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Electron-triggered processes in halogenated carboxylates: dissociation pathways in CF₃COCl and its clusters†

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Trifluoroacetyl chloride, CF₃COCl, is produced in the Earth's atmosphere by photooxidative degradation of hydrochlorofluorocarbons, and represents a potential source of highly reactive halogen radicals. Despite considerable insight into photochemistry of CF₃COCl, its reactivity towards electrons has not been addressed so far. We investigate the electron ionization and attachment in isolated CF₃COCl molecules and $(CF_3COCI)_N$, max. $N \ge 10$, clusters using a molecular beam experiment in combination with quantum chemical calculations. The ionization of the molecule at 70 eV electron energy leads to strong fragmentation: weakening of the C-C bond yields the CF_3^+ and $COCl^+$ ions, while the fission of the C-Cl bond produces the major CF₃CO⁺ fragment ion. The cluster spectra are dominated by $M_n \cdot COCl^+$ and $M_n \cdot CF_3CO^+$ ions (M = CF₃COCl). The electron attachment at energies between 1.5 and 11 eV also leads to the dissociation of the molecule breaking either the C-Cl bond at low energies below 3 eV yielding mainly Cl⁻ ions, or dissociating the C-C bond at higher energies above 4 eV leading mainly to CF_3^- ions. In the clusters, the intact M_n^- ions are stabilized after electron attachment at low energies with contribution of M_{ρ} ·Cl⁻ fragment ions. At higher energies, the M_{ρ} ·Cl⁻ fragments dominate the spectra, and C-C bond dissociation occurs as well yielding $M_0 \cdot CF_3^-$. Interestingly, $M_0 \cdot Cl_2^-$ ions appear in the spectra at higher energies. We briefly discuss possible atmospheric implications.

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1 Introduction

Certain ozone-depleting chlorofluorocarbons (CFCs) nowadays have been largely replaced by hydrochlorofluorocarbons (HCFCs). Chemicals within this class of compounds are considered to be acceptable temporary alternatives to CFCs^{1,2} (according to the Montreal Protocol, the definite phase-out of HCFCs is foreseen in 2030). HCFCs undergo further reactions in the atmosphere. Their photooxidative degradation produces trifluoroacetyl chloride (TFAC, CF₃COCl).^{3,4} Since CF₃COCl is a potential source of highly reactive halogen radicals, its

presence in the atmosphere creates a need for understanding the dissociation pathways of this molecule.

A lot of attention has been dedicated recently to the dissociation of CF₃COCl induced by UV photons³⁻¹¹ and its photochemical pathways are thus well understood. Its structure as well as vibrational spectra were also analyzed both experimentally 12-14 and theoretically. 15 There is, however, no information about electron-induced dissociation of CF3COCl even the electron ionization mass spectrum is not available in the NIST database. At the same time, energetic particle precipitation influences chemical changes in the atmosphere.¹⁶ Each type of naturally occurring ionizing radiation - galactic cosmic rays, solar energetic particles and energetic electron precipitation - produces an avalanche of secondary electrons¹⁶ which can effectively induce molecular dissociation. ¹⁷ For example, cosmic-ray driven electron-induced reactions of CFCs have been proposed to contribute considerably to ozone depletion. 18,19 They have been even suggested to be a driving force behind climate change, 20,21 however, this hypothesis has met with strong criticism.^{22,23} These proposed roles of halogenated species motivate the need for addressing our complete lack of knowledge about the electron-induced chemistry of HCFCs and the current target, TFAC. In this respect, CF₃COCl

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is interesting also from the chemical perspective due to the presence of five highly electronegative atoms, contributing to its considerable reactivity and possible instability upon ionization/electron attachment.

Similarly, there is no information available about how the aggregation of CF₃COCl influences its dissociation. This is relevant since heterogeneous processes, especially occurring on aerosol nanoparticles, are crucial for chemical transformations in the atmosphere. In laboratory investigations, such aerosols can be mimicked by clusters in molecular beams.²⁴ Apart from elucidating the atmospheric pathways, the clusterbeam studies offer the possibility of probing the gas-phase and aggregated molecules with exactly the same experimental techniques which facilitates the identification of dissociation mechanisms.25

In this study, we investigate the electron-induced dissociation of isolated CF₃COCl molecules and their clusters. We probed both the cationic pathways, induced by electrons with 70 eV kinetic energy, and anionic pathways, induced by the attachment of slow (less than 11 eV) electrons. We focus especially on the dissociative processes and how they are influenced by clustering. The experiments are complemented by quantum chemical calculations to provide insight into the structure, energetics and dissociation pathways of both singlemolecular ions and clusters.

2 Methods

2.1 Experiment

The experiments were performed with the cluster-beam (CLUB) apparatus in Prague, the general description of which can be found in our recent reviews²⁴⁻²⁶ and references cited therein. The beams of CF₃COCl molecules and clusters in a vacuum were generated by continuous supersonic expansions of CF₃COCl with helium and argon buffer gas, respectively. A 5% (by the partial pressure) mixture of CF₃COCl with the buffer gas was prepared in a pressurized cylinder and the gas mixture was expanded at a constant pressure p_S through a conical nozzle (50 µm diameter, 30° full opening angle, and 2 mm long). Different expansion pressures p_S were exploited for optimum cluster generation discussed below.

