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Designing barrier-free metal/MoS₂ contacts through electrene insertion†

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Transition-metal dichalcogenides (TMDCs), including MoS₂, have great potential in electronics applications. However, achieving low-resistance metal contacts is a challenge that impacts their performance in nanodevices due to strong Fermi-level pinning and the presence of a tunnelling barrier. As a solution, we explore a strategy of inserting monolayers of alkaline-earth sub-pnictide electrenes with a general formula of [M₂X]^{+e⁻} (M = Ca, Sr, Ba; X = N, P, As, Sb) between the TMDC and the metal. These electrenes possess two-dimensional sheets of charge on their surfaces that can be readily donated when interfaced with a TMDC semiconductor, thereby lowering its conduction band below the Fermi level and eliminating the Schottky and tunnelling barriers. In this work, density-functional theory (DFT) calculations were performed for metal/electrene/MoS₂ heterojunctions for all stable M₂X electrenes and both Au and Cu metals. To identify the material combinations that provide the most effective Ohmic contact, the charge transfer, band structure, and electrostatic potential were computed. Linear correlations were found between the charge donated to the MoS₂ and both the electrene surface charge and work function. Overall, Ca₂N appears to be the most promising electrene for achieving an Ohmic metal/MoS₂ contact due to its high surface charge density.

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1 Introduction

The field of nanoelectronics is poised to undergo a major revolution with the introduction of novel two-dimensional (2D) semiconductor materials. Monolayer transition-metal dichalcogenides (TMDCs) represent a pioneering family of materials in this regard, with MoS₂ serving as their leading member. The stunning features of TMDC-based transistors, such as high mobility,¹ extremely short channel lengths,² high on/off ratio,³ and low sub-threshold swing,⁴ make this class of materials well suited for electronic device applications. However, high metal/semiconductor contact resistance has inhibited the efficacy of TMDC-based transistors.^{1,5–7} The high contact resistance is primarily caused by (i) the Schottky barrier and Fermi-level pinning, despite the lack of dangling or incomplete bonds on the surface

of the TMDC,^{8–11} and (ii) the tunnelling barrier caused by the van der Waals (vdW) gap at the metal/TMDC interface.^{5,12–14}

The choice of metal plays an important role in the performance of 2D contacts. For some metal/TMDC interfaces, such as Pd/MoS₂, overlap of metal and semiconductor states contributes to electron or hole injection, so that negligible tunnelling barriers are feasible.^{6,13} One way to fully overcome the tunnelling barrier is to choose specific metals that can form covalent bonds with the TMDC layer, *e.g.*, Ti when interfaced with MoS₂.^{12,15–17} However, such bonding results from strong hybridization of the metal and semiconductor orbitals, which can potentially disrupt the original MoS₂ states and deteriorate the sheet resistivity of the semiconductor.^{5,13} That being said, hybridization of the MoS₂ energy bands with the semimetallic materials bismuth and antimony has been demonstrated as one of the best strategies to push the contact resistance towards the quantum limit.^{18,19}

Another current strategy to develop practical TMDC-based devices involves introducing an intermediate layer between the metal and 2D semiconductor to address both Fermi-level pinning and tunnelling barrier issues.^{9,20–25} One major cause of Fermi-level pinning in 2D metal contacts involving MoS₂ is the production of gap states, mainly of Mo d-orbital character.¹¹ Based on density-functional theory (DFT) calculations, we previously proposed that the insertion of a Ca₂N monolayer not only removes the Fermi-level pinning, but also eliminates both

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† Electronic supplementary information (ESI) available: Geometric details of all interfaces, computed Bader charge transfer, decomposition of the strain energies, summary of MoS₂ band edges and gaps, results of hybrid DFT calculations on the Ca₂N/MoS₂ interface, and results of WKB calculations. Also included are electronic band structures, DOS, and electrostatic potentials for all studied metal/electrene/MoS₂ interfaces, as well as band structures of the isolated materials at the strained interface geometries. See DOI: <https://doi.org/10.1039/d3cp06112d>



Fig. 1 Top: Schematic design of the Au/MoS₂ interface with monolayer Ca₂N as an intermediate layer. The associated band diagram indicates the removal of the Schottky barrier upon metallization of the MoS₂ bands. Bottom: List of all stable M₂X electrides, along with a 3D representation of the electrene's surface charge.

Schottky and tunnelling barriers, for a model Au/Ca₂N/MoS₂ heterostructure.²⁰ Fig. 1 (top) shows the atomic configuration and a schematic illustration of the electronic band diagram of a proposed Au/Ca₂N/MoS₂ interface. Here, Φ_{Au} and $\Phi_{\text{Ca}_2\text{N}}$ are the work functions of the metal and the electrene, X_{MoS_2} is the electron affinity of the semiconductor, and E_{vac} and E_{F} are the vacuum and Fermi levels, respectively. The conduction and valence bands of MoS₂ are also shown as E_{c} and E_{v} . Due to charge transfer from the Ca₂N monolayer, the MoS₂ is negatively doped so that its conduction band edge falls below the Fermi level ($E_{\text{c}} < E_{\text{F}}$), removing the Schottky barrier for electron transport.

