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Preparation of novel lithiated high-entropy spinel-type oxyhalides and their electrochemical performance in Li-ion batteries†

Olena Porodko,^a Ladislav Kavan,^b Martin Fabián,^a
Barbora Pitřna Lásková,^b Vladimír Šepelák,^{a,c} Hristo Kolev,^d
Klebson Lucenildo da Silva,^{c,e} Maksym Lisnichuk^f and Markéta Zukalová^{*b}

Compositionally complex doping of spinel oxides toward high-entropy oxides is expected to enhance their electrochemical performance substantially. We successfully prepared high-entropy compounds, *i.e.* the oxide $(\text{Zn}_{0.25}\text{Mg}_{0.25}\text{Co}_{0.25}\text{Cu}_{0.25})\text{Fe}_2\text{O}_4$ (HEOFe), lithiated oxyfluoride $\text{Li}_{0.5}(\text{Zn}_{0.25}\text{Mg}_{0.25}\text{Co}_{0.25}\text{Cu}_{0.25})_{0.5}\text{Fe}_2\text{O}_{3.5}\text{F}_{0.5}$ (LiHEOFeF), and lithiated oxychloride $\text{Li}_{0.5}(\text{Zn}_{0.25}\text{Mg}_{0.25}\text{Co}_{0.25}\text{Cu}_{0.25})_{0.5}\text{Fe}_2\text{O}_{3.5}\text{Cl}_{0.5}$ (LiHEOFeCl) with a spinel-based cubic structure by ball milling and subsequent heat treatment. The products exhibit particles with sizes from 50 to 200 nm with a homogeneous atomic distribution. The average elemental composition of the samples is close to the nominal value. ^{57}Fe Mössbauer spectroscopy revealed that incorporating Li and F or Cl and forming oxygen defects do not influence the redistribution of Fe^{3+} cations over the spinel lattice sites and result in their preferred octahedral coordination. Electrochemical measurements carried out using 2032-coin cells with a Li-metal anode have shown voltammetric charge capacities of 450, 694, and 593 mA h g^{-1} for HEOFe, LiHEOFeCl, and LiHEOFeF, respectively. The best electrochemical performance of LiHEOFeCl was ascribed to its smallest particle size. Galvanostatic chronopotentiometry at 1C rate confirmed high initial charge capacities for all the samples but galvanostatic curves exhibited capacity decay over 100 charging/discharging cycles. Raman spectroelectrochemical analysis conducted on the LiHEOFeF sample proved the reversibility of the electrochemical process for initial charging/discharging cycles. Electrochemical impedance spectroscopy revealed the lowest initial charge transfer resistance for LiHEOFeCl and its gradual decrease both for LiHEOFeCl and LiHEOFeF during galvanostatic cycling, whereas the charge transfer resistance of HEOFe slightly increases over 100 galvanostatic cycles due to the different mechanism of the electrochemical reduction.

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1 Introduction

The current technological boom, digitalization, and the development of portable devices supporting online services go

hand in hand with the high demand for high-capacity batteries with cycling stability enabling fast charging/discharging. Regardless of various newly developed battery systems, Li-ion batteries remain an essential energy storage system for the majority of applications. However, the charge capacity of electrode materials used for classical Li-ion batteries is limited to *ca.* 180 mA h g^{-1} . The intercalation anodes are limited by the capacity of graphite (372 mA h g^{-1} for LiC_6). To satisfy progress in advanced rechargeable lithium-ion batteries, a significant effort has been devoted to exploring electrode materials with optimal electrochemical properties. The novel concept of compositionally complex doping led to high-entropy compounds.¹ This allows the compositional variability needed to design new materials for next-generation batteries.² These compounds comprise the incorporation of multiple metal cations into single-phase crystal structures. Interactions among the various metal cations lead to unique electrochemical³ and magnetic⁴ properties. This concept was originally developed for high-

^aInstitute of Geotechnics, Slovak Academy of Sciences, Watsonova 45, 040 01 Košice, Slovak Republic

^bJ. Heyrovský Institute of Physical Chemistry, Czech Acad. Sci., Dolejškova 3, CZ-18200 Prague 8, Czech Republic. E-mail: marketa.zukalova@jh-inst.cas.cz

^cInstitute of Nanotechnology, Karlsruhe Institute of Technology, Hermann-von-Helmholtz-Platz 1, 76344 Eggenstein-Leopoldshafen, Germany

^dInstitute of Catalysis, Bulgarian Academy of Sciences, Acad. G. Bonchev St., Bldg. 11, Sofia, 1113, Bulgaria

^eDepartment of Physics, State University of Maringá, Av. Colombo 5790, 87020-900 Maringá, Brazil

^fInstitute of Physics, Pavol Jozef Šafárik University, Park Angelinum 9, 04154 Košice, Slovak Republic

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entropy alloys (HEAs) and the role that entropy plays in the phase stability of these compounds can be understood in terms of the Gibbs free energy of mixing (ΔG_{mix})^{2,5}

$$\Delta G_{\text{mix}} = \Delta H_{\text{mix}} - T\Delta S_{\text{mix}} \quad (1)$$

where ΔH_{mix} is the mixing enthalpy, ΔS_{mix} is the mixing entropy, and T is the absolute temperature. If the $T\Delta S_{\text{mix}}$ dominates the enthalpy term (ΔH_{mix}), the overall ΔG_{mix} becomes negative and the entropy stabilization is established.⁵ The mixing entropy (ΔS_{mix}) includes many entropic terms; however, the configurational entropy (ΔS_{conf}) is usually dominant.^{2,5} Another definition considers several elements, *i.e.* HEAs are formed by mixing equal or relatively large proportions of (usually) five or more elements.⁶ Like HEAs, entropy-based definitions have been proposed for high-entropy oxides (HEOs), where the configurational entropy from both cations and anions is considered.^{7,8} Typically, the entropy contribution of O^{2-} is zero, but the presence of oxygen vacancies or other anions (*e.g.*, F^-) can contribute to the system's entropy.⁷ HEOs with a spinel structure are suitable anode components in Li-ion batteries due to their high-rate capability and three-dimensional transport of lithium ions.^{9,10} In this context, several papers reported on the electrochemical properties of HEOs with a spinel structure of various compositions and morphologies, (see ref. 11–18 and references therein). The stabilization effect of entropy brings significant benefits for the storage capacity retention of high-entropy compounds and greatly improves their cycling stability. Additionally, the electrochemical behavior of high-entropy compounds depends on each of the metal cations present, thus providing the opportunity to tailor the electrochemical properties by simply changing the composition.³ To increase the discharge capacity, Lökçü *et al.*¹⁹ reported the synthesis of lithium-containing high-entropy ($\text{MgCoNiZn}_{1-x}\text{Li}_x\text{O}$) oxides. Subsequently, a variety of HEOs with incorporated lithium have been introduced. Duan *et al.*²⁰ reported the anode material ($\text{FeCoNiCrMnXLi}_3\text{O}_4$) where $\text{X} = \text{Cu, Mg, or Zn}$. ($\text{FeCoNiCrMnZnLi}_3\text{O}_4$) exhibited specific discharge capacities of 695, 577, 460, 336, 250, and 173 mA h g^{-1} at 50 mA g^{-1} , 100 mA g^{-1} , 200 mA g^{-1} , 500 mA g^{-1} , 1000 mA g^{-1} and 2000 mA g^{-1} , respectively. In a similar work,²¹ ($\text{CoNiZnFeMnLi}_3\text{O}_4$) provided an initial specific discharge capacity of 1104.3 mA h g^{-1} , good cycling stability (84% capacity retention after 100 cycles at 100 mA g^{-1}), and the rate performance of 293 mA h g^{-1} at 2000 mA g^{-1} . Ma *et al.*²² synthesized ($\text{FeMgNiCrMnLi}_3\text{O}_4$) using the solution combustion method combined with ball milling. This material showed a reversible discharge capacity of 850.7 mA h g^{-1} after 200 cycles at 2.0 A g^{-1} .

