

Cite this: *Org. Biomol. Chem.*, 2021, **19**, 4922

Dibismuthanes in catalysis: from synthesis and characterization to redox behavior towards oxidative cleavage of 1,2-diols†

Marc Magre,[‡] Jennifer Kuziola,[‡] Nils Nöthling and Josep Cornella^{‡*}

A family of aryl dinuclear bismuthane complexes has been successfully synthesized and characterized. The two bismuth centers are bonded to various xanthene-type backbones, which differ in ring-size and flexibility, resulting in complexes with different intramolecular Bi...Bi distances. Moreover, their penta-valent Bi(v) analogues have also been prepared and structurally characterized. Finally, the synergy between bismuth centers in catalysis has been studied by applying dinuclear bismuthanes **5–8** to the catalytic oxidative cleavage of 1,2-diols. Unfortunately, no synergistic effects were observed and the catalytic activities of dinuclear bismuthanes and triphenylbismuth are comparable.

Received 25th February 2021,
Accepted 24th April 2021

DOI: 10.1039/d1ob00367d

rsc.li/obc

Introduction

Organobismuthanes are a class of organometallic reagents where the Bi(III) is connected to C atoms. Such compounds have been largely studied and a wide variety of examples have been reported.¹ In particular, the triarylbismuthanes subclass has attracted the attention of chemists due to their rather high stability and facile preparation. As a result, several monometallic organobismuthanes have been explored as reagents for organic synthesis.² Compared to the vast literature on monometallic triarylbismuthanes, examples of bi- and dimetallic Bi complexes are much rarer and are mainly relegated to the low-valent counterparts Bi(I) and Bi(II). Dimerization in these complexes is highly favored as a result of the unpaired electron in Bi(II)^{3,4} or the stabilization of the highly reactive lone-pair in Bi(I).^{3,5} On the other hand, examples of dimetallic organobismuth complexes have been comparatively much less explored, and only a handful of examples exist in the literature.⁶ Indeed, compared to the lighter counterparts in the group 15 (e.g. N⁷ and P⁸), examples of dimetallic heavy pnictogens are really limited and mainly focus on As and Sb,^{9,10} leaving dimetallic Bi compounds as boutique examples (Fig. 1A).⁶ This can be attributed to the classical ligand redistribution of heteroleptic triarylbismuthanes,¹¹ which poses severe hurdles in the selective synthesis of unsymmetrical dinuclear bismuthanes.

Among the known examples, a common denominator is the high ligand flexibility and the long distances between the two Bi atoms (Fig. 1A). Yet, access to dibismuthanes bearing a more rigid ligand backbone would enable a systematic study of the Bi–Bi distance. Herein we report the synthesis, characterization and structural analysis of a family of new triaryl dinuclear bismuthanes, which bear distinct aromatic backbones on the tether unit. Systematic structural variations on the backbone permit an evaluation on how the rigidity of the tether influences the Bi–Bi distance (Fig. 1B).

Oxidation of the dibismuthanes with SO₂Cl₂ afforded high yields of dimetallic Bi(v) dichloro compounds, a class of compounds that remained elusive to date. The structure of both Bi(III) and Bi(v) dimetallic complexes was elucidated by single crystal X-ray diffraction and their behavior in solution analyzed by NMR. Due to our growing interest in exploring the redox properties of Bi,¹² the catalytic redox properties of these compounds have also been explored in the context of the oxidative cleavage of 1,2-diols.

Results and discussion

Synthesis of heteroleptic triaryldibismuthanes **5–8**

Based on the precedents in the synthesis of diphosphines and the availability of methods to obtain dihalogenated precursors,¹³ we set out to explore the heavy analogs of Xantphos, Homoxantphos, DBFphos and DPEphos, where the P atom has been replaced by Bi. Capitalizing on the important work from Hyvl on the efficient synthesis of heteroleptic triarylbismuthanes,¹⁴ we were able to access the desired dibismuthanes (**5–8**) in good yields (Scheme 1). Dinuclear bismuthanes

Max-Planck-Institut für Kohlenforschung, Mülheim an der Ruhr 45470, Germany.

