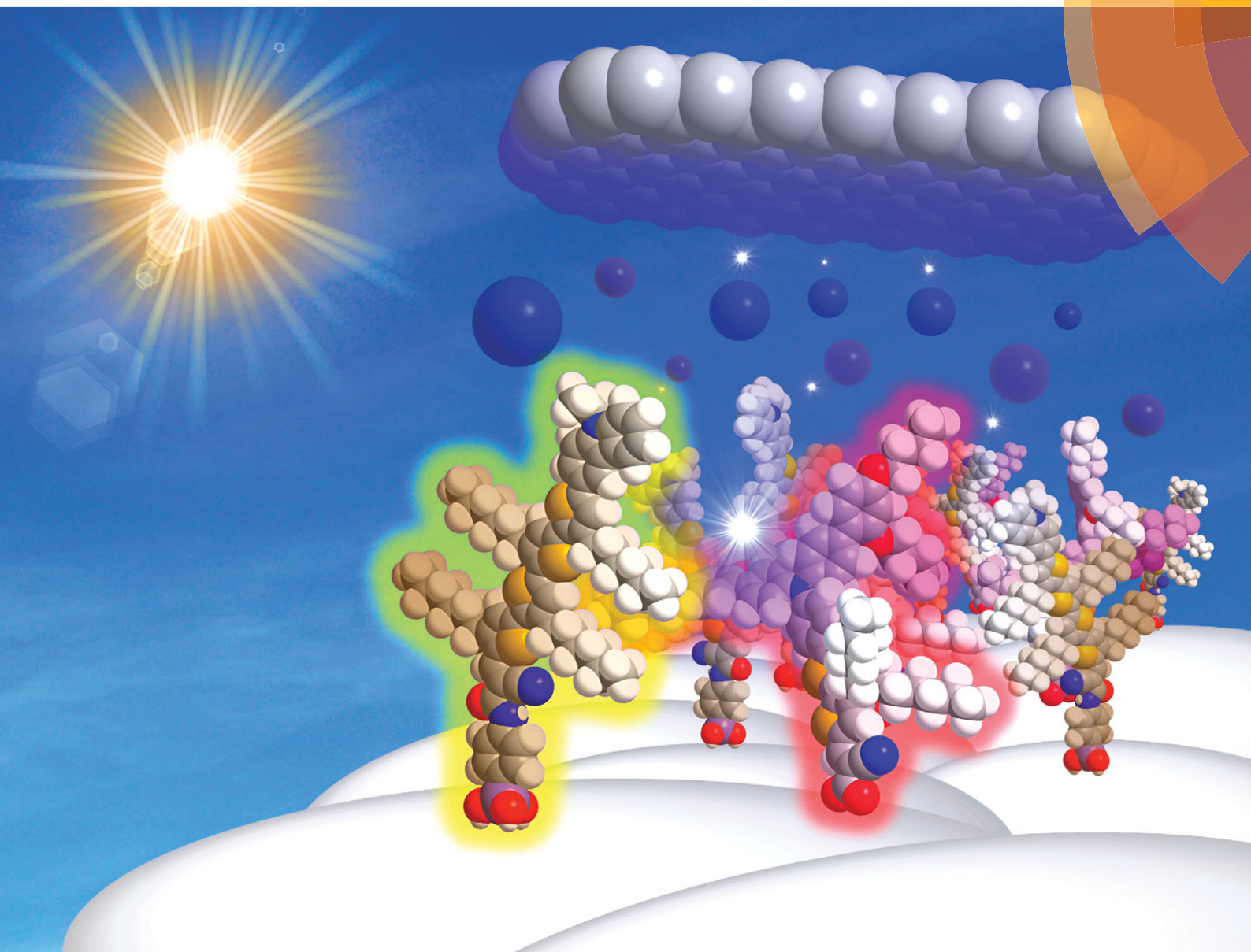


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Highly-efficient dye-sensitized solar cells with collaborative sensitization by silyl-anchor and carboxy-anchor dyes†

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In dye-sensitized solar cells co-photosensitized with an alkoxy-silyl-anchor dye ADEKA-1 and a carboxy-anchor organic dye LEG4 was revealed to work collaboratively by enhancing the electron injection from the light-excited dyes to the TiO₂ electrodes, and the cells exhibited a high conversion efficiency of over 14% under one sun illumination.