In our recent work,²³ we reported a novel lithiated high-entropy oxychloride $\text{Li}_{0.5}(\text{Zn}_{0.25}\text{Mg}_{0.25}\text{Co}_{0.25}\text{Cu}_{0.25})_{0.5}\text{Fe}_2\text{O}_{3.5}\text{Cl}_{0.5}$ (LiHEOFeCl) as an additive increasing the stability of a sulfur cathode in Li–sulfur batteries. A considerable effect of the LiHEOFeCl material was assigned to its excellent structural and electrochemical stability within the operating

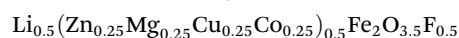
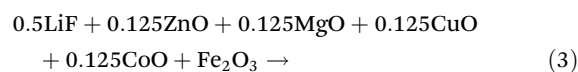
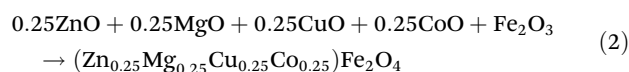
window of the Li–sulfur battery, where it had no inherent electrochemical activity and acted solely as an electrocatalyst. The charge capacity of LiHEOFeCl (evaluated by cyclic voltammetry in the potential window of 0.65/3.0 V vs. Li^+/Li at the scan rate of 0.1 mV s^{-1}) reached 648 mA h g^{-1} . Since this value exceeds almost four times the charge capacities of the state-of-the-art oxide anodes for Li-ion batteries and almost twice the capacity of a graphite anode, it is challenging to carry out a more detailed electrochemical study completed with galvanostatic charging/discharging, electrochemical impedance spectroscopy, and structural analysis. In addition, the same study is carried out for the parent ($\text{Zn}_{0.25}\text{Mg}_{0.25}\text{Co}_{0.25}\text{Cu}_{0.25}\text{Fe}_2\text{O}_4$) (HEOFe) and its lithium-containing fluorinated homologue $\text{Li}_{0.5}(\text{Zn}_{0.25}\text{Mg}_{0.25}\text{Co}_{0.25}\text{Cu}_{0.25})_{0.5}\text{Fe}_2\text{O}_{3.5}\text{F}_{0.5}$ (LiHEOFeF).

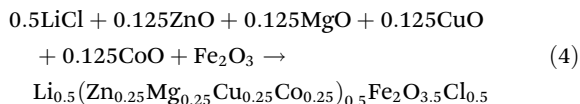
The objective of the present study is to establish the synthetic route of lithiated HEO oxides and oxyhalides and characterize their structure, composition, morphology, and crystallite size distribution by a combination of X-ray powder diffraction (XRPD), Atomic absorption spectroscopy (AAS), and Scanning transmission electron microscopy ((S)TEM). To determine the atomic structure at the local level, ^{57}Fe Mössbauer spectroscopy is employed to provide the local environment of Fe nuclei. Moreover, valence states of particular elements are investigated *via* X-ray photoelectron spectroscopy (XPS) to examine their detailed electrochemical properties, together with their electrochemical performances.

2 Experimental

2.1 Sample preparation

($\text{Zn}_{0.25}\text{Mg}_{0.25}\text{Co}_{0.25}\text{Cu}_{0.25}\text{Fe}_2\text{O}_4$) (HEOFe), $\text{Li}_{0.5}(\text{Zn}_{0.25}\text{Mg}_{0.25}\text{Co}_{0.25}\text{Cu}_{0.25})_{0.5}\text{Fe}_2\text{O}_{3.5}\text{F}_{0.5}$ (LiHEOFeF) and $\text{Li}_{0.5}(\text{Zn}_{0.25}\text{Mg}_{0.25}\text{Co}_{0.25}\text{Cu}_{0.25})_{0.5}\text{Fe}_2\text{O}_{3.5}\text{Cl}_{0.5}$ (LiHEOFeCl) were prepared by combination of ball milling and heat treatment. In the first step, a stoichiometric ratio of precursors zinc oxide (ZnO, 99.99% Aldrich), copper oxide (CuO, 99.99% Acros Organics), magnesium oxide (MgO, 99.99%, Acros Organics), iron oxide ($\alpha\text{-Fe}_2\text{O}_3$, 99.9%; Alfa Aesar) and lithium chloride (LiCl, 98%, Acros Organics) or lithium fluoride (LiF, 99.9%, Acros Organics) were used. The cobalt oxide CoO precursor was prepared by thermal decomposition of cobalt hydroxide ($\text{Co}(\text{OH})_2$, 95% Acros Organics) at 180 °C in a vacuum. The mixture of precursors (see eqn (2)–(4)) was milled at 600 rpm for 390 minutes in a planetary ball mill Pulverisette 7 premium line (Fritsch). A milling chamber (80 cm^3 in volume) and balls (10 mm in diameter) made of tungsten carbide (WC) were used. The ball-to-powder weight ratio was 30 : 1. Milling was carried out in an argon atmosphere.





In the second step, the mechanically activated mixtures were calcined at 600 °C in an argon atmosphere for 2 h.

2.2 Structure determination

XRPD patterns were obtained using a D8 Advance diffractometer (Bruker) operating with Cu K α radiation in the Bragg-Brentano configuration. The generator was set up at 40 kV and 40 mA and data were recorded in the range of 20 to 80° 2 θ . The ICSD²⁴ database was applied for the assignment of phases. Lattice parameters were determined by the Le Bail method in the space group *Fd3m* using the FullProf program²⁵ with the application of regular pseudo-Voigt function parameters.

The microstructure and morphology of the synthesized powders were investigated by (S)TEM (JEOL JEM-2100F), coupled with the Energy dispersive X-ray spectrometer (EDS) (Oxford Instruments). The sample was crushed in a mortar, dispersed in ethanol, and fixed on a copper-supported carbon grid. ImageJ software²⁶ was used to evaluate the particle size distribution from the (S)TEM micrographs.

The elemental composition was investigated by Atomic absorption spectroscopy AAS (Varian 240RS/240Z). Before the AAS measurement, the powdered samples were dissolved in aqua regia, filtered, and diluted.

⁵⁷Fe Mössbauer spectra were obtained in transmission geometry at room temperature. The ⁵⁷Co isotope in the Rh matrix was used as the γ -ray source. Mössbauer data were fitted using the spectral analysis software “Recoil”.²⁷ The Voigt-based fitting method was applied for all spectra. The experimental data were analyzed using a fitting procedure that considers hyperfine field distribution for the octahedrally coordinated ferric cations, Fe³⁺[O_h]. This results in the asymmetric broadening of the Fe³⁺[O_h] subspectrum arising from the different possible nearest-neighbor (T_d)-site configurations *via* the supertransfer mechanism in magnetic spinels.²⁸ In contrast, the Fe³⁺(T_d) subspectrum reflects the presence of a narrower field distribution in agreement with the usually smaller (T_d)-site supertransferred hyperfine field.²⁹ The derived isomer shifts (IS) are normalized to the IS of α -Fe.

XPS was performed on an ESCALAB MkII, VG Scientific (now Thermo Fisher Scientific) to analyze the elemental composition of the surface and oxidation states of the metal ions. Measurements were conducted at room temperature under 5 $\times$ 10^{−9} mbar. A twin anode MgK α /AlK α X-ray source was used with excitation energies of 1253.6/1486.6 eV, respectively. The spectra were recorded at the total instrumental resolution (as it was measured with the FWHM of the Ag 3d_{5/2} photoemission line) of 1.06 and 1.18 eV for MgK α and AlK α excitation sources, respectively. The binding energies of all the elements were calibrated relative to C 1s at 285.0 eV. The processing of the measured spectra included the subtraction of X-ray satellites and a Shirley-type background.³⁰ The peak posi-

tions and areas were determined by a symmetrical Voigt-based curve fitting.

2.3 Electrode preparation

The respective electrode material (HEOFe, LiHEOFeCl, or LiHEOFeF) was mixed with conductive carbon black C65 (Timcal) and a 2% aqueous solution of carboxymethylcellulose (CMC from Sigma). The mass ratio of the HEO-containing active material to C65 and CMC was 7 : 2 : 1. The mixture was then diluted with deionized water to a consistency of viscous paste and coated by doctor-blading on Al foil. After drying in air at ambient temperature and subsequently at 100 °C in a vacuum overnight, the coated Al foil was cut into disc electrodes of 15 mm in diameter. The electrodes were stored in a glove box with an argon atmosphere.

