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Water effect on the spin-transition behavior of Fe(II) 1,2,4-triazole 1D chains embedded in pores of MCM-41†

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The spin-crossover (SCO) compounds [Fe(Htrz)₃](BF₄)₂·H₂O (SCO-1) and [Fe(Htrz)₂trz]BF₄ (SCO-2) (Htrz = 1,2,4-triazole) were embedded in the pores of mesostructured silica MCM-41 to yield SCO@MCM composites as evidenced by electron microscopy, gas sorption studies, powder X-ray diffractometry, atomic absorption and infrared spectrometry. Studies of the temperature-induced spin crossover behavior of the composites by temperature-variable ⁵⁷Fe Mössbauer spectroscopy, magnetic and differential scanning calorimetry measurements and optical reflectivity indicate that the spin transition of the composites was significantly shifted for SCO-1@MCM to higher temperature in comparison to bulk SCO-1 compounds while the shift for SCO-2 was negligible. These shifts in the transition temperature for SCO-1@MCM [*versus* bulk SCO-1] amounted to $T_c^\uparrow = 371/376$ K [282/291 K] and $T_c^\downarrow = 340/345$ K [276/286 K] (magnetic/optical reflectivity data) with a broadening of the hysteresis by 25–26 K relative to bulk SCO-1 (varying slightly with the used method). The significant difference in the SCO behavior of the similar materials SCO-1 and SCO-2 when embedded in the MCM-41 matrix is assigned to the hydration of the SCO-1@MCM material. Water is apparently crucial in transmitting the confinement pressure or matrix effect on the spin transition when the SCO compound is embedded between the pore walls.

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Introduction

The spin crossover (SCO) phenomenon is the switching of a transition metal complex between two different stable ground states: the low-spin (LS) and the high-spin (HS) states.^{1,2} The switching between these two states may be invoked by different external stimuli such as light, temperature or pressure.^{3–5} Iron(II) SCO complexes typically based on N-donor azole/azolate or azine ligands not only lead to a change in magnetic moment but usually give rise to optical transitions from purple in the LS state (lower temperature) to off-white in the HS state when increasing the temperature (*cf.* Fig. 1).^{6,7} One interesting aspect in this area concerns the case of SCO hybrid materials where a given complex can feel the constraint of a given matrix which is equivalent to the application of additional pressure.^{8–10} Several SCO nanoparticles based on Fe(II) complexes

were prepared and embedded in matrices, such as SiO₂,¹¹ Ni,¹² and CsFe[Cr(CN)₆], to name a few.¹³ Another manifestation of indirect pressure effects could also be found in SCO nanoparticles for which theoretical descriptions were recently proposed.^{14,15}

In this work, two iron(II) SCO compounds were embedded in mesostructured silica MCM-41. The selected materials are the 1D coordination polymers [Fe(Htrz)₃](BF₄)₂·H₂O (SCO-1) and [Fe(Htrz)₂trz]BF₄ (SCO-2) with Htrz = 4H-1,2,4-triazole and trz = 1,2,4-triazolato (Scheme 1).^{16–18} Both are known to exhibit a hysteretic SCO behavior with sharp spin transition (ST) above room temperature. SCO-2 which displays a reproducible hysteresis loop of $T_c^\uparrow = 385$ K and $T_c^\downarrow = 345$ K¹⁷ has been selected for numerous studies, because it does not contain any water molecules which can give rise to unstable hysteresis loops due to solvent release. In particular, hybrid materials including SCO-2 were recently prepared with a mesoporous silica monolith,¹⁹ graphene,²⁰ and with silica-gold nanocomposites.²¹ Yet, the matrix effect with MCM-41 has not been investigated although this material is well known for its potential to accommodate various sorts of guests.²² The MCM-41 matrix effect, that is, the change in SCO behavior under the constraints of surrounding silicate walls has been studied herein by temperature-variable ⁵⁷Fe Mössbauer spectroscopy, differential scanning calorimetry, magnetic measurements and optical reflectivity.

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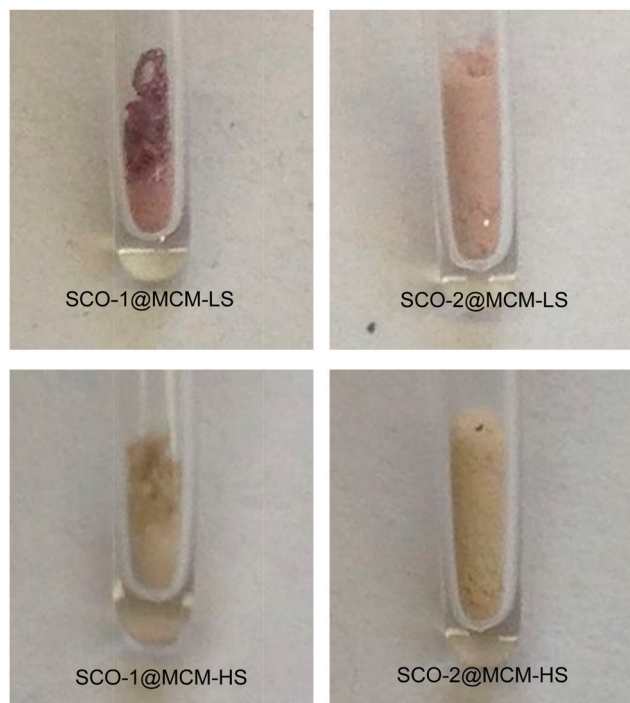
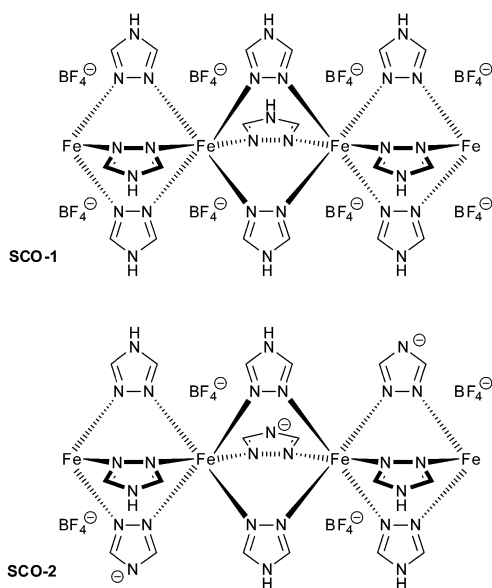


Fig. 1 Photographs showing the color change from LS (room temperature) to HS ($\sim 100^\circ\text{C}$) for each sample.

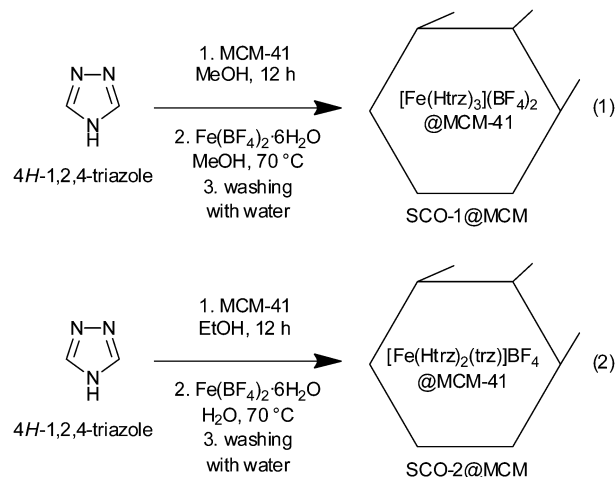


Scheme 1 Schematic drawings of 1D SCO materials to be inserted into MCM-41.