Dye-sensitized solar cells (DSSCs), which are composed of mesoporous nanocrystalline-TiO₂ thin layers modified with photosensitizing dyes as working electrodes, redox electrolytes and counter electrodes, have been actively investigated as photovoltaic devices in the next generation of alternatives to conventional silicon-based inorganic solar cells (Fig. S1, ESI†), because of their potentially low production costs, shorter energy and CO₂ payback times, low toxicity of the constituent elements and relatively high light-to-electric energy conversion efficiencies (η) especially under low-light intensities and scattered light conditions.^{1–4} In DSSCs, η values of 11–13% under the simulated sunlight of one sun have been reported up to now through photosensitization using polypyridyl and porphyrin complexes of metals such as ruthenium or zinc, and a few metal-free organic dyes with carboxy-anchor moieties for binding to the surface of the TiO₂.^{1–9}

Organosilicon compounds such as silanols and alkoxy-silanes have high metal-oxide surface bonding abilities by forming strong Si–O–metal bonds. While paying attention to the characteristics of silanols and alkoxy-silanes, we have focused on the development of photosensitizing dyes for DSSCs possessing silyl-anchor moieties,^{10–12} and recently we succeeded in achieving over 12% conversion efficiency in cells using a carbazole/alkyl-functionalized oligothiophene/alkoxy-silyl-anchor moiety type compound, ADEKA-1 (Fig. 1a), as the photosensitizer.¹³

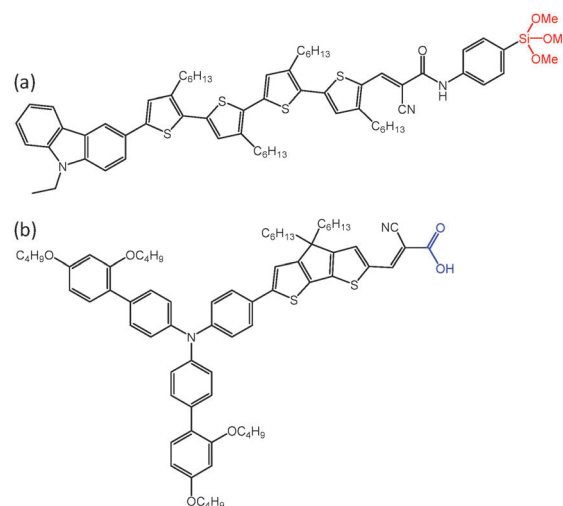


Fig. 1 Molecular structures of photosensitizing dyes: (a) silyl-anchor dye ADEKA-1 and (b) carboxy-anchor dye LEG4.

Besides the high photovoltaic performance, the TiO₂ photoelectrode sensitized with ADEKA-1 possesses much higher durability to solvents, *e.g.* nitrile, water and mixtures of them, and to surface modification using wet processes than those sensitized with carboxy-anchor dyes. The durability of the photoelectrode allows the co-adsorption of another sensitizing dye to the electrode for production of a co-sensitization effect, and actually we succeeded in improving the η to 12.8% by means of the co-sensitization of ADEKA-1 and a silyl-anchor coumarin dye SFD-5.¹⁴

For further improvement in the efficiency of ADEKA-1-sensitized DSSCs, we expanded the study of co-sensitizers for the cells to widely developed carboxy-anchor dyes, which have been demonstrated to possess high sensitizing properties as photosensitizers in DSSCs. In the investigation we found that the ADEKA-1-sensitized cells with the co-sensitizer LEG4¹⁵ (Fig. 1b) exhibited a considerably higher photovoltaic performance through collaborative sensitization by the dyes, and succeeded in achieving

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over 14% conversion efficiency in the cells. The efficiency shows the high potential of DSSCs to be practical light-to-electric energy conversion devices in the near future.

As co-sensitizers for the **ADEKA-1**-sensitized DSSCs, we selected carboxy-anchor organic sensitizing dyes which have been reported to have high sensitizing properties and an absorption band in the shorter wavelength region than **ADEKA-1**, *i.e.* **LEG4**, **D35**, **L0** and **D131** (Fig. S2–S6 and Table S1, ESI[†]). To check the potential of these dyes as co-sensitizers for **ADEKA-1**, we fabricated cells sensitized by **ADEKA-1** and by **ADEKA-1** with the dyes using an electrolyte solution containing an I_3^-/I^- redox mediator (cell-A; the fabrication procedures of the cells are described in the ESI[†]). Among the cells, a significant and the largest improvement in the incident monochromatic photon-to-current conversion efficiency (IPCE) was observed in the cell photosensitized by **ADEKA-1** with **LEG4** (Fig. 2a and Fig. S7 in the ESI[†]). The cell exhibited much

higher IPCE values of close to 90% compared to the cells sensitized only by **ADEKA-1** and only by **LEG4** in all of the visible region. The increase of the open-circuit photovoltage (V_{oc}) and the short-circuit photocurrent density (J_{sc}) compared to those of the cell sensitized only by **ADEKA-1** resulted in the improvement of the η by a factor of 1.3 under simulated sunlight at one sun (AM-1.5G, 100 mW cm⁻²; Table S2, ESI[†]).

