

## PAPER

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# Engineering Au single-atom sites embedded in TiO<sub>2</sub> nanostructures for boosting photocatalytic methane oxidation†

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Photocatalytic methane oxidation under mild conditions using single-atom catalysts remains an advanced technology. In this work, gold single atoms (Au SAs) were introduced onto TiO<sub>2</sub> nanostructures using a simple method. The resulting performance demonstrated effective conversion of methane into H<sub>2</sub> and C<sub>2</sub> products at room temperature. The as-synthesized Au SA/TiO<sub>2</sub> exhibited a high hydrogen production rate of 2190 μmol g<sup>-1</sup>, with selectivity reaching up to 58% under optimized conditions. The methane oxidation mechanism was investigated, revealing a methyl radical pathway for generating value-added chemicals. This research provides a strategy for photocatalytic methane conversion over single-atom-supported photocatalysts.

Received 18th November 2024

Accepted 10th January 2025

DOI: 10.1039/d4na00947a

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## 1. Introduction

Methane (CH<sub>4</sub>) is a crucial fossil fuel and an important component of natural resources, serving as a significant chemical raw material. It constitutes approximately 80% of natural gas, making it a massive contributor to the greenhouse effect and global warming, particularly during the combustion process compared to carbon dioxide (CO<sub>2</sub>).<sup>1,2</sup> Therefore, the sustainable harvesting of CH<sub>4</sub> and its conversion into valuable products is essential for the sustainable utilization of natural resources and environmental protection. However, conventional methods for CH<sub>4</sub> conversion typically require high operating conditions (>100 °C and 100 MPa), which can result in excessive oxidation and low product selectivity. This is due to the bond dissociation energy of the initial C–H bonds being higher than that of the subsequent C–H bonds, which prevents the activation of the first C–H bond when relying solely on thermal energy to overcome the activation barrier.<sup>3–5</sup> Thus, it remains a challenge to enhance the selectivity of CH<sub>4</sub> conversion to high-value-added products under mild conditions.<sup>6,7</sup> Photocatalysis has been recognized as a green technique and a unique method for harnessing light to generate electron–hole pairs in semiconducting materials.<sup>8–11</sup> The application of single-atom catalysts in methane oxidation shows significant

advancement over traditional catalysts, as they exhibit enhanced catalytic performance by uniformly decorating noble metal atoms. This arrangement facilitates active exposure and increases the surface area for the oxidation process. Additionally, single-atom catalysts reduce the amount of catalyst required, addressing concerns about high costs and the scarcity of precious metals, making them economically viable even in small quantities.<sup>12–14</sup> It is crucial to understand the overall reaction rate of photocatalysts with varying active atom loadings due to their low coverage, which helps prevent the formation of nanoparticles. Single-atom catalysts also integrate seamlessly with advanced nanomaterials, such as reduced graphene oxide and carbon nanotubes, further improving catalytic activity and stability. These catalysts have demonstrated the ability to reduce dependence on noble metals while enhancing selectivity, owing to the abundance of active sites. Several studies have focused on engineering noble metal atoms (*e.g.*, Pt and Ag) on titania catalysts, exploring the role of titanium cations and oxygen vacancies in the localization of single atoms on oxo ligands, thereby facilitating superior photocatalytic performance.<sup>15–20</sup>

Titanium dioxide (TiO<sub>2</sub>) has garnered significant attention within the scientific community and has become a prominent commercial product due to its unique properties and diverse applications. Its excellent chemical stability, strong catalytic ability, and wide band gap energy (*E<sub>g</sub>*) under light illumination contribute to its utility.<sup>21–23</sup> The application of TiO<sub>2</sub> as a catalyst in photocatalytic methane oxidation and water splitting presents a promising solution in response to the urgent demand for net-zero emissions and sustainable energy applications. However, one of the main disadvantages of TiO<sub>2</sub> is its wide band gap and the rapid recombination rate of electron–hole pairs, which poses challenges for researchers. Numerous strategies

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† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d4na00947a>

have been explored to enhance the photocatalytic performance of  $\text{TiO}_2$ , including metal doping, non-metal loading, and surface modification with other semiconductors or quantum dots.<sup>3,24,25</sup> These combinatorial techniques aim to reduce the electron-hole recombination rate, extend light harvesting capabilities, and improve surface reaction kinetics.

Building on this foundation, the highly selective methane oxidation using an  $\text{Fe}/\text{TiO}_2$  catalyst was investigated in 2018, which demonstrated a 15% conversion rate after 3 hours of reaction, with methanol selectivity exceeding 90%. The presence of  $\text{FeOOH}$  and  $\text{Fe}_2\text{O}_3$  plays a crucial role in enhancing charge transfer and separation in  $\text{TiO}_2$ , thereby reducing the overall overpotential of the reaction. Photocatalytic methane oxidation over the  $\text{TiO}_2$  nanostructure is commonly considered owing to its low cost, stable chemical properties, and non-toxic nature.<sup>26</sup>

In the same context, photocatalytic  $\text{CH}_4$  oxidation to  $\text{C}_1$  oxygenates over  $\text{Pd}/\text{TiO}_2$  composites was investigated using  $\text{O}_2$  and  $\text{H}_2\text{O}$  as oxidants. Palladium ( $\text{Pd}$ ) serves as a hole acceptor, while oxygen vacancies function as electron acceptors, enhancing charge separation efficiency. Under optimized conditions,  $\text{C}_1$  products were obtained with approximately 99% selectivity and a high yield of  $54.7 \text{ mmol g}^{-1} \text{ h}^{-1}$ .<sup>27</sup> Shuang *et al.* utilized  $\text{Au-ZnO}/\text{TiO}_2$  composites for the direct oxidation of  $\text{CH}_4$  to valuable  $\text{C}_2\text{H}_6$  in a flow reactor under mild conditions. A  $\text{C}_2\text{H}_6$  production rate of over  $5 \text{ mmol g}^{-1} \text{ h}^{-1}$  was achieved after 50 minutes of reaction, with a high selectivity of 90% attributed to the weak overoxidation ability of  $\text{ZnO}$ . In this system,  $\text{Au}$  facilitates the desorption of  $\cdot\text{CH}_3$  radicals in the gas phase, thereby preventing overoxidation to  $\text{CO}_2$ .<sup>28</sup>

In this work,  $\text{Au}$  single-atom ( $\text{Au SA}$ ) supported  $\text{TiO}_2$  nanostructures were investigated for driving  $\text{CH}_4$  oxidation under mild conditions. Due to energetic active sites of  $\text{Au SAs}$ , the optimized production yield of  $\text{H}_2$  reached  $2190 \mu\text{mol g}^{-1}$ , with a selectivity of 58%. The investigation into the plausible mechanism indicated that  $\text{Au}$  inhibits the charge recombination rate and activates the reactant gases. These results suggest that this approach can effectively enhance  $\text{H}_2$  production through improved photocatalytic performance at room temperature.