Results and discussion

Synthesis

Fe(II) 1,2,4-triazole compounds were embedded in MCM-41 host matrices by first impregnating the MCM-41 material with an alcoholic solution of the 1,2,4-triazole ligand over a time of at least 12 h. Then an alcoholic or an aqueous solution of the appropriate amount of the iron(II) salt $\text{Fe}(\text{BF}_4)_2 \cdot 6\text{H}_2\text{O}$, was added



Scheme 2 Preparation of SCO hybrid materials used in this work.

to form the composite SCO@MCM materials (eqn (1) and (2), respectively). For the subsequent investigations it was crucial to have the SCO material only inside the cylindrical pores of MCM-41. So, careful washing with water was done to remove any SCO precipitate outside of the pores. Both compounds SCO-1 and -2 readily dissolve in water with plausible fragmentation of the 1D chains (according to ref. 17 SCO-2 dissolves with decomposition) (Scheme 2).

The IR spectra of the composite materials SCO@MCM compared to the IR spectra of bulk SCO and MCM-41 show the presence of both components in the composite (Fig. 2).

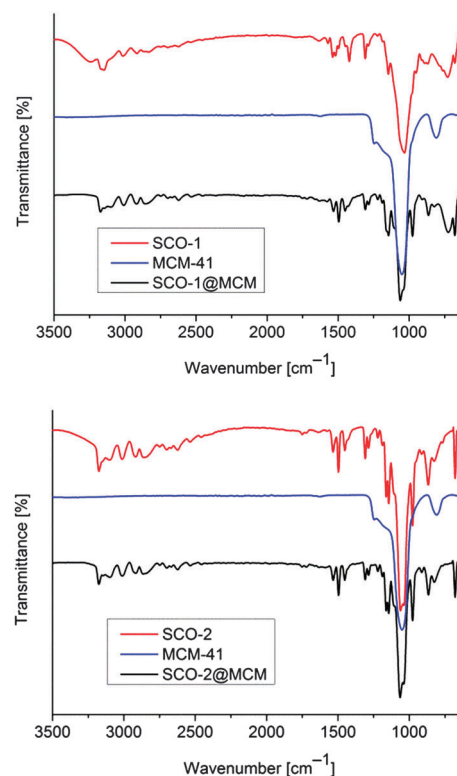


Fig. 2 IR spectra of SCO@MCM samples (black) in comparison to bulk SCO (red) and MCM-41 (blue).



The SCO amount in the composite was calculated from the iron analysis by atomic absorption spectroscopy to be 56.4 wt% for SCO-1@MCM and 43.8 wt% for SCO-2@MCM (see Table 3 for details).

Matching powder X-ray diffractograms (PXRD, Fig. 3) of bulk SCO and the SCO@MCM materials confirm the identity of the SCO materials formed in bulk and as a SCO@MCM composite for SCO-1 and SCO-2.

Porosity measurements

IR spectroscopy, elemental analysis and PXRD cannot distinguish between bulk and embedded SCO material and, hence, do not prove the presence of SCO inside the pores of MCM. However, with the pores of MCM filled by the SCO compound the porosity should decrease. The remaining porosity of the SCO@MCM composite materials was analyzed by N₂ sorption studies at 77 K (Fig. 4). The sample was degassed in vacuum before measurement at a temperature of 393 K for 2 h. In the composite materials the Brunauer–Emmett–Teller (BET) and Langmuir surface areas for SCO-1 have decreased considerably to less than 25% of the value found in MCM-41 (Table 1).

The adsorption isotherm of SCO-1@MCM is of type-IV, which is typical for many mesoporous adsorbents,^{23–25} with an H2 hysteresis loop which may be associated with pores with narrow necks and wide bodies (“bottle-neck” pores).²⁶ Such pore structures can result from the formation of SCO-1 microcrystals inside the MCM mesopores. The N₂ sorption isotherm for MCM-41 with

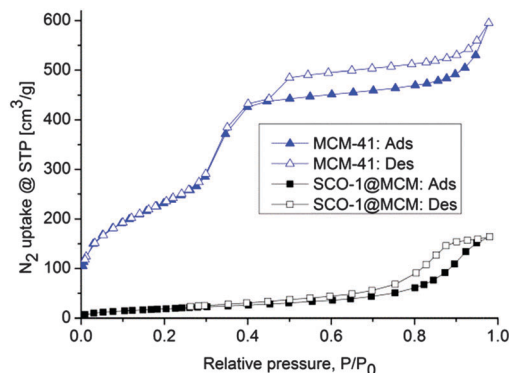


Fig. 4 N₂ sorption isotherms of MCM-41 and SCO@MCM samples; closed and open symbols refer to adsorption and desorption, respectively.

Table 1 Porosity data for SCO@MCM from N₂ isotherms at 77 K

Composite	S_{BET}^a (m ² g ⁻¹)	S_{Lang}^b (m ² g ⁻¹)	$V_{0.1}^c$ (cm ³ g ⁻¹)	V_{tot}^d (cm ³ g ⁻¹)	$V_{0.1}/V_{\text{tot}}^e$	V_{micro}^e (CO ₂) (cm ³ g ⁻¹)
MCM-41	857	1238	0.297	0.82	0.36	0.0175
SCO-1@MCM	70	104	0.0222	0.235	0.09	0.002
SCO-2@MCM	0	0	n.a.	n.a.	n.a.	0.002

^a BET surface area calculated at $0.05 < p/p_0 < 0.2$ from N₂ sorption isotherm at 77 K with a standard deviation of ± 50 m² g⁻¹. ^b Langmuir surface area over the pressure range 0–110 Torr. ^c Micropore volume calculated from N₂ adsorption isotherm at $p/p_0 = 0.1$ for pores with $d \leq 2$ nm (20 Å). ^d Calculated from N₂ sorption isotherm at 77 K ($p/p_0 = 0.95$) for pores ≤ 20 nm. ^e Pore volume for pores with $d \leq 1$ nm (10 Å, cf. Fig. S3, ESI) from the CO₂ NL-DFT model at 273 K.

Table 2 Summary of transition temperatures (T_c in K)

Material	⁵⁷ Fe Mössbauer		Magnetic		Optical reflect.		DS calorimetry		Ref.
	$T_c^{\uparrow}$	$T_c^{\downarrow}$	$T_c^{\uparrow}$	$T_c^{\downarrow}$	$T_c^{\uparrow}$	$T_c^{\downarrow}$	$T_c^{\uparrow}$	$T_c^{\downarrow}$	
Bulk SCO-1	n.a. ^a	n.a. ^a	345 ^b	323 ^b	336 ^b	323 ^b			17
SCO-1@MCM	330	310	282 ^c	276 ^c	291 ^c	286 ^c			This work
			389 ^d	340 ^d	394 ^d	345 ^d			
Bulk SCO-2	380	344	371 ^e	340 ^e	376 ^e	345 ^e	383 ^g	343 ^g	17
					373 ^f	343 ^f	387 ^h	343 ^h	This work
SCO-2@MCM	342	328	381	346	385	351	390 ^h	355 ^h	This work

^a n.a. = not available. ^b Hydrated material, [Fe(Htrz)₃](BF₄)₂·H₂O. ^c Dehydrated material. ^d First heating and cooling cycle on hydrated material. ^e Second to fourth cycle on dehydrated material. ^f Optical reflectance measurements were recorded in this work on SCO-2 prepared as nanoparticles in the liquid state according to ref. 40. ^g At 7 K min⁻¹. ^h At 10 K min⁻¹.

