

Showcasing research from Professor Wang's laboratory,
College of Chemistry, Nankai University, Tianjin 300071, China.

EDTA etching: a simple way for regulating the traps, size and
aqueous-dispersibility of Cr^{3+} -doped zinc gallate

The traps, size and aqueous-dispersibility are the most important parameters that affect the features and applications of persistent luminescent nanoparticles (PLNPs). However, the simultaneous controlling of those parameters is rather difficult and has not been reported yet. We present the first exploration on adjusting the traps, size and aqueous-dispersibility of PLNPs via a simple ethylenediaminetetraacetate (EDTA) etching. The EDTA etching of the sintered $\text{Cr}^{3+}_{0.004}\text{ZnGa}_2\text{O}_4$ results in effective reduction of the size, great improvement of the aqueous-dispersibility, but meanwhile altering the density of mediate traps and generating the new deep traps.

As featured in:



See He-Fang Wang *et al.*,
Chem. Sci., 2018, 9, 8923.



rsc.li/chemical-science

Registered charity number: 207890



Cite this: *Chem. Sci.*, 2018, 9, 8923

All publication charges for this article have been paid for by the Royal Society of Chemistry

EDTA etching: a simple way for regulating the traps, size and aqueous-dispersibility of Cr³⁺-doped zinc gallate†‡

He-Fang Wang,¹ Xi Chen,² Fan Feng,² Xia Ji² and Ye Zhang²

Traps, size and aqueous-dispersibility are the most important parameters that affect the features and applications of persistent luminescent nanoparticles (PLNPs). However, simultaneous controlling of these parameters is rather difficult and has not been reported yet. We present the first exploration on adjusting the traps, size and aqueous-dispersibility of PLNPs *via* simple ethylenediaminetetraacetate (EDTA) etching. Cr_{0.004}³⁺:ZnGa₂O₄ (ZGO) was used as the PLNP model. EDTA etching of the sintered ZGO results in effective reduction of the size and great improvement in the aqueous-dispersibility. In addition, EDTA etching alters the density of mediate traps and generates new deep traps, thus achieving the massive production of (ultra)small ZGO–EDTA with fine aqueous-dispersibility, suitable mediate/deep traps and superlong bright afterglows (51 days). As EDTA can interact with most metals, this simple EDTA etching strategy is prospectively amenable to other PLNPs, and the resulting PLNPs–EDTA have wide applications in both biological field and information storage.

Received 19th September 2018

Accepted 11th November 2018

DOI: 10.1039/c8sc04173c

rsc.li/chemical-science

Introduction

Persistent luminescent nanoparticles (PLNPs)^{1–5} have aroused considerable attention in pre-excitation optosensing,^{6–10} bio-imaging,^{11–28} drug delivery,^{29–31} imaging-guided therapy,^{6,32–36} and information storage.^{22,37–39} Traps, size and aqueous-dispersibility are the most important parameters that affect the features and applications of PLNPs. For instance, the density of mediate traps (0.5–0.8 eV) is considered to be inter-related with the room temperature (RT) afterglow intensity and durations. Deep traps of 0.8–1.6 eV are required for retaining the recorded information for long time,^{37,38} while the size and aqueous-dispersibility of PLNPs are the key factors in biological applications.⁴⁰ Some studies have reported the low-temperature-synthesis methods, including the hydrothermal (200–220 °C^{7,41,42}), biphasic (160 °C)⁴³ and microwave-assisted non-aqueous sol–gel (270 °C),⁴⁴ to obtain the small-sized PLNPs, but without controlling the traps. Others reported doping with different codopants for engineering the trap depth,^{37,38} but with no ideas on adjusting the size and dispersibility. To the best of our knowledge, simultaneous

controlling of the traps, size and dispersibility of PLNPs is a formidable task, and has not been reported yet.

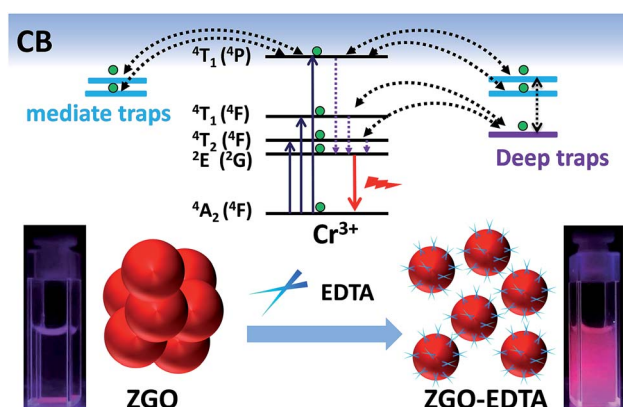
Herein, we present the first exploration on adjusting the traps, size and aqueous-dispersibility of PLNPs *via* a simple post-treatment of ethylenediaminetetraacetate (EDTA) etching. The widely reported Cr_{0.004}³⁺:ZnGa₂O₄ (ZGO) was used as the PLNP model. The ZGO synthesized by the biphasic method⁴³ was sintered at 1000 °C and then treated with EDTA solutions at different temperatures, pH values and amounts (Scheme 1, the details are elaborated in ESI†). EDTA etching prioritizes in dissolving the interparticle necks between the sintered ZGO nanoparticles, and thus breaks the large particles down into the small particles. Moreover, the EDTA etching drives some metals

¹Research Center for Analytical Sciences, College of Chemistry, Nankai University, Tianjin Key Laboratory of Biosensing and Molecular Recognition, State Key Laboratory of Medicinal Chemical Biology, Tianjin 300071, China. E-mail: wanghefang@nankai.edu.cn

²Collaborative Innovation Center of Chemical Science and Engineering (Tianjin), Tianjin 300071, China

† This article is dedicated to the 100th anniversary of Nankai University.

