepared ligand solutions (CDCl₃, AlOx-filtered, dried over 3 Å molecular sieves) were added to a solution of the porphyrin in an NMR tube. All NMR spectra were recorded on a 500 MHz spectrometer with $d1 = 5$ s and $nt = 32$. The precise host/ligand ratios were calculated from the integrals at the estimated 1:1 ratio. Most ligands were titrated to $[G]/[H] = 9$ unless the titration curve had already levelled out or the ligand showed no binding at all (Fig. 1).

The binding constants (K_a) for the complexes were determined by NMR titration utilizing the iterative fitting program published by P. Thordarsson²⁵ in Matlab R2012b. Both 1:1 and 1:2 complexation models were tested. K_a is calculated by (eqn (3) and (4)). The standard error (SE_y) is estimated by (eqn (6)).

$$SE_y = \sqrt{\frac{\sum (y_{data} - y_{calc})^2}{N - k}} \quad (6)$$

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

We thank Professor Helena Grennberg for valuable insights into porphyrin synthesis and purification. Financial support by the Swedish Research Council (grant nr. 621-2012-3379) and by the Carl Tryggers Foundation (CTS 16:156) is gratefully acknowledged. The computations were performed on resources provided by the Swedish National Infrastructure for Computing (SNIC) at National Supercomputer Centre (NSC), Linköping University.

References

- 1 R. J. P. Williams, The properties of metalloporphyrins, *Chem. Rev.*, 1956, **56**, 299–328.
- 2 C. H. Devillers, A. K. D. Dimé, H. Cattey and D. Lucas, Crystallographic, spectroscopic and electrochemical characterization of pyridine adducts of magnesium(II) and zinc (II) porphine complexes, *C. R. Chim.*, 2013, **16**, 540–549.
- 3 X. Huang and J. T. Groves, Oxygen Activation and Radical Transformations in Heme Proteins and Metalloporphyrins, *Chem. Rev.*, 2018, **118**, 2491–2553.
- 4 I. Beletskaya, V. S. Tyurin, A. Y. Tsivadze, R. Guillard and C. Stern, Supramolecular Chemistry of Metalloporphyrins, *Chem. Rev.*, 2009, **109**, 1659–1713.
- 5 X. Huang and J. T. Groves, Oxygen Activation and Radical Transformations in Heme Proteins and Metalloporphyrins, *Chem. Rev.*, 2018, **118**, 2491–2553.
- 6 M. Jurow, A. E. Schuckman, J. D. Batteas and C. M. Drain, Porphyrins as molecular electronic components of functional devices, *Coord. Chem. Rev.*, 2010, **254**, 2297–2310.
- 7 L. P. Cook, G. Brewer and W. Wong-Ng, Structural Aspects of Porphyrins for Functional Materials Applications, *Crystals*, 2017, **7**, 223.
- 8 H. Huang, W. Song, J. Rieffeland; and J. F. Lovell, Emerging applications of porphyrins in photomedicine, *Front. Phys.*, 2015, **3**, 23.
- 9 A. G. Petrovic, G. Vantomme, Y. Negrón-Abril, E. Lubian, G. Saielli, I. Menegazzo, R. Cordero, G. Proni, K. Nakanishi, T. Carofiglio and N. Berova, Bulky Melamine-Based Zn-Porphyrin Tweezer as a CD Probe of Molecular Chirality, *Chirality*, 2011, **23**, 808–819.
- 10 X. Huang, N. Fujioka, G. Pescitelli, F. E. Koehn, R. T. Williamson, K. Nakanishi and N. Berova, Absolute Configurational Assignments of Secondary Amines by CD-Sensitive Dimeric Zinc Porphyrin Host, *J. Am. Chem. Soc.*, 2002, **124**, 10320–10335.
- 11 (a) A. Dhamija, B. Saha and S. P. Rath, Metal-Center-Driven Supramolecular Chirogenesis in Tweezer Amino Alcohol Complexes: Structural, Spectroscopic, and Theoretical Investigations, *Inorg. Chem.*, 2017, **56**, 15203–15215; (b) S. Brahma, S. A. Iqbal and S. P. Rath, Synthesis, Structure, and Properties of a Series of Chiral Tweezer-Diamine Complexes Consisting of an Achiral Zinc(II) Bisporphyrin Host and Chiral Diamine Guest: Induction



