


 Cite this: *CrystEngComm*, 2018, 20, 385

Salts with titanyl and vanadyl phthalocyanine radical anions. Molecular design and effect of cations on the structure and magnetic and optical properties†

 Dmitri V. Konarev,^a Alexey V. Kuzmin,^b Salavat S. Khasanov,^b Mikhail S. Batov,^{ac} Akihiro Otsuka,^{de} Hideki Yamochi,^{de} Hiroshi Kitagawa^d and Rimma N. Lyubovskaya^a

Radical anion salts of metal phthalocyanines have been providing a new family of compounds which can potentially possess promising optical, magnetic and conducting properties. Essential progress in this field has been reached only recently when series of salts of phthalocyanines with different metals and axial substituents were obtained. In this work we summarize our recent results about the radical anion salts of titanyl (Ti^{IV}OPc) and vanadyl (V^{IV}OPc) phthalocyanines and nine new salts of these phthalocyanines are discussed. Different synthetic approaches for the preparation of these salts are presented, and the effect of the size and shape of counter cations on the crystal structures of the salts and consequently their magnetic and optical properties are considered. The {Ti^{IV}O(Pc³⁻)⁻} radical anions have an $S = 1/2$ spin on the Pc³⁻ macrocycle, whereas the {V^{IV}O(Pc³⁻)⁻} radical anions have two $S = 1/2$ spins on both the Pc³⁻ macrocycle and the central V^{IV} atom. Therefore, it is possible to compare the magnetic properties of isostructural compounds with different magnetic states of metal phthalocyanines. We also discuss the effect of reduction on the molecular structure and the optical and magnetic properties of negatively charged titanyl and vanadyl phthalocyanines.

 Received 5th November 2017,
Accepted 13th December 2017

DOI: 10.1039/c7ce01918a

rsc.li/crystengcomm

Introduction

Metal-free and metal-containing phthalocyanines can possess promising optical, conducting and magnetic properties.^{1–9} As a result, some phthalocyanine derivatives are used as sensors and as materials for optical, photoelectronic and photovoltaic devices.^{1,2} To obtain conducting compounds phthalocyanines should be partially oxidized to introduce an unpaired electron into the phthalocyanine macrocycle. A columnar or layered arrangement of such macrocycles together with effective π – π stacking provides semiconductivity or quasi-one-dimensional

metallic conductivity.^{3–7} The paramagnetic atoms in the center of the phthalocyanine macrocycle can induce a giant magnetoresistance.^{5–7} Molecular magnets containing manganese(III) phthalocyanine or substituted naphthalocyanine cations and tetracyanoethylene radical anions are also known.^{8,9} All these compounds were obtained by oxidation of metal phthalocyanines.

Potentially, reduced metal-free and metal-containing phthalocyanines can also exhibit interesting magnetic or conducting properties. Theoretical calculations showed the possibility of metallic conductivity or even superconductivity at electron doping of non-transition metal phthalocyanines.¹⁰ Compounds with effective magnetic coupling of spins were obtained. For example, ferrimagnetic ordering of spins below 4.5 K was found in (Cp*₂Cr⁺)[Fe^I(Pc²⁻)⁻] \cdot 4C₆H₄Cl₂ which contains chains of alternating Fe^I ($S = 1/2$) and Cr^{III} ($S = 3/2$) spins.¹¹ Magnetic interactions are manifested between the Pc³⁻ macrocycles in the case of their effective π – π interactions in dimers or layers.^{12,13} Reduction of metal phthalocyanines is realized at –0.4––0.8 V vs. SCE in solutions.^{14,15} Therefore, all manipulations of metal phthalocyanine radical anions require anaerobic conditions. Probably, for this reason this field has been developed only recently and the

^a Institute of Problems of Chemical Physics RAS, Chernogolovka, Moscow region, 142432 Russia. E-mail: konarev3@yandex.ru

^b Institute of Solid State Physics RAS, Chernogolovka, Moscow region, 142432 Russia

^c Moscow State University, Leninskie Gory, 119991 Moscow, Russia

^d Division of Chemistry, Graduate School of Science, Kyoto University, Sakyo-ku, Kyoto 606-8502, Japan

^e Research Center for Low Temperature and Materials Sciences, Kyoto University, Sakyo-ku, Kyoto 606-8501, Japan

† Electronic supplementary information (ESI) available: Data of magnetic measurements for salts 4–7. CCDC 1581401–1581403, 1583980, 1581407, 1581408, 1581410, 1581414 and 1581417. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c7ce01918a



number of reduced metal phthalocyanines obtained till now as crystals is still rather limited.^{16–28} Among them there are salts with cobalt(i) and iron(i) phthalocyanine anions^{16–20} as well as iron(i) hexadecachlorophthalocyanine anions.^{21,22} Their formation is accompanied by the reduction of metal atoms in the phthalocyanines. A series of salts with the [MPC][−] radical anions (M = H₂, Cu^{II}, Ni^{II}, Sn^{II}, Pb^{II}, Ti^{IV}O, V^{IV}O, Sn^{IV}Cl₂, Mg^{II})^{12,23–30} were also obtained. In this case the reduction is centered on the Pc macrocycle and is accompanied by the formation of paramagnetic Pc^{•−} radical trianions. Several neutral metal phthalocyanine compounds are known which contain phthalocyanine radical trianions and even tetraanions: Al^{III}(Pc^{•−}),^{31,32} Ga^{III}(Pc^{•−}),³² Ge^{IV}(Pc^{•−}),³³ and Sn^{IV}Ph(Pc^{•−}).³⁴ In these compounds, the negative charge on Pc^{•−} or Pc^{4−} is compensated for by the positive +3 or +4 charge on the central metal atom.

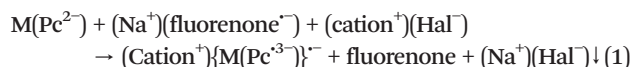
Reduced metal phthalocyanines can show strong magnetic coupling of spins especially when they have only one axial ligand assisting the formation of π -stacking dimers in the salts.^{12,35} In this work we summarize our results on the preparation and study of radical anion salts of titanyl (Ti^{IV}OPc) and vanadyl (V^{IV}OPc) phthalocyanines^{12,27,35} and present new salts of these phthalocyanines (Table 1). Different synthetic approaches for the preparation of these salts are presented. The effect of reduction on the molecular structure and properties of titanyl and vanadyl phthalocyanines is also discussed. In this work we discuss how the crystal structures of the salts can be affected by using cations of different size and shape. We found that it is possible to affect the overlapping degree of the Pc^{•−} macrocycles in the dimers thereby affecting the magnetic and optical properties of the salts. Based on these data we show how it is possible to control the properties of the compounds by methods of molecular design.

Results and discussion

1. Synthesis

Metal phthalocyanines can be reduced by sodium fluorenone ketyl which can be prepared by direct interaction of fluorenone

with metallic sodium in toluene.³⁶ The first oxidation potential of this reductant of −1.3 V vs. Ag/AgCl (or −1.345 vs. SCE) in CH₂Cl₂ (ref. 37) allows the generation of radical anions of nearly all metal phthalocyanines.^{14,15} The sodium metal cation of the reductant is substituted by different organic cations at the reduction and these cations become the counterions of reduced metal phthalocyanine (see reaction 1). The reaction is carried out in pure *o*-dichlorobenzene (or in some cases in the presence of benzonitrile). Sodium halides are insoluble in these solvents; they precipitate from the solution to yield pure phthalocyanine salt.



in *o*-dichlorobenzene, cation⁺ stands for Et₄N⁺, Bu₄N⁺ and others (Table 1).

It is important to use solvents which cannot solvate alkali metal cations, allowing their precipitation from the solution as chloride, bromide or iodide salts. The presence of cations of two types in the synthesis such as solvated alkali metal and organic cations generally does not allow the preparation of good quality single crystals. We used solvents purified of traces of oxygen and water by distillation over CaH₂ or sodium benzophenone under argon and then degassing prior to use. All manipulations for the synthesis of the salts were carried out in a glove box with a controlled atmosphere and content of water and oxygen less than 1 ppm. It should be noted that although pristine metal phthalocyanines are very sparingly soluble in organic solvents, their radical anion salts especially with large organic cations are highly soluble in *o*-dichlorobenzene. Slow mixing of their solutions with *n*-hexane in the diffusion tube for 1–2 months allowed the preparation of good quality single crystals which precipitated on the walls of the tube. Crystals were studied by X-ray diffraction. After thorough separation of the crystals from any admixtures under a microscope and determination of possible phases according to the unit cell parameters, we analyzed their optical and magnetic properties. The latter properties were studied for polycrystalline samples sealed under an

Table 1 Composition of the discussed salts of titanyl and vanadyl phthalocyanines

Cation	Neutral component	Phthalocyanine	Anion	Solvent	Ti ^{IV} OPc	V ^{IV} OPc
(Et ₄ N ⁺)		{M ^{IV} O(Pc ^{•−}) [−] }		C ₆ H ₄ Cl ₂	1 (ref. 12)	—
(Bu ₄ N ⁺)		{M ^{IV} O(Pc ^{•−}) [−] }			2 (ref. 12)	3 (ref. 12)
(Pent ₄ N ⁺)		{M ^{IV} O(Pc ^{•−}) [−] }			4 (this work)	5 (this work)
(Hex ₄ N ⁺)		{M ^{IV} O(Pc ^{•−}) [−] }		C ₆ H ₄ Cl ₂	6 (this work)	7 (this work)
(MDABCO ⁺)		{M ^{IV} O(Pc ^{•−}) [−] }			8 (this work)	
(<i>i</i> -Pr ₂ Im ⁺)		{M ^{IV} O(Pc ^{•−}) [−] }		0.75C ₆ H ₄ Cl ₂		9 (this work)
(Me ₄ P ⁺)	(TPC) _{0.5}	{M ^{IV} O(Pc ^{•−}) [−] }		C ₆ H ₄ Cl ₂	10 (ref. 35)	11 (ref. 35)
(MDABCO ⁺)	(TPC)	{M ^{IV} O(Pc ^{•−}) [−] }			12 (this work)	
(MDABCO ⁺) ₂		{M ^{IV} O(Pc ^{•−}) [−] }	(Γ)		13 (this work)	14 (this work)
{Cryptand(Na ⁺)}		{M ^{IV} O(Pc ^{•−}) [−] }		C ₆ H ₄ Cl ₂		15 (ref. 27)

Et₄N⁺ – tetraethylammonium; Bu₄N⁺ – tetrabutylammonium; Pent₄N⁺ – tetrapentylammonium; Hex₄N⁺ – tetrahexylammonium; MDABCO⁺ – N-methyl diazabicyclooctanium; *i*-Pr₂Im⁺ – 1,3-di(iso-propyl)imidazolium; Me₄P⁺ – tetramethylphosphonium; TPC – triptycene; cryptand: 4,7,13,16,21,24-hexaoxa-1,10-diazabicyclo[8.8.8]hexacosane; C₆H₄Cl₂ – *o*-dichlorobenzene.



ambient pressure of argon or under a 10^{-5} torr vacuum in quartz tubes of 2 mm diameter by EPR and SQUID techniques.

There are different synthetic approaches for the preparation of the salts with small organic cations as illustrated by the synthesis of metal phthalocyanine salts with *N*-methyl diazabicyclooctinium cations (MDABCO⁺). Reduction of titanyl phthalocyanine in the presence of exactly one molar equivalent of (MDABCO⁺)(I⁻) (with a slight excess of sodium fluorenone ketyl as a reductant) followed by precipitation of the salt by *n*-hexane yields the crystals of (MDABCO⁺){Ti^{IV}O(Pc³⁻)}⁻ (8). Since MDABCO⁺ is a small cation, addition of about 20% of benzonitrile is needed from the total volume of *o*-dichlorobenzene. Similarly, benzonitrile was added for the preparation of the salt with Et₄N⁺ (ref. 12) and other salts containing MDABCO⁺.