2.4 Coin cell preparation

Coin cells were assembled with a cathode described above, a Li-foil anode (14 mm in diameter), a glass microfiber separator (Whatman), and 20 μ L of electrolyte solution. The electrolyte solution consisted of 1 M LiPF₆ in ethylene carbonate/dimethyl carbonate (1 : 1 by volume). Electrolyte solutions contained 8–12 ppm H₂O as determined by Karl Fischer coulometric titration (Mettler Toledo). Electrolytes and solvents were of standard quality (p.a. or electrochemical grade) purchased from Aldrich or Merck. The electrolyte solution for Raman spectroelectrochemical measurement consisted of 1 M LiClO₄ in ethylene carbonate/dimethyl carbonate (1 : 1 by volume). Figure S1 (ESI†) depicts the scheme of the coin cell for Raman spectroelectrochemistry. The cell was equipped with a glass window to the cathode for *in situ* spectral measurement.

2.5 Electrochemical characterization

Cyclic voltammetry measurements were carried out with an Autolab 302N potentiostat (Metrohm) controlled by Nova SW in a potential window of 0.65 V/3.0 V vs. Li⁺/Li with a scan rate of 50, 20, 10, 5, 2, 1, 0.5, 0.2, and 0.1 mV s^{−1}. All potentials in this study are referred to the Li⁺/Li electrode. Galvanostatic chronopotentiometry at 1C was measured in 2032 coin-type test cells by the Neware Battery Testing System controlled by BTS 7.6 SW. Electrochemical impedance spectra (EIS) of the particular coin cell were obtained after cyclic voltammetry measurements, then after the 50th and 100th galvanostatic cycle by the FRA2 module of Autolab 302N. EIS was investigated over a frequency range from 100 kHz to 0.1 Hz (modulation amplitude 10 mV) at 2.4 V vs. Li⁺/Li (which is near the open-circuit potential). Before each EIS measurement, the potential was equilibrated at 2.4 V for 1 min. Spectra were evaluated using the Zview 4.0 (Scribner) software by fitting to a Randles-type circuit, in which *R*_s is the serial resistance and *R*_{CT} (the charge-transfer resistance) is parallel to the constant phase element (CPE) to account for non-ideal capacitive behavior of the cell. The circuit is completed by the Warburg impedance *Z*_w element for modeling the mass transport effects at lower frequencies.^{31,32} However, in contrast to the cited works, in which the impedances were normalized to the electrode



area, we present now the spectra normalized to the mass of the active electrode material.³³ Alternative normalization [to the physical (BET) areas] will give a similar result – the actual surface areas are all near 4 m² g^{−1} (ref. 23). Raman spectra were obtained on a MicroRaman system (LabRAM HR spectrometer, Horiba Jobin-Yvon) interfaced to an Olympus BX microscope. The spectra were excited with a He–Ne laser, 633 nm. The Raman spectrometer was calibrated using the F_{1g} line of Si at 520.2 cm^{−1}. *In situ* Raman spectra were obtained during chronoamperometry in a quasi-equilibrium state, which was attained after 2 min of arresting the electrode potential at the selected value. Chronoamperometry was carried out with Autolab PGSTAT30 apparatus controlled by Nova SW in the potential window of 3.0 V–0.7 V vs. Li⁺/Li.

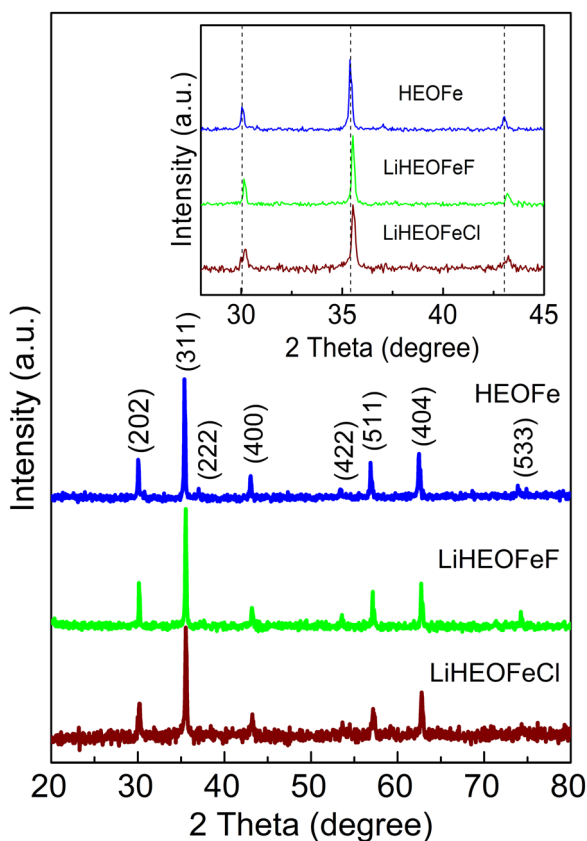


Fig. 1 XRPD patterns of HEOFe, LiHEOFeF, and LiHEOFeCl. The inset is a zoomed image to accent the shift of diffraction peaks by the structure modification. Vertical lines are guides to the eye.

3 Results and discussion

3.1 Structural and morphological properties

Fig. 1 shows the XRPD patterns of the as-prepared samples. They confirm the spinel-based phase with a cubic structure (*Fd3m*), without detectable impurities. A slight shift of the diffraction peaks for LiHEOFeF and LiHEOFeCl to the higher angular positions is observed (Fig. 1, inset). This shift indicates the incorporation of lithium and fluoride/chloride ions into the spinel lattice. As further summarized in Table 1, the calculated lattice parameters decrease with the incorporation of lithium and halide ions, which is reasonable considering the difference in the ionic radii of the particular elements.³⁴ The mean crystallite sizes of the studied samples, determined from Scherrer's equation, are from 50 to 70 nm. The morphology and particle size distribution were further investigated by (S)TEM. Fig. 2 shows the representative micrographs of the synthesized samples. They consist of agglomerated particles with a size distribution ranging from 100 to 200 nm, except for LiHEOFeF characterized by bigger particles. Approximately 100 particles were analyzed to determine the average crystallite size *D* (listed in Table 1). The average particle size determined by (S)TEM analysis is higher than that from XRPD. It is attributed to the agglomeration of particles. Moreover, the HR-TEM images show lattice fringes characteristic for the particular crystallographic planes (twenty atomic planes were analyzed to determine the average value). The observed values are in good agreement with those determined from the XRPD data according to the following formula:

$$\frac{1}{d_{hkl}^2} = (h^2 + k^2 + l^2)/a^2 \quad (5)$$

where *hkl* are Müller indices, and *a* is the lattice parameter.

EDS mapping of the particular elements Zn, Mg, Cu, Co, Fe, O, F, and Cl reveals their homogeneous distribution (Fig. 3). Contamination attributed to the abrasion during the milling process was below the detection limit of EDS (0.4 at% W). Even though, according to our recent experience,³⁰ the low concentration of tungsten has an insignificant effect on the electrochemical properties of oxide-based anodes. As listed in Table 2, the final chemical composition of the synthesized spinels is quasi-equi-molar and the average elemental composition determined by AAS is close to the nominal composition.