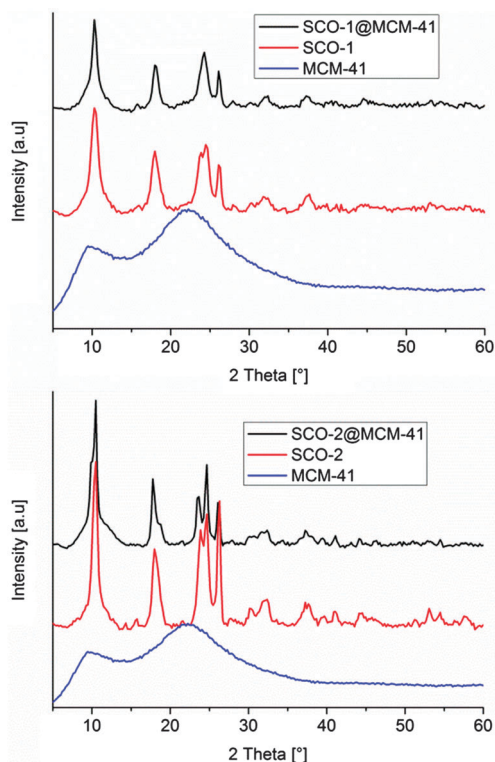


Fig. 3 Powder X-ray diffractograms of bulk SCO complexes, and SCO@MCM composite materials (background corrected). The MCM host matrix is amorphous.

its S-shape matches the literature reports.^{27–30} In summary, the decrease of surface area and porosity is a good indication that the MCM-41 matrix was filled by SCO materials as intended.

No observable surface area was found by N₂ sorption for SCO-2@MCM. However, at the cryogenic temperature of 77 K diffusion of N₂ molecules into micropores is very slow. Diffusion limitations at this temperature influences adsorption in ultramicropores (pores smaller than 7 Å).³¹ For porous materials which usually contain a wide range of pore sizes including ultramicropores, this would require time-consuming N₂ adsorption measurements and



Table 3 AAS analysis results (Fe content)^a

Sample	Fe content (wt%)	SCO content (wt%)
SCO-1@MCM	7.21	56.4
SCO-2@MCM	7.01	43.8

^a See Tables S2–S4 in ESI for details.

may still lead to under-equilibration of the adsorption isotherms, hence, will give erroneous results of the pore volume and pore size distribution analysis. Problems of this type can be eliminated by using CO₂ adsorption analysis at 273 K.³² The saturation pressure of CO₂ at 0 °C is very high (~26 141 Torr), therefore low relative pressure measurements necessary for micropore analysis are achieved in the range of moderate absolute pressures (1–760 Torr).³³ At 273 K and under higher absolute pressures CO₂ molecules can more easily access ultramicropores than N₂ at ~77 K and the kinetic diameter of CO₂ (3.3 Å) is also slightly smaller than for N₂ (3.64 Å). CO₂ micropore analysis at 273 K *versus* N₂ analysis at 77 K is advantageous because of (i) faster analysis and (ii) greater confidence that measured adsorption points are equilibrated (both due to higher diffusion rates) and (iii) extension of the range of analysis to pores of smaller sizes that are accessible to CO₂ molecules but not to N₂.³⁴

CO₂ sorption at 273 K shows that SCO-2@MCM has comparable porosity to SCO-1@MCM (Fig. S2 in ESI†). This indicates that pores or pore mouths of SCO-2@MCM are narrower than that of SCO-1@MCM. Thus, from CO₂ adsorption isotherms at 273 K (Fig. S2, ESI†), the pore size distribution (PSD) for SCO@MCM was derived between 4–10 Å by using nonlocal density functional theory (NLDFT) with a “slit-pore model” (Fig. S3, ESI†). CO₂ adsorption with the NLDFT model yields a better resolved PSD towards the ultramicropore end than from N₂ adsorption isotherms for SCO-1@MCM (Fig. S3 to S5 in ESI†). Here the two SCO@MCM materials give very similar pore size distributions for pores below 10 Å (1 nm).

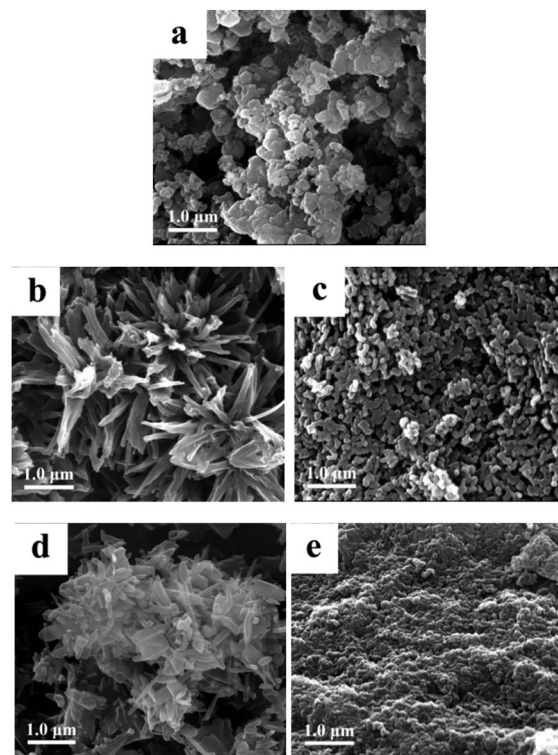
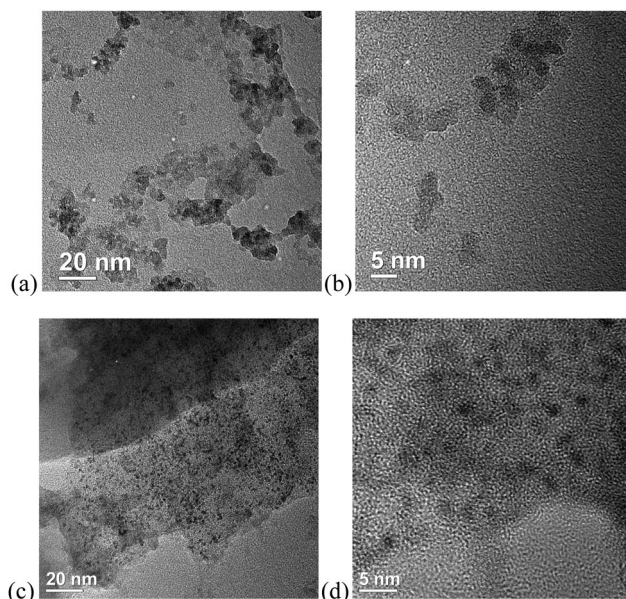
Electron microscopy

Scanning electron microscopy (SEM) images (Fig. 5) show a morphology of the SCO@MCM composite materials which is similar to native MCM. No characteristic SCO crystallites can be seen. This supports the formation of SCO inside the MCM mesopores and verifies the removal of any SCO formed outside through the washing procedures.

Transmission electron microscopy (TEM) images reveal the formation of nanocrystals of SCO-1 and -2 inside the MCM matrix (Fig. 6). The dark spots which are ~2 nm in diameter contain iron and at high resolution (5 nm gauge bar, Fig. 6b and d) it is visible that these dark spots consist of lattices planes which are indicative of crystalline material. The crystallinity of the SCO materials inside MCM is in agreement with the observation of PXRD patterns for the SCO@MCM composites (see Fig. 3).

Spin transition behavior

In this work, the ST behavior of the composites was investigated and compared to those of the guest materials using the same set of physical techniques. Such a careful comparison is important

**Fig. 5** SEM images of (a) MCM-41, (b) SCO-1, (c) SCO-1@MCM, (d) SCO-2 and (e) SCO-2@MCM.**Fig. 6** TEM images of SCO-1@MCM (a, b) and SCO-2@MCM (c, d) at different magnifications.

because of the diversity of techniques presented which can provide a different set of transition temperatures. The established SCO research practices foresee the use of several methods of investigation in order to cover the spread of ST parameters as found by different techniques. Bulk techniques: Mössbauer



measurements (run in static mode with long acquisition times) and DC magnetic measurements (in settle mode) and a surface technique: optical reflectivity (in scanning mode). Mössbauer spectroscopy is the only method which provides a careful evaluation of the spin population given that by reflectivity, only the surface of the sample is probed.

