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Pd single atoms on g-C₃N₄ photocatalysts: minimum loading for maximum activity†

Velu Jeyalakshmi,^{‡ab} Siming Wu,^{‡a} Shanshan Qin,^{Ⓜa} Xin Zhou,^a Bidyut Bikash Sarma,^{Ⓜc} Dmitry E. Doronkin,^{Ⓜd} Jan Kolařík,^e Miroslav Šoos^b and Patrik Schmuki^{Ⓜ*ae}

Noble metal single atoms (SAs) on semiconductors are increasingly explored as co-catalysts to enhance the efficiency of photocatalytic hydrogen production. In this study, we introduce a “spontaneous deposition” approach to anchor Pd SAs onto graphitic carbon nitride (g-C₃N₄) using a highly dilute tetraaminepalladium(II) chloride precursor. Maximized photocatalytic activity and significantly reduced charge transfer resistance can be achieved with a remarkably low Pd loading of 0.05 wt% using this approach. The resulting Pd SA-modified g-C₃N₄ demonstrates a remarkable hydrogen production efficiency of 0.24 mmol h⁻¹ mg⁻¹ Pd, which is >50 times larger than that of Pd nanoparticles deposited on g-C₃N₄ via conventional photodeposition. This significant enhancement in catalytic performance is attributed to improved electron transfer facilitated by the optimal coordination of Pd SAs within the g-C₃N₄ structure.

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Introduction

In recent years, the demand for hydrogen as a clean renewable energy carrier has grown rapidly, driven by the need to replace traditional fossil fuels. Photocatalytic water splitting is a promising and most elegant way to directly convert solar energy into clean, renewable hydrogen (H₂).^{1,2} Since the pioneering work of Fujishima *et al.*, titanium dioxide (TiO₂) has become the most extensively studied material for this purpose.^{3–5} More recently, graphitic carbon nitride (C₃N₄) is being increasingly explored as it shares many advantageous features with TiO₂, such as high stability, abundance, and a suitable band structure for water splitting but provides the key advantage of strong visible light absorption, resulting in significantly enhanced solar light utilization.^{6–9}

Nevertheless, due to strong kinetic hindrance for the photocatalytic hydrogen evolution reaction (HER) on the C₃N₄

surface, an enhanced efficiency highly depends on the development of highly active co-catalysts – typically noble metals such as Pt, Pd, and Rh need to be deposited on the semiconductor surface to obtain a reasonable H₂ production rate. Despite efforts made using non-precious metal HER electrocatalysts,¹⁰ noble metal-based materials remain nearly irreplaceable due to their superior activity and chemical stability. However, the scarcity and high cost of these metals pose significant challenges to the large-scale application of these noble metal-based co-catalysts. This has prompted considerable efforts to minimize noble metal usage while retaining high catalytic performance. In this context, single-atom catalysts have garnered wide attention due to their maximized atom utilization efficiency.^{11–13}

As for many other semiconductors and also for C₃N₄ substrates, a wide range of noble metal SAs have been successfully deposited as nanoparticles, clusters or SA, and accordingly improvement in photocatalytic H₂ production has been reported.^{14–21} In search of the most active SA species, Akinaga *et al.* conducted a remarkable study on ten transition metal elements, including Cu, Ni, Pd, Pt, Rh, Ru, Ag and Au, as SAs anchored on g-C₃N₄. Among the metals tested, Pd demonstrated significantly higher hydrogen evolution activity compared to other precious metals such as Pt, Rh as well as other transition metals.¹⁴ The authors ascribed the superior activity of Pd on C₃N₄ to the suitable electronic structure of this metal on C₃N₄. In their work, Akinaga *et al.* used a photodeposition approach to achieve a relatively high SA loading (>0.5 wt%).