To address the local structural arrangement of the as-prepared materials, the samples were investigated by ⁵⁷Fe Mössbauer spectroscopy. This method provides information on the chemical (valence) and magnetic states of the constitu-

Table 1 The lattice parameter (*a*), interplanar distance (*d*), and average crystallite size (*D*) for HEOFe, LiHEOFeF, and LiHEOFeCl

Sample	Lattice parameter, <i>a</i> (Å)	Interplanar distance, <i>d</i> observed (Å)	Interplanar distance, <i>d</i> calculated (Å)	Crystallite size, <i>D</i> XRD (nm)	Crystallite size, <i>D</i> TEM (nm)
HEOFe	8.4008(9)	2.8 for (<i>hkl</i> = 003)	2.8 for (<i>hkl</i> = 003)	54	140
LiHEOFeF	8.3662(1)	4.7 for (<i>hkl</i> = 111)	4.8 for (<i>hkl</i> = 111)	53	214
LiHEOFeCl	8.3690(4)	2.8 for (<i>hkl</i> = 222)	2.8 for (<i>hkl</i> = 222)	65	155



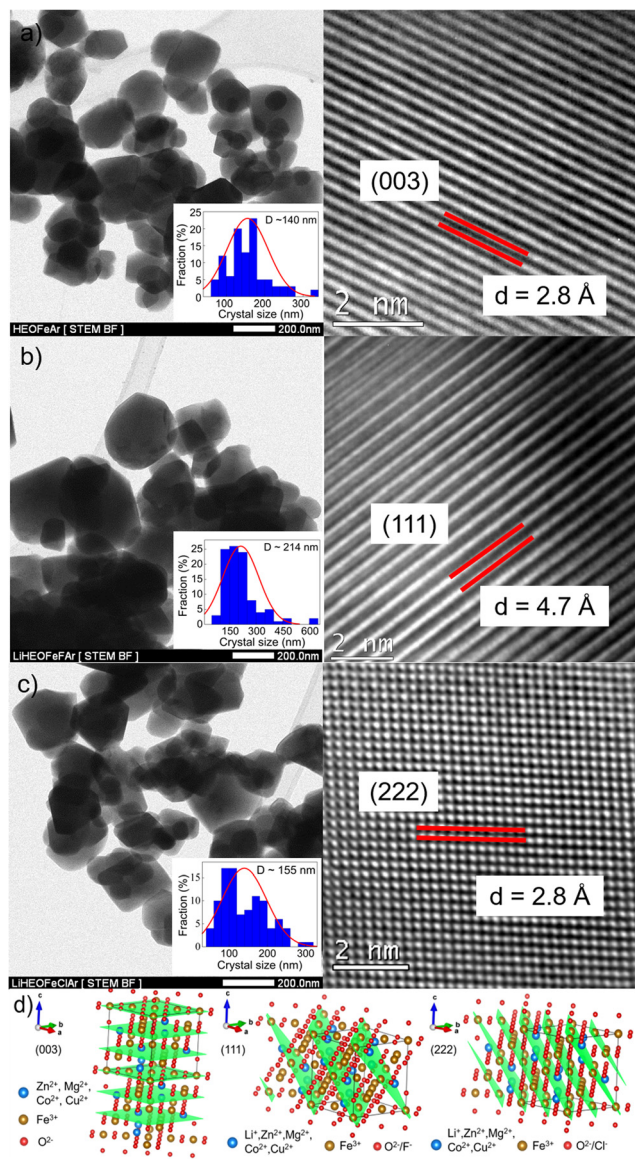


Fig. 2 (S)TEM and HR-TEM micrographs of (a) HEOFe, (b) LiHEOFef, and (c) LiHEOFefCl synthesized samples. Insets show the particle size distribution. Interplanar distances, d , correspond to those determined from eqn (5). (d) Idealized spinel structure of the investigated samples. Particular crystal planes for HEOFe (003), LiHEOFef (111), and LiHEOFefCl (222) are highlighted.

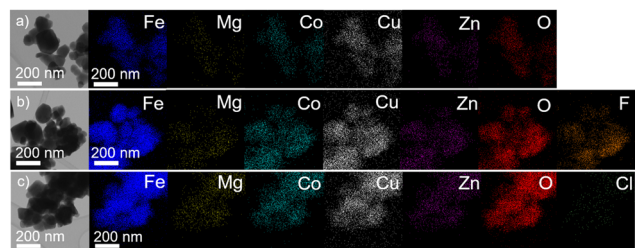


Fig. 3 (S)TEM micrograph of the synthesized (a) HEOFe, (b) LiHEOFef, and (c) LiHEOFefCl samples with the corresponding elemental maps determined by EDS.

Table 2 Elemental composition of HEOFe, LiHEOFef, and LiHEOFefCl samples represented by the M/Fe atomic ratios determined by AAS

HEOFe	Zn/Fe	Mg/Fe	Co/Fe	Cu/Fe	Li/Fe
Expected	0.1250	0.1250	0.1250	0.1250	—
Measured	0.1475	0.1350	0.1375	0.1238	—
LiHEOFef					
Expected	0.0625	0.0625	0.0625	0.0625	0.2500
Measured	0.0770	0.0702	0.0690	0.0616	0.1931
LiHEOFefCl					
Expected	0.0625	0.0625	0.0625	0.0625	0.2500
Measured	0.0575	0.0560	0.0628	0.0538	0.2292

ent iron ions including their local coordination and oxygen polyhedral distortions.³⁵ It is known that spinel ferrites with the general formula $M_1M_2O_4$ (M_1 and M_2 are divalent and trivalent metal cations, respectively) exhibit complex disordering phenomena involving the redistribution of cations over the sites of tetrahedral (T_d) and octahedral [O_h] coordination provided by the spinel structure.^{36–39} To emphasize the site occupancy at the atomic level, the structural formula of these materials may be written as $(M_{1-\lambda}M_{2\lambda})[M_{1\lambda}M_{2-\lambda}]O_4$, where the parentheses and square brackets denote (T_d) and [O_h] sites, respectively. The symbol λ represents the so-called degree of inversion defined as the fraction of the (T_d) sites occupied by trivalent (M_2^{3+}) cations. It varies from $\lambda = 0$ (normal spinel) to $\lambda = 1$ (fully inverse spinel). The value of $\lambda = 2/3$ corresponds to the random arrangement of cations. A non-equilibrium cation distribution has been evidenced in spinels prepared by various synthesis routes.^{40–44}

Fig. 4 shows the room-temperature ^{57}Fe Mössbauer spectra of HEOFe, LiHEOFef, and LiHEOFefCl. The spectra of all investigated samples are well-fitted using two overlapping sextets corresponding to Fe^{3+} ions located on both (T_d) and [O_h] sites provided by the spinel ferrite structure. Table 3 presents the hyperfine parameters of (T_d)- and [O_h]-site ferric ions in the investigated materials. From the intensities of spectral components, it is revealed that Fe^{3+} cations in the HEOFe sample incline to the random distribution over tetrahedral and octahedral positions ($I(T_d)/I(O_h) \sim 33/67$). As it is evident from the quantitative analysis of ^{57}Fe Mössbauer spectra, the presence of Li and F or Cl ions in the structures of the LiHEOFef and LiHEOFefCl does not influence the redistribution of Fe^{3+} cations and results in their preferred octahedral coordination ($I(T_d)/I(O_h) \sim 34/66$ for LiHEOFef and $I(T_d)/I(O_h) \sim 34/66$ for LiHEOFefCl; see Table 3). It should be emphasized that the presented values of the intensities of Mössbauer subspectra reflect only the distribution of Fe^{3+} . The estimation of the degree of inversion for the HEOFe, LiHEOFef, and LiHEOFefCl samples is not possible, because the distribution of other constituent cations among (T_d) and [O_h] sites in the spinel structures is unknown.

Fig. 5a–h show the XPS spectra (measured with $\text{AlK}\alpha$ excitation, see Experimental) of Zn 2p, Mg 1s, Co 2p_{1/2}, Cu 2p_{3/2}, Fe 2p, O 1s, F 1s, and Cl 2p for the samples HEOFe, LiHEOFef, and LiHEOFefCl, respectively. Fig. 5a compares the

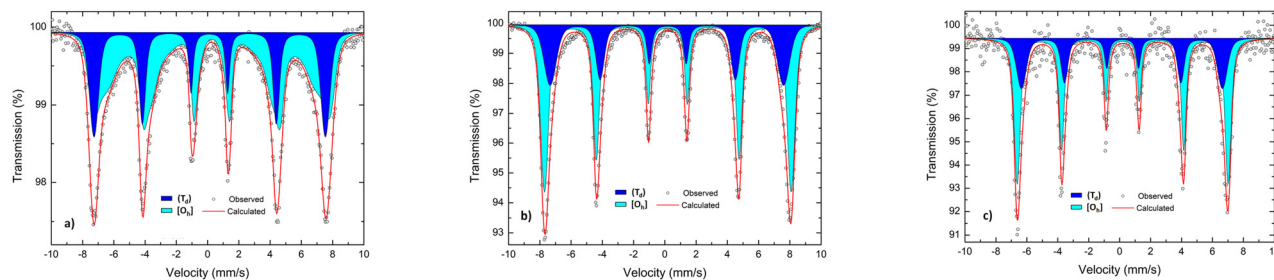


Fig. 4 The room-temperature ^{57}Fe Mössbauer spectra of (a) HEOFe, (b) LiHEOFeF, and (c) LiHEOFeCl. The dark blue and light blue spectra correspond to tetrahedrally (Td) and octahedrally [Oh] coordinated iron cations, respectively.