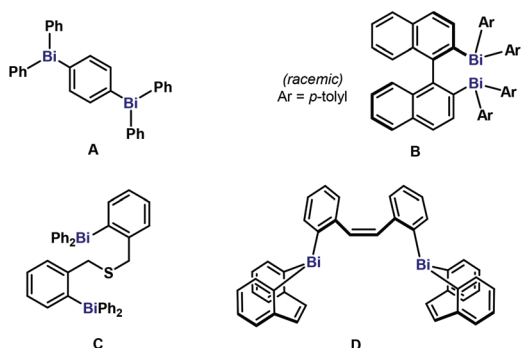
E-mail: cornella@kofo.mpg.de

† Electronic supplementary information (ESI) available: Crystallographic data for compounds **5–12**. CCDC 2063973 (5), 2063975 (6), 2063978 (7), 2063976 (8), 2063977 (9), 2063980 (10), 2063974 (11) and 2063979 (12). For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/d1ob00367d

‡ These authors contributed equally to this work.



A. Examples of dimetallic Bi(III) compounds tethered through a common aryl group



B. This work: Dimetallic Bi(III) and Bi(V) complexes with different Bi-Bi distances

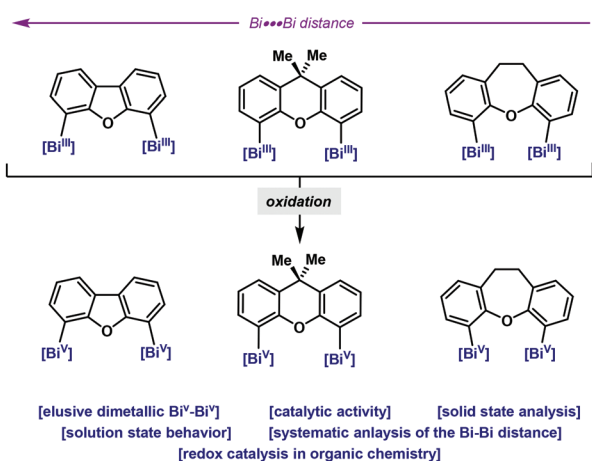
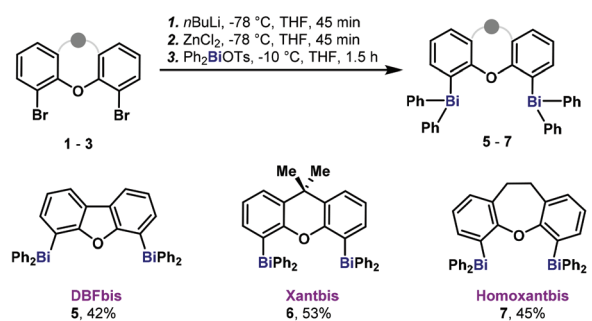
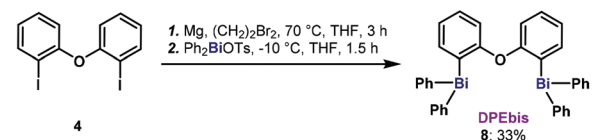


Fig. 1 (A) Representative examples of dimetallic triarylbi-muthanes. (B) Synthesis of dimetallic Bi(III) and Bi(V) complexes.

A. Synthesis of DBFbis, Xantbis and Homoxantbis: rigid ether backbones



B. Synthesis of DPEbis: flexible ether backbone



Scheme 1 Synthesis of dibismuthanes 5–8.

bearing a rigid xanthene-like backbone (5–7) could be isolated in good yields as white solids *via* an organozinc-transmetalation protocol (Scheme 1A). On the other hand, dibismuthane

8, bearing a more flexible backbone, required *in situ* Grignard formation followed by transmetalation to Ph₂BiOTs (Scheme 1B). It is important to mention that 5–8 were reluctant to hydrolysis and were purified by flash chromatography through silica gel. Crystals suitable for X-ray diffraction were obtained for all the compounds prepared. By analogy to the P counterparts and based on the already reported nomenclature for BINABI,^{6b} we named the new dibismuthanes Xantbis, Homoxantbis, DPEbis and DBFbis.