However, for many semiconductors, the activity of SAs in photocatalysis is extremely dependent on the deposition

^aDepartment of Materials Science WW4-LKO, Friedrich-Alexander-University of Erlangen-Nuremberg, Martensstrasse 7, 91058 Erlangen, Germany. E-mail: schmuki@www.uni-erlangen.de

^bDepartment of Chemical Engineering, University of Chemistry and Technology, Technická 3, Prague 160 00, Czech Republic

^cLaboratoire de Chimie de Coordination (LCC), CNRS, Université de Toulouse, INPT, UPR 8241, 205 Route de Narbonne, 31077 Toulouse Cedex 4, France

^dInstitute of Catalysis Research and Technology, KIT, Hermann-von-Helmholtz-Platz 1, 76344 Eggenstein-Leopoldshafen, Germany

^eRegional Centre of Advanced Technologies and Materials, Štechtitelů 27, 78371 Olomouc, Czech Republic

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‡ These authors contributed equally to this work.



approach.^{13,21–23} Namely, for Pt SAs on TiO₂, it was reported that “reactive” deposition leads to highly active SA configurations that can provide maximized H₂ production, *i.e.*, a very high catalytic efficiency can be reached at very low noble-metal loading. This approach relies on the surface reaction of highly dilute solutions of suitable noble-metal precursors.^{24–27}

In the present work, we first explore various Pd precursors for the feasibility of a reactive SA attachment on C₃N₄. We find that tetraaminopalladium(II) chloride – Pd(NH₃)₄Cl₂ as a Pd precursor solution with C₃N₄ allows for an adjustable Pd SA loading with a wide range of deposition concentrations from 0.04 wt% to 0.75 wt%. Our results show that by the reaction of a minimal amount of a 0.05 mM precursor, maximum photocatalytic efficiency can be obtained. The photocatalytic hydrogen production activity of such Pd SA-decorated C₃N₄ achieves a normalized H₂ production activity of 0.24 mmol h⁻¹ mg⁻¹ Pd, which is 55 times higher than that observed with Pd nanoparticle-decorated C₃N₄ at an effective loading that is more than 10 times lower than that typically reported in the literature for Pd on C₃N₄. The superior activity of Pd SAs/C₃N₄ is attributed to the strong coordination of Pd SAs within the C₃N₄ structure, forming a highly stable and catalytically effective configuration that drastically reduces the charge transfer resistance for the HER. These results illustrate how a refined anchoring of SAs on substrates can enable more cost- and production-effective use of precious metals in photocatalysis.

Results and discussion

Nanosheets of g-C₃N₄ were synthesized using a thermal polycondensation method starting from an equimolar mixture of melamine and dicyandiamide, followed by thermal exfoliation, as described in the literature.^{28–30} In order to explore the feasibility of direct deposition of a (reactive) SA such as Pd on C₃N₄,

we examined different precursor species, namely tetraaminopalladium(II) chloride (Pd(NH₃)₄Cl₂), palladium(II) chloride (PdCl₂) and ammonium hexachloropalladate(IV) ((NH₄)₂[PdCl₆]). To investigate the reactive deposition behavior, we used three different Pd precursors at a concentration of 2 mM to decorate Pd on C₃N₄. We then evaluated the general deposition behavior with electron microscopy and XPS and also evaluated the photocatalytic H₂ production performance. Among the samples, XPS results reveal that both PdCl₂ and (NH₄)₂[PdCl₆] lead to relatively high Pd loadings (>1 at%) (Fig. S1a and b[†]); however, the strong Cl 2p signals in the XPS spectra (Fig. S1c[†]) indicate that most of the Pd precursor did not react with C₃N₄, *i.e.*, the precursor is just physically adsorbed on the C₃N₄ surface. In the SEM images of these two samples (Fig. S2[†]), obvious Pd nanoparticles can be seen, due to the agglomeration caused by high loading. In contrast, the Pd(NH₃)₄Cl₂ sample shows no visible metal nanoparticle formation in SEM (Fig. 1a), non-metallic Pd position in XPS (Fig. S1a[†]) and no detectable Cl 2p signal (Fig. S1b[†]), indicating a complete reaction of this particular precursor with the C₃N₄ surface.

Fig. 1a shows the SEM image of Pd-deposited g-C₃N₄ (Pd SAs/C₃N₄) using Pd(NH₃)₄Cl₂ at a concentration of 0.002 mM, and Fig. S3[†] shows the SEM image of neat g-C₃N₄. The introduction of Pd SAs does not affect the morphology of C₃N₄ – both samples show a sheet-like structure with a thin layer thickness of approximately 16 nm. The high-angle annular dark-field scanning transmission electron microscopy (HAADF-STEM) image of Pd SAs/g-C₃N₄ is shown in Fig. 1b and c, which confirms the presence of individual Pd atoms (highlighted in Fig. 1c with red dots and yellow circles). Also, in the HAADF-STEM image (Fig. 1d), there are no observable Pd agglomeration on the g-C₃N₄ surface. Energy-dispersive X-ray spectroscopy (EDX) mapping (Fig. 1e–g) further proves the uniform dispersion of Pd SAs throughout the g-C₃N₄ structure. The density of



Fig. 1 (a) SEM image, (b) original HAADF-STEM image, (c) the HAADF-STEM image with individual Pd SAs highlighted by red dots and yellow circles, (d) HAADF-STEM image and the corresponding EDS mapping (e) C, (f) N, and (g) Pd of Pd SAs/C₃N₄.





