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Detection and identification of designer drugs by nanoparticle-based NMR chemosensing†

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Properly designed monolayer-protected nanoparticles (2 nm core diameter) can be used as nanoreceptors for selective detection and identification of phenethylamine derivatives (designer drugs) in water. The molecular recognition mechanism is driven by the combination of electrostatic and hydrophobic interactions within the coating monolayer. Each nanoparticle can bind up to 30–40 analyte molecules. The affinity constants range from 10^5 to 10^6 M⁻¹ and are modulated by the hydrophobicity of the aromatic moiety in the substrate. Detection of drug candidates (such as amphetamines and methamphetamines) is performed by using magnetization (NOE) or saturation (STD) transfer NMR experiments. In this way, the NMR spectrum of the drug is isolated from that of the mixture, allowing broad-class multianalyte detection and even identification of unknowns. The introduction of a dimethylsilane moiety in the coating monolayer allows performing STD experiments in complex mixtures. In this way, a detection limit of 30 µM is reached with standard instruments.

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Introduction

“Designer drugs” are psychotropic substances usually synthesized by modification of existing drugs.¹ Being new substances, they elude both classification as illicit substances and identification by standard analytical methods. For this reason, and because the emergence of such drugs has dramatically increased in the past few years, the threat they pose to society is relevant.² In 2014–2017, 265 new psychoactive substances were reported to the EU’s early warning system.³ Two of the most relevant families of designer drugs are amphetamines (2–3, Fig. 1) and cathinones (4), which share a quite similar chemical structure based on the 2-phenylethylamine (phenethylamine, 1, Fig. 1) backbone.

Detection and identification of these substances, which lack certified analytical standards, is a major challenge for forensic and customs laboratories.⁴ Standard procedures for the identification of new substances require time-consuming isolation and careful identification by NMR spectroscopy.^{4,5} On-site detection kits are commercially available for early screening and even “in-home” quality control. These are based on chromogenic chemical reactions⁶ or antibody-based immunoassays.⁷ However, they provide qualitative results which need

validation by more sophisticated analysis and may easily fail in identifying new substances.⁷

Chemosensors provide an alternative detection approach that can be applied both to “on field” testing and quantitative determination.⁸ In addition, they can usually operate directly on the sample under analysis, without the need for any pre-treatment.⁹ A few supramolecular chemosensors have recently been proposed for the detection of drugs of abuse.¹⁰ Most of the systems reported are based on cavitand hosts, such as cucurbituril or resorcinarene tetraphosphonate derivatives, capable of recognizing amphiphilic organic cations in water.^{11–13} Signal generation is obtained by the alteration of the properties of receptor-conjugated fluorescent dyes^{11a,b,13b} or materials, such as organic field-effect transistors,¹² or cantilevers of an atomic

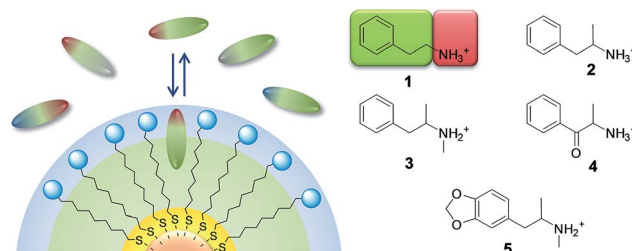


Fig. 1 Left: recognition of amphiphilic organic ions by MPGNs coated with amphiphilic thiols with complementary charge. Right: chemical structure of representative psychoactive molecules of the 2-phenylethylamine (1) family: amphetamine (2), methamphetamine (3), cathinone (4), and MDMA (5) (color code: blue, negative charge; red, positive charge; grey, neutral hydrophilic; green, neutral hydrophobic).

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S2 and **S3** a phenyl moiety was added *via* amide linkage. In **S4** a dimethylsilane group was introduced for STD-NMR purpose, since it allows selective NP saturation without overlapping with the analyte signals. **S1–S4** were used to synthesize MPGNs (**S1/S4**-AuNP) with an average gold core diameter of 1.6 ± 0.3 nm. The average molecular formula is $\text{Au}_{140}\text{SR}_{50}$ (see the ESI†). In all the cases the resulting nanoparticles were well soluble in water.