‡ Electronic supplementary information (ESI) available. See DOI: 10.1039/c8sc04173c



Scheme 1 The strategy of EDTA etching for regulating the traps, size and aqueous-dispersibility of ZGO.



in ZGO into the supernatants, resulting in further reduction of the size, and the departure of Ga and Zn takes away the linked oxygens and thus generates some oxygen vacancies (corresponding to deep traps) neighboring the remaining Ga and Zn on the resulting nanoparticles of ZGO-EDTA.

To demonstrate the hypothesis, we collected and analyzed the liquid supernatants after the EDTA-treatment by inductively coupled plasma-atomic emission spectroscopy (ICP-AES) to examine the etching amount of the metals. Moreover, the harvested solid particles were dried and evaluated in the aspects of the yield, size, aqueous-dispersibility, photoluminescence (PL), thermo-luminescence (TL) and afterglows. We show that EDTA etching can regulate the traps, size, and aqueous-dispersibility of the resultant ZGO-EDTA nanoparticles. We discuss the reason behind EDTA etching and applications of the ZGO-EDTA nanoparticles in *in vivo* bioimaging and information storage. To the best of our knowledge, EDTA etching is the first strategy that can simply generate the deep traps while reducing the size and improving the aqueous-dispersibility.

Results and discussion

EDTA etching

The treatments by EDTA solutions at different temperatures (RT and 220 °C), pH (4, 7, 10, 13) and amounts (1, 3, 5, 7 equivalent weight of ZGO) were examined. We first measured the contents of metals in the supernatants treated with EDTA under various conditions and with water, HNO₃-pH4, and NaOH-pH13 (the treatments are at 220 °C unless stated otherwise, Table S1† and Table 1, the corresponding percentages against the nominal amounts are also listed). The supernatants treated with water and HNO₃-pH4 have negligible contents of Ga, Zn and Cr, and the treatment with EDTA at RT (5EDTA-pH13-RT) results in a small quantity of Ga, Zn and Cr in the supernatant. On the contrary, the supernatants treated with 5EDTA-pH (4, 7, 10 and 13) and NaOH-pH13 own substantial concentrations and corresponding percentages of Ga and Zn. The highest and lowest erosion percentage of Ga and Zn are observed at 5EDTA-pH4 and 5EDTA-pH10, respectively. Compared with the data of HNO₃-pH4 and NaOH-pH13, the etching percentage of EDTA at

identical pH (4 or 13) is significantly high toward Ga and Zn. For Cr, the various amounts of EDTA solutions at pH 13 show the lowest etching percentage. Currently, the reason behind these erosion data is not well understood. One fact we can think of is that Ga, Zn and Cr are amphoteric; thus, ZGO can be partly etched by EDTA solutions at various pH. Among those treatments (Tables S1† and 1), EDTA solutions at pH 13 show substantial etching ability toward Ga and Zn (the matrix of ZGO), but very low erosion ability toward Cr, having the potential of reducing the size while maintaining the luminescent center. Thus, we systematically examined the etching cases of various amounts of EDTA at pH 13 in subsequent studies. We note that the etching percentage of Zn is lower than that of Ga (Table 1), resulting in the increase in Zn/(Ga + Cr) ratio of the resultant ZGO-*n*EDTA (*n* = 1, 3, 5, 7) nanoparticles, which leaves more chances for Cr³⁺ to neighbor the antisite defect and finally generate the more intense N₂ emission (695 nm) than R emission (688 nm) in PL spectra⁴⁵ (Fig. S1†).

We then tested the harvested amounts of ZGO-*n*EDTA after etching. To imitate the selective sedimentation, all solids after corrosion were re-dispersed in ultrapure-water and settled for 10 min, and the solid in suspensions (marked as ZGO-*n*EDTA-S) and sediments (marked as ZGO-*n*EDTA-B) were collected (ZGO-S and ZGO-B are the products from tradition grinding, without hydrothermal treatment). Table S2† lists the amounts and the corresponding yields. The traditional grinding gives 4.7% of ZGO-S and 90% of ZGO-B. In contrast, the EDTA etching produces more solid in suspensions (ZGO-*n*EDTA-S) and less solid in sediments (ZGO-*n*EDTA-B). The decreased total amounts of ZGO-*n*EDTA (the sum of S and B) are ascribed to the corrosion by EDTA, as listed in Table 1. The yields of ZGO-*n*EDTA-S gradually increased with the increase in amount of EDTA, giving the highest yield of 40% when 5EDTA was used. Further increase in EDTA results in slightly lower yield of ZGO-7EDTA-S (35.1%) than ZGO-5EDTA-S (40%), but higher yield of ZGO-7EDTA-B (34.3%) over ZGO-5EDTA-B (23.3%). It is worth noting that the solid in suspensions should have smaller size and/or more water-dispersible ligands than that in sediments. The following section gives the explanation.

Table 1 The concentration (conc., mean ± SD) and the corresponding erosion percentage (P, mean ± SD) of the metals existing in the supernatants from hydrothermal treatment of ZGO with different amounts of EDTA at pH 13 and water

Entry	Ga		Zn		Cr	
	Conc. ^a (mg L ⁻¹)	P ^b (%)	Conc. ^a (mg L ⁻¹)	P ^b (%)	Conc. ^a (mg L ⁻¹)	P ^b (%)
1EDTA	1813 ± 38	35.0 ± 0.7	542 ± 41	22.3 ± 1.7	0.022 ± 0.004	0.22 ± 0.04
3EDTA	2125 ± 67	41.0 ± 1.3	712 ± 21	29.3 ± 0.9	0.015 ± 0.005	0.15 ± 0.05
5EDTA	2233 ± 143	43.1 ± 2.8	714 ± 42	29.4 ± 1.7	0.030 ± 0.016	0.30 ± 0.16
7EDTA	1713 ± 38	33.0 ± 0.7	441 ± 1	18.1 ± 0.03	0.020 ± 0.001	0.21 ± 0.01
H ₂ O	0.19 ± 0.04	<0.01	<0.005	<0.01	<0.004	<0.05

^a The concentration of each metal is detected by ICP-AES. ^b The erosion percentage is calculated by $\frac{cV}{m \times \frac{nM_m}{M}} \times 100\%$, where *c* is the concentration of individual metal in the supernatant; *V* is the volume of the supernatant; *m* is the mass of raw ZGO; *M_m* is the molar mass of the individual metal; *n* is the number of the individual metal in nominal formula of ZGO; and *M* is the molar mass of ZGO calculated from the nominal formula.