Fig. 2 (a) X-ray diffraction pattern, (b–d) XPS spectra of (b) C 1s, (c) N 1s and (d) Pd 3d for C_3N_4 and Pd SAs/ C_3N_4 , (e) XANES spectra of Pd/ C_3N_4 at Pd K-edge, and (f) Fourier transform extended X-ray absorption fine structure (FT-EXAFS) spectra of Pd SAs/ C_3N_4 .

Pd SAs was calculated as $1.6 \times 10^6 \mu\text{m}^{-2}$ from HAADF-STEM images shown in Fig. S4.†

X-ray diffraction (XRD) patterns of $g-C_3N_4$ and Pd SAs/ $g-C_3N_4$ are presented in Fig. 2a. Both samples display two distinct diffraction peaks at 13° and 27.6° , corresponding to the (100) and (002) crystal planes of $g-C_3N_4$, respectively.³¹ Notably, no diffraction peaks related to metallic Pd are observed in the Pd SAs/ $g-C_3N_4$ sample (as is expected for the SA-decorated sample).¹⁴

X-ray photoelectron spectroscopy (XPS) was utilized to investigate the chemical state of $g-C_3N_4$ and Pd SAs on C_3N_4 (Fig. 2b–d and Table S1†). The high-resolution C 1s XPS spectrum (Fig. 2b) of both samples can be fitted by three peaks at

284.7 eV, 286.2 eV, and 288.1 eV corresponding to C–C, C–N and N–C=N of the heptazine ring carbon structure, respectively.¹⁶ The fitted N 1s spectra (Fig. 2c) exhibit peaks at 398.7 eV (N1), 400 eV (N2), and 401.4 eV (N3), representing the sp^2 hybridized aromatic two-coordinated (N2c) nitrogen of the triazine unit (C=N–C, *i.e.*, pyridinic N) and three coordinated (N3c) bridging N atoms connected to carbon as N–(C)₃ groups and amino functional (C–NH/NH_x) groups respectively.^{16,32,33} The Pd 3d spectrum of Pd SAs/ C_3N_4 shows two peaks (Fig. 2d), doublets at 337.6 and 342.9 eV corresponding to Pd⁰⁺ 3d_{5/2} and Pd⁰⁺ 3d_{3/2} ($0 < \delta < 2$), respectively.¹⁶ For comparison, Pd nanoparticles were deposited on $g-C_3N_4$ (Pd NPs/ C_3N_4) using an established photodeposition method described in the literature.^{34,35} SEM



images of Pd NPs/C₃N₄ (Fig. S5†) clearly show distinct Pd nanoparticles on the C₃N₄ surface with their typical diameter in the range of 7–15 nm. For this sample, the Pd 3d XPS spectra (Fig. S6†) of Pd NPs/C₃N₄ exhibit doublets at 335 eV and 340 eV, which are typically attributed to metallic Pd⁰.³³

The nature of Pd species was further investigated by X-ray absorption spectroscopy (XAS) and diffuse reflectance infrared Fourier transform spectroscopy (DRIFTS) measurements. Fig. 2e and f show the absorption near-edge structure (XANES) and extended X-ray absorption fine structure (EXAFS) spectra of Pd SAs/C₃N₄. The X-ray absorption spectrum of Pd SAs/C₃N₄ measured at the Pd K-edge (24 350 eV) indicates the presence of Pd atoms that are non-metallic and carry a positive charge.^{15,16,20} The EXAFS analysis and the corresponding Fourier transformed (FT) radial distribution function of Pd SAs/C₃N₄ show a peak at approximately 1.5 Å (without phase correction), attributed to the Pd–N bond, and no obvious scattering is observed for the metallic Pd–Pd bonding.^{16,20} CO-DRIFT spectra of Pd SAs/C₃N₄ (Fig. S7†) show a CO vibrational peak at 2125 cm⁻¹, which is characteristic of linearly bonded CO on a Pd single site (usually Pd²⁺).^{36,37} These results are well in line with the XPS results, *i.e.*, Pd SAs are N-coordinated in C₃N₄ with an oxidation state ≈ 2.

