

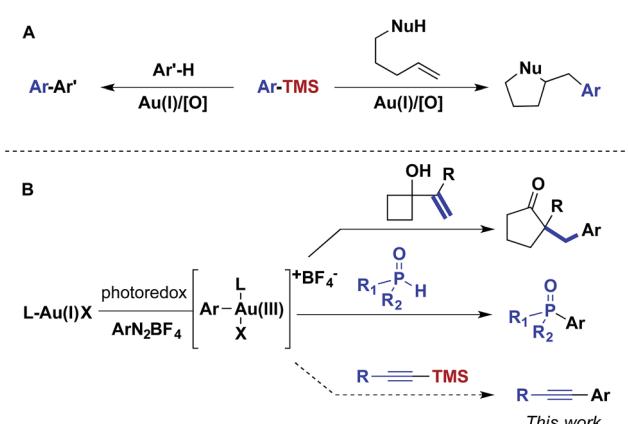
Cite this: *Chem. Sci.*, 2016, **7**, 85Received 14th August 2015  
Accepted 26th October 2015DOI: 10.1039/c5sc03025k  
[www.rsc.org/chemicalscience](http://www.rsc.org/chemicalscience)

## Introduction

For over a decade, homogeneous gold catalysis has been investigated because it provides access to novel modes of reactivity and enables rapid generation of complex molecular architectures.<sup>1</sup> However, in contrast to other low-valent late transition metal catalysts, the majority of gold(i) complexes are unreactive towards the oxidative addition of aryl and vinyl halides and pseudohalides.<sup>2</sup> Consequently, most gold-catalysed cross-coupling reactions have required sacrificial oxidants to access the +3 oxidation state.<sup>3</sup> For example, this reactivity platform has been exploited in a gold-catalysed oxidative coupling of organosilanes and 1,2-alkene functionalizations (Scheme 1A). As an alternative, visible light-mediated oxidative addition of

aryldiazoniums has emerged as a method for the generation of the requisite gold(iii) intermediates *via* photoredox catalysis, thereby obviating the need for sacrificial oxidants.<sup>4</sup> These gold(iii) intermediates have been intercepted by a variety of nucleophiles in dual catalytic photoredox-gold reactions (Scheme 1B).

Although aryldiazoniums and photoredox catalysts provide an efficient method to generate gold(iii) intermediates, several challenges must be addressed in order to facilitate carbon–carbon cross-coupling *via* these species. First, aryldiazonium salts are highly electrophilic and reactive towards many classes of nucleophilic coupling partners. Aryldiazoniums are generally decomposed under Kumada and Negishi coupling conditions and do not tolerate ligand additives and stoichiometric bases employed in the majority of Suzuki, Sonogashira and Hiyama coupling reactions.<sup>5</sup> Second, as visible light-mediated oxidative addition of aryldiazoniums is proposed to proceed through radical intermediates, nucleophilic coupling partners that can participate in single electron transfer, such as organotin and organotrifluoroborates compounds, are problematic.<sup>6</sup> Inspired by the oxidative coupling of organosilanes, we envisioned that the gold(iii) complexes generated *via* photoredox catalysis might undergo transmetalation with organosilanes, generating an intermediate poised for carbon–carbon bond formation through reductive elimination.<sup>7,8</sup>



Scheme 1 Photoredox catalyst and visible light-mediated  $\text{Au}^{\text{I}}/\text{Au}^{\text{III}}$ .

<sup>a</sup>Department of Chemistry, University of California, Berkeley, CA 94720, USA. E-mail: [fdtoste@berkeley.edu](mailto:fdtoste@berkeley.edu)

<sup>b</sup>Departamento de Química Orgánica (Módulo 01), Universidad Autónoma de Madrid, C/Francisco Tomás y Valiente 7, Cantoblanco, 28049 Madrid, Spain

† Electronic supplementary information (ESI) available. See DOI: [10.1039/c5sc03025k](https://doi.org/10.1039/c5sc03025k)

## Results and discussions

While aryldiazonium tetrafluoroborates have been shown to react with halo-, azido-, allyl- and thiotrimethylsilanes, their reactions with alkynyltrimethylsilanes has not been previously reported.<sup>9</sup> Additionally, previous reports of organosilane transmetalation with gold lead us to hypothesize that the transmetalation of alkynyltrimethylsilanes to gold(iii) intermediates might result in productive cross-coupling. On the basis of this hypothesis, we examined the combined gold/photoredox coupling of 1 equivalent of aryldiazonium salt **1a** and 1 equivalent of alkynyltrimethylsilane **2**.