⁵⁷Fe Mössbauer studies

SCO-1@MCM was investigated over the temperature range 78–358 K, first at 78 K and then on warming up to 358 K followed by a cooling cycle to 298 K. The associated Mössbauer parameters are gathered in Table 4. At 78 K, the spectrum of SCO-1@MCM consists of one quadrupole doublet, with an isomer shift of $\delta^{\text{LS}} = 0.43(1) \text{ mm s}^{-1}$ and a quadrupole splitting of $\Delta E_{\text{Q}}^{\text{LS}} = 0.26(2) \text{ mm s}^{-1}$ characteristic for the LS state of iron(II) ions (Table 4). The composite remains in the LS state on warming

up to 328 K (Fig. 7), showing no influence of the matrix on the spin state. SCO-1 was only partly investigated by Lavrenova *et al.* by Mössbauer spectroscopy.¹⁶ Nevertheless, the isomer shift δ^{LS} of SCO-1@MCM at room temperature nicely corresponds to the one reported for SCO-1 ($\delta = 0.43 \text{ mm s}^{-1}$),¹⁶ indicating that the coordination polymer structure is not affected by the matrix environment, *i.e.*, the 1D chain is not located nearby the walls of the MCM-41. No decomposition of the sample is noticed within the composite as seen by the absence of oxidation products, which was however clearly observed when studying a less stable derivative, $[\text{Fe}(\text{NH}_2\text{trz})_3](\text{NO}_3)_2$ ($\text{NH}_2\text{trz} = 4\text{-amino-1,2,4-triazole}$) within the same matrix. The latter material was synthesized and processed similarly to SCO-1,2@MCM. On warming to 333 K, a major signal is detected for the SCO-1@MCM composite with larger parameters ($\delta^{\text{HS}} = 0.97(1) \text{ mm s}^{-1}$ and $\Delta E_{\text{Q}}^{\text{HS}} = 2.30(1) \text{ mm s}^{-1}$) which are characteristic for the HS state of iron(II) ions (Fig. 7). The major signal is present in 80% population compared to the LS state. On warming further to 358 K, this population grows up to 87% indicating an incomplete spin transition. An asymmetric quadrupole doublet is noticed for the HS state at this temperature which is attributed to a texture effect. On cooling, a slight decrease of the HS population is observed down to 318 K after which a sharp transition to the LS state is observed. Interestingly, the ST is also incomplete at room temperature with 5% of ions remaining in the HS state. The temperature dependence of the HS molar fraction assuming equal Debye–Waller factors for HS and LS states, which is justified taking into account the sharpness of the ST,³⁵ is shown in Fig. 8. A hysteresis loop of 20 K width is delineated with transition temperatures $T_c^\uparrow = 330 \text{ K}$ and $T_c^\downarrow = 310 \text{ K}$. The ST is not complete in both HS and LS states.

SCO-2@MCM was investigated over the range 78–358 K by Mössbauer spectroscopy, first at 78 K and then on warming

Table 4 Overview of ⁵⁷Fe Mössbauer parameters for SCO-1@MCM^a

T [K]	δ [mm s ⁻¹]	ΔE_{Q} [mm s ⁻¹]	$\Gamma/2$ [mm s ⁻¹]	Relative area [%]	Sites
78	0.43(1)	0.26(2)	0.13(1)	100	LS
298 ↑	0.42(1)	0.26(1)	0.15(1)	100	LS
303 ↑	0.41(1)	0.26(1)	0.16(1)	100	LS
308 ↑	0.40(1)	0.26(1)	0.16(1)	100	LS
313 ↑	0.40(1)	0.26(1)	0.16(1)	100	LS
318 ↑	0.39(1)	0.25(1)	0.17(1)	100	LS
323 ↑	0.38(1)	0.26(1)	0.17(1)	100	LS
328 ↑	0.38(1)	0.26(1)	0.18(1)	100	LS
333 ↑	0.97(1)	2.30(1)	0.13(1)	80.0	HS
	0.46(1)	0.28(1)	0.29(1)	20.0	LS
338 ↑	0.96(1)	2.25(1)	0.13(1)	83.0	HS
	0.45(1)	0.28(1)	0.26(1)	17.0	LS
343 ↑	0.95(1)	2.21(1)	0.13(1)	83.0	HS
	0.45(1)	0.34(1)	0.31 ^b	17.0	LS
348 ↑	0.94(1)	2.18(1)	0.14(1)	85.0	HS
	0.44(1)	0.34 ^b	0.31 ^b	15.0	LS
353 ↑	0.93(1)	2.15(1)	0.13(1)	87.0	HS
	0.43 ^b	0.36(2)	0.28(1)	13.0	LS
358	0.93(1)	2.10(1)	0.13(1)	87.0	HS
	0.43(1)	0.38 ^b	0.26 ^b	13.0	LS
353 ↓	0.94(1)	2.14(1)	0.14(1)	87.0	HS
	0.43(1)	0.35 ^b	0.32(1)	13.0	LS
348 ↓	0.94(1)	2.13(1)	0.14(1)	87.0	HS
	0.44(1)	0.38(1)	0.33(1)	13.0	LS
343 ↓	0.95(1)	2.21(1)	0.13(1)	83.0	HS
	0.45 ^b	0.34 ^b	0.34(1)	17.0	LS
338 ↓	0.96(1)	2.25(1)	0.14(1)	83.0	HS
	0.45(1)	0.32 ^b	0.46(1)	17.0	LS
333 ↓	0.97(1)	2.29(1)	0.16(1)	82.0	HS
	0.46(1)	0.35 ^b	0.44(1)	18.0	LS
328 ↓	0.98(1)	2.32(1)	0.14(1)	82.0	HS
	0.47(1)	0.30 ^b	0.45(1)	18.0	LS
318 ↓	0.99(1)	2.39(1)	0.14(1)	81.0	HS
	0.49(1)	0.35 ^b	0.31(1)	19.0	LS
313 ↓	0.99(1)	2.42(1)	0.14(1)	75.0	HS
	0.54(1)	0.47(1)	0.26(1)	25.0	LS
308 ↓	0.99(1)	2.42(1)	0.15(1)	5.0	HS
	0.50(1)	0.31(1)	0.16(1)	95.0	LS
303 ↓	0.99(1)	2.40(1)	0.14(1)	5.0	HS
	0.50(1)	0.30(1)	0.16(1)	95.0	LS
298 ↓	1.07(1)	2.63(1)	0.16(1)	5.0	HS
	0.49(1)	0.30(1)	0.16(1)	95.0	LS

^a ↑ indicates warming and ↓ indicates cooling; δ = isomer shift relative to α -iron, ΔE_{Q} = quadrupole splitting, $\Gamma/2$ = half width at half maximum.

^b Fixed parameter.

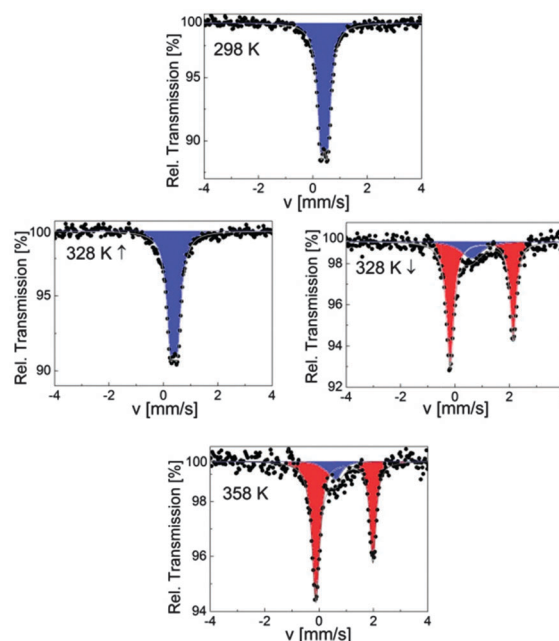


Fig. 7 Selected ⁵⁷Fe Mössbauer spectra of SCO-1@MCM; blue and red colors correspond to LS and HS doublets, respectively.



