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# Magnetic $CoFe_2O_4$ nanoparticles supported on graphene oxide ($CoFe_2O_4$/GO) with high catalytic activity for peroxymonosulfate activation and degradation of rhodamine B

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Herein, we report the preparation of magnetic $CoFe_2O_4$ nanoparticles and $CoFe_2O_4$/graphene oxide (GO) hybrids and evaluate their catalytic activity as heterogeneous peroxymonosulfate (PMS) activators for the decomposition of rhodamine B. The surface morphologies and structures of both $CoFe_2O_4$ nanoparticles and $CoFe_2O_4$/GO hybrids were investigated by powder X-ray diffraction (XRD), scanning electron microscopy (SEM), energy-dispersive X-ray spectroscopy (EDS), Fourier transform infrared spectroscopy (FTIR) and nitrogen adsorption–desorption isotherms. The magnetic properties of the samples were assessed using a SQUID magnetometer at 298 K. Catalytic oxidation experiments demonstrated that $CoFe_2O_4$/GO hybrids exhibited much better catalytic activity than $CoFe_2O_4$ nanoparticles or $CoFe_2O_4$/reduced graphene oxide (rGO) hybrids, suggesting that GO plays an important role in $CoFe_2O_4$/GO hybrids in the decomposition of rhodamine B. The influence of various reaction conditions such as temperature, concentration of PMS, pH and decomposition time of rhodamine B over the $CoFe_2O_4$/GO catalyst were investigated and optimized. The rhodamine B degradation process was found to fit a pseudo-first order kinetics model. The catalyst could be easily separated from the reaction mixture by applying an external magnet. In particular, the as-prepared $CoFe_2O_4$/GO hybrid exhibited good reusability and stability in successive degradation experiments in PMS solution.

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## 1. Introduction

The dye industry discharges large amounts of industrial wastewater and hence, it is one of the major sources of organic pollutants. Nowadays, there are more than 100 000 dyes belonging to various chemical classes with an annual production of 7.105 tons.[1] It is estimated that 10–15% of initial quantities are lost during dyeing procedures, which are discharged without prior treatment to the effluent; this contaminates groundwater and is toxic to humans and animals. Most of these compounds are chemically stable and have a complicated constitution, which makes them resistant to photo- and biological degradation. The Rhodamine (RhB) dye is one of fresh peach of synthetic dyes that is widely used as a colorant in the manufacturing of textiles and foodstuffs. It has been medically proven that rhodamine dye is harmful and toxic to humans and animals, and causes irritation of the skin, eyes and respiratory tract.[2] Due to its high toxicity and negative effects on public health, various physical, chemical and biological approaches have been extensively explored and investigated for the removal of organic dyes from wastewater including adsorption, coagulation, biological degradation and filtration processes as well as chemical oxidation by hydro chlorite and Fenton methods.[3] However, these methods suffer from different drawbacks that are primarily associated with the cost-intensive production of oxidants, instability during long reaction times, short lifetime and pH adjustments.

In recent years, sulfate radical-based oxidation processes have received much attention[4] for its efficient degradation of organic contaminants. The sulfate radical ($SO^{\bullet 4-}$) generated from peroxymonosulfate, as an alternative to the hydroxyl radical ($OH^{\bullet}$), is a strong oxidant with a high redox potential. It can react with many organic contaminants to yield a degradation performance similar to that expected for the hydroxyl radical ($OH^{\bullet}$). The activation processes of PMS can be achieved using heat, ultraviolet irradiation, transition metals, or metal oxides,[5–7] which are similar to the cases involving hydrogen peroxide. Recently, the $Co^{2+}$ ion coupled with a PMS system for the degradation of organic contaminants has attracted tremendous interest since it exhibits better efficiencies than the Fenton reaction.[8,9] Despite the advantages of this homogeneous activation process, the application of this method in water treatment is limited due to

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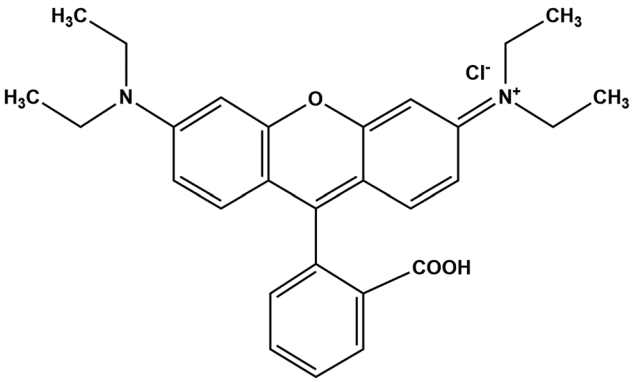
pollution caused by the high solubility and significant toxicity of transition metal ions and the complex recovery of metal ions from the reaction medium. One of the most favorable ways to overcome these drawbacks is through activation of PMS *via* heterogeneous systems,[8–10] such as Co/activated carbon,[11] Co/carbon aerogel,[12] Co/carbon xerogels,[13] $Co_3O_4$,[14–16] Co-exchanged zeolites,[17,18] Co/SBA-15,[19–21]Co/mesoporous $MnO_2$,[22] and Co/MCM-41.[23] Among the various heterogeneous catalysts, magnetic cobalt ferrites ($CoFe_2O_4$), belonging to the family of spinel-type ferrites, have attracted extensive attention due to their large surface area, high catalytic activity, stable crystalline structure and particularly their easy separation from the reaction system by utilizing magnetic fields derived from their ferromagnetic properties.[24–26] The cobalt ferrites prepared through conventional methods generally consist of highly agglomerated particles with low specific area, which reduces their catalytic performance.[27] To solve this agglomeration problem, several synthetic routes have been developed. Among the developed approaches, dispersing agglomerated particles onto the various supports was found to be an effective method to enhance the catalytic activity of $CoFe_2O_4$.[28,29] The enhancement of the catalytic activity of the $CoFe_2O_4$-supported catalyst was due to the synergic effect between $CoFe_2O_4$ and the support. The role of the support was not negligible in this case. Recently, immobilization of $CoFe_2O_4$ nanoparticles onto exfoliated graphite oxide has been the subject of intense research due to the excellent properties and functionalities of the resultant hybrid material as well as its wide spectrum of applications, which include catalysis, biomedicine and decontamination of waste water.[30–33]

In the present study, we report a facile approach for preparing magnetic $CoFe_2O_4$ nanoparticles, $CoFe_2O_4$/reduced graphene oxide (rGO) and $CoFe_2O_4$/graphene oxide (GO) and their catalytic performance toward activating PMS for the removal of rhodamine B. The physicochemical properties of all samples were characterized by various techniques, such as nitrogen adsorption–desorption, SEM, XRD and FTIR. The catalytic activities of all prepared samples were investigated in terms of the reaction kinetics, reaction temperature, concentration of RhB and catalytic stability.

## 2. Experimental

### 2.1. Materials

High purity rhodamine B ($C_{28}H_{31}ClN_2O$), oxone ($KHSO_4 \cdot K_2SO_4 \cdot KHSO_5$), iron chloride hexahydrate ($FeCl_3 \cdot 6H_2O$), cobalt chloride hexahydrate ($CoCl_2 \cdot 6H_2O$), sodium hydroxide (NaOH), graphite, sodium nitrate ($NaNO_3$), sulfuric acid $H_2SO_4$ (98% w/w), potassium permanganate $KMnO_4$, hydrogen peroxide $H_2O_2$ (30% w/w) and $C_2H_6O$ were purchased from Aldrich chemical company. All the reagents were used without further purification. Water used in all experiments was deionized. The molecular structure of rhodamine B is shown in Scheme 1.