Evaluation of reaction conditions showed that the combination of  $\text{Ar}_3\text{PAuCl}$  and  $\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$ , in acetonitrile gave the highest yield of aryl alkyne **3a** (Table 1). Changing either the solvent or the photoredox catalyst proved detrimental to the yield of the desired product. With respect to the ligands on the gold catalyst, both electronic and steric factors impacted the efficiency of the cross-coupling. For example, electron-deficient triarylphosphine ligands resulted in significantly lower yield of **3a** (Table 1, entry 10). The sterics of the ligand showed an even more dramatic effect on reaction yield. The reaction conducted with  $(p\text{-tol})_3\text{PAuCl}$  provided 72% yield of the desired product, while the much more hindered  $(o\text{-tol})_3\text{PAuCl}$ -catalysed reaction only provided 17% yield of **3a** (Table 1, entry 11 and 12). The yield of product was lower when tricyclohexylphosphinegold(i) chloride was used as catalyst instead of triphenylphosphine-gold(i) chloride (Table 1, entry 13). Gold(i) complexes of dialkylbiarylphosphine were also examined, but all showed less than 5% yield due to the combined effect of dialkyl groups and the steric effect of biaryl groups. Finally, no coupling product was observed in the absence of the ruthenium catalyst (entry 14) or when the reaction was conducted in the dark (entry 15).

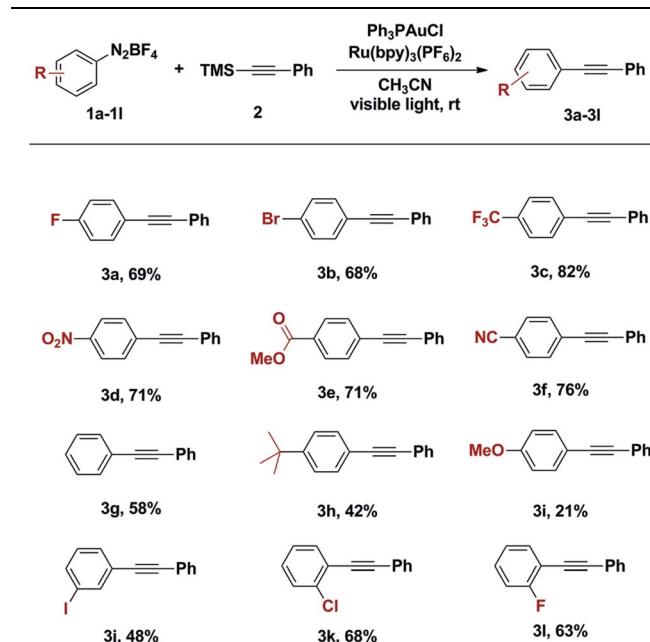
With an optimized catalyst system in hand, the scope with respect to the aryl diazonium coupling partner was examined (Table 2). The yields of alkynylation reactions of electron-poor aryl diazonium salts were generally higher than those of reactions with electron-rich aryl diazonium coupling partners. It should be noted that the intrinsic instability of *ortho*-substituted aryl diazonium salts limited their use in this transformation (Table 2, entry **3k** and **3l**). It is also noteworthy that all halogen substitutions on the aryl diazonium coupling partner

Table 1 Optimization of reaction conditions

Entry	Gold cat.	Photoredox cat.	Solvent	Yield <sup>a</sup>	gold cat. photoredox cat. solvent visible light, rt	
					1a	2
1	$\text{Ph}_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	$\text{CH}_3\text{CN}$	70		
2	$\text{Ph}_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	Acetone	39		
3	$\text{Ph}_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	$\text{CH}_3\text{NO}_2$	15		
4	$\text{Ph}_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	DMF	10		
5	$\text{Ph}_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	MeOH	18		
6	$\text{Ph}_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	EtOH	4		
7	$\text{Ph}_3\text{PAuCl}$	$\text{Ir}(\text{ppy})_3$	$\text{CH}_3\text{CN}$	21		
8	$\text{Ph}_3\text{PAuCl}$	$\text{Ir}(\text{ppy})_2(\text{dtbbpy})\text{PF}_6$	$\text{CH}_3\text{CN}$	45		
9	$(p\text{-MeOPh})_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	$\text{CH}_3\text{CN}$	73		
10	$(p\text{-CF}_3\text{Ph})_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	$\text{CH}_3\text{CN}$	43		
11	$(p\text{-tol})_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	$\text{CH}_3\text{CN}$	72		
12	$(o\text{-tol})_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	$\text{CH}_3\text{CN}$	17		
13	$\text{Cy}_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	$\text{CH}_3\text{CN}$	48		
14	$\text{Ph}_3\text{PAuCl}$	—	$\text{CH}_3\text{CN}$	0		
15 <sup>b</sup>	$\text{Ph}_3\text{PAuCl}$	$\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$	$\text{CH}_3\text{CN}$	0		

