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Anion exchange membrane water electrolysis over superparamagnetic ferrites†

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The oxygen evolution reaction (OER) is usually the bottleneck in water electrolysis due to its sluggish kinetics, resulting in increased costs in the production of green hydrogen. Therefore, there is a need for more efficient, stable, and ideally, critical-raw-material-free catalysts. To this end, we have synthesized nanosized spinel ferrites CoFe_2O_4 , NiFe_2O_4 , and ZnFe_2O_4 , and a high-entropy spinel ferrite $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ through a simple coprecipitation reaction in an automated reactor on a gram scale. The powder X-ray diffraction and transmission electron microscopy studies revealed crystallite sizes of 20–35 nm. Insight into the oxidation states and cation distribution in the mixed spinel systems was gained through X-ray photoelectron and Mössbauer spectroscopy studies. The activity of all spinel ferrites was tested for the OER through half-cell laboratory measurements and full-cell anion exchange membrane electrolysis (AEMEL), where $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ showed the lowest overpotential of 432 mV at a current density of 10 mA cm^{-2} . All the synthesized ferrites demonstrated good stability up to 20 h, with NiFe_2O_4 being the most active in high current density experiments up to 2 A cm^{-2} . In addition, studies on the magnetic properties at room temperature revealed a largely superparamagnetic response of the prepared materials, indicating that quantum spin-exchange interactions facilitate oxygen electrochemistry. Computational calculations shed light on the superior catalytic activities of NiFe_2O_4 and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$, the two strongly correlated oxides that exhibit the highest magnetization and the smallest band gaps, corroborating the recent principles determining the activity of magnetic oxides in electron transfer reactions.

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Introduction

Hydrogen is an important energy vector with regard to achieving the targeted net-zero CO_2 emissions by 2050, allowing for the elimination of our dependence on fossil fuels.¹ When powered by renewable electricity, water electrolysis is of profound and long-standing interest for green H_2 production.^{2,3} In the last few decades, the industry has focused on two electrolysis systems that operate at low temperatures. Traditional alkaline water electrolysis (ALK) typically operates in the temperature range of 60–80 °C with potassium hydroxide as the electrolyte, Ni-based electrodes, and a diaphragm (e.g., Zirfon).^{4,5} The usual maximum operating geometric current density is less than 0.4 A cm^{-2} , with an energy efficiency of about 60%.⁶ ALK is recognized as a mature technology, which is commercially available due to its low cost. However, its major limitation is related to its poor dynamic operation capability, e.g., the ability of the electrolysis system to rapidly adjust its operation to match the fluctuating power output of renewable

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energy sources.⁷ Proton exchange membrane water electrolysis (PEMEL) is the other key electrolysis technology in the near-term. PEMEL is advantageous in that it uses a semi-solid electrolyte in the form of proton exchange membranes (PEMs) with high proton conductivity (e.g., Nafion),⁸ pressurized set-up, and good operation under dynamic load.^{4,9} Despite its inherent advantages, PEMEL is defined by high CAPEX, due to the need for rare and expensive platinum-group-metal (PGM) electrodes,¹⁰ along with costly titanium gas-diffusion layers¹¹ and bipolar plates.¹²

The relatively recent anion exchange membrane electrolysis (AEMEL) offers a great opportunity to generate green H₂ in a cost-efficient and environmentally benign fashion. The use of anion exchange membranes (AEMs) may mitigate some of the limitations of ALK. Similar to PEMEL, AEMEL is characterized by pressurized operation and high dynamic responsiveness, but it allows for the use of cheap, non-PGM catalysts and low-cost cell/stack materials.^{13,14} For instance, various hydroxide/oxyhydroxide/oxide nanoparticles (NPs) based on abundant non-PGM 3d transition metals, such as Mn, Fe, Co, Ni, Cu, Zn, and Mo, are commonly investigated and employed as catalysts for the alkaline oxygen evolution reaction (OER) over AEMEL anodes. The fact that these types of non-PGM catalysts are more abundant and less costly than PGM-containing ones, while also featuring good chemical stability under alkaline conditions, shows the potential for the large-scale implementation of AEMEL.^{15–18} Despite great efforts in the recent years to develop non-PGM catalysts, a systematic analysis regarding their physicochemical properties such as the effect of composition, particle size, surface area, and physical properties is of high relevance for the future implementation of AEMEL technology at the industrial level.⁴