Scheme 1 Structure of rhodamine B.

### 2.2. Preparation of GO and $CoFe_2O_4$/GO

Graphene oxide (GO) was prepared from natural graphite *via* the modified Hummers method.[34] In the preparation of $CoFe_2O_4$–GO, 0.16 g of GO was first dispersed in water and sonicated for 30 min to produce a homogeneous brown dispersion of graphene oxide nanosheets. Separately, 2.7 g of $FeCl_3 \cdot 6H_2O$ and 1.19 g of $CoCl_2 \cdot 6H_2O$ were dissolved in 10 mL of distilled water under vigorous stirring. The bimetallic Fe–Co solution was then added to the GO dispersion and the resultant mixture was further stirred for 1 h. Then, 25 mL of an aqueous solution of NaOH (3 mol $L^{-1}$) was added dropwise to the above dispersion under vigorously stirring. Later, the dispersion was heated in a sand bath for 1 h at 100 °C. Finally, the resultant precipitate was magnetically separated, washed with water and ethanol until the pH was neutral (pH = 7) and dried at 60 °C for 24 h. For comparison, $CoFe_2O_4$ NPs were also prepared according to the above procedure.

### 2.3. Characterization

XRD measurements were recorded on a Bruker AXS D-8 diffractometer using Cu-K$\alpha$ radiation in Bragg–Brentano geometry ($\theta$–$2\theta$). All samples were also characterized by Fourier-transform infrared spectroscopy in the range of 4000–400 $cm^{-1}$ using an ABB Bomem FTLA 2000 spectrometer with 16 $cm^{-1}$ resolution. SEM and STEM micrographs were obtained on a Tecnai G2 microscope at 120 kV. The elemental composition of the $CoFe_2O_4$/GO nanocomposite was confirmed from energy dispersive X-ray analysis (EDAX). The surface areas of the prepared materials were measured using the Brunauer–Emmett–Teller (BET) method on a 3Flex automatic analyzer. Prior to $N_2$ sorption, all samples were degassed at 250 °C for 8 h. The magnetic properties of $CoFe_2O_4$ nanoparticles and $CoFe_2O_4$/GO nanocomposite were investigated in a MPMS-XL-7AC superconducting quantum interference device (SQUID) magnetometer. The magnetic measurements were performed from −15 000 to 15 000 Oe at room temperature. Total organic carbon (TOC) was determined by the Shimadzu TOC-L Series. Cobalt and iron within $CoFe_2O_4$/GO were determined by inductively coupled plasma atomic emission spectroscopy (ICP-AES) from Jab in Yvan and the carbon content was determined by a carbon/sulfur analyser (C.A) using HORIBA EMIA-320V2.

### 2.4. Catalytic test procedure

The catalytic degradation of RhB by $CoFe_2O_4$ and $CoFe_2O_4$/GO catalysts with oxone was performed in a 50 mL beaker

containing 50 mL of RhB solution at room temperature (25 °C). In a typical procedure, 0.005 g of oxone was first added to RhB solution under constant stirring. Then, 0.010 g of catalyst was added to start the reaction. At a given time interval, a predetermined amount (2.0 mL) of solution was withdrawn into a vial fitted with a micro-filter (45 μm) for solid catalyst removal. The concentrations of RhB were determined by monitoring the decrease in absorbance at the maximum wavelength (554 nm) with UV-vis spectroscopy.

The degradation efficiency was calculated according to the following equation:

$$\text{Degradation}(\%) = \frac{(C_0 - C) \times 100}{C_0} = \frac{(A_0 - A) \times 100}{A_0}$$

where, $C_0$ and $C$ are the initial and final concentrations of RhB, respectively, and $A_0$ and $A$ represent the initial and the final absorbance of RhB at 554 nm, respectively.

## 3. Results and discussion

XRD was employed to analyze the crystalline phases of as-prepared GO, $CoFe_2O_4$ and $CoFe_2O_4$/GO samples (Fig. 1). The XRD patterns (Fig. 1a) confirmed the successful oxidation of natural graphite to graphite oxide, which exhibits a strong diffraction peak at $2\theta = 10.28$, corresponding to the (001) interplanar spacing of 0.87 nm, which is much larger than the $d$-spacing of graphite (0.34 nm). It is evident that $CoFe_2O_4$ and $CoFe_2O_4$/GO exhibit similar XRD patterns. The diffraction peaks for the two samples at $2\theta = 30.3$, 35.6, 43.3, 53.6, 57.1 and 62.8 are consistent with the (220), (311), (400), (422), (333) and (440) reflections, respectively, of the cubic spinel-type structure of $CoFe_2O_4$ (JCPDS 75-0033). It should be noted that typical diffraction peaks of GO could not be detected in the XRD pattern of $CoFe_2O_4$/GO, suggesting the destruction of the regular layered structure of GO due to the crystal growth of $CoFe_2O_4$ between its inter-layers. The average crystallite size of $CoFe_2O_4$ and $CoFe_2O_4$/GO, estimated from the Debye–Scherrer formula, are 10.37 and 11.06 nm, respectively.

The FT-IR spectra of GO, $CoFe_2O_4$ and $CoFe_2O_4$/GO are shown in Fig. 2. As illustrated in Fig. 2, several characteristic bands of the functional groups of GO can be observed. The two peaks located at 1725 and 1618 $cm^{-1}$ are assigned to the anti-symmetric and symmetric stretching vibrations of carboxylic groups ($COO^-$). The absorption peaks at 1035 $cm^{-1}$ is attributed to the stretching vibrations of alkoxy C–O from the oxygen-containing functional groups such as carbonyl, carboxylic and epoxy groups.[35] Indeed, the FTIR spectrum of $CoFe_2O_4$ shows a peak at 577 $cm^{-1}$, which is ascribed to the Co–O and Fe–O vibrations. It is worth noting that in the spectrum of $CoFe_2O_4$/GO, the peaks at 1618 $cm^{-1}$, corresponding to the functional groups of COO in GO, shifts to 1567 $cm^{-1}$.

The calculated textural parameters of $CoFe_2O_4$ and $CoFe_2O_4$/GO are summarized in Table 1. From the data in Table 1, it can be seen that the BET surface area of $CoFe_2O_4$ is 124 $m^2\ g^{-1}$ and its pore volume is 0.1662 $cm^3\ g^{-1}$, while the surface area of the $CoFe_2O_4$/GO is 142 $m^2\ g^{-1}$ and its pore volume is 0.1868 $cm^3\ g^{-1}$. The addition of GO nanosheets should be the responsible for this increase in the surface area of $CoFe_2O_4$/GO compared to that of $CoFe_2O_4$. The nitrogen sorption isotherm of $CoFe_2O_4$, as shown in Fig. 3, is of type III with a distinct hysteresis loop of type H2 in the relative pressure range of 0.5 and extending almost to 1, which is characteristic of mesoporous textural porosity. The corresponding pore size distribution curve indicates that $CoFe_2O_4$ exhibited a pore size distribution in the range of 2 to 10 nm with a central value of 6.5 nm, corresponding to mesoporous materials. As for $CoFe_2O_4$/GO, there is a slight difference as the nitrogen sorption isotherms of $CoFe_2O_4$/GO are of the type III (Fig. 3) and the hysteresis loop is of type H3 according to the IUPAC classification. The corresponding pore size distribution curve indicates that $CoFe_2O_4$/GO has a centralized pore size distribution within two areas toward 4 and 6 nm, which confirm the existence of textural mesopores.

