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2024, 12, 7605Chemically driven dimensionality modulation of
hybrid tin(II) halide perovskite microcrystals†Raúl I. Sánchez-Alarcón, ^a Omar E. Solís,^a María Cristina Momblona-Rincon, ^a
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The lattice arrangement of lead-free perovskites can be altered by synthesis parameters such as concentration, temperature, choice of solvents, and the length of monovalent cations. This effect, often described as dimensionality modulation, can be exploited to develop new materials with better optical properties than their three-dimensional counterpart. This work explores the versatility of the hot-injection synthesis for obtaining 2-thiopheneethylammonium (TEA⁺) based tin halide perovskite microcrystals. The dimensionality of TEA-based tin bromide perovskite microcrystals can be modulated through changes in the TEA concentration, giving rise to the formation of highly luminescent 0D-TEA₄SnBr₆ and low-emissive 2D-TEA₂SnBr₄ microplates. In contrast, 0D-[TEASnCl₃][TEACl] and 2D-TEA₂SnI₄ are thermodynamically preferred as main products when using the chloride and iodide analogues, respectively, limiting the dimensional versatility. It is noteworthy that 0D-TEA₄SnBr₆ shows the highest PLQY from the low dimensional TEA tin halide family (PLQY = 75% in thin-film) owing to a more confined configuration based on isolated [SnBr₆]⁴⁻ octahedra moieties separated by TEA⁺ cations, being attractive for optoelectronic applications.

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Introduction

Lead halide perovskites show promising properties for semiconductor-based applications, such as efficient energy conversion in solar cells.¹ However, one of their major drawbacks is the toxicity of lead, a key component usually present in most efficient compositions. Researchers are actively exploring alternative materials to replace lead in perovskites, aiming to mitigate environmental and health concerns. This has become a key area of research with the aim of developing environmentally friendly and sustainable technologies. The most common approach to form lead-free perovskites is to replace Pb(II) with Sn(II). Since both of them present the same valence electron configuration (ns²sp²), the resulting perovskites can offer similar properties, maintaining the perovskite structure and optoelectronic behaviour, such as high charge carrier mobility, high light absorption properties, and small exciton binding energy.² However, tin-based perovskites exhibit lower stability than their

lead-containing counterparts.³ Even though Sn and Pb present multiple oxidation states, tin forms more commonly Sn(II) and Sn(IV) ions, while lead is often found as Pb(II). The reason for this is attributed to the lanthanide contraction effect in Pb by the presence of 4f¹⁴ electrons in its electronic configuration. Conversely, Sn lacks lanthanide shrinkage, leading to easy loss of all valence electrons, and consequently, forming Sn⁴⁺ ions.⁴

Reducing the dimensionality of perovskite materials by disaggregation of inorganic moieties with bulky organic cations is a promising approach to mitigate the Sn-based perovskite stability issues.⁵ For example, in two dimensional perovskites, 2D, layered perovskite moieties are hydrogen bonded with voluminous organic spacers, which are self-assembled by van der Waals forces and π -stacking. If layers of 3D perovskites are stuck within the voluminous organic spacers, the physical properties depend on the thickness of inorganic perovskite moieties, being possible to tune the exciton binding energy and band gap of the resulting materials.⁶ On the other hand, if the density of bulky organic cations is high enough to separate the structure on individual polyhedron moieties, these conduce to the formation of zero-dimensional perovskite structures, 0D.⁷

This bulky organic layer can act as a protection from external agents, increasing environmental stability. It also leads to a higher confinement of bounded excitons than their three-dimensional material counterpart,⁸ which can be advantageous

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for light emitting applications.⁹ From the fundamental point of view, manipulating the dimensionality of tin-based perovskites can also impact their bandgap, carrier mobility, and other electronic properties due to the spatial confinement of charge carriers. For example, 2D tin halide perovskites present radiative recombination of free excitons by increasing the exciton 2D binding energy.^{10,11} On the other hand, 0D tin halide perovskites exhibit radiative recombination of self-trapped excitons (STEs) at energy levels originating from the optical distortion of Sn–X bonds.^{12,13} In particular, STEs refer to the strong electron (exciton) phonon coupling when a soft lattice is distorted upon photoexcitation.¹⁴ Interestingly, 0D-tin halide perovskites show higher PLQYs, better environmental stability, higher Stokes shift, and a major exciton binding energy compared to their 2D and 3D counterparts.^{13,15,16}

However, inorganic and hybrid tin halide compounds can adopt multiple dimensionalities depending strongly on the polyhedron type (owing to the high coordination number of Sn(II) ions), inorganic or organic cations, oxidation number, and preparation conditions (pH, temperature, atmosphere, etc.).^{17,18} For example, zero-dimensional tetrabutylammonium (TBAC) tin(II) chloride could be obtained under acidic conditions with stable pyramidal $[\text{SnCl}_3]^-$ anions.¹⁹ If the monovalent cation is changed to Cs^+ , 3D- CsSnCl_3 with $[\text{SnCl}_6]^{4-}$ octahedrons is stable at high temperature because $5s^2$ orbitals hybridize with $5p_z$, resulting in hybrid orbitals with the same energy as that of np halogen orbitals.^{18,20} It should be mentioned that the octahedral tin chloride configuration is also stable for Sn(IV) ions.²¹ Different inorganic and organic tin(II/IV) halide compounds and dimensionalities are summarized in Table 1.