Ca₂N is a 2D electride;^{26,27} this is an ionic material in which the anion is an excess electron that is localized to the interstitial regions between the atomically thin, positively charged, layers of the material. Upon exfoliation, the anionic electrons form delocalized 2D sheets of charge surrounding the surface of the monolayer, now termed an electrene.^{28,29} The bottom-right panel of Fig. 1 illustrates the surface charge density of an electrene. These surface states could potentially lead to high charge transport, allowing electrines to find utility in device applications. Ca₂N is only one member of the family of 2D electrides with a general chemical formula of M₂X (M = Ca, Sr, Ba; X = N, P, As, Sb). The crystal structures of bulk Ca₂N, Sr₂N, and Ba₂N,^{27,30} as well as few-layer Ca₂N,²⁸ have been experimentally characterized, while the others are proposed to be layered electrides on the basis of DFT studies.^{28–41} We have previously investigated periodic trends in the electronic, structural, and transport properties of the corresponding monolayer and bilayer electrines.⁴²

In this work, we employ DFT to investigate various electrines in contact with both Au and Cu metals and MoS₂ to

determine which material is the most promising for eliminating the Schottky and tunnelling barriers in metal/MoS₂ contacts. In particular, we determine the extent to which MoS₂ conduction bands become metallized following electrene insertion and analyze the electrostatic potential at the interfaces, which should shed light on electron transport across these contacts. We also identify various correlations between monolayer electride characteristics, such as surface charge densities, and the improved charge transfer and barrier lowering that occurs at the metal/electrene/MoS₂ interfaces.

2 Computational methods

2.1 DFT calculations

All DFT calculations were performed with the Quantum ESPRESSO program,^{43,44} using the Perdew–Burke–Ernzerhof (PBE) exchange–correlation functional,⁴⁵ the projector augmented-wave (PAW) method,^{46,47} and periodic boundary conditions. Dispersion interactions were treated with Grimme's D3BJ pairwise dispersion scheme⁴⁸ with Becke–Johnson damping.^{49,50} Calculations for all metal/electrene/MoS₂ structures used plane-wave cutoff energies of 80 Ry and 1200 Ry for the wave function and electron density, respectively, and a k -point grid density of $0.03 \times 2\pi \text{ \AA}^{-1}$. In order to avoid interactions between layers, we adopt a 20 \AA vacuum region. The Methfessel–Paxton (MP) smearing method⁵¹ was used to integrate the Brillouin zone with a smearing width of 0.1 eV. Bader charge analysis was performed using the Bader^{52–54} code *via* post-processing of the DFT electron density.

2.2 Heterostructure design

After fully relaxing the atomic positions and lattice constants of the isolated MoS₂ monolayer, electrene monolayers, and metal slabs, we employed VASPKIT⁵⁵ to build all the metal/MoS₂ and metal/electrene/MoS₂ heterostructures. Two criteria were considered: (i) the MoS₂ layer in all heterostructures is not strained to allow direct comparison between all the interfaces and (ii) the strain on both the electrene layer and the metal slab is under 3.5% for all cases, except for Au/Ba₂Sb/MoS₂, for which there is 4.3% strain for the electrene. Tables S1 and S2 (ESI†) summarize the number of atoms and the in-plane lattice constants for each heterojunction, as well as the percent strains on the electrene and metal slabs. The heterostructure geometries were relaxed with respect to the atomic positions, subject to fixed lattice constants, until the force on each atom was less than 0.1 meV \AA^{-1} . Additional geometry relaxations were also performed for the free metal slabs and electrene monolayers using the same lattice constants as the corresponding heterojunctions.

3 Results and discussion

3.1 Exfoliation energies

The structural properties of the studied metal/electrene/MoS₂ heterostructures, including the average interlayer distances, are summarized in Table S2 (ESI†). Overall, the electrene/MoS₂ interlayer distances are found to be smaller than the sum of

the van der Waals radii of the sulfur and alkaline earth metal atoms, taking values of 1.73, 2.41, 2.63, and 2.71 Å, for S, Ca, Sr, and Ba respectively.⁵⁶ This reflects the ionic (rather than vdW) nature of these contacts.

To assess the stabilities of the various heterostructures, as well as to quantify the relative importance of strain, ionic binding (from the base density functional), and dispersion interactions, their exfoliation energies were computed. Here, the unstrained exfoliation energy is defined as the energy required to infinitely separate the three materials in the heterojunction:

$$E_{\text{exfo}}^{\text{unstr}} = E_{\text{metal}}^{\text{unstr}} + E_{\text{electrene}}^{\text{unstr}} + E_{\text{MoS}_2} - E_{\text{metal/electrene/MoS}_2} \quad (1)$$

This involves the energies of the isolated MoS₂ monolayer, electrene monolayer, and metal slab, all at their fully relaxed geometries. As the electrene and metal are strained in the heterostructure, it is reasonable to instead use these strained lattice constants for relaxations of only the atomic positions (and not the lattice constants) of the isolated materials, giving rise to the strained exfoliation energy:

$$E_{\text{exfo}}^{\text{str}} = E_{\text{metal}}^{\text{str}} + E_{\text{electrene}}^{\text{str}} + E_{\text{MoS}_2} - E_{\text{metal/electrene/MoS}_2} \quad (2)$$

This can be further broken down into contributions from the base density functional and dispersion correction. Finally, the strain energy is

$$E_{\text{strain}} = E_{\text{metal}}^{\text{str}} + E_{\text{electrene}}^{\text{str}} - E_{\text{metal}}^{\text{unstr}} - E_{\text{electrene}}^{\text{unstr}} \quad (3)$$

The results in Table S3 (ESI[†]) show that virtually all of the strain energy arises from strain on the metal slab, rather than on the electrene. Note that the MoS₂ layer in all heterostructures is not strained.

The exfoliation energies in Table 1 show that formation of the metal/electrene/MoS₂ interfaces is energetically favourable, and the exfoliation energies are much higher than for the Au/MoS₂ and Cu/MoS₂ interfaces. This is expected due to the high

charge transfer from the electrene to the MoS₂ and the metal slab (*vide infra*) resulting in strong ionic bonding, as is evident from the large base-functional contributions to the exfoliation energies (ranging from 49–77% for the strained exfoliation energies). The relative importance of the ionic interactions is largest for the light electrines (Ca₂N, Sr₂N, and Sr₂P), with the dispersion contributions being more significant for the heavier electrines containing barium, particularly for Ba₂Sb. From the results in Table 1, we also note that there is generally reduced dispersion binding for Cu as opposed to Au, as expected from the greater atomic size and polarizability of Au. Finally, the strain energies are typically smaller for the Cu interfaces, due to reduced strain on the metal.