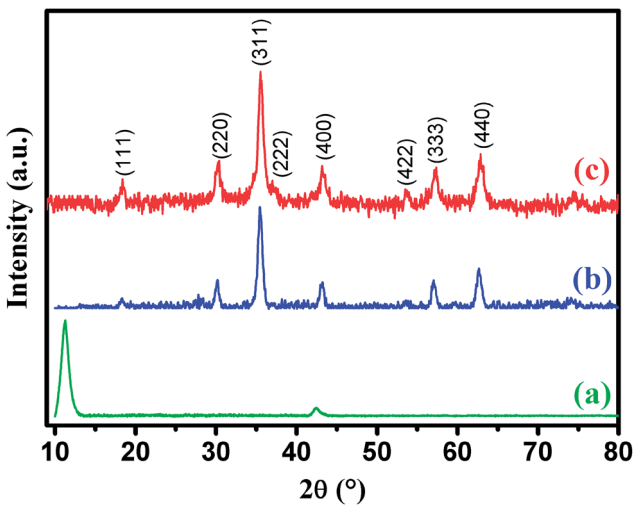



Fig. 1 XRD patterns of graphene oxide (a), $CoFe_2O_4$ (b) and $CoFe_2O_4$/GO (c).

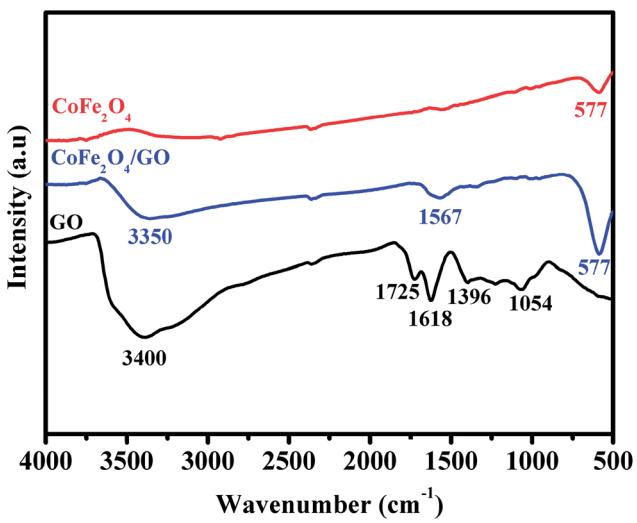



Fig. 2 FT-IR spectra of GO, $CoFe_2O_4$ and $CoFe_2O_4$/GO.

Table 1 Lattice parameter, crystallite size, BET surface area and pore volume of the $CoFe_2O_4$ and $CoFe_2O_4$/GO samples

| Samples | Lattice parameter $a$ (Å) | Cell volume (Å$^3$) | Crystallite size[a]/nm | $S_{BET}$[b]/m$^2$ g$^{-1}$ | $V_{meso}$[b]/mL g$^{-1}$ | $D_{meso}$[b]/nm |
|---|---|---|---|---|---|---|
| $CoFe_2O_4$ | 8.399 | 592.50 | 10.37 | 124 | 0.1662 | 6.5 |
| $CoFe_2O_4$/GO | 8.354 | 583.02 | 11.06 | 142 | 0.1868 | 4.6 |

[a] Calculated by the Debye–Scherer equation. [b] From nitrogen sorption analysis.



The surface morphology and elemental composition of the as-prepared $CoFe_2O_4$ and $CoFe_2O_4$/GO were investigated by SEM and EDX. From the SEM images, as shown in Fig. 4, it could be observed that $CoFe_2O_4$ exhibits a heterogeneous microstructure that consisted of crystallites of various sizes. It should be noted that particles of $CoFe_2O_4$ are strongly agglomerated as shown in the FE-SEM images in (Fig. 4b), which can be attributed to the powerful inherent magnetic interaction of $CoFe_2O_4$ magnetic particles. Furthermore, as shown in Fig. 4c and d, when $CoFe_2O_4$ was supported on the surface of GO nanosheets, the agglomeration phenomenon reduced, suggesting that GO can prevent the aggregation of the $CoFe_2O_4$ nanoparticles. Moreover, energy dispersive X-ray (EDX) analyses were recorded and are shown Fig. 4e. The EDX analysis confirmed the elemental composition of the $CoFe_2O_4$–GO material, which primarily consists of C, Co, Fe, and O. The chemical composition of the as-synthesized products was further analyzed by inductively coupled plasma atomic emission spectroscopy (ICP-AES). Table 2 shows the elemental composition of $CoFe_2O_4$/GO measured by EDX, ICP-AES and C.A. The stoichiometric elemental composition for Co, Fe, and O is 22.79, 44.44, and 4.32%, respectively. In all samples, Co and Fe are in the ratio of 1 : 2, which further confirm the structure of $CoFe_2O_4$/GO.

The morphology of GO and $CoFe_2O_4$/GO was characterized by STEM. From Fig. 5(a and b), it can be observed that $CoFe_2O_4$ NPs prepared in the absence of GO show spherical particles with severe aggregation due to the magnetic dipolar interaction among the magnetite NPs. In comparison, uniform $CoFe_2O_4$ NPs are deposited and well-dispersed on the basal planes of graphene (Fig. 5(c and d)). In addition, we noticed that $CoFe_2O_4$ NPs were still tightly anchored on the surface of GO even after sample preparation (mechanical stirring and sonication) for STEM analysis, suggesting a strong interaction between $CoFe_2O_4$ NPs and GO. Moreover, the graphene sheets distributed between the $CoFe_2O_4$ NP can prevent the aggregation of