For the synthesis of the radical anion salts it is possible to use a 2–3-fold molar excess of (MDABCO⁺)(I⁻) relative to metal phthalocyanine. The (MDABCO⁺)₂{M^{IV}O(Pc³⁻)}⁻(I⁻) salts (M =

Ti for 13 and M = V for 14) containing two MDABCO⁺ cations and one I⁻ anion per metal phthalocyanine were obtained by such a procedure. If *o*-dichlorobenzene molecules are involved in the salts, they contain solvent molecules instead of the second cation as in (Et₄N⁺){Ti^{IV}O(Pc³⁻)}⁻·C₆H₄Cl₂ (1) and (*i*-Pr₂Im⁺){V^{IV}O(Pc³⁻)}⁻·0.75C₆H₄Cl₂ (13). In this case it is not possible to prepare salts with two cations. It should be noted that titanyl and vanadyl phthalocyanines form only 1:1 salts with a large Bu₄N⁺ cation even with an excess of (Bu₄N⁺)(Br⁻) in the reaction mixture despite of planar phthalocyanines (H₂Pc, Cu^{II}Pc, Ni^{II}Pc) form salts with two Bu₄N⁺ cations and one Br⁻ anions.¹²

In the studies on fullerene salts, the authors derived the concept of multi-component approach in which a neutral component is added along with the source of the counterion to define the crystal packing at the salt.⁴¹ This method is successfully applied also in the anionic phthalocyanine chemistry. The aromatic hydrocarbon triptycene (TPC) can form layers with hexagonal vacancies which can accommodate



Fig. 1 Molecular structures of the components in salts 1–14. Carbon atoms are reported in brown, nitrogen atoms in blue, oxygen atoms in red, metal atoms in grey, iodine atoms in violet, chlorine atoms in green, and phosphorus atom in orange.



small cations like MDABCO⁺ or Me₄P⁺ to form the salt (MDABCO⁺)(TPC){Ti^{IV}O(Pc³⁻)⁻} (12) or (Me₄P⁺)(TPC)_{0.5}{M^{IV}O(Pc³⁻)⁻}·C₆H₄Cl₂ (M = Ti (10) and V (11)). In this case metal phthalocyanines are also packed in layers alternating with the cationic layers containing neutral TPC molecules. As a result, this method is suitable for the preparation of layered structures. Only small cations can be used together with TPC to form multi-component complexes. Even the Et₄N⁺ cations are too large to be incorporated into the TPC layers and synthesis of the salt with Et₄N⁺ and TPC yields (Et₄N⁺){M^{IV}O(Pc³⁻)⁻}·C₆H₄Cl₂ (1) only.

Components that cocrystallized with titanyl and vanadyl phthalocyanine radical anions in salts 1–14 are shown in Fig. 1. It is seen that the size of the cations varies from small (MDABCO⁺ and Me₄P⁺) to large (Hex₄N⁺ and cryptand[2,2,2]-Na⁺) (the latter cation is not shown in Fig. 1). The variation of the cations size allows one to study how the size of cations affects the crystal structure of the salts and packing of the macrocycles in the π -stacking [$\{M^{IV}O(Pc^{3-})\}^-$]₂ dimers.

2. Crystal structures

The structure formed by titanyl phthalocyanine radical anions in (Et₄N⁺){Ti^{IV}O(Pc³⁻)⁻}·C₆H₄Cl₂ (1)¹² is shown in Fig. 2. It is a layered structure in which phthalocyanine layers alternate with the Et₄N⁺·C₆H₄Cl₂ layers. Both the Et₄N⁺ cations and solvent C₆H₄Cl₂ molecules have a small size and a total of four such units are positioned over each phthalocyanine plane (Fig. 2b). The presence of only one axial substituent in {Ti^{IV}O(Pc³⁻)⁻} allows the formation of π - π stacking [$\{Ti^{IV}O(Pc^{3-})\}^-$]₂ dimers within the phthalocyanine layers. The Pc planes are arranged in these dimers very close to each other with an interplanar distance of only 3.13 Å, and 18 short van der Waals C...C,N contacts are formed between

the phthalocyanines. There are several side-by-side C...C contacts in the layers between phthalocyanines belonging to neighboring dimers. These contacts are shown by green dashed lines in Fig. 2a. Thus, weak interactions between the dimers having a square arrangement in the layers are observed. The intermolecular overlap integral for the singly occupied molecular orbitals (SOMOs) of {Ti^{IV}O(Pc³⁻)⁻} in 1 was calculated by the extended Hückel method.¹² The SOMO-SOMO overlap integral of 14.3×10^{-3} is large to show essential π - π interactions between the Pc³⁻ macrocycles in 1. This is due to the small shift of phthalocyanines relative to each other in the dimers and a short interplanar distance between them. Indeed, macrocycles are positioned in such a way that the titanium atom of one phthalocyanine is positioned over the pyrrole nitrogen atom of the neighboring phthalocyanine (the shift of these macrocycles is 2.005 Å).

Crystal structures of multi-component salts (Me₄P⁺)(TPC)_{0.5}{M^{IV}O(Pc³⁻)⁻}·C₆H₄Cl₂ (M = Ti (10) and V (11)) were described previously.³⁵ They also contain layers formed by π -stacking [$\{M^{IV}O(Pc^{3-})\}^-$]₂ dimers alternating with Me₄P⁺·TPC·C₆H₄Cl₂ layers along the *c* axis. The shift of the Pc³⁻ macrocycles in the dimers of 10 and 11 is larger than that in 1, providing essentially smaller SOMO-SOMO overlap integrals of 2.2×10^{-3} and 1.7×10^{-3} , respectively.³⁵

In previously described¹² crystal structures of 2 and 3, the size of the Bu₄N⁺ cations is essentially larger than that of Et₄N⁺. As a result, only two Bu₄N⁺ cations without solvent molecules are arranged near each Pc³⁻ macrocycle, forming cationic layers (Fig. 3b). In this case, phthalocyanines are also packed in layers consisting of π -stacking [$\{M^{IV}O(Pc^{3-})\}^-$]₂ dimers. However, the increase in cation size provides an essential shift (5.12 Å) of two Pc³⁻ macrocycles relative to each other in the dimers. This effect can be explained by the fact that the increase in size of the cationic layer proportionally increases the size of the

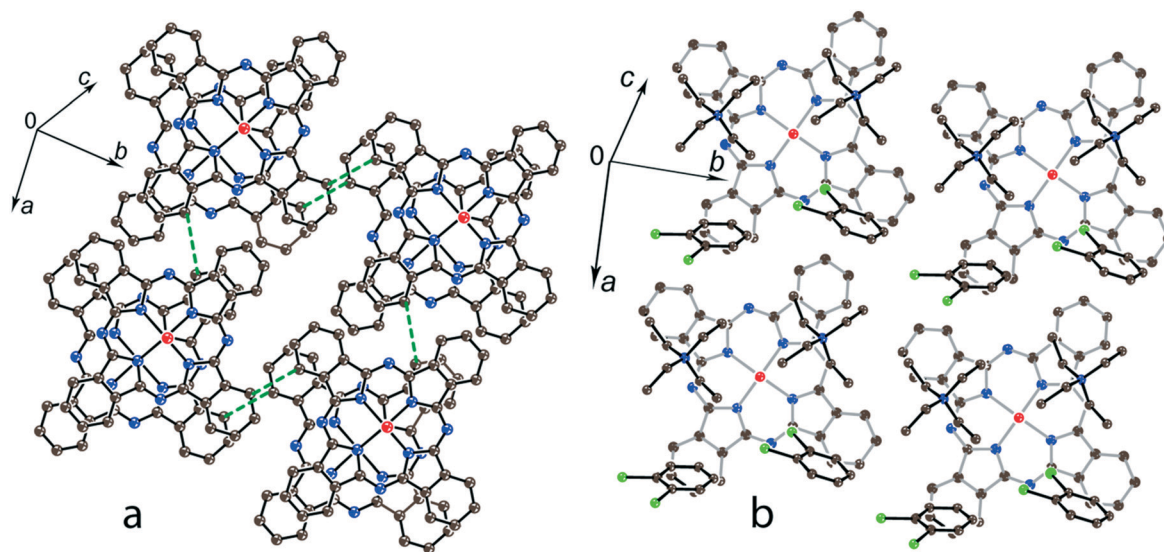


Fig. 2 (a) View of the phthalocyanine layer composed of the π -stacking [$\{Ti^{IV}O(Pc^{3-})\}^-$]₂ dimers in 1.¹² Short van der Waals C...C,N contacts are shown by green dashed lines. (b) Projection of the Et₄N⁺·C₆H₄Cl₂ layer on the phthalocyanine layer. Only one phthalocyanine from the dimer is shown (bonds are shown in grey).



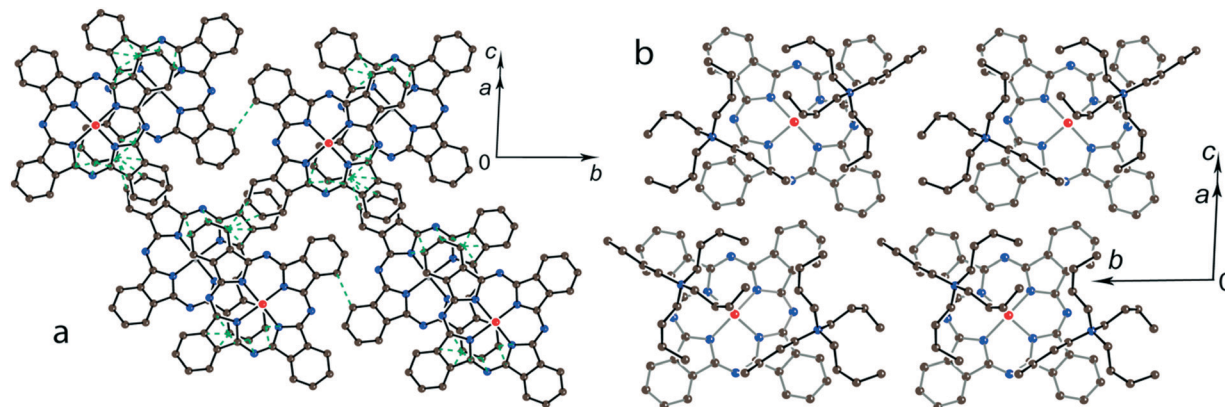


Fig. 3 (a) View of the phthalocyanine layer composed of the π -stacking $[(\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-}))_2]_2$ dimers in **2**.¹² Short van der Waals C,N...C,N contacts are shown by green dashed lines. (b) Projection of the Bu_4N^+ layer on the phthalocyanine layer. Only one phthalocyanine from the dimer is shown (bonds are shown in grey).

phthalocyanine layers. As a result, phthalocyanine macrocycles should be shifted relative to each other in the dimers to fit a larger area of the layers. The interplanar distance in the dimers increases up to 3.19 Å for both **2** and **3**. The increase in cation size essentially decreases the π - π overlapping between the Pc^{3-} macrocycles since the intermolecular SOMO-SOMO overlap integrals are only 0.1×10^{-3} and 0.3×10^{-3} for **2** and **3**, respectively.¹² Therefore, a several orders of magnitude decrease of the overlap integrals is observed for **2** and **3** compared with **1**.

The structures of **4**–**7** have not been published previously. The Pent_4N^+ cations form layers parallel to the bc plane in the isostructural $(\text{Pent}_4\text{N}^+)\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ (**4**) and $(\text{Pent}_4\text{N}^+)\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ (**5**) salts. Two large Pent_4N^+ cations are arranged near each phthalocyanine plane, and the π -stacking $[(\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-}))_2]_2$ dimers are also packed in layers (Fig. 4a). The extension of the alkyl substituents of the cations by one CH_2 unit affects the packing of the macrocycles in the dimers. The interplanar distance in the dimers of **4** and **5** is increased up to 3.22–3.24 Å. The Pc^{3-} macrocycles are shifted by 5.61 Å relative to each other in **4** and **5**, providing only slight overlapping of the Pc^{3-} macrocycles. These data show

that the use of the Pent_4N^+ cations instead of Bu_4N^+ impairs the conditions for π - π interactions between Pc^{3-} in the dimers. The size of the Pent_4N^+ cations can be considered as a boundary for the structures of metal phthalocyanines with one axial ligand at which the formation of π -stacking $[(\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-}))_2]_2$ dimers is still possible.

The study of the structures of $(\text{Hex}_4\text{N}^+)\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- \cdot \text{C}_6\text{H}_4\text{Cl}_2$ (**6**) and $(\text{Hex}_4\text{N}^+)\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- \cdot \text{C}_6\text{H}_4\text{Cl}_2$ (**7**) shows that the increase in the cation size by one more CH_2 unit does not allow the formation of π -stacking $[(\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-}))_2]_2$ dimers. Instead, the $\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ radical anions form one-dimensional chains with one side-by-side C...C contact (Fig. 5). These chains are completely isolated in a crystal by long alkyl substituents of the Hex_4N^+ cations.