## 2. Experimental

### 2.1. Materials

Titania powder ( $\text{TiO}_2$ , P25, 99%), gold(III) chloride tetrahydrate ( $\text{HAuCl}_4 \cdot 4\text{H}_2\text{O}$ ), sodium borohydride powder ( $\text{NaBH}_4$ , 98%), and sodium hydroxide ( $\text{NaOH}$ ) were purchased from Sigma-Aldrich. Deionized (DI) water and ethanol ( $\text{C}_2\text{H}_5\text{OH}$ ) were used as a solvent for material cleaning steps. All chemicals were used without further purification.

### 2.2. Preparation of catalysts

The  $\text{Au SA}$ -supported photocatalysts were synthesized using a wet precipitation process, as presented in Fig. 1. Specifically, 0.3 g of  $\text{TiO}_2$  was added to an aqueous solution of  $\text{HAuCl}_4$  ( $1.5 \text{ mg mL}^{-1}$ ), with the pH pre-adjusted to approximately 9

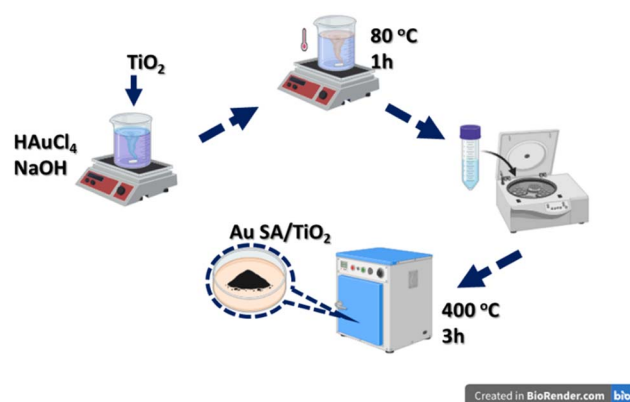


Fig. 1 The schematic illustration of  $\text{Au}$  single-atom ( $\text{Au SA}$ ) – supported catalysts.

using 0.1 M  $\text{NaOH}$  solution and  $\text{NaBH}_4$  solution, followed by stirring at  $80^\circ\text{C}$  for 1 hour. After the reaction, the precipitates were collected, thoroughly washed with deionized water, and dried at  $60^\circ\text{C}$  overnight. The loading content of  $\text{Au SAs}$  was estimated at 2 wt% as the standard dose. The final product was calcined at  $400^\circ\text{C}$  for 3 hours, ground, and denoted as  $\text{Au SA}/\text{TiO}_2$ , which was then stored for subsequent photocatalytic experiments. For comparison,  $\text{Au NPs}/\text{TiO}_2$  and  $\text{Pt NPs}/\text{TiO}_2$  were also prepared *via* the traditional impregnation technique without pH control, as in previous publications.<sup>29,30</sup>

### 2.3. Characterization

The crystal structure of the prepared photocatalysts and pure  $\text{TiO}_2$  was examined by using an X-ray diffractometer (XRD), equipped with a  $\text{Cu-K}\alpha$  radiation source working at 40 kV and 30 mA. The presence of single atoms was observed by using a scanning transmission electron microscope (S/TEM) equipped with an HAADF detector (electron energy of 300 keV). The optical absorption and radiative recombination of photocatalysts were studied by ultraviolet-visible diffuse reflectance spectroscopy (UV-vis DRS) and photoluminescence (PL) using a spectrometer (PerkinElmer Lambda 650S) and HR Labram, Horiba, respectively. The optical properties of selected samples were confirmed by using Fourier transform infrared spectroscopy (FTIR/JASCO-4600), Raman spectroscopy (Olympus-BX5), a semiconductor device analyzer (Keysight B1500A), and a CH instruments electrochemical workstation.

The analysis of paramagnetic species was performed by continuous-wave electron paramagnetic resonance (EPR). These experiments were performed on a Bruker ELEXSYS E500 spectrometer operating in the X-band (9.5 GHz). The EPR spectra were recorded at 94 K to avoid electron-hole recombination. The detection of radical species was conducted by spin-trapping tests using 5-*tert*-butoxycarbonyl-5-methyl-1-pyrroline-*N*-oxide (BMPO) as the scavenger. The atomic-scale and electronic structure of the sample were investigated by scanning tunneling microscopy (STM) measurement. For this, the  $\text{Au SA}/\text{TiO}_2$  solution was drop-cast on a gold plate, that was quickly loaded in the load lock chamber of a UHV system (base pressure  $< 1 \times 10^{-10}$  Torr) to be heated for twelve hours at  $85^\circ\text{C}$ . The STM



experiments were performed at room temperature with a tungsten tip and under tunneling conditions:  $V_{\text{sample}} = 1.8 \text{ V}$ ,  $I_t = 200 \text{ pA}$ .

#### 2.4. Photocatalytic methane oxidation

The photocatalytic oxidation of  $\text{CH}_4$  by  $\text{CO}_2$  molecules was carried out in a batch stainless-steel reactor (100 mL), with a circular quartz window attached on top. In brief, 50 mg of the sample was placed in the reactor, which was closed and evacuated for 10 min prior to the provision of 10 bars of the  $\text{CH}_4$  :  $\text{CO}_2$  gas mixture (3 : 1 pressure ratio) and kept under dark conditions for 30 min. Afterwards, a 300 W xenon lamp (200–1000 nm) was switched on as the light source. The wavelength of light irradiation could be varied by specific filters. The chamber was kept at room temperature by using a chiller. The gaseous products were analyzed and quantified using a gas chromatography system (HP-6890, Agilent).

### 3. Results and discussion

#### 3.1. Structural and morphological properties

The crystal structures of the samples were characterized by X-ray diffraction (XRD). As shown in Fig. 2a, the  $\text{TiO}_2$

nanostructures exhibit mixed crystal phases of rutile and anatase, corresponding to the standard JCPDS data (no. 21-1272 for rutile and no. 21-1276 for anatase).<sup>31–33</sup> Upon calcination at 400 °C, the well-defined diffraction peaks at 25.8°, 28.1°, 37.1°, 48.3°, and 55.3° were assigned to the (101), (110), (004), (220), and (105) planes of  $\text{TiO}_2$ , respectively. Additionally, the broadened peaks at 44.1° and 64.1° were observed, indicating the presence of Au SAs on the  $\text{TiO}_2$  nanostructures. The relatively small size of the Au diffraction peaks suggests high dispersion of the Au SA. Fig. 2b–d illustrate the Ti 2p, Au 4f, and O 1s region spectra of the selected Au SA/ $\text{TiO}_2$  composite. The main photoelectron peaks at binding energies of 463.8 eV and 458.3 eV are assigned to  $\text{Ti}^{4+}$  2p<sub>1/2</sub> and 2p<sub>3/2</sub>, respectively (Fig. 2b).<sup>34</sup> The Au 4f spectra were deconvoluted into two major peaks at 86.5 eV and 82.9 eV, corresponding to the characteristic peaks of Au 4f<sub>5/2</sub> and Au 4f<sub>7/2</sub>, respectively (Fig. 2c). These findings suggest that Au is reduced from  $\text{Au}^{3+}$  to  $\text{Au}^{1+}$  upon the formation of single atoms.<sup>35,36</sup> As observed in Fig. 2d, the O 1s peak at 529.4 eV can be deconvoluted into two peaks, with primary peaks at 529.1 eV and 529.6 eV corresponding to lattice oxygen ( $\text{O}_\text{L}$ ) and surface-adsorbed oxygenated species ( $\text{OH}_\text{s}/\text{O}_\text{s}$ ), respectively.<sup>34</sup> This confirms the creation of oxygen vacancies. These characterization studies confirm the phase and surface chemical states of the Au SA/ $\text{TiO}_2$  catalyst.