About 2.5 cm downstream from the nozzle, the beam passed through a skimmer (0.8 mm diameter) into the next differentially pumped vacuum chamber and subsequently passed through two additional chambers on a flight path of about 1.5 m before entering a reflectron time-of-flight mass spectrometer (TOF-MS) chamber for the cluster ionization and detection. Our TOF-MS was first described elsewhere. 27,28 It can detect either positive ions as described in the above cited publications, or it can work in the negative ion mode.^{29,30} Here, we use both modes. The clusters are ionized by an electron beam pulsed at 10 kHz frequency. The positive ions were generated by 70 eV electrons. Since we have not found the mass spectrum of the CF₃COCl molecule in the NIST database, where the standard mass spectra measured at 70 eV electron

energy can be found, we have measured the mass spectrum at this energy to provide the standard spectrum. The ionization pulse width was 5 µs. After a 0.5 µs delay, the ions were extracted by a 4 kV pulse and further accelerated to 8 keV into the time-of-flight region. In the negative ion mode, the electron energy was scanned from 0 to 11 eV in steps of 0.2 eV. The same settings were used with the voltages of opposite polarities. After the flight path of approximately 95 cm in the reflectron TOF-MS, the ions were detected with a microchannel plate detector and the mass spectra were recorded.

It ought to be mentioned that our electron source was not designed specifically for producing low-energy electrons, therefore our electron energy dependent ion yields provide reliable values at energies higher than approximately 1.5 eV. 29,30 The electron energy scale was calibrated using 4.4 eV and 8.2 eV resonances in dissociative electron attachment to CO2 molecules.

From the combination of the positive and negative ion mass spectra generated from the same cluster beam, some information about the neutral cluster size distribution can be derived (see the ESI†). For the present experiment we estimate that the $(CF_3COCI)_N$ clusters in the beam have maximum sizes, only slightly larger than $N \geq 10$.

2.2 Theory

The structure of all molecules and ions was optimized at the B3LYP-D3/aug-cc-pVTZ level of theory employing the D3 dispersion correction as proposed by Grimme.31 To obtain more reliable electronic energies, single-point recalculation at the coupled cluster singles and doubles level, also with noniteratively included triplets, CCSD/aug-cc-pVDZ and CCSD(T)/ aug-cc-pVDZ, respectively, was performed, with the zero-point energy included as calculated at the B3LYP-D3/aug-cc-pVTZ level (when not stated otherwise, we refer to CCSD(T) values in the text). We include both CCSD and CCSD(T) reaction energies as CCSD(T) calculations were too demanding in the case of trimer clusters. Charges on atoms were calculated employing the CHELPG scheme at the B3LYP-D3/aug-cc-pVTZ level.³² Wave function stabilization was performed prior to every calculation. A search for the structure of neutral and charged CF₃COCl dimers, trimers and tetramers was performed using our in-house genetic algorithm code. 33 Namely, we create a pool of 20 initial structures and run 20 (dimers) and 30 (trimers, tetramers) cycles with 10 recombinations in each. The semi-empirical PM6 approach was used within genetic algorithm runs, 30 structures were subsequently re-optimized at BLYP-D3/6-31g* and B3LYP-D3/aug-cc-pVDZ levels to select four most stable or representative structures. For a cationic tetramer, two runs with 16 cycles were performed due to convergence issues. Tetramer structures were calculated only up to the B3LYP-D3/aug-cc-pVDZ level and are summarized in the ESI.† Although the employed technique is incapable of providing an exhaustive exploration of the rich potential energy surface of the investigated species, it offers an overview of its most important features as discussed below. All quantum chemical calculations were performed using the Gaussian 16 program.34

3 Results and discussion

Neutral cluster structure

We start our discussion with the calculated structures of the neutral CF₃COCl molecule and its molecular clusters, see Fig. 1a. The CF_3COCl molecule is calculated to have C_8 symmetry, the highest negative charge is predicted on the oxygen atom, -0.33e, fluorine atoms have a charge of about -0.2e. The calculated structure shows eclipsed configuration of CF3 with respect to the CO group, in agreement with the structure determined by electron diffraction.¹⁴

When complexed in a dimer, the most stable structure found is OCF₂···CF₃ CFClCOCl, i.e., with considerable rearrangement of the two CF₃COCl units. Such reorganization is a typical feature of the highly reactive CF₃COCl molecules as will be also discussed below. However, we do not expect this structure to be present in the experiment due to kinetic reasons. Considering a simple (CF₃COCl)₂ dimer structure, the most stable structure maximizes the interaction between CF₃ and COCl groups, probably due to slight polarization of the groups, with the COCl group becoming slightly negatively charged. For neutral trimers and tetramers, clusters with intact CF₃COCl units are predicted to be the most stable ones (see ESI,† for tetramers). However, the presence of, e.g., the (CF₃COCl)₂·CF₃Cl·CO structure among the most stable trimer isomers further emphasises the low thermodynamic stability of the CF₃COCl molecule with highly electronegative atoms.