We then used the above deposition approach to place Pd SAs from Pd(NH₃)₄Cl₂ solutions in the concentration range of

0.0005 mM to 10 mM on C₃N₄. Fig. 3a shows the Pd 3d XPS spectra for these Pd SAs loaded on C₃N₄ samples. Notably, neither metallic Pd peaks nor Cl 2p peaks (Fig. S8†) are observed under any of the deposition conditions. Instead, the incorporation of Pd SAs is evident across all samples, as indicated by the Pd 3d doublet peaks at 337.6 and 342.9 eV. The XPS data align with the SEM images shown in Fig. S9,† where no Pd nanoparticles are observed in any of the samples, even at the highest precursor concentration of 10 mM.

The bulk loading of the samples prepared using different concentrations of Pd(NH₃)₄Cl₂ solutions was further quantified by atomic absorption spectroscopy (AAS) – the results are shown in Table S1.† As the concentration of the Pd precursor increases, the loading of Pd SAs increases (Fig. S10†), which is consistent with the XPS data (Table S2†). For the highest precursor concentration of 10 mM, the Pd SA loading reaches 0.75 wt%.

We then examined for all samples the photocatalytic H₂ evolution using a 365 nm LED light source with an intensity of 65 mW cm⁻² and an aqueous solution of 10% triethanolamine (TEOA) as a hole scavenger.^{19,38,39} From the results shown in Fig. 3b, it is evident that the decoration of either Pd NPs or Pd SAs significantly enhances the H₂ production activity compared to bare C₃N₄. In the concentration range of 0.0005 mM to 0.05 mM, corresponding to Pd SA loadings from 0.04 wt% to



Fig. 3 (a) Pd 3d XPS spectra of Pd SAs/C₃N₄ at different concentrations of tetraaminepalladium chloride, (b) photocatalytic H₂ evolution, (c) evolved H₂ at different Pd SA loadings, and (d) normalized H₂ evolution rates for different concentrations Pd SAs using tetraaminepalladium chloride.



0.26 wt% (Fig. S10[†]), the photocatalytic H₂ production activity increases with loading (Fig. 3c). It is worth mentioning that Pd SAs deposited through our method, utilizing the Pd(NH₃)Cl₂-precursor, lead to the notable finding that even a minimal loading of 0.05 wt% outperforms Pd NPs synthesized *via* photodeposition, the latter having a much higher loading of 1.5 wt% (Table S2[†]). As a side note, this loading is also much more efficient than using PdCl₂ and (NH₄)₂PdCl₆ precursors (see Fig. S11[†]). Also, in this comparison, the photocatalytic H₂ production results show that an SA deposition approach using Pd(NH₃)₄Cl₂ leads to much higher activity, despite the significantly lower Pd loading compared with PdCl₂ and (NH₄)₂PdCl₆.

In general, the data clearly show that the H₂ production amount increases with Pd precursor concentration loading until reaching a maximum at 0.05 mM (with a Pd SA loading of 0.26 wt%). Beyond this point, a further increase in Pd loading does not increase the activity any further but even leads to a slight drop in the activity and in the 10 mM case a very obvious decrease (Fig. 3c). This is likely due to the decrease in the density of Pd SAs and the formation of Pd agglomerates, which may increase charge recombination.⁴⁰

To further assess and compare the effectiveness of Pd as a co-catalyst in both single-atom and nanoparticle forms, we normalized the data from Fig. 3b relative to Pd loading (Table S1[†]); the results are shown in Fig. 3d. The analysis reveals that the highest mass-specific photocatalytic efficiency, resulting in an H₂ production rate of 0.24 mmol h⁻¹ mg⁻¹ Pd, is achieved with 0.002 mM Pd precursor (0.05 wt%). This efficiency is 55 times higher than that obtained through conventional photodeposition of Pd nanoparticles on g-C₃N₄, highlighting the superiority of our reactive deposition method for optimizing photocatalytic H₂ production. The exceptional performance of low Pd SA loading on g-C₃N₄, prepared using our direct deposition method, is evident when compared to other Pd SA-loaded g-C₃N₄ structures reported in the literature for photocatalytic H₂ generation. As shown in Table S3,[†] our work demonstrates the highest photocatalytic hydrogen evolution per Pd atom. Notably, even when compared to studies with similar or higher Pd SA loadings, the Pd SAs obtained through our reactive

deposition method using Pd(NH₃)₄Cl₂ exhibit the highest efficiency.