Reaction conditions: 4-fluorobenzene diazonium tetrafluoroborate **1a** (0.2 mmol, 1 equiv.), 1-trimethylsilyl-2-phenylacetylene **2** (0.2 mmol, 1 equiv.), gold catalyst (5 mol%), photoredox catalyst (1 mol%), solvent (2 mL),  $\text{N}_2$  atmosphere, visible light, rt for 6 h.<sup>a</sup> GC yields. <sup>b</sup> The reaction was run in the dark. bpy = 2,2'-bipyridine, ppy = 2-phenylpyridine, dtbbpy = 4,4'-di-*tert*-butyl-2,2'-bipyridine.

Table 2 Scope of aryl diazonium tetrafluoroborates



Reaction conditions: **1a-1l** (0.24 mmol, 1.2 equiv.), **2** (0.2 mmol, 1.0 equiv.),  $\text{Ph}_3\text{PAuCl}$  (10 mol%),  $\text{Ru}(\text{bpy})_3(\text{PF}_6)_2$  (2 mol%), acetonitrile (2 mL),  $\text{N}_2$  atmosphere, visible light, rt for 3 h. Isolated yields.

were preserved during the coupling reaction. Bromo- and iodoaryl diazonium salts were readily coupled and leaving halides intact for use in further reactions (Table 2, **3b** and **3j**); this chemoselectivity is challenging under typical palladium-catalysed Sonogashira coupling reactions.<sup>10</sup>

The difference in reactivity of electron-rich and -poor aryl diazoniums may be the result of competing reaction pathways. Oxidative addition is believed to proceed through single electron reduction of aryl diazoniums and this process is preferred with electron-poor aryl diazonium salts.<sup>4</sup> On the other hand, the generation of aryl cations, through loss of dinitrogen, is facile with electron-rich aryl diazoniums. Correspondingly, we observed trace amount of desired product when benzenediazonium tetrafluoroborate was heated at 60 °C for 9 hours in the presence of 1-trimethylsilyl-2-(4-methoxy)phenylacetylene; however the same reactivity was not observed with 4-fluorophenyl diazonium salts on the same conditions.

The scope of alkynyltrimethylsilanes was also investigated (Table 3). Both aryl- and alkylethynyltrimethylsilanes were coupled with modest to high product yields. *o*-Isopropylphenyl, biphenyl and naphthylalkynyltrimethylsilanes participated in the gold-catalysed coupling (Table 3, **3m**, **3n** and **3o**). Potentially sensitive benzylic and propargylic C–H and C–X bonds were well tolerated under the reaction conditions (Table 3, **3q**, **3aa** and **3s–v**). Additionally, alkynyltrimethylsilane **3w** was prepared in 78% yield from the coupling of aryl diazonium salt **1d** and bis(trimethylsilyl)acetylene, providing a compliment to using trimethylsilylacetylene in a traditional Sonogashira coupling (Table 3, **3w**).

Another interesting feature of this reaction is the effect of the silyl group on yield (Table 4). When trimethylsilylacetylene was



Table 3 Scope of alkynyltrimethylsilanes

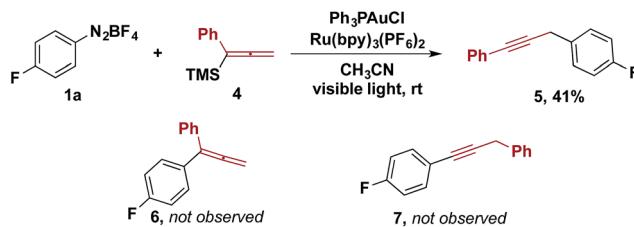

Reaction conditions: **1d** (0.24 mmol, 1.2 equiv.), **2m-aa** (0.2 mmol, 1.0 equiv.), Ph<sub>3</sub>PAuCl (10 mol%), Ru(bpy)<sub>3</sub>(PF<sub>6</sub>)<sub>2</sub> (2 mol%), acetonitrile (2 mL), N<sub>2</sub> atmosphere, visible light, rt for 3 h. Isolated yields. Ar = p-nitrophenyl. <sup>a</sup> 5 equiv. of bis(trimethylsilyl)acetylene was used with 1 equiv. of **1d**. <sup>b</sup> GC yields.