Size, zeta-potential and aqueous-dispersibility of ZGO-*n*EDTA

Fig. 1 describes the size (the hydrodynamic diameter from dynamic light scattering (DLS)) and zeta potential of ZGO, ZGO-S, ZGO-B and ZGO-*n*EDTA (both S and B, $n = 1, 3, 5, 7$). Compared with the raw ZGO, all ZGO-*n*EDTA (both S and B) have smaller size and more negative zeta potential. Moreover, all ZGO-*n*EDTA-S are sub-10 nm in size, and are smaller than ZGO-S, while all ZGO-*n*EDTA-B are 22–30 nm in size, and are smaller than ZGO-B (Table S3 in ESI†). The sizes of ZGO-*n*EDTA-S and ZGO-*n*EDTA-B gradually reduced with the increase in the n value ($n = 1, 3$ and 5 , Table S3†). Owing to the smaller size and more negative zeta potential, all ZGO-*n*EDTA ($n = 1, 3, 5, 7$) exhibit much better aqueous-dispersibility than ZGO, ZGO-S and ZGO-B (Fig. S2†), and ZGO-*n*EDTA-S have better aqueous-dispersibility than the corresponding ZGO-*n*EDTA-B. The HRTEM images of ZGO and typical ZGO-5EDTA (Fig. S3†) further support the changes in the size and dispersibility resulted from EDTA etching. The ZGO from high-temperature calcination exists as ‘hard’ agglomerates,⁴⁶ as revealed by the large particles composed of many adherent small particles with crystal lattices of different orientations. In contrast, the ZGO-5EDTA (both S and B) exists as the individual smaller nanoparticle of unidirectional lattice. The better dispersibility and more negative zeta potential are related to the more amount of EDTA attached on the nanoparticles. As revealed by FTIR spectra (Fig. S4†), ZGO-*n*EDTA-S particles have more EDTA than the corresponding ZGO-*n*EDTA-B particles, and ZGO has negligible EDTA. All these data demonstrate the merit of EDTA etching over traditional grinding in massively producing the (ultra)small and aqueous-dispersed ZGO-*n*EDTA nanoparticles.

TL, traps and afterglows of ZGO-*n*EDTA

For evaluation of the traps, we conducted TL measurements (Fig. 2, S5 and S6a–c†). The amount of ZGO-S and ZGO-1EDTA-S

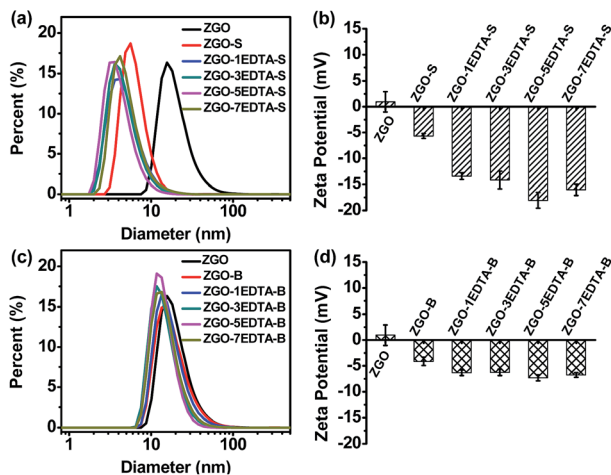


Fig. 1 Hydrodynamic diameter (from DLS) and zeta-potential of (a, b) the raw ZGO, ZGO-S and ZGO-*n*EDTA-S harvested from suspensions, and (c, d) the raw ZGO, ZGO-B and ZGO-*n*EDTA-B harvested from sediments ($n = 1, 3, 5, 7$).

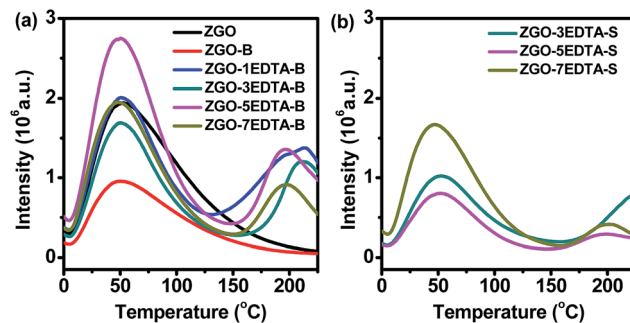


Fig. 2 TL of solid (a) ZGO, ZGO-B, ZGO-*n*EDTA-B and (b) ZGO-*n*EDTA-S. 60 mg of solid ZGO, ZGO-B and ZGO-*n*EDTA were irradiated at 0 °C by a 254 nm UV lamp for 5 min. Two minutes later, the solid sample was heated to 237 °C at a heating rate of 10 °C min⁻¹ and the afterglow at 695 nm was synchronously recorded using a FLS920 fluorescence spectrophotometer.