Table 3 Hyperfine parameters (IS: isomer shift, B_{hf} : the average magnetic hyperfine field, I : relative intensity of the spectral component) obtained by fitting the room-temperature ^{57}Fe Mössbauer spectra of the HEOFe, LiHEOFeF, and LiHEOFeCl samples

Sample	Spectral component	IS (mm s^{-1})	B_{hf} (T)	I (%)
HEOFe	(Td) (dark blue)	0.12(3)	45.9(1)	32.9(5)
	[Oh] (light blue)	0.26(4)	41.1(9)	67.1(5)
LiHEOFeF	(Td) (dark blue)	0.15(6)	46.3(1)	34.4(6)
	[Oh] (light blue)	0.18(8)	48.9(4)	65.6(3)
LiHEOFeCl	(Td) (dark blue)	0.16(8)	40.2(8)	33.5(1)
	[Oh] (light blue)	0.19(4)	42.2(8)	66.5(1)

Zn2p energy range of the samples. For all three samples, the shape of the line is identical with a binding energy of 1021.2 eV, indicating a Zn oxidation state of Zn^{2+} .^{45,46} Fig. 5b presents the Mg 1s photoelectron lines. Dashed lines indicate the positions of photoemission lines of Mg^0 (1303.0 eV) and Mg^{2+} (1304.3 eV) reported in the literature.⁴⁷ Variations in the line shape suggest Mg–F bonding in LiHEOFeF and possible Mg–O and Mg–Mg bonding in all three samples. In LiHEOFeCl, the line shape is influenced by the Auger peak of chlorine (ClKMM).

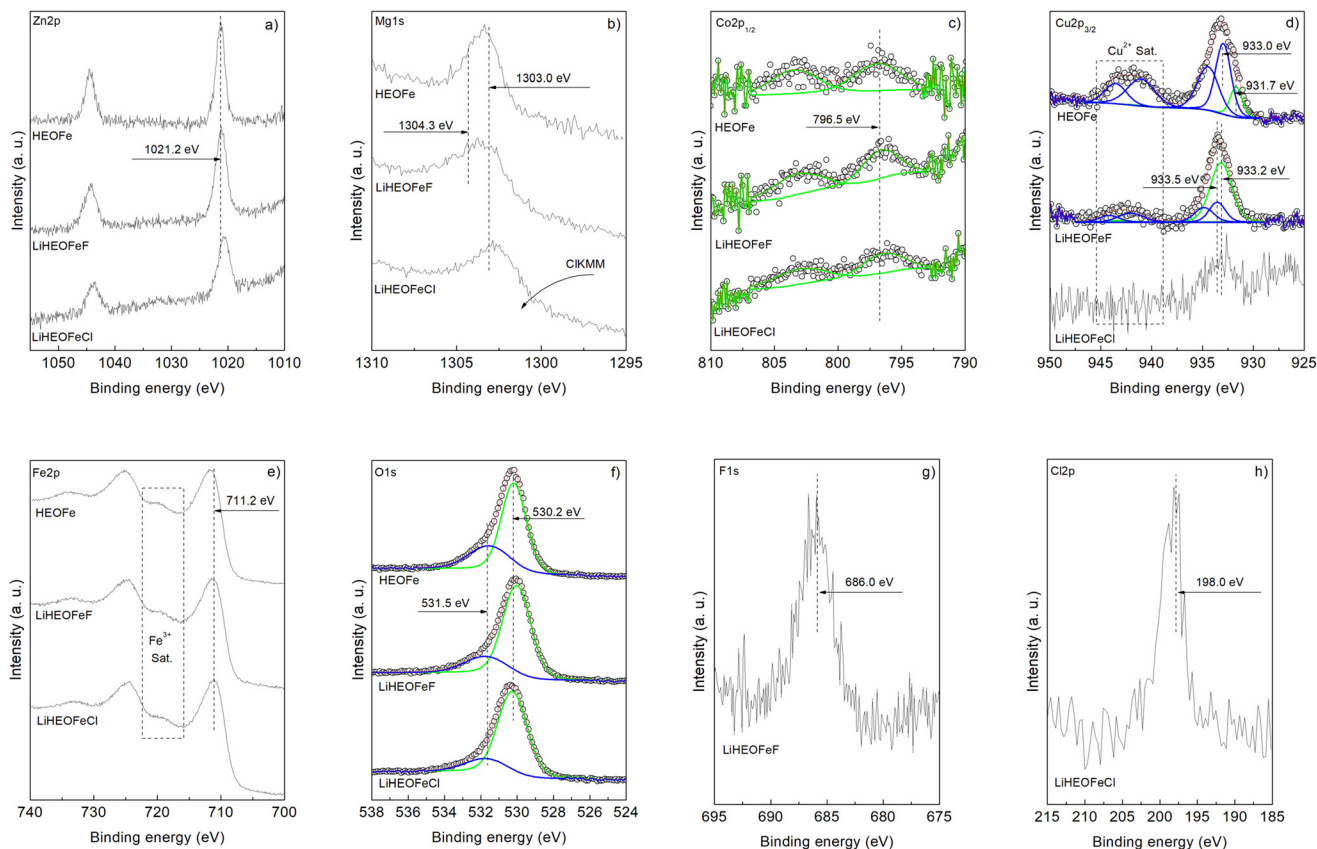


Fig. 5 High-resolution XPS spectra of the HEOFe, LiHEOFeF, and LiHEOFeCl compounds showing the signals corresponding to (a) Zn 2p, (b) Mg 1s, (c) Co 2p_{1/2}, (d) Cu 2p_{3/2}, (e) Fe 2p, (f) O 1s, (g) F 1s and (h) Cl 2p.



The Co 2p_{1/2} spectra in Fig. 5c confirm the presence of Co²⁺ with a peak at 796.5 eV binding energy, accompanied by an intense 3d → 4s “shake-up” satellite at 802.5 eV.^{45,48} Fig. 5d shows the Cu 2p_{3/2} signal. The satellite structure in the 938–946 eV range indicates a Cu oxidation state of Cu²⁺.^{45,49,50} Curve fitting suggests Cu⁺/Cu²⁺ ratios of 0.4 and 4.1 for HEOFe and LiHEOFeF, respectively. The noisy spectrum of LiHEOFeCl complicates the analysis, but a Cu²⁺ oxidation state is likely. Thus, the lithiated samples stabilize Cu⁺ ions on the surface, whereas Cu²⁺ predominates in HEOFe. It is interesting to note two different chemical shifts for Cu²⁺. The Fe 2p core level spectra in Fig. 5e show a binding energy of 711.2 eV with a satellite structure at 716–722 eV, typical of Fe³⁺ in agreement with our Mössbauer spectra.⁵¹ The O 1s spectra in Fig. 5f were fitted into two peaks at 530.2 eV and 531.5 eV, corresponding to metal–oxide bonds and defect sites with low oxygen coordination, respectively.^{19,45} The defect sites with low oxygen coordination to metal–oxide ratios are 0.45, 0.29, and 0.22 for HEOFe, LiHEOFeF, and LiHEOFeCl, respectively, indicating that adding lithium, chloride, and fluoride reduces defect sites, with the lowest value for LiHEOFeCl. Fig. 5g and h show the binding energies of F 1s (686.0 eV) and Cl 2p (198.0 eV), respectively, corresponding to metal–fluoride and metal–chloride bonds.^{19,52}

3.2 Electrochemical properties

The electrochemical performance of HEOFe, LiHEOFeCl, and LiHEOFeF in 2032-coin cells was evaluated by cyclic voltammetry of Li insertion in the potential window of 0.65 V/3.0 V vs. Li⁺/Li at the scan rate of 50, 20, 10, 5, 2, 1, 0.5, 0.2 and 0.1 mV s^{−1} and by galvanostatic chronopotentiometry at 1C. In addition, each cell was evaluated by electrochemical impedance spectroscopy after cyclic voltammetry and then after 50 and 100 cycles of galvanostatic charging/discharging. Fig. 6a

shows the cyclic voltammograms of HEOFe, LiHEOFeCl, and LiHEOFeF. All three cyclic voltammograms contain one reduction and one main oxidation peak at 1.0 V and 1.7 V, respectively. The charge capacities calculated from the oxidation branch of the cyclic voltammogram are 450, 694, and 593 mA h g^{−1} for HEOFe, LiHEOFeCl, and LiHEOFeF, respectively. Such a high charging capacity speaks for a conversion reaction of lithiation for all three high-entropy compounds with a spinel structure, although the conversion mechanism is mainly suggested for high-entropy compounds with a rock salt structure with less space for Li accommodation.³