As shown in Table 1, all metal/electrene/MoS₂ contacts have much higher binding energies than other high-efficiency 2D TMDC contacts. For example, PtTe₂/MoS₂ and graphene/MoS₂ have binding energies of 12.6 meV Å⁻² and 0.7 meV Å⁻²,⁵⁷ respectively. This indicates that the metal/electrene/MoS₂ contact should have higher thermal stability, and the stronger interlayer interactions should promote better charge-carrier injection.

3.2 Bands, densities of states, and charge transfer

Images of the interfaces with inserted electrines, as well as plots of their electronic band structures, densities of MoS₂ states, and electrostatic potentials, are collected in Fig. S4–S17 (ESI[†]). We take the two metal/Ca₂N/MoS₂ heterostructures as illustrative examples, with the results shown in Fig. 2. The interface structures are shown in Fig. 2(a) and (e), for Au and Cu, respectively. Panels (b) and (f) show the computed band structures, decomposed according to contributions from each component material. The densities of MoS₂ states are shown in Fig. 2(c) and (g). Integrating the density of MoS₂ states from a particular minimum energy (E_{min} , below which D_{MoS_2} is nearly zero, selected to be -0.5 eV in this work) up to the Fermi energy can reveal the extent of charge doping of the MoS₂ layer (Q_{MoS_2}):

$$Q_{\text{MoS}_2} = \frac{1}{A_{\text{Total}}} \int_{E_{\text{min}}}^{E_{\text{F}}} D_{\text{MoS}_2}(E) dE, \quad (4)$$

where $D_{\text{MoS}_2}(E)$ is the density of MoS₂ states and A_{Total} is the unit cell area (*i.e.* the total contact area). The integrated part of the DOS is highlighted in light blue in Fig. 2(c). Finally, panels (d) and (h) show the averaged value of the electrostatic potential with respect to the Fermi level, as well as the computed Bader charge transfer upon electrene insertion. The magnified region of Fig. 2(h) shows the tunnelling barrier height (TBH) and tunnelling barrier width (TBW). The TBH is defined as the energy difference between the peak of the averaged electrostatic potential at the interface and the Fermi energy, while the TBW is defined as the distance between the two successive intersections of the averaged potential curve with the Fermi energy.

To understand the changes in electronic structure upon electrene insertion, comparison must be made with data for the Au/MoS₂ and Cu/MoS₂ interfaces, shown in Fig. S1 (ESI[†]). These contacts have already been extensively studied

Table 1 Total exfoliation energies for metal/electrene/MoS₂ heterostructures. Values were computed using relaxed geometries of the isolated materials with either unstrained (unstr) or strained (str) lattice constants. The strain energies and the decomposition of the strained exfoliation energy into base functional and D3BJ dispersion contributions are also reported. All energies are in meV Å⁻², using the area of the strained unit cell

Electrene	None	Ca ₂ N	Sr ₂ N	Sr ₂ P	Ba ₂ N	Ba ₂ P	Ba ₂ As	Ba ₂ Sb
Au/electrene/MoS₂								
$E_{\text{exfo,unstr}}$	3.5	196.6	181.9	147.0	237.1	169.0	172.0	208.8
E_{strain}	48.2	49.7	47.7	46.6	3.0	46.7	49.5	52.5
$E_{\text{exfo,str}}$	51.7	245.7	229.6	193.6	240.1	215.7	221.5	261.3
$E_{\text{base,str}}$	-7.8	179.5	162.6	144.7	142.2	135.8	132.1	128.5
$E_{\text{disp,str}}$	59.5	66.2	67.0	48.9	97.9	79.9	89.5	132.8
Cu/electrene/MoS₂								
$E_{\text{exfo,unstr}}$	59.8	178.1	198.2	155.5	192.2	151.1	162.8	240.2
E_{strain}	36.7	17.0	8.8	16.3	23.5	35.8	33.9	7.5
$E_{\text{exfo,str}}$	96.5	195.2	207.0	171.8	215.7	187.0	196.8	247.6
$E_{\text{base,str}}$	21.9	145.1	139.3	133.9	127.9	108.5	113.2	122.9
$E_{\text{disp,str}}$	74.6	50.0	83.3	37.9	87.9	78.5	83.6	124.8