With dibismuthanes 5–8 in hand, we can structurally compare the effect of the ring size of the backbone (5 *vs.* 6 and 7) and the flexibility (5 *vs.* 8). Dibismuthanes 5–8 are stable towards air and moisture, both in solution and in solid state. By means of NMR spectroscopy, we could observe that dinuclear bismuthanes 5–8 present a C₂-symmetric backbone in solution, which suggest both Bi atoms are equivalent.

The solid state structure of dibismuthane 5 (Fig. 2) bears a close resemblance to the monometallic triphenylbismuth, both in the Bi–C distances and the trigonal pyramidal geometry around the Bi atoms.¹⁵ The resemblance in structure between triphenylbismuth and 5 is attributed to the rigidity of the backbone in 5, which limits the torsion and results in a large Bi...Bi distance of 5.544(1) Å. This led us to speculate that both Bi in 5 do not interact with each other, and the complex can be conceived as two independent Bi centers, both electronically and structurally. However, as the Bi get closer in space, this resemblance to the monometallic Ph₃Bi slowly disappears.

For example, the solid state structure of dibismuthane 6 (Fig. 3) shows the expected trigonal pyramidal geometry at both Bi atoms. However, compared to the geometry of triphenylbismuth,¹⁵ dinuclear bismuthane 6 shows a slight distortion comparing both Bi atoms [C2–Bi1–C16: 96.16(19), C2–Bi1–C22: 93.60(19), C16–Bi1–C22: 93.82(18)] compared to C12–Bi2–C28: 90.36(19), C12–Bi2–C34: 94.80(2), C28–Bi2–C34: 93.06(19)], probably as a result of the steric hindrance between both centers. A closer look at the xanthene backbone revealed

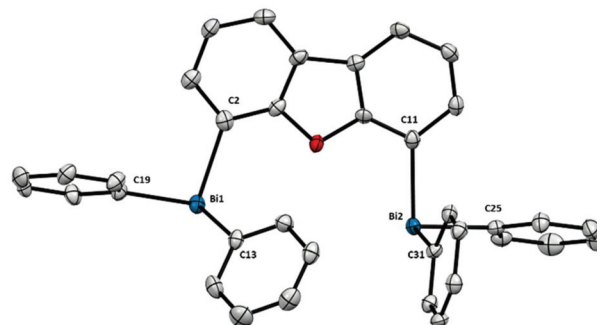


Fig. 2 Structure of 5. Ellipsoids are drawn at the 50% probability level and H atoms are omitted for clarity. Selected bond lengths (Å) and angles (°): Bi1–C2: 2.262(7), Bi1–C13: 2.244(6), Bi1–C19: 2.253(7), Bi2–C11: 2.252(7), Bi2–C25: 2.243(7), Bi2–C31: 2.248(7), C2–Bi1–C13: 93.7(2), C2–Bi1–C19: 96.3(3), C13–Bi1–C19: 93.8(2), C11–Bi2–C25: 95.8(2), C11–Bi2–C31: 94.0(2), C25–Bi2–C31: 96.4(2); Bi1...Bi2: 5.544(1) Å.



using strong oxidants,^{20,21} it is worth mentioning that oxidation of dinuclear bismuthanes have not been reported. Then, we decided to oxidize 5–8 with the commonly utilized SO_2Cl_2 . In all cases, smooth conversion of all dibismuthanes to the parent pentavalent dimetallic 9–12 were obtained (Scheme 2). Due to the high yields obtained, evaporation of the solvent led to pure Bi(v) compounds, whose structures were elucidated by single crystal X-ray analysis.

As depicted in Fig. 6, each Bi atom in 9 adopts a trigonal bipyramidal geometry with the two chloride ligands in apical

position and the aromatic rings in equatorial, in agreement with the apicophilicity of electronegative ligands.²² The structure of pentavalent dibismuthane 9 bears a close resemblance to Ph_3BiCl_2 .²³ Due to the larger distance between both bismuth centers ($\text{Bi}\cdots\text{Bi}$: 6.290(1) Å) the steric pressure between them is highly diminished. This structural feature is reflected on the distances between the Bi and the C in the backbone [Bi1-C2 : 2.184(2) Å and Bi2-C11 : 2.186(3) Å] showing minimal differences among them and to Ph_3BiCl_2 (Fig. 6).