Fig. 6b shows the dependence of the charge capacity calculated from the discharge (oxidation) branch of the cyclic voltammogram on the scan rate. The corresponding cyclic voltammograms are presented in Fig. S2.† Whereas LiHEOFeCl already reaches its maximum charge capacity at the scan rate of 0.5 mV s^{−1}, the charge capacities of both HEOFe and LiHEOFeF exhibit a gradual increase with the decreasing scan rate, which evidences a slower conversion reaction. This effect is most probably caused by different particle sizes of the tested materials. The lithiation of high-entropy oxides is a complex process involving a conversion reaction between divalent metals in the structure with Li.³ This conversion reaction can be accelerated by increasing the interfacial area where the reaction occurs. Consequently, smaller particles with a higher surface-to-volume ratio are beneficial in promoting this process. Although the average crystallite size *D*((S)TEM) of HEOFe is slightly smaller than that of LiHEOFeCl (140 nm vs. 155 nm) (Fig. 2), the size distribution of the latter exhibits its maximum for particles with less than 100 nm in diameter. In contrast, the maximum of the HEOFe particle size distribution lies between 100 and 200 nm.

After cyclic voltammetry, the electrochemical impedance spectrum of a particular coin cell was recorded. Then the per-

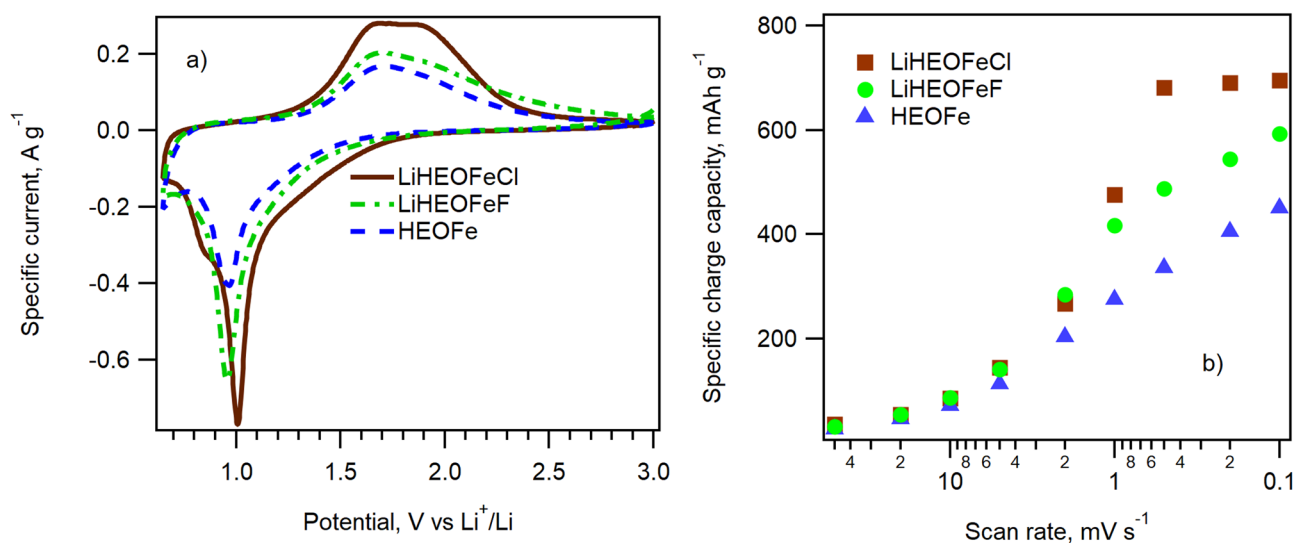


Fig. 6 (a) Cyclic voltammograms measured on 2032-coin cells containing LiHEOFeCl, LiHEOFeF, and HEOFe with a scan rate of 0.1 mV s^{−1}. (b) The scan-rate dependence of the charge capacities calculated from the discharge branch of the cyclic voltammogram on the scan rate for HEOFe (blue triangles), LiHEOFeCl (brown squares), and LiHEOFeF (green circles).



formance of all three materials during long-term cycling was evaluated by galvanostatic chronopotentiometry at the 1C charging/discharging rate. In addition, the electrochemical impedance spectrum was measured after the 50th and 100th cycles of charging/discharging. Fig. 7 shows the course of the galvanostatic curve for all three materials together with the calculated coulombic efficiency. Due to the interruption of galvanostatic measurements for all samples after the 50th cycle (for electrochemical impedance spectroscopy measurement) the coulombic efficiency curves show a small discontinuity. All the materials exhibit charge capacity decay, which is most pronounced for LiHFeOFeCl. This capacity decay is typical of conversion electrode materials, which often show substantial capacity degradation at high currents due to kinetic limitations of the diffusion-driven processes during de-lithiation.³

The charge capacity is stabilized after 50 cycles at *ca.* 200 mA h g⁻¹ for all the materials. Since the battery capacity decay is commonly accompanied by irreversible changes in the electrode materials, Raman spectroelectrochemistry on the LiHFeOFeF sample (with the most uniform particle size distribution) was conducted to reveal morphology changes during charging/discharging. The coin cell equipped with a glass window near the LiHFeOFeF electrode was assembled for *in situ* Raman spectroelectrochemical analysis. The coin cell open circuit potential was 1.75 V before measurement. Chronoamperometry was applied to the freshly prepared coin cell in the potential range from 3 V to 0.7 V *vs.* Li⁺/Li. Fig. 8 shows the Raman spectra of the electrode material recorded

for each potential step during the lithiation (Fig. 8a) or de-lithiation (Fig. 8b). The reference Raman spectra of pristine LiHFeOFeF and carbon C65 are also shown in Fig. 8a (red line) and (Fig. 8b) (blue line), respectively. The Raman features of LiHFeOFeF disappear during electrochemical charging and only the carbon additive C65 and electrolyte are seen at a potential of 0.7 V (Fig. 8a). These data indicate that the structural disruption and changes of LiHFeOFeF during charging are not associated with any new Raman modes. Fig. S3 (ESI[†]) shows current fluctuations occurring in the potential range around 2 V–1.6 V. These fluctuations are associated with the disappearance and rediscovery of Raman features of LiHFeOFeF during charging in the discussed potential range (Fig. 8a). This behavior can be attributed to the first cycle on a freshly prepared electrode, where the electrode response is inhomogeneous, and the solid/electrolyte interphase is formed. The discharging process from 0.7 V to 3 V leads to gradual recovery of the LiHFeOFeF Raman features (Fig. 8b). The spectra at 2 V and 2.8 V in (Fig. 8b) were collected at two different places of the cathode. The different rate of discharging in different places at an electrode is observed. Besides the inhomogeneity of the amount of the LiHFeOFeF material, the reason for this different recovery rate can be the different particle sizes in both places (see the particle size distribution in Fig. 2). However, the spectrum at 3 V is obtained in the identical place of the cathode as the spectrum at 2.8 V and the LiHFeOFeF Raman feature recovery is confirmed.

The data show good reversibility of the system during the first charging/discharging cycle.