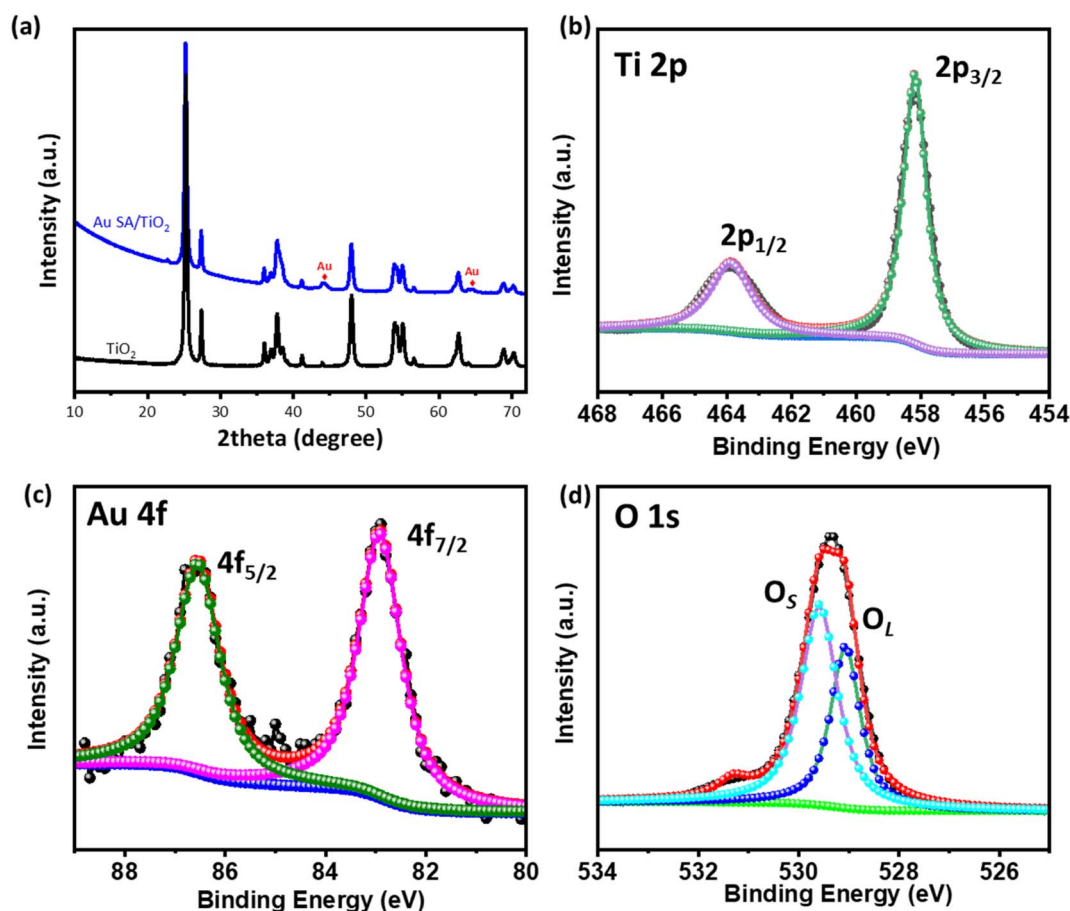


Fig. 2 (a) XRD patterns of pristine  $\text{TiO}_2$  and the Au SA/ $\text{TiO}_2$  composite and (b)–(d) XPS spectra of Ti 2p, Au 4f, and O 1s of the selected Au SA/ $\text{TiO}_2$  composite.



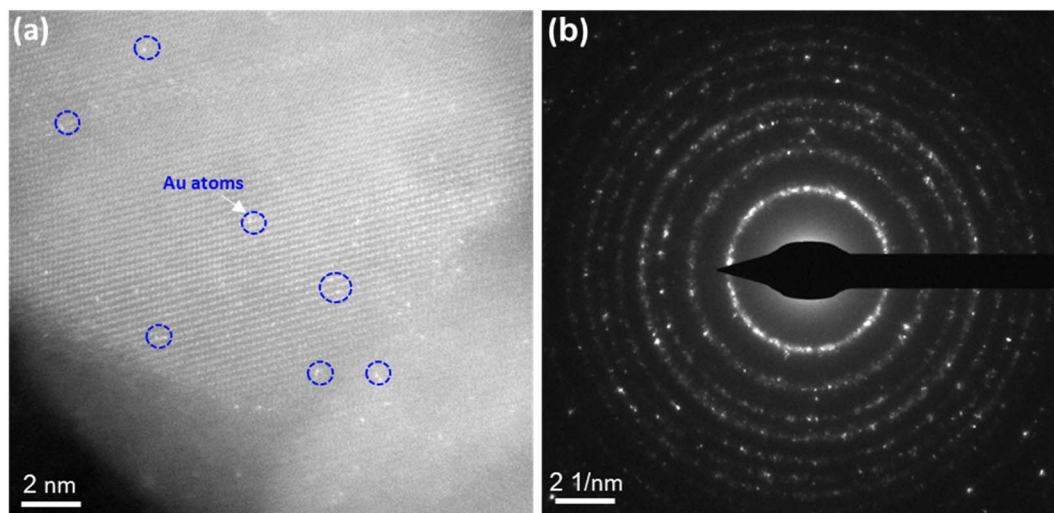


Fig. 3 The HR-TEM image and SAED images of Au SA/TiO<sub>2</sub> powder, respectively (a) and (b). The bright spots in (a) indicate Au single atoms (SA) over the TiO<sub>2</sub> matrix.