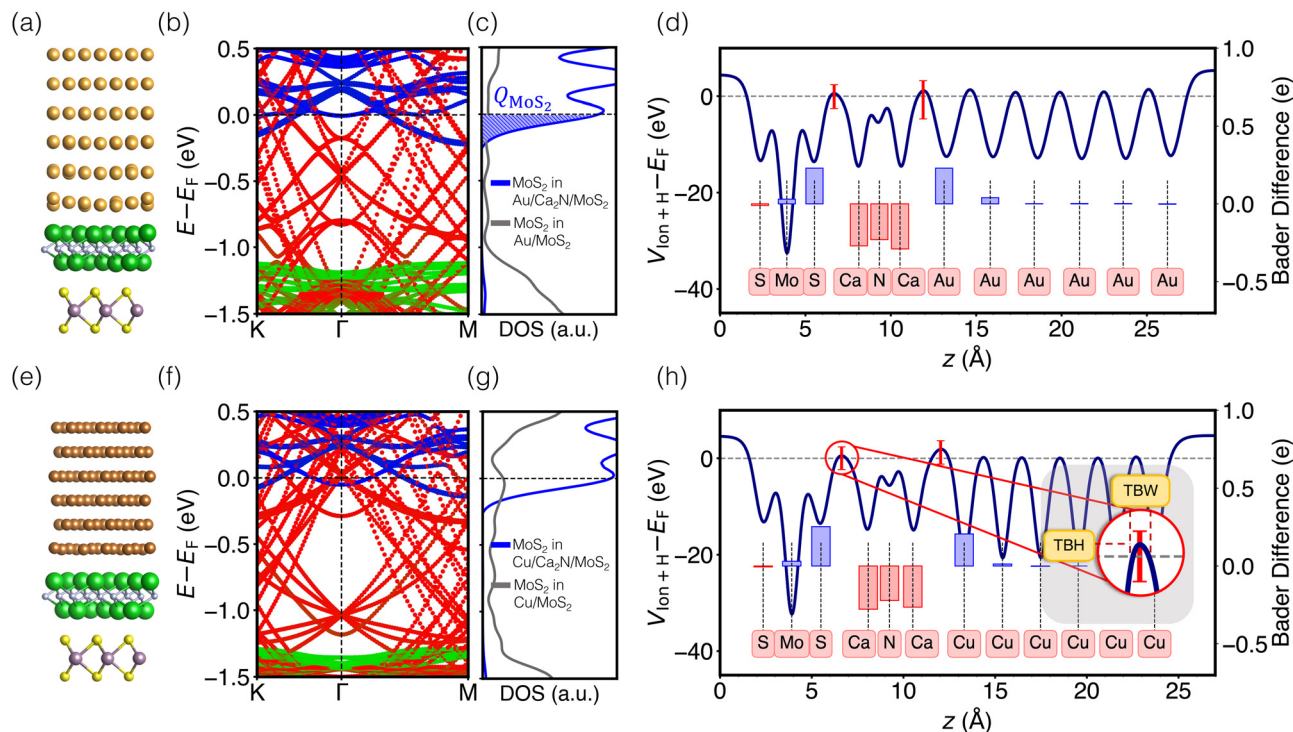


Fig. 2 (a) Optimized structure, (b) material-projected band structure, (c) densities of MoS₂ states per MoS₂ formula unit, and (d) electrostatic potential for the Au/Ca₂N/MoS₂ interface. Panels (e)–(h) present the same quantities for the Cu/Ca₂N/MoS₂ interface. Metal bands are coloured red, Ca₂N bands are coloured green, and MoS₂ bands are coloured blue. The shaded region of the DOS plot corresponds to Q_{MoS_2} . In the potential plots, the red bars indicate the range of values at the maximum point within the interface regions. The results are overlaid with the differences in Bader atomic charges for the heterostructures, relative to the separated materials. Positive (negative) values indicate accumulation (depletion) of electron density.

using density-functional calculations,^{9,11,58,59} so we only briefly discuss some of their most important features. Consistent with previous findings, our results indicate that MoS₂ physisorbs on Au, but chemisorbs on Cu. Carriers at the Au/MoS₂ interface experience a 0.5 eV Schottky barrier and a 3.0 eV tunnelling barrier. For the Cu/MoS₂ interface, electrons that overcome a 0.4 eV Schottky barrier face only a 0.7 eV tunnelling barrier and can travel through some fraction of the interface without tunnelling.

As shown in Fig. 2(b) and (f), Ca₂N insertion shifts the MoS₂ conduction band minimum to -0.2 eV below the Fermi level, meaning that the Schottky barrier is eliminated. This occurs due to electron transfer from the Ca₂N to the MoS₂, which is apparent from comparison with the isolated Ca₂N band structure (Fig. S18–S21, ESI†), as the heterojunction band structures no longer include any Ca₂N surface states near the Fermi level. For both Au- and Cu-based heterojunctions, evaluation of the Bader charges (Table S4, ESI†) shows that *ca.* $0.8 e^-$ per Ca₂N formula unit is transferred, partially to the metal and partially to the MoS₂. The observation of fractional charge transfer is potentially a consequence of delocalization error,^{60,61} which stems from using a generalized gradient approximation (PBE) in this work. Based on a comparison of PBE and HSE06⁶² band structures for the bare Ca₂N/MoS₂ interface without any metal (see Fig. S3, ESI†), it is expected that the charge transfer would increase to nearly a full electron per Ca₂N formula unit if hybrid

functionals were to be employed. However, this increase in charge transfer is not expected to qualitatively impact our results and, beyond being far too computationally expensive with planewave basis sets to apply to the full metal/electrene/MoS₂ interfaces, hybrid functionals are not recommended for metals.

Overall, the findings discussed for Ca₂N above are general based on our analysis of band structures and Bader charges. For all interfaces considered, electrene insertion at the metal/MoS₂ interfaces results in electron transfer from the electrene to the MoS₂ (see Table S4, ESI†), and a concomitant energy lowering of the MoS₂ conduction band edge (see Table S5, ESI†) to below the Fermi level. Thus, for all metal/electrene/MoS₂ heterostructures, electrene insertion negatively dopes the MoS₂, resulting in the removal of the Schottky barrier and unpinning of the Fermi level with no mid-gap states introduced, as shown in Fig. S4–S17 (ESI†).

In addition to providing a nearly barrierless interface to charge carriers, a good metal contact should also result in minimal distortion of the TMDC states, so as to not deteriorate the semiconductor sheet resistivity. Upon contact formation, the MoS₂ band gap is found to change by up to only ± 0.15 eV relative to that of the pristine MoS₂ monolayer (calculated to be 1.72 eV), depending on the choice of the electrene and metal (see Table S5, ESI†). This relatively small modification of the band gap should preserve the resistivity of contacted MoS₂.

