

Dalton Transactions

Accepted Manuscript



This is an *Accepted Manuscript*, which has been through the Royal Society of Chemistry peer review process and has been accepted for publication.

Accepted Manuscripts are published online shortly after acceptance, before technical editing, formatting and proof reading. Using this free service, authors can make their results available to the community, in citable form, before we publish the edited article. We will replace this *Accepted Manuscript* with the edited and formatted *Advance Article* as soon as it is available.

You can find more information about *Accepted Manuscripts* in the [Information for Authors](#).

Please note that technical editing may introduce minor changes to the text and/or graphics, which may alter content. The journal's standard [Terms & Conditions](#) and the [Ethical guidelines](#) still apply. In no event shall the Royal Society of Chemistry be held responsible for any errors or omissions in this *Accepted Manuscript* or any consequences arising from the use of any information it contains.

**A Force Field for Tricalcium Aluminate to Characterize
Surface Properties, Initial Hydration, and Organically
Modified Interfaces in Atomic Resolution**

by

Ratan K. Mishra,¹ Lucia Fernández-Carrasco,² Robert J. Flatt,^{1*} and Hendrik
Heinz^{3*}

¹ Department of Civil, Environmental and Geomatic Engineering, ETH Zurich, CH-
8093 Zürich, Switzerland

² Department of Architectural Technology I, Universitat Politècnica de Catalunya,
08028 Barcelona, Spain

³ Department of Polymer Engineering, University of Akron, Akron, OH 44325,
USA

* Corresponding authors: flatt@ethz.ch, hendrik.heinz@uakron.edu

Abstract

Tricalcium aluminate (C_3A) is a major phase of Portland cement clinker and dental root filling cements. An accurate all-atom force field is introduced to examine structural, surface, and hydration properties as well as organic interfaces to overcome challenges using current laboratory instrumentation. Molecular dynamics simulation demonstrates excellent agreement of computed structural, thermal, mechanical, and surface properties with available experimental data. The parameters are integrated into multiple harmonic energy expressions, including the PCFF, CHARMM, AMBER, OPLS, and INTERFACE force fields, to enable simulations of a wide range of inorganic-organic interfaces at a million times lower computational cost than DFT methods at the 1 to 100 nm scale. Molecular models of dry and partially hydrated surfaces are introduced to examine cleavage, agglomeration, and the role of adsorbed organic molecules. Cleavage of crystalline tricalcium aluminate requires approximately 1300 mJ/m^2 and superficial hydration introduces an amorphous calcium hydroxide surface layer that reduces the agglomeration energy from 850 to 250 mJ/m^2 . The adsorption of several alcohols and amines was examined to understand their action as grinding aids and setting modifiers in cement. The molecules mitigate local electric fields through complexation of calcium ions, hydrogen bonds, and introduction of hydrophobicity upon binding. Molecularly thin layers of about 0.5 nm thickness reduce agglomeration energies to between 100 and 30 mJ/m^2 . Molecule-specific trends were found to be similar for tricalcium aluminate and tricalcium silicate. The models are a starting point to provide fundamental understanding of the role of C_3A and organic additives in cement. Extensions to impure phases and advanced hydration stages are feasible.

1. Introduction

Concrete is a heterogeneous and complex multi-scale material “par excellence”.¹ An enormous annual production of 1.6 billion tons of cement,^{2,3} however, also contributes 5-8% of worldwide CO₂ emissions.^{4,5} In addition, particle grinding consumes 26% of the total required electrical energy.³ The use of less energy-intensive, supplementary cementitious materials can lower carbon emissions, and small amounts of organic additives reduce electricity consumption during grinding.⁶⁻⁸ Nevertheless, rational design of cement properties such as particle cleavage, hydration reactions, setting kinetics, and strength development with modified composition remains a key challenge facing the industry.⁹ The basic chemistry and physics of grinding and hydration from nanoscale to macroscale still poses many open questions. In particular, nanometer-scale phenomena elude direct observation and require new methods of control including guidance by simulation.^{5,10}

A wide array of characterization techniques has thus been applied to cement phases, for example, X-ray diffraction (XRD), small angle neutron scattering (SANS), in situ imaging by transmission electron microscopy (TEM), atomic force microscopy (AFM) and nanoindentation, zeta potential measurements, nuclear magnetic resonance spectroscopy (NMR), Infrared (IR), and Raman spectroscopy.¹¹⁻¹⁸ In addition, models and simulations have been developed to enhance understanding and materials development.^{8, 10, 19-30} Such approaches include electronic structure calculations, molecular dynamics, Monte Carlo, and continuum methods, as well as multi-scale combinations of methods.

Quantum-mechanical techniques such as density functional theory (DFT) can provide insight into the kinetics and thermodynamics of cement hydration at the scale of chemical bonds.^{21, 22, 26, 28} However, the methods cannot reach beyond hundreds of atoms and picosecond time scales. Simulations of actual nanocrystals, aqueous interfaces, surface reactions, selective adsorption of organic modifiers, and assembly of nano-objects thus require classical molecular dynamics simulations.^{8, 10, 19, 23, 24} Such models retain all-atomic detail for up to a million atoms and access time scales up to microseconds. Access to yet larger systems of micrometer-size particles, grinding, sedimentation, and aggregate mechanical properties is feasible using coarse-grain and continuum methods.^{27, 31, 32}

To date, modeling studies of cement-related materials have primarily focused on silicates and sulfates at various length scales, and few examples for aluminates are found.^{10, 22, 26} Prior reports on atomistic simulations of tricalcium aluminate ($\text{Ca}_3\text{Al}_2\text{O}_6$, C_3A) at the scale of 1 to 100 nm are not known to us, and a needed force field was not available. This contribution aims at the introduction and thorough validation of an atomistic force field for tricalcium aluminate along with first applications to hydration and adsorption of organic modifiers.^{10, 33-35}

The outline of this paper is as follows. Key properties of tricalcium aluminate, the approach toward a force field, and parameters are described in section 2. Validation of the model and insight into structural, thermal, mechanical, and vibration properties is discussed in section 3. In section 4, surface properties, model representations of the hydration reaction, and interactions between cleaved surfaces before and after initial hydration are explored. In section 5, the structure of organic thin films on the surfaces and their impact on agglomeration is analyzed. Conclusions are presented in section 6.

Full computational details and additional information can be found in the Supporting Information.

2. Derivation of the Force Field

2.1. Key Properties of Tricalcium Aluminate. Tricalcium aluminate ($\text{Ca}_3\text{Al}_2\text{O}_6$, C_3A) is the main aluminate phase in cement clinker (5%-10%) and also used in dental root filling cements based on mineral trioxide aggregate (MTA).³⁶ Pure tricalcium aluminate possesses a cubic crystal structure (Figure 1). It is highly hygroscopic, forms poorly crystalline hydration products, and contributes to rapid setting of cement.³⁷⁻³⁹ The exact concentration, or mass percentage, of tricalcium aluminate strongly influences cement setting properties and is important for the early development of compressive strength.^{40, 41} C_3A also forms a wide range of solid solutions, for example by replacement of Al^{3+} by Fe^{3+} and Si^{4+} in the range of 0-5%, and substitution of Ca^{2+} by Na^+ , K^+ , and Mg^{2+} in the range 0-2%. An impurity content over 3% can modify the crystal structure from cubic to orthorhombic and several other polymorphs.^{37, 42, 43} In the derivation of the model, data for pure C_3A ^{26, 44-51} as well as $\text{Ca}(\text{OH})_2$ are relied upon.⁵² Customized local substitutions can be implemented as described earlier for clay minerals.^{20, 53, 54}

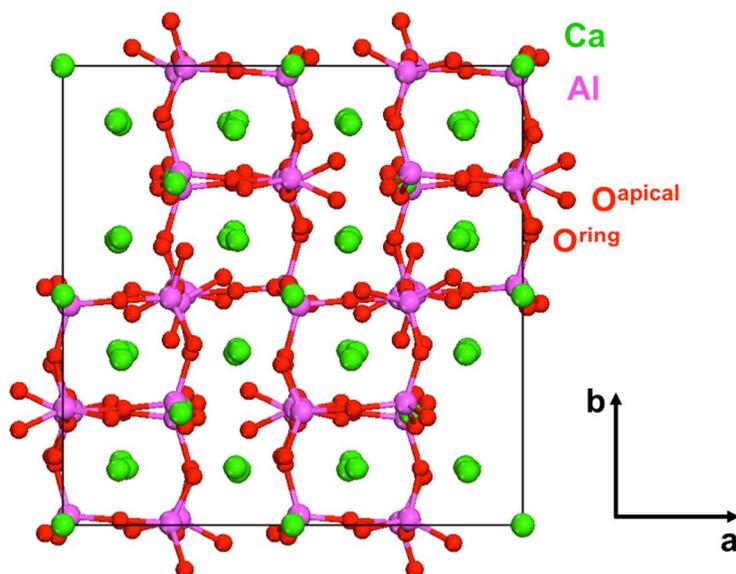


Figure 1. Unit cell of tricalcium aluminate. The lattice constant is 1.5263 nm and distinct atom types are shown (cubic, space group $Pa\bar{3}$, ref. ⁴⁸). Viewed from the c axis onto the ab plane.

2.2. Approach Toward an Atomistic Force Field. Molecular models of several cement minerals are available that reproduce structural and mechanical properties in good agreement with experiment.^{8, 10, 19, 21-28} The largest portion of cement minerals is covered by the INTERFACE force field,^{8, 10} which is also designed for compatibility with aqueous and organic interfaces and accurately reproduces interfacial properties in comparison to experiment. Such properties include cleavage energies, solid-water interface tensions, and binding energies of organic molecules to solid surfaces that are central for material performance. The uncertainties are in the low percent range in comparison to tens of percent in density functionals and deviations up to hundreds of percent in other force fields. Compounds in the INTERFACE force field are parameterized individually and typically reproduce structural properties with <1%

deviation from experiment, surface/interface properties with <5% deviation from experiment, mechanical properties with <10% deviation from experiment, and vibration spectra in good agreement with experiment.