Compared to Ph_3BiCl_2 ,²³ the geometry of complex 10 (Fig. 7) is slightly more distorted, probably due the steric hindrance between both Bi units. This geometry distortion is more evident when comparing both Bi centers. Whereas Bi1 atom adopts a geometry similar to the Ph_3BiCl_2 [C12-Bi1-C34A : 137.74(16)°, C12-Bi1-C28 : 116.05(10)°, C34A-Bi1-C28 : 106.19(16)°], the geometry of Bi2 atom presents a larger degree of distortion [C2-Bi2-C16 : 155.63(9)°, C2-Bi2-C22 : 102.11(9)°, C16-Bi2-C22 : 102.17(9)°]. The large angle observed between C2-Bi2-C16 is the consequence of the effort of Bi2 to accommodate the incoming Cl3 from the neighboring Bi1. The steric hindrance between both Bi centers is also manifested in the distance between the Bi and the carbon at the backbone. For example, whereas Bi1-C12 distance is 2.190(2) Å, the distance of Bi2-C2 is elongated up to 2.212(3) Å. This repulsion is also reflected in the torsion between the Bi2 and the ligand backbone. For example, the C1-C2-Bi2 angle is 128.98(18)°, compared to the C13-C12-Bi1 angle of 124.27(19)°. Hence, the Bi2-



Scheme 2 Synthesis of pentavalent dibismuth(v) compounds 9–12.

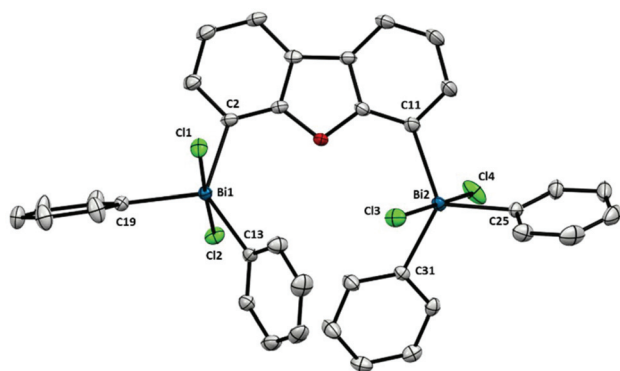


Fig. 6 Structure of 9. Ellipsoids are drawn at the 50% probability level and H atoms are omitted for clarity. Selected bond lengths (Å) and angles (°): Bi1-C2 : 2.184(2), Bi1-C13 : 2.213(2), Bi1-C19 : 2.213(2), Bi1-C11 : 2.5816(6), Bi1-C12 : 2.5827(6), Bi2-C11 : 2.186(3), Bi2-C25 : 2.197(3), Bi2-C31 : 2.194(2), Bi2-Cl3 : 2.5496(6), Bi2-Cl4 : 2.5769(6), C2-Bi1-C13 : 130.24(9), C2-Bi1-C19 : 113.08(9), C13-Bi1-C19 : 116.51(8), Cl1-Bi1-Cl2 : 177.46(2), C2-Bi1-Cl1 : 86.64(7), C2-Bi1-Cl2 : 91.17(7), C13-Bi1-Cl1 : 89.26(7), C13-Bi1-Cl2 : 91.19(7), C19-Bi1-Cl1 : 90.29(7), C19-Bi1-Cl2 : 91.75(7), C11-Bi2-C25 : 112.86(9), C11-Bi2-C31 : 128.63(9), C25-Bi2-C31 : 118.32(9), C13-Bi2-Cl4 : 174.32(2), C11-Bi2-Cl3 : 92.41(7), C11-Bi2-Cl4 : 87.53(7), C25-Bi2-Cl3 : 93.52(7), C25-Bi2-Cl4 : 91.74(7), C31-Bi2-Cl3 : 88.73(6), C31-Bi2-Cl4 : 86.88(6); $\text{Bi1}\cdots\text{Bi2}$: 6.290(1) Å.