Further increase of cation size does not allow the formation of one-dimensional phthalocyanine chains. Crystals of the $\{\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ salts with tetraoctylammonium cations were not obtained. However, our previous study shows that this cation assists in the formation of structures with a perpendicular arrangement of the Pc planes and large channels occupied by cations and solvent molecules.²³ The structure of

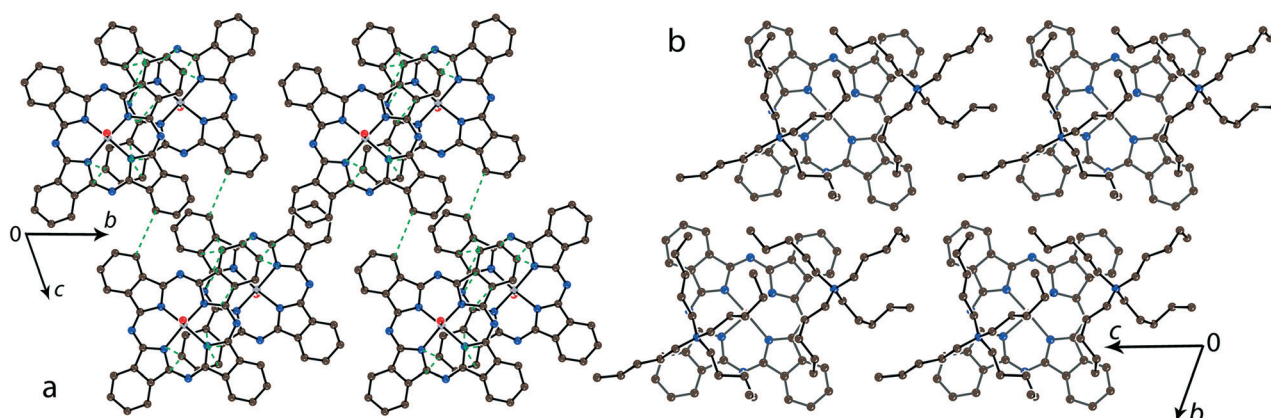


Fig. 4 (a) View of the phthalocyanine layer composed of the π -stacking $[(\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-}))_2]_2$ dimers in **4**. Short van der Waals C,N...C,N contacts are shown by green dashed lines. (b) Projection of the Pent_4N^+ layer on the phthalocyanine layer. Only one phthalocyanine from the dimer is shown (bonds are shown in grey).



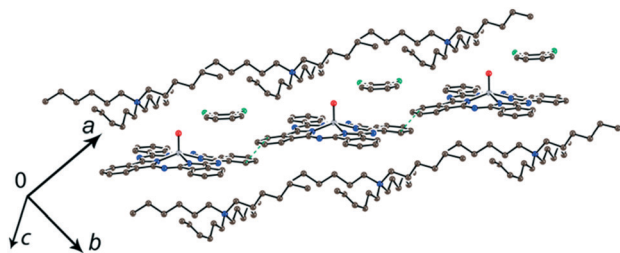


Fig. 5 View of the chains from $\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ in $(\text{Hex}_4\text{N}^+)\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- \cdot \text{C}_6\text{H}_4\text{Cl}_2$ (6).

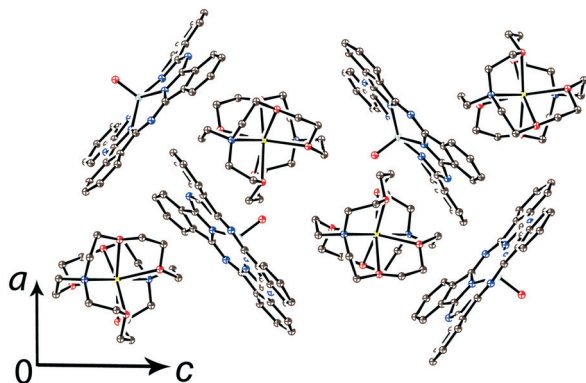


Fig. 6 View of the crystal structure of $(\text{cryptand}(\text{Na}^+)\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- \cdot \text{C}_6\text{H}_4\text{Cl}_2$ (15) along the b axis and channels occupied by solvent $\text{C}_6\text{H}_4\text{Cl}_2$ molecules (not shown for clarity) and the cryptand(Na^+) cations.²⁷

$\{\text{cryptand}[2,2,2](\text{Na}^+)\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- \cdot \text{C}_6\text{H}_4\text{Cl}_2$ (15) with large nearly spherical cryptand[2,2,2](Na^+) cations also shows a perpendicular arrangement of the Pc^{3-} macrocycles with channels occupied by cations and solvent molecules (Fig. 6).²⁷ Any π - π interactions between the Pc planes are absent in such structures in spite of the presence of several side-by-side $\text{C}\cdots\text{C}$ contacts between Pc^{3-} . Therefore, large cations do not yield salts with noticeable π - π interactions between Pc^{3-} . Layered structures with π -stacking $\{[\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-})]_2\}^-$ dimers are formed by using cations of relatively small or moderate size (smaller than Hex_4N^+). It is seen that the size of the cations defines the shift of the Pc^{3-} macrocycles in the dimers affecting the overlap integrals between SOMO orbitals of $\{\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ and, correspondingly, the physical properties of the salts.

3. Magnetic properties

Titanyl and vanadyl phthalocyanine radical anions have different magnetic states. Titanyl phthalocyanine contains diamagnetic Ti^{IV} , and the magnetic properties of $\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ are defined by the $S = 1/2$ spin delocalized over Pc^{3-} . This system can be considered as an organic π -radical having an unpaired $S = 1/2$ spin only on the π -orbitals of the macrocycle. Vanadyl phthalocyanine radical anions, $\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$, have two $S = 1/2$ spins per radical anion since one $S = 1/2$ spin is delocalized over Pc^{3-} and the other $S = 1/2$ spin is localized mainly on the V^{IV} d-orbitals. Therefore, it is possible

to study how the presence of paramagnetic metal atoms in the center of the macrocycle affects the magnetic properties of isostructural salts. Previously it was shown that paramagnetic metal centers like Fe^{III} in some cases affect the conductivity of the $\{\text{M}^{\text{III}}(\text{CN})_2(\text{Pc})\}^-$ salts with partially oxidized Pc macrocycles due to effective interactions of d-electrons of Fe^{III} with conducting π -electrons of the macrocycles. This provides an anisotropic magnetoresistance according to the direction of magnetic field.⁵⁻⁷

Data of magnetic measurements for the discussed salts are listed in Table 2. For all the salts with titanyl phthalocyanine radical anions, effective magnetic moments are in the 1.71–1.80 μ_{B} range at 300 K (Table 2). A theoretically calculated value for the system with one non-interacting $S = 1/2$ spin is equal to 1.73 μ_{B} and the observed magnetic moments of the salts are close to this value.

For salt 1 with the smallest Et_4N^+ cations the decrease of molar magnetic susceptibility is observed below 300 K and dimers are transferred to the diamagnetic singlet state below 150 K (ref. 12) (Fig. 7a). Such a behavior can be explained by triplet-singlet transitions in the dimers at very strong antiferromagnetic coupling between spins.¹² However, when we try to fit the experimental data by the Heisenberg model for pairs of isolated antiferromagnetically interacting spins,⁴⁰ this model does not describe well the magnetic behavior of 1. At the same time the observed magnetic behavior can be fitted well by the Heisenberg model for quasi-two-dimensional square layers³⁸ with a very large exchange interaction of $J/k_{\text{B}} = -255$ K (Fig. 7a). Since the $\{[\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})]_2\}^-$ dimers have a square arrangement in 1 and there is a weak overlapping between Pc^{3-} macrocycles of the neighboring dimers (Fig. 2), we suppose the antiferromagnetic ordering of spins within phthalocyanine layers to be possible. Taking into account that magnetic coupling is realized only between π -electrons of Pc^{3-} , the observed exchange interaction is very large for such a type of interactions.

Introduction of neutral TPC molecules into the cationic Me_4P^+ layers in $(\text{Me}_4\text{P}^+)(\text{TPC})_{0.5}\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- \cdot \text{C}_6\text{H}_4\text{Cl}_2$ (10) increases their size and additionally shifts the Pc^{3-} macrocycles in the dimers.³⁵ The magnetic behavior of 10 is described well by the Heisenberg model for pairs of isolated antiferromagnetically coupled spins with an exchange interaction of -123 K. This interaction is still rather large, providing transition of the dimers to the diamagnetic state below $T < 130$ K, but it is nearly two times smaller than that in 1. The use of the model for pairs is supported by the isolation of dimers in 10 since SOMO-SOMO overlap integrals with the neighboring dimers within the layers are smaller than 0.1×10^{-3} .³⁵

Large Bu_4N^+ cations in $(\text{Bu}_4\text{N}^+)\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ (2) provide an even stronger shift of the Pc^{3-} macrocycles in the dimers (Fig. 3) and a small SOMO-SOMO overlap integral showing weak π - π interactions in 2. As a result, no transition of the dimers to the diamagnetic singlet state is observed and spins are only antiferromagnetically coupled but with a rather large Weiss temperature of -57 K.¹²



Table 2 Magnetic properties of selected salts (exchange interaction (J/k_B), Weiss temperature (θ), effective magnetic moment at 300 K (μ_{eff}) and calculated SOMO-SOMO overlap integrals between the macrocycles in the π -stacking [$\{\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^{\cdot-}$] dimers

Cation	Ti ^{IV} OPc	V ^{IV} OPc
Et ₄ N ⁺	1, $J/k_B = -255$ K, —, $\mu_{\text{eff.}} = 1.72 \mu_B$, 14.3×10^{-3} , dimers ¹²	—
Me ₄ P ⁺ (TPC) _{0.5}	10, $J/k_B = -123$ K, $\theta = -131$ K, $\mu_{\text{eff.}} = 1.71 \mu_B$, 2.2×10^{-3} , dimers ³⁵	11, $J/k_{B,\text{inter}} = -105$ K, $\theta = -86$ K, $J/k_{B,\text{intra}} = -15.2$ K, $\mu_{\text{eff.}} = 2.19 \mu_B$, 1.7×10^{-3} , dimers ³⁵
Bu ₄ N ⁺	2, —, $\theta = -57$ K, $\mu_{\text{eff.}} = 1.72 \mu_B$, 0.1×10^{-3} , dimers ¹²	3, —, $\theta = -9.6$ K, $\mu_{\text{eff.}} = 2.37 \mu_B$, 0.3×10^{-3} , dimers ¹²
Pent ₄ N ⁺	4, —, $\theta = -1$ K, $\mu_{\text{eff.}} = 1.80 \mu_B$, dimers	5, $\theta = -1.4$ K, $J/k_{B,\text{intra}} = +5.6$ K, $\mu_{\text{eff.}} = 2.38 \mu_B$, dimers
Hex ₄ N ⁺	6, —, $\theta = -3$ K, $\mu_{\text{eff.}} = 1.78 \mu_B$, no dimers	7, $\theta = -2.3$ K, $J/k_{B,\text{intra}} = +3.3$ K, $\mu_{\text{eff.}} = 2.41 \mu_B$, no dimers
Cryptand(Na ⁺)	—	15, $\theta = -1.3$ K, $J/k_{B,\text{intra}} = +7.7$ K, $\mu_{\text{eff.}} = 2.22 \mu_B$, no dimers ²⁷

The magnetic properties of salts 4–7 are presented in the ESI† (Fig. S15–S18). The Pent₄N⁺ cations in (Pent₄N⁺){Ti^{IV}O(Pc³⁻)^{·-}} (4) are larger than the Bu₄N⁺ ones. The overlapping between the Pc³⁻ macrocycles is still preserved in [$\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^{\cdot-}$]₂ but becomes weaker due to an additional shift of the macrocycles (Fig. 4). This salt shows a nearly paramagnetic behavior with a Weiss temperature close to zero ($\theta = -1$ K, Fig. S15†) which shows the absence of magnetic coupling between spins in the dimers. It is seen that the use of Pent₄N⁺ cations suppresses magnetic coupling between spins within the dimers due to the essential shift of the macrocycles. Thus, a simple variation of cation size in such systems allows regulation of magnetic coupling between spins ranging from very strong up to nearly the absence of any magnetic coupling.