In the cell photosensitized by **ADEKA-1** with **LEG4**, in which the relative amount of the dyes adsorbed on the TiO₂ electrode was estimated to be 1.0:0.25 for **ADEKA-1**:**LEG4**, the improvement of the IPCE values compared to those of the **ADEKA-1**-sensitized cell was observed not only in the light-absorption wavelength region of **LEG4** but also in the longer wavelength region where the light absorption by **LEG4** was absent (Fig. S8, ESI[†]), which is different to the other co-sensitized cells. In order to clarify the origin of the peculiar and large improvement in the IPCE by the co-sensitization with **LEG4**, we examined MO calculations of the dyes (Fig. S9 and S10, and Tables S3 and S4 in the ESI[†]).

The light-to-electric energy conversion in DSSCs proceeds by light excitation of the sensitizing dye followed by charge separation produced through electron injection from the LUMO of the light-excited dye to the conduction band of TiO₂. In **ADEKA-1** the alkoxy-silyl-anchor moiety links to the chromophore (carbazole/alkyl-functionalized oligothiophene moiety) *via* the phenyl-amide moiety, and the LUMO has a small electron distribution around the silyl-anchor moiety. On the other hand, the LUMO of **LEG4** has a large electron distribution around the carboxy-anchor moiety and thus **LEG4** is expected to have a higher electron injection efficiency from the LUMO to the TiO₂ conduction band than **ADEKA-1** (Fig. S9, ESI[†]). When comparing the energy levels of the LUMOs of the dyes, only **LEG4** has a lower LUMO than **ADEKA-1** which is different to the other dyes, **D35**, **L0** and **D131** (Fig. S6, ESI[†]), and emission analyses using an Al₂O₃ porous film modified by **ADEKA-1** with **LEG4** showed that the emission from **ADEKA-1** was quenched almost completely by the existence of **LEG4** as the co-adsorbent (Fig. S11, ESI[†]). From the MO properties and the results of the emission analyses, the large improvement of the photovoltaic performance in the cell photosensitized by **ADEKA-1** with **LEG4** is considered to be brought about by the collaborative sensitization by the dyes through an electron-injection enhancement effect due to the existence of the **LEG4** molecules near the **ADEKA-1** molecules on the TiO₂ electrode; the electron transfers from the light-excited **ADEKA-1** to the co-adsorbent **LEG4** and immediate electron injection occurs from **LEG4** to the conduction band of the TiO₂ with much higher efficiency than the direct electron injection from the light-excited **ADEKA-1** (Fig. 2b). Internal quantum efficiency (IQE) measurements revealed a considerably higher electron injection efficiency in the cell photosensitized by **ADEKA-1** with **LEG4**, and the maximum IQE was evaluated to be 99 ± 2% (Fig. S12, ESI[†]). The increase of the V_{oc} (Table S2, ESI[†]), the decrease of the dark-current (Fig. S13, ESI[†]) and the elongation of the electron lifetime in the TiO₂ conduction band estimated from the transient open-circuit voltage decays (Fig. S14, ESI[†]) observed in the co-sensitization with **LEG4** indicate that the