was too small (Table S2†) for TL measurement. Thus, the TL profiles of ZGO-S and ZGO-1EDTA-S are not included in Fig. 2b. The TL profiles of ZGO and ZGO-B are similar, with the extremely broad tailing peaks maximizing at ca. 50 °C (corresponding to mediate traps of 0.64 eV and 0.73 eV, Fig. S5†), but the TL peak of ZGO-B is lower than that of ZGO. The TL of ZGO-water-B (from hydrothermal treatment by water) also has the similar profile of ZGO, but with the slightly narrower peak and higher intensity at around 50 °C (Fig. S6a† vs. Fig. 2a). In contrast, the ZGO-*n*EDTA (both B and S) displays two distinct peaks (one peak at 50 °C as ZGO, ZGO-B and ZGO-water-B, another at around 175–227 °C) and one shoulder between those peaks (Fig. 2). The EDTA treatment at RT also results in the significant increase in TL peak at 175–227 °C of ZGO-5EDTA-RT-B (Fig. S6a†). These data demonstrate that the EDTA treatments are the cause of the evolution of TL peaks at the high temperature range of 175–227 °C, which are corresponding to the deep traps of 0.83–1.00 eV (Fig. S5, S6b and c†). Moreover, the TL intensities varied with the amount of EDTA, of which the ZGO-5EDTA-B has the highest TL peaks.

Afterglows are usually related to the traps (reflected by TL peaks). The effective traps for RT afterglows are considered to be around 0.5–0.8 eV, but many reports suggested that the electrons in the deep traps could be retrapped by the mediate or shallow traps *via* the quantum tunneling process;¹⁵ thus, the deep traps also would have some contributions to the RT afterglows. The EDTA etching of ZGO not only alters the density of mediate traps, but also generates the new deep traps; thus, it is a new way to regulate the traps and afterglows of ZGO-EDTA. For instance, the ZGO-5EDTA-B has the highest density of mediate traps and deep traps (Fig. 2a and S5†), and thus the brightest afterglows for 51 days (Fig. 3b and S7†). The ZGO-7EDTA-B owns similar density of 0.64 eV traps as ZGO, but displays brighter afterglows than ZGO since the 15th day due to the existence of deep traps. The ZGO-3EDTA-B, which has lower density of mediate traps but higher density of deep traps than ZGO, exhibits higher afterglows over ZGO since the 30th day. The slightly more mediate traps of ZGO-water-B over ZGO-5EDTA-RT-B lead to the slightly higher afterglows of ZGO-water-



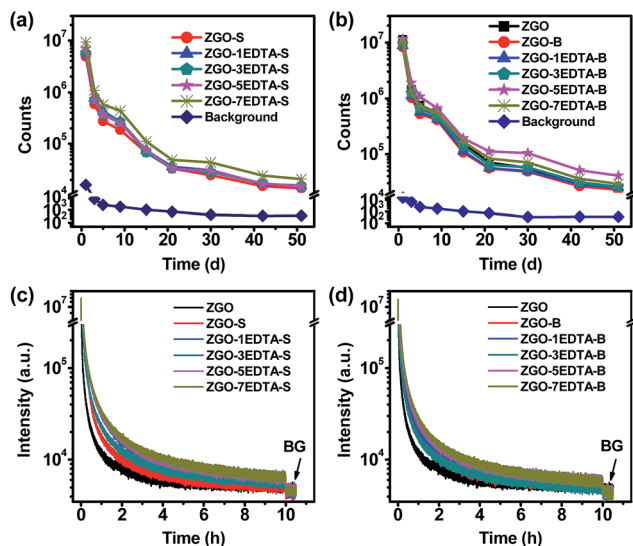


Fig. 3 The afterglow decays of (a) solid ZGO-S and ZGO-*n*EDTA-S, (b) solid ZGO, ZGO-B and ZGO-*n*EDTA-B, (c) aqueous dispersions of ZGO, ZGO-S and ZGO-*n*EDTA-S and (d) aqueous dispersions of ZGO, ZGO-B and ZGO-*n*EDTA-B ($n = 1, 3, 5, 7$). For (a) and (b), 3 mg of solid was irradiated by a 254 nm UV lamp (6 W) for 5 min, and the afterglow images and intensity were recorded using a CCD camera at different durations after stoppage of irradiation. For (c) and (d), 3 mL of 2 g L⁻¹ aqueous dispersions was pre-lighted by a 254 nm UV lamp for 5 min and then was put into the sample holder immediately for recording the afterglow at 695 nm using a PTI fluorescence spectrophotometer.

B over ZGO-5EDTA-RT-B during the first 30 days, while the deep traps of ZGO-5EDTA-RT-B result in the reverse on the 42nd day (Fig. S6d[†]). All these data suggest that the mediate traps are related to the afterglow intensity at early durations. After depletion of the electrons in mediate traps, the electrons in deep traps can be released and retrapped by mediate traps, and generate the afterglows of the late stage (the electrons trapped in mediate traps are emptied faster than those in deep traps, as revealed by the decay time-dependent TL measurements in Fig. S8f[†]). It should be noted that the maintained or enhanced afterglows of the ZGO-*n*EDTA are accompanied with the reduced size, which is quite different from the reported PL enhancement of the hydrothermal treatments (where the particle size keeps invariable).

It is worth noting that the afterglow of all solid ZGO, ZGO-S, ZGO-B and ZGO-*n*EDTA is bright at the 51st day (Fig. S7f[†]), with the highest signal-to-noise ratio (SNR) of 115 (ZGO-5EDTA-B) and the lowest SNR of 38 (ZGO-S). The ZGO-*n*EDTA-S ($n = 1, 3, 5, \text{ and } 7$) have smaller size than ZGO-S (Fig. 1a), but own higher afterglow than ZGO-S (Fig. 3a and S7f[†]). Similarly, ZGO-*n*EDTA-B ($n = 1, 3, 5, \text{ and } 7$) also have smaller size (Fig. 1b) but higher afterglow (Fig. 3b) than ZGO-B. These data demonstrate the superiority of EDTA etching in maintaining or enhancing the afterglow when reducing the size. It should be noted that the ZGO-*n*EDTA-B ($n = 3, 5, 7$) even own higher afterglow intensity than the raw ZGO, which is the first case of reducing the size but increasing the afterglow of nanoparticles.