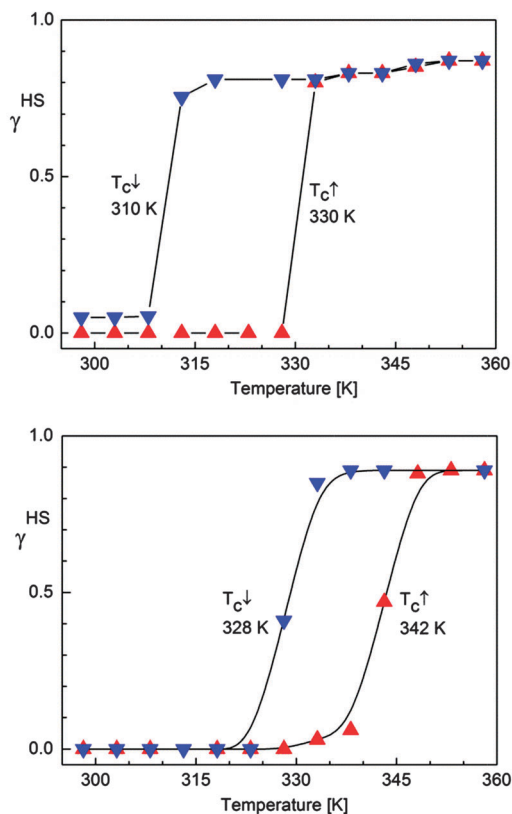


Fig. 8 Temperature dependence of the HS molar fraction (γ^{HS}) deduced from ^{57}Fe Mössbauer spectroscopy over the temperature range 298–360 K for SCO-1@MCM (top) and for SCO-2@MCM (bottom).

up to 358 K followed by a cooling cycle to 298 K. The spectrum at 78 K shows one quadrupole doublet attributed to LS Fe^{II} ions ($\delta^{\text{LS}} = 0.42(1) \text{ mm s}^{-1}$ and $\Delta E_{\text{Q}}^{\text{LS}} = 0.23(1) \text{ mm s}^{-1}$) (Fig. 9). At room temperature, the hyperfine parameters of the LS state (Table 5) are in exact agreement with the ones of the pure material with ($\delta^{\text{LS}} = 0.43(1) \text{ mm s}^{-1}$ and $\Delta E_{\text{Q}}^{\text{LS}} = 0.29(1) \text{ mm s}^{-1}$),¹⁷ thus indicating that the inclusion of $[\text{Fe}(\text{Htrz})_2\text{trz}]\text{BF}_4$ in the MCM matrix neither changed its microstructure nor resulted in an oxidation. However, the spin state population is modified because 11% of HS ions detected at room temperature for SCO-2 are now switched to the LS state in the composite. Usually, this fraction of HS ions in the LS state is attributed to chain ends, *i.e.*, the formation of short chains which end with a few percent of remaining water molecules.³⁶ Thus the coordination polymerization would tend to proceed completely within the matrix. Another explanation could be that the remaining paramagnetic Fe spins having a FeN_6 environment within the matrix would switch to the LS state due to a local pressure effect caused by the matrix. In both cases, a matrix influence is therefore identified. On warming up to 338 K, the composite remains in the LS state where a second doublet attributed to HS Fe^{II} ions grows slightly in intensity (7%), $\delta^{\text{HS}} = 0.98(1) \text{ mm s}^{-1}$ and $\Delta E_{\text{Q}}^{\text{HS}} = 2.53(1) \text{ mm s}^{-1}$ (Fig. 9). This behavior indicates the onset of a SCO behavior, a spin situation which differs from the pure material which only switches below 380 K,¹⁷ confirming a matrix influence on the spin state. On warming further, the HS signal ($\delta^{\text{HS}} = 0.95(1) \text{ mm s}^{-1}$

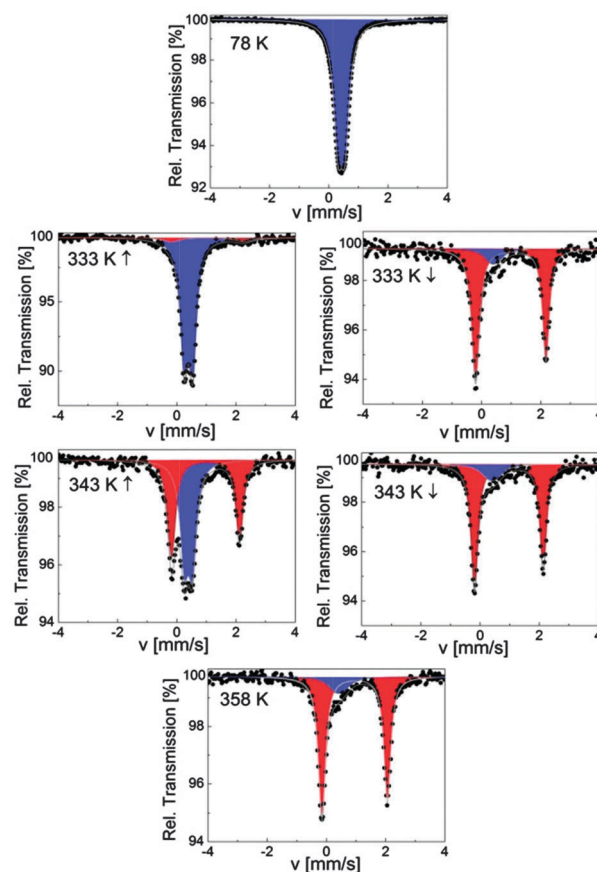


Fig. 9 Selected ^{57}Fe Mössbauer spectra of SCO-2@MCM; red and blue colors correspond to HS and LS doublets, respectively.

and $\Delta E_{\text{Q}}^{\text{HS}} = 2.21(1) \text{ mm s}^{-1}$) increases, revealing an asymmetric character, to reach 85% at 358 K. Worthwhile to note that the fraction of LS ions in the HS state is dramatically increased in the composite (Table 5) compared to the 3% detected in the pure material.¹⁷ On cooling, a hysteresis behavior is detected for the composite as seen in Fig. 9. The temperature dependence of the HS molar fraction shown in Fig. 8 allows to determine the width of the hysteresis loop $\Delta T = 14 \text{ K}$ as well as the transition temperatures ($T_{\text{c}}^{\uparrow} = 342 \text{ K}$ and $T_{\text{c}}^{\downarrow} = 328 \text{ K}$) which differ from the ones of the pure material ($T_{\text{c}}^{\uparrow} = 380 \text{ K}$ and $T_{\text{c}}^{\downarrow} = 344 \text{ K}$) recorded by Mössbauer spectroscopy.¹⁷ Thus the hysteresis width is considerably reduced compared to the genuine material. In addition, the sharpness of the ST is also affected indicating a loss of cooperativity, and the hysteresis loop is shifted downwards.

DC magnetic measurements

The magnetic properties of both composite materials are presented for comparison in Fig. 10. For SCO-1@MCM, the first cycle, from 200 K to 400 K, followed by a cooling to 200 K, reveals an abrupt hysteresis loop as large as 49 K, with $T_{\text{c}}^{\uparrow} = 389 \text{ K}$ and $T_{\text{c}}^{\downarrow} = 340 \text{ K}$. This hysteresis is not stable, due to the expected water release on warming. Therefore, for the second cycle, the hysteresis falls down to 31 K and is fully reversible, with $T_{\text{c}}^{\uparrow} = 371 \text{ K}$ and $T_{\text{c}}^{\downarrow} = 340 \text{ K}$. Compared to the magnetic data recorded on the pure material in the crystalline state, a clear