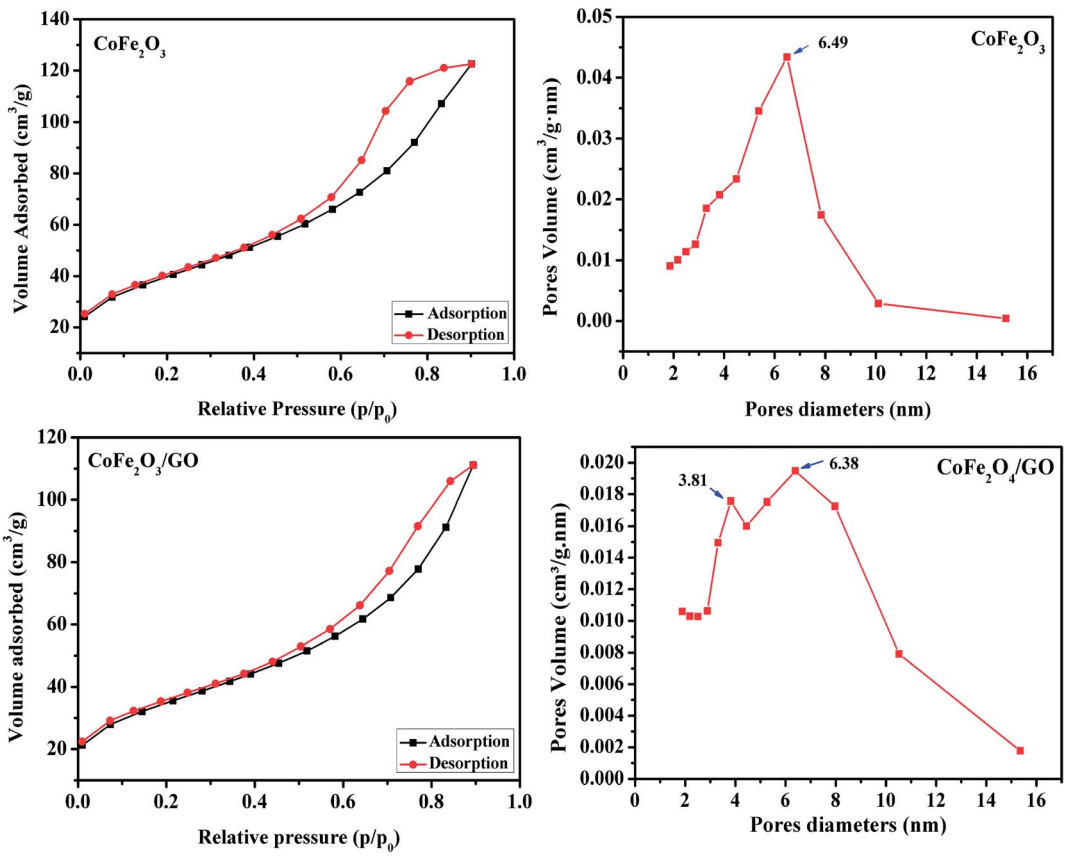



Fig. 3 Nitrogen adsorption/desorption isotherms and BJH pore size distribution of $CoFe_2O_4$ and $CoFe_2O_4$/GO.

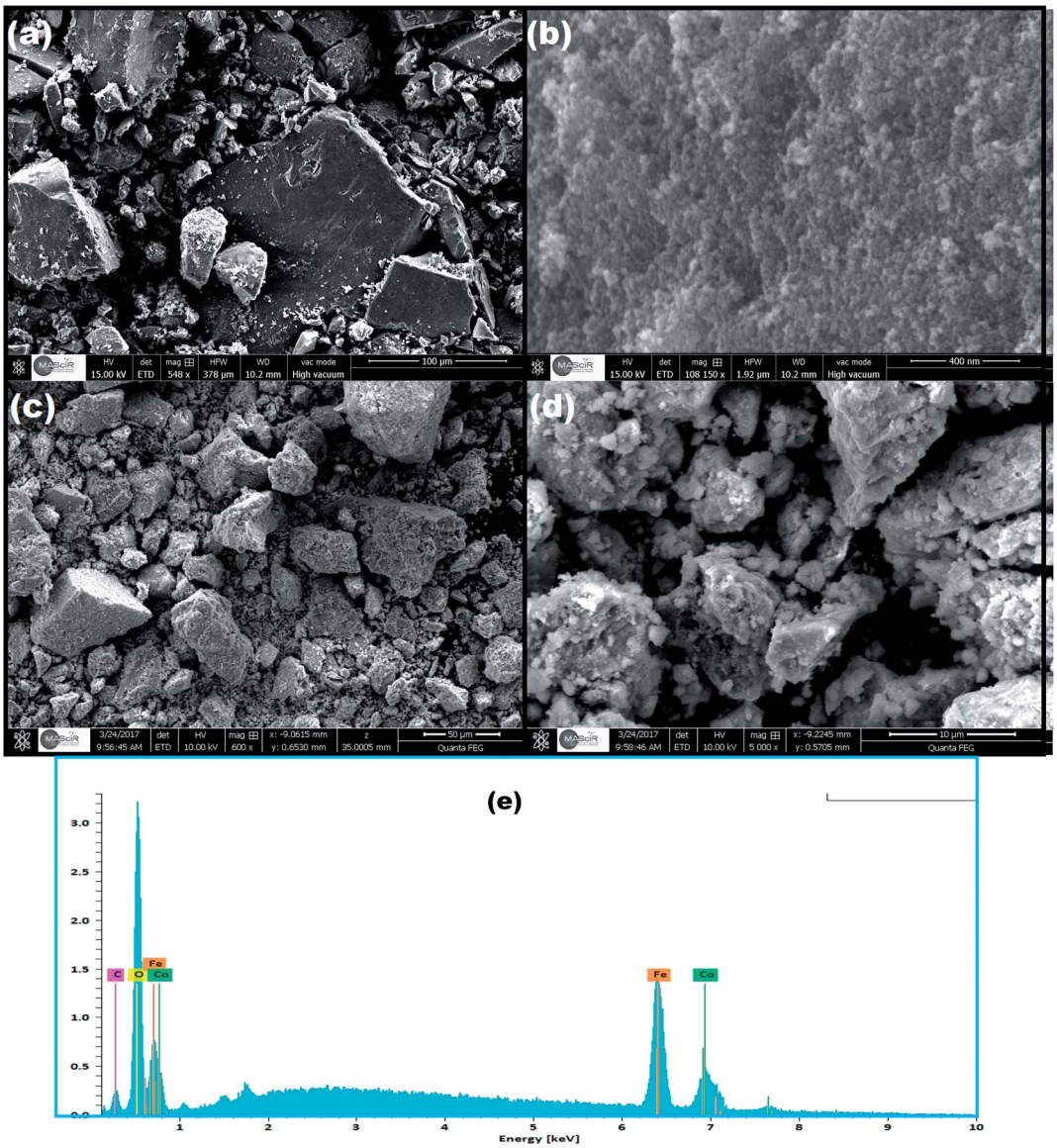



Fig. 4 SEM images of (a, b) $CoFe_2O_4$ and (c, d) $CoFe_2O_4$/GO, (e) EDX of $CoFe_2O_4$/GO.

Table 2 Elemental composition of $CoFe_2O_4$/GO measured by EDX, ICP-AES and C.A

| | Weight (%) | | |
|---|---|---|---|
| Element | EDX | ICP-AES | C.A |
| Co | 22.15 | 22.79 | — |
| Fe | 45.94 | 44.44 | — |
| C | 4.24 | — | 4.32 |
| O | 37.65 | — | — |

$CoFe_2O_4$ NP to a certain extent, which can be of great benefit to reactions.