coupled with **1d** to generate terminal alkyne **3x**, the yield was 15% with none of the corresponding silylacetylene **3w** (Table 3). This observation suggests that, under the current base free

Table 4 Effect of counteranion and silyl group

Entry	Counteranion (X)	Silyl group (Y)
1	BF <sub>4</sub> <sup>-</sup>	TMS
2	TsO <sup>-</sup>	TMS
3	PF <sub>6</sub> <sup>-</sup>	TMS
4	BF <sub>4</sub> <sup>-</sup>	TES
5	BF <sub>4</sub> <sup>-</sup>	TBDMS
6	BF <sub>4</sub> <sup>-</sup>	TIPS
7	BF <sub>4</sub> <sup>-</sup>	TBDPS
8	BF <sub>4</sub> <sup>-</sup>	H

Reaction conditions: aryl diazonium salt (0.2 mmol, 1 equiv.), phenylethynylsilane (0.2 mmol, 1 equiv.), Ph<sub>3</sub>PAuCl (5 mol%), Ru(bpy)<sub>3</sub>(PF<sub>6</sub>)<sub>2</sub> (1 mol%), acetonitrile (2 mL), N<sub>2</sub> atmosphere, visible light, rt for 6 h. <sup>a</sup> GC yields.



Scheme 2 Reaction with allenyltrimethylsilanes.

reaction conditions, the silyl group plays a critical role in the coupling. To investigate this role, a set of reactions was performed by varying counteranions of aryl diazonium salts and silyl group identities of alkynylsilanes. The yield of coupling product was diminished as the steric hindrance of silyl group was increased (Table 4, entry 1 and entries 4–7). While both tetrafluoroborate and tosylate diazonium salts proved efficient coupling partners, the yield was diminished when hexafluorophosphate salts were used.<sup>11</sup> Moreover, when the corresponding terminal alkyne was used instead of alkynyltrimethylsilanes, only 3% of the coupling product was observed (Table 4, entry 8).<sup>12</sup> Taken together, these results suggests that transmetalation from the organosilane is critical for high efficiency of the current coupling reaction.

Other organotrimethylsilanes were tested under the optimized reaction conditions. While the reactions of aryl and vinylsilanes were complicated by competing reactions,<sup>13,14</sup> the gold/photoredox-catalysed coupling of allenyltrimethylsilane **4** provided propargylic compound **5** as the sole coupling product (Scheme 2). The product from aryl-allenyl reductive elimination and the other aryl-propargylic coupling isomer were prepared independently (Scheme 2, **6** and **7**). We examined whether **6** was an intermediate that underwent isomerization to **5** facilitated by visible light and the photoredox catalyst. However, no isomerization was observed of **6** to either **5** or **7** under the reaction conditions. Therefore, the formation of **5** is best rationalized by transmetalation of **4** to the gold(III) intermediate, followed by 1,3-migration and aryl-propargyl reductive elimination.

## Conclusions

In summary, we have demonstrated that trimethylsilylalkynes and aryl diazonium tetrafluoroborates can be coupled *via* dual photoredox and gold catalysis strategy. The reaction proceeds under mild conditions and shows excellent functional group tolerance, including aryl halides that may be reactive under traditional coupling conditions. This process complements the Sonogashira coupling, especially when aryl-alkynyl coupling is required under non-basic conditions and may present an advantage when TMS-alkynes are available rather than corresponding terminal alkynes.

## Acknowledgements

We are grateful for financial support from the NIHGMs (RO1 GM073932). Jaime Rojas-Martin thanks MINECO for the EBB-



FPI fellowship. We thank Mark D. Levin, William J. Wolf and Matthew S. Winston for helpful discussions.