Thus, lead-free low-dimensional tin(II) halide perovskites are attracting scientific attention owing to their unique combination of interesting physical properties and lower toxicity.³¹ They show great potential as visible light emitters on optoelectronic devices³² as well as for enhancing the performance of photovoltaic devices.³³ Among multiple organic cations tested on low-dimensional hybrid metal halide perovskites, 2-thiopheneethylammonium (TEA^+) has drawn attention as a heteroatomic organic spacer and a passivating agent for lead iodide perovskites, with a stable power conversion efficiency (PCE) of 18.75% for 1000 h^{34,35} when used in solar cells. Concerning tin-based perovskites,

the synthesis of 2D- TEA_2SnI_4 has been reported at different scales, from nanocrystals³⁶ and thin films³⁷ to single crystals,³⁸ showing a PLQY around 19% for nanodisks in solution prepared by the LARP method³⁶ and 23% for thin films prepared using the antisolvent approach.³⁷ However, to the best of our knowledge, the synthesis and characterization of TEA-based Sn(II) chloride and bromide remain unexplored.

The dimensionality of perovskite nanocrystals synthesized by the hot injection method can be modulated by means of changes in the precursor concentrations.^{39–41} Dimensional modulation improves the optical properties of lead-free perovskites and their PLQY owing to the separation and distortion of metal halide octahedrons with lone electron pair ns^2 .⁴² For instance, Cs-oleate solution precursor concentration is a cornerstone parameter that determines whether 0D- Cs_4SnBr_6 or 3D- CsSnBr_3 nanocrystals are obtained.^{43,44} Appropriate dimensionality manipulation of lead-free perovskites expands the range of these less toxic materials with a full gamut of wavelength emission depending on the type of inorganic moiety and organic cation. Therefore, the hot injection approach could be re-imagined as a controlled method to achieve high luminescent perovskite nanostructures.^{39,43}

Here, we report the synthesis of 2D and 0D TEA-based Sn(II) halide perovskite microplates by the hot injection method. The control of the $[\text{TEA}]/[\text{Sn}^{2+}]$ ratio for the bromide-based perovskite adjusts the resulting perovskite dimensionality from 2D to 0D. The PLQY was strongly affected by this factor, ranging from 0.1% for 2D- $\text{TEA}_2\text{SnBr}_4$ to 75% for orange emitting 0D- $\text{TEA}_4\text{SnBr}_6$. Interestingly, TEA/Sn(II)-driven dimensionality control only takes place with TEA-based tin(II) bromide derivatives. Conversely, the reaction between SnCl_2 and SnI_2 with TEA provides main products pure 0D- $[\text{TEASnCl}_3][\text{TEACl}]$ and 2D- TEA_2SnI_4 microplates, respectively.

Materials and methods

Chemicals

Mesitylene (97% Pure) and iodotrimethylsilane (TMSI, 95–97% stabilized) were purchased from Thermo Scientific.

Table 1 Dimensionality of different inorganic and hybrid tin halide compounds

Cation	Compound	Polyhedral unit	Dimensionality	Ref.
TBAC ⁺	(TBAC)SnCl ₃	$[\text{SnCl}_3]^-$	0D	19
R/S- α -PEA ⁺	R/S- α -PEA SnX ₃ , (X = Cl, Br)	Distorted $[\text{SnX}_6]$	1D	22
MA ⁺ , FA ⁺	ASnI ₃ (A = MA ⁺ , FA ⁺)	$[\text{SnI}_6]^{4-}$	3D	23
PEA ⁺	PEA ₂ SnI ₄	$[\text{SnI}_6]^{4-}$	2D	24
H ₄ BAPP ³⁺	H ₄ BAPP ₂ SnX ₅ (X = Cl, Br)	$[\text{SnX}_5]^{3-}$	0D	25
C ₇ H ₈ N ₃ ⁺	(C ₇ H ₈ N ₃) ₄ SnBr ₆	$[\text{SnBr}_6]^{4-}$	0D	26
Cs ⁺	CsSn ₂ Br ₅	$[\text{Sn}_2\text{Br}_5]^-$	1D	27
DAO ²⁺	DAOSn ₂ I ₆	$[\text{Sn}_2\text{I}_6]^{2-}$	1D	28
IPA ⁺	μ -[IPA] ₃ Sn ₂ I ₇	$[\text{Sn}_2\text{I}_7]^{3-}$	2D	29
(CH ₃) ₃ S ⁺	$[(\text{CH}_3)_3\text{S}]_2\text{SnCl}_6 \cdot \text{H}_2\text{O}$	$[\text{SnCl}_6]^{2-}$	0D	21,27
Cs ⁺ , MA ⁺ , FA ⁺	A ₂ SnI ₆	$[\text{SnI}_6]^{2-}$	Undefined	30