3.2 Positive ions

Fig. 2 shows the positive ion mass spectrum of the CF₃COCl molecule in co-expansion with helium, measured at 70 eV

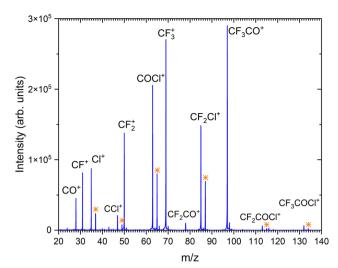


Fig. 2 The mass spectrum of CF₃COCl molecule at 70 eV electron energy. The assigned ions are indicated and the isotope peaks corresponding to the isotpe ³⁷Cl are labeled by the orange stars.

electron energy. To our best knowledge, no mass spectrum of the gas phase molecule could be found in the literature. The molecule fragments significantly upon the ionization exhibiting only very little abundance of the parent ion CF_3COCl^+ at m/z = 132. The major fragment ions (in order of their abundances) are: $CF_3CO^+(m/z = 97)$, $CF_3^+(69)$, $COCl^+(63)$, CF_2Cl^+ (85) and CF_2^+ (50).

Our calculations show that the vertical and adiabatic ionization energies of the CF₃COCl molecule are 12.6 eV and 11.6 eV, respectively. In the cation ground state, the electron is removed from an orbital localized mostly on the C-C bond. Subsequently,

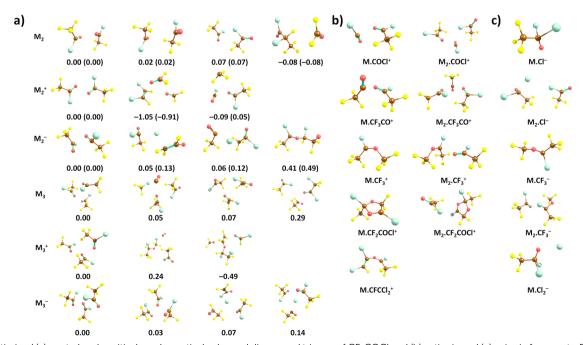


Fig. 1 Optimized (a) neutral and positively and negatively charged dimers and trimers of CF₃COCl and (b) cationic and (c) anionic fragments. For clusters, zero of energy is set to the structure considered to be present in the experiment, energy is given in eV. Calculated at the CCSD/aug-cc-pVDZ//B3LYP-D3/aug-cc-pVTZ level, the CCSD(T)/aug-cc-pVDZ//B3LYP-D3/aug-cc-pVTZ values are given in parentheses.

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Table 1 Energies of various reactions including CF₃COCl. Calculated at the CCSD/aug-cc-pVDZ//B3LYP-D3/aug-cc-pVTZ level, CCSD(T)/aug-ccpVDZ//B3LYP-D3/aug-cc-pVTZ values are given in parenthesis. For reactions with oligomers, clusters with zero relative energy in Fig. 1 are used

Reaction	Energy (eV)
$2CF_3COCl \rightarrow (CF_3COCl)_2$	-0.19 (-0.23)
$3CF_3COCl \rightarrow (CF_3COCl)_3$	-0.48
$CF_3COCl \rightarrow CF_3COCl^+ + e^-$	11.68 (11.62)
$(CF_3COCl)_2 \rightarrow (CF_3COCl)_2^+ + e^-$	11.29 (11.18)
$(CF_3COCl)_3 \rightarrow (CF_3COCl)_3^+ + e^-$	11.07
$CF_3COCl + e^- \rightarrow CF_3COCl^-$	-1.00 (-1.04)
$(CF_3COCl)_2 + e^- \rightarrow (CF_3COCl)_2^-$	-1.32(-1.43)
$(CF_3COCl)_3 + e^- \rightarrow (CF_3COCl)_3^-$	-1.58
$CF_3COCl^+ \rightarrow CF_3CO^+ + Cl$	0.04 (0.20)
$CF_3COCl^+ \rightarrow CF_3^+ + COCl$	0.88 (0.96)
$CF_3COCl^+ \rightarrow COCl^+ + CF_3$	0.08 (0.18)
$CF_3COCl^+ \rightarrow CF_2Cl^+ + OCF$	0.74(0.81)
$CF_3COCl^+ \rightarrow CF_2^+ + COClF$	2.01(2.10)
$(CF_3COCl)_2^+ \rightarrow CF_3COCl \cdot COCl^+ + CF_3$	0.13(0.29)
$\begin{array}{l} (\text{CF}_3\text{COCl})_3^+ \rightarrow (\text{CF}_3\text{COCl})_2\text{-}\text{COCl}^+ + \text{CF}_3 \\ (\text{CF}_3\text{COCl})_2^+ \rightarrow \text{CF}_3\text{COCl} \cdot \text{CF}_3\text{CO}^+ + \text{Cl} \\ (\text{CF}_3\text{COCl})_3^+ \rightarrow (\text{CF}_3\text{COCl})_2\text{-}\text{CF}_3\text{CO}^+ + \text{Cl} \end{array}$	0.38
$(CF_3COCl)_2^+ \rightarrow CF_3COCl \cdot CF_3CO^+ + Cl$	-0.05 (0.09)
$(CF_3COCl)_3^+ \rightarrow (CF_3COCl)_2 \cdot CF_3CO^+ + Cl$	-0.09
$(CF_3COCl)_2^{T} \to CF_3COCl \cdot CF_3^{T} + COCl$	0.12(0.24)
$(CF_3COCl)_3^+ \rightarrow (CF_3COCl)_2 \cdot CF_3^+ + COCl$	0.04
$ \begin{array}{l} \left(\text{CF}_3\text{COCl}\right)_2^+ \to \text{CF}_3\text{COCl}\cdot\text{CF}_2\text{COCl}^+ + F \\ \left(\text{CF}_3\text{COCl}\right)_3^+ \to \left(\text{CF}_3\text{COCl}\right)_2\cdot\text{CF}_2\text{COCl}^+ + F \end{array} $	1.20(1.36)
$(CF_3COCl)_3^+ \rightarrow (CF_3COCl)_2 \cdot CF_2COCl^+ + F$	1.42
$(CF_3COCl)_3^{+} \rightarrow CF_3COCl \cdot CFCCl_2^{+} + CF_4 + CO_2 + F$	1.82
$CF_3COCl + e^- \rightarrow Cl^- + CF_3CO$	-0.23 (-0.15)
$CF_3COCl + e^- \rightarrow CF_3^- + CO + Cl$	1.53(1.62)
$CF_3COCl + e^- \rightarrow CF_3Cl^- + CO$	-0.55 (-0.47)
$CF_3COCl + e^- \rightarrow COCl^- + CF_3$	$-0.20\ (-0.08)$
$CF_3COCl + e^- \rightarrow F^- + CF_2COCl$	1.60(1.61)
$(CF_3COCl)_2 + e^- \rightarrow CF_3COCl_2^- + CF_3CO$	-0.79 (-0.85)
$(CF_3COCl)_3 + e^- \rightarrow (CF_3COCl)_2 \cdot Cl^- + CF_3CO$	-0.78
$(CF_3COCl)_2 + e^- \rightarrow CF_3COCl \cdot CF_3^- + COCl$	0.27(0.31)
$(CF_3COCl)_3 + e^- \rightarrow (CF_3COCl)_2 \cdot CF_3^- + COCl$	0.22
$(CF_3COCl)_3 + e^- \rightarrow CF_3COCl \cdot Cl_2^- + 2CF_3CO$	2.27