3.3 Electrostatic potential profiles and tunneling

While the band structure and DOS plots reveal the promise of the electrone insertion strategy to remove the Schottky barrier, analysis of the electrostatic potential profile of the interface provides information regarding the presence or absence of a tunnelling barrier. Comparison of the Au/Ca₂N/MoS₂ electrostatic potential profile in Fig. 2(d) with the result for the Au/MoS₂ heterojunction in Fig. S1 (ESI[†]) illustrates how Ca₂N insertion reduces the tunneling barrier at the interface to *ca.* 0.5 eV, compared to *ca.* 3.0 eV for Au/MoS₂ without the electrone. Ca₂N insertion at Cu/MoS₂ leads to a similarly small tunneling barrier of *ca.* 0.5 eV.

In the electrostatic potential plots in Fig. 2(d) and (h), the red bars indicate the range of values (in the *x, y* plane) that occur at the *z* position having the highest average potential within the interface regions. The minima of the red bars extend below the Fermi level, indicating Ohmic contact behaviour for some regions of the interface. By projecting the potential profile at the interface into the real-space 2D plane, it is revealed that there is a correlation between the location of the alkaline-earth metal atoms and the Ohmic areas of the interface. For example, for the Au/Ca₂N/MoS₂ heterostructure, the minima of the electrostatic potential occur at the same *x, y* positions as the calcium sites, while the secondary minima correspond to sulfur locations (see Fig. S4(f), ESI[†]). This is the case for all metal/electrone/MoS₂ heterostructures considered, as shown in Fig. S4–S17 (ESI[†]), which allows electrons injected from the metal contact to more easily make their way through to the TMDC.

To ensure that our predictions of high charge transfer and low tunneling barriers are not artifacts of a particular interface geometry, we performed additional calculations for the Cu/Ca₂N/MoS₂ interface, in which the MoS₂ layer was slid over the Ca₂N. Specifically, the MoS₂ layer was shifted uniformly in the *a* and *b* cell directions in increments of 0.2, 0.4, 0.6, and 0.8 times the in-plane lattice vectors and the resulting structures geometry optimized. Relative energies, Fermi energies, charge transfer, and Ohmicity fraction (*vide infra*), as well as tunneling barrier heights and widths, are given in Table S6 (ESI[†]). All show only very minor variations with the fractional sliding coordinate. These results provide confidence that our calculations have used the most stable configurations in terms of MoS₂ sliding, and that our predictions of favourable interface properties are general and not just a feature of one possible geometry.

3.4 Comparison of electrines

Next, we consider trends involving the extent of charge transfer for all seven electrines to assess which is the most promising. Fig. 3(a) and (b) show the correlation of Q_{MoS_2} values with the surface charges ($Q_{\text{M}_2\text{X}}^{\text{surf}}$) and the work functions (Φ_{m}) of various monolayer electrines computed in our previous work.⁴² The surface charges of the monolayer electrines were obtained from the planar average of the integrated local density of states (ILDOS) along the *z*-axis, integrated over energy from the



Fig. 3 (a) Correlation of Q_{MoS_2} and $Q_{\text{M}_2\text{X}}^{\text{surf}}$ in Au and Cu-based interfaces. (b) Correlation of Q_{MoS_2} and the monolayer electrone work function (Φ_{m}). Also shown are correlations between the base-functional (ionic) component of the exfoliation energy and Q_{MoS_2} for Au (c) and Cu (d) interfaces, respectively.

bottom of the surface bands up to the Fermi level. The averaged ILDOS was then integrated along the *z*-axis from the outermost metal atom to the vacuum region (where the ILDOS reaches zero) to give $Q_{\text{M}_2\text{X}}^{\text{surf}}$. The almost linear correlation between Q_{MoS_2} and $Q_{\text{M}_2\text{X}}^{\text{surf}}$ indicates that electrines with higher surface charge densities can donate more charge to the adjacent semiconductor layer.

Fig. 3(b) shows a positive correlation between Q_{MoS_2} and the work function, Φ_{m} , computed⁴² for the monolayer electrone. This result is somewhat counterintuitive, since lower work function materials may be expected to promote more electron transfer to the semiconductor. However, we previously found that electrines with higher work functions have a greater surface charge density⁴² (*i.e.* more charge available in the surface states) that is available to donate. Overall, electrines composed of smaller and chemically harder metal and pnictogen atoms show stronger electrostatic stabilization of the anionic surface charge by the cationic layer. This gives rise to greater surface charge density and tighter binding of the surface states, as quantified by smaller average distances of the surface charge from the alkaline earth metal cations and by larger work functions. Thus, a higher work function implies that the electrone has more charge available to donate to the MoS₂, explaining the greater Q_{MoS_2} values for the corresponding metal/electrone/MoS₂ interfaces.

It should be noted that separate correlations are observed for Au and Cu in Fig. 3(a) and (b) due to the differing

electronegativities of these metals. Since Cu is less electronegative than Au, it will accept less charge. Indeed, the computed Bader charges in Table S4 (ESI†) illustrate that, for any specified electrene, $Q_{\text{Au}}^{\text{Bader}} > Q_{\text{Cu}}^{\text{Bader}}$. As a result, more charge is available to be donated to the MoS₂ layer in the case of a Cu surface, such that Q_{MoS_2} values are higher for Cu/electrene/MoS₂ heterostructures compared to their Au counterparts, regardless of the choice of electrene. Hence, the MoS₂ layer receives a greater fraction of the electrene's surface charge when in contact with the less electronegative metal.