To better examine the origin of high activity of our SAs on C₃N₄, we evaluated the charge transfer properties of the Pd-decorated C₃N₄ photocatalysts by electrochemical impedance spectroscopy (EIS). The measurements were performed in the 0.1 M Na₂SO₄ electrolyte at -0.5 V *vs.* Ag/AgCl, *i.e.*, close to flat band conditions (details are outlined in the ESI-Experimental section[†]). Fig. 4a presents the Nyquist plots for bare g-C₃N₄, and g-C₃N₄ decorated with varying amounts of Pd SAs, and g-C₃N₄ decorated with Pd NPs (see the zoomed-in spectra shown in Fig. S12[†]). The Nyquist plots were fitted using the classic Randle's equivalent circuit model (inset of Fig. 4a).^{41,42} The significantly smaller radius of the fitted curve for Pd SAs/C₃N₄, compared to bare C₃N₄, indicates a substantial reduction in charge transfer resistance (*R*_{ct}) upon Pd SA loading. Quantitative fitting data in Table S4[†] show a 98-fold decrease in *R*_{ct} due to Pd SA incorporation (already at a concentration of 0.05 mM). Fig. 4b shows the *R*_{ct} values plotted against Pd SA loading, showing that even a minimal amount of Pd SAs (0.03 wt%) can dramatically enhance the charge transfer of C₃N₄ to the electrolyte. This aligns with the low loading required to achieve peak efficiency in photocatalytic H₂ production. Conversely, Pd NPs on C₃N₄ also reduce *R*_{ct} compared to bare C₃N₄ (Fig. 4b) but require a much higher loading (1.5 wt%) to achieve a similar reduction in charge transfer resistance, as compared to the Pd SA-loaded sample, which achieves this with just 0.03 wt%. PEIS measurements for all the samples were measured using a 365 nm LED (as described in the ESI Experimental section[†]). As shown in Fig. S13,[†] the results indicated a similar trend to the EIS data collected in the dark (Fig. 4 and Table S4[†]), although *R*_{ct} values were different. Under illumination, *R*_{ct} values decreased due to enhanced charge transfer dynamics in the presence of light (Fig. S13 and Table S5[†]). Notably, Pd SAs demonstrated lower *R*_{ct} values compared to Pd nanoparticles, indicating the superior performance of Pd SAs. These results underscore the effectiveness of small Pd SA quantities in significantly improving the charge transfer characteristics of C₃N₄.



Fig. 4 (a) EIS plots of C₃N₄, Pd SAs/C₃N₄ and Pd NPs/C₃N₄ at the voltage of -0.5 V (*vs.* Ag/AgCl) in 0.1 M Na₂SO₄ aqueous electrolyte. The equivalent circuit model used for fitting is depicted in the inset of (a). (b) *R*_{ct} vs. Pd loading plot of Pd SAs/C₃N₄ samples.



Incident photon-to-current conversion efficiency (IPCE) measurements were also conducted to assess the photoelectrochemical characteristics of C_3N_4 and Pd SAs/ C_3N_4 . Fig. S14† shows photocurrent spectra for both samples with a photocurrent onset in the visible range. The bandgap (determined from a replot of the photocurrent data according to an indirect transition, Fig. S14-inset†) was 2.7 eV, consistent with the typical bandgap of $g-C_3N_4$.^{14,43} Additionally, the photocurrent (Fig. S15†) increases after the decoration of Pd SAs on $g-C_3N_4$ (under near flat band conditions), which confirms the beneficial effect of Pd SAs in promoting charge transfer under illumination conditions.