- and Rationalization of Supramolecular Chirality, *Inorg. Chem.*, 2014, **53**, 49–62; (c) S. Brahma, S. A. Iqbal, A. Dhamija and S. P. Rath, Highly Enhanced Bisignate Circular Dichroism of Ferrocene-Bridged Zn(II) Bisporphyrin Tweezer with Extended Chiral Substrates due to Well-Matched Host-Guest System, *Inorg. Chem.*, 2014, **53**, 2381–2395.
- 12 S. Norrehed, P. Polavarapu, W. Yang, A. Gogoll and H. Grennberg, Conformational Restriction of Flexible Molecules in Solution by a Semirigid Bis-Porphyrin Molecular Tweezer, *Tetrahedron*, 2013, **69**, 7131–7138.
 - 13 M. Blom, S. Norrehed, C.-H. Andersson, H. Huang, M. E. Light, J. Bergquist, H. Grennberg and A. Gogoll, Synthesis and Properties of Bis-Porphyrin Molecular Tweezers: Effects of Spacer Flexibility on Binding and Supramolecular Chirogenesis, *Molecules*, 2016, **21**, 16, DOI: 10.3390/molecules21010016.
 - 14 S. Norrehed, H. Johansson, H. Grennberg and A. Gogoll, Improved Stereochemical Analysis of Conformationally Flexible Diamines by Binding to a Bisporphyrin Molecular Clip, *Chem. – Eur. J.*, 2013, **19**, 14631–14638.
 - 15 (a) D. Dolphin, T. G. Traylor and L. Y. Xie, Polyhaloporphyrins: Unusual Ligands for Metals and Metal-Catalyzed Oxidations, *Acc. Chem. Res.*, 1997, **30**, 251–259; (b) E. K. Woller and S. G. DiMaggio, 2,3,7,8,12,13,17,18-Octafluoro-5,10,15,20-tetraarylporphyrins and Their Zinc Complexes: First Spectroscopic, Electrochemical, and Structural Characterization of a Perfluorinated Tetraarylmetalloporphyrin, *J. Org. Chem.*, 1997, **62**, 1588–1593; (c) X. Li, M. Tanasova, C. Vasileiou and B. Borhan, Fluorinated Porphyrin Tweezer: A Powerful Reporter of Absolute Configuration for erythro and threo Diols, Amino Alcohols, and Diamines, *J. Am. Chem. Soc.*, 2008, **130**, 1885–1893.
 - 16 K. Ermanis, K. E. B. Parkes, T. Agback and J. M. Goodman, Doubling the Power of DP4 for Computational Structure Elucidation, *Org. Biomol. Chem.*, 2017, **15**, 8998–9007.
 - 17 X. Li, C. E. Burrell, R. J. Staples and B. Borhan, Absolute Configuration for 1,n-Glycols: A Nonempirical Approach to Long-Range Stereochemical Determination, *J. Am. Chem. Soc.*, 2012, **134**, 9026–9029.
 - 18 G. P. Dechaine, Y. Maham, X. Tan and M. R. Gray, Regular solution theories are not appropriate for model compounds for petroleum asphaltenes, *Energy Fuels*, 2011, **25**, 737–746.
 - 19 K. P. Kepp, A Quantitative Scale of Oxophilicity and Thiophilicity, *Inorg. Chem.*, 2016, **55**, 9461–9470.