The accessibility of accurate bulk and interfacial properties depends mainly on three premises. (1) The balance of ionic versus covalent bonding is accurately represented by atomic charges.^{10, 20} This condition is fulfilled by assignments consistent with measured X-ray deformation electron densities, dipole moments, an Extended Born model, and known chemical reactivity for electronically and structurally similar compounds across the periodic table. (2) Computed surface properties are validated in comparison to available experimental data.^{10, 54} This aspect is critical to avoid errors up to 500% as seen in earlier models and force fields derived from ab-initio methods. (3) Simplicity and broad applicability of the energy expression is desired. Simplicity of the energy expressions leads to fewer parameters, enables full interpretation of the parameters with a chemical and physical rationale, and eases extensions to related compounds as well as to chemical reactions. Broad applicability of the energy expression facilitates the usage of existing, thermodynamically consistent parameters for solvents and macromolecules in other force fields via standard combination rules to simulate inorganic-organic interfaces. These conditions of the INTERFACE approach enable the coverage of compounds across the periodic table (e.g. cement minerals, metals, biopolymers)^{8, 10, 54, 55} in high accuracy and full compatibility with harmonic force fields for biopolymers and organic molecules (CFF, PCFF,⁵⁶⁻⁵⁸ COMPASS,⁵⁹ CHARMM,⁶⁰ AMBER,⁶¹ OPLS-AA,⁶² CVFF⁶³). The new parameters for tricalcium aluminate and partially hydrated phases contribute to this platform.¹⁰

2.3. Choice of Energy Expressions. Therefore, to facilitate broad applicability, multiple harmonic energy expressions were chosen. These include common force fields for organic molecules, biopolymers, solvents, and other inorganic phases, namely, CFF, PCFF,⁵⁶⁻⁵⁸ COMPASS (eq 1),⁵⁹ CHARMM,⁶⁰ AMBER,^{61, 64} CVFF,⁶³ and OPLS-AA⁶² (eq 2):

$$E_{pot} = \sum_{ij \text{ bonded}} K_{r,ij} (r_{ij} - r_{0,ij})^2 + \sum_{ijk \text{ bonded}} K_{\theta,ijk} (\theta_{ijk} - \theta_{0,ijk})^2 + \frac{1}{4\pi\epsilon_0} \sum_{\substack{ij \text{ nonbonded} \\ (1,3 \text{ excl})}} \frac{q_i q_j}{r_{ij}} + \sum_{\substack{ij \text{ nonbonded} \\ (1,3 \text{ excl})}} \epsilon_{ij} \left[2 \left(\frac{\sigma_{ij}}{r_{ij}} \right)^9 - 3 \left(\frac{\sigma_{ij}}{r_{ij}} \right)^6 \right] \quad (1)$$

$$E_{pot} = \sum_{ij \text{ bonded}} K_{r,ij} (r_{ij} - r_{0,ij})^2 + \sum_{ijk \text{ bonded}} K_{\theta,ijk} (\theta_{ijk} - \theta_{0,ijk})^2 + \frac{1}{4\pi\epsilon_0} \sum_{\substack{ij \text{ nonbonded} \\ (1,3 \text{ excl})}} \frac{q_i q_j}{r_{ij}} + \sum_{\substack{ij \text{ nonbonded} \\ (1,3 \text{ excl})}} \epsilon_{ij} \left[\left(\frac{\sigma_{ij}}{r_{ij}} \right)^{12} - 2 \left(\frac{\sigma_{ij}}{r_{ij}} \right)^6 \right] \quad (2)$$

The energy expressions contain harmonic bond stretching and angle bending terms for bonded interactions, as well as Coulomb and Lennard-Jones terms for nonbonded interactions (electrostatic and van-der Waals contributions). This combination offers a suitable description of molecular features in minerals, as previously shown,^{8, 10, 53, 54} and the application of standard combination rules enables the simulation of interfaces with solvents, organic, and inorganic molecules in high accuracy.¹⁰

2.4. Derivation and Interpretation of Individual Parameters. The parameterization procedure followed the INTERFACE approach that involves (1) retrieval of the crystal structure of tricalcium aluminate from X-ray data, (2) definition of

atom types, (3) assignment of atomic charges, (4) assignment of initial bonded and Lennard-Jones parameters, (5) computational tests of density and geometry followed by refinements, (6) computational tests of surface properties followed by refinements, and (7) secondary validation (vibration, thermal, and mechanical properties).¹⁰

X-ray data reveal that pure tricalcium aluminate consists of Al_6O_{18} rings and calcium ions with a net stoichiometry $\text{Ca}_3\text{Al}_2\text{O}_6$ (Figure 1 and Figure 2).⁴⁸ Four chemically distinct atom types can be identified similar to the concept of chemical equivalence in NMR spectroscopy. These include calcium, aluminum, tetrahedral oxygen, and ring oxygen. This level of differentiation proved to be sufficient for a well-performing force field. Four further atom types are generated upon partial hydration (Table 1), including oxygen and hydrogen in aluminol groups as well as oxygen and hydrogen in hydroxide ions. These atom types share similarities with AlOH groups in ettringite and with OH^- ions in calcium hydroxide, respectively.^{8, 10}

Atomic charges q_i were then assigned on the basis of multiple experimental properties according to Heinz et al (Figure 2 and Table 1).²⁰ Key data are atomic charges derived from X-ray deformation densities for closely related compounds, dipole moments, and an extended Born model. The extended Born model relies on available atomization energies, ionization potentials, and coordination numbers, as well as comparisons to structurally and electronically similar compounds across the periodic table. Melting points, acid-base reactivity, affinity to hydrolysis, and ab-initio charges were also taken into account as previously described, aiding in convergence testing of the derived atomic charges.^{10, 20} The resulting aluminum charge in the Al_6O_{18} rings in tetrahedral oxygen coordination is $+1.2\pm 0.1e$, slightly lower than $+1.45e$ in octahedral

oxygen coordination as found in diaspore and in pyrophyllite.^{10,20,54} Accordingly, Al—O bonds in tetrahedral oxygen coordination have more covalent character and are slightly shorter than in octahedral oxygen coordination.⁴⁸ Calcium is embedded in coordination environments similar to tricalcium silicate and hydroxyapatite with a charge of $+1.5 \pm 0.1e$, resulting in partly covalent bonding that is also common in other calcium salts.^{10,20} The corresponding negative charges on oxygen were distributed over ring and apical sites under the condition of overall charge neutrality. A higher negative charge is concentrated on the monovalent apical oxygen atoms ($-1.35e$) and a lower negative charge on the divalent ring oxygen atoms ($-0.75e$), in accordance with less covalent character and closer contact of apical oxygen atoms with calcium ions (Figure 2 and Table 1).⁴⁸ The exact balance was confirmed through best structural fit in the simulation. Upon initial reaction with water, apical oxygen atoms form aluminol groups and hydroxide ions. Aluminol hydrogen is less acidic than silanol hydrogen ($+0.40e$); therefore it carries a charge of $+0.30e$ and leads to a charge of $-0.90e$ on aluminol oxygen.^{20,54} Hydroxide ions are approximated by charges of $-1.0e$ and $+0.25e$ for O and H for a total charge of $-0.75e$.

Bonded terms were assigned between neighbor atoms that carry less than or equal to half the formal charge and have a short distance, corresponding to a predominantly covalent bond (Table 1).¹⁰ In addition, all atoms participate in nonbond interactions. Accordingly, only atoms in the Al_6O_{18} rings possess bonded terms and the initial bonded parameters $r_{0,ij}$ and $\theta_{0,ijk}$ were chosen from the X-ray structure.⁴⁸ Final parameters $r_{0,ij}$ for the Al—O bonds were obtained by refinement so that the average bond length of 175 pm in X-ray data is reproduced in NPT molecular dynamics simulation (Table 1). The

precise values $r_{0,ij}$ are up to 13% higher to offset strongly attractive Coulomb interactions that depend on scaling conventions of nonbond interactions between 1,4 bonded atoms. Accordingly, $r_{0,ij}$ is 198 pm (ring) and 181 pm (apical) in PCFF and CHARMM with 100% inclusion of 1,4 Coulomb forces, but only 181.5 pm and 177.5 pm in OPLS-AA with 50% inclusion of 1,4 Coulomb forces. Final parameters $\theta_{0,ijk}$ for the bond angles were obtained upon refinement in NPT molecular dynamics simulation so that average bond angles from X-ray data are reproduced. Modifications were <3%. Vibration constants $K_{r,ij}$ and $K_{\theta,ijk}$ were initially estimated from clay minerals and refined to reasonably reproduce Infrared (IR) and Raman spectra during molecular dynamics.^{45-47, 49, 51, 54} The resulting bond stretching constants $K_{r,ij}$ for Al–O bonds are uniformly 180 kcal/(mol·Å²) and angle bending constants $K_{\theta,ijk}$ for Al–O–Al and O–Al–O angles are uniformly 110 kcal/(mol·rad²).

Initial Lennard-Jones (LJ) parameters σ_{ii} and ε_{ii} for each atom type were assigned on the basis of van-der-Waals radius, polarizability, and coordination environment (Table 1) (see refs. ^{10, 54, 65}). The LJ parameters were subsequently refined for best reproduction of lattice constants and structural match with X-ray data for each energy expression, including extensive tests of the well depths ε_{ii} and nonbond diameters σ_{ii} . For a highly polar mineral such as C₃A, larger well depths ε_{ii} increase interatomic repulsion by increasing the slope of the repulsive portion of the LJ potential and have a similar effect as a larger nonbond diameter σ_{ii} on computed cell parameters. Secondary validation of surface, thermal, and mechanical properties provided instantly a good match with available experimental data. This is a consequence of the chemically informed

assignment, interpretation of all parameters in the context of verifiable experimental properties, and overall thermodynamic consistency. Finally, the influence of all parameters on key properties was individually tested in more than 100 further test calculations and minor adjustments were made.

The parameters using a 9-6 LJ potential (eq 1) were then transcribed to energy expressions with a 12-6 LJ potential (eq 2) using known trends^{8, 54, 55} and fine-tuned to reproduce lattice constant and structural parameters. Some additional differentiation in LJ parameters σ_{ii} , ϵ_{ii} , equilibrium bond lengths $r_{0,ij}$, and angles $\theta_{0,ijk}$ was made due to specific scaling conventions of nonbond interactions between 1,4 bonded atoms and combination rules. Finally, the influence of all parameters on key properties was individually tested for each energy expression in over 500 further test calculations and minor adjustments were made. All energy expressions (eq 1 and eq 2) compute essentially the same properties. The parameters for C₃A are consistent with those for other compounds in the INTERFACE force field, including layered silicates, silica, and tricalcium silicate that underwent similarly rigorous evaluation.^{8, 10, 54}

The assignment and refinement of force field parameters on the basis of chemical knowledge is a great advantage over “black-box” assignments from ab-initio simulations, which typically lead to poor results and lack rational explanation. Chemical accuracy in atomic charges was achieved for pure and hydrated phases, as well as full understanding of all parameters and relative simplicity of the force field, which is paramount to examine chemical reactions. Therefore, the force field can be seen as a “molecular code” to understand nanoscale properties of tricalcium aluminate and its initial hydrates.

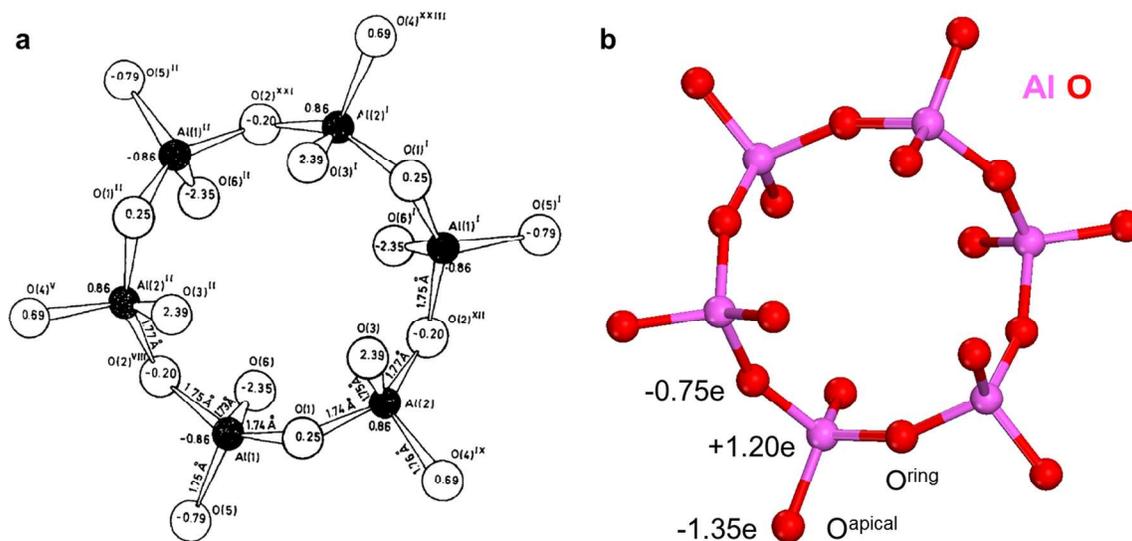


Figure 2. The structure of aluminate rings in tricalcium aluminate according to X-ray data and computation with the force field. (a) Al_6O_{18} rings consist of six AlO_4 tetrahedra connected by bridging O atoms according to experiment (from ref. ⁴⁸). Ring atoms exhibit minor variations in bond length, and bond angles deviate somewhat from ideal tetrahedral values due to a puckered structure. (b) Equilibrium geometry of the aluminate ring according to MD simulation in the NPT ensemble. A close match to experiment can be seen and atomic charges in the force field are indicated.