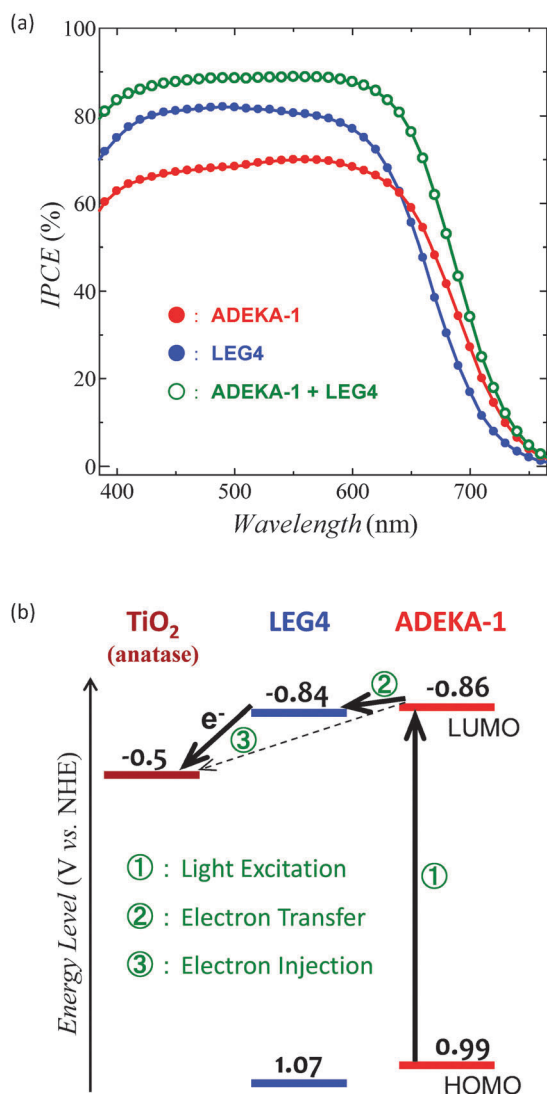


Fig. 2 (a) IPCE spectra of the cells photosensitized by **ADEKA-1**, by **LEG4** and by **ADEKA-1** with **LEG4** (cell-A) and (b) a schematic drawing of the charge separation processes for the TiO₂ electrode sensitized collaboratively by **ADEKA-1** and **LEG4**.

Table 1 Photovoltaic parameters of the cells sensitized collaboratively by **ADEKA-1** and **LEG4** (cell-B) under the illuminations of the simulated sunlight (AM-1.5G)

Entry	Electrolyte:redox ^a	Counter electrode	Light intensity (mW cm ⁻²)	<i>J</i> _{sc} (mA cm ⁻²)	<i>V</i> _{oc} (V)	FF	<i>η</i> (%)
1	A:I ₃ ⁻ /I ⁻	FTO/Pt	100	19.11	0.783	0.748	11.2
2	F:[Co(phen) ₃] ^{3+/2+}	FTO/Pt	100	17.77	1.018	0.765	13.8
3 ^b	F:[Co(phen) ₃] ^{3+/2+}	FTO/Au/GNP	100	18.27	1.014	0.771	14.3
4	F:[Co(phen) ₃] ^{3+/2+}	FTO/Au/GNP	50	9.55	0.994	0.776	14.7

^a Electrolyte: (A) 0.07 M I₂, 0.05 M LiI, 0.05 M NaI, 0.50 M DMPImI, 0.10 M EMImI, 0.05 M TBAl, 0.05 M THAl, 0.40 M TBP, 0.10 M MP, and 0.10 M GuSCN in MeCN/VN/THF (8 : 1 : 1 in volume); (F) 0.20 M [Co²⁺(phen)₃](PF₆⁻)₂, 0.05 M [Co³⁺(phen)₃](PF₆⁻)₃, 0.07 M LiClO₄, 0.02 M NaClO₄, 0.03 M TBAPF, 0.01 M TBPPF, 0.01 M HMImPF, 0.30 M TBP, 0.10 M TMSP, 0.10 M MP, 0.05 M CPrBP, 0.10 M CPeBP, and 0.05 M CCoBP in MeCN. The data for the cells with other electrolytes (B–E) are listed in Table S6 (ESI). ^b The values are the averages of the results of the four cells which were prepared separately (Table S5, ESI).

adsorbed **LEG4** on the TiO₂ electrode also works as a suppressor, preventing back electron transfer from the TiO₂ electrode to the electrolyte by covering the naked surface of the TiO₂ electrode with its plural alkyl-chain substituents.^{15–18} By using an I₃⁻/I⁻ redox electrolyte solution with an experimentally optimized composition, the cell photosensitized collaboratively by **ADEKA-1** and **LEG4** (cell-B; the fabrication procedures of the cell are described in the ESI†) exhibited the *η* of 11.2% under AM-1.5G one sun illumination (entry 1 in Table 1).

The maximum photovoltage (*V*_{max}) obtained in the DSSC is attributed to the energy gap between the quasi-Fermi level of the TiO₂ [approximately the energy level of the conduction-band edge (*E*_{C.B.})] and the redox potential of the electrolyte, and the improvement of the efficiency of DSSCs is possible by the increase of the photovoltage through using an electrolyte having a more positive (lower) redox potential than I₃⁻/I⁻.^{1–4,7–9,11,13–15} The redox potential of the cobalt(III/II) tris(1,10-phenanthroline) complex ([Co(phen)₃]^{3+/2+}) is lower than that of I₃⁻/I⁻ by ca. 0.2 V,¹⁹ and the values of the HOMO levels of **ADEKA-1** and **LEG4** are still more positive than the redox potential of the cobalt(III/II) complex (Fig. S15, ESI†), which provides the thermodynamic driving force for the dye regeneration reaction by electron transfer from the Co²⁺-complex electrolyte to the oxidized dye.^{13–15} Thus we employed [Co(phen)₃]^{3+/2+} as the redox electrolyte for the **ADEKA-1** and **LEG4** co-sensitized cell for further improvement of the *η* of the cells.