the bond weakens considerably, prolonging from 1.56 Å to 2.09 Å (in agreement with the high experimental yield of CF₃⁺ and COCl⁺ ions). In the optimized ion, the charge is distributed almost equally, with 0.47e and 0.53e predicted for CF₃ and for COCl moieties, respectively. The energies of the most important dissociation channels are collected in Table 1, the low dissociation energies match nicely the experimental observations. The appearance of CF₃CO⁺, CF₃⁺ and COCl⁺ can be explained by direct dissociation of a covalent bond, CF₂Cl⁺ and CF₂⁺ require a transfer of a chlorine or fluorine atom prior to the dissociation.

The electron-impact ionization can be described by the Binary-Encounter-Bethe (BEB) model in the formalism developed by Kim and co-workers.35 In this model, the total ionization cross section is given as the sum of cross sections for

ionization from occupied molecular orbitals, $\sigma_{\mathrm{tot}} = \sum_{i} \sigma_{i}$. The

individual σ_i 's can be evaluated from orbital energies and kinetic energies by a semi-empirical formula. 35,36 These orbital σ_i 's contributions, evaluated at 70 eV electron energy, are shown as Fig. S10 in the ESI.† A similar effect as was observed for a different fluorinated molecule in ref. 36 can be seen: the ionization from HOMO does not dominate the cross section, rather, the ionization from many molecular orbitals contributes significantly. This means that the cation is produced in many electronically excited states. Most probably, these undergo an

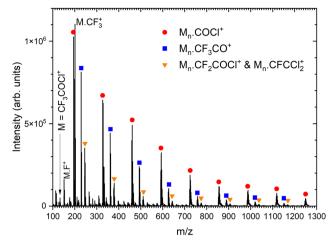


Fig. 3 The mass spectrum of (CF₃COCl)_N clusters ionized at 70 eV electron energy. The clusters are produced in co-expansion of CF₃COCl molecules with Ar buffer gas at the expansion pressure of 3.8 bar.

ultrafast internal conversion into the ground cation state which is left vibrationally hot and fragments statistically.³⁶ Together with the low reaction energy for Cl evaporation (0.2 eV), this explains the low yield of the parent ion in the experimental spectrum.

Next, we generate $(CF_3COCI)_N$ clusters by co-expanding the CF₃COCl molecules with Ar buffer gas. Several expansion pressures were exploited and the spectra were qualitatively the same (see Fig. S1 in ESI†), therefore we show only the spectrum recorded at the highest used stagnation pressure of 3.8 bar corresponding to the largest produced clusters, see Fig. 3. In the mass range up to the monomer mass m/z = 132, the cluster spectra are very similar to the spectrum of the molecule (see Fig. S2 in ESI†). Three significant cluster ion series are labeled in Fig. 3: red circles denote the M_n ·COCl⁺ series (M = CF₃COCl), blue squares label $M_n \cdot CF_3CO^+$, and orange triangles can correspond to $M_n \cdot CF_2COCl^+$ and/or $M_n \cdot$ CFCCl₂⁺. The ambiguity in the assignment of the last series arises from the mass coincidence (19F plus 16O coincides with ³⁵Cl). This could, in principle, be resolved by isotope contributions, namely 35Cl and 37Cl.