Table 1. Force field parameters for tricalcium aluminate and initially hydrated surfaces. Different energy expressions such as CFF/PCFF/COMPASS, CHARMM, CVFF, AMBER, OPLS-AA have some parameters in common ($q_i, K_{r,ij}, K_{\theta,ijk}$) while others are specific due to minor differences in functional form ($\sigma_{ii}, \varepsilon_{ii}, r_{0,ij}, \theta_{0,ijk}$).

		σ_{ii} (pm), ε_{ii} (kcal/mol)		
I. Nonbond	q_i (e)	CFF, PCFF, COMPASS ^a	CHARMM, CVFF ^{b,c}	AMBER, OPLS- AA ^{b,c}
Ca	+1.50	355, 0.24	330, 0.15	330, 0.15
Al	+1.20	480, 0.50	460, 0.50	460, 0.50
O ^{ring}	-0.75	362, 0.06	350, 0.08	342, 0.08
O ^{apical}	-1.35	352, 0.06	340, 0.08	332, 0.08
O in AlOH	-0.90	360, 0.06	347, 0.08	347, 0.08
H in AlOH	+0.30	109.8, 0.013	0, 0 ^d	0, 0 ^d
O in OH ⁻	-1.00	360, 0.12	347, 0.12	347, 0.12
H in OH ⁻	+0.25	109.8, 0.013	0, 0 ^d	0, 0 ^d
II. Bonds		$r_{0,ij}$ (pm)	$K_{r,ij}$ [kcal/(mol·Å ²)]	
Al-O ^{ring}		198 (AMBER: 194; OPLS-AA: 181.5) ^c	180	
Al-O ^{apical}		181 (OPLS-AA: 177.5) ^c	180	
O-H (in AlOH and OH ⁻)		92.9	495	
III. Angles		$\theta_{0,ijk}$ (°)	$K_{\theta,ijk}$ [kcal/(mol·rad ²)]	
O ^{ring} -Al-O ^{ring}		111.5	110	
Al-O ^{ring} -Al		156.7 (OPLS-AA: 152.0) ^c	110	
O ^{ring} -Al-O ^{apical}		105.0	110	

$\text{O}^{\text{apical}}-\text{Al}-\text{O}^{\text{apical}}$	121.0	110
$\text{H}^{\text{ap}}-\text{O}^{\text{ap}}-\text{Al}$ (in AlOH)	115.0	40

^a Eq 1 with 9-6 LJ potential.

^b Eq 2 with 12-6 LJ potential.

^c Specific scaling of nonbond interactions between 1,4 bonded atoms requires some differences in parameters (CFF, PCFF, COMPASS, CHARMM, CVFF: fully included; AMBER: $\frac{1}{2}$ vdW, $\frac{5}{6}$ Coulomb; OPLS-AA: $\frac{1}{2}$ vdW, $\frac{1}{2}$ Coulomb). Specific combination rules to obtain $\sigma_{0,ij}$ (arithmetic in CHARMM versus geometric in CVFF) require no further differentiation in parameters for C₃A.

^d The zero 12-6 LJ potential for hydrogen equals hydrogen in SPC and TIP3P water. Nonzero parameters, e.g., (108.5, 0.016), change computed interfacial energies <5%.

3. Validation of Structural, Thermal, Mechanical, and Vibration Properties

The force field computes structural, thermal, mechanical, and vibration properties in very good agreement with experiment, which helps establish the reliability of the force field and explains molecular details. Surface properties will be discussed in section 4.

3.1. Structural and Thermal Properties. NPT molecular dynamics simulation of tricalcium aluminate under standard conditions reproduces bond lengths and angles of the Al₆O₁₈ rings with less than 1% average deviation from X-ray data (Figure 2, Table 2). The root mean square (rms) deviation in atomic positions from X-ray measurements is between 0.03 and 0.05 Å. Computed unit cell parameters and density agree better than 0.5% with X-ray data for all energy expressions (Table 3).

The crystal structure is also stable in NPT simulations up to 1500 °C. Increasing fluctuations in atomic positions are noted towards the melting and decomposition point at 1542 °C in experiment. The computed temperature dependence of the lattice parameter indicates near-linear thermal expansion between 300 K and 800 K (Figure 3), and the linear thermal expansion coefficient averages $11.5 \pm 1.5 \cdot 10^{-6} \text{ K}^{-1}$ in agreement with measurements ranging from $10 \cdot 10^{-6}$ to $15 \cdot 10^{-6} \text{ K}^{-1}$.⁴⁴ The small number of force field types and parameters thus reproduces structural and thermal properties very well.

Table 2. Comparison of average bond lengths and angles of the puckered Al_6O_{18} ring according to X-ray data (ref. ⁴⁸) and in NPT simulation at 298 K (PCFF).

Bond lengths (pm)			Bond angles (°)		
Bond	Exptl.	Sim.	Angle	Exptl.	Sim.
$\text{Al-O}^{\text{ring}}$	175.5 ± 1.5	174.5 ± 2.0	$\text{O}^{\text{ring}}-\text{Al}-\text{O}^{\text{ring}}$	111.8 ± 2.6	112.7 ± 3.0
$\text{Al-O}^{\text{apical}}$	175.0 ± 2.0	174.5 ± 2.0	$\text{Al-O}^{\text{ring}}-\text{Al}$	152.3 ± 2.9	151.0 ± 4.0
			$\text{O}^{\text{ring}}-\text{Al}-\text{O}^{\text{apical}}$	105.4 ± 3.7	105.5 ± 4.5
			$\text{O}^{\text{apical}}-\text{Al}-\text{O}^{\text{apical}}$	123.1 ± 1.3	123.3 ± 3.0

Table 3. Cell parameters of tricalcium aluminate according to X-ray data (ref. ⁴⁸) and molecular dynamics simulation with different force fields in the NPT ensemble under standard temperature and pressure.

Method	Cell	a (nm)	b (nm)	c (nm)	α (°)	β (°)	γ (°)	ρ (kg/m ³)
	dim.							
Experiment	1×1×1	1.5263	1.5263	1.5263	90	90	90	3027 (2)
PCFF, CFF, COMPASS	1×1×1	1.5265	1.5260	1.5260	90	90	90	3028
CHARMM	1×1×1	1.5280	1.5280	1.5280	90	90	90	3017
CVFF	1×1×1	1.5240	1.5240	1.5240	90	90	90	3040
AMBER	1×1×1	1.5267	1.5267	1.5267	90	90	90	3025
OPLS-AA	1×1×1	1.5258	1.5258	1.5258	90	90	90	3030

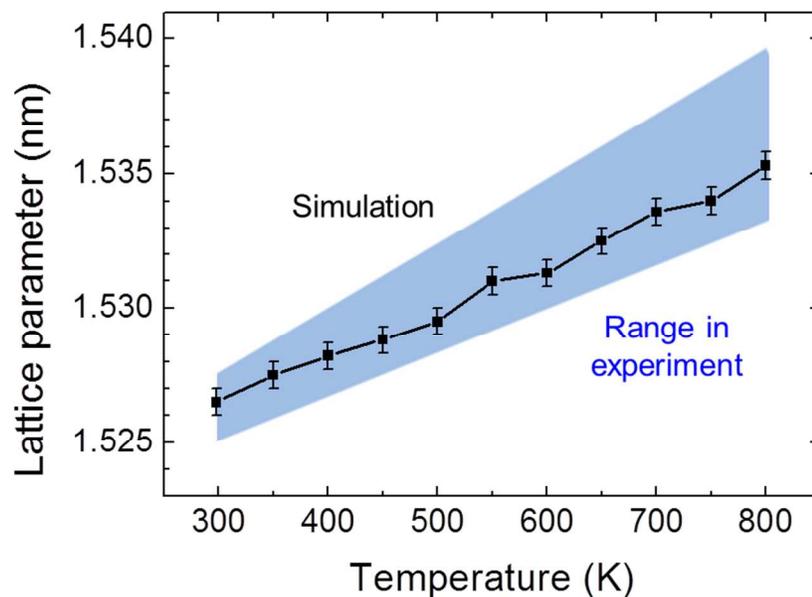


Figure 3. Thermal expansion of the lattice parameter a of C_3A under standard pressure in NPT molecular dynamics simulation and in experimental measurements (shaded area according to ref. ⁴⁴).

3.2. Mechanical Properties. Tricalcium aluminate is known to play an important role in the early development of compressive strength of cement through the various hydrates that form in the reaction with water.^{22, 26, 50} Besides, also the mechanical properties of pure C_3A have been studied using nanoindentation, high-pressure X-ray diffraction, and resonance frequency methods,^{26, 50} as well as by density functional theory (Table 4).^{22, 26} Results with the force field fall fully within the experimental uncertainty and can be applied to length and time scales a million times larger than with ab-initio methods (Figure 4 and Table 4).

Computed isotropic compression and stress-strain curves proceed approximately linear as a function of stress. The volume decrement upon compression slightly decreases under higher pressure, leading to an increase in bulk modulus at higher pressure (Figure 4a). These observations agree with experimental compression data, including a bulk modulus K near 100 GPa at low stress (<1 GPa) and a higher bulk modulus over 120 GPa at high stress (>4 GPa).²⁶ The trend is also consistent with a decreasing slope of the experimental compression curve and with resonance frequency measurements of Young's moduli E up to 160 GPa,⁵⁰ which correspond to Bulk moduli up to 127 GPa according to the relationship $K = E/[3(1 - 2\nu)]$ for isotropic elastic solids (Table 4).⁶⁶ Computed uniaxial stress-strain curves along the three principal axes x , y , and z converge to the same Young's modulus E of 134 ± 5 GPa within small deviations (Figure 4b). The Poisson

ratio ν was identified as ~ 0.29 . Measurements indicate Young's moduli in the range 135 to 145 GPa at low stress.^{26, 50}

The capability of quantitative predictions of mechanical properties can also be applied to heterogeneous mineral structures in dry and hydrated cements along with existing parameters for other cement phases.¹⁰

Table 4. Experimental and computed elastic properties of tricalcium aluminate, using molecular dynamics simulation in the NPT ensemble.

Elastic Properties	Experiment	Computed (PCFF)
Bulk modulus (K), GPa	106 ± 8^a (DFT: 102.9^b)	98 ± 3
Young's modulus (E), GPa	135 ± 7^a , 145 ± 10^c	134 ± 5
Poisson ratio	0.285 ± 0.01^a	0.29 ± 0.03

^a Ref. ²⁶. ^b Ref. ²². ^c Ref. ⁵⁰. The Young's modulus was identified as (145 ± 10) GPa by nanoindentation and up to higher values of (160 ± 10) GPa by resonance frequency techniques.