Salt (Hex₄N⁺){Ti^{IV}O(Pc³⁻)^{·-}}·C₆H₄Cl₂ (6) has a different packing of {Ti^{IV}O(Pc³⁻)^{·-}} which are arranged in the one-

dimensional chains (Fig. 5). In this case only weak antiferromagnetic coupling between spins is realized with a Weiss temperature of -3 K (Table 2 and Fig. S16†). Similar weak magnetic interactions were found for different salts with the radical anions of copper(II), nickel(II), tin(II), lead(II) and metal-free phthalocyanines having nearly isolated packing of the Pc³⁻ (or H₂Pc^{·-}) macrocycles.¹² The largest magnetic coupling among these salts with a Weiss temperature of -17 K was found for (Bu₄N⁺)₂{Pb^{II}(Pc³⁻)^{·-}}³⁻(Br⁻).¹²

Several salts (3, 5, 7, 11, and 15) synthesized with the {V^{IV}O(Pc³⁻)^{·-}} radical anions are isostructural to the {Ti^{IV}O(Pc³⁻)^{·-}} salts (Table 2). These salts have effective magnetic moments from 2.19 to 2.41 μ_B at 300 K in accordance with the presence of two $S = 1/2$ spins per one radical anion (Table 2). The calculated value for the system with two noninteracting $S = 1/2$ spins is equal to 2.45 μ_B . To understand the magnetic behavior of individual {V^{IV}O(Pc³⁻)^{·-}} radical anions we consider the magnetic behavior of salts with isolated Pc³⁻ packing: (Pent₄N⁺){V^{IV}O(Pc³⁻)^{·-}} (5), (Hex₄N⁺){V^{IV}O(Pc³⁻)^{·-}}·C₆H₄Cl₂ (7) and {cryptand[2,2,2](Na⁺)}{V^{IV}O(Pc³⁻)^{·-}}·C₆H₄Cl₂ (15). All of them show similar magnetic behavior with a slight increase of the magnetic moment below 60 K and its abrupt decrease at low temperatures ($T < 10$ K) (Fig. 7b and c for salts 5 and 7, respectively). Such a magnetic behavior can be described well by a modified singlet-triplet (S-T) model³⁹ with a rather strong ferromagnetic intramolecular exchange interaction between the Pc³⁻ and V^{IV} spins. This intramolecular exchange interaction (J/k_B) ranges from +3.3 up to +7.7 K, while weak antiferromagnetic intermolecular coupling between the Pc³⁻ macrocycles is found with a Weiss temperature (θ) from -1.3 up to -2.3 K. In the case of 5, 7 and 15 intermolecular coupling is weak due to the absence of π - π interactions between Pc³⁻. When more effective π - π interactions are realized in 3 and 11, antiferromagnetic intermolecular exchange interactions can attain essentially larger values (up to $J/k_B = -105$ K for 11).³⁵ Since this value is nearly 10 times larger than those for the ferromagnetic intramolecular interactions, the total magnetic behavior of these salts is antiferromagnetic. Ferromagnetic intramolecular interactions and antiferromagnetic intermolecular

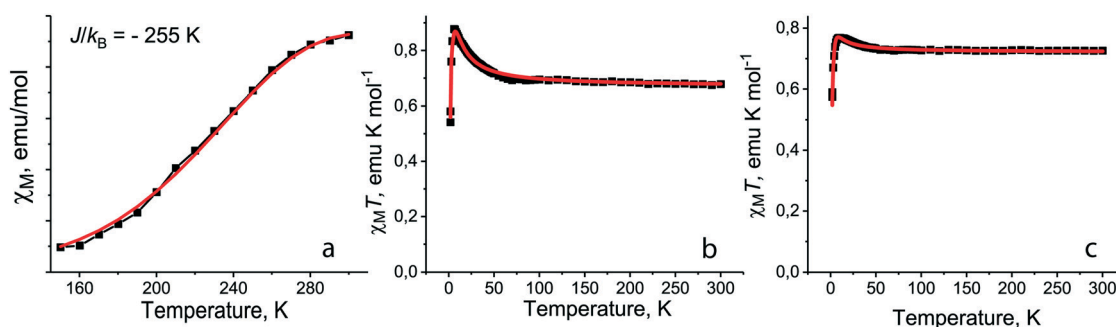


Fig. 7 (a) Temperature dependence of molar magnetic susceptibility of 1 in the 150–300 K range (black squares) and fitting the data by the Heisenberg model for quasi-square layered arrangement of spins³⁸ with $J/k_B = -255$ K (red curve); fitting the experimental $\chi_M T$ data (black squares) by the modified S-T model³⁹ for 5 (b) with $J_{\text{intra}}/k_B = +5.6$ K, $\theta = -1.4$ K, $g = 2$ (fix), $f = 0.996$ (red curve) and for 7 (c) with $J_{\text{intra}}/k_B = +3.3$ K, $\theta = -2.3$ K, $g = 2$ (fix), $f = 0.999$ (red curve).



interactions suppress each other. As a result, a comparative analysis of magnetic properties of salts **2**, **3** and **10**, **11** shows that the total antiferromagnetic coupling is noticeably weaker for the $\{V^{IV}(Pc^{3-})\}^-$ salts in comparison with the $\{Ti^{IV}(Pc^{3-})\}^-$ salts (Table 2).

The EPR spectra of newly prepared compounds are shown in Fig. S19–S29† and the parameters of the signals are listed in Table 3. All salts with the $\{Ti^{IV}(Pc^{3-})\}^-$ radical anions manifest EPR signals containing a main component with $g = 1.9994$ – 2.0008 and a linewidth (ΔH) of 1.50 – 3.67 mT. Only in case of salt **4** was a two-component spectrum found with the components of nearly equal intensity and linewidth (Fig. S19†). In spite of the $S = 1/2$ spin is delocalized over the Pc^{3-} macrocycles in $\{Ti^{IV}(Pc^{3-})\}^-$, the parameters of the EPR signal in salts **1**, **2**, **4**, **6**, **10** and **13** are noticeably different from those of the salts with the metal-free phthalocyanine ($H_2Pc^{\cdot-}$) radical anions ($g = 2.0030$ – 2.0035 and $\Delta H = 0.1$ – 0.3 mT).^{12,18} Therefore, the central metal Ti^{IV} atom, in spite of diamagnetism, affects the parameters of the EPR signal from the macrocycle. A temperature decrease generally narrows the

EPR signals and increases their asymmetry since they are split into two or three lines at low temperatures (Table 3). Salts containing the $\{V^{IV}(Pc^{3-})\}^-$ radical anions show essentially broader EPR signals with $g = 1.9878$ – 1.9954 and $\Delta H = 7.80$ – 25.30 mT. Since these species contain two paramagnetic centers (V^{IV} and Pc^{3-} with an $S = 1/2$ spin state), the observed signal originates from both centers having an exchange interaction. As a result, the observed signal has approximately intermediate parameters between those characteristic of V^{IV} and Pc^{3-} . In fact, for example, the room-temperature spectrum of pristine $[V^{IV}OPc^2]_0$ contains a broad main component ($\Delta H = 24.9$ mT) from V^{IV} with $g = 1.9858$ (ref. 12), whereas the EPR spectra of the salts with $\{Ti^{IV}(Pc^{3-})\}^-$ containing Pc^{3-} manifest EPR signals with the main component at $g = 1.9994$ – 2.0008 and $\Delta H = 1.50$ – 3.67 mT (Table 3). Generally, a temperature decrease results in narrowing of the EPR signals from the salts and asymmetry increase as in case of the $\{Ti^{IV}(Pc^{3-})\}^-$ salts. However, in several cases (**11** and **14**) essential broadening and shift to the higher g -factors are observed for the main components. This can be explained (as was shown previously for **11**)³⁵ by a strong antiferromagnetic coupling of spins delocalized over Pc^{3-} . The contribution from Pc^{3-} disappears at low temperatures due to the antiparallel arrangement of these spins, providing the manifestation of contribution only from weaker interacting V^{IV} spins which have broad EPR signals at low temperatures.

Table 3 Data of EPR measurements for salts **1–7**, **10–11** and **13–15**

N	Room temperature g -Factor (linewidth (mT), relative intensity (%))	Low temperature g -Factor (linewidth (mT), relative intensity (%))
Salts with $\{Ti^{IV}(Pc^{3-})\}^-$		
1 (ref. 12)	2 lines	2 lines, 4.2 K
	2.0006 (3.54, 97%)	1.9991 (1.0, main)
	2.0025 (0.78, 3%)	2.0023 (0.52)
2 (ref. 12)	1 line	2 lines, 4.2 K
	2.0003 (1.50)	2.0010 (0.57, 50%) 1.9981 (1.00, 50%)
4	2 line	3 lines, 50 K
	2.0033 (0.77, 50%)	2.0037 (0.32, 20%)
	1.9994 (1.17, 50%)	2.0008 (0.62, 50%) 1.9963 (0.55, 30%)
6	1 line	2 lines, 50 K
	2.0008 (3.58)	1.9952 (1.37, 80%) 2.0017 (0.74, 20%)
10 (ref. 35)	2 lines	—
	2.0001 (3.67, 93.2%)	
	2.0031 (1.82, 6.8%)	
13	1 line	2 lines, 4.2 K
	2.0007 (2.62)	2.0024 (0.54, 45%) 1.9981 (1.25, 55%)
Salts with $\{V^{IV}(Pc^{3-})\}^-$		
3 (ref. 12)	1 line	2 lines, 4.2 K
	1.9918 (7.8)	1.9886 (7.12, 80%) 1.9928 (4.04, 20%)
5	1 line	2 lines, 50 K
	1.9902 (10.2)	1.9849 (7.86, 10%) 1.9868 (29.7, 90%)
7	1 line	1 line
	1.9919 (9.17)	1.9893 (3.74)
11 (ref. 35)	2 lines	2 lines, 6.9 K
	1.9954 (25.32, 90%)	1.9832 (32.2, 40%)
	2.0002 (10.49, 10%)	2.0111 (27.3, 60%)
14	1 line	2 lines
	1.9878 (11.19)	1.9851 (9.97, 30%) 2.1323 (55.23, 70%)
15 (ref. 27)	1 line	1 line
	1.9923 (8.91)	1.9912 (4.37)

4. Control of the crystal packing of the macrocycles.

It is possible to control the crystal packing of the macrocycles by using different synthetic approaches. In the presence of small cations the $\{[M^{IV}(Pc^{3-})]_2\}^-$ dimers are packed in layers when the area size occupied by couples of the cations and solvent molecules (a total of 4 units) are close to that of the phthalocyanine macrocycle. For example, in $(Et_4N^+)\{Ti^{IV}(Pc^{3-})\}^- \cdot C_6H_4Cl_2$ (**1**) two positions are occupied by the Et_4N^+ cations and two positions are occupied by solvent $C_6H_4Cl_2$ molecules as described above (Fig. 2b). Synthesis with an excess of $(MDABCO^+)(I^-)$ results in the formation of titanyl and vanadyl phthalocyanine salts with four $MDABCO^+$ cations per one $\{[M^{IV}(Pc^{3-})]_2\}^-$ dimer. In this case two I^- anions are involved in the salt to compensate for an excess of positive charge of the cations in $(MDABCO^+)_2[M^{IV}(Pc^{3-})]_2(I^-)$ ($M = Ti$ for **13** and $M = V$ for **14**). As a result, a total of four $MDABCO^+$ cations are arranged near the phthalocyanine plane in both salts but two I^- anions are inserted between the cations (Fig. 8b). Such a packing yields the layered structure in which phthalocyanine layers composed of the $\{[M^{IV}(Pc^{3-})]_2\}^-$ dimers are formed (Fig. 8a). Since $MDABCO^+$ cations have a small size like Et_4N^+ , the arrangements of the macrocycles in the dimers of salt **1** and both salts **13** and **14** are similar. They have the same small shift of the Pc^{3-} macrocycles relative to each other (2.006 (**13**) and 2.003 Å (**14**)) and slightly larger interplanar distances of 3.205 Å for **13** and **14**. It is possible to suppose that similarly to salt **1**, salts **13** and **14** can show a





Fig. 8 (a) View of the phthalocyanine layer composed of the π -stacking $[\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}]_2$ dimers in **13**. Short van der Waals C,N...C,N contacts are shown by green dashed lines. (b) Projection of the $(\text{MDABCO})_2^-$ (I^-) layer on the phthalocyanine layer. Only one phthalocyanine from the dimer is shown (bonds are shown in grey). Salt **14** is isostructural to **13**.

rather strong magnetic coupling between the Pc^{3-} spins in the dimers.