In the fabrication of the cells using the cobalt(III/II) complex redox electrolytes (cell-B), the compositions of the electrolyte solutions, *i.e.* the Co²⁺/Co³⁺ ratio, the kind of cobalt(III/II) complex counter anion and the electrolyte additives, were optimized experimentally according to the literature^{7,13,14,20,21} using a platinum-deposited F-doped SnO₂ (FTO)-coated glass plate as the counter electrode. The cell using an electrolyte solution with the optimized composition exhibited a high *V*_{oc} of above 1 V and the *η* was improved to 13.8% under AM-1.5G one sun illumination (entry 2 in Table 1) as was expected from the more positive redox potential of [Co(phen)₃]^{3+/2+}. However, a decrease of the *J*_{sc} was also observed in the cell compared to the cell with the I₃⁻/I⁻ redox electrolyte solution. In order to recover the *J*_{sc}, we employed graphene nanoplatelets (GNPs) as the material for the counter electrode and prepared the counter electrode on a FTO-coated glass plate with a structure of FTO/Au/GNP, because the counter electrode has been reported to produce

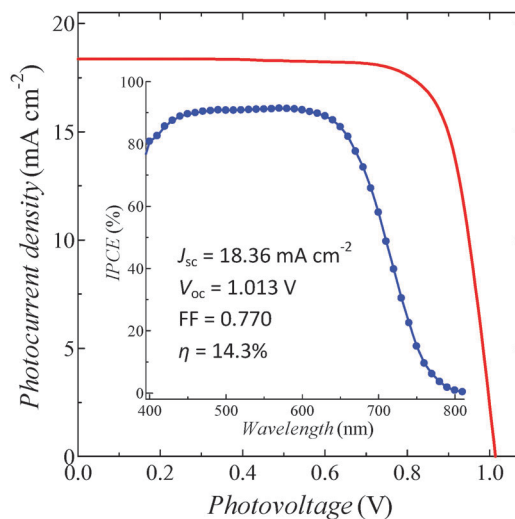


Fig. 3 A typical *J*–*V* curve of the cell photosensitized collaboratively by **ADEKA-1** and **LEG4** with an efficiency of over 14% (entry 3a in Table S5, ESI†) under the illumination of simulated sunlight (AM-1.5G, 100 mW cm⁻²). Inset shows the IPCE spectrum of the cell.

higher *J*_{sc} and fill factor (FF) in photocurrent–voltage properties than the standard platinum electrodes.^{8,22,23} Fig. 3 shows an example of the *J*–*V* curve under AM-1.5G one sun illumination (100 mW cm⁻²) and the IPCE spectrum of the cell co-sensitized by **ADEKA-1** and **LEG4**. The photovoltaic parameters, assessed as the averaged values from the *J*–*V* curves of the four separately prepared cells, are listed in Table 1 as entry 3 (Table S5, ESI†). The *J*_{sc} was actually improved in the cell from 17.8 to 18.3 ± 0.1 mA cm⁻² by using the FTO/Au/GNP counter electrode and the maximum value in the IPCE spectrum reached 91%, resulting in the *η* of 14.3% with the *V*_{oc} above 1 V. The better photovoltaic performance in the lower light intensity is a characteristic of DSSCs. This is also observed in the present cell and the cell exhibited an *η* of close to 15% under simulated sunlight with a 50 mW cm⁻² intensity (entry 4 in Table 1, and Fig. S16 and S17 in the ESI†).

In conclusion, a carboxy-anchor organic dye **LEG4** was revealed to work effectively as the collaborative sensitizer to the silyl-anchor dye **ADEKA-1** in DSSCs, and we succeeded in obtaining a high IPCE of up to 91%, *V*_{oc} of above 1 V and 14.3% conversion efficiency in the cell with the optimized cobalt(III/II) complex redox electrolyte solution and the GNP counter electrode.

The result is attributed basically to the strong adsorption properties of ADEKA-1 to the TiO₂ electrode and shows the validity of silyl-anchor dyes as photosensitizers for DSSCs. The observation of conversion efficiencies of over 14% in these DSSCs indicates the high potential of DSSCs as light-to-electric energy conversion devices. The collaborative sensitization by plural organic dyes including silyl-anchor dyes, which would bring a further improvement in the photovoltaic performance of DSSCs, is considered as a promising way to produce practical DSSCs.

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