List of acronyms. TBAC⁺ = tetrabutylammonium; R/S- α -PEA⁺ = R/S- α -phenethylammonium; MA⁺ = methylammonium; FA⁺ = formamidinium; H₄BAPP = 1,4-bis(3ammoniumpropyl)piperazinium; C₇H₈N₃⁺ = 2-amino benzimidazolium; Cs⁺ = cesium; DAO²⁺ = 1,8-octyldiammonium; IPA⁺ = isopropylammonium; (CH₃)₃S⁺ = trimethyl sulfonium.



Bromotrimethylsilane (TMSBr, 98%) and *n*-octane were acquired from Acros Organics. Chlorotrimethylsilane (TMSCl, 98%), trioctylphosphine (TOP, 97%), tin(II) ethylhexanoate (95–100%), *n*-hexane and oleic acid (OLAC, technical grade of 90%) were obtained from Sigma Aldrich. 2-Thiopheneethylamine (TEA, > 98%) was purchased from TCI Chemicals. All the chemicals were used without further purification.

Experimental procedure

Synthesis of SnX₂-TOP precursor solution. In a standard procedure, 1 mL of tin(II) 2-ethylhexanoate and 2 mL TOP were mixed in a 25 mL round bottom flask and degassed in a vacuum with magnetic stirring for 45 min at room temperature. After that, the atmosphere was changed by nitrogen flux, and three mmol of TMSX was injected, keeping a molar ratio of [TMSX]/[Sn²⁺] of 2. Finally, the SnX₂-TOP solution was stored under a nitrogen atmosphere for further use.

Synthesis of 0D-[TEASnCl₃][TEACl], 0D-TEA₄SnBr₆ and 2D-TEA₂SnX₄ (X = Br, I) microplates. In a 100 mL 3-neck round bottom flask, 20 mL of mesitylene and 1 mL of oleic acid were degassed under vacuum for 1 h at room temperature. Subsequently, the SnX₂-TOP solution was injected into the reaction system at room temperature to obtain a precursor solution with a nominal concentration of 0.12 M of SnX₂-TOP. After that, the temperature was increased until 160 °C, and TEA was injected. Different [TEA]/[Sn²⁺] molar ratios between 0.3 and 4 were tested to modulate the dimensionality of the final product. After 5 s, the reaction mixture was chilled in an ice bath with continuous magnetic stirring. TEA-based tin(II) halide microplates were isolated by centrifugation and washed three times with *n*-hexane. The final product was re-dispersed in *n*-octane and stored in a glovebox for further characterisation. All proceedings (synthesis and purification) were carried out under normal atmospheric conditions of moisture and oxygen.

Characterisation

X-ray diffraction (XRD) diffractograms were measured with an Empyrean diffractometer from PANalytical and the measurements were conducted in the range of $2\theta = 4^\circ$ – 60° with a 2θ step of 0.026° and using a Cu target as an X-ray source ($K\alpha$ Cu radiation, $\lambda = 0.15406$ nm). For XRD measurements, 10 μ m films were prepared by drop-casting lead-free microcrystal dispersions on *n*-octane over 0.5 cm $\times$ 0.5 cm glass substrates and drying at room-temperature under an inert atmosphere.

¹H and ¹³C nuclear magnetic resonance (NMR) spectra in the liquid phase for 0D-TEA₄SnBr₆ and 2D-TEA₂SnX₄ (X = Br, I) were recorded with a Bruker AV400 spectrometer, with an Oxford Magnet of 400 MHz (9.4 T). ¹H and ¹³C NMR spectra are reported in ppm (δ) relative to the chemical shift of solvent residual signals for DMSO-*d*₆ at 2.51 ppm and 39.99 ppm, respectively.

¹³C and ¹¹⁹Sn magic angle solid state nuclear magnetic resonance (MAS-NMR) spectra of 0D-[TEASnCl₃][TEACl], 0D-TEA₄SnBr₆, and 2D-TEA₂SnX₄ (X = Br, I) were measured with a Bruker Advance III 400 WB spectrometer, with an ultra-shielded Widemouth magnet (89 mm), 400 MHz (9.4T),

two RF channels and high-power amplifiers for solid-state experiments. For ¹¹⁹Sn MAS-NMR spectra measurements, the SnO powder sample was used as a reference. On the other hand, for ¹³C and ¹¹⁹Sn MAS-NMR and liquid state NMR, perovskite microcrystals were isolated from *n*-octane suspension by centrifugation and dried under an inert atmosphere overnight. ¹¹⁹Sn MAS-NMR spectra of 0D-[TEASnCl₃][TEACl] were measured at different spinning rates (8 kHz and 10 kHz).