Due to the inherent resolution limitations of the SEM system, these images cannot definitively confirm the formation of single atoms or nanoparticle assemblies. Therefore, high-resolution TEM (HR-TEM) analysis provides clear evidence of the distribution of Au SAs on the surface of TiO<sub>2</sub> nanostructures. As illustrated in Fig. 3, the Au SAs are atomically dispersed across the TiO<sub>2</sub> surface, visible through their brighter contrast (highlighted by the blue circle). This confirms the enhanced distribution of Au SAs as a result of the calcination process, with no noticeable agglomeration compared to the Au NPs/TiO<sub>2</sub> synthesized *via* traditional impregnation methods. In other words, TiO<sub>2</sub>, as a reducible oxide, is well-suited for the atomic dispersion of metal catalysts due to its inherent surface

defects, which help stabilize metal atoms by forming metal-oxygen-support bonds.<sup>37,38</sup>

To further investigate the Au SA/TiO<sub>2</sub> heterostructure at the atomic scale, the STM analysis has been carried out (schematized in Fig. 4a). The topography STM image indicated the atomic arrangement of TiO<sub>2</sub> powder, although the signals of Au SAs were unable to be detected (Fig. 4b). By selecting a specific area on the catalyst surface, STM measurement was conducted and revealed the electronic structure of Au/TiO<sub>2</sub> (Fig. 4c). By measuring the distance between peaks of the valence band ( $E_V$ ) and conduction band ( $E_C$ ), the bandgap energy of Au SA/TiO<sub>2</sub> was estimated to be 2.71 eV, which is smaller than 3.2 eV of pristine TiO<sub>2</sub> P25, leading to favorable benefits for

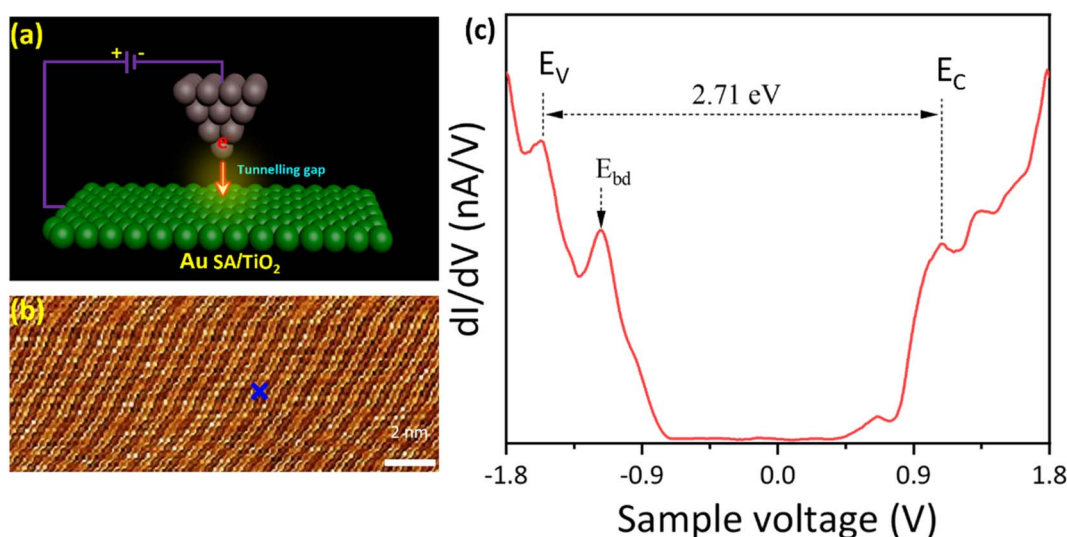


Fig. 4 (a) The schematic illustration of scanning tunneling microscopy (STM) analysis, (b) and (c) topography STM image and tunneling spectroscopy of the Au SA/TiO<sub>2</sub> sample recorded from the green cross in (b). Tunneling conditions:  $V_{\text{sample}} = 1.8$  V,  $I_t = 200$  pA. The valence and conduction band edges of Au SA/TiO<sub>2</sub> are indicated by vertical segments, labeled as  $E_V$  and  $E_C$ . The  $E_{bd}$  stands for bright defects in the TiO<sub>2</sub> structure.





photocatalytic reactions.<sup>39</sup> Besides, the bright defects ( $E_{bd}$ ) near the valence band were also observed, indicating shallow trapping states of  $\text{TiO}_2$  crystals.

### 3.2. Optical properties

To understand the effect of Au loading on the light absorption properties of  $\text{TiO}_2$ , UV-vis spectra were analyzed. As shown in Fig. 5a, apart from strong absorption of pristine  $\text{TiO}_2$  below 400 nm, the Au SA/ $\text{TiO}_2$  powder exhibited another absorption peak at 550 nm, indicating the plasmonic effects, which were absent in the case of the pristine  $\text{TiO}_2$  spectrum. The plasmon resonance of Au SAs could significantly contribute to photocatalytic performance later. Moreover, the steady-state photoluminescence (PL) spectra were obtained to confirm the separation properties of photogenerated carriers (Fig. 5b). The pure  $\text{TiO}_2$  samples exhibited stronger photoluminescence emission signals compared to Au SA/ $\text{TiO}_2$ , indicating a higher recombination rate of electrons and holes. Notably, Au SA/ $\text{TiO}_2$  displayed the lowest fluorescence intensity, suggesting efficient separation of photocarriers. This finding indicates that the transfer of excited electrons occurred rapidly between Au and  $\text{TiO}_2$ , effectively reducing the recombination rate in the photocatalysts.<sup>40,41</sup> The photocatalysts were analyzed using the photocurrent response and electrochemical impedance spectroscopy. Fig. S1a† displays the current–voltage ( $I$ – $V$ ) curve for the Au SA/ $\text{TiO}_2$  photocatalyst under both dark and light conditions, with the applied voltage ranging from  $-1$  V to  $+1$  V. Under light illumination, the slope of the  $I$ – $V$  curve begins to increase, indicating that the generated charge carriers significantly influence the electrical behavior of the optimized sample. Furthermore, the Nyquist plot in Fig. S1b† shows that the electron transport capability of the electrode surface has been slightly enhanced in the Au SA/ $\text{TiO}_2$  sample compared to pure  $\text{TiO}_2$ .

The interaction between  $\text{TiO}_2$  and Au SAs was analyzed using Raman spectroscopy and FTIR. As shown in Fig. 6a and S2a,† the Raman spectrum of the selected samples displays five characteristic modes located at 128, 182, 390, 514, and

$643\text{ cm}^{-1}$ , which correspond to the  $E_g$ ,  $E_g$ ,  $B_g$ ,  $A_g$ , and  $E_g$  modes, respectively. The main peak at  $128\text{ cm}^{-1}$ , characteristic of the Ti–O stretching mode in pristine  $\text{TiO}_2$ , positively shifts to  $133\text{ cm}^{-1}$ , confirming that Au has interacted with the Ti–O bond. This shift indicates the strengthening of the Ti–O bond, as well as phonon confinement effects associated with the creation of oxygen vacancies.<sup>42,43</sup> The FTIR spectra of all samples show bands corresponding to  $\text{TiO}_2$ , with a broad band from 400 to  $900\text{ cm}^{-1}$  assigned to Ti–O–Ti vibrations, as well as bands at around  $1621\text{ cm}^{-1}$  and  $3416\text{ cm}^{-1}$ , which are attributed to O–H bending and O–H stretching vibrations, respectively (Fig. 6b and S2b†).<sup>44,45</sup> The FTIR spectrum of  $\text{TiO}_2$  after modification with Au does not exhibit significant changes compared to pristine  $\text{TiO}_2$ , likely due to the formation of Au single atoms on the overall  $\text{TiO}_2$  structure.