The magnetic properties of $CoFe_2O_4$ and $CoFe_2O_4$/GO were investigated by SQUID from −15 000 to 15 000 Oe at room temperature as shown in Fig. 6. The saturation magnetization (Ms) of $CoFe_2O_4$ is close to 47.33 emu $g^{-1}$, which is higher than that corresponding to the $CoFe_2O_4$/GO material (30.15 emu $g^{-1}$). This increase in Ms value could be due to (i) the existence of the GO nanosheets, (ii) the surface defect of $CoFe_2O_4$ crystallites and (iii) the strong interfacial interaction between $CoFe_2O_4$ nanoparticles and GO surfaces. Moreover, the field-dependent magnetization curves of $CoFe_2O_4$ and $CoFe_2O_4$/GO show non-negligible remanence (Mr) and coercivity (Hc), indicating paramagnetic behavior of the two samples at room temperature (Table 3). Furthermore, $CoFe_2O_4$/GO hybrids can be easily separated from the reaction by applying an external magnetic field. The $CoFe_2O_4$/GO catalyst can be dispersed in deionized water to form a stable brown homogenous suspension before magnetic separation (inset of Fig. 6a). However, when a magnet was placed close to the reaction vessel, it could be observed that the synthesized samples were rapidly attracted,

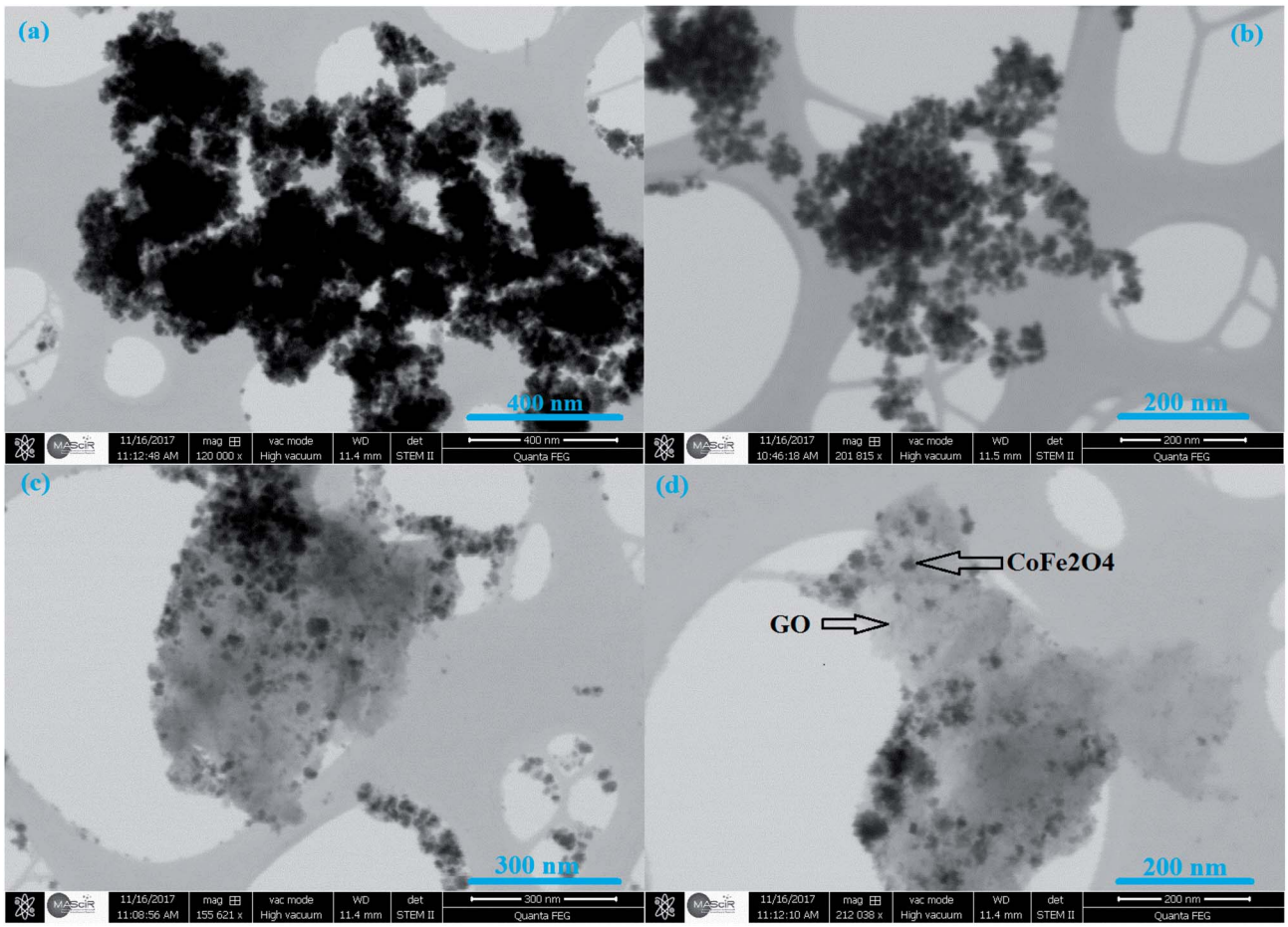



Fig. 5 STEM images of $CoFe_2O_4$ (a, b) and $CoFe_2O_4$/GO (c, d).

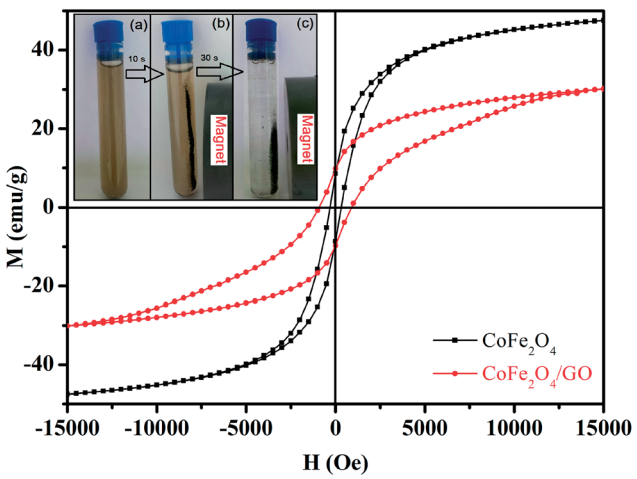



Fig. 6 Magnetization curves of $CoFe_2O_4$ and $CoFe_2O_4$/GO. The inset shows the photographs of the separation processes of $CoFe_2O_4$/GO: (a) without external magnetic field, and (b, c) with external magnetic field.

and a nearly colorless solution was obtained (inset of Fig. 6(b and c)).

To evaluate the catalytic performance of our materials, the degradation of RhB in the presence of the PMS was selected as a model catalytic reaction. The degradation kinetics of RhB in aqueous solution was studied by monitoring the decrease of its absorption peak at 554 nm in the UV-vis spectra with time as exemplified in Fig. 7a. No significant shift in the absorption peak ($\lambda_{max}$ = 554 nm) was observed before 12 min except for a quick reduction in the absorbance at 554 nm. Since the absorbance at 554 nm was primarily attributed to RhB remaining in the solution under our experimental conditions, the linear relationship between the concentration of the residual RhB and its absorbance at 554 nm was further verified by the standard calibration curve ($R^2$ = 0.9970), and the corresponding absorbance was then converted to the RhB concentration in the solution used for kinetic analysis. To further investigate the degradation of RhB, total organic carbon (TOC), which has been widely used to evaluate the degree of

Table 3 The values of saturation magnetization (Ms), remanent magnetization (Mr) and coercivity (Hc) of $CoFe_2O_4$ and $CoFe_2O_4$/GO, extracted from Fig. 5

| Sample | Ms (emu $g^{-1}$) | Mr (emu $g^{-1}$) | Hc (Oe) |
|---|---|---|---|
| $CoFe_2O_4$ | 47.33 | 8.55 | 348.53 |
| $CoFe_2O_4$/GO | 30.15 | 9.78 | 897.16 |