By considering only the base-functional contribution, $E_{\text{base, str}}$, we can also correlate the ionic (*i.e.* non-dispersion) part of the exfoliation energy and the metallization of the semiconductor (Q_{MoS_2}), as shown in Fig. 3(c) and (d). Strong linear correlations are observed for both Au/electrene/MoS₂ and Cu/electrene/MoS₂ heterostructures. Ca₂N and Sr₂N, which possess the largest electrene surface charge densities, also have the highest ionic contributions to the exfoliation energies for these interfaces. The results are in line with the periodic trend in the exfoliation energies of the corresponding bulk electrides.⁴²

Fig. 4 schematically depicts the measures used for assessing contacts in this study, which are also collected in Table S7 (ESI†). Panel (a) shows that the extent of charge transfer, Q_{MoS_2} , is greatest for Ca₂N, followed by Sr₂N and Ba₂N. This is the same trend observed for the monolayer surface charge, as noted above. Values are also shown for the two metal/MoS₂ interfaces for comparison, where there is still a Schottky barrier present

and the Q_{MoS_2} values quantify the charge occupying mid-gap states up to 0.5 eV below the Fermi level. This is in contrast to the metal/electrene/MoS₂ interfaces, where there are no mid-gap states present and Q_{MoS_2} quantifies the metallization of the MoS₂ conduction band.

The Ohmicity of metal/electrene/MoS₂ contacts can be quantified from analysis of the electrostatic potential at the electrene/MoS₂ interface. The Ohmicity fraction is defined as

$$F_{\text{Ohmic}} = \frac{A_{V < E_F}}{A_{\text{Total}}}, \quad (5)$$

where $A_{V < E_F}$ and A_{Total} are the barrier-free contact area and total contact area of the interface, respectively. Values of F_{Ohmic} for all metal/electrene/MoS₂ contacts are summarized in Fig. 4(b). Regardless of metal, Ca₂N shows superior contact behaviour in terms of Ohmicity compared to the other electrines, followed by Sr₂N and Sr₂P. Furthermore, since copper is less electronegative than gold, more charge per formula unit is donated to the MoS₂ layer. This results in greater Q_{MoS_2} and F_{Ohmic} values for the Cu/electrene/MoS₂ contacts, relative to their Au-based counterparts.

Tunnelling characteristics of the contacts are also compared in Fig. 4(c) and (d) for all metal/electrene/MoS₂ configurations. It is imperative that the tunnelling barrier be small enough to maximize the transmission probability. Regardless of the choice of the metal, Ca₂N manifests the lowest TBH and TBW, followed by Sr₂N and Sr₂P. Also, consistent with the F_{Ohmic} results, Cu-based heterostructures show slightly smaller TBWs compared to their Au counterparts. Using the WKB approximation⁶³ (see Section 4 of the ESI†), the calculated tunnelling specific resistivity of Au/MoS₂ is $1.2 \times 10^{-9} \Omega \text{ cm}^2$, with a tunnelling probability of $\leq 1\%$. As such, insertion of monolayer Ca₂N, which raises the tunneling probability to $> 20\%$, lowers the contact resistivity by almost two orders of magnitude. However, since the Cu/MoS₂ tunnelling barrier is negligible, the electrene insertion strategy is most helpful with the elimination of the Schottky barrier in this case. Following the same trends as extracted from the electrostatic potential plots, Ca₂N shows the highest tunnelling probability and lowest tunnelling specific resistivity of all the electrines considered, followed by Sr₂N.

Finally, it is insightful to compare the metal/electrene/MoS₂ tunnelling characteristics with those of the elemental semimetal contacts of MoS₂, which currently set the gold standard of 2D contacts. For example, Sb(0112)/MoS₂ and Sb(0001)/MoS₂ show TBWs of 1.35 Å and 1.39 Å, respectively,¹⁸ whereas the TBWs of {Au,Cu}/Ca₂N/MoS₂ are ~ 0.5 Å. The tunnelling barrier height of {Au,Cu}/Ca₂N/MoS₂ is also significantly lower than that of Sb(0112)/MoS₂ (3.05 eV) and Sb(0001)/MoS₂ (3.13 eV).¹⁸ Our metal/electrene/MoS₂ interfaces also show better tunnelling characteristics than a Bi(0001)/MoS₂ interface with a TBW and TBH of 1.66 Å and 3.6 eV, respectively.¹⁹

4 Conclusion

To summarize our results, considering all four figures of merit employed (Q_{MoS_2} , F_{Ohmic} , TBW, and TBH), we find that Ca₂N is consistently the most favourable electrene material to insert to

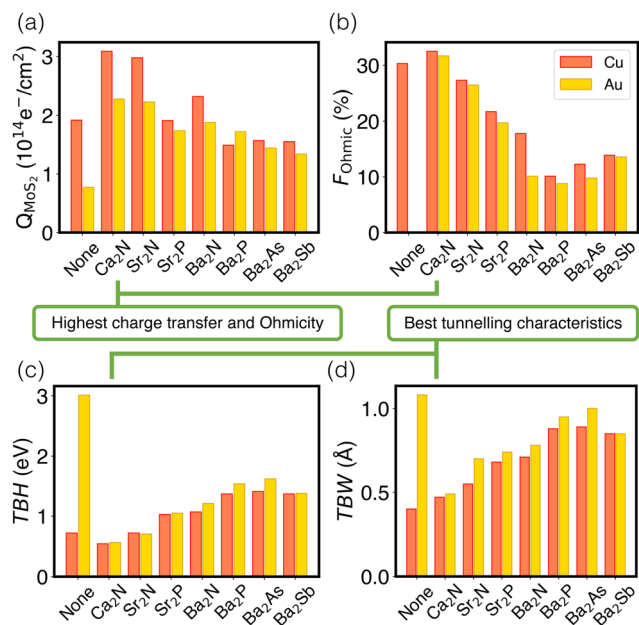


Fig. 4 Bar graphs showing the variation of four figures of merit as a function of the electrene in the metal/electrene/MoS₂ contacts. The individual plots show: (a) the charge transfer to MoS₂, Q_{MoS_2} ; (b) the fraction of the contact area that is Ohmic, F_{Ohmic} ; (c) the tunneling barrier height, TBH; and (d) tunneling barrier width, TBW. Results for the metal/MoS₂ interfaces with no electrene are shown for comparison; note that the Q_{MoS_2} for these two cases results from mid-gap states rather than from the MoS₂ conduction band.