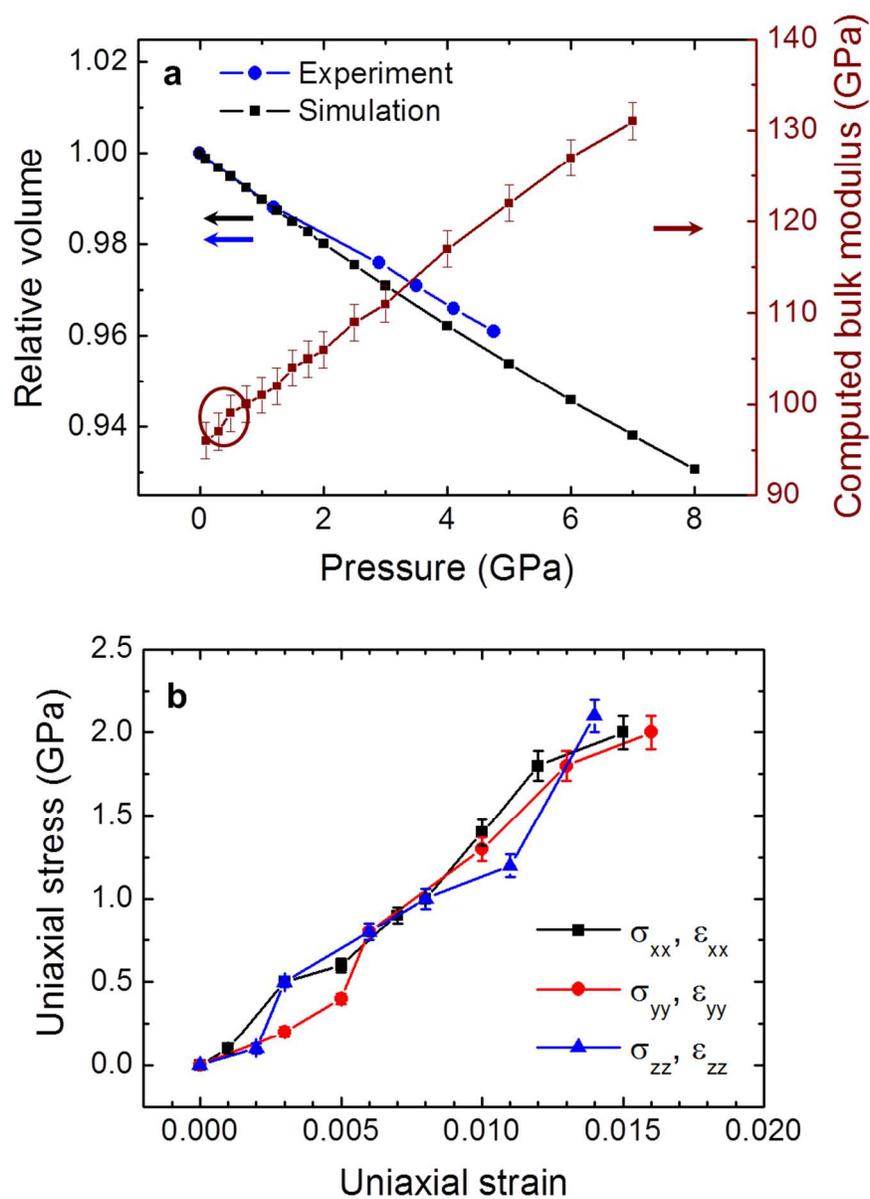


Figure 4. Isothermal compression and stress-strain curves of tricalcium aluminate along the principal directions in NPT molecular dynamics simulation at 298 K. (a) The volume decreases approximately linear as a function of pressure, with a softening slope, in agreement with experimental compression data (from ref. ²⁶). The bulk modulus increases approximately linear under higher pressure. (b) Computed stress-strain curves in x, y, and z direction. The average slope at low strain equals Young's modulus.

3.3. Vibration Properties. IR⁵¹ and Raman spectra³⁸ of tricalcium aluminate are also approximated in molecular dynamics simulation (Figure 5). Absorption bands in the frequency range of 920 to 700 cm^{-1} correspond to stretching vibrations of Al–O bonds in the aluminate rings^{45-47, 49, 51} and a very good fit within 20 cm^{-1} between model and experiment is seen. Intensities are only approximately reproducible in a non-quantum mechanical model.⁵⁴ The frequency range of 550 to 300 cm^{-1} corresponds to bending vibrations of Al–O–Al and O–Al–O angles. These bands are shifted by approximately 100 cm^{-1} upwards in molecular dynamics simulation to support structural stability and accuracy of bond angles up to high temperature (see section S1.5 for details).

Bands in the range of 700 to 500 cm^{-1} were also observed in measurements although they are likely associated with condensed AlO_6 octahedra upon progressive hydration.^{45-47, 49, 51, 67} Such structural elements are not present in this model although they can be introduced using structural features in ettringite or monosulphate.¹⁰

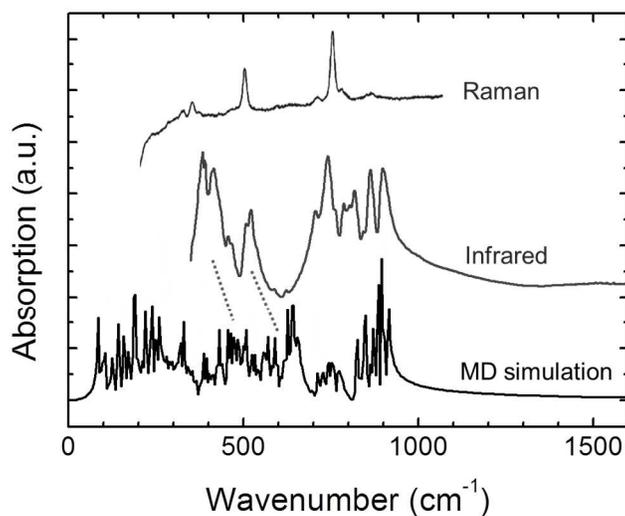


Figure 5. Vibration spectra of tricalcium aluminate in experiment and in molecular dynamics simulation (Raman data from ref. ³⁸ and IR data from ref. ⁵¹). Regions for angle bending ($300\text{-}550\text{ cm}^{-1}$) and bond stretching ($700\text{-}920\text{ cm}^{-1}$) can be seen. A slight mismatch in angle bending vibrations can be seen.

3.4. Performance of Different Energy Expressions. The performance of the force field does not significantly depend on the energy expression, as shown by benchmarks for structural properties, surface properties, and mechanical properties (Table 5). Structural properties exhibit only minor variations; only the force fields AMBER and OPLS exhibit a slightly wider distribution of the Ca^{2+} ions within the channel-like features (Figure 1). 12-6 LJ parameters are slightly more repulsive than 9-6 LJ parameters so that cleavage energies are 5-10% higher and the bulk modulus is approximately 107 GPa instead of 98 GPa, even closer to experiment (Table 4). These differences are small and reflect performance limits due to the mathematical form of the energy expressions, including differences in the type of Lennard-Jones potential, combination rules, and scaling of nonbond interactions between 1,4 bonded atoms.

Table 5. Comparison of key properties of tricalcium aluminate in experiment and molecular dynamics simulation with different energy expressions.^a

Property	Expt	PCFF	CHARMM	CVFF	AMBER	OPLS-AA
Lattice parameter <i>a</i> (nm)	1.5263 ^b	1.5265	1.5280	1.5240	1.5267	1.5258
Al□O ^{ring} bond length (pm)	175.5 ^b	174.5	174.0	173.5	176.0	175.5
Al□O ^{apical} bond length (pm)	175.0 ^b	174.5	172.5	172.0	174.5	174.5
Cleavage energy (mJ/m ²)	1250 ±150 (±12%) ^c	1260 (±0%)	1325 (+5.2%)	1321 (+4.1%)	1341 (+6.4%)	1380 (+9.5%)
Bulk modulus (GPa) at 0-1 GPa	106 ±8 ^d	98 ±3	107 ±3	108 ±3	107 ±3	108 ±3

^a In addition to the listed properties, computed bond angles and computed vibration spectra are ±1% the same for all energy expressions. ^b Ref. ⁴⁸. ^c Ref. ⁵². ^d Ref. ²⁶.

4. Cleavage, Hydration, and Agglomeration of Tricalcium Aluminate Surfaces

Particle cleavage plays a major role in processing of minerals, rocks, ores, as well as composite materials.^{7, 8, 68-71} The associated energies of cleavage and agglomeration characterize the individual surfaces and can be challenging to measure for specific (h k l) facets, especially when accompanied by chemical reactions.^{52, 68, 72} Molecular models can

monitor such processes in atomic resolution to provide further validation, gain mechanistic insight, and elucidate experimentally unavailable data.

4.1. Cleavage Energy. Many possible cleavage planes of tricalcium aluminate, such as (100), (010), (001), (200), (020), and (002) are identical due to cubic crystal symmetry (Figure 6). These cleavage planes are of lower energy than others, such as (111), due to the orientation of calcium ions and aluminate rings in “channels” along the coordinate axes. The cleavage energy of (100) planes upon equilibrium cleavage at 25 °C was computed as 1260 ± 50 mJ/m² (Table 6). An increase in temperature to 90 °C, typically employed for particle grinding, lowered the cohesive energy by approximately 5% to 1210 ± 50 mJ/m². Experimentally determined cleavage energies for Ca₃Al₂O₆ are not available, however, the computational results are close to reported measurements of 1100 to 1400 mJ/m² for similar minerals such as Ca(OH)₂ and CaO that are of comparable polarity.⁵²

It was thereby assumed that the two model surfaces possess equal amounts of Ca²⁺ cations in a low-energy distribution after cleavage, identified by temperature-gradient protocols in the simulation (Figure 6).^{8, 68, 72} In this manner, each surface remains electrically neutral and exhibits minimal local electric fields. Higher, distance-dependent cleavage energies were also observed by rapid, non-equilibrium cleavage.

Table 6. Computed equilibrium cleavage energy and agglomeration energy of C₃A. The agglomeration energy depends on temperature, initial hydration (H0, H1) and surface displacement after cleavage (D0, D1).^a The (100) cleavage plane is of lowest energy and equivalent to several other planes (see text).

Cleavage plane	Temperature (K)	Cleavage energy (mJ/m ²)
(1 0 0)	298	1260 ± 50
(1 0 0)	363	1210 ± 50
Expt: Ca(OH) ₂ , CaO	298	1180 ± 100, 1310 ± 200 ^b
Agglomeration plane ^a	Temperature (K)	Agglomeration energy (mJ/m ²)
(1 0 0), H0, D0	298	860 ± 40
(1 0 0), H0, D0	363	720 ± 45
(1 0 0), H0, D1	363	300 ± 50
(1 0 0), H1, D0	298	520 ± 30
(1 0 0), H1, D0	363	500 ± 40
(1 0 0), H1, D1	363	250 ± 30

^a “H0” indicates no hydration and “H1” initial hydration after cleavage. “D0” indicates no surface displacement and “D1” a specified surface displacement after cleavage (Figure 9). Data for another displacement D1’ range between D0 and D1 (Figure 9).

^b Ref. ⁵².