The efficiency of AEMEL relies highly on that of the OER due to its slower reaction kinetics and associated high overpotentials as compared to the hydrogen evolution reaction (HER), the other half-reaction of water electrolysis.^{19–21} Notably, the most active catalysts for the OER have strongly correlated electrons,²² associated with itinerant spin conduction.^{23,24} For example, ferrites AFe₂O₄ with a spinel structure are strongly correlated oxides with favorable spin-orbital orderings, and thus these have been postulated as promising catalysts for the OER.²⁵ High activity and durability of ferrites are commonly associated with their open-shell electronic configurations. Moreover, the structural and chemical stability of ferrites under alkaline conditions combined with their multiple cation valences (*i.e.*, A³⁺/A²⁺ and Fe³⁺/Fe²⁺) renders them interesting catalysts for AEMEL anodes.^{14,26–30} While several ferrites with various compositions, structures, and nanostructures have been synthesized, characterized, and investigated in the laboratory in a half-cell alkaline OER experimental setting,³¹ their performance in industrially relevant full-cell AEMEL has not been evaluated in great detail to date.

Our laboratories have been exploring the chemical synthesis of ferrite NPs to modulate their structural and magnetic properties.^{32,33} Herein, we report the preparation of a series of ferrite AFe₂O₄ (A = Fe²⁺, Ni²⁺, Co²⁺, Zn²⁺, and Mn²⁺) NPs on a

gram scale and their performance in alkaline OER, and importantly, in full-cell AEMEL. The spin-dependent theory in catalysis²² and computational calculations provided comprehensive insight into the structural, magnetic, and electronic properties of the prepared ferrites, leading to structure–property relationships that will assist in the future design of active catalysts.

Experimental section

Materials

Iron(III) chloride hexahydrate (FeCl₃·6H₂O, >98%, Sigma-Aldrich), nickel(II) chloride hexahydrate (NiCl₂·6H₂O, >98%, Thermo Scientific), cobalt(II) chloride (CoCl₂·6H₂O, >98%, Thermo Scientific), zinc(II) chloride anhydrous (ZnCl₂, >98%, Thermo Scientific), manganese(II) chloride tetrahydrate (MnCl₂·4H₂O, 99%, Thermo Scientific), sodium hydroxide (NaOH, pellets, EMSURE[®], Supelco), potassium hydroxide (KOH, flakes, 90%, Sigma-Aldrich), 30% hydrogen peroxide solution (H₂O₂, Sigma-Aldrich), potassium permanganate (KMnO₄, ≥99.0%, Sigma-Aldrich), isopropanol (≥99.5%, Sigma-Aldrich), carbon black Vulcan XC 72 (Fuel Cell Store) and ethanol (≥99.5%, Honeywell) were used as received. Ultrapure water (18.2 MΩ cm) was produced using a Milli-Q Advantage A10 system (Millipore).

Synthesis of ferrites

All ferrites, namely, CoFe₂O₄, NiFe₂O₄, and ZnFe₂O₄, as well as high-entropy Zn_{0.2}Mn_{0.2}Ni_{0.2}Co_{0.2}Fe_{2.2}O₄ ferrite [4A_{1/4}Fe_{2.2}O₄] were prepared using the robust coprecipitation method adapted from elsewhere¹² using NaOH as a precipitating agent. Typically, the molar ratio of base to metal precursors was fixed as 10:1. The reproducible gram-scale synthesis of the ferrites (>5 g) was achieved using a 2 L automated Atlas potassium synthesis system (Syriris) equipped with two syringe pumps (Syriris), a temperature control system LH85 PLUS (Julabo), an overhead stirrer, and a temperature probe. In a typical reaction, 11.0 g of NaOH (0.225 mol) was dissolved in 400 mL of ultrapure water while heating at 80 °C under mechanical stirring of 300 rpm for 30 min. This step was followed by a controlled addition of the metal precursors previously dissolved in 200 mL of ultrapure water at a feeding rate of 5 mL min⁻¹. Typically, for two-metal ferrites, 42 mmol of Fe(III) salt and 21 mmol of the corresponding A(II) salt were mixed and used as the precursor solution. For the synthesis of the high-entropy ferrite containing five metals, an aqueous solution was prepared with 42 mmol of Fe(III) salt and 4.2 mmol of each M(II) salt. After complete addition of the precursor solution to the NaOH solution in a 2 L reactor, the reaction mixture was allowed to gradually cool to room temperature. The resulting solid products were either collected magnetically with the aid of a permanent NdFeB magnet for magnetically active products or through centrifugation at 12 000 rpm for 10 min for weakly magnetic products. The resulting ferrites were extensively washed with water and ethanol and dried under vacuum. The obtained powders were calcined at 400 °C in air with a heating rate of 3 °C min⁻¹ and holding time of 5 h.