As many applications require using the aqueous dispersions, we measured the afterglow spectra (Fig. S1f[†]) and decays (Fig. 3c

and d) of the aqueous dispersions of ZGO-*n*EDTA, ZGO-S, ZGO-B and ZGO using a fluorometer. As shown in Fig. S1,† ZGO-*n*EDTA ($n = 1, 3, 5, 7$) displays higher intensity in both PL and afterglow spectra. Furthermore, smaller size, better aqueous-dispersibility, and higher intensity are observed. More significant difference is observed in the afterglow intensity because the difference in aqueous-dispersibility is magnified when the aqueous suspensions are settled for longer time in measurements of afterglow spectra. Similarly, the afterglow decays of the aqueous dispersions measured using a fluorometer are mainly related to aqueous-dispersibility and not only the traps. ZGO-*n*EDTA has much better aqueous-dispersibility. Thus, all the aqueous dispersions of the small-sized ZGO-*n*EDTA display higher afterglow intensity than that of the large-sized ZGO during the decays (Fig. 3c and d), which is quite different from the results of the measurements in solid form (Fig. 3a and b). For the same reason, the aqueous dispersions of the ZGO-*n*EDTA-S have higher afterglow intensity than the corresponding ZGO-*n*EDTA-B. For the aqueous dispersions of S or B series, the more the EDTA used, the higher the afterglow observed in afterglow decays. This elucidates the regulation role of EDTA in adjusting the size, aqueous-dispersibility and afterglows of ZGO-*n*EDTA. It is worth noting that the afterglows of aqueous dispersions of ZGO-*n*EDTA are higher than the background (BG) at 10 h.

The erosion mechanism

The sintering process greatly enhances the afterglows, but with the price of converting the primary powder into the hard-agglomerated particles *via* the formation of interparticle necks with different curvature and crystallinity degree.⁴⁶ The erosion by EDTA prioritizes the dissolution of such interparticle necks between the agglomerated ZGO nanoparticles, and thus breaks the large particles down into the small ones. Moreover, EDTA etching drives some metals (mainly Ga and Zn) into the supernatants (Table 1 and S1f[†]), which results in further reduction of the size (Fig. 1), but without significant change in the spinel phase of ZnGa₂O₄ (JCPDS 38-1240, Fig. S9f[†]). As all Ga and Zn atoms were connected to several oxygen atoms, the departure of Ga and Zn inevitably takes away the linked oxygen atoms and thus generates some oxygen vacancies neighboring the remaining Ga and Zn sites on the resultant ZGO-*n*EDTA.

To verify the oxygen vacancies generated by EDTA erosion, we compared the XPS (Fig. 4) and ESR spectra (Fig. 5) of ZGO and the typical ZGO-5EDTA. The symmetric peaks of Ga 2p_{3/2} (1117.5 eV) and Zn 2p_{3/2} (1021.2 eV) of ZGO suggest that the Ga³⁺ ions are located at the octahedral sites and Zn²⁺ ions are located at the tetrahedral sites⁴⁷ (Fig. 4). The corresponding peaks of ZGO-5EDTA (both S and B) are still symmetric, but positively shift to around 1117.8 (Ga) and 1021.8 eV (Zn). Moreover, the O_{1s} peak in ZGO is at 530.4 eV, but that in ZGO-5EDTA (both S and B) is observed at 531.0 eV. Similarly, the binding energy of Cr 2p_{3/2} is also enlarged in ZGO-5EDTA. These positively shifted binding energies of Ga, Zn, Cr and O in ZGO-5EDTA are ascribed to the neighboring electron-attracting oxygen vacancies.⁴⁸ The extra tailing peaks at around 533 eV of



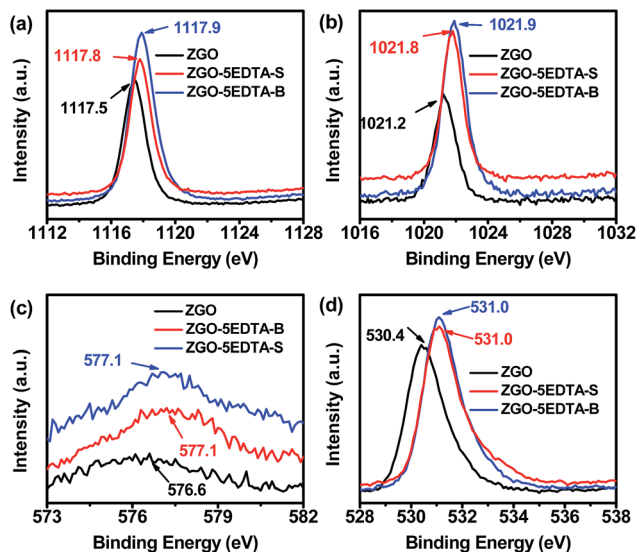


Fig. 4 XPS spectra of (a) Ga $2p_{3/2}$, (b) Zn $2p_{3/2}$, (c) Cr $2p_{3/2}$ and (d) O $1s$ of ZGO, ZGO-5EDTA-S and ZGO-5EDTA-B.

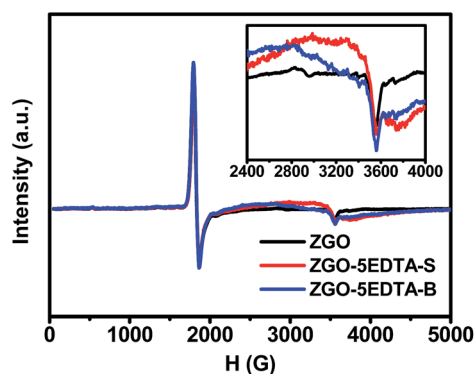


Fig. 5 ESR spectra of ZGO, ZGO-5EDTA-S and ZGO-5EDTA-B.