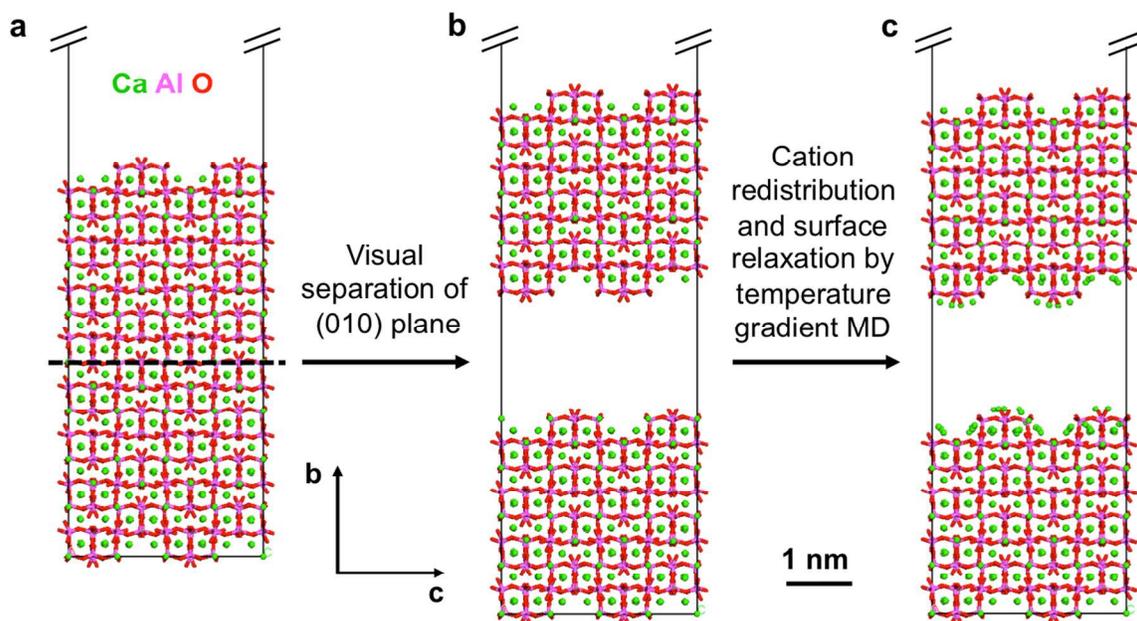
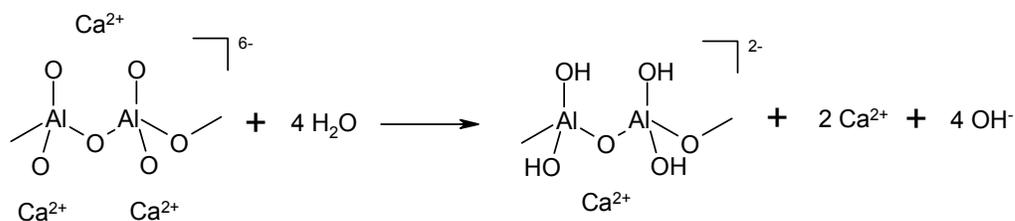


Figure 6. Generation of model surfaces of lowest energy upon cleavage for the example of the (010) cleavage plane. (a) An electroneutral slab of C₃A. (b) Initial separation of (010) surfaces by at least 2 nm with an arbitrary distribution of surface cations. (c) Equilibrium surfaces with electroneutral distribution of cations after temperature gradient molecular dynamics (ref. ⁷²).

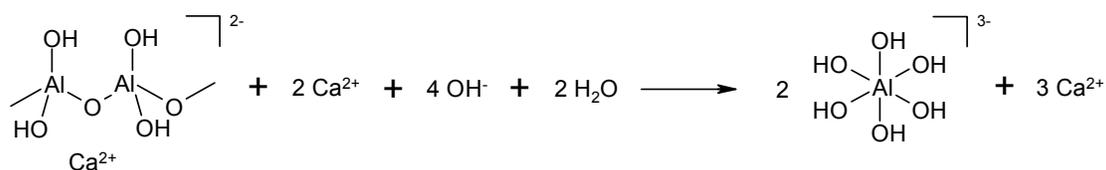
4.2. Initial Hydration of the C₃A Surface. Water reacts instantly with tricalcium aluminate.³⁷ Hydration likely begins already during grinding of cement clinker at 90 °C due to residual humidity and dehydration of gypsum additives. Initial contact with water causes protonation of single bonded oxygen atoms on aluminate rings and leads to the formation of calcium hydroxide (Scheme 1):



Scheme 1. Initial hydration of an aluminate ring fragment.

The aim in this contribution is modeling this very first step of the hydration reaction, which drastically changes the surface chemistry and morphology (Figure 7). For simplicity, we assume that this process initially affects the first molecular layer as well as the top portion of the second molecular layer.

The next steps in the sequence of the hydration reaction involve the generated hydroxide ions as strong nucleophiles, which fully hydrolyze the ring structure in combination with water (Scheme 2):



Scheme 2. Subsequent hydration of an aluminate ring fragment to octahedral aluminates such as hydrocalumite.

The coordination number of aluminum then increases from four to six and individual $\text{Al}(\text{OH})_6$ ions eventually organize into layered double hydroxides, such as hydrocalumite, Friedel's salt, monosulphate and ettringite depending on the present ions.³⁷ Force field parameters for the simulation of these phases are available in the INTERFACE force

field so that also further steps of the reaction can be investigated.¹⁰ In the present work, however, we focus on the initial hydration stage since very little information is available about the transformation at the molecular scale for each step. Full hydration in water is thus not the subject of this contribution because a cascade of reactions needs to be evaluated.

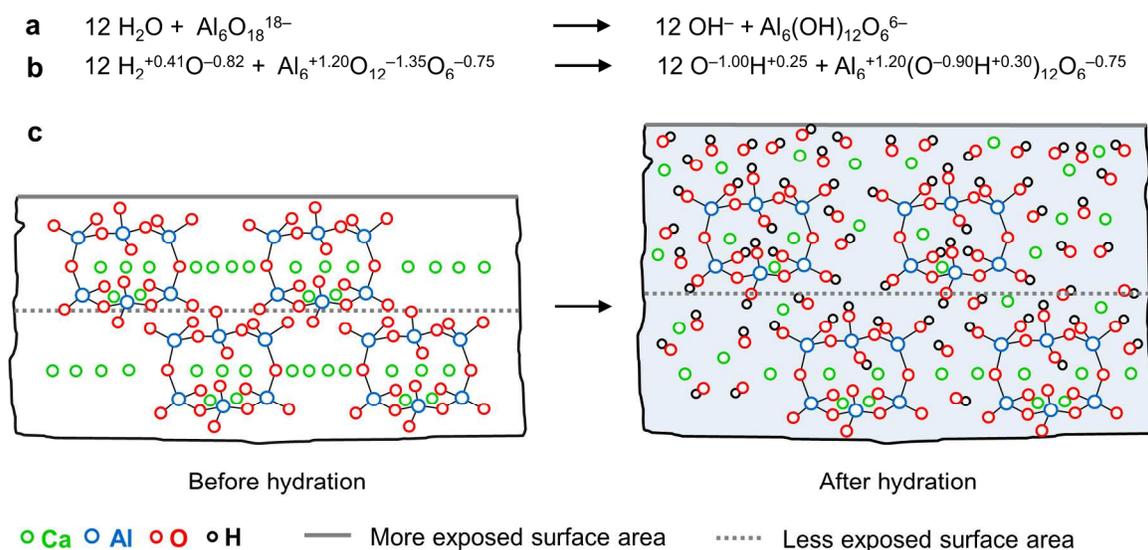


Figure 7. Details of the initial hydration reaction of tricalcium aluminate. (a) Net reaction using formal charges. (b) Net reaction showing atomic charges that represent the balance of covalent versus ionic bonding truthfully for simulations (refs.^{10, 20}). (c) Schematic of structural changes upon protonation of the first layer of aluminate rings and of the top portion of the second layer. The hydration reaction causes swelling and deposition of an amorphous layer of $\text{Ca}(\text{OH})_2$ (portlandite) on the aluminate surface.

4.3. Agglomeration of Pure and Initially Hydrated C_3A Surfaces. Agglomeration is the reverse process of cleavage (Figures 8 and 9). The distance dependence of the

agglomeration energy shows that over 90% of recoverable attractive energy is felt at distances closer than 1 nm, consistent with detailed analyses of similar ionic surfaces of clay minerals and tricalcium silicate.^{8, 68, 72, 73} The presented analysis thus focuses on the total energy difference between closest contact and large separation of 20 nm (Figure S2).

Agglomeration energies of pure and hydrated C₃A surfaces are computed significantly lower than cleavage energies (Table 6). The maximum recoverable energy upon agglomeration is 860 mJ/m², i.e., about 70% of the original cleavage energy. The agglomeration energy also decreases slightly at higher temperature similar to the cleavage energy (Table 6). The decrease amounts to 6±4% upon change in temperature from 25 °C to 90 °C.

Another major reduction in cohesion between the cleaved surfaces occurs upon initial hydration (Table 6). Hydrated surfaces are of lower polarity, e.g., due to AlOH groups instead of semi-ionic Al–O bonds, and Coulomb interactions are thus weaker (Figure 7). In addition, the hydrated surfaces are covered with an amorphous layer of calcium hydroxide that widens the gap between cleavage planes and mitigates attraction (Figure 8). After initial hydration, the remaining energy of agglomeration is therefore approximately 500 mJ/m², i.e., about 40% of the initial cleavage energy.

The fit between cleaved surfaces can further be diminished by lateral displacement (Figure 9, Table 6). The inherent surface roughness of tricalcium aluminate, represented by step-like features every 0.8 nm, causes a substantial reduction in cohesion as a consequence of lateral offsets of surfaces upon agglomeration, in contrast to other minerals with less surface corrugation. It is thus difficult that two surface after cleavage would match in surface corrugation. These effects are illustrated for three representative

cases (Figure 9a-c). An unfavorable offset alone can reduce the maximum agglomeration energy of 860 mJ/m^2 to 300 mJ/m^2 (Table 6). The combination of hydration and lateral displacement of surfaces after cleavage leads to a maximum reduction in agglomeration energy to $\sim 250 \text{ mJ/m}^2$ at 90°C , which corresponds to $\sim 20\%$ of the original cleavage energy. Especially for real particle surfaces with imperfect cleavage planes and changing angular orientation of the surfaces after cleavage, the effects of surface displacement onto agglomeration are likely significant.