## Notes and references

- For recent reviews on gold catalysis: (a) A. S. K. Hashmi, *Acc. Chem. Res.*, 2014, **47**, 864; (b) C. Obradors and A. M. Echavarren, *Chem. Commun.*, 2014, **50**, 16; (c) D. Garayalde and C. Nevado, *ACS Catal.*, 2012, **2**, 1462; (d) M. Rudolph and A. S. K. Hashmi, *Chem. Soc. Rev.*, 2012, **41**, 2448; (e) L. P. Liu and G. B. Hammond, *Chem. Soc. Rev.*, 2012, **41**, 3129; (f) N. D. Shapiro and F. D. Toste, *Synlett*, 2010, 675.
- (a) T. Lauterbach, M. Livendahl, A. Rosellón, P. Espinet and A. M. Echavarren, *Org. Lett.*, 2010, **12**, 3006. For recent breakthroughs, (b) M. D. Levin and F. D. Toste, *Angew. Chem., Int. Ed.*, 2014, **53**, 6211; (c) M. S. Winston, W. J. Wolf and F. D. Toste, *J. Am. Chem. Soc.*, 2014, **136**, 7777; (d) J. Guenther, S. Mallet-Ladeira, L. Estevez, K. Miqueu, A. Amgoune and D. Bourissou, *J. Am. Chem. Soc.*, 2014, **136**, 1778; (e) M. Joost, A. Zeineddine, L. Estevez, S. Mallet-Ladeira, K. Miqueu, A. Amgoune and D. Bourissou, *J. Am. Chem. Soc.*, 2014, **136**, 14654.
- For early publications on gold-catalysed oxidative coupling, (a) A. Kar, N. Mangu, H. M. Kaiser, M. Beller and M. K. Tse, *Chem. Commun.*, 2008, 386; (b) G. Zhang, Y. Peng, L. Cui and L. Zhang, *Angew. Chem., Int. Ed.*, 2009, **48**, 3112; (c) A. D. Melhado, W. E. Brenzovich, A. D. Lackner and F. D. Toste, *J. Am. Chem. Soc.*, 2010, **132**, 8885; (d) W. E. Brenzovich, D. Benitez, A. D. Lackner, H. P. Shunatona, E. Tkatchouk, W. A. Goddard and F. D. Toste, *Angew. Chem., Int. Ed.*, 2010, **49**, 5519; (e) W. E. Brenzovich, J. F. Brazeau and F. D. Toste, *Org. Lett.*, 2010, **12**, 4728; (f) G. Zhang, L. Cui, Y. Wang and L. Zhang, *J. Am. Chem. Soc.*, 2010, **132**, 1474; (g) T. de Haro and C. Nevado, *J. Am. Chem. Soc.*, 2010, **132**, 1512; (h) J. P. Brand, J. Charpentier and J. Waser, *Angew. Chem., Int. Ed.*, 2009, **48**, 9346; (i) J. P. Brand and J. Waser, *Angew. Chem., Int. Ed.*, 2010, **49**, 7304; (j) M. N. Hopkinson, A. Tessier, A. Salisbury, G. T. Giuffredi, L. E. Combettes, A. D. Gee and V. Gouverneur, *Chem.-Eur. J.*, 2010, **16**, 4739; (k) L. T. Ball, G. C. Lloyd-Jones and C. A. Russell, *Org. Lett.*, 2010, **12**, 4724; (l) L. T. Ball, G. C. Lloyd-Jones and C. A. Russell, *Science*, 2012, **337**, 1644.
- For visible-light mediated oxidative addition of aryl diazoniums on gold catalysis, (a) B. Sahoo, M. N. Hopkinson and F. Glorius, *J. Am. Chem. Soc.*, 2013, **135**, 5505; (b) M. N. Hopkinson, B. Sahoo and F. Glorius, *Adv. Synth. Catal.*, 2014, **356**, 2794; (c) X. Z. Shu, M. Zhang, Y. He, H. Frei and F. D. Toste, *J. Am. Chem. Soc.*, 2014, **136**, 5844; (d) Y. He, H. Wu and F. D. Toste, *Chem. Sci.*, 2015, **6**, 1194.
- A. Roglans, A. Pla-Quintana and M. Moreno-Manas, *Chem. Rev.*, 2006, **106**, 4622.
- For single-electron transmetalation of organotrifluoroborates (a) J. C. Tellis, D. N. Primer and G. A. Molander, *Science*, 2014, **345**, 433; (b) D. N. Primer, I. Karakaya, J. C. Tellis and G. A. Molander, *J. Am. Chem. Soc.*, 2015, **137**, 2195.
- W. J. Wolf, M. S. Winston and F. D. Toste, *Nat. Chem.*, 2013, **6**, 159.