### 3.3. Photocatalytic performance

According to unique properties, the Au SA/ $\text{TiO}_2$  sample was employed for photocatalytic  $\text{CH}_4$  oxidation by  $\text{CO}_2$  molecules, as depicted in Fig. 7. Note that there were no products generated in the absence of catalysts and light irradiation (Fig. S3†). By contrast, under light conditions, the Au SA/ $\text{TiO}_2$  catalyst was able to convert  $\text{CH}_4$  and  $\text{CO}_2$  into several gaseous products as a function of reaction time (Fig. 7a). Obviously, Au SA/ $\text{TiO}_2$  could produce a significant number of products after 8 hours (around  $3800\text{ }\mu\text{mol g}^{-1}$  in total) including  $\text{H}_2$ ,  $\text{C}_2\text{H}_6$ ,  $\text{C}_3\text{H}_8$  and  $\text{CO}$ , with a  $\text{H}_2$  selectivity of about 58%. This performance surpassed that of pure  $\text{TiO}_2$  powder (Fig. S4†) and other reported composites (Table 1). Although the Au SA/ $\text{TiO}_2$  photocatalyst did not exhibit the highest performance in comparison with other composites listed in Table 1, it remains a promising candidate for  $\text{CH}_4$  oxidation reactions.

To determine the role of Au single atoms, the photocatalytic activities of Au SA/ $\text{TiO}_2$  catalysts were compared with those of Au nanoparticle (NP)-based samples. The morphology of typical Au NP-supported  $\text{TiO}_2$  is depicted in Fig. S5.† Under the same reaction conditions, Au SA/ $\text{TiO}_2$  exhibited the highest product yield, outperforming Pt NPs/ $\text{TiO}_2$  and Au NPs/ $\text{TiO}_2$  (Fig. 7b).

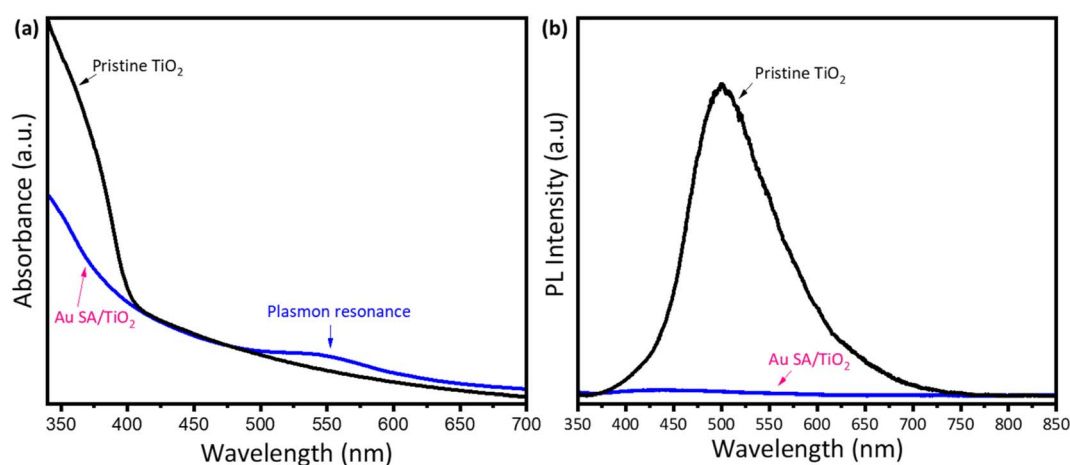


Fig. 5 The UV-vis absorption spectra of pristine  $\text{TiO}_2$  and Au SA/ $\text{TiO}_2$ , respectively (a) and (b).



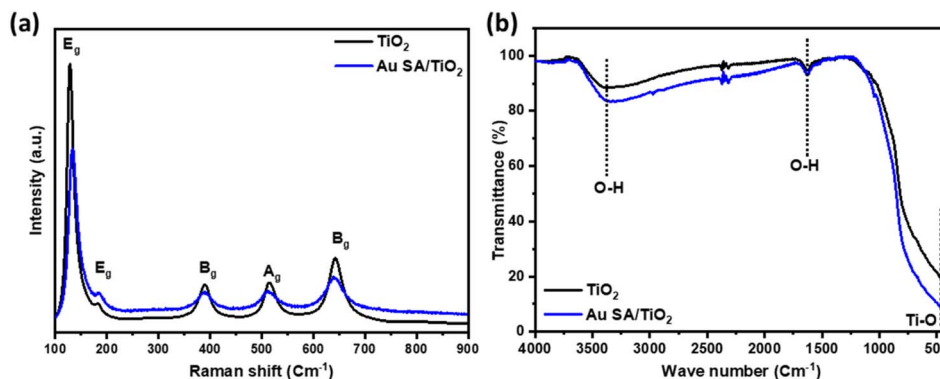


Fig. 6 (a) The Raman and (b) FTIR spectra of pristine  $\text{TiO}_2$  and Au SA/ $\text{TiO}_2$ , respectively.

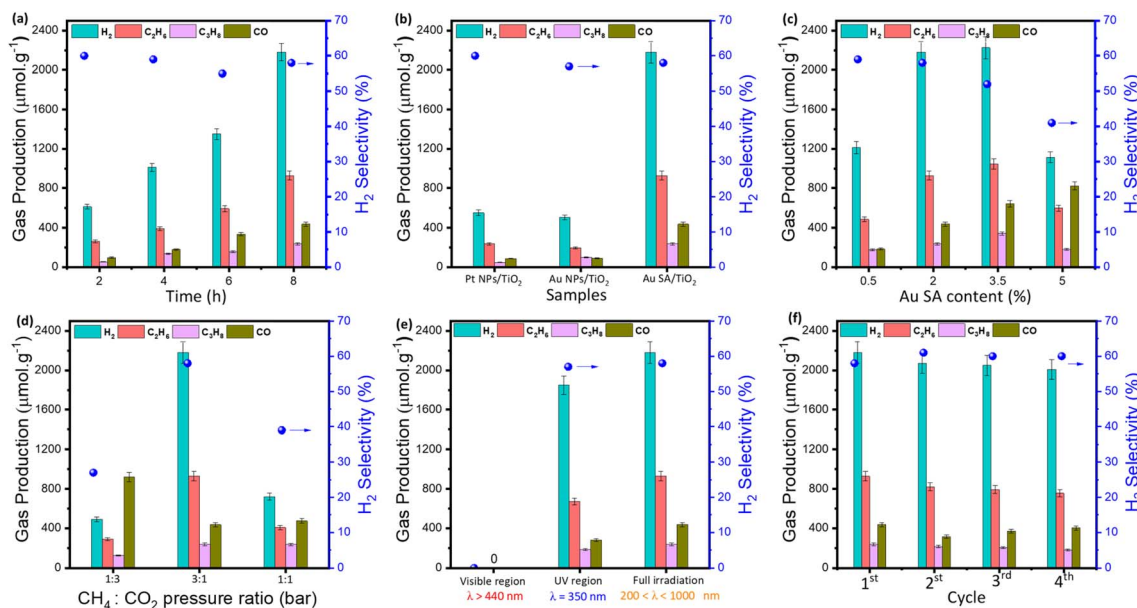


Fig. 7 Photocatalytic methane oxidation performances and the  $\text{H}_2$  selectivity as a function of reaction time over the standard Au SA/ $\text{TiO}_2$  sample (2 wt% Au) (a), comparable photocatalytic experiments of different catalysts (b), comparison of photocatalytic activity with different Au loading contents, reactant ratios, and irradiation regions, respectively (c)–(e), and stability tests over four cycles (f). Reaction conditions: 50 mg of the standard Au SA/ $\text{TiO}_2$  catalyst with 2% Au loading,  $\text{CH}_4 : \text{CO}_2 = 3 : 1$  (stand pressure ratio), xenon lamp 300 W. In figures (a), (d) and (f), standard Au SA/ $\text{TiO}_2$  was used.