ZGO-5EDTA are ascribed to the oxygen of the attached EDTA, which is in accordance with the results of zeta potentials (Fig. 1) and FT-IR (Fig. S4†). Further calculations of the surface atomic concentration from XPS spectra reveal that ZGO has the ratio of $\text{Cr}_{0.005}\text{Zn}_{0.365}\text{GaO}_{2.00}$, while the atom ratios of ZGO-5EDTA-S and ZGO-5EDTA-B are $\text{Cr}_{0.006}\text{Zn}_{0.438}\text{GaO}_{1.85}$ and $\text{Cr}_{0.006}\text{Zn}_{0.428}\text{GaO}_{1.80}$, respectively. Compared with ZGO, the surface of ZGO-5EDTA has more Zn atoms (in accordance with Table 1), but less O atoms. All these XPS data prove that the ZGO-5EDTA nanoparticles (both B and S) have oxygen vacancies.

In ESR (Fig. 5), the signals near 3555 G ascribed to the Cr^{3+} ions in the undisturbed sites of ZnGa_2O_4 host⁴⁹ are both observed in ZGO and ZGO-5EDTA, but ZGO-5EDTA shows the extra conspicuous broad peaks in the range of 2400–4000 G (the g -factor of the peak at 3560 G is 1.9745), which identify the existence of oxygen vacancies.⁵⁰ These oxygen vacancies, related to the deep traps⁵¹ in TL (Fig. 2), are responsible for the maintained or enhanced afterglow of ZGO- n EDTA with reduced size, as discussed above.

Application in afterglow imaging *in vivo*

We examined the *in vivo* afterglow imaging and the corresponding biodistribution of ZGO and ZGO-5EDTA in normal mice (Fig. S10†) and MCF-7 tumor-bearing mice (Fig. 6). The afterglow images with high SNR of both ZGO and ZGO-5EDTA are clearly detected without any external illuminations within 6 h. At 24 h, the afterglow images with high SNR were observed with the help of *in vivo* activation using a 650 nm LED for 1 min. For normal mice, a great difference in the biodistribution of ZGO-5EDTA and ZGO (Fig. S10†) is observed because ZGO displays much higher existence in the lung and the spleen, while the ZGO-5EDTA has slightly higher existence in the liver. For MCF-7 tumor-bearing mice, the SNRs in the liver of the mice injected with ZGO and ZGO-5EDTA are very similar (54 vs. 56) at 3 h. However, the SNR in tumor of the mice injected with ZGO-5EDTA (26) is higher than that of the mice injected with ZGO (15). With passage of time, ZGO-5EDTA displays gradual accumulation in the tumor and gradual elimination in the liver. For example, at 6 h, the SNR in the tumor is 1174, but the SNR in the liver is only 23. The SNR of the tumor and the liver of the mice injected with ZGO-5EDTA gradually increased with time. In contrast, the mice injected with ZGO exhibited gradually decreased SNR ratio in the tumor and the liver (at 3, 6 and 24 h, the ratio is 15/54, 3.7/21 and 4.2/171, respectively). The biodistribution of ZGO and ZGO-5EDTA in the organs and tumor at 24 h (Fig. 6c and d) was in accordance with the *in vivo* images (Fig. 6a and b). The vital biodistribution difference between the mice injected with ZGO-5EDTA and ZGO is most probably ascribed to the much smaller size and much better aqueous dispersibility of ZGO-5EDTA than ZGO,⁴⁰ demonstrating the great superiority of ZGO-5EDTA in biological applications.

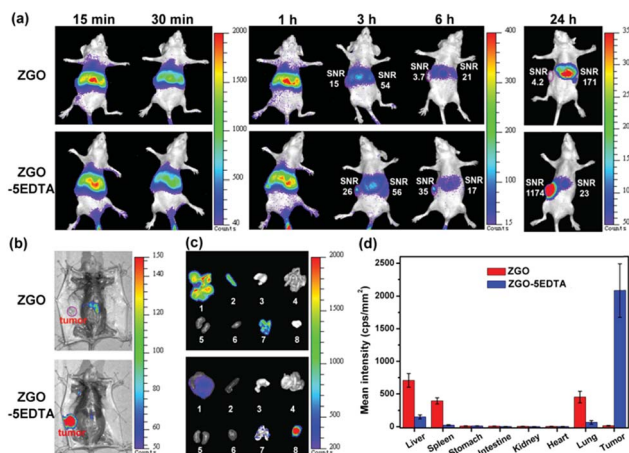


Fig. 6 (a) Afterglow images of MCF-7 tumor-bearing mice after intravenous injection of ZGO and ZGO-5EDTA (0.1 mg, 5 min of irradiation using a 254 nm UV lamp before injection). (b) NIR afterglow images of the mice at 24 h in panel (a), without the superficial skin. (c) Representative afterglow images of isolated organs from the mice in panel (a): (1) liver, (2) spleen, (3) stomach, (4) intestine, (5) kidney, (6) heart, (7) lung, and (8) tumor. (d) Mean afterglow intensity of the organs in panel (c).



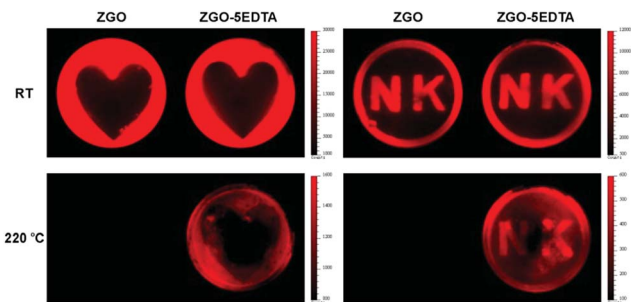


Fig. 7 The information readout by ZGO and ZGO-5EDTA at RT and 220 °C.