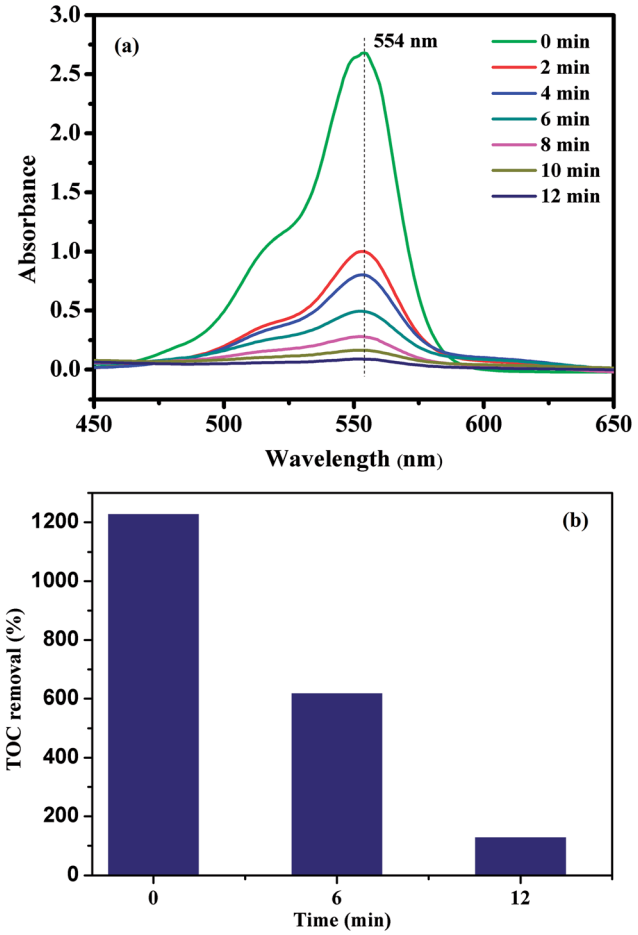



Fig. 7 UV-vis spectra of RhB solution at different times using $CoFe_2O_4$/GO as a catalyst (a); the TOC removal efficiency of RhB using $CoFe_2O_4$/GO as a catalyst.

mineralization of organic species, was measured in the RhB degradation process by the $CoFe_2O_4$/GO/PMS system as shown in Fig. 7b. The TOC removal efficiency of RhB reached 89.34% after 12 min in the presence of the $CoFe_2O_4$/GO/PMS system, which confirms that RhB could be mineralized to residual organic molecules by the as-prepared samples.

As shown in Fig. 8, no RhB degradation was observed *via* PMS oxidation alone. Similarly, $CoFe_2O_4$, $CoFe_2O_4$/rGO and $CoFe_2O_4$/GO cannot catalyze RhB degradation in the absence of PMS, which reveals that the contribution from simple physical adsorption is negligible in this case. Moreover, in the absence of a catalyst, the concentration of RhB remained unchanged over time, suggesting that it is difficult for the degradation of RhB to proceed without a catalyst. Therefore, the degradation of RhB is very sensitive to the presence of both PMS and catalyst in the reaction system. The catalytic performances of $CoFe_2O_4$, $CoFe_2O_4$/rGO and $CoFe_2O_4$/GO for the degradation of RhB with PMS were clearly different in the three samples and the degradation efficiencies were 78, 90 and 98%, respectively. Moreover, it can be clearly observed that the degradation rate of RhB over $CoFe_2O_4$/GO was much faster than that corresponding to $CoFe_2O_4$/rGO and $CoFe_2O_4$ and it took around 12 min for

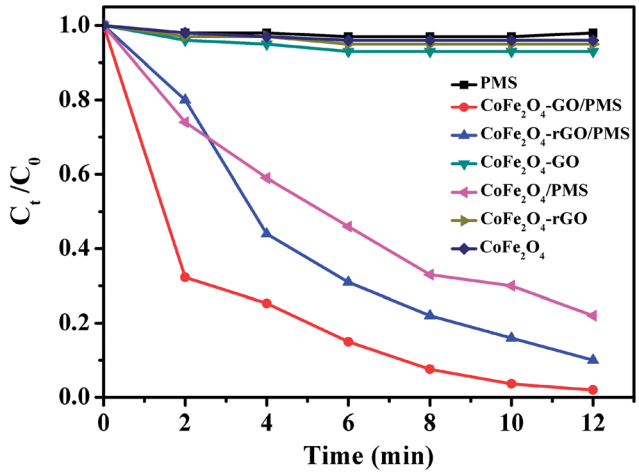



Fig. 8 RhB degradation by various catalytic PMS systems. Conditions: $[RhB]_0$ = 0.03 mmol $L^{-1}$, [PMS] = 0.10 mg $L^{-1}$, amount of catalyst = 10 mg, temperature = 25 °C.

complete removal of RhB. This superior catalytic activity of $CoFe_2O_4$/GO could be related to the electronic structure and the presence of functional hydroxyl groups in GO, which could be involved in the degradation mechanism, thus enhancing the catalytic activity of the $CoFe_2O_4$/GO catalyst for the degradation of RhB. As compared with $CoFe_2O_4$, GO can offer an environment to prevent aggregation of $CoFe_2O_4$ nanoparticles and also a higher surface area (142 $m^2$ $g^{-1}$) (124 $m^2$ $g^{-1}$), which can provide more active sites for catalytic degradation of RhB.

Fig. 9 shows the linear kinetic fitting plots of $\ln(C_0/C_t) = f(t)$ for RhB photodegradation, in which $C_0$ and $C_t$ are the initial concentration of RhB and its concentration at time $t$, respectively. In general, the degradation of organic dyes obeys a pseudo-first order kinetics model. As shown in Fig. 9, RhB degradation by $CoFe_2O_4$, $CoFe_2O_4$/GO and $CoFe_2O_4$/rGO can follow a pseudo-first order kinetics model. Furthermore, the constant rate values indicate that $CoFe_2O_4$/GO possesses

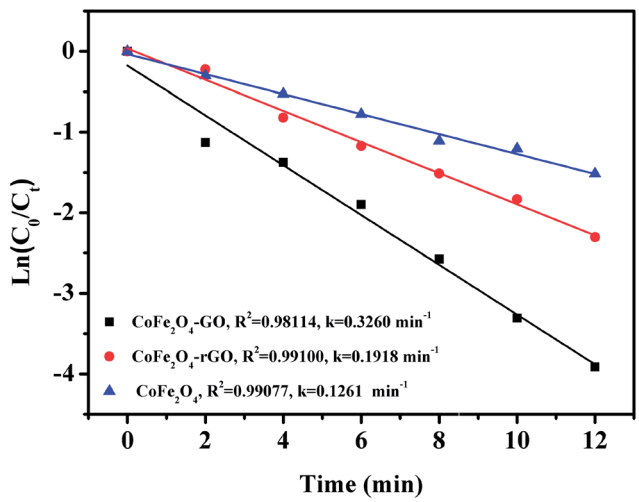



Fig. 9 Linear kinetic fitting plots of $\ln(C_0/C_t) = f(t)$ based on a first-order kinetic model.

a better rate constant (0.3260 $min^{-1}$) than $CoFe_2O_4$/rGO (0.1918 $min^{-1}$) and $CoFe_2O_4$ (0.1261 $min^{-1}$), suggesting that graphene oxide plays a significant role in the enhancement of PMS catalytic degradation of RhB. The strong interfacial interaction between graphene oxide and $CoFe_2O_4$ generates a synergistic function of large surface area and improves the electron transport ability and chemical reaction sites.