These results confirm that photocatalytic activities of the single-atom catalyst are more efficient for driving  $\text{CH}_4/\text{CO}_2$  conversion compared to that of nanoparticle-based samples, which could be attributed to its abundant active sites and highly energetic features. Moreover, the effect of Au SA loading on product formation was then investigated. As shown in Fig. 7c, the catalysts with Au SA loading contents in the range of 2–3.5% showed outstanding performances, which were remarkably higher than that of the 0.5% and 5% Au SA loadings. This enhanced performance can be attributed to the improved charge separation rate in the photocatalyst powders.

On top of that, the reactant composition is also a key element and essential to study (Fig. 7d). By varying the pressure ratios of  $\text{CH}_4$  to  $\text{CO}_2$ , the photocatalytic yields of the catalyst were changed drastically. Particularly, by feeding  $\text{CH}_4$  and  $\text{CO}_2$  with

ratios of 1 : 3 and 1 : 1, the yields of  $\text{H}_2$  and hydrocarbon products were relatively low, while the amount of O-based products increased to around  $900 \mu\text{mol g}^{-1}$  for CO formation, compared to the ratio of 3 : 1, which contains more  $\text{CH}_4$  composition. It is evident that the main products were preferentially generated from  $\text{CH}_4$  rather than  $\text{CO}_2$ .

Next, the wavelength-dependent yields were investigated in the visible region and UV region, and under full irradiation for the standard Au SA/ $\text{TiO}_2$  photocatalyst. The catalyst exhibited impressive activities under full illumination and in the UV range, while the visible region was unable to activate  $\text{CH}_4$  and  $\text{CO}_2$  molecules (Fig. 7e). The nonappearance of photocatalytic activities in the visible range could be explained by its deficient energy to excite charge carriers in deep 3d orbitals of the  $\text{TiO}_2$  crystal for  $\text{CH}_4$  activation. As another important factor, the



**Table 1** Comparison of the methane oxidation performance over the Au SA/TiO<sub>2</sub> catalyst with other reported composites

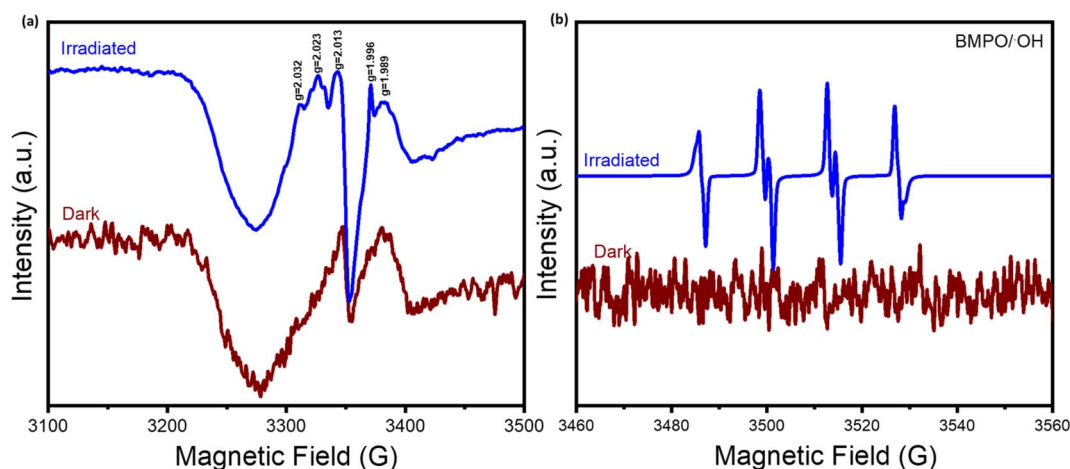
Photocatalysts	Amount of catalyst	Reaction conditions	Performance	Ref.
Cu@TiO <sub>2</sub>	10 mg	3 MPa CH <sub>4</sub> H <sub>2</sub> O <sub>2</sub> 1 M 10 mL UV-vis light	CH <sub>3</sub> OH: 20 000 $\mu\text{mol g}^{-1} \text{h}^{-1}$	46
Ru/TiO <sub>2</sub> -H <sub>2</sub>	5.0 mg	8 vol% CO <sub>2</sub> 8 vol% CH <sub>4</sub> 84 vol% Ar 50 mL min <sup>-1</sup> 300 W Xe lamp	CO: 708.4 mmol g <sup>-1</sup> h <sup>-1</sup> , H <sub>2</sub> : 645.5 mmol g <sup>-1</sup> h <sup>-1</sup>	47
Pt@TiO <sub>2</sub>	2 mg	3 MPa CH <sub>4</sub> 300 W Xe lamp	CH <sub>3</sub> OH: 300 $\mu\text{mol g}^{-1} \text{h}^{-1}$	48
Rh/TiO <sub>2</sub>	8.6 mg	1% CH <sub>4</sub> 1% CO <sub>2</sub> 98% Ar 150 W Hg-Xe lamp	H <sub>2</sub> : 21.5 mmol g <sup>-1</sup> h <sup>-1</sup> , CO: 21.2 mmol g <sup>-1</sup> h <sup>-1</sup>	49
Cr@TiO <sub>2</sub>	10 mg	3 MPa CH <sub>4</sub> 9.5 mL H <sub>2</sub> O <sub>2</sub> (30%) Mild conditions	CH <sub>3</sub> OH: 3.45 $\mu\text{mol g}^{-1} \text{h}^{-1}$	50
2% Rh/TiO <sub>2</sub>	20 mg	10% CH <sub>4</sub> 3% H <sub>2</sub> O vapor 10 mL min <sup>-1</sup> LA-251 Xe lamp, 260 °C	H <sub>2</sub> : 117 $\mu\text{mol g}^{-1} \text{min}^{-1}$	51
Au SA/TiO <sub>2</sub>	50 mg	CH <sub>4</sub> : CO <sub>2</sub> ~ 3 : 1 H <sub>2</sub> O 10 mL 300 W Xe lamp	H <sub>2</sub> : 2190 $\mu\text{mol g}^{-1}$	This work

cycling catalytic performance tests were conducted to study the lifetime of the catalyst. As shown in Fig. S6,<sup>†</sup> there are virtually no changes in the surface morphology of the sample before and after four recycling cycles. The activity of Au SA/TiO<sub>2</sub> during photocatalytic conversion remained almost stable, with a high yield of 2000  $\mu\text{mol g}^{-1}$  after four cycles (Fig. 7f). A slight decrease in product yield was observed, which was attributed to the reduction of catalyst efficiency during collection and washing processes between cycles.