Application in information storage

We also tested the potential application of ZGO and ZGO-5EDTA in information storage. The information (the heart shape and N K letters) was first restored by the photomask-protected illumination of 254 nm and then, the afterglow was read using a CCD camera. As shown in Fig. 7, both ZGO and ZGO-5EDTA display fine shapes at room temperature, but only ZGO-5EDTA with deep traps has a recognizable readout at a high temperature of 220 °C. These data further support the existence of deep traps in ZGO-5EDTA, and demonstrate that the ZGO-EDTA from EDTA etching has great potential for information storage.

Conclusions

We present a novel and simple EDTA-etching strategy for regulating the traps, size and aqueous-dispersibility of ZGO-EDTA. The EDTA etching ability of the sintered ZGO results in effective reduction in size and great improvement in the aqueous-dispersibility. Moreover, EDTA etching alters the density of mediate traps and generates new deep traps, thus achieving massive production of (ultra)small ZGO-EDTA with fine aqueous-dispersibility and suitable mediate/deep traps. As EDTA can interact with most metals, this simple EDTA etching strategy would be prospectively applied to other PLNPs, and the resultant PLNPs-EDTA would find wide applications in biological field and information storage.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The authors appreciate Dr Hongwu Zhang and Yanan Liu at Institute of Urban Environment, Chinese Academy of Sciences for help in TL measurements, and the support by the National Natural Science Foundation of China (No. 21575070 and 21435001), the Open Funds of the State Key Laboratory of Electroanalytical Chemistry (SKLEAC201805) and the Fundamental Research Funds for the Central Universities (China). All animal experiments were in accordance with the Guidelines for

the Care and Use of Laboratory Animals of the Chinese Animal Welfare Committee and approved by the Institutional Animal Care and Use Committee, Nankai University Center for Animal Experiment, Tianjin, China.

Notes and references

- 1 Y. Li, M. Gecevicius and J. R. Qiu, *Chem. Soc. Rev.*, 2016, **45**, 2090–2136.
- 2 T. Lecuyer, E. Teston, G. Ramirez-Garcia, T. Maldiney, B. Viana, J. Seguin, N. Mignet, D. Scherman and C. Richard, *Theranostics*, 2016, **6**, 2488–2524.
- 3 S. K. Singh, *RSC Adv.*, 2014, **4**, 58674–58698.
- 4 J. Wang, Q. Q. Ma, Y. Q. Wang, H. J. Shen and Q. Yuan, *Nanoscale*, 2017, **9**, 6204–6218.
- 5 S.-K. Sun, H.-F. Wang and X.-P. Yan, *Acc. Chem. Res.*, 2018, **51**, 1131–1143.
- 6 L.-J. Chen, S.-K. Sun, Y. Wang, C.-X. Yang, S.-Q. Wu and X.-P. Yan, *ACS Appl. Mater. Interfaces*, 2016, **8**, 32667–32674.
- 7 Z. Zhou, W. Zheng, J. Kong, Y. Liu, P. Huang, S. Zhou, Z. Chen, J. Shi and X. Chen, *Nanoscale*, 2017, **9**, 6846–6853.
- 8 B.-Y. Wu, H.-F. Wang, J.-T. Chen and X.-P. Yan, *J. Am. Chem. Soc.*, 2011, **133**, 686–688.
- 9 B.-Y. Wu and X.-P. Yan, *Chem. Commun.*, 2015, **51**, 3903–3906.
- 10 N. Li, Y. H. Li, Y. Y. Han, W. Pan, T. T. Zhang and B. Tang, *Anal. Chem.*, 2014, **86**, 3924–3930.
- 11 A. Abdulkayum, J.-T. Chen, Q. Zhao and X.-P. Yan, *J. Am. Chem. Soc.*, 2013, **135**, 14125–14133.
- 12 Y.-J. Li and X.-P. Yan, *Nanoscale*, 2016, **8**, 14965–14970.
- 13 Y.-C. Lu, C.-X. Yang and X.-P. Yan, *Nanoscale*, 2015, **7**, 17929–17937.
- 14 H.-X. Zhao, C.-X. Yang and X.-P. Yan, *Nanoscale*, 2016, **8**, 18987–18994.
- 15 Z. Pan, Y.-Y. Lu and F. Liu, *Nat. Mater.*, 2012, **11**, 58–63.
- 16 T. Maldiney, A. Bessiere, J. Seguin, E. Teston, S. K. Sharma, B. Viana, A. J. J. Bos, P. Dorenbos, M. Bessodes, D. Gourier, D. Scherman and C. Richard, *Nat. Mater.*, 2014, **13**, 418–426.
- 17 A. Abdulkayum, C.-X. Yang, Q. Zhao, J.-T. Chen, L.-X. Dong and X.-P. Yan, *Anal. Chem.*, 2014, **86**, 4096–4101.
- 18 S. Q. Wu, C. W. Chi, C. X. Yang and X. P. Yan, *Anal. Chem.*, 2016, **88**, 4114–4121.
- 19 Y.-J. Chuang, Z. Zhen, F. Zhang, F. Liu, J. P. Mishra, W. Tang, H. Chen, X. Huang, L. Wang, X. Chen, J. Xie and Z. Pan, *Theranostics*, 2014, **4**, 1112–1122.
- 20 Z. Xue, X. Li, Y. Li, M. Jiang, G. Ren, H. Liu, S. Zeng and J. Hao, *Nanoscale*, 2017, **9**, 7276–7283.
- 21 S. K. Sharma, D. Gourier, E. Teston, D. Scherman, C. Richard and B. Viana, *Opt. Mater.*, 2017, **63**, 51–58.
- 22 F. Liu, W. Z. Yan, Y. J. Chuang, Z. P. Zhen, J. Xie and Z. W. Pan, *Sci. Rep.*, 2013, **3**, 1554.
- 23 Z. Li, Y. Zhang, X. Wu, X. Wu, R. Maudgal, H. Zhang and G. Han, *Adv. Sci.*, 2015, **2**, 1500001.
- 24 Y. Li, S. Zhou, Y. Li, K. Sharafudeen, Z. Ma, G. Dong, M. Peng and J. Qiu, *J. Mater. Chem. C*, 2014, **2**, 2657–2663.