Nanoreceptor molecular recognition ability

We selected the simplest **S1**-AuNP as the front-line candidate to test the ability of anionic nanoparticles to detect phenethylamine derivatives. The nanoparticles, at 15 μM concentration (corresponding to 1 mM concentration of coating molecules), were mixed with 2.0 mM phenethylamine (**1**) in deuterated water buffered at pD 7.0. Then, an “NOE pumping-CPMGz” experiment was performed (Fig. 3).^{18b,24} In this experiment, a diffusion filter, an NOE sequence and a T_2 relaxation filter are concatenated before recording the signal. The diffusion filter dephases the magnetization of all the fast diffusing species in the sample but not that of the slow-diffusing nanoparticles (Fig. 3b). During the following NOE step, a part of the residual magnetization is transferred from the nanoparticles to any molecular species interacting (in a fast exchange regime) with them (Fig. 3c). Eventually, the T_2 relaxation filter attenuates the signal of the slowly tumbling and conformationally hindered nanoparticle-coating ligands.²⁵ As a result, only the signals of the molecules recognized by the nanoparticle receptor are present in the spectrum (Fig. 3d).

The results reported in Fig. 3 confirmed the ability of **S1**-AuNP to recognize and detect phenethylamine. The four signals belonging to the five groups of magnetically equivalent protons in the substrate (two of them overlap to form a multiplet at 7.3 ppm) are the sole signals present in the NOE pumping-CPMGz spectrum and allow easy identification of the detected

molecule. Interestingly, signals of other species present in the sample, such as water, residual solvents, and in particular the HEPES buffer, are not present in the final spectrum. Consequently, the two triplets at 2.9 and 3.2 ppm arising from the ethyl residue of **1**, which are not visible in the ^1H -NMR spectrum due to the overlap of the broad and intense HEPES signals, are clearly extracted in the final spectrum.

The ability of **S1**-AuNP nanoparticles to recognize the cationic amphiphilic structure of many designer drugs was confirmed by investigating the detection of other molecules with a similar structure (Fig. 4), including some neurotransmitters and drug precursors. Phenethylamine derivatives such as *N*-methylphenethylamine (**6**), 4-fluoro-phenethylamine (**7**), tyramine (**8**), dopamine (**9**), 3-methoxyphenethylamine (**10**), 4-methoxyphenethylamine (**11**), 3,4-methylenedioxyphenethylamine (**12**), serotonin (**13**), 4-nitro-phenethylamine (**14**) and ephedrine (**15**) were all detected and identified from their distinctive ^1H -NMR signals. On the other hand, molecules with similar structures but devoid of the cationic head group, such as phenylalanine (**16**), as well as hydrophilic organic molecules, such as HEPES, glucose and lactose (not shown), did not produce any signal in the NOE pumping experiments.

Surprisingly, phloretic acid (**17**) is detected, even if with lower sensitivity. This suggests that the electrostatic repulsion between the anionic headgroups of the nanoparticles and the carboxylate moiety of the amphiphilic substrate may not be sufficient to completely prevent the interaction. Still, even if the NMR spectrum of **17** is quite similar to that of the corresponding phenethylamine **8**, identification of the compound as a carboxylate is quite simple, as the signals of the aliphatic methylenes shift from the region typical of phenethylamines (2.7–3.3 ppm) to that of phenethylcarboxylates (2.3–2.7 ppm).

To gain more insight into the recognition properties of **S1**-AuNP, we measured their affinity for different analytes by fluorescence titrations, taking advantage of the luminescence properties of molecules **7–13** and **17–19**. The experiments were performed by adding increasing amounts of analytes to a 1.4 μM (0.1 mM concentration of coating molecules) solution of **S1**-AuNP in water at pH 7 (HEPES buffer 10 mM).²⁶ Gold nanoparticles effectively quench the emission of dyes bound to the coating monolayer. When the affinity of the analytes for **S1**-AuNP was high enough, an initial quenching of the emission was observed (indicating the binding of the dye to the AuNPs) followed by a linear emission increase after saturation was reached (see ESI, Fig. S17 and S18†).