For the 1:1 salts with small organic cations like $(\text{MDABCO}^+)\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ (**8**) and $(i\text{-Pr}_2\text{Im}^+)\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- \cdot 0.75\text{C}_6\text{H}_4\text{Cl}_2$ (**9**) another packing mode of the macrocycles is observed. In this case only two cations are positioned near each phthalocyanine plane, and these cations are alternated with the $[\{\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}]_2$ dimers in the chains as shown in Fig. 9. The structure is organized in such a way that two other positions above the phthalocyanine plane are occupied by phenylene substituents of two $\{\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ from the dimers belonging to the neighboring chains. In **8**, such chains form layers parallel to the ab plane and the direction of the chains

changes at the transition from one layer to another between $[1\ 1\ 0]$ and $[1\ -1\ 0]$. The $[\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}]_2$ dimers in **8** have a small shift of the Pc^{3-} macrocycles (2.005 Å) like that observed for **1** and **13** and the interplanar distance between them is 3.231 Å. Thus, a rather strong magnetic coupling between spins can be realized within the isolated $[\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}]_2$ dimers in **8**.

Salt $(i\text{-Pr}_2\text{Im}^+)\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- \cdot 0.75\text{C}_6\text{H}_4\text{Cl}_2$ (**9**) also contains alternating $[\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}]_2$ dimers and pairs of $i\text{-Pr}_2\text{Im}^+$ cations (Fig. 9b). This salt has some peculiarities due to the formation of multiple hydrogen bonds between hydrogen atoms of two $i\text{-Pr}_2\text{Im}^+$ cations and oxygen atoms of $\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ (Fig. 9b). The shortest bonds are formed between the H atom of the imidazole ring and two oxygen atoms of 2.487 and 2.525 Å lengths. Moreover, two hydrogen atoms of two methyl groups of $i\text{-Pr}$ of each $i\text{-Pr}_2\text{Im}^+$ approach close to the oxygen atoms. Each oxygen atom forms a total of four such short contacts with $\text{H}(i\text{-Pr})\cdots\text{O}$ distances of 2.58–2.88 Å (not shown in Fig. 9b). As a result, two oxygen atoms involved in the hydrogen bonding approach very close to each other at a 2.958 Å distance. Chains of alternating phthalocyanine dimers and cations in **9** are directed along the a axis. However, phthalocyanine planes have a different incline in the neighboring chains within the layer, forming a parquet structure. The dihedral angle between the 24-atom Pc planes of two $\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ from the neighboring chains is 57.88°. The shift of the Pc^{3-} macrocycles in the $[\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}]_2$ dimers is equal to 5.60 Å. This shift is closer to that in $(\text{Pent}_4\text{N}^+)\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ (**5**) which has a weak macrocycle overlapping and weak magnetic coupling of spins.

The multi-component approach allows $(\text{MDABCO}^+)(\text{TPC})\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ (**12**) to be synthesized. In this salt, phthalocyanine layers composed of the $[\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}]_2$ dimers



Fig. 9 View of the chains composed of π -stacking $[\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}]_2$ dimers and the MDABCO^+ cations in **8** (a) and the $[\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}]_2$ dimers and $i\text{-Pr}_2\text{Im}^+$ cations in **9** (b). Van der Waals C,N...C,N contacts and hydrogen $\text{H}(i\text{-Pr}_2\text{Im}^+)\cdots\text{O}(\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-)$ bonds are shown by green and red dashed lines, respectively.



alternate with the cationic MDABCO⁺ layers containing additional neutral TPC molecules (Fig. 10b). One-dimensional chains from the $[\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-]_2$ dimers with close interphthalocyanine contacts can be outlined in the layers of **12** along the *c* axis (Fig. 11a). These chains are formed in such a way that the planes of phthalocyanines in each dimer are located obliquely relative to the cationic layers and dimers form stacks along the *c* axis with effective overlapping between phthalocyanines (Fig. 11a). The view along the phthalocyanine layers approximately along the *a* axis shows that the overlapping between the Pc^{3-} macrocycles in this direction is realized as in other salts at the shift of Pc^{3-} in the dimers (Fig. 11b). As a result, the packing of phthalocyanine layers in **12** is closer to a two-dimensional one (Fig. 10a). Previously, a multi-component layered fullerene complex (MDABCO⁺)-(TPC)(C₆₀⁻) including the MDABCO⁺ and TPC components was obtained.⁴¹ However, the MDABCO⁺-TPC layers in **12** have essentially looser packing compared with (MDABCO⁺)-(TPC)(C₆₀⁻). The distances between positively charged nitrogen atoms of MDABCO⁺ in the layers of **12** are 13.50 and 13.84 Å and these distances are essentially larger than 10 Å in the fullerene complex.⁴¹ This is due to the insertion of oxygen atoms of $\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ into the cationic layers and the orientation of MDABCO⁺ cations lying sideways in **12**. The shift of the Pc^{3-} macrocycles in the $[\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-]_2$ dimers of **12** is 4.29 Å and the interplanar distance is 3.251 Å. The observed shift is smaller than that in **2** (5.12 Å) but essentially larger than that in **1** (2.005 Å). It is seen that using different synthetic approaches and small cations of different sizes and shapes allows one to essentially modify the crystal structures of the salts, potentially affecting the magnetic and optical properties.

To conclude on the overlapping of Pc^{3-} in the phthalocyanine dimers, four types of π - π interactions between the 24-atom Pc planes can be considered (Fig. 12). The most effective overlapping with the shift of the Pc^{3-} macrocycles by

2.005 Å is observed in **1** (Fig. 12a) as well as in **8** and **13**. Partial overlapping with the shift of the Pc^{3-} macrocycles by 5.12 Å is observed in **2** (Fig. 12b). Weak overlapping with the shift of the Pc^{3-} macrocycles by 5.61 Å is found for **4** and **9** (Fig. 12c). Salt **12** has intermediate overlapping with the shift of the Pc^{3-} macrocycles by 4.29 Å (Fig. 12d), but in this case a different packing mode of the dimers in the layers is observed. Isostructural salts with the $[\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-]_2$ dimers have a similar arrangement of the 24-atom Pc planes within the dimers.

5. Jahn–Teller distorted molecular structure and effect of reduction on the optical properties of titanyl and vanadyl phthalocyanines

In this section we consider the effect of reduction on the geometry of titanyl and vanadyl phthalocyanines (Table 4) and their optical properties. The structure of pristine $\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{2-})$ was studied by the Rietveld method⁴² but that of $\text{V}^{\text{IV}}\text{O}(\text{Pc}^{2-})$ was studied by single crystal X-ray diffraction⁴³ (Table 4). There are two types of C–N bonds in the Pc macrocycles with pyrrole (pyr) and imine (im) nitrogen atoms. The C–N_{pyr} bond lengths are 1.374(6) and 1.387(4) Å in $\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{2-})$ and $\text{V}^{\text{IV}}\text{O}(\text{Pc}^{2-})$, respectively,^{42,43} which are longer than those of the C–N_{im} bonds. No alternation of the C–N_{pyr} is manifested at the formation of $\{\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ but they are slightly elongated in the salts (Table 4). Shorter C–N_{im} bonds neither alternate in pristine phthalocyanines.^{42,43} However, the formation of the $\{\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ radical anions provides obvious alternation of these bonds since four bonds belonging to two oppositely located isoindole units become short but four other bonds belonging to two other oppositely located isoindole units become long (Table 4 and Fig. 13a). The difference between the short and long bonds of 0.022–0.038 (Table 4) essentially exceeds the error in the determination of the lengths of these bonds. Additionally, the alternation of



Fig. 10 (a) View of the phthalocyanine layer composed of the π -stacking $[\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-]_2$ dimers in **12**. Short van der Waals C,N...C,N contacts are shown by green dashed lines. (b) Projection of the (MDABCO⁺)-TPC layer on the phthalocyanine layer. Only one phthalocyanine from the dimer is shown (bonds are shown in grey).



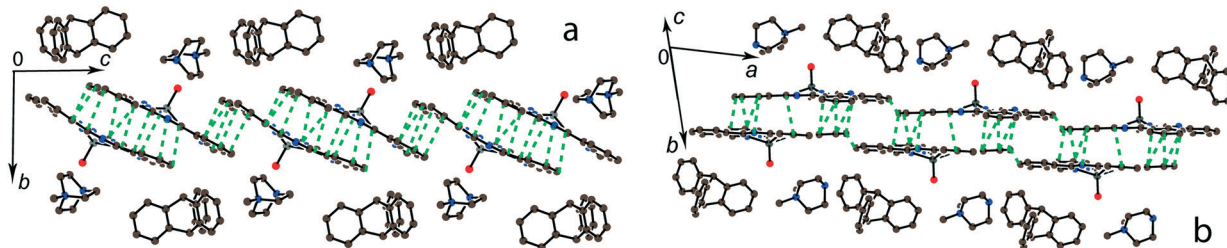


Fig. 11 View along the phthalocyanine layers composed of the $\{Ti^{IV}(O)(Pc^{3-})\}_2$ dimers in **12**: view along the *a* (a) and approximately along the *c* (b) axes. Short van der Waal contacts between phthalocyanines are shown by green dashed lines.

the C–C bonds in the pyrrole rings (which are not involved in the phenylene substituents) is also observed, and among eight bonds four bonds are short and four other bonds are long as shown in Fig. 13a. The difference between the short and long bonds is in the 0.020–0.029 Å range (Table 4).

The reason for the bond alternation can be clearly explained in terms of the Jahn–Teller (JT) theory. According to the theoretical works^{44–46} the calculated potential surface of phthalocyanine radical anions has four special points: two equivalent minima and two saddle points (two barriers between these minima). The potential minima correspond to rectangular deformations of the Pc macrocycles while transition states correspond to rhombic ones. All these deformations reduce the symmetry from C_{4v} to C_{2v} but they are characterized by different bond structures as shown in Fig. 13b and c. Recent studies show relatively small energy barriers between the JT deformations for free phthalocyanines in the gas phase, suggesting the dynamical JT phenomena. However, in the case of monoclinic crystals, the pres-

ence of a low-symmetry environment for the radical anions stabilizes the lower-symmetry JT configurations. It can be concluded from a comparative analysis of experimental data and the data of theoretical calculations that rectangular distortions are observed for the Pc^{3-} macrocycles in all the salts with the alternation of the C–N_{im} and the C–C bonds in the pyrrole rings (Fig. 13a and b). However, no predicted alternation of the N_{pyr}–M^{IV} bonds (Fig. 13b) was found. Rhombic type distortions (Fig. 13c) were not found in the experimental structures of the $\{M^{IV}(O)(Pc^{3-})\}_2$ radical anions.

Disruption of aromaticity can give an additional effect on the alternation of the bonds in the macrocycles. The Pc^{2-} macrocycles in pristine phthalocyanines have a stable aromatic 18 π -electron system. The formation of a less stable 19 π -electron system in Pc^{3-} should be accompanied by partial disruption of their aromaticity, which enhances the effect of alternation. From this point of view it is possible to explain the increase in the difference between short and long C–N_{im} bonds (0.060–0.071 Å) in the Pc^{4-} macrocycles in the salts with substituted titanil phthalocyanines and porphyrazines⁴⁷ since in this case an even less stable 20 π -electron system is formed.

In addition to distortions, shortening of the M–N_{pyr} bonds is found at the formation of $\{M^{IV}(O)(Pc^{3-})\}_2$ (Table 4). Therefore, the appearance of an additional electron on the Pc^{3-} macrocycles slightly enhances the bonding ability of a Pc ligand. In accordance with the shortening of the M–N_{pyr} bonds the displacement of metal atoms from the 24-atom Pc plane decreases at the formation of salts (Table 4). It is also seen that the M=O bonds are slightly elongated in the salts (Table 4) most probably due to stronger bonding of Pc^{3-} to the metal atoms.

The spectra of the salts in the IR region are presented in Fig. S1–S9,[†] whereas those in the UV-visible-NIR range are presented in Fig. S10–S14.[†] The optical properties of salts **1–3**, **10**, **11**, and **15** were published previously.^{12,27,35} Pristine $Ti^{IV}(O)(Pc^{2-})$ shows the Soret band at 350 nm and the Q-band with a maximum at 713 nm (Fig. 14). Obviously, the Q-band is split into three bands but other bands are not well pronounced. There is no absorption in the spectrum of $Ti^{IV}(O)(Pc^{2-})$ over 900 nm (Fig. 14). The spectra of the salts are similar and strongly different from those of pristine $Ti^{IV}(O)(Pc^{2-})$ (Fig. 14). New bands are manifested in the spectra of the salts in the NIR range with maxima at 1000 nm (Fig. 14). These bands are associated with the population of



Fig. 12 View of the π -stacking $\{Ti^{IV}(O)(Pc^{3-})\}_2$ dimers along the M=O bonds: (a) optimal overlapping observed in **1**, **8** and **13**; (b) partial overlapping observed in **2**; (c) weak overlapping observed in **4** and **9**; (d) intermediate overlapping in **12**. Only 24-atom Pc planes are shown.