To elucidate the reaction pathway and mechanism, the electron paramagnetic resonance analysis was carried out. As shown in Fig. 8, unlike under dark conditions, obvious signals

of radical oxygen species (ROS) were detected under light irradiation, which are responsible for reaction activation. Importantly, by conducting the spin-trapping test with the BMPO scavenger, the presence of the hydroxyl ( $\cdot\text{OH}$ ) radical was definitely confirmed, which played an important role in the CH<sub>4</sub> oxidation process.

Based on the energy band structure and the aforementioned photocatalytic methane oxidation performance, we propose a plausible reaction pathway for CH<sub>4</sub> on Au SA/TiO<sub>2</sub>. Under dark conditions, CH<sub>4</sub> adsorbs onto the catalyst surface. Upon illumination, electrons are excited and jump from the valence band to the conduction band, leaving holes behind. As a result,



**Fig. 8** (a) EPR spectra of the Au SA/TiO<sub>2</sub> sample under subsequential dark and light conditions at 90 K. (b) The spin trapping tests of Au SA/TiO<sub>2</sub> for detecting hydroxyl ( $\cdot\text{OH}$ ) radical formation by using 5-*tert*-butoxycarbonyl-5-methyl-1-pyrroline-*N*-oxide (BMPO) as a spin trapping agent.



dehydrogenation of CH<sub>4</sub> occurs on the catalyst surface, forming methyl radicals ( $\cdot\text{CH}_3$ ). Furthermore, during the photo-generated charge transfer process, Au plays a crucial role in preventing electron-hole pair recombination, which facilitates the oxidation to OH $\cdot$  and H $^+$ . In subsequent steps, the generated radicals combine with other species to produce H<sub>2</sub>, C<sub>2</sub>H<sub>6</sub>, and C<sub>3</sub>H<sub>8</sub>.<sup>52–54</sup>

## 4. Conclusion

In summary, we have prepared an atomically dispersed Au catalyst supported on TiO<sub>2</sub> nanostructures with a high loading capacity. Photocatalytic CH<sub>4</sub> oxidation measurements confirmed that the as-synthesized composite exhibited superior catalytic activity and long-term durability compared to pure TiO<sub>2</sub> under mild conditions. The main products, including H<sub>2</sub> and C<sub>2</sub>H<sub>6</sub>, reached a high yield of 2190  $\mu\text{mol g}^{-1}$  with approximately 58% selectivity after 8 hours of operation. We anticipate that this simple approach holds great promise for enhancing single-atom engineering without the formation of nanoclusters.

## Data availability

The data supporting this article have been included as part of the ESI.†

## Conflicts of interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

## References

- 1 M. Filonchik, M. P. Peterson, L. Zhang, V. Hurynovich and Y. He, *Sci. Total Environ.*, 2024, **935**, 173359.
- 2 P. Martin, I. B. Ocko, S. Esquivel-Elizondo, R. Kupers, D. Cebon, T. Baxter and S. P. Hamburg, *Energy Sci. Eng.*, 2024, **12**, 3995–4009.
- 3 Y. Jiang, S. Li, Y. Fan and Z. Tang, *Angew. Chem.*, 2024, **136**, e202404658.
- 4 M. C. Alvarez-Galvan, N. Mota, M. Ojeda, S. Rojas, R. M. Navarro and J. L. G. Fierro, *Catal. Today*, 2011, **171**, 15–23.
- 5 M. M. Zain and A. R. Mohamed, *Renewable Sustainable Energy Rev.*, 2018, **98**, 56–63.
- 6 P. Kumar, T. A. Al-Attas, J. Hu and M. G. Kibria, *ACS Nano*, 2022, **16**, 8557–8618.
- 7 Q. T. H. Ta, L. T. Nhiem, D. T. Y. Oanh, N. H. Hieu and P. K. T. Nguyen, *Vietnam J. Chem.*, 2024, DOI: [10.1002/VJCH.202400155](#).
- 8 N. Sun, X. Si, L. He, J. Zhang and Y. Sun, *Int. J. Hydrogen Energy*, 2024, **58**, 1249–1265.
- 9 M. J. Molaei, *Fuel*, 2024, **365**, 131159.
- 10 F. Mohamadpour and A. M. Amani, *RSC Adv.*, 2024, **14**, 20609–20645.
- 11 V. H. Dang, T. A. Nguyen, M. V. Le, D. Q. Nguyen, Y. H. Wang and J. C. S. Wu, *Chem. Eng. J.*, 2024, **484**, 149213.
- 12 A. Dong, Q. Jiang and Y. Zhou, *Int. J. Hydrogen Energy*, 2023, **48**, 29542–29551.
- 13 T. H. Nguyen, D. T. Tran, N. H. Kim and J. H. Lee, *Int. J. Hydrogen Energy*, 2023, **48**, 32294–32303.
- 14 H. Xu, Y. Zhao, G. He and H. Chen, *Int. J. Hydrogen Energy*, 2022, **47**, 14257–14279.
- 15 Y. Wang, F. Chu, J. Zeng, Q. Wang, T. Naren, Y. Li, Y. Cheng, Y. Lei and F. Wu, *ACS Nano*, 2021, **15**, 210–239.
- 16 S. S. A. Shah, T. Najam, M. S. Bashir, L. Peng, M. A. Nazir and M. S. Javed, *Energy Storage Mater.*, 2022, **45**, 301–322.
- 17 A. Kumar, V. Q. Bui, J. Lee, L. Wang, A. R. Jadhav, X. Liu, X. Shao, Y. Liu, J. Yu, Y. Hwang, H. T. D. Bui, S. Ajmal, M. G. Kim, S. G. Kim, G. S. Park, Y. Kawazoe and H. Lee, *Nat. Commun.*, 2021, **12**(1), 1–10.
- 18 S. Weon, M. J. Suh, C. Chu, D. Huang, E. Stavitski and J. H. Kim, *ACS ES&T Eng.*, 2021, **1**, 512–522.
- 19 D. Wang, B. Zhang, H. Ding, D. Liu, J. Xiang, X. J. Gao, X. Chen, Z. Li, L. Yang, H. Duan, J. Zheng, Z. Liu, B. Jiang, Y. Liu, N. Xie, H. Zhang, X. Yan, K. Fan and G. Nie, *Nano Today*, 2021, **40**, 101243.
- 20 S. Hejazi, H. Mehdi-pour, C. O. Otieno, J. Müller, S. Pour-Ali, M. Shahsanaei, S. Sarabadani Tafreshi, B. Butz, M. S. Killian and S. Mohajernia, *Int. J. Hydrogen Energy*, 2024, **51**, 222–233.
- 21 R. Rashid, I. Shafiq, M. R. H. S. Gilani, M. Maaz, P. Akhter, M. Hussain, K. E. Jeong, E. E. Kwon, S. Bae and Y. K. Park, *Chemosphere*, 2024, **349**, 140703.
- 22 Y. Zhao, Y. Kondo, Y. Kuwahara, K. Mori and H. Yamashita, *Appl. Catal., B*, 2024, **351**, 123945.
- 23 M. Chandra Sekhar, B. Purusottam Reddy, C. Kuchi, C. Kamal Basha, F. A. M. Al-Zahrani and R. Mangiri, *Ceram. Int.*, 2024, **50**, 38679–38687.
- 24 R. Kaveh, H. Alijani, E. Falletta, C. L. Bianchi, M. Mokhtarifar and D. C. Boffito, *Prog. Org. Coat.*, 2024, **190**, 108347.
- 25 N. Nair, V. Gandhi, A. Shukla, S. Ghotekar, V. H. Nguyen and K. Varma, *J. Phys.: Condens. Matter*, 2024, **36**, 413003.
- 26 J. Xie, R. Jin, A. Li, Y. Bi, Q. Ruan, Y. Deng, Y. Zhang, S. Yao, G. Sankar, D. Ma and J. Tang, *Nat. Catal.*, 2018, **1**(11), 889–896.
- 27 Z. Gong, L. Luo, C. Wang and J. Tang, *Sol. RRL*, 2022, **6**, DOI: [10.1002/SOLR.202200335](#).
- 28 S. Song, H. Song, L. Li, S. Wang, W. Chu, K. Peng, X. Meng, Q. Wang, B. Deng, Q. Liu, Z. Wang, Y. Weng, H. Hu, H. Lin, T. Kako and J. Ye, *Nat. Catal.*, 2021, **4**(12), 1032–1042.
- 29 A. Das, P. Dagar, S. Kumar and A. K. Ganguli, *J. Solid State Chem.*, 2020, **281**, 121051.
- 30 S. Wu, X. Tan, J. Lei, H. Chen, L. Wang and J. Zhang, *J. Am. Chem. Soc.*, 2019, **141**, 6592–6600.
- 31 T. Wang, J. Zhou, D. Li and Z. Ao, *Environmental Surfaces and Interfaces*, 2024, **2**, 19–25.
- 32 F. Ariaeinezhad, G. Mohammadnezhad, M. Zare, O. Akintola and W. Plass, *J. Mater. Chem. A*, 2024, **12**, 6488–6506.
- 33 T. D. N. Thi, X. H. Nguyen, H. V. Phung, M. T. Le and L. H. Nguyen, *Vietnam Journal of Science, Technology and Engineering*, 2023, **65**, 81–88.