The calcination products were converted into fine powders through ball milling in an 8000 M high-energy mixer/mill (SPEX) using a zirconia vial and spheres at a sample-to-sphere *w/w* ratio of 9:1 (1 cycle at 1425 rpm of 30 min).

Characterization

The materials were characterized by powder X-ray diffraction (XRD, X'Pert PRO diffractometer, PANalytical), Raman spectroscopy (alpha300 R confocal microscope, WITec), Mössbauer spectroscopy (MS4 spectrometer, SEE Co.), transmission electron microscopy (TEM, JEM-2100 microscope, Jeol), high-angle annular dark field scanning TEM (HAADF-STEM), energy-dispersive X-ray spectroscopy (EDX) in STEM mode (STEM-EDX, Titan ChemiSTEM microscope at 200 kV; FEI, 0.08 nm point resolution, Super-X EDX System), N₂ physisorption at 77 K (Autosorb IQ2 apparatus, Quantachrome), scanning electron microscopy (SEM, Quanta 650 FEG microscope, FEI), energy dispersive X-ray spectroscopy (EDX, INCA 350 spectrometer, Oxford Instruments), vibrating sample magnetometry (VSM, EV9 instrument, Lot-Oriel), UV/vis diffuse reflectance spectroscopy (LAMBDA 950 UV/Vis/NIR spectrophotometer, PerkinElmer), X-ray photoelectron spectroscopy (XPS, ESCALAB 250 Xi, Thermo Scientific), and inductively coupled plasma-optical emission spectroscopy (ICP-OES, ICPE-9000 spectrometer, Shimadzu); all powders were digested using a concentrated HCl 37% solution. The NiFe₂O₄ nanoparticles were further digested under heat treatment at 170 °C using an autoclave reactor.

Half-cell alkaline OER measurements

The electrochemical OER activity of the synthesized ferrites was studied by cyclic voltammetry (CV) and linear sweep voltammetry (LSV) in a conventional three-electrode glass cell. Before use, the glassware was cleaned following a standard cleaning procedure: first, the glassware was submerged overnight in a saturated KMnO₄ solution. Next, the glassware was rinsed with a dilute H₂O₂ solution and then boiled three times in ultrapure water (5 minutes each time). The working electrode consisted of a catalyst ink drop-casted on a glassy carbon (GC) rotating disk electrode (RDE) (Metrohm, $\varnothing = 3$ mm, $S = 0.0071$ cm²). The counter electrode and reference electrode were a GC rod and a Hg|HgO electrode, respectively. All potentials E in the half-cell study are quoted with respect to the reversible hydrogen electrode (RHE), converted *via* $E_{\text{RHE}} = E(\text{RHE}) = E(\text{Hg/HgO}) + 0.059 \times \text{pH} + E_0(\text{Hg/HgO}) = E(\text{Hg/HgO}) + 0.924$ V.³⁴ The applied potential was controlled using an Autolab PGSTAT302N potentiostat (Metrohm) while the RDE rotation rate was controlled using an Autolab RRDE rotator (Metrohm).

The catalyst ink was prepared by mixing ultrapure water, isopropanol (volume ratio of water/isopropanol = 3:1), Sustainion ionomer solution (5% in solvent, Dioxide Materials) at a 5% volume ratio to the overall volume, a physical mixture of the synthesized ferrite powder (8 mg mL⁻¹), and conductive carbon black (16 mg mL⁻¹). Before drop-casting the catalyst ink, the GC RDE was polished with alumina powder for 5 min, and then rinsed with water and ethanol. This procedure was repeated three times. After ultrasonication treatment for 15 min, the as-

formulated catalyst ink was drop-casted on the GC RDE, and then left in a solvent-vapor-saturated water + isopropanol atmosphere overnight for drying. The mass loading (0.3 mg cm⁻² ferrite + 0.6 mg cm⁻² conductive carbon black) was precisely controlled by the volume of the drop-casted ink.