The cluster ions contain the ³⁵Cl and ³⁷Cl isotopes in a ratio corresponding to their natural abundances, which is ³⁵Cl: ³⁷Cl = 0.76: 0.24. Depending on the number of Cl atoms in the cluster, we can calculate the relative intensities of the isotope peaks that follow the binomial distribution (see ESI,† for details). In Fig. 4, we show the details of the ambiguous series $\mathbf{M}_n \cdot \mathbf{R}^+$, $\mathbf{R} =$ CF_2COCl or $CFCCl_2$ for n = 1-3, and we compare the peak intensities to the calculated abundances of the isotopes (symbols) assuming both assignments. The calculated isotope abundances are normalized to the maximum peak intensity in the corresponding group. The first mass peak group for n = 1 agrees better with the $R = CF_2COCl$ assignment, while for n = 2 and 3, the isotope abundances for $M_n \cdot R^+$ with $R = CFCCl_2$ fit the spectrum better. In Fig. S3 and S4 in the ESI,† we demonstrate that the assignment using isotope ratios works well for the unambiguously assigned series $M_n \cdot COCl^+$ and $M_n \cdot CF_3CO^+$.

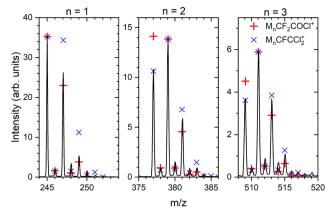


Fig. 4 Details of the mass spectrum of (CF3COCI)_N clusters for the ambiguous series $M_n \cdot R^+$ where $R = CF_2COCl$ or $CFCCl_2$ (orange triangles in Fig. 3). Symbols indicate the calculated abundances of the isotopes normalized to the maximum peak within each group

Therefore, we assume that the first group of mass peaks labeled by orange triangles in Fig. 3 corresponds mainly to M·CF₂COCl⁺ and the other mass peaks can be assigned to M_n ·CFCCl₂⁺, $n \ge$ 2, series, although a contribution of both series in all mass peaks cannot be excluded.

There is also a series corresponding to $M_n \cdot CF_3^+$. The first member of the series, $M \cdot CF_3^+$ at m/z = 201, is the labeled maximum peak in Fig. 3. The following $M_n \cdot CF_3^+$, n > 1, ion peaks have intensities comparable with the M_n ·CFCCl₂⁺ series (orange triangles). The analysis is complicated by the isotopologue contribution of the major $M_n \cdot COCl^+$ ions as shown in ESI† (Fig. S5), but the contribution of the $M_n \cdot CF_3^+$ series is relatively small with the major contribution originating just from the first M·CF₃⁺ peak. The relative intensities of the cluster ion series $M_n \cdot COCl^+$, $M_n \cdot CF_3CO^+$, $M_n \cdot CF_3^+$, and M_n ·CFCCl₂⁺ (M_n ·CF₂COCl⁺) could be estimated to 46%, 25%, 20%, and 9%, respectively. These estimates should be considered with caution in view of relatively high spectra congestion, mass coincidences and ambiguities in the peak assignments. Nevertheless, the three major molecular fragments CF₃CO⁺, CF₃⁺, and COCl⁺ are reflected in the cluster ion series, though in different ratios.

There are other minor series in the spectra, which could not be assigned, however, their contribution is small (see Fig. S6 in ESI,† for the fully resolved spectrum). Worth noting is the absence of non-fragmented cluster ions M_n^+ , i.e., $(CF_3COCl)_n^+$, in the spectrum, which points to the fact that the molecule is prone to a strong fragmentation after the electron ionization apparent from the mass spectrum of the isolated molecule, and even the cluster environment does not hinder this fragmentation.

Quantum chemical calculations show that the most stable structures of positively charged oligomers are often considerably reconstructed, see Fig. 1. Apart from simple dissociation of a CF₃COCl unit into CF₃ and COCl that are then stabilized in the cluster, we see also more fundamental reconstruction, e.g., formation of FClC-CFCl or formation of a CO molecule. Still, we consider intact CF₃COCl oligomers to be the most probable moieties present in the experiment due to kinetic reasons. The adiabatic ionization energy of the neutral dimer and trimer is predicted to be 11.3 eV and 11.1 eV, respectively (CCSD//B3LYP-D3 level), Table 1.

With respect to the observed fragments, the presence of two main fragments, namely $M_n \cdot COCl^+$ and $M_n \cdot CF_3CO^+$, can be explained by direct dissociation of CF₃ and Cl, respectively, with reaction energies close to 0 eV. The electronic structure of the products can be described as COCl⁺ and CF₂CO⁺ stabilized by the cluster environment; the fact that no cluster reconstruction is required can contribute to the high experimental yield of these ions. As to $M_n \cdot CF_2COCl^+$, formation of a five-membered ring is predicted upon F atom dissociation as shown in (Fig. 1b); the respective formation energies lie below 2 eV (Table 1) for both dimers and trimers. Our calculations show that $M_n \cdot CF_2COCl^+$ and $M_n \cdot CFCCl_2^+$ fragments could be connected as they share the dissociation of F in the first place. For the second ion fragment, CF₄ and CO₂ molecules have to leave the cluster in the subsequent reaction.