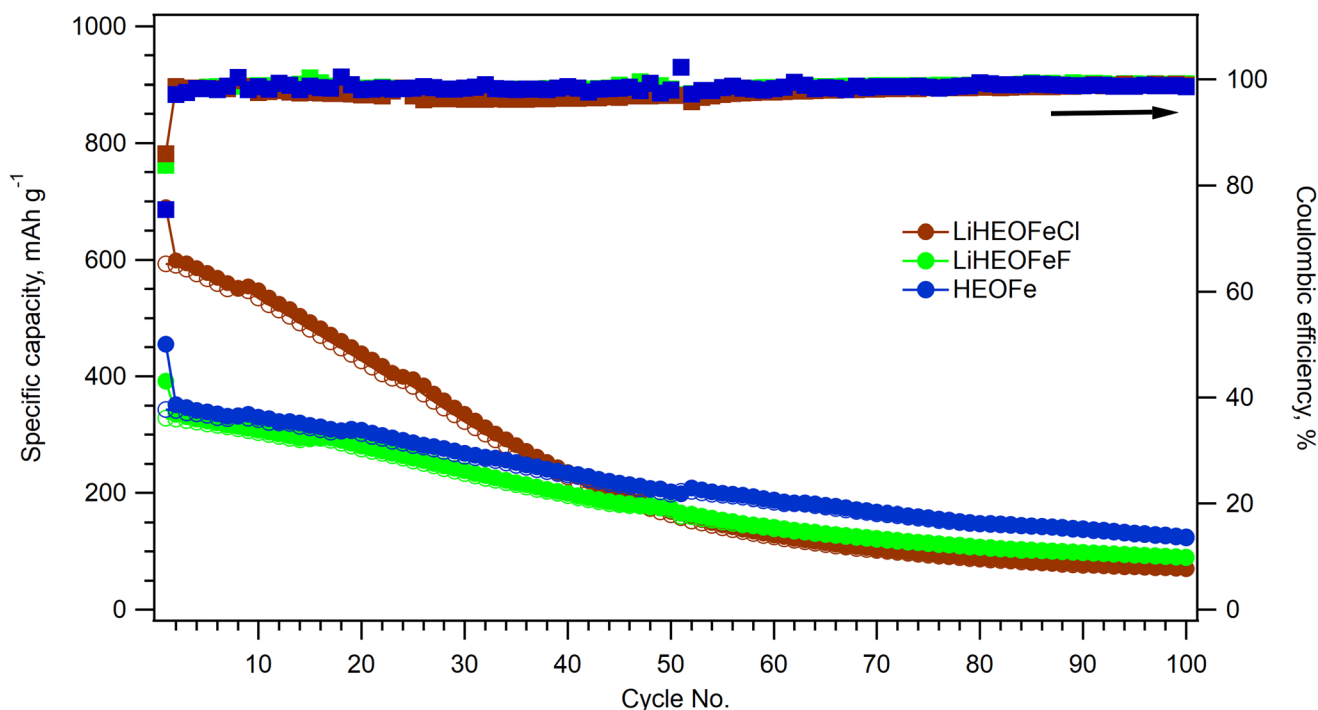


Fig. 7 Galvanostatic chronopotentiometry at 1C rate for HFeOFe (blue circles), LiHFeOFeCl (brown circles), and LiHFeOFeF (green circles). The charge capacity within charging is marked with full circles and the charging capacity within discharging with empty circles. The corresponding coulombic efficiencies for HFeOFe, LiHFeOFeCl, and LiHFeOFeF are depicted by squares of the same color coding.



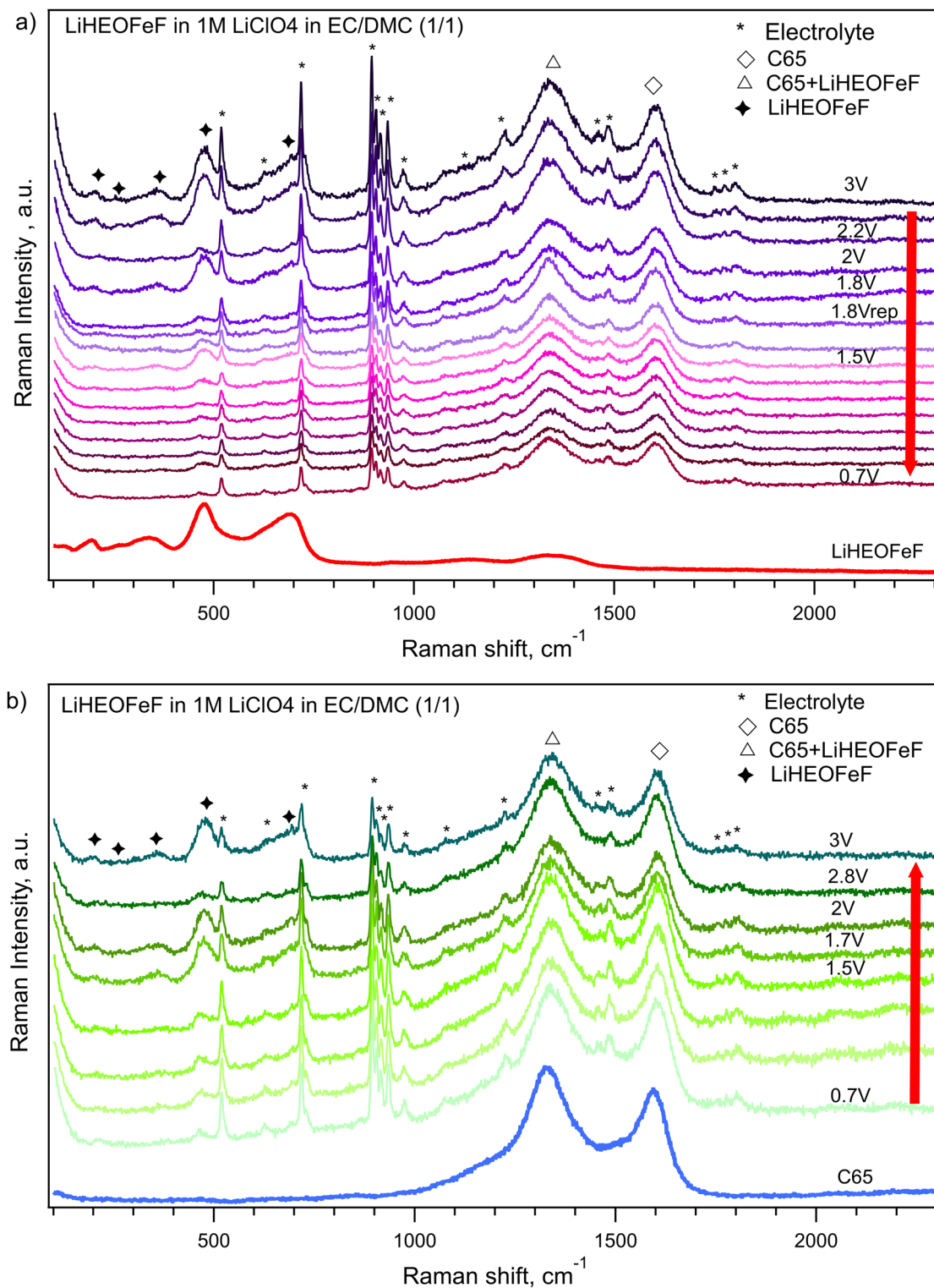


Fig. 8 *In situ* Raman spectroelectrochemical analysis of coin cells with the LiHFeF cathode and 1M LiClO_4 in ethylene carbonate/dimethyl carbonate (1 : 1 by volume) as the electrolyte solution. (a) Charging of the coin cell from 3 V to 0.7 V vs. Li^+/Li . The Raman spectrum of blank LiHFeF is also shown (red line). The label rep stands for repeated measurement at 1.8 V. (b) Discharging of coin cells from 0.7 V to 3 V vs. Li^+/Li . The Raman spectrum of the blank C65 additive is also shown (blue line). Spectra are offset, but the intensity scale is identical for all plots.