Emission intensity *versus* analyte concentration plots were fitted with a 1 : 1 binding model. It is important to point out that this model assumes that multiple, equivalent and independent binding sites are present in the nanoparticle-coating monolayer. The results of the fittings are summarized in Table 1. Binding constant (K) values in the range 1×10^5 to $1.3 \times 10^6 \text{ M}^{-1}$ were found. Such values are consistent with those previously measured for the interaction of cationic nanoparticles with organic anions.^{18c} Remarkably, a good linear correlation (Fig. 5) was found between $\log(K)$ values of the different analytes and their *n*-octanol/water partition coefficients computationally predicted at pH 7.4 ($\log D$). This



Fig. 3 (a) ^1H -NMR spectrum of phenethylamine (**1**) and **S1**-AuNP in HEPES buffered D_2O ; (b) diffusion filter spectrum of the same sample; (c) NOE-pumping spectrum of the same sample (3072 scans, 4 h). (d) NOE-pumping-CPMGz spectrum of the same sample (60 ms CPMGz filter, 3072 scans, 4 h). Conditions: $[\text{1}] = 2.0 \text{ mM}$, $[\text{S1-AuNP}] = 15 \mu\text{M}$, HEPES buffer 10 mM, and pD = 7.0.



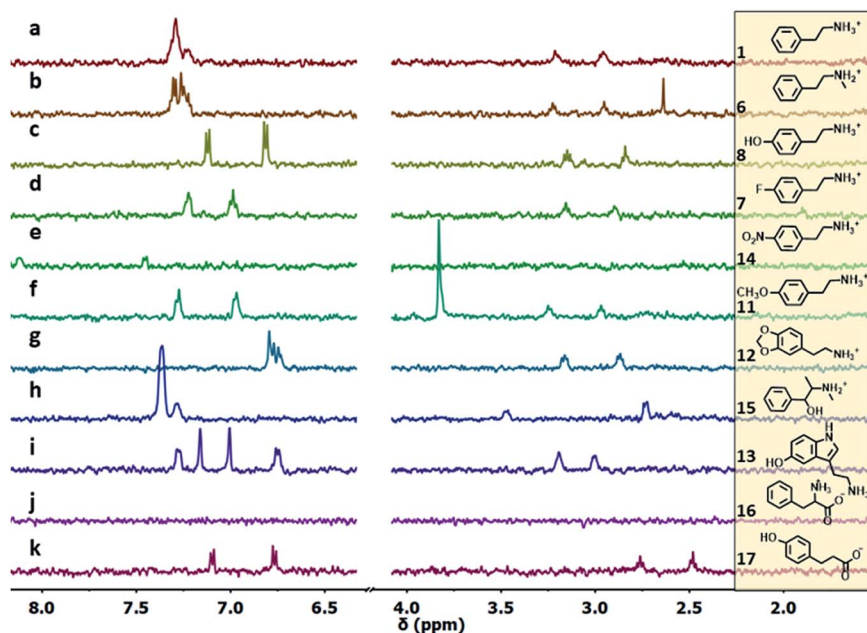


Fig. 4 ^1H -NMR NOE pumping-CPMGz sub-spectra (3072 scans, 4 h) of AuNP-S1 (14 μM in D_2O), HEPES buffer (10.0 mM) and different analytes (2 mM): (a)–(k). For 4-nitrophenethylamine (e), the NOE pumping spectrum is shown (same acquisition parameters). For 12 (g) and 15 (h), the signals respectively at 5.92, 5.11 and 1.04 ppm, present in the spectrum, are outside the spectral window shown for clarity (full spectra are reported in Fig. S23†).

Table 1 Binding parameters of analytes 1 and 7–19 to S1-AuNPs in water (values in red are estimated, see note c). Errors reported are the standard deviations obtained from the fittings^a