The formation of the $M_n \cdot CF_3^+$ ions is predicted to proceed through a direct dissociation process and reconstruction into the CF₃CClOCF₃⁺ ion as shown in (Fig. 1b), possibly leading to the lower $M_n \cdot CF_3^+$ abundance despite its low reaction energy and the fact that the CF₃⁺ ion is the most abundant fragment in the monomer ionization. Finally, the very low reaction energies that even reach negative values for Cl dissociation explain the absence of M_n^+ clusters in the experimental mass spectra.

3.3 Negative ions

3.3.1 DEA of the CF₃COCl molecule. We start again with the results obtained for the beam of isolated molecules produced in co-expansion with helium. Since the electron attachment is very sensitive to the electron energy, the energy was scanned as the mass spectra were measured, as outlined in the Experimental section. Therefore, the results are presented in the form of a 2D intensity map shown in Fig. 5a. Some of the vertical traces corresponding to particular m/z are labeled by the corresponding negative fragment ion assignments. Although the 2D maps provide a convenient overview of the mass spectra energy dependence, a better quantitative analysis is obvious from the mass spectra shown in (Fig. 5b and c) where the long acquisitions at particular electron energies of 1.8 eV and 5.6 eV are shown.

There are two pronounced fragments produced in the DEA to the CF₃COCl molecule; namely Cl⁻ generated preferentially at low electron energies below 3 eV and CF₃⁻ that appear above 3 eV with a maximum between 5 and 6 eV, Fig. 6. Some much weaker fragments appear in the spectra shown in Fig. 5 upon closer inspection, especially at low electron energy, e.g., CF₃Cl⁻, COCl and F with their energy dependence given also in Fig. 6. Further ions labeled in Fig. 5b are either negligible (CFCOCl⁻, C₂F₃⁻) or are obviously generated from dimers present in the molecular beam since they contain two oxygen atoms (grey labels in Fig. 5b). The energy-dependent ion yields in Fig. 6 show that dominant Cl⁻ (black) are produced at very low energies possibly with maximum at zero. There is a second much lower maximum at around 5.4 eV and possibly a weak

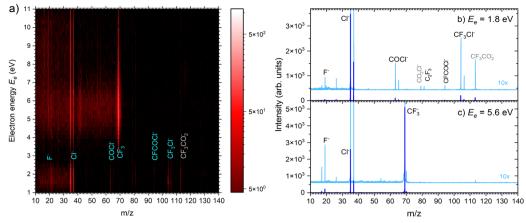


Fig. 5 (a) 2D mass spectra of negative ions produced by electron attachment to CF₃COCI molecules at different electron energies. (b) and (c) Mass spectra at two different electron energies

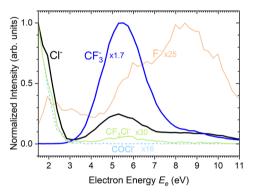


Fig. 6 Electron energy dependence of the negative ion yield in the DEA to CF3COCI molecules

shoulder at around 9-10 eV. CF₃ ions (blue) are produced only at the higher energies of around 5.4 eV with a tail towards 9 eV suggesting a similar process to the Cl- second peak. The other fragments are much less populated. Note that the energydependent ion yields shown in Fig. 6 have been normalized to the maximum (the multiplication factors are outlined by the corresponding lines). The CF₃Cl⁻ (light green) and COCl⁻ (dashed light blue) ions also exhibit the maximum at low energies, and the former ion is also observed around 5.4 eV. The F⁻ ion (light orange) seems to have a different dependence with a wide maximum at around 8-9 eV and a shoulder at 6 eV (a low-energy peak at around 2 eV is uncertain due to low intensity).

The calculated adiabatic electron affinity of CF₃COCl is 1.04 eV. Upon electron attachment, an electron is added into the π^* orbital of the C=O bond. The additional charge leads to the C-Cl bond prolongation to 2.18 Å. In the resulting moiety, [CF₃CO···Cl]⁻, the charge is almost equally distributed on CF₃CO (-0.45e) and Cl (-0.55e). Such a molecular anion is not experimentally observed in the binary electron collisions with isolated CF₃COCl since excess energy leads either to electron detachment or to the fragmentation of the complex (dissociative electron attachment, DEA).

According to the calculated reaction energies (Table 1), there are three exothermic DEA channels. Interestingly, the most energetically favorable one, formation of CF₃Cl⁻, is experimentally observed to be rather weak. We ascribe its low cross section to the necessity for a complex atomic rearrangement on the resonant potential energy surface. On the other hand, Cl⁻ and COCl⁻ can be formed by breaking a single bond and have similar calculated exothermicities (0.15 eV vs. 0.08 eV). Also, CF₃Cl⁻ can dissociate into CF₃ + Cl⁻, requiring only 0.53 eV.