The enhanced performance of Pd-SAs deposited by our decoration technique, compared with the literature, must be attributed to the used precursor $Pd(NH_3)_4^{2+}$ that leads to the direct formation of active Pd–N configurations (as confirmed by EXAFS, Fig. 2e and f) – in the literature, such sites are regarded as premier active sites in C_3N_4 .^{16,20,44,45}

The presence of NH_3 groups in $Pd(NH_3)_4^{2+}$ ions may promote the formation of such a Pd–N coordination structure compared to the chloride-coordinated Pd precursors due to suitable ligand exchange energetics.^{46,47} Furthermore, the Pd^{2+} in $Pd(NH_3)_4Cl_2$ enables stronger interactions with nitrogen atoms in C_3N_4 than the Pd^{4+} in $(NH_4)_2PdCl_6$ or Pd^{2+} in $PdCl_2$.^{48,49} Evidently, chloro-coordinated Pd precursors in the Pd^{2+} or Pd^{4+} state are either adsorbed onto the C_3N_4 surface at less specific sites, as shown by the XPS Pd 3d spectra (Fig. S1a†) and the significant presence of Cl detected in the XPS Cl 2p spectra (Fig. S1b†), or reduced and form metallic Pd agglomerates, as clearly observed in the SEM images (Fig. S2†). The tetraammonium complex, on the other hand, leads without any evident change in the reduction state to active $Pd^{\delta+}$ ($\delta \approx 2$) N-coordinated SAs, accompanied by the complete loss of Cl coordination during the reaction (Fig. S8†). Therefore, $Pd(NH_3)_4Cl_2$ is identified as the most suitable precursor for direct reactive deposition of Pd SAs on C_3N_4 .

Considering that many literature studies, particularly those involving DFT calculations, suggest that Pd in an N_4 -coordination on C_3N_4 exhibits the highest degree of stability and activity, one may conclude that Pd^{2+} undergoes a ligand exchange process to form this active Pd– N_4 configuration. This means that the process is of self-homing nature (as described for the reactive deposition of Pt)^{23–25,35} *i.e.*, the Pd precursor reacts and deposits Pd SAs at most active surface sites on $g-C_3N_4$, these sites then provide a maximized electron transfer and thus are highly catalytically active. This explains why such a low loading of Pd SAs is sufficient in our work to achieve maximised photocatalytic H_2 production efficiency compared to the higher Pd loadings required, as reported in most literature studies (Table S3†).^{14,45,49,50}

The remarkable activity becomes particularly clear if the present data are compared to the work of Akinaga *et al.*,¹⁴ where a 0.5 wt% Pd SA loading was required to maximize the photocatalytic H_2 production activity, *i.e.*, our very low Pd SA loading of 0.05 wt% demonstrates a tenfold increase in efficiency. This superior performance highlights the importance of the attachment chemistry and process of Pd SAs within C_3N_4 , *i.e.*,

processes that lead to Pd SAs located at the most active sites can lead to maximized efficiency with minimal Pd usage, avoiding the waste of Pd associated with random Pd SA or NP deposition.

Conclusion

In this work, we successfully integrated Pd SAs onto/into exfoliated $g-C_3N_4$ *via* a reactive deposition method, achieving a controllable uniform loading of highly active Pd SAs in a Pd–N configuration on $g-C_3N_4$. Notably, using reactive deposition from a $Pd(NH_3)_4Cl_2$ precursor a low loading of 0.05 wt% Pd SAs on C_3N_4 with a density of $1.6 \times 10^6 \mu m^{-2}$ can achieve a maximum H_2 production rate of $0.24 \text{ mmol h}^{-1} \text{ mg}^{-1} \text{ Pd}$, significantly higher than that of Pd nanoparticles decorated on $g-C_3N_4$ and also ten times higher than that of Pd SAs decorated on $g-C_3N_4$ using other reported approaches. Other tested precursors may also deliver SA attachment but lack the high co-catalytic activity. These results underline the importance of the attachment mechanism in creating a SA/substrate coupling with minimized charge transfer resistance and thus maximized co-catalytic activity – in a most effective way, the process is self-homing, *i.e.*, activation takes place where it is most effective.

Data availability

The data that support the findings of this study are available within the article and ESI.†

Author contributions

Velu Jeyalakshmi conducted the majority of the experimental work, including synthesis, characterization, data analysis, and drafting of the manuscript. Siming Wu contributed significantly to data analysis and revision of the original draft. Shanshan Qin participated in data analysis, while Xin Zhou performed HAADF-STEM analysis. Bidyut Bikash Sarma and Dimitry E. Doronkin conducted XAS and EXAFS analyses and Jan Kolařík carried out AAS analysis. Miroslav Šoos supervised the experiments and managed funding acquisition. Patrik Schmuki conceived the idea, supervised the experiments, analyzed the data, revised the manuscript, and secured funding. All authors contributed to the preparation of the manuscript and approved the final manuscript.

Conflicts of interest

There are no conflicts to declare.

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