The impedance spectra shown in Fig. 9 confirm the high electrochemical activity of freshly assembled cells with a LiHFeOFeCl-based cathode. For easy comparison of individual

electrodes, the spectra are normalized to the mass of the active material.³³ The EIS fitting provided charge-transfer resistances of 5.36 mΩ g, 33.1 mΩ g, and 3.36 mΩ g for the cells with

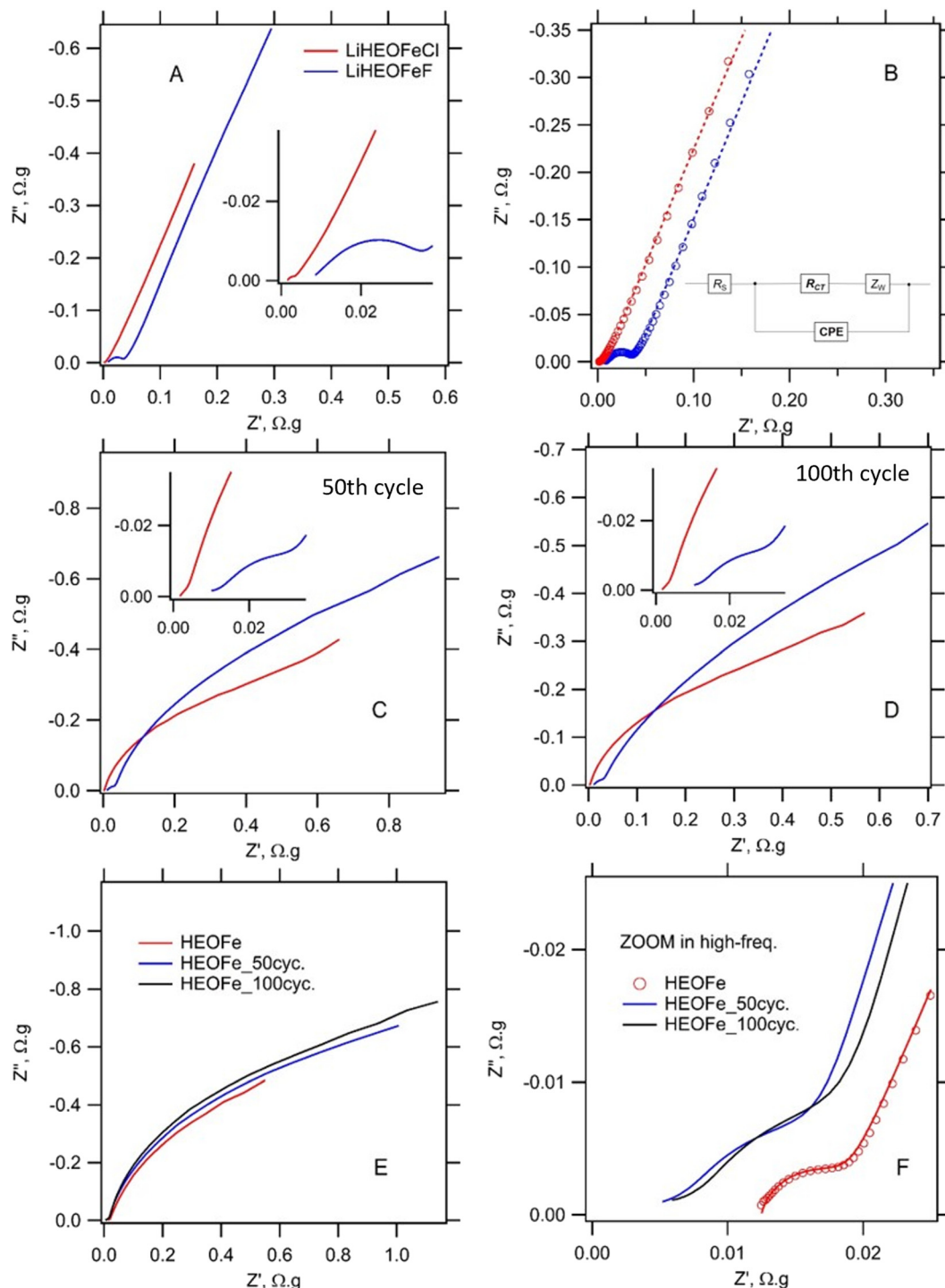


Fig. 9 Electrochemical impedance spectra of our coin cells with LiHFeOFeX- and HFeOFe-based cathodes. (A) Freshly assembled cells after nine forming cycles by cyclic voltammetry. (Inset: details of the spectra measured at the high-frequency cutoff). (B) The spectra as in chart A but fitted to the equivalent circuit shown. (C) Spectra of the same cell after 50 galvanostatic charge/discharge cycles at 1C. (D) Spectra of the same cell after 100 galvanostatic charge/discharge cycles at 1C. For color coding of curves see the annotation in chart A. (E and F) Impedance spectra for freshly assembled and cycled cells (as in charts A–D, but with the HFeOFe-based cathode).



HEOFe, LiHEOFeF and LiHEOFeCl, respectively. Galvanostatic cycling causes a decrease of R_{CT} of both high-entropy oxyhalides, which is reminiscent of the similar behavior of the LiHEOFeCl-based cathode in an electrolyte standardly used for Li-sulfur cells (LiTFSI + LiNO₃ in 1,3 dioxolane/1,2-dimethoxyethane).²³ In contrast, the cycling of HEOFe leads to a small enhancement of R_{CT} , which is accompanied by a marked improvement of the serial resistance, R_s . At this stage of our research, we have no persuasive explanation for the difference. The mechanism of electrochemical reduction of HEOFe and the high-entropy oxyhalides can be different, which is mirrored in the specific EIS features of the former. Furthermore, the extraordinary electrochemical activity of LiHEOFeCl, observed at early stages of voltammetric and galvanostatic cycling (Fig. 6–8), is in accord with the smallest observed charge-transfer resistance of this particular material.

4. Conclusions

The phase pure high-entropy oxide (Zn_{0.25}Mg_{0.25}Co_{0.25}Cu_{0.25})Fe₂O₄ (HEOFe), lithiated oxyfluoride Li_{0.5}(Zn_{0.25}Mg_{0.25}Co_{0.25}Cu_{0.25})_{0.5}Fe₂O_{3.5}F_{0.5} (LiHEOFeF), and lithiated oxychloride Li_{0.5}(Zn_{0.25}Mg_{0.25}Co_{0.25}Cu_{0.25})_{0.5}Fe₂O_{3.5}Cl_{0.5} (LiHEOFeCl) were prepared by combination of mechanical milling and heat treatment. X-ray powder diffraction confirmed their spinel-based cubic structure (*Fd3m*) with an average crystallite size of 50–70 nm. The corresponding particle sizes from (S)TEM are 3–4 times larger due to the agglomeration of the crystallites. EDS mapping confirmed a homogeneous distribution of constituent elements in all the prepared samples. In addition, the average elemental composition determined by AAS is close to the nominal composition. Mössbauer spectroscopy revealed that Li and F or Cl ions do not influence the redistribution of Fe³⁺ cations. Photoelectron spectroscopy provided deeper insight into the electronic states of particular cations and anions.

Electrochemical measurements carried out in the 2032-coin cell with a Li-metal anode indicated voltammetric charge capacities of 450, 694, and 593 mA h g^{−1} for HEOFe, LiHEOFeCl, and LiHEOFeF, respectively. They exceed almost four times the theoretical capacity of classical intercalation electrodes of Li-ion batteries. The best electrochemical performance of LiHEOFeCl was ascribed to its smallest particle size enabling a faster electrochemical conversion reaction. Galvanostatic chronopotentiometry at 1C rate confirmed high initial charge capacities for all the samples but a capacity decay over 100 charging/discharging cycles. This decay was the most pronounced for the sample LiHEOFeCl with the highest initial charge capacity.

Raman spectroelectrochemical analysis of LiHEOFeF proved the reversibility of the electrochemical process for the initial charging/discharging cycles. Electrochemical impedance spectroscopy data agree with cyclic voltammetry and galvanostatic chronopotentiometry for all three samples. The initial charge transfer resistance of LiHEOFeCl exhibits the lowest value and decreases gradually during galvanostatic

cycling, analogous to that of LiHEOFeF. In contrast, galvanostatic cycling causes an enhancement of the charge transfer resistance of the HEOFe sample, obviously due to the different mechanisms of the electrochemical reduction.

Author contributions

Conceptualization, M. Z., L. K., and M. F.; methodology, M. Z., M. F., and L. K.; investigation, M. Z., M. F., O. P., L. K., B. P. L., V. Š., H. K., K. L. S., and M. L.; data curation, M. Z., M. F., O. P., B. P. L., H. K., and L. K.; writing – original draft preparation, M. Z., L. K., and M. F.; writing – review and editing, M. Z., L. K., M. F., and B. P. L.; project administration, M. Z., and M. F.; funding acquisition, M. Z., M. F., H. K., O. P., and V. Š. All authors have read and agreed to the published version of the manuscript.

Data availability

Data are available at the following link <https://doi.org/10.5281/zenodo.13827515>.

Conflicts of interest

There are no conflicts to declare.

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