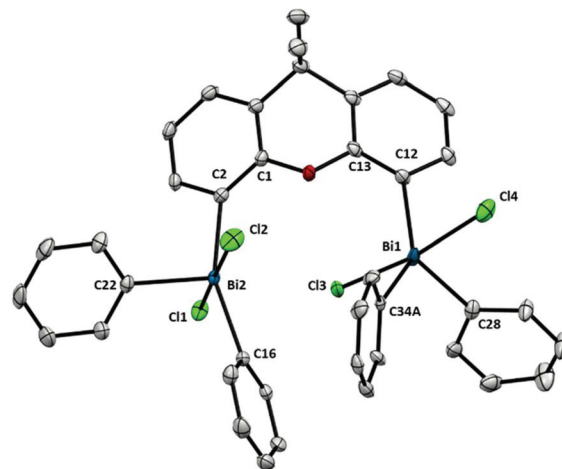
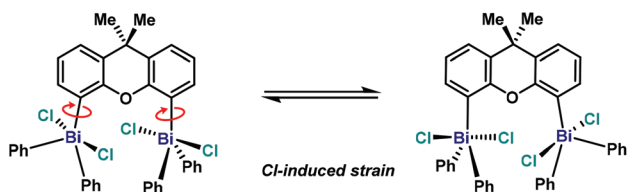


Fig. 7 Structure of 10. Ellipsoids are drawn at the 50% probability level and disordered parts, solvent and H atoms are omitted for clarity. Selected bond lengths (Å) and angles (°): Bi2-C2 : 2.212(3), Bi2-C16 : 2.223(2), Bi2-C22 : 2.212(2), Bi2-Cl2 : 2.5702(7), Bi2-Cl1 : 2.5977(6), Bi1-C12 : 2.190(2), Bi1-C34A : 2.189(5), Bi1-C28 : 2.214(3), Bi1-Cl4 : 2.5825(7), Bi1-Cl3 : 2.5828(7), C2-Bi2-C16 : 155.63(9), C2-Bi2-C22 : 102.11(9), C16-Bi2-C22 : 102.17(9), Cl1-Bi2-Cl2 : 176.01(2), C2-Bi2-Cl2 : 89.03(7), C2-Bi2-Cl1 : 87.53(7), C16-Bi2-Cl2 : 92.14(7), C16-Bi2-Cl1 : 90.16(7), C22-Bi2-Cl2 : 92.19(7), C22-Bi2-Cl1 : 90.51(7), C12-Bi1-C34A : 137.74(16), C12-Bi1-C28 : 116.05(10), C34A-Bi1-C28 : 106.19(16), Cl3-Bi1-Cl4 : 172.07(2), C12-Bi1-Cl4 : 85.19(7), C12-Bi1-Cl3 : 87.13(7), C34A-Bi1-Cl4 : 93.92(16), C34A-Bi1-Cl3 : 90.43(15), C28-Bi1-Cl4 : 93.43(7), C28-Bi1-Cl3 : 91.72(7); $\text{Bi1}\cdots\text{Bi2}$: 5.031(1) Å.



center is out-of-plane due to the steric interaction with the neighboring Bi1. The high repulsion and steric hindrance in **10** results in the increase of the Bi...Bi distance (5.031(1) Å vs. 4.187(1) Å from the trivalent analogue **6**). Further insights into the repulsion and torsion between both Bi centers in **10** can be obtained from the VT-NMR analysis (from 25 °C to -90 °C, see ESI† for details). In solution at ambient temperature, the fluxionality between both bismuth centers is apparent, as **10** reveals itself as a C_2 -symmetric compound, with both Bi being equivalent. However, at lower temperatures, the two Bi atoms are no longer equivalent, consistent with the solid-state analysis of **10**. The solution behavior of **10** can be explained by a chloride-induced steric repulsion between the two TBP bismuth centers (Scheme 3). At low temperature, the Cl-induced structural strain can be observed.