- 34 N. Roy, K. Bhunia, C. Terashima, A. Fujishima and D. Pradhan, *ACS Omega*, 2017, **2**, 1215–1221.
- 35 Y. Du, M. Guo, Y. Chen, X. Mo, J. Cao and F. Hu, *Anal. Chim. Acta*, 2024, **1303**, 342462.
- 36 L. Zeng, C. Dai, B. Liu and C. Xue, *J. Mater. Chem. A*, 2019, **7**, 24217–24221.
- 37 S. Hejazi, H. Mehdi-pour, C. O. Otieno, J. Müller, S. Pour-Ali, M. Shahsanaei, S. Sarabadani Tafreshi, B. Butz, M. S. Killian and S. Mohajernia, *Int. J. Hydrogen Energy*, 2024, **51**, 222–233.
- 38 J. Chen, Y. Wanyan, J. Zeng, H. Fang, Z. Li, Y. Dong, R. Qin, C. Wu, D. Liu, M. Wang, Q. Kuang, Z. Xie and L. Zheng, *ACS Sustainable Chem. Eng.*, 2018, **6**, 14054–14062.
- 39 C. Zhang, T. Uchikoshi and T. Akashi, *Mater. Lett.*, 2024, **367**, 136604.
- 40 L. A. Al-Hajji, A. A. Ismail, A. Bumajdad, M. Alsaidi, S. A. Ahmed, F. Almutawa and A. Al-Hazza, *J. Mater. Sci.: Mater. Electron.*, 2021, **32**, 19764–19777.
- 41 F. E. Osterloh, *Chem. Mater.*, 2008, **20**, 35–54.
- 42 N. Lopez and J. K. Nørskov, *Surf. Sci.*, 2002, **515**, 175–186.
- 43 L. Torrente-Murciano, B. Solsona, S. Agouram, R. Sanchis, J. M. López, T. García and R. Zanella, *Catal. Sci. Technol.*, 2017, **7**, 2886–2896.
- 44 N. González-Ballesteros, P. M. Martins, C. J. Tavares and S. Lanceros-Méndez, *J. Ind. Eng. Chem.*, 2025, **143**, 526–537.
- 45 F. Boccuzzi, A. Chiorino, M. Manzoli, D. Andreeva and T. Tabakova, *J. Catal.*, 1999, **188**, 176–185.
- 46 W. Li, Y. Ren, Z. Xie, Y. Wang, H. Zhang, D. Peng, H. Shen, H. Shi, J. Cai, P. Wang, T. Zhang and Z. Zhao, *Nano Res.*, 2024, **17**, 3844–3852.
- 47 Q. Li, H. Wang, M. Zhang, G. Li, J. Chen and H. Jia, *Angew. Chem., Int. Ed.*, 2023, **62**, e202300129.
- 48 Y. Sun, C. Bo, Z. Cheng, X. Zhang, J. Liu and L. Piao, *Nano Res.*, 2023, **16**, 12942–12948.
- 49 M. Kushida, A. Yamaguchi and M. Miyauchi, *J. Energy Chem.*, 2022, **71**, 562–571.
- 50 Q. Shen, C. Cao, R. Huang, L. Zhu, X. Zhou, Q. Zhang, L. Gu and W. Song, *Angew. Chem., Int. Ed.*, 2020, **59**, 1216–1219.
- 51 H. Song, X. Meng, Z. J. Wang, Z. Wang, H. Chen, Y. Weng, F. Ichihara, M. Oshikiri, T. Kako and J. Ye, *ACS Catal.*, 2018, **8**, 7556–7565.
- 52 X. Zhang, Y. Wang, K. Chang, S. Yang, H. Liu, Q. Chen, Z. Xie and Q. Kuang, *Appl. Catal., B*, 2023, **320**, 121961.
- 53 X. Lu, H. Luo, B. Xu, Z. Liu, Y. Cao, K. Li, X. Yang, L. Xie, T. Guan, W. Zhu and Y. Zhou, *Chem. Commun.*, 2024, **60**, 14276–14279.
- 54 T. M. Vo, T. M. H. Nguyen and C. W. Bark, *ACS Appl. Electron. Mater.*, 2024, **6**, 4391–4405.