Analyte	K , M^{-1}	[Binding sites], M	$\log D$ (pH 7.4) ^b
1	$3.6 \times 10^{5c,d}$	— ^c	−0.84
7	$[(7.9 \pm 0.8) \times 10^5]^e$	$[(3.8 \pm 0.1) \times 10^{-5}]^e$	−0.85
8	$[(4.1 \pm 0.4) \times 10^5]^e$	$[(4.2 \pm 0.1) \times 10^{-5}]^e$	−2.01
9	$(1.3 \pm 0.2) \times 10^5$	$(3.5 \pm 0.2) \times 10^{-5}$	−2.18
10	$(1.2 \pm 0.2) \times 10^5$	$(2.8 \pm 0.2) \times 10^{-5}$	−2.18
11	$(4.1 \pm 0.4) \times 10^5$	$(4.2 \pm 0.1) \times 10^{-5}$	−1.04
12	$(4.8 \pm 0.5) \times 10^5$	$(3.9 \pm 0.1) \times 10^{-5}$	−1.04
13	$(3.9 \pm 0.3) \times 10^5$	$(4.1 \pm 0.1) \times 10^{-5}$	−0.87
14	$(2.7 \pm 0.3) \times 10^5$	$(2.9 \pm 0.1) \times 10^{-5}$	−1.71
15	$4.6 \times 10^{5c,d}$	— ^c	−0.43
16	$3.7 \times 10^{5c,d}$	— ^c	−0.75
17	— ^c	— ^c	−1.46
18	— ^f	— ^f	−1.24
19	$(6.1 \pm 1.4) \times 10^5$	$(3.3 \pm 0.1) \times 10^{-5}$	0.55
	$(2.2 \pm 0.1) \times 10^6$	$(5.5 \pm 0.1) \times 10^{-5}$	2.16

^a $[\text{1-AuNP}] = 10 \times 10^{-5} \text{ M}$, pH = 7.0 (HEPES buffer 10 mM). ^b Predicted with the ACD/Labs Percepta module, <http://www.chemspider.com>. ^c Not measurable as the substrate is not fluorescent. ^d Estimated from the plot in Fig. 5. ^e Values obtained by 9-displacement titration. ^f No binding observed.

confirms that the interaction between the nanoparticles and the analytes is modulated by the accommodation of the aromatic moiety in the hydrophobic portion of the monolayer (being the ion-pairing headgroup interaction similar in all the cases). The relatively small slope (0.26) suggests that hydrophobic “stabilization”²⁷ provided by the alkyl chains in the monolayer is less

effective than *n*-octanol solvation. The linear plot of Fig. 5 may in principle be used to extrapolate affinities for non-fluorescent analytes as 1, 14 and 15 (Table 1). To verify the reliability of the correlation and of the estimated data, we investigated the affinity of these analytes for S1-AuNPs by diffusion-ordered NMR spectroscopy (DOSY). The success of these experiments is based on the ability of the bulky nanoparticles to effectively perturb the apparent diffusion coefficients of the interacting analytes. In particular, being in a fast exchange regime on the NMR timescale, the apparent diffusion coefficient of each analyte will be the

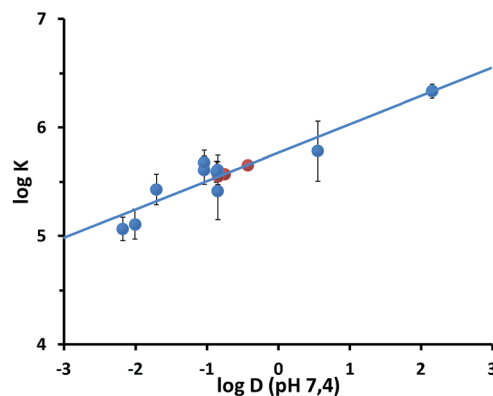


Fig. 5 Plot of the $\log K$ vs. $\log D$ (pH = 7.4) values relative to the binding of the luminescent analytes 8–13 and 17–19 to S1-AuNP. The lines represent the linear fit of the data ($R = 0.885$). Red circles report the affinity values estimated for substrates 1, 14 and 15 on the basis of their $\log D$ values and the fitting parameters. The error bars reported represent the confidence intervals (3σ) calculated from standard deviations reported in Table 1.



average between those of the free species and of the nanoparticles, weighted on the relative populations of bound and unbound analyte.^{18a,28} We selected phenethylamine (7), 4-nitrophenethylamine (14) and, as a control, tyramine (8), whose affinity for S1-AuNP had been measured by fluorescence titration.