The information on nanoscale agglomeration is newly generated by the force field and experimental measurements are currently not available for comparison. The data are expected to be about $\pm 10\%$ accurate similar to validated cleavage energies (section 4.2), and supported by predictions for similar minerals.^{8, 10, 68, 72, 73}

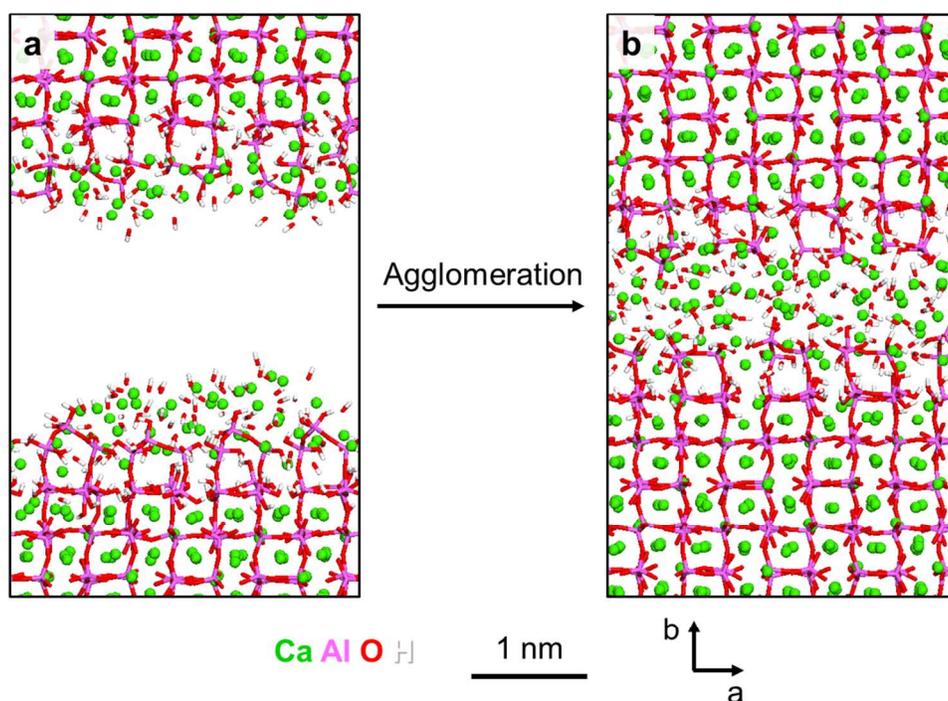


Figure 8. Agglomeration of initially hydrated surfaces of tricalcium aluminate in air. (a) Cleaved (100) surfaces upon initial hydration in side view. (b) Agglomerated (100) surfaces upon initial hydration. An amorphous interfacial layer of calcium hydroxide can be seen.

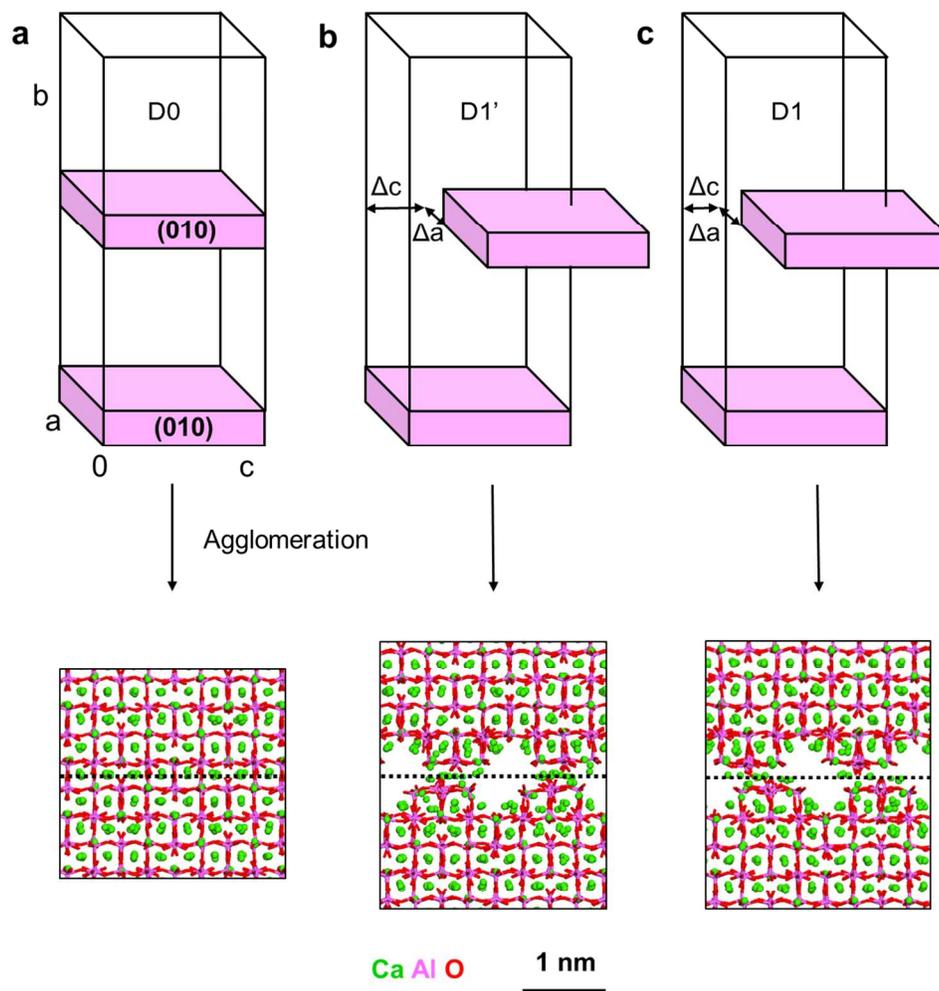
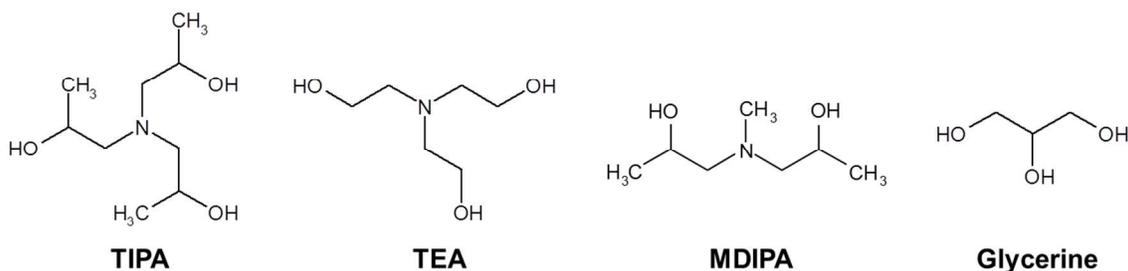


Figure 9. Effect of surface displacement on contact area and agglomeration energy. (a) The highest agglomeration energy is found in the absence of lateral offsets (D0). (b) Example of a horizontal displacement $\Delta a = -1.0$ nm, $\Delta c = 1.0$ nm of the two surfaces, which reduces agglomeration (D1'). (c) Example of another displacement $\Delta a = -1.0$ nm,

$\Delta c = 0.5$ nm, leading to even lower agglomeration (D1). Lower gravimetric density related to lower area density of calcium ions in the contact zone decreases agglomeration (see dotted black lines).

5. Influence of Adsorbed Organic Thin Films

5.1. Mechanism. Agglomeration of the mineral surfaces in air undergoes drastic further changes in the presence of adsorbed organic molecules. Amino alcohols such as triisopropanol amine (TIPA), triethanol amine (TEA), MDIPA (N-Methyl-diisopropanolamine), and glycerine are often added during particle grinding (Scheme 3).^{2, 3, 8, 33-35} A typical surface coverage of ~65% of a molecular monolayer, or 0.20 mg/m², of these compounds was assumed on individual cleaved surfaces (details in Table S1).⁸ The interfacial region between two surfaces in contact with each other therefore contains an amount between a molecular monolayer and a molecular bilayer (130%, Figure 10).



Scheme 3. Structure of triisopropanolamine (TIPA), triethanolamine (TEA), N-methyl-diisopropanolamine (MDIPA), and glycerine used as additives for grinding in cement.

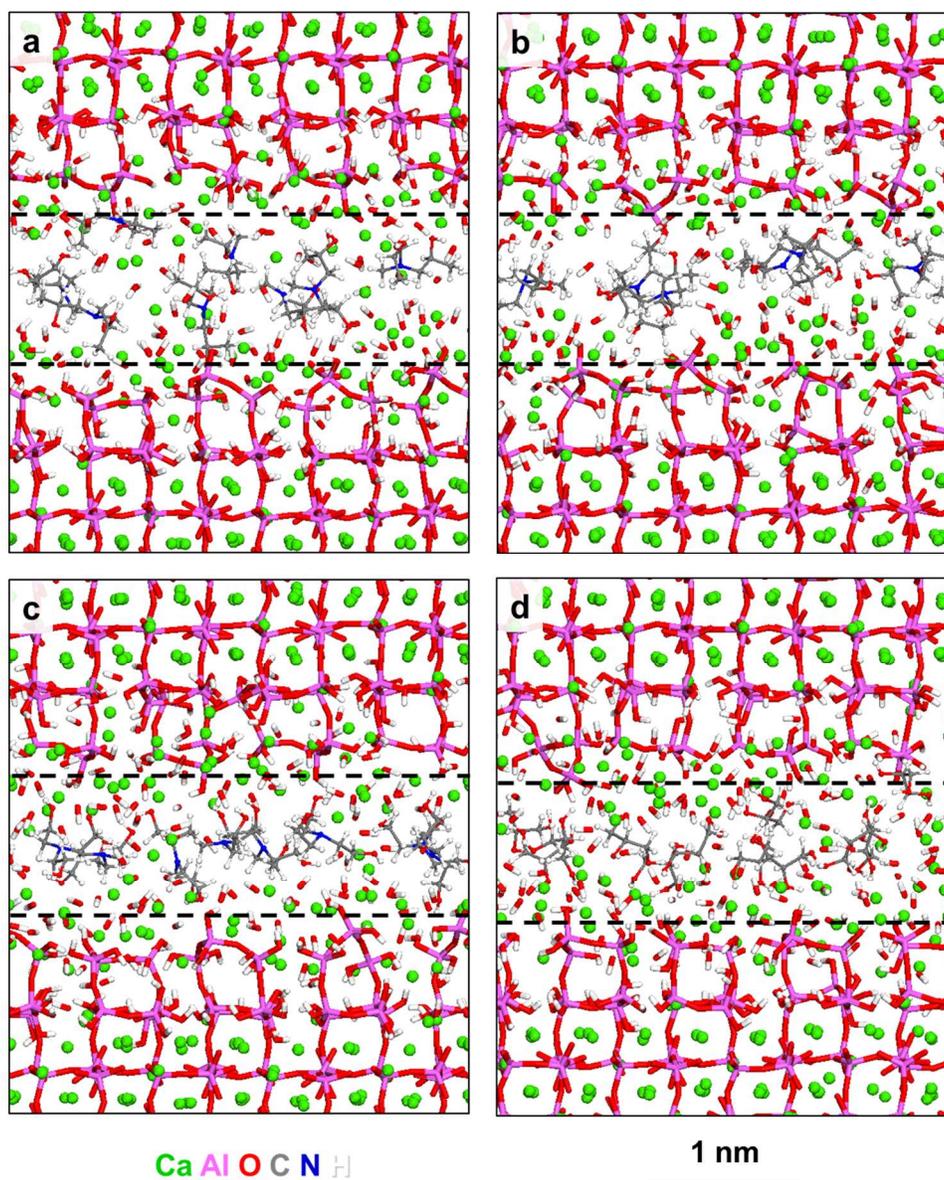


Figure 10. Conformation of alcohol molecules confined between initially hydrated surfaces of tricalcium aluminate according to molecular dynamics simulation. (a) MDIPA, (b) TIPA, (c) TEA, and (d) glycerine. The organic layer acts as a spacer to keep the surfaces apart and mitigates attractive polar interactions.