The effect of oxone concentration on RhB degradation was studied and the results are presented in Fig. 10. It is worth mentioning that increasing oxone concentration from 0.02 to 0.20 g $L^{-1}$ led to a faster and more efficient degradation of RhB from 54.4 to 99%. Meanwhile, the kinetic rate constant also increased from 0.0654 to 0.1918 $min^{-1}$. These results could be explained by the high concentration of free sulfate radicals formed at higher oxone concentration, which result in an increase in the rate of RhB degradation. Therefore, the optimal oxone concentration was about 0.1 g $L^{-1}$ for the degradation of RhB on the $CoFe_2O_4$/GO catalyst.

The effect of reaction temperature on RhB degradation was also investigated by varying the reaction temperatures from 20 to 30 to 40 °C; the experimental findings are presented in Fig. 11. It can be observed that the degradation rate of RhB increases from 0.135 to 1.085 $min^{-1}$ with increasing temperature from 20 to 40 °C. These results could be explained by the high concentrations of free sulfate radicals generated at high temperature.

To further evaluate the effect of the $CoFe_2O_4$/GO nanocatalyst on the process kinetics, important parameters associated with the energetic aspects of the reaction, such as activation energy, ($E_a$) play a crucial role. The activation energy ($E_a$) of the degradation of RhB over the $CoFe_2O_4$/GO catalyst was evaluated by plotting $\ln(k)$ *versus* $1000/T$ according to the Arrhenius equation of $\ln(k) = \ln(A) - E_a/RT$, where $k$ is the rate constant, $R$ is the universal gas constant (8.314 J $mol^{-1}$ K) and $A$ is the pre-exponential. The $E_a$ value of 79.24 kJ $mol^{-1}$ was obtained from the slope of the fitted equation by linear regression

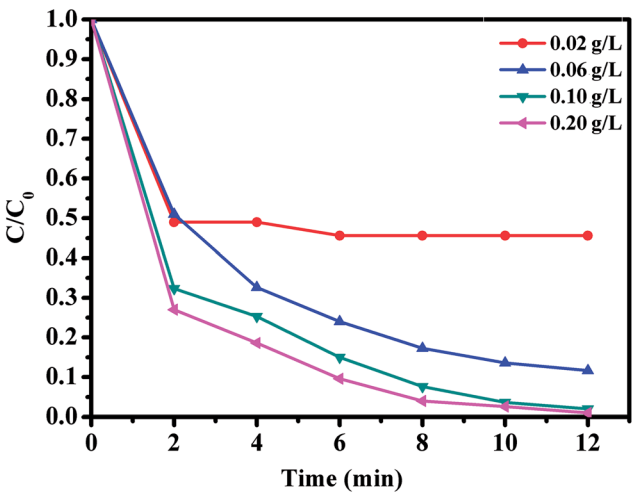



Fig. 10 Effect of oxone concentration on RhB degradation over $CoFe_2O_4$/GO. Reaction conditions: [RhB] = 0.03 mmol $L^{-1}$, amount of catalyst = 10 mg, $T$ = 25 °C.

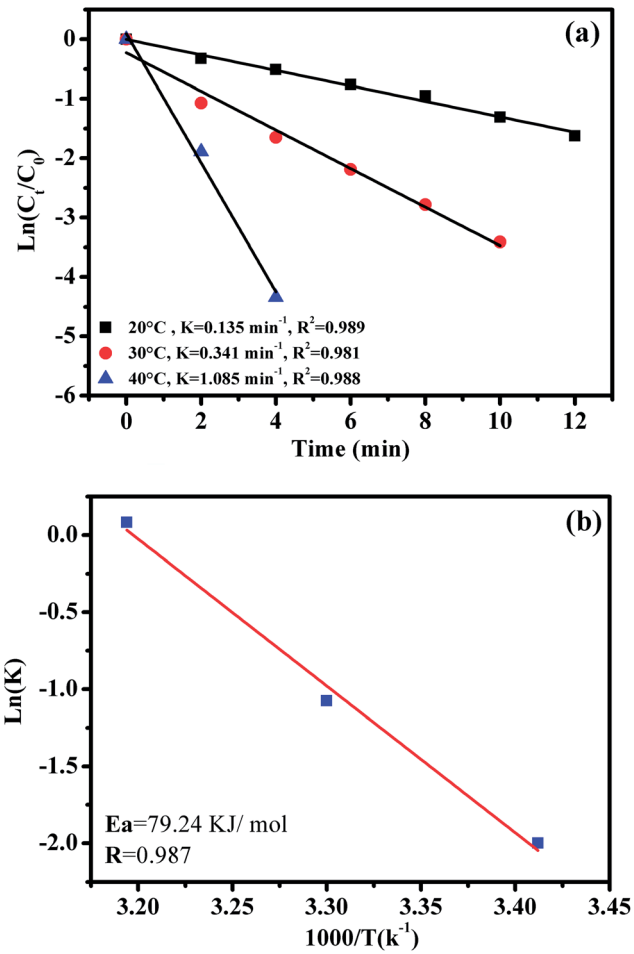



Fig. 11 The effect of reaction temperature on RhB degradation over the $CoFe_2O_4$/GO catalyst (a); the Arrhenius plot with linear regression (b); reaction conditions: [RhB] = 0.03 mmol $L^{-1}$, amount of catalyst = 10 mg and 0.1 g $L^{-1}$ of oxone at different temperatures.

($R^2$ = 0.987). This $E_a$ value is comparable to those of the highest active heterogeneous catalysts ever reported, *e.g.*, Co/active carbon (59.7 kJ $mol^{-1}$),[11] $Co_3O_4/SiO_2$ (61.7–75.5 kJ $mol^{-1}$)[36] and Co/ZSM-5, (69.7 kJ $mol^{-1}$),[17] which indicates that $CoFe_2O_4$–GO can be considered a promising heterogeneous catalyst for the PMS oxidation process.

The effect of RhB concentration was studied by varying the concentration of RhB and the experimental findings are presented in Fig. 12. Upon decreasing RhB concentration from 0.01 to 0.03 mmol $L^{-1}$, the degradation rate of RhB over $CoFe_2O_4$/GO increased significantly. Indeed, RhB was almost completely removed within 6 min at the RhB concentration of 0.01 mmol $L^{-1}$, while it was removed within 12 min at the concentration of 0.03 mmol $L^{-1}$.