Table 4 Geometric parameters of titanyl and vanadyl phthalocyanines in pristine compounds and selected radical anion salts

Compound	C–N _{pyr}	Average bond length in the macrocycle, Å (see Fig. 13a)			C–C bonds in the pyrrole ring			Average M–N(Pc) bond, Å	The M=O bond, Å	Displacement, ^a Å
		C–N _{im}	Short	Long	Difference	Short	Long			
Ti ^{IV} O(Pc ²⁻) ^b (ref. 42b)	1.375(5)	1.329(5)		0	1.442(4)		0	2.071(5)	1.643(3)	0.742
(MDABCO ⁺) {Ti ^{IV} O(Pc ³⁻) ⁻ } ⁻ (8)	1.382(3)	1.312(3)	1.351(3)	0.038	1.429(3)	1.458(3)	0.029	2.053(2)	1.645(2)	0.605
(MDABCO ⁺)(TPC) {Ti ^{IV} O(Pc ³⁻) ⁻ } ⁻ (12)	1.388(3)	1.320(3)	1.343(3)	0.023	1.438(3)	1.458(3)	0.020	2.055(2)	1.649(2)	0.683
(MDABCO ⁺) ₂ {Ti ^{IV} O(Pc ³⁻) ⁻ } ⁻ (Γ) (13)	1.390(3)	1.319(4)	1.357(4)	0.038	1.442(3)	1.467(3)	0.025	2.061(2)	1.650(2)	0.572
(Pent ₄ N ⁺) ₁ {M ^{IV} O(Pc ³⁻) ⁻ } ⁻ (4) 2	1.386(3)	1.313(4)	1.351(3)	0.038	1.439(3)	1.463(3)	0.024	2.067(2)	1.647(2)	0.754
{M ^{IV} O(Pc ³⁻) ⁻ } ⁻ (4) 2	1.386(3)	1.315(4)	1.352(4)	0.038	1.434(3)	1.458(3)	0.024	2.064(2)	1.649(2)	0.768
V ^{IV} O(Pc ²⁻) (ref. 43)	1.387(4)	1.334(4)		0	1.442(4)		0	2.044(4)	1.595(4)	0.696
(MDABCO ⁺) ₂ {V ^{IV} O(Pc ³⁻) ⁻ } ⁻ (Γ) (14)	1.392(2)	1.318(2)	1.355(2)	0.037	1.439(2)	1.467(2)	0.028	2.031(2)	1.613(1)	0.522
(Pent ₄ N ⁺) ₁ {V ^{IV} O(Pc ³⁻) ⁻ } ⁻ (5) 2	1.393(3)	1.315(3)	1.352(3)	0.037	1.440(3)	1.466(3)	0.026	2.037(2)	1.605(2)	0.690
{V ^{IV} O(Pc ³⁻) ⁻ } ⁻ (5) 2	1.388(3)	1.318(3)	1.354(3)	0.036	1.439(3)	1.462(3)	0.023	2.038(2)	1.610(2)	0.699

^a Displacement of metal atoms from the 24-atom plane. ^b Made by the Rietveld analysis.

the SOMO orbital of {Ti^{IV}O(Pc³⁻)⁻}⁻ from which a new electronic transition takes place to the above located orbitals. An essential blue shift of both Soret and Q-bands is seen at the formation of Pc³⁻. This indicates the increase of the energy of transition associated with the Soret and the Q-bands. The Soret band is positioned at 337–344 nm in the spectra of the salts but the Q-band is split into three bands observed at 592–599, 630–636 and 700–706 nm (Fig. 14). The intensity of relatively weak and broad charge transfer (CT) bands between {Ti^{IV}O(Pc³⁻)⁻}⁻ in the dimers is defined by the π - π interactions between Pc³⁻. Weak overlapping of Pc³⁻ in 2, 4 and 6 does not allow the observation of CT bands. These bands are manifested in the spectra of salts 1 and 10 with noticeable overlap integrals. These bands are observed in the spectra of 1 (Fig. 14) and 10 at 1484 and 1467 nm, respectively.

The spectrum of pristine V^{IV}O(Pc²⁻) (Fig. 15) shows features similar to those of Ti^{IV}O(Pc²⁻) with the single Soret

band at 351 nm and the split Q-band with a maximum at 723 nm. There is no absorption in the spectrum of V^{IV}O(Pc²⁻) over 900 nm (Fig. 15). New bands appear in the spectra of the salts at 1000–1006 nm. Both Soret and Q-bands are noticeably blue shifted and appear in the spectra of the salts as a single band at 339 nm (the Soret band) and the split Q-band at 593–604 and 634–640, and the third band is well pronounced in the spectrum of 11 at 693 nm (Fig. 15). There is also a set of weak bands between 700 and 1000 nm (Fig. 15). Similarly to the {Ti^{IV}O(Pc³⁻)⁻}⁻ salts there are no CT bands in the spectra of 5 and 7 which have relatively weak Pc³⁻ overlapping. At the same time an intense CT band with a maximum at 1447 nm is observed in the spectrum of 11 which has the largest π - π overlapping between Pc³⁻ among the studied {V^{IV}O(Pc³⁻)⁻}⁻ salts and shows effective magnetic coupling between spins. Thus, the presence of relatively

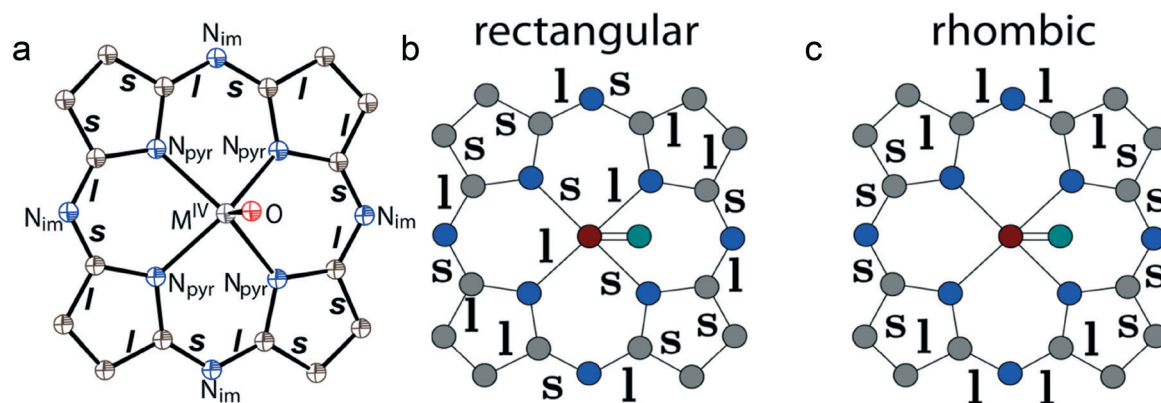


Fig. 13 (a) Experimentally observed alternation of the C–N(imine) and C–C pyrrole bonds in the {M^{IV}O(Pc³⁻)⁻}⁻ radical anions. Schematic diagrams of the theoretically predicted rectangular (a global minimum) (b) and rhombic (a saddle point). (c) The C_{2v} JT deformations of the central part of the {M^{IV}O(Pc³⁻)⁻}⁻ radical anions. The symbols l and s indicate C,N...C,N bonds which are elongated and shortened, respectively, at the reduction of Pc²⁻ to Pc³⁻.



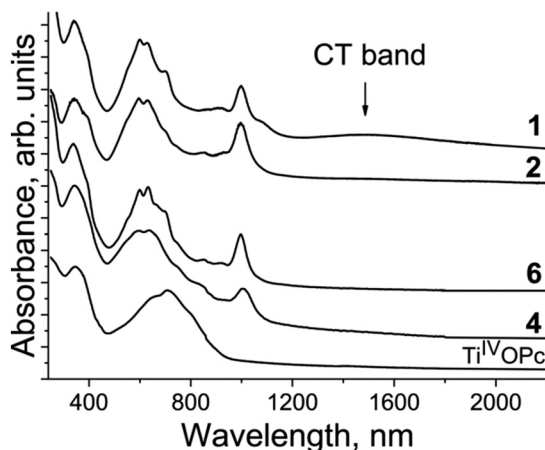


Fig. 14 Spectra of pristine $\text{Ti}^{\text{IV}}(\text{Pc}^{2-})$ and selected salts 1, 2, 4 and 6 in the UV-visible-NIR ranges in KBr pellets. Pellets for the salts were prepared under anaerobic conditions.

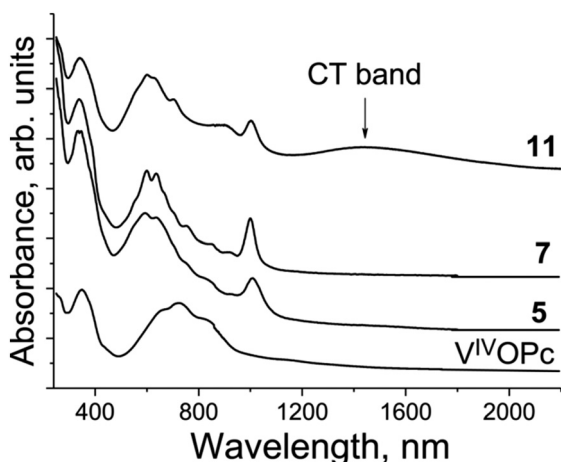


Fig. 15 Spectra of pristine $\text{V}^{\text{IV}}(\text{Pc}^{2-})$ and selected salts 5, 7 and 11 in the UV-visible-NIR ranges in KBr pellets. Pellets for the salts were prepared under anaerobic conditions.

intense CT bands in the spectra of the salts with phthalocyanine dimers can guarantee the observation of rather effective magnetic coupling between spins.

Conclusion

We developed several approaches for the synthesis of salts with titanyl and vanadyl phthalocyanine radical anions: the synthesis of 1:1 salts as solvent-free or solvent-containing phases with the cations of different sizes, the insertion of neutral components into the radical anion salts and finally the synthesis of the salts containing two cations and an additional counter anion per one phthalocyanine molecule. The reduction of these metallophthalocyanines centered on the macrocycles is accompanied by the formation of the Pc^{3-} radical trianions. Titanyl and vanadyl phthalocyanines containing one axial oxygen ligand tend to form π -stacking $[\{\text{M}^{\text{IV}}(\text{Pc}^{3-})\}]_2^-$ dimers with effective π - π interactions between the Pc^{3-}

macrocycles. In most cases these dimers are packed in layers which alternate with the cationic layers, while the too large-sized tetrahexylammonium cation did not allow the dimer formation. The size of cationic layers defines the size of phthalocyanine layers composed of the dimers. As a result, the increase of the cation size provides the shift of the Pc^{3-} macrocycles in the dimers. The decrease of the π - π interactions between the Pc^{3-} macrocycles results in the decrease of magnetic coupling between spins and suppresses CT between Pc^{3-} macrocycles. In this case it is possible to tune the magnetic and optical properties of the salts by a simple variation of cation size. Titanyl and vanadyl phthalocyanine radical anions have different magnetic states since Ti^{IV} is diamagnetic and the $S = 1/2$ spin is delocalized on Pc^{3-} (π -radical), whereas both V^{IV} and Pc^{3-} have an $S = 1/2$ spin in $\{\text{V}^{\text{IV}}(\text{Pc}^{3-})\}^-$ (d- and π -electrons). It is shown that in the latter case spins are ferromagnetically coupled within the radical anion with an exchange interaction of +3.3 to +7.7 K. Intermolecular coupling between the macrocycles is antiferromagnetic and at effective π - π interaction between Pc^{3-} is several times stronger than intramolecular ferromagnetic coupling. As a result, due to mutual suppression of ferro- and antiferromagnetic interactions, the resulting antiferromagnetic coupling is weaker for the $\{\text{V}^{\text{IV}}(\text{Pc}^{3-})\}^-$ salts in comparison with the $\{\text{Ti}^{\text{IV}}(\text{Pc}^{3-})\}^-$ salts. Reduction provides alternation of the C- N_{im} bonds and the C-C bonds of the pyrrole rings in the Pc^{3-} macrocycle, the appearance of new bands in the NIR range and the essential blue shift of both Soret and Q-bands. The salts obtained in some cases show effective magnetic coupling between spins especially when only π -electrons are present on the macrocycles. Further enhancement of this coupling at the extension of the π -systems of titanyl and vanadyl macroheterocycles is in progress.