- For cross-coupling reactions of alkynyltrimethylsilanes, (a) S. A. Lermontov, I. M. Rakov, N. S. Zefirov and P. J. Stang, *Tetrahedron Lett.*, 1996, **37**, 4051; (b) G. W. Kabalka, L. Wang and R. M. Pagni, *Tetrahedron*, 2001, **57**, 8017; (c) S. Kang, H. Ryu and Y. Hong, *J. Chem. Soc., Perkin Trans. 1*, 2007, **736**; (d) C. Yang and S. P. Nolan, *Organometallics*, 2002, **21**, 1020; (e) L. K. Hwang, Y. Na, J. Lee, Y. Do and S. Chang, *Angew. Chem., Int. Ed.*, 2005, **44**, 6166; (f) W. J. Sommer and M. Weck, *Adv. Synth. Catal.*, 2006, **348**, 2101; (g) N. Sakai, R. Komatsu, N. Uchida, R. Ikeda and T. Konakahara, *Org. Lett.*, 2010, **12**, 1300; (h) Y. Nishihara, S. Noyori, T. Okamoto, M. Suetsugu and M. Iwasaki, *Chem. Lett.*, 2011, **40**, 972; (i) Y. Nishihara, D. Ogawa, S. Noyori and M. Iwasaki, *Chem. Lett.*, 2012, **41**, 1503; (j) L. Peng, F. Xu, Y. Suzuma, A. Orita and J. Otera, *J. Org. Chem.*, 2013, **78**, 12802.
- For reactions of aryl diazonium tetrafluoroborate with halo-, azido-, allyl- and thiotrimethylsilanes, (a) T. Keumi, T. Umeda, Y. Inoue and H. Kitajima, *Bull. Chem. Soc. Jpn.*, 1989, **62**, 89; (b) H. Mayr and K. Grimm, *J. Org. Chem.*, 1992, **57**, 1057; (c) G. K. S. Prakash, D. Hoole, D. S. Ha, J. Wilkinson and G. A. Olah, *ARKIVOC*, 2002, **13**, 50.
- For Sonogashira cross-coupling of aryl diazonium salts, (a) S. Darses, G. Michaud and J. Genet, *Eur. J. Org. Chem.*, 1999, 1875; (b) G. Fabrizi, A. Goggiamani, A. Sferrazza and S. Cacchi, *Angew. Chem., Int. Ed.*, 2010, **49**, 4067; (c) B. Panda and T. K. Sarkar, *Chem. Commun.*, 2010, **46**, 3131; (d) X. Wu, H. Neumann and M. Beller, *Chem. Commun.*, 2011, **47**, 7959; (e) This challenge was overcome by employing *o*-benzenedisulfonimide as a counteranion. ; M. Barbero, S. Cadamuro and S. Dughera, *Eur. J. Org. Chem.*, 2014, 598. (f) While preparing a manuscript, gold-catalysed cross-coupling of aryl diazonium salts with terminal alkynes and boronic acids was published. ; R. Cai, M. Lu, E. Y. Aguilera, Y. Xi, N. G. Akhmedov, J. L. Petersen, H. Chen and X. Shi, *Angew. Chem., Int. Ed.*, 2015, **54**, 8772.
- After the completion of coupling reaction with tetrafluoroborate salts, stoichiometric amounts of trimethylsilyl fluoride were observed by <sup>1</sup>H NMR.
- To examine the robustness of this reaction towards terminal alkynes, 1 equiv. of phenylacetylene and 1 equiv. of phenylethynyltrimethylsilane were reacted with 1 equiv. **1a** surprisingly, less than 5% of **3a** was detected.
- When (Z)-styryltrimethylsilane was used the corresponding Z-product was obtained from gold-catalysed cross coupling and E-product was obtained from uncatalysed addition of the aryl diazonium to the vinylsilane (see: A. Benkeser, E. W. Bennet and R. A. Hickner, *J. Am. Chem. Soc.*, 1957, **79**, 6253.) After consumption of the starting material, photoredox promoted isomerization of E-stilbene to Z-stilbene occurred.
- In the case aryltrimethylsilanes, further optimization is required due to the formation of significant amounts of competing homo-coupling reactions.