eliminate Schottky and tunnelling barriers across the Au/MoS₂ or Cu/MoS₂ interfaces, owing to its high surface charge density. Sr₂N takes second place for all figures of merit and also appears to be promising, although somewhat less so than Ca₂N. Finally, since Cu is less electronegative than Au, it accepts less charge from the electrene layer, resulting in a more metallized MoS₂, regardless of the choice of electrene. Insertion of electrides appears to be a highly promising strategy to generate n-type (but not p-type) Ohmic contacts since they are excellent reducing agents (but not oxidizing agents). Provided the practical experimental challenges of electrene exfoliation can be overcome,²⁸ including addressing problems with air and water stability,⁶⁴ insertion of 2D electrines should be a promising approach to improve electron transport across metal/TMDC heterojunctions.

Conflicts of interest

There are no conflicts to declare.

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References

- S. Das, H.-Y. Chen, A. V. Penumatcha and J. Appenzeller, *Nano Lett.*, 2013, **13**, 100–105.
- L. Xie, M. Liao, S. Wang, H. Yu, L. Du, J. Tang, J. Zhao, J. Zhang, P. Chen and X. Lu, *et al.*, *Adv. Mater.*, 2017, **29**, 1702522.
- B. Radisavljevic, A. Radenovic, J. Brivio, V. Giacometti and A. Kis, *Nat. Nanotechnol.*, 2011, **6**, 147–150.
- X.-F. Wang, H. Tian, Y. Liu, S. Shen, Z. Yan, N. Deng, Y. Yang and T.-L. Ren, *ACS Nano*, 2019, **13**, 2205–2212.
- A. Allain, J. Kang, K. Banerjee and A. Kis, *Nat. Mater.*, 2015, **14**, 1195–1205.
- W. Liu, J. Kang, W. Cao, D. Sarkar, Y. Khatami, D. Jena and K. Banerjee, *2013 IEEE Intl. Electron Device Mtg.*, 2013, pp. 19–4.
- H. Liu, A. T. Neal and P. D. Ye, *ACS Nano*, 2012, **6**, 8563–8569.
- C. Kim, I. Moon, D. Lee, M. S. Choi, F. Ahmed, S. Nam, Y. Cho, H.-J. Shin, S. Park and W. J. Yoo, *ACS Nano*, 2017, **11**, 1588–1596.
- M. Farmanbar and G. Brocks, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2015, **91**, 161304.
- P. Bampoulis, R. van Bremen, Q. Yao, B. Poelsema, H. J. Zandvliet and K. Sotthewes, *ACS Appl. Mater. Interfaces*, 2017, **9**, 19278–19286.
- C. Gong, L. Colombo, R. M. Wallace and K. Cho, *Nano Lett.*, 2014, **14**, 1714–1720.
- I. Popov, G. Seifert and D. Tománek, *Phys. Rev. Lett.*, 2012, **108**, 156802.
- J. Kang, W. Liu, D. Sarkar, D. Jena and K. Banerjee, *Phys. Rev. X*, 2014, **4**, 031005.
- T. Shen, J.-C. Ren, X. Liu, S. Li and W. Liu, *J. Am. Chem. Soc.*, 2019, **141**, 3110–3115.
- K. Stokbro, M. Engelund and A. Blom, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2012, **85**, 165442.
- J. Kang, D. Sarkar, W. Liu, D. Jena and K. Banerjee, *2013 IEEE Intl. Electron Device Mtg.*, 2012, pp. 17–4.
- L.-Y. Gan, Y.-J. Zhao, D. Huang and U. Schwingenschlögl, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2013, **87**, 245307.
- W. Li, X. Gong, Z. Yu, L. Ma, W. Sun, S. Gao, Ç. Köroğlu, W. Wang, L. Liu and T. Li, *et al.*, *Nature*, 2023, **613**, 274–279.
- P.-C. Shen, C. Su, Y. Lin, A.-S. Chou, C.-C. Cheng, J.-H. Park, M.-H. Chiu, A.-Y. Lu, H.-L. Tang and M. M. Tavakoli, *et al.*, *Nature*, 2021, **593**, 211–217.
- F. Kaadou, J. Maassen and E. R. Johnson, *J. Phys. Chem. C*, 2021, **125**, 11656–11664.
- Y. Liu, J. Guo, Y. Wu, E. Zhu, N. O. Weiss, Q. He, H. Wu, H.-C. Cheng, Y. Xu and I. Shakir, *et al.*, *Nano Lett.*, 2016, **16**, 6337–6342.
- H.-J. Chuang, B. Chamlagain, M. Koehler, M. M. Perera, J. Yan, D. Mandrus, D. Tomanek and Z. Zhou, *Nano Lett.*, 2016, **16**, 1896–1902.
- M. Farmanbar and G. Brocks, *Adv. Electron. Mater.*, 2016, **2**, 1500405.
- Y. Du, L. Yang, J. Zhang, H. Liu, K. Majumdar, P. D. Kirsch and D. Y. Peide, *IEEE Electron Dev. Lett.*, 2014, **35**, 599–601.
- W. S. Leong, X. Luo, Y. Li, K. H. Khoo, S. Y. Quek and J. T. Thong, *ACS Nano*, 2015, **9**, 869–877.
- J. L. Dye, *Acc. Chem. Res.*, 2009, **42**, 1564–1572.
- K. Lee, S. W. Kim, Y. Toda, S. Matsuishi and H. Hosono, *Nature*, 2013, **494**, 336–340.
- D. L. Druffel, K. L. Kuntz, A. H. Woomer, F. M. Alcorn, J. Hu, C. L. Donley and S. C. Warren, *J. Am. Chem. Soc.*, 2016, **138**, 16089–16094.
- D. L. Druffel, A. H. Woomer, K. L. Kuntz, J. T. Pawlik and S. C. Warren, *J. Mater. Chem. C*, 2017, **5**, 11196–11213.
- G. V. Vajenine, A. Grzechnik, K. Syassen, I. Loa, M. Hanfland and A. Simon, *C. R. Chim*, 2005, **8**, 1897–1905.
- T. Kocabas, A. Özden, I. Demiroğlu, D. Çakr and C. Sevik, *J. Phys. Chem. Lett.*, 2018, **9**, 4267–4274.
- S. Guan, S. A. Yang, L. Zhu, J. Hu and Y. Yao, *Sci. Rep.*, 2015, **5**, 1–14.
- S. G. Dale and E. R. Johnson, *Phys. Chem. Chem. Phys.*, 2017, **19**, 27343–27352.
- S. Liu, W. Li, S. W. Kim and J.-H. Choi, *J. Phys. Chem. C*, 2019, **124**, 1398–1404.
- N. E. Brese and M. O'Keeffe, *J. Solid State Chem.*, 1990, **87**, 134–140.
- T. Inoshita, S. Takemoto, T. Tada and H. Hosono, *Phys. Rev. B*, 2017, **95**, 165430.
- W. Ming, M. Yoon, M.-H. Du, K. Lee and S. W. Kim, *J. Am. Chem. Soc.*, 2016, **138**, 15336–15344.
- X.-L. Qiu, J.-F. Zhang, H.-C. Yang, Z.-Y. Lu and K. Liu, *Phys. Rev. B*, 2022, **105**, 165101.
- P. H. Souza, D. Kuritza, J. E. Padilha and R. H. Miwa, *Phys. Rev. B*, 2022, **105**, 235301.
- B. Mortazavi, G. R. Berdiyrov, M. Shahrokhi and T. Rabczuk, *J. Alloys Compd.*, 2018, **739**, 643–652.
- X. Zeng, S. Zhao, Z. Li and J. Yang, *Phys. Rev. B*, 2018, **98**, 155443.