Experimental

Materials

(Pent_4N^+)(Br^-) (Aldrich, >99%), (Hex_4N^+)(Br^-) (Aldrich, >99%), triptycene (TPC, Aldrich, 98%), 1,3-di-*i*-propylimidazolium chloride, ($(i\text{-Pr}_2\text{Im}^+)(\text{Cl}^-)$, Strem, >97%), and $[\text{Ti}^{\text{IV}}(\text{Pc}^{2-})]^0$ (Acros 85%) were used as received. $[\text{Ti}^{\text{IV}}(\text{Pc}^{2-})]^0$ was obtained by 2 hour boiling of $[\text{Ti}^{\text{IV}}\text{Cl}_2(\text{Pc}^{2-})]^0$ (95%, Aldrich) in wet pyridine (5% of H_2O) as reported.⁴⁸ Sodium fluorenone ketyl was obtained according to the literature.³⁶ (MDABCO^+)(Γ^-) was obtained as previously described.⁴⁹ Solvents were purified in an argon atmosphere and degassed. *o*-Dichlorobenzene ($\text{C}_6\text{H}_4\text{Cl}_2$) was distilled over CaH_2 under reduced pressure, benzonitrile was distilled over Na under reduced pressure, and *n*-hexane was distilled over Na/benzophenone. All manipulations for the syntheses of 4–9 and 12–14 were carried out in an MBraun 150B-G glove box with a controlled argon atmosphere and the content of H_2O and O_2 less than 1 ppm. The solvents and crystals were stored in the glove box. Polycrystalline samples of 4–7 were placed in



quartz tubes of 2 mm diameter in anaerobic conditions (under argon) and sealed at ambient pressure or under a 10^{-5} torr vacuum for SQUID and EPR measurements. KBr pellets for IR- and UV-visible-NIR measurements were prepared in the glove box.

Synthesis

Syntheses of crystals of 1–3,¹² 10–11 (ref. 35) and 15 (ref. 27) were described previously. The crystals of other salts were obtained by diffusion technique. The reaction mixture was cooled to room temperature and filtered into a glass tube for diffusion, of 1.8 cm diameter and 50 mL volume with a ground glass plug, and then 30 mL of *n*-hexane was layered over the solution. Slow mixing of the solutions resulted in the precipitation of crystals over 1–2 months. The solvent was then decanted from the crystals, and they were washed with *n*-hexane. All crystals had a characteristic copper luster. The compositions of the obtained salts were determined from X-ray diffraction analysis on a single crystal. Several crystals from one synthesis were found to consist of a single crystal-line phase. Due to the high air sensitivity of 4–9 and 12–14, elemental analysis could not be used to confirm the composition because these salts reacted with oxygen in the air before the quantitative oxidation procedure could be performed.

The crystals of $(\text{Pent}_4\text{N}^+)\{\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ ($\text{M} = \text{Ti}$ (4), V (5)) and $(\text{Hex}_4\text{N}^+)\{\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- \cdot \text{C}_6\text{H}_4\text{Cl}_2$ ($\text{M} = \text{Ti}$ (6), V (7)) were obtained by the following procedure. The reduction of titanyl or vanadyl phthalocyanines (24.4 mg, 0.042 mmol) in 16 ml of $\text{C}_6\text{H}_4\text{Cl}_2$ with sodium fluorenone ketyl (14 mg, 0.069 mmol) in the presence of one equivalent of $(\text{Pent}_4\text{N}^+)(\text{Br}^-)$ (15.8 mg, 0.042 mmol for 4 and 5) or $(\text{Hex}_4\text{N}^+)(\text{Br}^-)$ (18.1 mg, 0.042 mmol for 6 and 7) for one day at 80 °C yielded a deep blue solution of the salts. Crystals were obtained as black prisms (4), blocks (5, 6) and black trapezoids (7) in 74%, 83%, 66% and 72% yields, respectively.

The crystals of $(\text{MDABCO}^+)\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ (8) were obtained by the reduction of titanyl phthalocyanine (24.4 mg, 0.042 mmol) in 14 ml of $\text{C}_6\text{H}_4\text{Cl}_2$ with sodium fluorenone ketyl (12 mg, 0.059 mmol) in the presence of exactly one equivalent of $(\text{MDABCO}^+)(\Gamma^-)$ (10.6 mg, 0.042 mmol) for one day at 80 °C. The reduction yielded a deep blue solution of the salt. To increase the solubility of the salt benzonitrile was added (3 ml) to the solution. Salt 8 was obtained as black irregular crystals in 42% yield.

The crystals of $(i\text{-Pr}_2\text{Im}^+)\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- \cdot 0.75\text{C}_6\text{H}_4\text{Cl}_2$ (9) were obtained by the reduction of vanadyl phthalocyanine (24.4 mg, 0.042 mmol) in 16 ml of $\text{C}_6\text{H}_4\text{Cl}_2$ with sodium fluorenone ketyl (14 mg, 0.069 mmol) in the presence of a double excess of $(i\text{-Pr}_2\text{Im}^+)(\text{Cl}^-)$ (15.75 mg, 0.082 mmol) for one day at 100 °C. The reaction yielded a deep blue solution of the salt. Crystals of 9 were obtained as black blocks in 34% yield. The crystal structure of 9 contains six solvent *o*- $\text{C}_6\text{H}_4\text{Cl}_2$ molecules per eight $(i\text{-Pr}_2\text{Im}^+)\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ units. Therefore, the composition of 9 is $(i\text{-Pr}_2\text{Im}^+)\{\text{V}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- \cdot 0.75\text{C}_6\text{H}_4\text{Cl}_2$.

The crystals of $(\text{MDABCO}^+)(\text{TPC})\{\text{Ti}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^-$ (12) were obtained by the reduction of titanyl phthalocyanine (24.4 mg,

0.042 mmol) in 14 ml of $\text{C}_6\text{H}_4\text{Cl}_2$ with sodium fluorenone ketyl (12 mg, 0.059 mmol) in the presence of exactly one equivalent of $(\text{MDABCO}^+)(\Gamma^-)$ (10.6 mg, 0.042 mmol) for one day at 100 °C. The reaction yielded a deep blue solution of the salt. To increase its solubility benzonitrile was added (3 ml) to the solution. The solution was cooled to room temperature and 150 mg of TPC (0.590 mmol) was dissolved in the obtained solution at room temperature for 4 hours. The final solution was filtered into a tube for diffusion. Crystals of 12 were obtained as black blocks in 27% yield.

The crystals of $(\text{MDABCO}^+)\{\text{M}^{\text{IV}}\text{O}(\text{Pc}^{3-})\}^- (\Gamma^-)$ ($\text{M} = \text{Ti}$ (13), V (14)) were obtained by the following procedure. The reduction of titanyl or vanadyl phthalocyanines (24.4 mg, 0.042 mmol) in 14 ml of $\text{C}_6\text{H}_4\text{Cl}_2$ with sodium fluorenone ketyl (14 mg, 0.069 mmol) in the presence of the excess of $(\text{MDABCO}^+)(\Gamma^-)$ (24 mg, 0.094 mmol) for one day at 100 °C yielded a deep blue solution of the salts. To increase the solubility of the salts benzonitrile was added (3 ml) to the solution. Crystals were obtained as dark black blocks (13) and plates (14) in 38% and 43% yields.

General

UV-visible-NIR spectra for 4–9 and 12–14 were measured in KBr pellets on a Perkin-Elmer Lambda 1050 spectrometer in the 250–2500 nm range. FT-IR spectra of 4–9 and 12–14 were obtained in KBr pellets with a Perkin-Elmer Spectrum 400 spectrometer ($400\text{--}7800\text{ cm}^{-1}$). EPR spectra were recorded for sealed polycrystalline samples of 4–7, 13, and 14 from 4 up to 295 K with a JEOL JES-TE 200 X-band ESR spectrometer equipped with a JEOL ES-CT470 cryostat. A Quantum Design MPMS-XL SQUID magnetometer was used to measure the static magnetic susceptibility of 4–7 at 100 mT magnetic field under cooling and heating conditions in the 300–1.9 K range. A sample holder contribution and core temperature independent diamagnetic susceptibility (χ_d) were subtracted from the experimental values. The χ_d values were estimated by the extrapolation of the data in the high-temperature range by fitting the data with the expression $\chi_M = C/(T - \theta) + \chi_d$, where C is the Curie constant and θ is the Weiss temperature. Effective magnetic moment (μ_{eff}) was calculated with the following formula: $\mu_{\text{eff}} = (8 \cdot \chi_M T)^{1/2}$.

Crystal structure determination

X-ray diffraction data for 4–9 and 12–14 were collected on an Oxford diffraction “Gemini-R” CCD diffractometer with graphite monochromated MoK_α radiation using an Oxford Instrument Cryojet system (Table 5). Raw data reduction to F^2 was carried out using CrysAlisPro, Oxford Diffraction Ltd. The structures of 4–9 and 12–14 were solved by direct methods and refined by the full-matrix least-squares method against F^2 using SHELX 2016/6, SHELX 2013 and Olex2 programs.⁵⁰ Non-hydrogen atoms were refined in the anisotropic approximation. The positions of hydrogen were calculated geometrically. Views of the unit cells of salts of 4–9 and 12–14 are shown in Fig. S30–S38.† One pentyl substituent of





Table 5 Crystallographic data for salts 4–9 and 12–14

Compound	4	5	6	7	8	9	12	13	14
Emp. formula	C ₁₀₄ H ₁₂₀ N ₁₈ O ₂ Ti ₂	C ₁₀₄ H ₁₂₀ N ₁₈ O ₂ Y ₂	C ₆₂ H ₇₂ Cl ₂ N ₉ OTi	C ₆₂ H ₇₂ Cl ₂ N ₉ OV	C ₃₀ H ₃₁ N ₁₀ OTi	C ₄₁ H ₄₆ N ₁₀ OV	C ₅₉ H ₄₅ N ₁₀ OTi	C ₄₆ H ₄₆ N ₁₂ OTi	C ₄₆ H ₄₆ N ₁₂ OV
M _r [g mol ⁻¹]	1749.97	1756.05	1078.08	1081.12	703.64	732.71	957.95	957.75	960.79
Color and shape	Black, prism	Black, block	Black, block	Black, trapezoid	Black, irregular	Black, block	Black, block	Black, block	Black, plate
Crystal system	Triclinic	Triclinic	Monoclinic	Monoclinic	Monoclinic	Orthorhombic	Monoclinic	Monoclinic	Monoclinic
Space group	<i>P</i> $\bar{1}$	<i>P</i> $\bar{1}$	<i>Pc</i>	<i>Cc</i>	<i>C2/c</i>	<i>Pccn</i>	<i>P2₁/n</i>	<i>P2₁/c</i>	<i>P2₁/c</i>
<i>a</i> , Å	11.8378(3)	11.8287(4)	23.2181(5)	22.9200(6)	19.5955(11)	11.8705(5)	13.4984(3)	12.5787(7)	12.5922(3)
<i>b</i> , Å	13.0103(3)	13.0419(4)	11.4301(2)	11.7027(3)	17.1171(10)	31.689(2)	24.7177(5)	13.0731(5)	13.0726(2)
<i>c</i> , Å	16.5916(3)	16.6492(5)	21.4414(4)	21.2051(6)	19.1812(13)	23.8325(9)	13.8468(3)	25.4452(17)	25.4070(7)
α , °	67.7290(19)	67.700(3)	90	90	90	90	90	90	90
β , °	71.634(2)	71.527(3)	96.059(2)	95.783(3)	99.444(6)	90	98.086(2)	91.733(5)	91.730(2)
γ , °	78.1997(19)	78.147(2)	90	90	90	90	90	90	90
<i>V</i> , Å ³	2233.71(9)	2243.27(13)	5658.44(19)	5658.8(3)	6346.5(7)	8964.9(8)	4574.04(17)	4182.4(4)	4180.41(17)
<i>Z</i>	1	1	4	4	8	8	4	4	4
ρ_{calc} [g cm ⁻³]	1.301	1.300	1.266	1.269	1.473	1.086	1.391	1.521	1.527
μ [mm ⁻¹]	0.242	0.271	0.296	0.320	0.322	0.261	0.244	0.996	1.029
<i>F</i> (000)	930	932	2284	2288	2920	3040	1996	1956	1960
<i>T</i> [K]	117	117	100	110	150	293	150	130	130
Max. 2 θ , °	61.016	61.016	58.552	62.054	58.222	52.75	58.108	59.248	61.746
Refins measured	46 042	44 912	71 322	22 958	19 120	46 378	45 629	50 066	71 075
Unique refins	26 916	26 754	27 307	12 469	7 370	8768	11 107	10 563	12 334
Parameter	1163	1183	1360	692	585	496	640	552	552
Restraints	3	699	2	206	346	12	0	0	0
Refins [<i>F</i> _o > 2(<i>F</i> _σ)]	24 196	22 238	20 292	10 847	4546	5782	4654	8487	10 408
<i>R</i> ₁ [<i>F</i> _o > 2σ(<i>F</i> _σ)]	0.0379	0.0428	0.0544	0.0406	0.0613	0.1261	0.0455	0.0485	0.0373
<i>wR</i> ₂ (all data)	0.0963	0.1019	0.1282	0.0898	0.1467	0.3212	0.0814	1.102	1.094
G.O.F.	1.041	1.014	1.026	1.029	0.997	1.064	0.733	0.1006	0.0833
CCDC number	1581401	1581402	1583980	1581403	1581407	1581408	1581410	1581414	1581417