 - 20 O. I. Koifman and T. A. Ageeva, Metalloporphyrins in macromolecular chemistry, *Russ. Chem. Bull., Int. Ed.*, 2015, **64**, 2001–2011.
 - 21 J. S. Lindsey and J. N. Woodford, A Simple Method for Preparing Magnesium Porphyrins, *Inorg. Chem.*, 1995, **34**, 1063–1069.
 - 22 K. M. Kadish and L. A. Bottomley, Substituent effects on the formation constants of iron(III) and iron(II) tetraphenylporphyrin-pyridine complexes, *J. Am. Chem. Soc.*, 1977, **99**, 2380–2382.
 - 23 E. V. Goldammer, Interactions of some nitrogenous bases with low-spin ferriporphyrins, *J. Mol. Struct.*, 1980, **60**, 127–130.
 - 24 V. P. Andreev, P. S. Sobolev and V. A. Tafeenko, Coordination of Zinc Tetraphenylporphyrin with Pyridine Derivatives in Chloroform Solution and in the Solid Phase, *Russ. J. Gen. Chem.*, 2017, **87**, 1070–3632.
 - 25 P. Thordarsson, Determining association constants from titration experiments in supramolecular chemistry, *Chem. Soc. Rev.*, 2011, **40**, 1305–1323.
 - 26 M. Sankaralingam, Y.-M. Lee, S. Hyun Jeon, M. Sook Seo, K.-B. Cho and W. Nam, A mononuclear manganese(III)–hydroperoxo complex: synthesis by activating dioxygen and reactivity in electrophilic and nucleophilic reactions, *Chem. Commun.*, 2018, **54**, 1209–1212.
 - 27 T. K. Chandrashekar and V. Krishnan, Coordinative Interaction of Morpholine With Divalent Metallo Tetraphenyl Porphyrins, *J. Inorg. Nucl. Chem.*, 1981, **43**, 328–3290.
 - 28 A. D. Becke, Density-Functional Thermochemistry. III. The Role of Exact Exchange, *J. Chem. Phys.*, 1993, **98**, 5648–5652.
 - 29 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, Ö. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski and D. J. Fox, *Gaussian 09, Revision D.01*, Gaussian, Inc., Wallingford CT, 2009.
 - 30 P. C. Hariharan and J. A. Pople, The Influence of Polarization Functions on Molecular Orbital Hydrogenation Energies, *Theor. Chim. Acta*, 1973, **28**, 213–222.
 - 31 D. Andrae, U. Häußermann, M. Dolg, H. Stoll and H. Preuß, Energy-Adjusted ab Initio Pseudopotentials for the Second and Third Row Transition Elements, *Theor. Chim. Acta*, 1990, **77**, 123–141.
 - 32 S. Grimme, Semiempirical GGA-Type Density Functional Constructed with a Long-Range Dispersion Correction, *J. Comput. Chem.*, 2006, **27**, 1787–1799.
 - 33 W. Wei, F.-Q. Bai, B.-H. Xia, J. Wang and H.-X. Zhang, A density functional theory investigation of the stability, aromaticity, and photophysical behavior for the highly conjugated macrocycles containing 4 pyrroles, *J. Phys. Org. Chem.*, 2017, **30**, e3617.