- 42 M. R. Diznab, E. R. Johnson and J. Maassen, *Nanoscale*, 2023, **15**, 12038–12047.
- 43 P. Giannozzi, S. Baroni, N. Bonini, M. Calandra, R. Car, C. Cavazzoni, D. Ceresoli, G. L. Chiarotti, M. Cococcioni and I. Dabo, *et al.*, *J. Phys.: Condens. Matter*, 2009, **21**, 395502.
- 44 P. Giannozzi, O. Andreussi, T. Brumme, O. Bunau, M. B. Nardelli, M. Calandra, R. Car, C. Cavazzoni, D. Ceresoli and M. Cococcioni, *et al.*, *J. Phys.: Condens. Matter*, 2017, **29**, 465901.
- 45 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865.
- 46 P. E. Blöchl, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **50**, 17953.
- 47 G. Kresse and D. Joubert, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1999, **59**, 1758.
- 48 S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, **132**, 154104.
- 49 S. Grimme, S. Ehrlich and L. Goerigk, *J. Comput. Chem.*, 2011, **32**, 1456–1465.
- 50 E. R. Johnson and A. D. Becke, *J. Chem. Phys.*, 2006, **124**, 174104.
- 51 M. Methfessel and A. Paxton, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1989, **40**, 3616.
- 52 G. Henkelman, A. Arnaldsson and H. Jónsson, *Comput. Mater. Sci.*, 2006, **36**, 354–360.
- 53 E. Sanville, S. D. Kenny, R. Smith and G. Henkelman, *J. Comput. Chem.*, 2007, **28**, 899–908.
- 54 W. Tang, E. Sanville and G. Henkelman, *J. Phys.: Condens. Matter*, 2009, **21**, 084204.
- 55 V. Wang, N. Xu, J.-C. Liu, G. Tang and W.-T. Geng, *Comput. Phys. Commun.*, 2021, **267**, 108033.
- 56 S. S. Batsanov, *Inorg. Mater.*, 2001, **37**, 871–885.
- 57 S. Hao, S. Han, J. Zhang, J. Li, N. Wang and X. Li, *ACS Appl. Nano Mater.*, 2023, **6**, 2285–2291.
- 58 M. Farmanbar and G. Brocks, *Phys. Rev. B*, 2016, **93**, 085304.
- 59 B. A. Helfrecht, D. M. Guzman, N. Onofrio and A. H. Strachan, *Phys. Rev. Mater.*, 2017, **1**, 034001.
- 60 K. R. Bryenton, A. A. Adeleke, S. G. Dale and E. R. Johnson, *Wiley Interdiscip. Rev.: Comput. Mol. Sci.*, 2023, **13**, e1631.
- 61 D. Hait and M. Head-Gordon, *J. Phys. Chem. Lett.*, 2018, **9**, 6280–6288.
- 62 A. V. Krukau, O. A. Vydrov, A. F. Izmaylov and G. E. Scuseria, *J. Chem. Phys.*, 2006, **125**, 224106.
- 63 J. G. Simmons, *J. Appl. Phys.*, 1963, **34**, 1793–1803.
- 64 S. H. Kang, D. Thapa, B. Regmi, S. Ren, Y.-M. Kim, S.-G. Kim and S. W. Kim, *J. Am. Chem. Soc.*, 2022, **144**, 4496–4506.