one of two independent Pent_4N^+ cations is disordered in 4 between two positions with the 0.564(7)/0.436(7) occupancies. There are two independent Pent_4N^+ cations in 5. One of the pentyl substituents of the cation is disordered between two positions with the 0.839(4)/0.161(4) occupancies whereas two pentyl substituents of the second cation are disordered between two positions with the 0.681(4)/0.319(4) and 0.546(4)/0.454(4) occupancies. One of the hexyl substituents of Hex_4N^+ is disordered between two positions with the 0.588(5)/0.412(5) occupancies in 7. The MDABCO⁺ cation is disordered in 8 between three orientations with the 0.597(2)/0.246(3)/0.157(3) occupancies. The crystal structure of **9** contains six strongly disordered solvent *o*-C₆H₄Cl₂ molecules per unit cell which were deleted by a standard SQUEEZE procedure. The unit cell contains eight (*i*-Pr₂Im⁺){V^{IV}O(Pc³⁻)⁻} units. Therefore, the corrected composition of **9** should be (*i*-Pr₂Im⁺){V^{IV}O(Pc³⁻)⁻}·0.75C₆H₄Cl₂. One of two *i*-Pr substituents of *i*-Pr₂Im⁺ is disordered between two positions with the 0.522(11)/0.448(11) occupancies. Only major occupied orientations are shown for all disordered components in Fig. 2–5 and 8. To keep the anisotropic thermal parameters of the disordered atoms within reasonable limits the displacement components were restrained using ISOR, SIMU and DELU SHELXL instructions. This resulted in 699, 206, and 346 restraints used for the refinement of the crystal structures of **5**, **7** and **8**, respectively.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The work was supported by the Russian Science Foundation (project N 17-13-01215), the JSPS KAKENHI Grant Number 26288035, and the JST (ACCEL) 27 (100150500010) project.

Notes and references

- C. G. Claessens, W. J. Blau, M. Cook, M. Hanack, R. J. M. Nolte, T. Torres and D. Wöhrle, *Monatsh. Chem.*, 2001, **132**, 3.
- D. Wöhrle, G. Schnurpfeil, S. G. Makarov, A. Kazarin and O. N. Suvorova, *Macroheterocycles*, 2012, **5**, 191.
- J. L. Petersen, C. S. Schramm, D. R. Stojakovic, B. M. Hoffman and T. J. Marks, *J. Am. Chem. Soc.*, 1977, **99**, 286.
- H. Hasegawa, T. Naito, T. Inabe, T. Akutagawa and T. Nakamura, *J. Mater. Chem.*, 1998, **8**, 1567.
- M. Matsuda, T. Naito, T. Inabe, N. Hanasaki, H. Tajima, T. Otsuka, K. Awaga, B. Narymbetov and H. Kobayashi, *J. Mater. Chem.*, 2000, **10**, 631.
- M. Matsuda, T. Naito, T. Inabe, N. Hanasaki and H. Tajima, *J. Mater. Chem.*, 2001, **11**, 2493.
- T. Inabe and H. Tajima, *Chem. Rev.*, 2004, **104**, 5503.
- J. S. Miller, C. Vazquez, J. C. Calabrese, R. S. McLean and A. J. Epstein, *Adv. Mater.*, 1994, **6**, 217.
- D. K. Rittenberg, L. Baars-Hibbe, A. B. Böhm and J. S. Miller, *J. Mater. Chem.*, 2000, **10**, 241.
- E. Tosatti, M. Fabrizio, J. Tóbiik and G. E. Santoro, *Phys. Rev. Lett.*, 2004, **93**, 117002.
- D. V. Konarev, L. V. Zorina, S. S. Khasanov, E. U. Hakimova and R. N. Lyubovskaya, *New J. Chem.*, 2012, **36**, 48.
- D. V. Konarev, A. V. Kuzmin, M. A. Faraonov, M. Ishikawa, Y. Nakano, S. S. Khasanov, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Chem. – Eur. J.*, 2015, **21**, 1014.
- D. V. Konarev, A. V. Kuzmin, Y. Nakano, M. A. Faraonov, S. S. Khasanov, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Inorg. Chem.*, 2016, **55**, 1390.
- (a) A. B. P. Lever, E. P. Milaeva and G. Speier, in *Phthalocyanines, properties and applications*, ed. C. C. Lezuoff and A. B. P. Lever, VCH: NewYork, 1993, vol. 3, pp. 1–70; (b) R. H. Campbell, G. A. Heath, G. T. Hefter and R. C. S. McQueen, *J. Chem. Soc., Chem. Commun.*, 1983, 1123.
- (a) D. W. Clack and J. R. Yandle, *Inorg. Chem.*, 1972, **11**, 1738; (b) A. Louati, M. Elmeray, J. J. Andre, J. Simon, K. M. Kadish, M. Gross and A. Giraudeau, *Inorg. Chem.*, 1985, **24**, 1175.
- M. Tahiri, P. Doppelt, J. Fischer and R. Weiss, *Inorg. Chim. Acta*, 1987, **127**, L1.
- H. Huckstadt and H. Homborg, *Z. Anorg. Allg. Chem.*, 1998, **624**, 715.
- D. V. Konarev, A. V. Kuzmin, S. S. Khasanov and R. N. Lyubovskaya, *Dalton Trans.*, 2013, **42**, 9870.
- D. V. Konarev, S. S. Khasanov, M. Ishikawa, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Inorg. Chem.*, 2013, **52**, 3851.
- D. V. Konarev, A. V. Kuzmin, S. V. Simonov, S. S. Khasanov, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Dalton Trans.*, 2012, **41**, 13841.
- D. V. Konarev, L. V. Zorina, M. Ishikawa, S. S. Khasanov, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Cryst. Growth Des.*, 2013, **13**, 4930.
- D. V. Konarev, A. V. Kuzmin, M. Ishikawa, Y. Nakano, M. A. Faraonov, S. S. Khasanov, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Eur. J. Inorg. Chem.*, 2014, 3863.
- D. V. Konarev, L. V. Zorina, S. S. Khasanov, A. L. Litvinov, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Dalton Trans.*, 2013, **42**, 6810.
- D. V. Konarev, S. S. Khasanov, M. Ishikawa, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Dalton Trans.*, 2017, **46**, 3492.
- D. V. Konarev, A. V. Kuzmin, S. S. Khasanov, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Dalton Trans.*, 2014, **43**, 13061.
- D. V. Konarev, S. S. Khasanov, M. Ishikawa, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Chem. – Asian J.*, 2017, **12**, 910.
- D. V. Konarev, M. A. Faraonov, A. V. Kuzmin, S. S. Khasanov, Y. Nakano, M. S. Batov, S. I. Norko, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *New J. Chem.*, 2017, **41**, 6866.
- D. V. Konarev, S. I. Troyanov, M. Ishikawa, M. Faraonov, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *J. Porphyrins Phthalocyanines*, 2014, **18**, 1157.



- 29 E. W. Y. Wong and D. B. Leznoff, *J. Porphyrins Phthalocyanines*, 2012, **16**, 154.
- 30 E. W. Y. Wong, C. J. Walsby, T. Storr and D. B. Leznoff, *Inorg. Chem.*, 2010, **49**, 3343.
- 31 J. A. Cissell, T. P. Vaid and A. L. Rheingold, *Inorg. Chem.*, 2006, **45**, 2367.
- 32 D. V. Konarev, A. V. Kuzmin, S. S. Khasanov, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Eur. J. Inorg. Chem.*, 2016, 4099.
- 33 J. A. Cissell, T. P. Vaid, A. G. DiPasquale and A. L. Rheingold, *Inorg. Chem.*, 2007, **46**, 7713.
- 34 D. V. Konarev, A. V. Kuzmin, S. S. Khasanov, M. Ishikawa, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Dalton Trans.*, 2016, **45**, 10780.
- 35 D. V. Konarev, S. S. Khasanov, A. V. Kuzmin, Y. Nakano, M. Ishikawa, A. Otsuka, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Cryst. Growth Des.*, 2017, **17**, 753.
- 36 D. V. Konarev, S. S. Khasanov, E. I. Yudanov and R. N. Lyubovskaya, *Eur. J. Inorg. Chem.*, 2011, 816.
- 37 R. O. Loutfy, C. K. Hsiao, B. S. Ong and B. Keoshkerian, *Can. J. Chem.*, 1984, **62**, 1877.
- 38 M. E. Lines, *J. Phys. Chem. Solids*, 1970, **31**, 101.
- 39 Y. Nakano, T. Yagyu, T. Hirayama, A. Ito and K. Tanaka, *Polyhedron*, 2005, **24**, 2141.
- 40 J. S. Smart, in *Magnetism III*, ed. G. T. Rado and H. Suhl, Academic Press, NY, 1963, p. 63.
- 41 (a) D. V. Konarev, S. S. Khasanov, A. Otsuka, M. Maesato, G. Saito and R. N. Lyubovskaya, *Angew. Chem., Int. Ed.*, 2010, **49**, 4829; (b) D. V. Konarev, S. S. Khasanov, A. Otsuka, M. Maesato, M. Uruichi, K. Yakushi, A. Shevchun, H. Yamochi, G. Saito and R. N. Lyubovskaya, *Chem. – Eur. J.*, 2014, **20**, 7268.
- 42 (a) K. Oka, O. Okada and K. Nukada, *Jpn. J. Appl. Phys.*, 1992, **31**, 2181; (b) O. Okada, K. Oka and M. Iijima, *Jpn. J. Appl. Phys.*, 1993, **32**, 3556.
- 43 A. J. Ramadan, L. A. Rochford, D. S. Keeble, P. Sullivan, M. P. Ryan, T. S. Jones and S. Heutz, *J. Mater. Chem. C*, 2015, **3**, 461.
- 44 J. Tóbiš and E. Tosatti, *J. Phys. Chem. A*, 2007, **111**, 12570.
- 45 M. G. Cory, H. Hirose and M. C. Zerner, *Inorg. Chem.*, 1995, **34**, 2969.
- 46 L. Andjelković, S. Stepanović, F. Vlahović, M. Zlatar and M. Gruden, *Phys. Chem. Chem. Phys.*, 2016, **18**, 29122.
- 47 M. A. Faraonov, D. V. Konarev, S. S. Khasanov, S. I. Troyanov and R. N. Lyubovskaya, *Dalton Trans.*, 2017, **46**, 3547.
- 48 L. A. Tomachynsk, V. Y. Chernii and S. V. Volkov, *Russ. J. Inorg. Chem.*, 2002, **47**, 208.
- 49 D. V. Konarev, S. S. Khasanov, A. Otsuka, G. Saito and R. N. Lyubovskaya, *Chem. – Eur. J.*, 2006, **12**, 5225.
- 50 (a) G. M. Sheldrick, *Acta Crystallogr., Sect. A: Found. Crystallogr.*, 2008, **64**, 112; (b) O. V. Dolomanov, L. J. Bourhis, R. J. Gildea, J. A. K. Howard and H. Puschmann, *J. Appl. Crystallogr.*, 2009, **42**, 339.