As represented in Fig. 8, the presence of bulky BiPh₂Cl₂-moieties in **11** has an influence on the backbone, as observed also in its analogue Bi(III) **7** (Fig. 4). In agreement with the increase of steric hindrance of the Bi-centers (-BiPh₂Cl₂ in **11** vs. -BiPh₂ **7**) the torsion in the ligand backbone is more accen-



Scheme 3 Proposed structural distortion observed in **10** at low temperature.

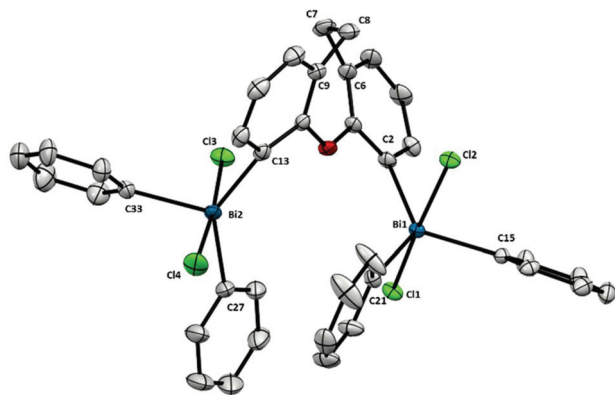


Fig. 8 Structure of **11**. Ellipsoids are drawn at the 50% probability level and solvent and H atoms are omitted for clarity. Selected bond lengths (Å) and angles (°): Bi1–C2: 2.192(2), Bi1–C15: 2.200(2), Bi1–C21: 2.212(2), Bi1–Cl1: 2.5972(6), Bi1–Cl2: 2.5962(6), Bi2–C13: 2.218(3), Bi2–C27: 2.216(3), Bi2–C33: 2.206(3), Bi2–Cl3: 2.5881(7), Bi2–Cl4: 2.5856(7), C2–Bi1–C15: 119.38(9), C2–Bi1–C21: 127.71(9), C15–Bi1–C21: 112.80(9), Cl1–Bi1–Cl2: 170.725(19), C2–Bi1–Cl1: 88.91(7), C2–Bi1–Cl2: 85.48(7), C15–Bi1–Cl1: 88.26(6), C15–Bi1–Cl2: 88.06(6), C21–Bi1–Cl1: 96.08(7), C21–Bi1–Cl2: 93.19(7), C13–Bi2–C27: 138.46(10), C13–Bi2–C33: 110.94(10), C27–Bi2–C33: 110.60(11), Cl3–Bi2–Cl4: 175.59(2), C13–Bi2–Cl3: 91.39(7), C13–Bi2–Cl4: 89.29(7), C27–Bi2–Cl3: 90.69(9), C27–Bi2–Cl4: 91.67(9), C33–Bi2–Cl3: 87.62(7), C33–Bi2–Cl4: 88.07(7), C9–C8–C7: 109.6(2), C6–C7–C8: 110.1(2); Bi1...Bi2: 5.195(1) Å.

tuated, reflected in the angle between the planes of the benzo groups (in **11** is 46.39° and in **7** is 56.09°). This steric interaction between both bismuth centers is also shown in the distance between the Bi and the carbon at the backbone. For example, although attached to the same aryl unit, the Bi1–C2 distance is 2.192(2) Å whereas the distance of Bi2–C13 is elongated up to 2.218(3) Å. This Bi–C elongation is also observed in its 6-membered ring analogue **10** (Fig. 7) but not in the more rigid 5-member ring **9** (Fig. 6). The high repulsion and steric hindrance in **11** results in the increase of the Bi...Bi distance [5.195(1) Å vs. 3.807(1) Å from the trivalent analogue **7**]. Due to the higher degree of backbone flexibility in **11** compared to **10**, the Bi...Bi distance in **11** is slightly longer than in **10**, an opposite structural feature observed in the less sterically congested trivalent analogues (**7** and **6**).