According to the estimated values of the binding constants, affinity of the three analytes for S1-AuNP should follow the order $14 > 1 > 7$ (Table 1). When the mixture of the three molecules, each at 0.5 mM concentration, was analysed by DOSY-NMR in the presence of S1-AuNPs (45 μ M), the three components were nicely separated according to their apparent diffusion coefficients, which decreased in perfect agreement with the predicted affinity order (Fig. 6). Nicely enough, the DOSY experiment reported in Fig. 6 also proves that S1-AuNP allows a multianalyte detection by solution-state "chromatographic NMR" as well.²⁹

The number of binding sites in S1-AuNPs estimated with the luminescence titrations is in most cases between 30% and 40% of the number of coating thiols (Table 1), suggesting that each binding pocket in the monolayer is formed by about 3 thiols and that each nanoparticle can bind between 20 and 30 analyte molecules.

A displacement titration performed with 7 (Table 1) in the presence of 9 provided a number of binding sites similar to that obtained with the direct titration. This suggests that the incoming guest molecules occupy the same binding pockets of the leaving ones. The affinity constant measured with the displacement experiment is slightly larger than that obtained with direct titration.

Affinity tuning

NOE-pumping NMR and fluorescence titrations on selected analytes were then used to investigate the different recognition abilities of S1-, S2-, S3- and S4-AuNPs (Fig. 7). Modification of the nanoparticle coating thiols by introducing either an aromatic or

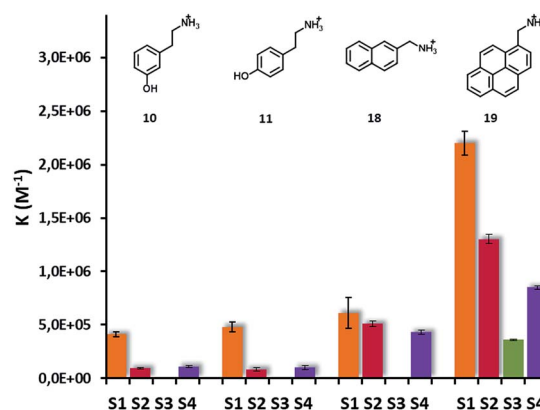


Fig. 7 Graphical representation of binding constants of analytes 10, 11, 18, 19 and S1-, S2-, S3- and S4-AuNPs. Values of association constants and binding sites with their uncertainties are reported in Table S1 in the ESI.†

a dimethylsilane moiety resulted in a reduction of the affinity for the analytes with respect to S1-AuNP. $\log K$ versus $\log D$ plots (see ESI, Fig. S21†) remain linear, indicating that also with these nanoparticles the ion pairing and the hydrophobic interaction are the main, if not the sole, interactions at play. In the case of S4-AuNP, the slope of the plot (0.29) is very close to that obtained with S1-AuNP (0.26), indicating a similar sensitivity to substrate hydrophobicity and suggesting that the affinity decrease might arise from a weaker ion pairing interaction. The electron-donating propensity of the silicon atom may reduce the charge density of the sulfonate. In the case of S2-AuNP the slope is 0.37, indicating a stronger sensitivity to substrate hydrophobicity. Still, affinities measured are always smaller than those of S1-AuNPs. The aromatic residue produces apparently two different effects: on one hand it decreases, *via* delocalization, the charge density at the sulfonate; on the other hand, it increases the stabilization of the hydrophobic portion of the substrates. The reduction of ion pairing ability is presumably stronger with S3-AuNP due to the presence of the electron-withdrawing carbonyl group in the *para* position. Indeed, affinity generally decreases below the threshold that can be detected by fluorescence titrations, which is around $5 \times 10^4 \text{ M}^{-1}$ (see ESI, Fig. S19†). Only with substrate 19, where the strong hydrophobic interaction maintains the affinity for the nanoparticles in the measurable interval, we could determine the binding constant for S3-AuNPs, which is 6-fold smaller than that for S1-AuNP.

The affinity reduction observed has a small effect on the sensitivity of the NOE pumping experiments (see the ESI†). Indeed, the detection limit remains at around 0.5 mM concentration for all the nanoparticles. On the other hand, the binding selectivity is substantially improved. Signals of phloric acid (17), which was detected under these conditions with S1-AuNP (Fig. 4), are not present in the NOE pumping spectra with S2-AuNP (see ESI, Fig. S24†). Clearly the affinity decrease brought about by thiol S2 on the likely already small binding constant of 17 prevents an effective interaction with the nanoparticles.