- 25 T. Maldiney, B.-T. Doan, D. Alloeyau, M. Bessodes, D. Scherman and C. Richard, *Adv. Funct. Mater.*, 2015, **25**, 331–338.
- 26 E. Teston, Y. Lalatonne, D. Elgrabli, G. Autret, L. Motte, F. Gazeau, D. Scherman, O. Clement, C. Richard and T. Maldiney, *Small*, 2015, **11**, 2696–2704.
- 27 Y. Wang, C.-X. Yang and X.-P. Yan, *Nanoscale*, 2017, **9**, 9049–9055.
- 28 L. Song, X. H. Lin, X. R. Song, S. Chen, X. F. Chen, J. Li and H. H. Yang, *Nanoscale*, 2017, **9**, 2718–2722.
- 29 J. Shi, X. Sun, J. Li, H. Man, J. Shen, Y. Yu and H. Zhang, *Biomaterials*, 2015, **37**, 260–270.
- 30 T. Maldiney, B. Ballet, M. Bessodes, D. Scherman and C. Richard, *Nanoscale*, 2014, **6**, 13970–13976.
- 31 L.-J. Chen, C.-X. Yang and X.-P. Yan, *Anal. Chem.*, 2017, **89**, 6936–6939.
- 32 B. Zheng, H.-B. Chen, P.-Q. Zhao, H.-Z. Pan, X.-L. Wu, X.-Q. Gong, H.-J. Wang and J. Chang, *ACS Appl. Mater. Interfaces*, 2016, **8**, 21603–21611.
- 33 S.-Q. Wu, C.-X. Yang and X.-P. Yan, *Adv. Funct. Mater.*, 2017, **27**, 1604992.
- 34 R. Abdurahman, C.-X. Yang and X.-P. Yan, *Chem. Commun.*, 2016, **52**, 13303–13306.
- 35 J. Shi, X. Sun, S. Zheng, J. Li, X. Fu and H. Zhang, *Biomaterials*, 2018, **152**, 15–23.
- 36 Y.-J. Li, C.-X. Yang and X.-P. Yan, *Anal. Chem.*, 2018, **90**, 4188–4195.
- 37 Y. Zhuang, Y. Lv, L. Wang, W. Chen, T.-L. Zhou, T. Takeda, N. Hirosaki and R.-J. Xie, *ACS Appl. Mater. Interfaces*, 2018, **10**, 1854–1864.
- 38 Y. Zhuang, L. Wang, Y. Lv, T.-L. Zhou and R.-J. Xie, *Adv. Funct. Mater.*, 2018, **28**, 1705769.
- 39 X. Xu, L. Yan, X. Yu, H. Yu, T. Jiang, Q. Jiao and J. Qiu, *Mater. Lett.*, 2013, **99**, 158–160.
- 40 T. Maldiney, C. Richard, J. Seguin, N. Wattier, M. Bessodes and D. Scherman, *ACS Nano*, 2011, **5**, 854–862.
- 41 Z. Li, Y. Zhang, X. Wu, L. Huang, D. Li, W. Fan and G. Han, *J. Am. Chem. Soc.*, 2015, **137**, 5304–5307.
- 42 J. P. Shi, X. Sun, J. F. Zhu, J. L. Li and H. Zhang, *Nanoscale*, 2016, **8**, 9798–9804.
- 43 B. B. Srivastava, A. Kuang and Y. Mao, *Chem. Commun.*, 2015, **51**, 7372–7375.
- 44 E. Teston, S. Richard, T. Maldiney, N. Lievre, G. Y. Wang, L. Motte, C. Richard and Y. Lalatonne, *Chem.–Eur. J.*, 2015, **21**, 7350–7354.
- 45 A. Bessiere, S. K. Sharma, N. Basavaraju, K. R. Priolkar, L. Binet, B. Viana, A. J. J. Bos, T. Maldiney, C. Richard, D. Scherman and D. Gourier, *Chem. Mater.*, 2014, **26**, 1365–1373.
- 46 S. P. Yeap, *Powder Technol.*, 2018, **323**, 51–59.
- 47 Y. Yuan, W. Du and X. Qian, *J. Mater. Chem.*, 2012, **22**, 653–659.
- 48 M. Xing, J. Zhang, F. Chen and B. Tian, *Chem. Commun.*, 2011, **47**, 4947–4949.
- 49 D. Gourier, A. Bessiere, S. K. Sharma, L. Binet, B. Viana, N. Basavaraju and K. R. Priolkar, *J. Phys. Chem. Solids*, 2014, **75**, 826–837.
- 50 J. S. Kim, H. I. Kang, W. N. Kim, J. I. Kim, J. C. Choi, H. L. Park, G. C. Kim, T. W. Kim, Y. H. Hwang, S. I. Mho, M. C. Jung and M. Han, *Appl. Phys. Lett.*, 2003, **82**, 2029–2031.
- 51 J. Su, S. Ye, X. Yi, F. Q. Lu, X. B. Yang and Q. Y. Zhang, *Opt. Mater. Express*, 2017, **7**, 734–743.