Transmission electron microscopy (TEM) images were acquired from a Hitachi HT7800 microscope with a high-resolution filament of LaB₆ with an acceleration voltage of 100 kV. TEA-based tin(II) halide microcrystal suspensions in *n*-octane were dropped on Cu/C TEM grids.

Energy-dispersive X-ray spectroscopy (EDS) microanalysis was performed using an Oxford Ultim Max 170 detector monitored with Aztec software coupled to a SCIOS2 FESEM microscope from Thermo Scientific. For these measurements, TEA-based tin(II) halide microcrystals were deposited on Si substrates, and sample borders were coated with Ag paste.

Photoluminescence (PL) and photoluminescence excitation (PLE) measurements were conducted in an FLS1000 spectrophotometer from Edinburgh Instruments. A xenon lamp (Xe2) was used as an excitation source. PL spectra were recorded with a PMT-980 detector from Edinburgh Instruments. PL and PLE spectra were measured with a dwell time of 1 s and a wavelength step of 1 nm. The PL quantum yield (PLQY) was measured with an integrating sphere coupled to an FLS1000 spectrophotometer from Edinburgh Instruments. TEA based tin(II) halide microcrystal films were prepared over 0.25 cm² quartz substrates by drop cast deposition to measure PLE and PL spectra and PLQY.

Results and discussion

The hot injection method was employed to synthesize 0D-TEA₄SnBr₆ and 2D-TEA₂SnBr₄ microcrystals by adjusting the precursor concentration, as detailed in the Experimental section, with [TEA]/[Sn²⁺] ratios of 1 and 0.3, respectively. Thin films fabricated from 2D-TEA₂SnBr₄ microcrystals exhibited flakes with an average particle size of $7.66 \mu\text{m} \pm 5.2 \mu\text{m}$, and a mean thickness of 280 nm (Fig. 1a–c). On the other hand, quasi-rectangular plates with a length of $\sim 32 \pm 8 \mu\text{m}$ and a thickness of $\sim 3.55 \mu\text{m}$ were predominantly observed on 0D-TEA₄SnBr₆ thin films (Fig. 1d–f).

Fig. 2a shows the XRD diffractograms of TEA-based Sn(II) bromide perovskite films for different [TEA]/[Sn²⁺] ratios. From [TEA]/[Sn²⁺] < 0.5, the XRD diffractogram shows the multiple reflections peaks at 2θ of 5.62° , 11.32° , 17.02° , 22.78° , 28.64° and 34.56° . These reflections are spaced apart by a 2θ interval of 5.7° , as expected from a Ruddlesden–Popper phase for pure 2D TEA₂SnBr₄.³⁶ Additionally, EDS analysis (Table S1, ESI[†]) reveals that Br/Sn and S/Sn atomic ratios are in good agreement with the expected stoichiometry. As the [TEA]/[Sn²⁺] ratio increases, isolated metal halide octahedra [SnBr₆]^{4–} or clusters are generated. From [TEA]/[Sn²⁺] > 0.5, new non-periodically XRD



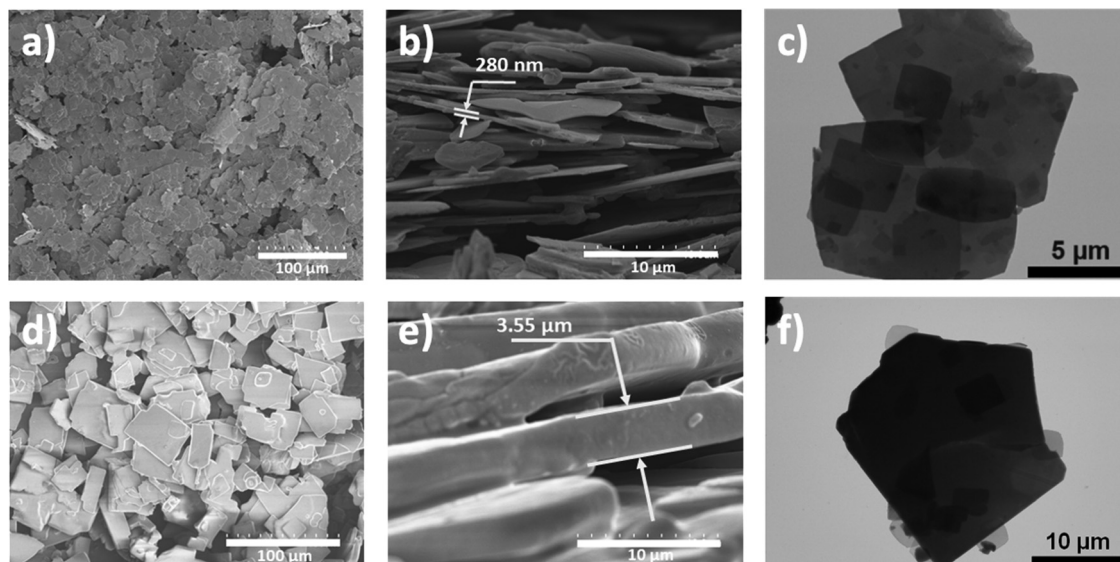


Fig. 1 Top-view SEM (a) and (d), cross-sectional SEM (b) and (e) and TEM images (c) and (f) of 2D-TEA₂SnBr₄ (a)–(c) and 0D-TEA₄SnBr₆ (d)–(f) microcrystals.