Table 5 Overview of ^{57}Fe Mössbauer parameters for SCO-2@MCM^a

T [K]	δ [mm s ⁻¹]	ΔE_Q [mm s ⁻¹]	$\Gamma/2$ [mm s ⁻¹]	Relative area [%]	Sites
78	0.42(1)	0.23(1)	0.18(1)	100	LS
298 ↑	0.43(1)	0.28(1)	0.14(1)	100	LS
303 ↑	0.42(1)	0.29(1)	0.14(1)	100	LS
308 ↑	0.42(1)	0.29(1)	0.14(1)	100	LS
318 ↑	0.42(1)	0.29(1)	0.14(1)	100	LS
323 ↑	0.42(1)	0.29(1)	0.14(1)	100	LS
328 ↑	0.41(1)	0.29(1)	0.14(1)	100	LS
333 ↑	0.40(1)	0.29(1)	0.15(1)	100	LS
338 ↑	0.98(1)	2.53(1)	0.19 ^b	7.0	HS
	0.39(1)	0.29(1)	0.15(1)	93.0	LS
343 ↑	0.97(1)	2.29(1)	0.15(1)	47.0	HS
	0.39(1)	0.27(1)	0.16(1)	53.0	LS
348 ↑	0.96(1)	2.27(1)	0.14(1)	86.0	HS
	0.41(1)	0.29 ^b	0.37(1)	14.0	LS
353 ↑	0.96(1)	2.25(1)	0.15(1)	86.0	HS
	0.41(1)	0.28 ^b	0.41(1)	14.0	LS
358 ↑	0.95(1)	2.21(1)	0.14(1)	85.0	HS
	0.41(1)	0.28(3)	0.32(2)	14.0	LS
343 ↓	0.97(1)	2.32(1)	0.14(1)	85.0	HS
	0.41(1)	0.28 ^b	0.32(1)	14.0	LS
338 ↓	0.97(1)	2.35(1)	0.15(1)	85.0	HS
	0.41(1)	0.28 ^b	0.32 ^b	14.0	LS
333 ↓	0.98(1)	2.37(1)	0.14(1)	85.0	HS
	0.42(1)	0.30 ^b	0.32 ^b	15.0	LS
328 ↓	0.99(1)	2.41(1)	0.14(1)	41.0	HS
	0.40(1)	0.30(1)	0.16(1)	59.0	LS
323 ↓	0.44(1)	0.27(1)	0.11(1)	100	LS
318 ↓	0.42(1)	0.30(1)	0.14(1)	100	LS
313 ↓	0.43(1)	0.29(1)	0.12(1)	100	LS
308 ↓	0.43(1)	0.29(1)	0.13(1)	100	LS
303 ↓	0.41(1)	0.31(1)	0.16(1)	100	LS
298 ↓	0.41(1)	0.31(1)	0.16(1)	100	LS

^a ↑ indicates warming and ↓ indicates cooling; δ = isomer shift relative to α -iron, ΔE_Q = quadrupole splitting, $\Gamma/2$ = half width at half maximum. ^b Fixed parameter.

shift of the ST upwards is observed. Indeed, $[\text{Fe}(\text{Htrz})_3](\text{BF}_4)_2 \cdot \text{H}_2\text{O}$ (SCO-1) displays a sharp ST at $T_c^\uparrow = 345$ K and $T_c^\downarrow = 323$ K while a lower hysteresis width is obtained for the dehydrated material with $T_c^\uparrow = 282$ K and $T_c^\downarrow = 276$ K.¹⁷

The transition temperatures recorded for the Mössbauer studies of SCO-1@MCM show a hysteresis loop shifted downwards compared to SCO-1@MCM recorded by DC magnetic measurements which is expected due to the time necessary to record each Mössbauer spectrum (one or two days in air atmosphere) which is enough to slowly partially dehydrate the material. As a result the hysteresis loop width by Mössbauer studies decreases to 20 K, with transition temperatures $T_c^\uparrow = 330$ K and $T_c^\downarrow = 310$ K (Fig. 8).

The SCO-2@MCM nanocomposite shows a fully reversible hysteresis loop of width 35 K, with $T_c^\uparrow = 381$ K and $T_c^\downarrow = 346$ K. Compared to the magnetic properties of the pure material, a slight shift downwards of the ST curve along with an hysteresis width decrease is noticed (bulk SCO-2: $T_c^\uparrow = 385$ K and $T_c^\downarrow = 345$ K).¹⁷

This result is consistent with theoretical predictions about a hysteresis reduction with particle size lowering.³⁷ The shift was more pronounced for the ST of nanoparticles of SCO-2 coated with gold recorded by DC magnetic measurements with $T_c^\uparrow = 373$ K and $T_c^\downarrow = 344$ K.^{38,39} These two results however contradict

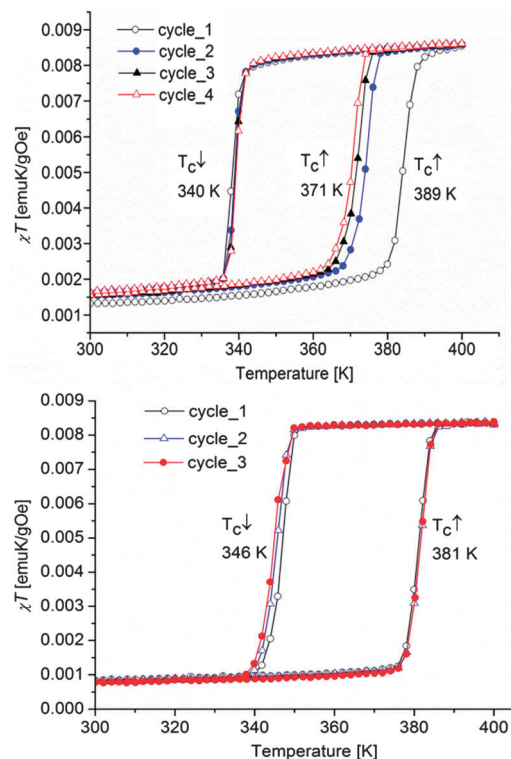


Fig. 10 Temperature-variable magnetic curves for SCO-1@MCM (top) and SCO-2@MCM (bottom).

earlier literature reports which indicate that the ST properties of this material are not modified when prepared as nanoparticles.⁴⁰

Optical reflectivity measurements

The SCO of iron(II) complexes with azole-based ligands is usually accompanied by a color change from purple in the LS state (lower temperature) to white in the HS state (higher temperature) (*cf.* Fig. 1).⁷ Hence, the change in spin state can also be followed by optical reflectance measurements in the solid state.^{41,42} The spin transitions of the SCO@MCM composites were studied optically under a nitrogen atmosphere to exclude any potential oxidation of the samples in air. The reflectance measurements were carried out using an optical disk cryostat mounted on the stage of a fluorescence microscope linked to a CCD camera. Before each set of measurements, the cryostat was purged for 2 h at 10 °C.

The first cycle recorded on SCO-1@MCM reveals an abrupt transition upon heating and cooling with a hysteresis as large as 49 K (black curve in Fig. 11). However, the hysteresis loop is reduced to 31 K in the next cycles (blue and red curves in Fig. 11), most presumably due to dehydration. These results are consistent with literature data reporting a large hysteresis loop for the hydrated SCO-1 phase ($T_c^\uparrow = 336$ K and $T_c^\downarrow = 323$ K) and a narrow one for the dehydrated phase, ($T_c^\uparrow = 291$ K and $T_c^\downarrow = 286$ K).¹⁷ For the SCO-1@MCM material, a hydration/dehydration phenomenon was observed in this work (see also below). However, the phase transitions between the alpha and beta forms, known for the bulk SCO-1 were not observed.¹⁷



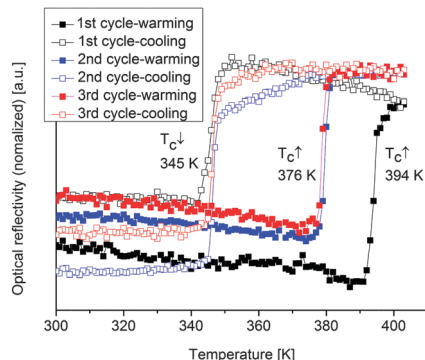


Fig. 11 Thermal dependence of the normalized optical reflectivity of SCO-1@MCM at a scan rate of 2 K min⁻¹. The transition temperatures recorded in warming and cooling modes are indicated.