A combination of effects can contribute to the fact that the Cl⁻ yield is much higher (approximately by a factor of 16). The first one is the involved resonant state, especially its dissociative direction. We can roughly estimate the energy of the π^* (C=O) resonance (i.e., the vertical attachment energy at the geometry of the neutral) using the empirical relation between the energy of the unoccupied molecular orbital and the energy of the corresponding resonance. The empirical relation of Chen and Gallup³⁷ [$E_{res} = (E_{MO} - 2.33 \text{ eV})/1.31$] predicts that adding an electron to the LUMO (E_{MO} = 2.31 eV) yields an E_{res} value very close to zero. It thus corresponds to a state which is very weakly bound or to a resonance with energy close to 0 eV. This is clearly the state involved in the low-energy Cl and COCl production. Upon visual inspection of the LUMO (Fig. S11 in ESI†), it is clear that the antibonding plane crosses the C-Cl bond, while the C-C bond is in a bonding plane. This hints that upon the formation of the π^* (C=O) resonance, the C-Cl bond will be promptly cleaved. The cross section for the cleavage of this bond can be further enhanced by the bending motion of the Cl atom from the plane defined by the CCO atoms. Such bending will mix the π^* (C=O) resonance with a higher-lying σ^* (C=O) state, such mixing strongly increases the DEA cross section. 38-40 Also, the reduced mass of the CF₃CO···Cl⁻ system is 25.7 amu, while that of the CF₃···COCl⁻ system is 32.9 amu (for the ³⁵Cl isotope). The dissociation in the latter case will be thus slower, providing more time for the electrons to autodetach.41 This leads to lowering of the DEA cross section in the COCl channel. Finally, the COCl ion is a weakly bound complex, CO···Cl-, and might easily dissociate to CO + Clwith a calculated reaction energy of 0.14 eV.

Formation of CF₃⁻ has a calculated threshold of 1.62 eV. The experimental signal peaks appear at around 5.4 eV. The UV-VIS spectrum⁵ supported by theoretical calculations¹⁰ exhibits a maximum at around 4.9 eV. Thus, the incoming electron can excite the molecule and the observed peak can correspond to a core-excited shape resonance (their formation can be possibly enhanced by a dipole moment of the excited states^{42,43}). The DEA peaks visible at about 8-10 eV could correspond to the core-excited resonances that have higher excited states¹⁰ as the parent states.

3.3.2 Electron attachment to clusters. Fig. 7 shows the 2D mass spectrum and electron energy dependence after electron attachment to the (CF₃COCl)_N clusters generated under the same expansion conditions as the positive ion spectrum in Fig. 3. The molecular part of the spectrum is essentially the same as that for the DEA of the isolated molecule shown in Fig. 5. The major cluster fragments, labeled in Fig. 7, are $M_n \cdot Cl^-$ that prevail at higher electron energies but one can see in Fig. 8 that the spectrum is dominated by the cluster ions M_n^- (i.e., $(CF_3COCl)_n^-$) at lower energies below 3 eV.

Fig. 8 shows the mass spectra at specific electron energies of (a) 1.7 eV, (b) 5.0 eV and (c) 8.0 eV. The major cluster ion fragment series are $M_n \cdot Cl^-$ and M_n^- mentioned above. Also an interesting series with molecular chlorine $M_n \cdot Cl_2^-$ exhibits appreciable abundances at higher energies. However, $M_n \cdot Cl_2$ overlaps with $M_n \cdot CF_3$ due to isotope contributions (Cl₂ corresponds to m/z = 70 for the ³⁵Cl isotope, and CF_3 to m/z = 69). The two series differ by the number of Cl atoms, therefore a detailed analysis is possible using the natural isotope abundances as outlined for the positive ions. It shows that both $M_n \cdot Cl_2$ and $M_n \cdot CF_3$ series are present in the spectra (see Fig. S7 and S8 in the ESI†).

We identify further less abundant ion series in the spectra upon close analysis as well. For example, a group of mass peaks corresponding to $M_n \cdot CF_3CO^-$ at m/z = 229 is labelled in (Fig. 8b) at 5.0 eV. This peak and further smaller fragment series are shown in detailed spectra as shown in Fig. S9 in the ESI.† However, the abundance of these ions is relatively small.

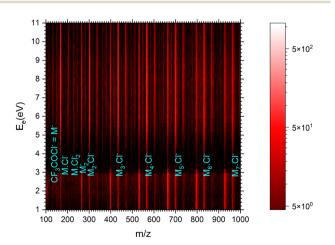


Fig. 7 The 2D mass spectrum and electron energy dependence after the electron attachment to the (CF₃COCI)_N clusters.

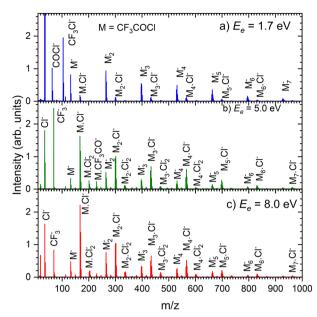


Fig. 8 The mass spectra at specific electron energies of (a) 1.7 eV, (b) 5.0 eV and (c) 8.0 eV.