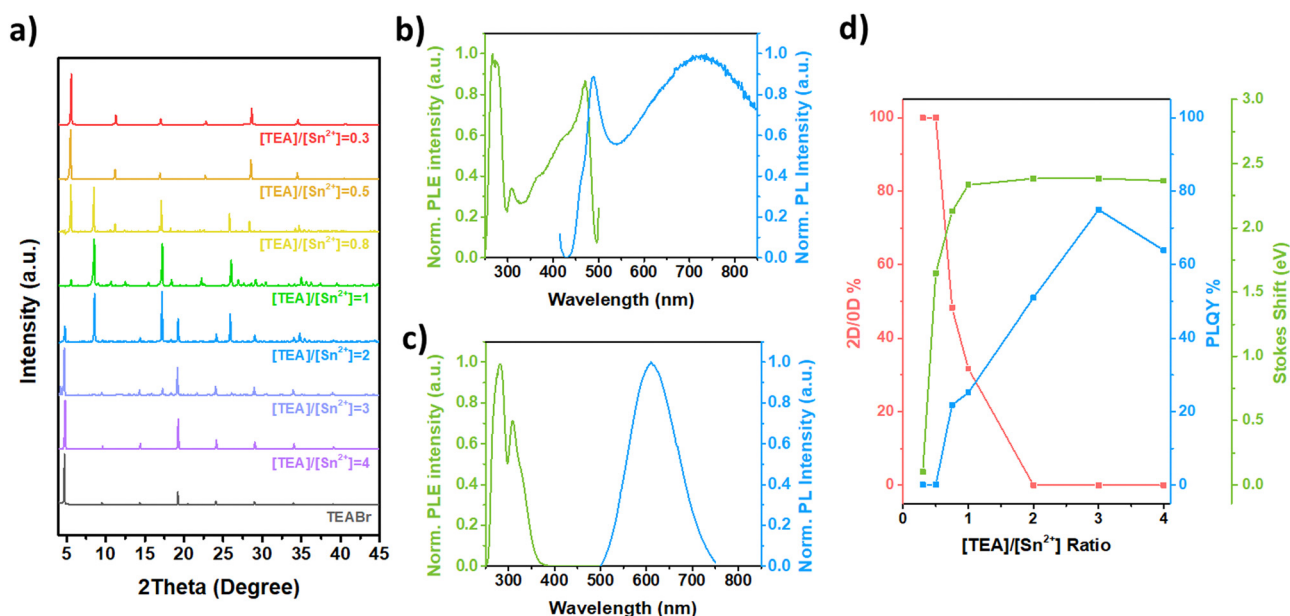


Fig. 2 (a) XRD patterns of 2D-TEA₂SnBr₄ and 0D-TEA₄SnBr₆ samples synthesized at different molar ratios [TEA]/[Sn²⁺]; PL and PLE spectra of: (b) 2D-TEA₂SnBr₄ and (c) 0D-TEA₄SnBr₆; (d) 2D/0D ratio (red line), PLQY % (blue line) and Stokes shift (green line) as a function of [TEA]/[Sn²⁺] ratio. PLQYs of 2D-TEA₂SnBr₄ and 0D-TEA₄SnBr₆ thin film samples were measured at a fixed wavelength of 280 nm.

peaks are observed at 8.52°, 17.13°, 18.23° and 25.84°, which are attributed to 0D-TEA₄SnBr₆. Atomic percentages of Br are higher than those measured for 2D-TEA₂SnBr₄ microplates (Table S1, ESI[†]), which supports the proposed chemical formula.

Concurrently, the increase of TEA molarities results in a decrease in the formation of the 2D-TEA₂SnBr₄ phase. From [TEA]/[Sn²⁺] > 2, the diffractograms reveal only reflection peaks of 0D-TEA₄SnBr₆ along with those of pure TEABr at 2θ values of

4.5°, 9°, 14°, 19°, 24°, 29° and 34°. As expected, 0D-TEA₄SnBr₆ reflection peaks do not exhibit the same periodic ordering observed in the 2D-TEA₂SnBr₄.