The influence of initial pH values on RhB degradation over the $CoFe_2O_4$/GO catalyst was explored by adjusting the solution pH to 3.5, 6, 7, 8, 10 and 11. As shown in Fig. 13, the initial pH has a significant influence on the degradation efficiency of RhB. Based on the results obtained, we observed that the RhB degradation conducted at an initial pH of 3.5 and 7.0 was faster (decolorization efficiency of 98% at about 12 min) due to the

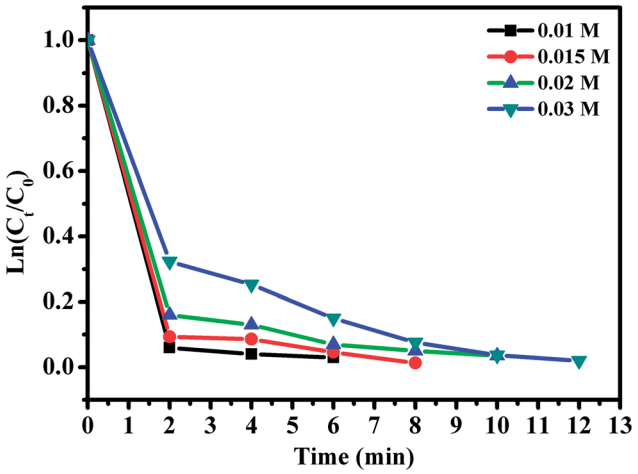



Fig. 12 Effect of RhB concentration. Reaction conditions: [PMS] = 0.1 mmol $L^{-1}$; amount of catalyst = 10 mg, $T$ = 25 °C.

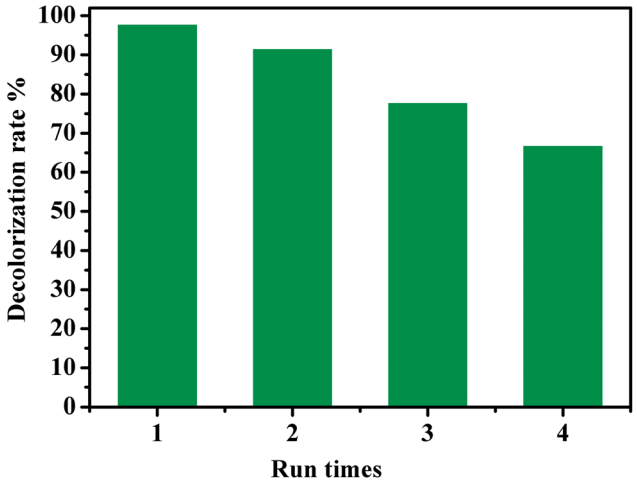



Fig. 14 Reuse performance of the $CoFe_2O_4$/GO catalyst in RhB degradation. Reaction conditions: [RhB] = 0.03 mmol $L^{-1}$, [PMS] = 0.10 mmol $L^{-1}$, amount of catalyst = 10 mg, $T$ = 25 °C.

electrostatic attraction between the negative charge of the $CoFe_2O_4$/GO catalyst at low pH (Fig. 13c) and the positive charge of RhB. It is also noteworthy that there was no obvious impact on the RhB degradation when the initial pH value changed in the range of 3.5–7. Moreover, at high initial pH values of 10.0 and 11.0, the decolorization rates were 40% and 17%, respectively, which can be explained by the deprotonation of the carboxyl group of RhB and the transformation of the cationic form of RhB into zwitterionic form.

The recyclability of a catalyst is advantageous for its commercialization and industrialization. To explore the reusability of our catalytic system, $CoFe_2O_4$/GO nanocomposite was employed as a recyclable catalyst in the degradation of RhB by PMS over four cycles. After each cycle, the catalyst was easily separated by an external magnet and washed successively by water and dichloromethane. As shown in Fig. 14, it was found that the catalytic performance of the recovered catalyst remained nearly the same for the second successive run. Although the catalytic activity slightly diminished, 67% of decolorization rate was still achieved in the fourth run, indicating that $CoFe_2O_4$/GO nanocatalysts exhibited good

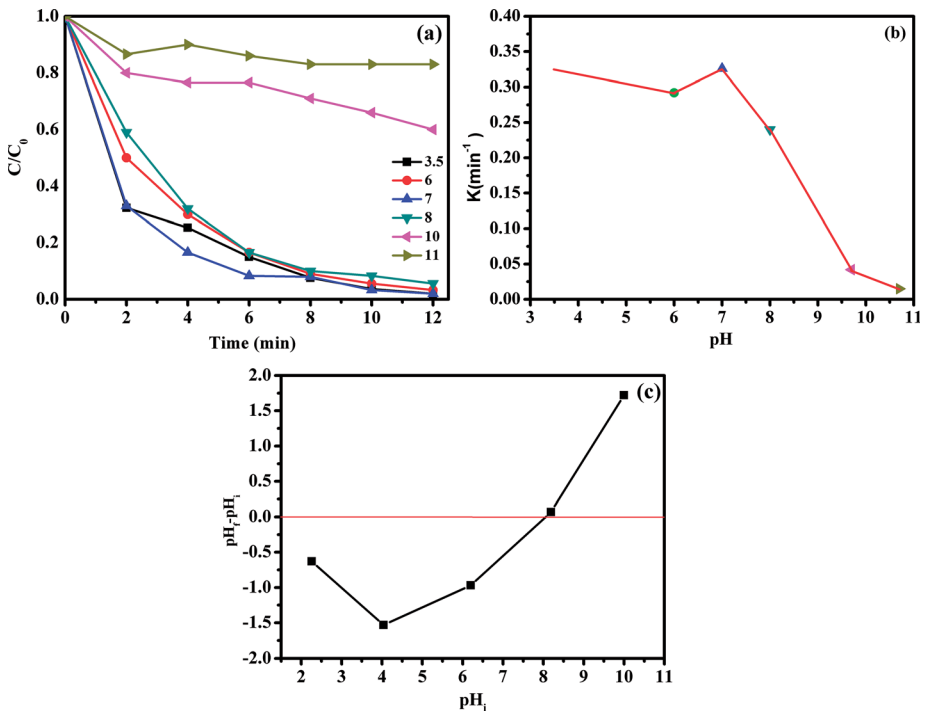



Fig. 13 Effect of pH on RhB degradation over the $CoFe_2O_4$/GO catalyst (a, b); point zero charge of $CoFe_2O_4$/GO. Reaction conditions: [RhB] = 0.03 mmol $L^{-1}$, [PMS] = 0.10 mmol $L^{-1}$, amount of catalyst = 10 mg, $T$ = 25 °C.

recyclability. Metal leaching was studied by ICP-AES analysis of the $CoFe_2O_4$/GO catalyst after the reaction. The Co and Fe concentrations in the catalyst was 20.96% and 39.51%, respectively, after one catalytic cycle of RhB degradation, which confirms negligible metal leaching.

## 4. Conclusion

In summary, magnetic $CoFe_2O_4$ and $CoFe_2O_4$/GO catalysts were successfully prepared *via* a facile approach. The physico-chemical properties of these materials were evaluated by FT-IR, XRD, SEM, and BET. These catalysts showed potential capability for catalytic degradation of rhodamine B using PMS as an oxidant. Catalyst screening revealed that $CoFe_2O_4$/GO exhibited superior catalytic activity for the removal of RhB when compared with $CoFe_2O_4$ and $CoFe_2O_4$/rGO. Then, we studied the effect of several parameters on RhB degradation. It was found that the degradation rate was dependent on pH, temperature, the concentration of oxone and the initial concentration of RhB. Furthermore, this catalyst can be easily separated by an external magnet and reused.

## Conflicts of interest

There are no conflicts to declare.

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