The alcohol molecules are attracted to the mineral surfaces through chelation of calcium ions by hydroxyl groups and through the formation of hydrogen bonds with

hydroxide ions and aluminol groups of the hydrated tricalcium aluminate. The organic layer reduces Coulomb interactions between the cleaved surfaces and thereby their cohesion (Figure 11). It has previously been shown that layered minerals with interlayer cations require approximately 0.5 nm layer separation to distribute cations between the two cleaved layers and eliminate over 90% of the attractive Coulomb energy.^{8, 68, 72, 73} The polar organic molecules are attracted to the surfaces and assume essentially the role of a molecular spacer of similar thickness. The reduction in agglomeration energy varies depending on the chemistry of the alcohol. Assuming no surface displacement upon cleavage, the agglomeration energy of initially hydrated surfaces is 500 mJ/m². Most alcohols reduce the agglomeration energy to values between 100 and 60 mJ/m² (Figure 11). The largest reduction was observed for N-methyl-diisopropanolamine (MDIPA) with an agglomeration energy of only ~30 mJ/m². Minimal agglomeration in case of MDIPA may be associated with a larger proportion of hydrophobic groups versus hydrophilic groups (Scheme 3) that help separate ionic groups on the two surfaces better.

Once organic molecules are present, also the effect of surface displacements on the agglomeration energy (Figure 9) becomes much less significant. The reduction of surface forces between the cleaved mineral surfaces by organic inclusions reduces energy barriers of lateral surface displacement and enables higher interfacial mobility.

5.2. Comparison to Other Surfaces. The role of similar organic modifiers has also been previously investigated on other substrates, for example, tricalcium silicate,⁸ clay minerals,^{68, 72} metal nanostructures,^{74, 75} and various oxides.^{76, 77} On tricalcium aluminate, the molecules are loosely grafted without specific chemical bonds, ionic bonds, and regular coordination patterns. The situation is similar on tricalcium silicate, which is the

major phase with >50% content in Portland cement (Figure 12). On common cleavage planes of both tricalcium aluminate (C_3A) and tricalcium silicate (C_3S), the molecules show qualitatively the same trend in reduction of the agglomeration energy at the same surface coverage, except for TIPA as an outlier.

The similar behavior on these surfaces leads to practical implications. The computed ranking of reduction in agglomeration energy on tricalcium silicate was shown to correlate with the performance in full-scale trials at cement plants.⁸ The force field and results for tricalcium aluminate are of the same quality. Since tricalcium aluminate (C_3A) and tricalcium silicate (C_3S) exhibit a similar trend, modeling the behavior of molecular additives on only one surface might suffice to predict the anticipated performance on both surfaces, offering a saving in time (Figure 12). However, possibly meaningful outliers such as TIPA also need consideration. It is known that TIPA has a major effect on long-term strength development of cement, and possible relations to the unexpected behavior will be examined in follow-on studies.

5.3. Applications of the Model. This work focuses on establishing credibility of the force field and on first new insights into hydration and surface properties of tricalcium aluminate. The models for tricalcium aluminate become meaningful in the context of the entire set of cement minerals and through full compatibility,¹⁰ and are thus applicable to impure phases, mixtures of minerals, hydrated systems at any stage, and polymer interfaces. Nanostructures up to 100 nm size upon grinding and hydration of cement, far beyond single nanocrystals, can be reliably explored by simulations in comparison to measurements by X-Ray scattering, IR, NMR, AFM, kinetic measurements, and mechanical characterization. Such efforts can accurately identify structural patterns,

associated energies, and correlations with observable properties. Molecular modeling and simulation is chiefly a visualization tool that allows the quantitative prediction of associated thermodynamic properties to guide in control over the assembly of molecules to nanocrystals and glassy structures, helping to unravel the mystery of cement materials.

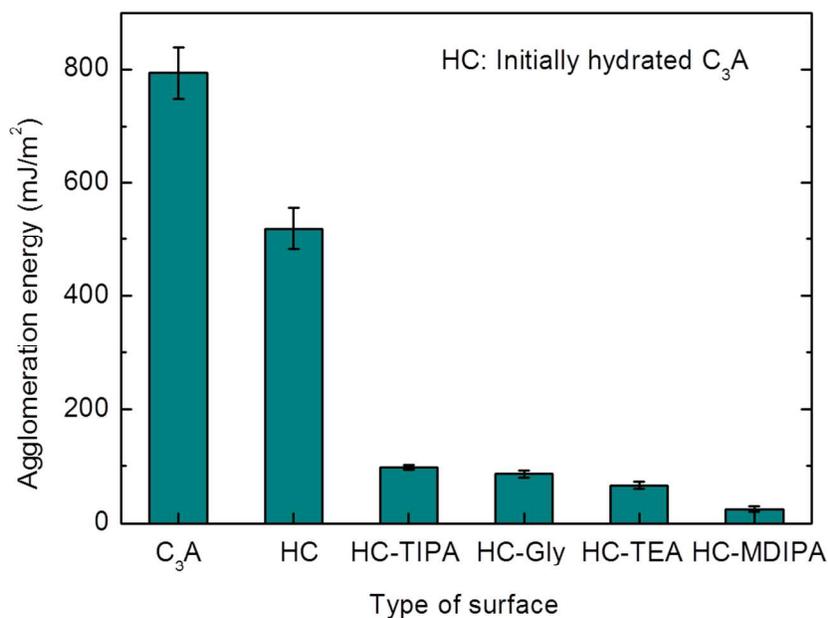


Figure 11. Computed agglomeration energy of tricalcium aluminate modified with surface-adsorbed alcohols at ~130% surface coverage at 363 K. A decrease up to 95% is found in comparison to the initially hydrated surface (HC).

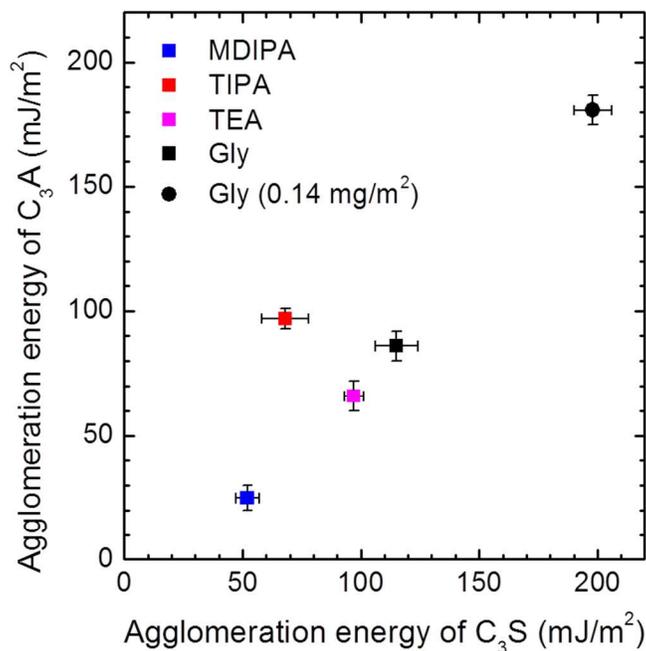


Figure 12. Comparison of the agglomeration energy in the presence of grinding aids for initially hydrated C₃A and C₃S surfaces (data for C₃S from ref. ⁸). With the exception of TIPA, the action of the additives is similar with a trend towards lower agglomeration energy for C₃A.

6. Conclusion

This contribution introduces a force field for tricalcium aluminate (C₃A) and provides understanding of the initial hydration reaction, agglomeration of surfaces, and interactions with organic molecules. The force field parameters for the mineral and the partially hydrated surfaces can be applied to the study of structural, interfacial, and mechanical properties of cement at the scale of 1 to 100 nm, and facilitate access to a range of problems that cannot be addressed by computationally expensive DFT methods. The force field reproduces atomic coordinates, elastic moduli, thermal expansion, as well as cleavage energies in excellent agreement with experimental data. The parameters are

embedded in the force fields PCFF, CVFF, CHARMM, AMBER, OPLS-AA, as well as INTERFACE and can be applied to a wide range of aqueous, inorganic, organic, and biological interfaces in building materials and dental root filling cements. The simulation of impure solid solutions of C_3A and advanced hydration products is feasible using parameters for related compounds in the INTERFACE force field, and extensions to chemical reactions are possible (see section S1.7).

The common (001) cleavage plane in the highly symmetric C_3A crystal possesses a cleavage energy of $1260 \pm 50 \text{ mJ/m}^2$, of which about two thirds are recoverable upon agglomeration. Lateral displacement of the corrugated surfaces and initial hydration after cleavage substantially lower the agglomeration energy, reaching values as low as 250 mJ/m^2 . Quantitative atomistic models of the initial hydration process were introduced, including the formation of an amorphous layer of calcium hydroxide on the mineral surface and emerging disordered morphologies. The modified interactions between the mineral surfaces significantly affect the agglomeration energy

Adsorbed organic molecules such as alcohols and amines coordinate superficial calcium ions and form hydrogen bonds with hydroxide and aluminol groups on the hydrated surfaces. The presence of these molecules further reduces agglomeration energies. At a thickness of a partial molecular monolayer per surface (surface coverage of 65%), corresponding to typical conditions during cement grinding in a ball mill, the agglomeration energy could be remarkably reduced to 30 mJ/m^2 , i.e., more than an order of magnitude compared to an initially hydrated surface. Thereby, N-methyldiisopropanolamine (MDIPA) showed a stronger reduction of surface forces than glycerol and other amines at the same surface coverage. The relative order of surface

forces on tricalcium aluminate is similar to tricalcium silicate and suggests the possibility of predictions using a single reference surface. The force field and initial results are a step forward to understand cement hydration and the role of adsorbed organic molecules at the 1 to 100 nm scale. In follow-on studies, simulations may effectively monitor complex mineral phases and hydration mechanisms in the presence of organic and polymeric additives to improve specific cement properties in comparison with experiment.

Acknowledgements

The authors acknowledge support by Sika Technology AG, ETH Zurich, the National Science Foundation (DMR 0955071), the University of Akron, and the Swiss Commission for Technological Innovation (KTI 13703.1 PFFLR-IW). The allocation of computational resources at the Ohio Supercomputing Center is gratefully acknowledged.

Electronic supplementary information (ESI) available: Additional figures, tables, full computational details, limitations, and opportunities.

References

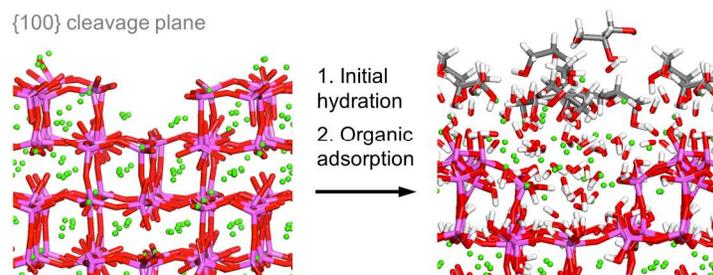
1. R. J. Flatt, N. Roussel and C. R. Cheeseman, *J. Eur. Ceram. Soc.*, 2012, 32, 2787-2798.
2. M. Katsioti, P. E. Tsakiridis, P. Giannatos, Z. Tsibouki and J. Marinos, *Constr. Build. Mater.*, 2009, 23, 1954-1959.
3. A. Atmaca and M. Kanoglu, *Energy*, 2012, 42, 261-269.
4. K. L. Scrivener, *Chem. World*, 2008, 5, 62-66.
5. K. L. Scrivener and A. Nonat, *Cem. Concr. Res.*, 2011, 41, 651-665.
6. D. W. Fuerstenau, K. S. Venkataraman and B. V. Velamakanni, *Int. J. Miner. Process.*, 1985, 15, 251-267.
7. D. W. Fuerstenau and A. Z. M. Abouzeid, *Int. J. Miner. Proc.*, 2002, 67, 161-185.
8. R. K. Mishra, R. J. Flatt and H. Heinz, *J. Phys. Chem. C*, 2013, 117, 10417-10432.