The differences between 2D-TEA₂SnBr₄ and 0D-TEA₄SnBr₆ crystal phases and band structures also generate remarkable differences in their optical properties. PL emission and PL excitation (PLE) spectra of TEA-Sn(II) bromide films synthesized at different molar ratios are shown in Fig. S1 (ESI[†]). In particular, the [TEA]/[Sn²⁺] ratio of 0.5, for 2D-TEA₂SnBr₄ films,



is represented in Fig. 2b. Its PL spectra can be deconvoluted into two bands; one sharp band at 470 nm, and another broader band centered at 730 nm. The shape of the PL emission bands is similar to that reported for 2D-PEA₂PbI₄,⁴⁵ thus mainly determined by the radiative recombination of free excitons. Additionally, the PLE spectrum exhibits a sharp band at 470 nm. From the spectral difference in the position of the band maximum of the PL and PLE spectra, we measured the Stokes shift energy, which is around 113 meV. The sharp PL and PLE bands shown in Fig. 2b, along with the high order observed in XRD, may be associated with the synthesis of lamellar-like microcrystals. These structures consist of intercalated layers comprising [SnBr₆]^{4−} octahedra and organic layers of TEA⁺ cations, as reported for 2D-TEA₂SnI₄ nanocrystals and thin films.⁴⁶ In contrast, the PL emission and PLE spectra of 0D-TEA₄SnBr₆ films synthesized with a [TEA]/[Sn²⁺] ratio of 3 are displayed in Fig. 2c. When the perovskite was exposed to UV light (wavelength of 282 nm), 0D-TEA₄SnBr₆ exhibited a PL spectrum characterized by a band centered at 610 nm and an FWHM of 130 nm. The PLE spectrum shows two sharp bands with a maximum at 282 nm and 310 nm. From PL and PLE spectra, we measured a strong Stokes shift of 2.38 eV, which is attributed to the radiative recombination of self-trapped excitons.^{47,48}

The larger bandgaps and Stokes shifts when transitioning from the 2D to 0D TEA-based Sn bromide perovskite are attributed to a reduction of the orbital overlap between [SnX₆]^{4−} octahedra as the 0D phase becomes more predominant with increasing [TEA]/[Sn²⁺] ratio. This occurs because TEA⁺ cations separate the [SnBr₆]^{4−} octahedra of 2D layers to form isolated [SnBr₆]^{4−} octahedra. This results in a higher exciton binding energy, lower carrier mobility, and increased stability. The isolated

structure of 0D perovskites also imparts a strong quantum confinement effect, contributing to a higher PL emission compared to that of higher dimensionality.

Fig. 2d shows the 0D/2D phase ratio, PLQY, and Stokes shift of TEA-based Sn(II) bromide perovskite films depending on the [TEA]/[Sn²⁺] molar ratio. To estimate the phase ratio between 2D and 0D materials, we used eqn (1):

$$R_{2D/0D} = \frac{A_{2D}}{A_{0D} + A_{2D}} \quad (1)$$

where $R_{0D/2D}$ is the 2D/0D phase ratio in the sample, A_{0D} and A_{2D} are the integrated area under XRD peaks (Fig. 2a) corresponding to 2D-TEA₂SnBr₄ ($2\theta_{2D} = 5.62^\circ, 11.32^\circ, 17.02^\circ, 22.78^\circ, 28.64^\circ$ and 34.56°) and 0D-TEA₄SnBr₆ ($2\theta_{0D} = 8.52^\circ, 17.13^\circ, 18.23^\circ$ and 25.84°), respectively. In the transition ratios, as the presence of 0D phase increases, we observe a dramatic increase in the PLQY from 0.15% for the pure 2D perovskite to 75% at $\lambda_{ex} = 280$ nm and 62.75% at $\lambda_{ex} = 280$ nm for a [TEA]/[Sn²⁺] ratio of 3 (with a slight PLQY reduction at higher ratios). Moreover, the Stokes shift increases noticeably (Fig. 1d) from 113 meV from 2D-TEA₂SnBr₄ to a maximum of 2.36 eV for 0D-TEA₄SnBr₆.

Based on our experimental results, the growth of 0D-TEA₄SnBr₆ is strongly influenced by the isolation of [SnBr₆]^{4−} octahedra layers from the 2D-TEA₂SnBr₄ material onto individual clusters resulting from the increase of TEA⁺ concentration. Also, the growth of 0D-TEA₄SnBr₆ is not influenced by coupling with mesitylene or oleic acid molecules as it is the case of the phenethylammonium tin(II) bromide perovskite, and thus the separation of inorganic moieties is influenced mostly by the increment in the TEA⁺ concentration.⁴⁹ As a consequence,