The structure of pentavalent dibismuthane **12** (Fig. 9) shows similarities to compound **9** and Ph₃BiCl₂.²³ Therefore, each bismuth center adopts a slightly distorted trigonal bipyramidal geometry with the Cl in the apical positions. Similar to the parent Bi(III) (**8**), the flexibility of the ligand backbone allows to diminish the steric constraints between both bismuth centers. Indeed, the two aryl groups from the backbone are placed almost perpendicular to each other [86.28(1)°]. Nevertheless, despite the free rotation of the ligand, the distance between both bismuth atoms is 6.288(1) Å, very similar to that found for **9**, which contains a rigid 5-member ring scaffold in the backbone.

Catalytic oxidative cleavage of 1,2-diols

We have seen that dinuclear bismuthanes **5–8** show excellent air and moisture stability and high reactivity with strong oxi-

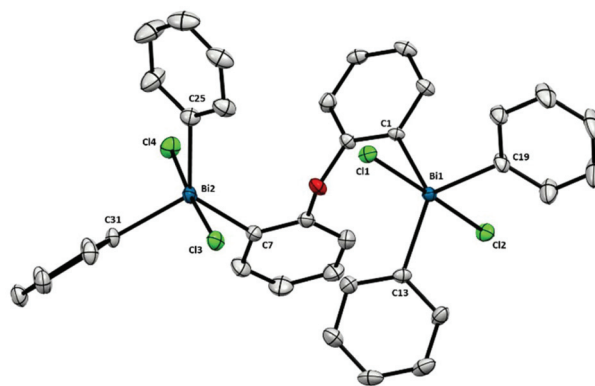


Fig. 9 Structure of **12**. Ellipsoids are drawn at the 50% probability level and H atoms are omitted for clarity. Selected bond lengths (Å) and angles (°): Bi1–C1: 2.184(3), Bi1–C13: 2.199(3), Bi1–C19: 2.204(3), Bi1–Cl1: 2.5862(8), Bi1–Cl2: 2.5892(8), Bi2–C7: 2.189(3), Bi2–C25: 2.205(3), Bi2–C31: 2.214(3), Bi2–Cl3: 2.5677(8), Bi2–Cl4: 2.6191(8), C1–Bi1–C13: 130.36(11), C1–Bi1–C19: 109.56(12), C13–Bi1–C19: 120.09(12), Cl1–Bi1–Cl2: 175.42(3), C1–Bi1–Cl1: 86.12(8), C1–Bi1–Cl2: 90.01(8), C13–Bi1–Cl1: 92.04(8), C13–Bi1–Cl2: 88.59(8), C19–Bi1–Cl1: 92.47(9), C19–Bi1–Cl2: 91.13(9), C7–Bi2–C25: 128.79(12), C7–Bi2–C31: 114.23(12), C25–Bi2–C31: 116.77(12), C13–Bi2–Cl4: 176.06(3), C7–Bi2–Cl3: 93.06(9), C7–Bi2–Cl4: 83.75(9), C25–Bi2–Cl3: 89.99(9), C25–Bi2–Cl4: 90.22(9), C31–Bi2–Cl3: 91.45(8), C31–Bi2–Cl4: 91.96(9); Bi1...Bi2: 6.288(1) Å.