9. M. Schneider, M. Romer, M. Tschudin and H. Bolio, *Cem. Concr. Res.*, 2011, 41, 642-650.
10. H. Heinz, T.-J. Lin, R. K. Mishra and F. S. Emami, *Langmuir*, 2013, 29, 1754-1765.
11. P. M. Dove, N. Z. Han and J. J. De Yoreo, *Proc. Natl. Acad. Sci. U. S. A.*, 2005, 102, 15357-15362.
12. P. J. McDonald, J. P. Korb, J. Mitchell and L. Monteilhet, *Phys. Rev. E*, 2005, 72, 011409.
13. A. J. Allen, J. J. Thomas and H. M. Jennings, *Nature Mater.*, 2007, 6, 311-316.
14. I. G. Richardson, *Cem. Concr. Res.*, 2008, 38, 137-158.
15. R. J. Flatt, I. Schober, E. Raphael, C. Plassard and E. Lesniewska, *Langmuir*, 2009, 25, 845-855.
16. L. B. Skinner, S. R. Chae, C. J. Benmore, H. R. Wenk and P. J. M. Monteiro, *Phys. Rev. Lett.*, 2010, 104, 195502.
17. J. W. Bullard, H. M. Jennings, R. A. Livingston, A. Nonat, G. W. Scherer, J. S. Schweitzer, K. L. Scrivener and J. J. Thomas, *Cem. Concr. Res.*, 2011, 41, 1208-1223.
18. B. J. Smith, A. Rawal, G. P. Funkhouser, L. R. Roberts, V. Gupta, J. N. Israelachvili and B. F. Chmelka, *Proc. Natl. Acad. Sci. U. S. A.*, 2011, 108, 8949-8954.
19. A. G. Kalinichev and R. J. Kirkpatrick, *Chem. Mater.*, 2002, 14, 3539-3549.
20. H. Heinz and U. W. Suter, *J. Phys. Chem. B*, 2004, 108, 18341-18352.
21. R. J. Pellenq, A. Kushima, R. Shahsavari, K. J. Van Vliet, M. J. Buehler, S. Yip and F. J. Ulm, *Proc. Natl. Acad. Sci. U. S. A.*, 2009, 106, 16102-16107.
22. H. Manzano, J. S. Dolado and A. Ayuelaz, *J. Am. Ceram. Soc.*, 2009, 92, 897-902.
23. J. S. Dolado, M. Griebel, J. Hamaekers and F. Heber, *J. Mater. Chem.*, 2011, 21, 4445-4449.
24. S. Galmarini, A. Aimable, N. Ruffray and P. Bowen, *Cem. Concr. Res.*, 2011, 41, 1330-1338.
25. H. Manzano, R. J. M. Pellenq, F.-J. Ulm, M. J. Buehler and A. C. T. van Duin, *Langmuir*, 2012, 28, 4187-4197.
26. J. Moon, S. Yoon, R. M. Wentzcovitch, S. M. Clark and P. J. M. Monteiro, *J. Am. Ceram. Soc.*, 2012, 95, 2972-2978.
27. E. Masoero, E. Del Gado, R. J. M. Pellenq, F. J. Ulm and S. Yip, *Phys. Rev. Lett.*, 2012, 109, 155503.
28. E. Durgun, H. Manzano, R. J. M. Pellenq and J. C. Grossman, *Chem. Mater.*, 2012, 24, 1262-1267.
29. H. Heinz, R. A. Vaia, H. Koerner and B. L. Farmer, *Chem. Mater.*, 2008, 20, 6444-6456.
30. S. V. Patwardhan, F. S. Emami, R. J. Berry, S. E. Jones, R. R. Naik, O. Deschaume, H. Heinz and C. C. Perry, *J. Am. Chem. Soc.*, 2012, 134, 6244-6256.
31. N. Fernandez, R. Mani, D. Rinaldi, D. Kadau, M. Mosquet, H. Lombois-Burger, J. Cayer-Barrioz, H. J. Herrmann, N. D. Spencer and L. Isa, *Phys. Rev. Lett.*, 2013, 111, 108301.
32. D. P. Bentz, *J. Am. Ceram. Soc.*, 1997, 80, 3-21.
33. F. G. Serafin, *Amide Grinding Aid*, U.S. Patent No. 3,459,570, August 5, 1969.

34. S. Sohoni, R. Sridhar and G. Mandal, *Powder Technol.*, 1991, 67, 277-286.
35. B. Marazzani, C. Kurz, C. Burge and T. Muller, *Dialkanolamine als Additive zum Mahlen von Feststoffen* European Patent No. 2,527,307 A1, November 28, 2012.
36. M. Torabinejad and N. Chivian, *J. Endodontics*, 1999, 25, 197-205.
37. H. F. W. Taylor, *Cement Chemistry*, Academic Press, London, 1997.
38. L. Black, C. Breen, J. Yarwood, C. S. Deng, J. Phipps and G. Maitland, *J. Mater. Chem.*, 2006, 16, 1263-1272.
39. H. Minard, S. Garrault, L. Regnaud and A. Nonat, *Cem. Concr. Res.*, 2007, 37, 1418-1426.
40. S. H. Oh, R. Finones, S. Jin, S. Y. Choi and K. N. Kim, *J. Mater. Res.*, 2004, 19, 1062-1067.
41. W. Liu, W. Peng, Y. Zhu and J. Chang, *J. Biomed. Mater. Res. B-Appl. Biomater.*, 2012, 100B, 1257-1263.
42. F. C. Lee, H. M. Banda and F. P. Glasser, *Cem. Concr. Res.*, 1982, 12, 237-246.
43. K. Fukuda, S. Inoue and H. Yoshida, *J. Am. Ceram. Soc.*, 2003, 86, 112-114.
44. S. L. Meyers, *J. Ind. Eng. Chem.*, 1940, 32, 1107-1112.
45. R. A. Schroeder and L. L. Lyons, *J. Inorg. Nucl. Chem.*, 1966, 28, 1155-1163.
46. P. Tarte, *Spectrochim. Acta a-Mol. Spectr.*, 1967, A 23, 2127-2143.
47. S. N. Ghosh and A. K. Chatterjee, *J. Mater. Sci.*, 1974, 9, 1577-1584.
48. P. Mondal and J. W. Jeffery, *Acta Cryst. B-Struct. Sci.*, 1975, 31, 689-697.
49. T. L. Hughes, C. M. Methven, T. G. J. Jones, S. E. Pelham, P. Fletcher and C. Hall, *Adv. Cem. Based Mater.*, 1995, 2, 91-104.
50. K. Velez, S. Maximilien, D. Damidot, G. Fantozzi and F. Sorrentino, *Cem. Concr. Res.*, 2001, 31, 555-561.
51. L. Fernández-Carrasco, D. Torrens-Martín, M. L. M. and M.-R. S., in *Infrared Spectroscopy - Materials Science, Engineering and Technology*, ed. T. Theophanides, InTech Press, <http://www.intechopen.com>, 2012, ch. 19.
52. S. Brunauer, D. L. Kantro and C. H. Weise, *Can. J. Chem.*, 1956, 34, 729-742.
53. H. Heinz, H. J. Castelijns and U. W. Suter, *J. Am. Chem. Soc.*, 2003, 125, 9500-9510.
54. H. Heinz, H. Koerner, K. L. Anderson, R. A. Vaia and B. L. Farmer, *Chem. Mater.*, 2005, 17, 5658-5669.
55. H. Heinz, R. A. Vaia, B. L. Farmer and R. R. Naik, *J. Phys. Chem. C*, 2008, 112, 17281-17290.
56. H. Sun, *J. Comput. Chem.*, 1994, 15, 752-768.
57. H. Sun, S. J. Mumby, J. R. Maple and A. T. Hagler, *J. Am. Chem. Soc.*, 1994, 116, 2978-2987.
58. H. Sun, *Macromolecules*, 1995, 28, 701-712.
59. H. Sun, *J. Phys. Chem. B*, 1998, 102, 7338-7364.
60. A. D. MacKerell, D. Bashford, M. Bellott, R. L. Dunbrack, J. D. Evanseck, M. J. Field, S. Fischer, J. Gao, H. Guo, S. Ha, D. Joseph-McCarthy, L. Kuchnir, K. Kuczera, F. T. K. Lau, C. Mattos, S. Michnick, T. Ngo, D. T. Nguyen, B. Prodhom, W. E. Reiher, B. Roux, M. Schlenkrich, J. C. Smith, R. Stote, J. Straub, M. Watanabe, J. Wiorkiewicz-Kuczera, D. Yin and M. Karplus, *J. Phys. Chem. B*, 1998, 102, 3586-3616.

61. D. A. Pearlman, D. A. Case, J. W. Caldwell, W. S. Ross, T. E. Cheatham, S. Debolt, D. Ferguson, G. Seibel and P. Kollman, *Comput. Phys. Comm.*, 1995, 91, 1-41.
62. W. L. Jorgensen, D. S. Maxwell and J. TiradoRives, *J. Am. Chem. Soc.*, 1996, 118, 11225-11236.
63. P. Dauber-Osguthorpe, V. A. Roberts, D. J. Osguthorpe, J. Wolff, M. Genest and A. T. Hagler, *Proteins: Struct., Funct., Genet.*, 1988, 4, 31-47.
64. J. M. Wang, R. M. Wolf, J. W. Caldwell, P. A. Kollman and D. A. Case, *J. Comput. Chem.*, 2004, 25, 1157-1174.
65. S. S. Batsanov, *Inorg. Mater.*, 2001, 37, 871-885.
66. H. Heinz, R. A. Vaia, B. L. Farmer and R. R. Naik, *J. Phys. Chem. C*, 2008, 112, 17281-17290.
67. V. L. Burdick and D. E. Day, *J. Am. Ceram. Soc.*, 1967, 50, 97-101.
68. H. Heinz, R. A. Vaia and B. L. Farmer, *J. Chem. Phys.*, 2006, 124, 224713.
69. G. D. Zartman, H. Liu, B. Akdim, R. Pachter and H. Heinz, *J. Phys. Chem. C*, 2010, 114, 1763-1772.
70. Y. T. Fu, G. D. Zartman, M. Yoonessi, L. F. Drummy and H. Heinz, *J. Phys. Chem. C*, 2011, 115, 22292-22300.
71. C. E. Powell and G. W. Beall, *Curr. Opinion Solid State Mater. Sci.*, 2006, 10, 73-80.
72. Y. T. Fu and H. Heinz, *Chem. Mater.*, 2010, 22, 1595-1605.
73. Y. T. Fu and H. Heinz, *Philos. Mag.*, 2010, 90, 2415-2424.
74. J. Feng, J. M. Slocik, M. Sarikaya, R. R. Naik, B. L. Farmer and H. Heinz, *Small*, 2012, 8, 1049-1059.
75. K. C. Jha, H. Liu, M. R. Bockstaller and H. Heinz, *J. Phys. Chem. C*, 2013, 117, 25969-25981.
76. L. H. Dubois and R. G. Nuzzo, *Ann. Rev. Phys. Chem.*, 1992, 43, 437-463.
77. H. Heinz, R. A. Vaia and B. L. Farmer, *Langmuir*, 2008, 24, 3727-3733.

Table of Contents Entry



A force field for tricalcium aluminate is presented to generate insight into the hydration reaction, surface forces, and adsorbed molecules.