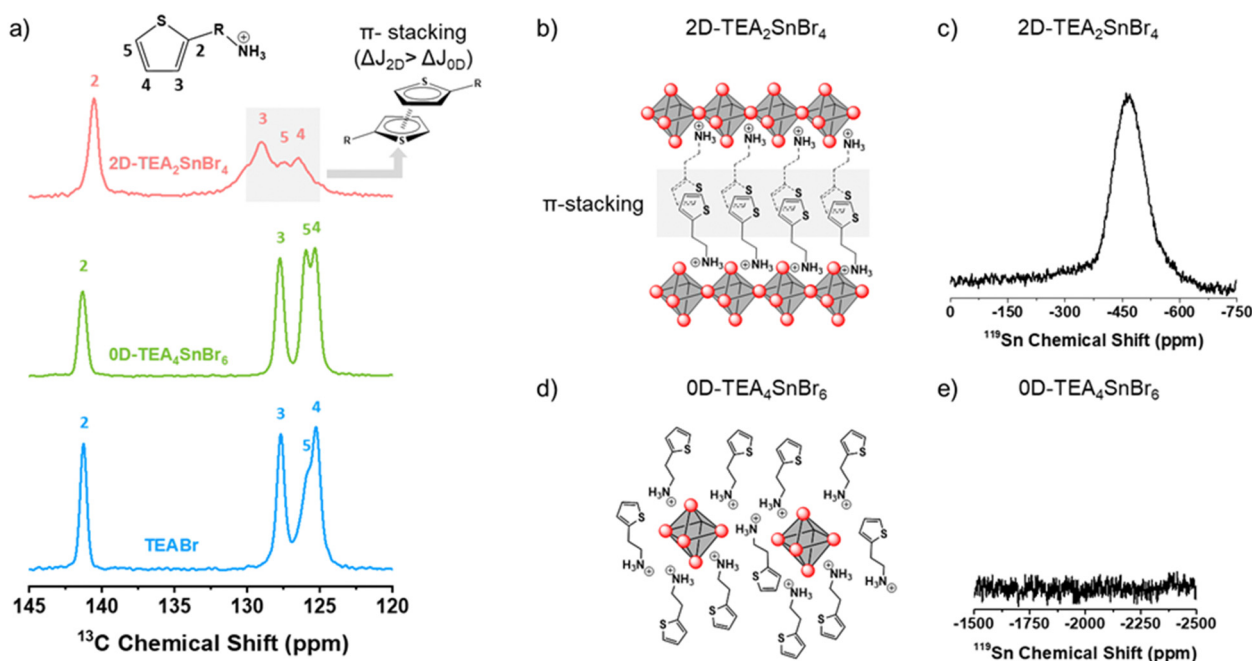


Fig. 3 (a) ¹³C MAS-NMR spectra of TEABr, 0D-TEA₄SnBr₆, and 2D-TEA₂SnBr₄ powders. Crystalline structures and ¹¹⁹Sn NMR of (b) and (c) 2D-TEA₂SnBr₄ and (d) and (e) 0D-TEA₄SnBr₆.



highly luminescent 0D-TEA₄SnBr₆ microcrystals can be synthesized by controlling the concentration of TEA.⁵⁰

A qualitative magic angle solid state (MAS) NMR analysis was carried out to investigate the local environment and chemical bonding of carbon (¹³C) and tin (¹¹⁹Sn) nuclei in solid-state materials (Fig. 3). 2D-TEA₂SnBr₄ and 0D-TEA₄SnBr₆ are compared with TEABr to determine the internal structure of low dimensional TEA-based Sn(II) bromide perovskites.

The ¹³C NMR spectra of 2D-TEA₂SnBr₄, 0D-TEA₄SnBr₆, and TEABr are shown in Fig. 3a. A set of four signals, ranging from 145 to 120 ppm, has been ascribed to the aromatic carbons of TEA⁺. J₁, J₂, and J₃ signals observed in 2D-TEA₂SnBr₄ are shifted to a higher field ($\Delta J = 1.21$ ppm) than those of 0D-TEA₄SnBr₆ ($\Delta J = 0.41$ ppm) in comparison with TEABr. A characteristic signature of π -stacking interactions often manifests in changes to signal intensity and peak broadening in NMR spectra. The π -stacking interactions among aromatic rings lead to dynamic processes, causing an exchange between stacked and unstacked conformations. This equilibrium results in broader peaks in the NMR spectra due to variations in the chemical environments experienced by the interacting nuclei.

At a low [TEA]/[Sn²⁺] ratio, TEA⁺ can be distributed between layers of [SnBr₆]⁴⁻ octahedra, leading to the formation of a 2D tin bromide perovskite. In this structure, [SnBr₆]⁴⁻ octahedra are interconnected through four corner-sharing, creating 2D perovskite monolayers. Under these conditions, we can assume that thiophene rings self-assemble by π - π stacking interaction between thiophenes of adjacent 2D perovskite monolayers in 2D-TEA₂SnBr₄.⁵¹ This combination of structural features contributes to the unique properties and characteristics of the resulting material.⁵² Conversely, when the [TEA]/[Sn²⁺] ratio is high, the TEA⁺ cations form a surrounding layer around individual [SnBr₆]⁴⁻ octahedra, resulting in a lack of interconnection

between the octahedra. The high [TEA]/[Sn²⁺] ratio influences the structural organization, and the isolation between [SnBr₆]⁴⁻ octahedra enables their strong quantum confinement effect, which has implications for the properties of the material, such as radiative recombination by self-trapped excitons.⁴⁴ The TEA⁺ cations randomly distributed around the [SnBr₆]⁴⁻ octahedra reduce the π stacking interaction. As a result, the chemical shifts of aromatic carbons observed in 0D-TEA₄SnBr₆ are very similar to those of TEABr.⁵³