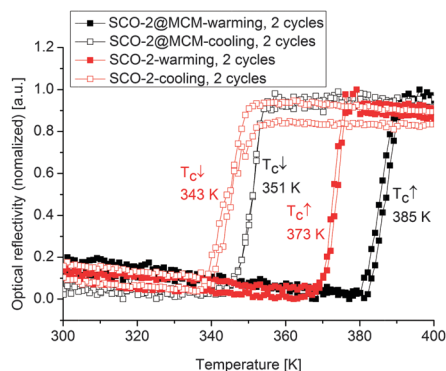


Fig. 12 Thermal dependence of the normalized optical reflectivity recorded at 2 K min⁻¹ on both SCO-2@MCM (black curves) and SCO-2 prepared as nanoparticles (red curves) using the reverse micelle method. The transition temperatures in the warming and cooling modes are given.

In the previous literature on SCO-1¹⁷ the hydration/dehydration might have been ascribed to an alpha/beta phase transition. Compared to the literature data on optical reflectivity for the free SCO-1 compound, the transition temperatures for SCO-1@MCM are shifted upwards with $T_c^\uparrow = 376$ K and $T_c^\downarrow = 345$ K,¹⁷ which suggests a significant matrix effect.

Optical reflectance measurements were recorded here for the first time on SCO-2 prepared as nanoparticles in the liquid state thanks to the reverse micelle method.⁴⁰ A hysteresis loop with $T_c^\uparrow = 373$ K and $T_c^\downarrow = 343$ K was observed (Fig. 12), which differs for the transition temperature in the heating mode compared to previous reflectivity measurements in the bulk sample (not in a nano-form) in the crystalline state $T_c^\uparrow = 381$ K and $T_c^\downarrow = 347$ K.¹⁷

The SCO-2 compound embedded in the MCM matrix shows a ST at $T_c^\uparrow = 385$ K and $T_c^\downarrow = 351$ K with a hysteresis of about 35 K (Fig. 12), hence of the same width as the bulk sample but slightly shifted upwards. This behavior is different to what was observed in the case of SCO-1@MCM.

Differential scanning calorimetry measurements

DSC measurements of SCO-2 reveal a sharp endothermic peak on warming and an exothermic peak on cooling at $T_{\text{max}}^\uparrow = 387$ K

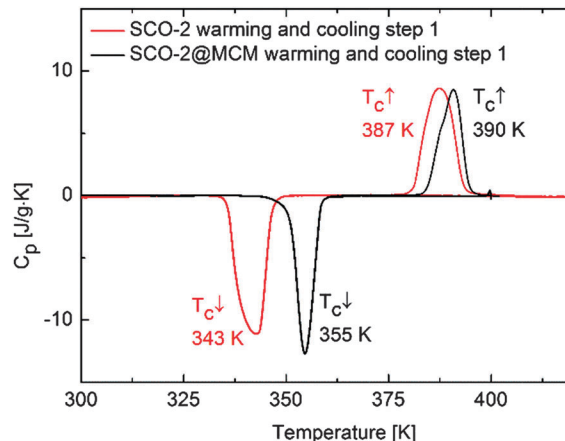


Fig. 13 Heat capacity versus temperature for SCO-2 and SCO-2@MCM recorded by differential scanning calorimetry at 10 K min⁻¹.

and $T_c^\downarrow = 343$ K, respectively. SCO-2@MCM reveals a similar pattern with a shift in temperature upwards at $T_{\text{max}}^\uparrow = 390$ K and $T_c^\downarrow = 355$ K, both at 10 K min⁻¹ (Fig. 13).

Comparison

We attribute the strong matrix effect in SCO-1@MCM to the compound being embedded between the pore walls with a pressure mediator (see below). The HS state has elongated Fe–ligand bond lengths (Fe–N $\approx$ 2.1–2.2 Å) compared to the LS state (Fe–N $\approx$ 1.9–2.0 Å) thereby requiring more space. Thus, the ST has to operate against the higher pressure exerted by the pore walls of the MCM matrix. In other words, higher external pressure favors the LS state so that the LS $\rightarrow$ HS transition occurs at higher molecular energy level, *i.e.*, temperature.

Table 2 provides a summarizing overview of the spin transition temperatures for the SCO and SCO@MCM materials from

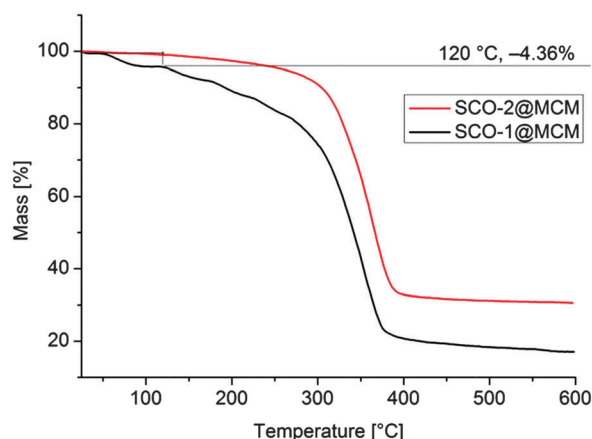


Fig. 14 Thermogravimetric analysis (TGA) shows that SCO-1@MCM possesses crystal water, whereas SCO-2@MCM nearly does not contain solvent. Up to 120 °C, 4.36 and 0.95 wt% are lost for SCO-1 and SCO-2@MCM, respectively. 4.36 wt% correlated to a Fe content of 7.2 wt% correspond to about 2H₂O molecules per SCO-1 formula unit (that is per Fe atom) in SCO-1@MCM. Further heating shows that the SCO-1@MCM composite contains more water which is lost above 120 °C.

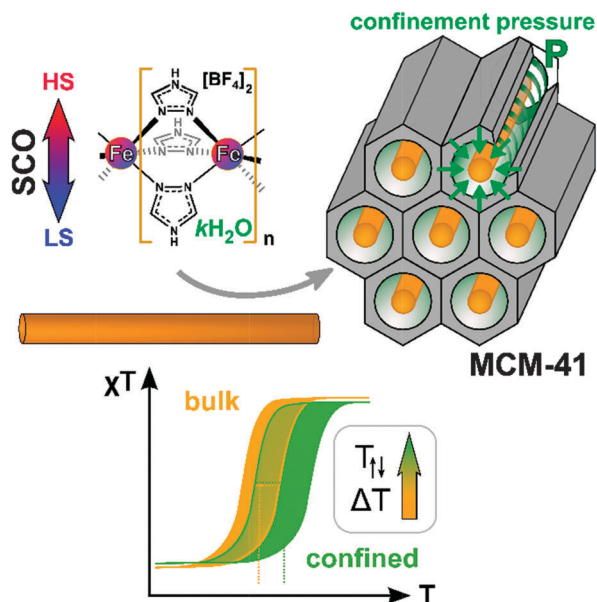


Fig. 15 Schematic presentation of the confinement pressure on the SCO-1 material mediated by water.

the different methods (for the hysteresis value included see Table S1 in ESI†).

It is worth to notice that, from Mössbauer, DC magnetic and optical reflectivity measurements, a strong matrix effect is found only for SCO-1@MCM, whereas the matrix effect for SCO-2@MCM is much weaker. We trace this to the hydration of SCO-1, which contains crystal water (Fig. 14) and suggest that water is important as a “pressure mediator” for delivering the “hydrostatic pressure” to induce a significant matrix effect. However, the specific reasons of the action of crystal water still need further investigation. The crystal water in SCO-1@MCM is partly lost upon heating to 127 °C (400 K) (*cf.* Fig. 14) which explains the shifts from the first to the second cycle in the warming temperatures ($T_c^\uparrow$) in Fig. 10 and 11.