Finally, we discuss the electron energy dependent ion yield for several cluster ion fragments shown in Fig. 9. The top panels show the negative ion yield of (a) Cl⁻ and (b) CF₃⁻ fragments (orange lines) in comparison with the yields of the same ions from the DEA of isolated molecules (black). The high-energy tails of the cluster spectra suggest that in clusters, processes at somewhat higher energies can also contribute to the generation of these ions. Panels (c) and (d) below show the spectra for the $M_n \cdot Cl^-$ and $M_n \cdot CF_3^-$ ions respectively. The shape of the spectra is independent of the cluster ion size n. A comparison of Fig. 9c and d with Fig. 9a and b, respectively, suggests that the main maxima in the top spectra are due to the isolated molecules in the beam and the Cl and CF₃ fragments from the clusters are produced at higher energies. The low-energy maximum for Cl⁻ seems to be strongly suppressed in the clusters. Finally, the bottom row in Fig. 9e and f shows the spectra for M_n^- and $M_n \cdot Cl_2^-$ fragments. The M_n^- ions are produced at low energies. The broad peak at higher energies might be due to the self-scavenging processes in clusters,²⁹ where the incoming electron excites a CF₃COCl molecule, loses its kinetic energy and attaches as the slow electron. The $M_n \cdot Cl_2^-$ ions are formed at higher electron energies due to the extra energy required to rearrange the molecules (see below).

The calculated adiabatic electron affinity of the CF₃COCl dimer and trimer is 1.32 eV and 1.58 eV, respectively (CCSD// B3LYP-D3). As already hinted through the experiments, the cluster environment might prevent the Cl pre-dissociation. Although Cl⁻ could be formed in the cluster, a cluster conformation with intact CF₃COCl units is slightly more stable

In the case of CF₃COCl₂⁻, our calculations suggest that a C-Cl bond is formed, see Fig. 1, the formation of the ion upon electron attachment is exothermic in both dimer and trimer

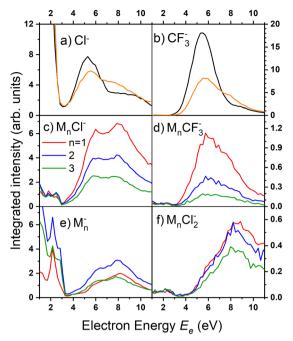


Fig. 9 The electron energy dependence of the negative ion yield for several fragments of the electron attachment to the (CF3COCI)_N clusters. The black lines in the top panels correspond to the spectra of isolated molecules for comparison.

(Table 1). For the $M_2 \cdot Cl^-$ ion, a Cl^- anion is predicted to lie next to the CF₃COCl molecules. The formation of $M_n \cdot CF_3^-$ is again predicted to be connected to cluster reconstruction, forming a $CF_3C(Cl)OCF_3^-$ unit. The $M_n \cdot Cl_2^-$ anion can be produced by the dissociation of two CF₃CO units, forming a CF₃COCl₂····Cl⁻ cluster at an energy of about 2.3 eV.

4 Conclusions

We performed a molecular beams study of trifluoroacetyl chloride and its clusters, analyzing their behavior by interaction with electrons.

For an isolated molecule, ionization leads to the weakening the C-C bond, yielding CF₃⁺ or COCl⁺ fragment ions. Breaking of the C-Cl bond leads to the major fragment ion CF₃CO⁺, however, the sum of CF3⁺ and COCl⁺ ion yields exceeds the CF₃CO⁺ yield significantly, suggesting that the C-C bond dissociation is the primary channel. On the other hand, the electron attachment primarily breaks the C-Cl bond leading to Cl⁻ dominating at low electron energies below 3 eV, and CF₃, which is the major channel at higher energies above 4 eV.

For clusters, a typical feature is their thermodynamic instability, with cluster reconstruction often predicted to be favorable by our genetic algorithm search. In cationic clusters, two main series are observed upon ionization, M_n . $COCl^+$ and $M_n \cdot CF_3CO^+$, both formed through direct dissociation. In the case of minor fragments, considerable cluster reconstruction is predicted, most probably contributing to their low experimental yield.

For anions, stabilized molecular M_n^- clusters dominate the mass spectra at low energies below 3 eV. Also dissociation of the C-Cl bond plays a role at low energies forming M_n ·Cl⁻ fragment ions, which become the dominant species at higher electron energies above 4 eV. The dissociation of the C-C bond can occur as well at these higher energies, forming $M_n \cdot CF_3^-$. At higher energies, $M_n \cdot Cl_2^-$ cluster ions also occur.

In terms of atmospheric relevance we stress two findings: (i) the gas-phase molecule is very fragile with respect to the electron impact, both the positive ionization and electron attachment mass spectra are fragmentative, and (ii) the major cationic fragment in the gas phase is CF₃CO⁺, with its necessary neutral counterpart being the Cl radical. Upon aggregation, the anion dissociation pathways are considerably suppressed and the intact molecular cluster anions are dominant. The cationic pathways are not influenced by clustering as strongly as the anionic ones, even though the chlorine-producing channel $(M_n \cdot CF_3CO^+)$ is not the most dominant fragment progression in clusters.

Author contributions

B. Kocábková: data curation, investigation, formal analysis, and visualization; J. Ďurana: investigation, formal analysis, and visualization; J. Rakovský: investigation; A. Pysanenko: investigation; V. Poterya: investigation; J. Fedor: writing - review and editing, and validation; M. Ončák: conceptualization, writing original draft, formal analysis, visualization, validation; and M. Fárník: conceptualization, writing original draft, formal analysis, visualization, validation, and funding acquisition.

Conflicts of interest

There are no conflicts to declare.

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