Regarding the ¹¹⁹Sn MAS-NMR spectra, 2D-TEA₂SnBr₄ (Fig. 3d and Fig. S2, ESI†) shows two different signals at -465 ppm and -1983 ppm assigned to Sn(II) and Sn(IV) respectively,²⁰ with a lower intensity for the latter. In the case of 0D-TEA₄SnBr₆, no significant signal was observed between -1000 and -2000 ppm. This is fully aligned with the better ambient stability displayed by the 0D-TEA₄SnBr₆ compared to its bidimensional counterpart.

Optical and structural properties of 0D-[TEASnCl₃][TEACl] and 2D-TEA₂SnI₄ microcrystals

It is key to understand if a similar approach can be carried out to selectively control the dimensionality of the expanded TEA-based Sn(II) halide family. Microcrystals based on chloride can be synthesized with different [TEA]/[Sn²⁺] ratios. The XRD patterns of TEA-based Sn(II) chloride films (Fig. 4a) displayed reflection peaks at 8.63°, 17.26°, 18.09°, and 26.03° for ratios ranging from 0.125 to 1. These peaks show slight shifts compared to those of 0D-TEA₄SnBr₆ (2θ values of 8.52°, 17.13°, 18.23°, and 25.84°) (Fig. 2a), which is attributed to a decrease in the lattice parameters resulting from the use of chloride.⁵⁴ For [TEA]/[Sn²⁺] ratios between 2 and 4, XRD is dominated by the reflections of the TEACl salt, which is in excess. Noticeably, PL spectra show a broad emission band between 450 and 850 nm,

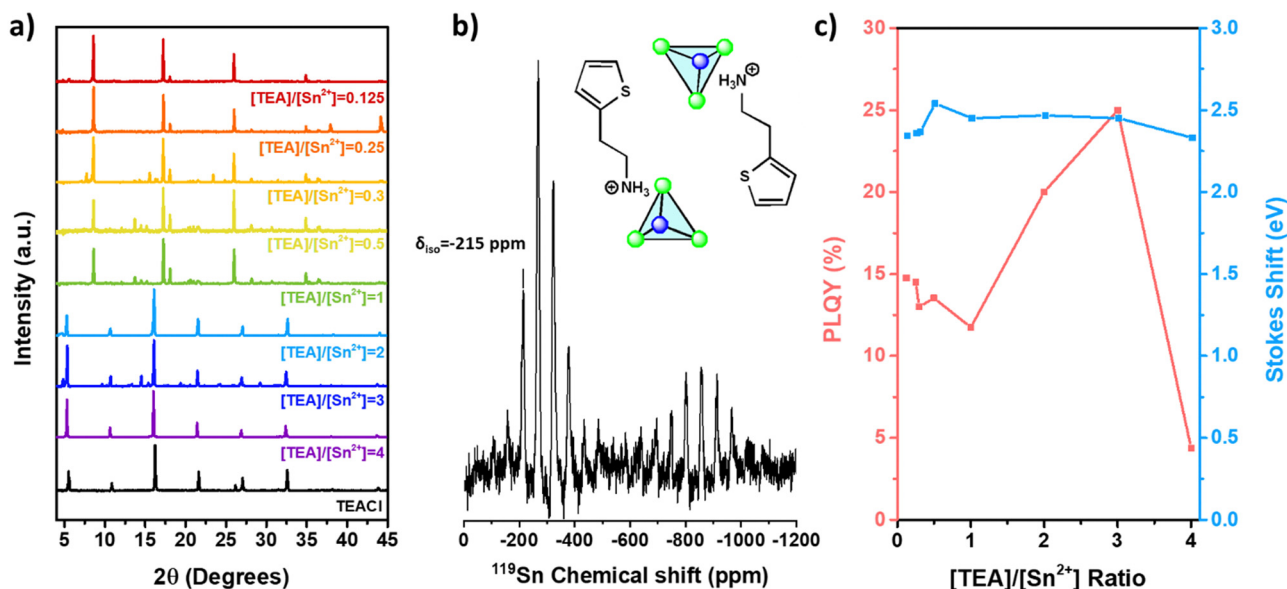


Fig. 4 (a) XRD patterns, (b) ¹¹⁹Sn MAS-NMR spectra measured at room temperature (Inset: probable crystalline structure) and (c) variation of the PLQY and Stokes shift of 0D-[TEASnCl₃][TEACl] microcrystals as a function of the molar ratio [TEA]/[Sn²⁺].



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