Conclusions

Composite materials made of MCM-41 and of two 1D SCO compounds, namely $[\text{Fe}(\text{Htrz})_3](\text{BF}_4)_2 \cdot \text{H}_2\text{O}$ (SCO-1) and $[\text{Fe}(\text{Htrz})_2\text{trz}]\text{BF}_4$ (SCO-2) were successfully prepared in a form of embedded nanoobjects with 2 nm cross-section. Although, as concluded from the ^{57}Fe Mössbauer spectroscopy, the 1D chain is separated from the walls of the host matrix, a pressure effect was identified for SCO-1@MCM by temperature-variable magnetic and optical reflectance measurements, causing the spin transition to be shifted to higher temperatures, compared to the bulk material. For SCO-1@MCM the LS $\rightarrow$ HS transition ($T_c^\uparrow$) was increased by 44–58 K and the HS $\rightarrow$ LS transition ($T_c^\downarrow$) by 17–22 K in comparison to bulk SCO-1 during the first warming-cooling cycle. During the first warming to about 400 K partial dehydration occurs. For the partially dehydrated SCO-1@MCM material this shift was even more pronounced with $T_c^\uparrow$ increased by 85–89 K and $T_c^\downarrow$ increased by 59–64 K relative to bulk SCO-1.

The range given reflects the values from two different methods (temperature-variable magnetic and optical reflectance measurements). Also, the hysteresis between $T_c^\uparrow$ and $T_c^\downarrow$ increases by 26–33 K when placing SCO-1 into the MCM-41 matrix (Table S1 in ESI†). In the case of SCO-2@MCM this pressure or matrix effect only becomes evident upon comparison to a nanoSCO-2 reference material which was done by the optical reflectance measurements. The significant difference in matrix effect on the SCO behavior of the similar materials SCO-1 and SCO-2 is traced to the hydration of the SCO-1 and SCO-1@MCM material. This hydration is only partially lost during the first heating cycle which is additional evidence to the change of transition parameters and the role of water. Water is apparently crucial in exerting a confinement pressure or matrix effect on the spin transition (Fig. 15).

Experimental section

Materials

Iron(II) tetrafluoroborate hexahydrate ($\text{Fe}(\text{BF}_4)_2 \cdot 6\text{H}_2\text{O}$, Aldrich), 1,2,4-*4H*-triazole (99%, Alfa Aesar), and mesostructured silica (MCM-41 type, Aldrich; unit cell size: 4.5–4.8 nm; $0.98 \text{ cm}^3 \text{ g}^{-1}$ pore volume; 2.1–2.7 nm pore size; BET surface area $\sim 1000 \text{ m}^2 \text{ g}^{-1}$) were used as received without further purification.

Instrumentation

Powder X-ray diffractograms. Powder X-ray diffractograms were acquired at ambient temperature on a Bruker D2 Phaser using a flat low background sample holder and $\text{Cu-K}\alpha$ radiation ($\lambda = 1.54182 \text{ \AA}$) at 30 kV covering 2 theta angles 5–80° over a time of 2 h, that is $0.01^\circ \text{ s}^{-1}$.

Nitrogen adsorption isotherms. Nitrogen adsorption isotherms were acquired on a Quantachrome Nova[®], with 2 h degassing at a temperature of 120 °C in vacuum prior to each measurement.

CO₂ sorption isotherms. CO₂ sorption isotherms were measured using a Micromeritics ASAP 2020 automatic gas sorption analyzer at 0 °C, with 4 h degassing at a temperature of 120 °C in vacuum prior to each measurement.

Thermogravimetric analysis (TGA). Thermogravimetric analysis (TGA) was measured on a Netzsch TG 209 F3 at $5^\circ \text{C min}^{-1}$ heating rate using aluminum sample holders and nitrogen as carrier gas.

FT-IR measurements. FT-IR measurements were carried out on a Bruker TENSOR 37 IR spectrometer at ambient temperature in the range of 4000 to 500 cm^{-1} with an ATR unit (Platinum ATR-QL, Diamond).

Synthesis of SCO@MCM and SCO materials

Bulk SCO-1 and SCO-2 was synthesized according to ref. 17; see in ESI† for the analytical data. Nanoparticles of SCO-2 for optical reflectivity measurements were prepared as a purple solution following a procedure described in ref. 40, see in ESI† for the analytical data.

SCO-1@MCM. 1,2,4-*4H*-Triazole (208 mg, 3.0 mmol) and MCM-41 (100 mg) were stirred in 100 mL of methanol for 12 h. A solution of 337.5 mg (1.0 mmol) of $\text{Fe}(\text{BF}_4)_2 \cdot 6\text{H}_2\text{O}$ in 60 mL of



methanol was added to the slurry and the solvent was removed rapidly at 70 °C using a rotary evaporator. The formed product was collected and thoroughly washed with water 3 times and 10 mL each. The product was dried in an evacuated desiccator over silica gel. Yield ~253 mg of a red-pink powder.

SCO-2@MCM. To a solution of 416 mg (6.0 mmol) of 1,2,4-4H-triazole in 2 mL ethanol, 100 mg of MCM-41 powder was added and the slurry was stirred for 12 h. The formed suspension was transferred to a solution of 675 mg (2.0 mmol) of $\text{Fe}(\text{BF}_4)_2 \cdot 6\text{H}_2\text{O}$ in 4 mL of water and the mixture was stirred for 24 h at room temperature. The solid product was filtered, washed with ethanol and water for three times (10 mL each), and dried in an evacuated desiccator over silica gel. Yield ~235 mg of a light-pink powder.

Washing procedure. Any SCO precipitate, which formed on the outer MCM surface, was removed by washing procedures. SCO-1@MCM was washed three times with water because SCO-1 is very easy dissolved in water.

For SCO-2@MCM, the solid product was first placed in a beaker, stirred for 2 h with water and separated by centrifugation. This step was repeated 3 times. Then the solvent was changed to ethanol, stirred for 2 h again and centrifuged. The ethanol washing was repeated until no pink color appeared in the supernatant.

Atomic absorption spectrometry (AAS). Atomic absorption spectrometry (AAS) was used to measure the iron content. A few carefully weighted milligrams of each sample were completely dissolved in 2 mL HNO_3 (65%) solution. These solutions were transferred into 20 mL volumetric flasks. The iron content of each sample was determined by comparison to standard solution (Table 3).

Mössbauer studies

^{57}Fe Mössbauer spectra were recorded in transmission geometry using a Wissel spectrometer, equipped with a $^{57}\text{Co}(\text{Rh})$ source from Cyclotron Ltd and fitted to an Oxford Instruments bath cryostat for low temperature measurements. Samples were inserted into aluminium foils for the low temperature measurements. Measurements above room temperature were carried out with a Wissel Furnace MBF-1100. The sample was fixed between two round plates of B_4C , made of heat- and corrosion-resistant steel with high nickel content. The sample holder and heat screens were positioned in quartz tubes in air atmosphere. Spectra were fitted to a sum of Lorentzian line shapes by least-squares refinement using Recoil 1.05 Mössbauer Analysis Software.⁴³

Magnetic measurements

DC magnetic measurements were performed by using a Quantum Design MPMS XL-5 SQUID magnetometer. For both samples, magnetization measurements were carried out in a magnetic field of 1000 Oe, starting from 200 K up to 400 K (heating), and then cooling back to 200 K. No attempt was made to evaluate the molar magnetic susceptibility.

Optical reflectance measurements

Thermal dependence of the optical reflectance has been carried out at 2 K min^{-1} , with a Linkam optical cryostat mounted on

the stage of an Olympus BX51 fluorescence microscope linked to a CCD camera. Data were treated using ImageJ software (Wayne Rasband, National Institute of Mental Health, Bethesda, Maryland, USA).

Differential scanning calorimetry measurements

DSC were carried out in a $\text{He}(\text{g})$ atmosphere using a Perkin-Elmer DSC Pyris 1 instrument equipped with a cryostat and operating down to 98 K. Aluminum capsules were loaded with 20–50 mg of sample and sealed. The heating and cooling rates were fixed at 10 K min^{-1} . Temperatures and enthalpies were calibrated over the temperature range of interest (298–400 K) using the solid-liquid transitions of pure indium (99.99%)⁴⁴ over the range 78–298 K.

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