- 34 Z. Cao, Q. Chen, Y. Lu, H. Liu and Y. Hu, Density functional theory study on the interaction between metalloporphyrins and NH₃, *Int. J. Quantum Chem.*, 2013, **113**, 1137–1146.
- 35 A. Ghosh, Transition metal spin state energetics and non-innocent systems: Challenges for DFT in the bioinorganic arena, *J. Biol. Inorg. Chem.*, 2006, **11**, 712–724.
- 36 A. E. Reed, R. B. Weinstock and F. Weinhold, Natural-population analysis, *J. Chem. Phys.*, 1985, **83**, 735–746.
- 37 T. H. Dunning Jr. and P. J. Hay, in *Modern Theoretical Chemistry*, ed. H. F. Schaefer III, Plenum, New York, 1977, vol. 3, pp. 1–28.
- 38 P. J. Hay and W. R. Wadt, Ab initio effective core potentials for molecular calculations – potentials for the transition-metal atoms Sc to Hg, *J. Chem. Phys.*, 1985, **82**, 270–283.
- 39 S. Cho, Y. Lee, H. S. Han, H. K. Lee and S. Jeon, Characteristic electronic perturbation by asymmetric arrangements of p -aminophenyl substituents in free-base porphyrins, *J. Phys. Chem. A*, 2014, **118**, 4995–5001.
- 40 G. A. Spyroulias, A. Despotopoulos, C. P. Raptopoulou, A. Terzis and A. G. Coutsolelos, First X-ray crystal structure of a lanthanide(III) monoporphyrinic complex using non-planar porphyrin rings, *Chem. Commun.*, 1997, 783–784.
- 41 C. C. de Melo, W. da, C. Moreira, T. J. Martins, M. R. Cordeiro, J. Ellena, F. F. Guimarães and F. T. Martins, Saddle-shaped macrocycle distortion and symmetry decrease in cobalt(II) meso-tetraphenylporphyrin: Structure of a dichloromethane solvate and DFT calculations, *J. Mol. Struct.*, 2014, **1076**, 468–474.
- 42 A. J. Cohen, P. Mori-Sánchez and W. Yang, Insights into current limitations of density functional theory, *Science*, 2008, **321**(5890), 792–794.
- 43 P. Chen, M. H. Chisholm, J. C. Gallucci, X. Zhang and Z. Zhou, Binding of Propylene Oxide to Porphyrin- and Salen-M(III) Cations, Where M=Al, Ga, Cr, and Co, *Inorg. Chem.*, 2005, **44**, 2588–2595.
- 44 C. E. Anderson, S. I. Vagin, W. Xia, H. Jin and B. Rieger, Cobaltoporphyrin-Catalyzed CO₂/Epoxide Copolymerization: Selectivity Control by Molecular Design, *Macromolecules*, 2012, **45**, 6840–6849.
- 45 Y. Qin, X. Wang, S. Zhang, X. Zhao and F. Wang, Fixation of Carbon Dioxide into Aliphatic Polycarbonate, Cobalt Porphyrin Catalyzed Regio-Specific Poly(propylene carbonate) with High Molecular Weight, *J. Polym. Sci., Part A: Polym. Chem.*, 2008, **46**, 5959–5967.
- 46 D. F. O'Shea, M. A. Miller, H. Matsueda and J. S. Lindsey, Investigation of the Scope of Heterogeneous and Homogeneous Procedures for Preparing Magnesium Chelates of Porphyrins, Hydroporphyrins, and Phthalocyanines, *Inorg. Chem.*, 1996, **35**, 7325–7338.
- 47 V. F. Slagt, P. W. N. M. van Leeuwen and J. N. H. Reek, Supramolecular bidentate phosphorus ligands based on bis-zinc(II) and bis-tin(IV) porphyrin building blocks, *Dalton Trans.*, 2007, 2302–2310.
- 48 A. M. Stolzenberg and G. S. Haymond, Activation Barriers to meso-Aryl Group Rotation in Titanyl Tetraaryltetrapyrroles. An Investigation of the Out-of-Plane Flexibility of Hydroporphyrins, *Inorg. Chem.*, 2002, **41**, 300–308.
- 49 S. A. Vail, D. I. Schuster, D. M. Guldi, M. Isosomppi, N. Tkachenko, H. Lemmetyinen, A. Palkar, L. Echegoyen, X. Chen and J. Z. H. Zhang, Energy and Electron Transfer in β -Alkynyl-Linked Porphyrin-[60]Fullerene Dyads, *J. Phys. Chem. B*, 2006, **110**, 14155–14166.
- 50 R. Wang, Z. Hailin, B. Gao, R. Wang, Á. H. Zhu and Á. B. Gao, The catalytic activity of poly(N-vinylimidazole)/SiO₂-supported metalloporphyrins in ethyl benzene oxidation, *React. Kinet., Mech. Catal.*, 2011, **103**, 431–441.