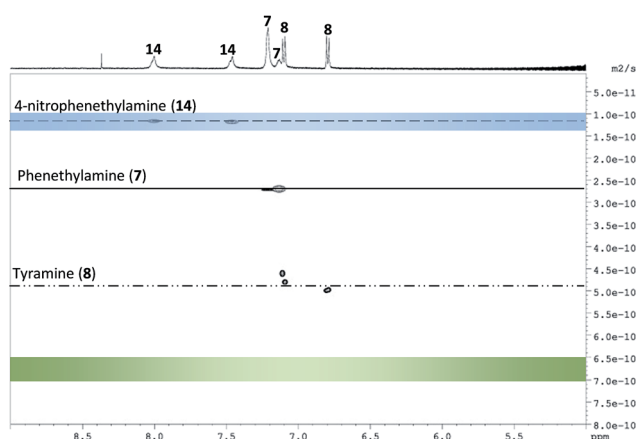


Fig. 6 DOSY experiment performed on a mixture of phenethylamine (7), tyramine (8) and 4-nitro-phenethylamine (14) in water in the presence of S1-AuNP (32 transients, 64 scans per transient, 2 h). Conditions: [analytes] = 0.5 mM, [AuNP] = 45 μ M, HEPES buffer 10 mM, and pD = 7.0. The blue and green bars represent the diffusion coefficients respectively of the nanoparticles and unbound analytes under the experimental conditions used.

designed, where the dimethylsilane protons resonate at 0.1 ppm. Indeed, this signal lies at the edge of the spectral window for most organic species, thus allowing a selective saturation of the nanoparticle monolayer. The effectiveness of **S4**-AuNP in the detection of drug models at low concentration was tested by analysing HEPES buffered D₂O solutions (pD 7.0) containing 1 mM phenylalanine (**14**) and 50 μ M *N*-methylphenethylamine (**6**). The severe spectral crowding and the low signal intensity prevent the detection of **6** in the sample with a standard ¹H experiment. However, a STD spectrum featuring a 2 s saturation time (Fig. 9a) selectively reveals the presence of **6**. No interference is observed from the overlapping signal of phenylalanine, even when it is present in 20-fold excess. Concentration dependent STD experiments confirmed that the integrated intensities of the signal from **6** increased linearly with dopamine concentration in the physiologically relevant concentration range 10–200 μ M, which allows the quantitative determination of the analyte (Fig. 9b and S26†). The lowest concentration of **6** that could be detected under these conditions is 30 μ M.

Conclusions

In this work, we designed monolayer protected nanoparticles capable of broad-class recognition of amphiphilic organic cations in water by exploiting strong ion pairing and hydrophobic interactions. The association of nanoparticle-based receptors with magnetization or saturation transfer NMR experiments allows the straightforward detection and identification of *N*-phenethylamine derivatives even in complex mixtures, at micromolar concentrations and without any pre-treatment. “Cocktails” of different drugs can be analysed using DOSY experiments. No other chemosensing-based system can reach a similar performance. In principle, by using this protocol it should be possible to propose a tentative chemical structure of a new “designer” drug a few hours after the seizure of a single tablet. The introduction of a dimethylsilane group in proximity to the sulfonate moiety allowed the synthesis of AuNPs tailored for STD-NMR experiments, where the monolayer can be saturated without interference from the sample's signals. This approach successfully decreased the detection limit down to the low micromolar range, thus opening the way to the screening of biological fluids such as urine or blood. It must be underlined that STD is expected to deliver a further relevant sensitivity increase when performed on spectrometers equipped with cryoprobes.

Simple thiols such as those used here can be synthesized in a few days and assembled on the nanoparticles in a few hours. Even if the insertion of aromatic residues apparently did not result in the establishment of additional interactions with the analyte, modifications in the hydrophobic layer resulted in a modified selectivity. This suggests that the next issue to address is to improve the design of the nanoreceptor in order to reach more sophisticated molecular recognition. More studies are needed in order to deeply understand the effect of thiol modifications on the monolayer properties and structure and in particular on how coating molecules interact with vicinal ones

and with external analytes as well. Yet, a binding site reproducing the structure of natural receptors could provide not only structural but also functional evidence against new illicit drugs.

Conflicts of interest

The authors declare no competing financial interests.

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