- (e) K. Materne, S. Hoof, N. Frank, C. Herwig and C. Limberg, *Organometallics*, 2017, **36**, 4891–4895.
- 11 (a) F. Challenger, *J. Chem. Soc., Trans.*, 1914, **105**, 2210–2218; (b) H. Gilman, H. L. Yablunsky and A. C. Svigoon, *J. Am. Chem. Soc.*, 1939, **61**, 1170–1172; (c) H. Gilman and H. L. Yablunsky, *J. Am. Chem. Soc.*, 1941, **63**, 207–211.
- 12 (a) F. Wang, O. Planas and J. Cornella, *J. Am. Chem. Soc.*, 2019, **141**, 4235–4240; (b) O. Planas, F. Wang, M. Leutzsch and J. Cornella, *Science*, 2020, **367**, 313–317; (c) O. Planas, V. Peciukenas and J. Cornella, *J. Am. Chem. Soc.*, 2020, **142**, 11382–11387; (d) Y. Pang, M. Leutzsch, N. Nöthling and J. Cornella, *J. Am. Chem. Soc.*, 2020, **142**, 19473–19479.
- 13 (a) P. W. N. M. van Leeuwen, P. C. J. Kamer, J. N. H. Reek and P. Dierkes, *Chem. Rev.*, 2000, **100**, 2741–2770; (b) M.-N. Birkholz, Z. Freixa and P. W. N. M. van Leeuwen, *Chem. Soc. Rev.*, 2009, **38**, 1099–1118.
- 14 T. Louis-Goff, A. L. Rheingold and J. Hyvl, *Organometallics*, 2020, **39**, 778–782.
- 15 D. M. Hawley and G. Ferguson, *J. Chem. Soc. A*, 1968, 2059–2063.
- 16 (a) M. Kranenburg, Y. E. M. van der Burgt, P. C. J. Kamer, P. W. N. M. van Leeuwen, K. Goubitz and J. Fraanje, *Organometallics*, 1995, **14**, 3081–3089; (b) S. Hillebrand, J. Bruckmann, C. Krüger and M. W. Haenel, *Tetrahedron Lett.*, 1995, **36**, 75–78.
- 17 L. Ouyang, C. Qi, H. He, Y. Peng, W. Xiong, Y. Ren and H. Jiang, *J. Org. Chem.*, 2016, **81**, 912–919.
- 18 The intramolecular Bi...Bi distance in **B** (ref. 6b) is 4.690 Å, in dinuclear bismuthane **C** (ref. 6c) is 6.368 Å, and in dinuclear bismuthane **D** (ref. 6d) is 4.435 Å. The intermolecular Bi...Bi distance in **A–D** is 5.559–13.058 Å (ref. 6).
- For examples of short intermolecular Bi...Bi distances see: (a) A. M. Toma, A. Pop, A. Silvestru, T. Rüffer, H. Lang and M. Mehring, *Dalton Trans.*, 2017, **46**, 3953–3962; (b) T. Dunaj, K. Dollberg, C. Ritter, F. Dankert and C. von Hänisch, *Eur. J. Inorg. Chem.*, 2021, **2021**, 870–878; (c) S. Schulz, A. Kuczkowski, D. Bläser, C. Wölper, G. Jansen and R. Haack, *Organometallics*, 2013, **32**, 5445–5450; (d) M. Hébert, P. Petiot, E. Benoit, J. Dansereau, T. Ahmad, A. Le Roch, X. Ottenwaelder and A. Gagnon, *J. Org. Chem.*, 2016, **81**, 5401–5416.
- 19 A. Pintado-Alba, H. de la Riva, M. Nieuwhuyzen, D. Bautista, P. R. Raithby, H. A. Sparkes, S. J. Teat, J. M. López-de-Luzuriaga and M. C. Lagunas, *Dalton Trans.*, 2004, 3459–3467.
- 20 (a) S. A. Lermontov, I. M. Rakov, N. S. Zefirov and P. J. Stang, *Tetrahedron Lett.*, 1996, **37**, 4051–4054; (b) T. Ooi, R. Goto and K. Maruoka, *J. Am. Chem. Soc.*, 2003, **125**, 10494–10495.
- 21 D. H. R. Barton, N. Y. Bhatnagar, J.-P. Finet and W. B. Motherwell, *Tetrahedron*, 1986, **42**, 3111–3122.
- 22 R. S. McDowell and A. Streitwieser Jr., *J. Am. Chem. Soc.*, 1985, **107**, 5849–5855.
- 23 D. M. Hawley and G. Ferguson, *J. Chem. Soc. A*, 1968, 2539–2543.
- 24 (a) D. H. R. Barton, W. B. Motherwell and A. Stobie, *J. Chem. Soc., Chem. Commun.*, 1981, 1232–1233; (b) D. H. R. Barton, J.-P. Finet, W. B. Motherwell and C. Pichon, *Tetrahedron*, 1986, **42**, 5627–5636.
- 25 For a comprehensive review on oxidative cleavage of glycols, see: A.-K. C. Schmidt and C. B. W. Stark, *Synthesis*, 2014, 3283–3308.