Aqueous 1 M NaOH electrolyte solution was used, and all experiments were performed under N₂-saturated (5.0, Linde-Gas) conditions at room temperature. Before electrochemical characterization, the working electrode containing the catalyst was carefully rinsed with ultrapure water, followed by electrochemical cleaning for 15 min by performing CV from 0 to 1.2 V_{RHE} at a scan rate of 200 mV s⁻¹. All data shown are *iR*-corrected for a voltage drop of 80%. The electrochemical surface area (ECSA) of the working electrode was determined from the double-layer capacitance (C_{dl}) measured at 1 V_{RHE}, assuming a uniform specific double layer of 40 μF cm⁻².

OER properties of the synthesized ferrites in KOH electrolyte were also investigated in non-RDE mode at a higher catalyst mass loading of 1 mg cm⁻² on a 1 cm² glassy carbon plate. The catalyst ink was prepared by dispersing a mixture of 4 mg of catalyst and 1 mg of conductive carbon black in a 3:2 ethanol/water solution while adding 60 μL of Nafion ionomer solution (5% in aliphatic alcohols and water, Sigma-Aldrich) as a binder, to a total volume of 1 mL.

Full-cell AEMEL testing

Full-cell testing was conducted in an in-house-built AEMEL electrolyzer using a cell with a 4 cm² active area (Fuel Cell Store). The catalyst was coated on the gas diffusion/porous transport layers (GDL/PTL) using the catalyst coated-on-substrate method, thus affording gas diffusion/porous transport electrodes (GDE/PTE). The cathode GDE was carbon paper GDL SGL 39BC (Sigracet) loaded with 0.3 mg cm⁻² of 40%Pt/C catalyst (Premetek), while the anode PTE was nickel PTL 2Ni 18-0.5 (Bekaert) loaded with 3 mg cm⁻² of the synthesized ferrite catalysts. Proprietary Membrane B 70 μm (HYDROLITE) was employed as the AEM, and PTFE gaskets of 200 μm and 400 μm were used for sealing the cathode and the anode, respectively. After the assembly, the AEMEL cell was connected to the electrolyte inlets and outlets, thermocouple, as well as voltage and current sensors in the testing workstation. The 1 M KOH electrolyte was supplied on the OER side at a flow rate of 170 mL h⁻¹. Membrane electrode assembly (MEA) activation was performed by gradually increasing the voltage from 1.4 to 2 V to achieve the maximum current density for the cell at a specific temperature. All catalysts were tested at both 60 °C and 80 °C. After activation, the current-voltage (I - V) curve was measured from 1.4 to 2 V and backwards. The short-term durability of the AEMEL cells was evaluated through chronoamperometric measurements under a constant current density of 0.5 A cm⁻² for 20 h.

Results

CoFe₂O₄, NiFe₂O₄, ZnFe₂O₄, and high-entropy Zn_{0.2}Mn_{0.2}Ni_{0.2}-Co_{0.2}Fe_{2.2}O₄ were prepared following a straightforward

coprecipitation method using NaOH as the precipitating agent. We avoided the use of aqueous ammonia solution for precipitation because of the possible formation of soluble complexes with several transition metals (e.g., Ni and Zn). An automated benchtop synthesis system was employed, which allows for the synthesis of NPs on a gram scale in a highly controlled and reproducible fashion.^{32,35} The as-synthesized products were calcined at a moderate temperature to clean the surface of the catalysts and remove any residuals from the synthesis. Finally, the catalysts were subjected to ball milling to decrease their degree of agglomeration to facilitate the formulation of catalytic inks.

Powder XRD patterns of the catalysts (Fig. S1, ESI[†]) evidenced all final products to be single-phase nanocrystalline ferrite powders exhibiting a cubic spinel structure (space group $Fd\bar{3}m$, no. 227). No diffraction peaks corresponding to secondary phases were detected. The average size of the crystallites and reliable unit cell parameters were estimated from the collected XRD data and are summarized in Table S1 (ESI[†]). Next, Raman spectra were recorded for the prepared catalysts (Fig. S2, ESI[†]), which exhibited broad peaks with a low signal-to-noise ratio, which is common for nanocrystalline powders. The catalysts displayed typical Raman band profiles associated with the spinel structure of the O^7h ($Fd\bar{3}m$) space group while featuring characteristic Raman-active bands arising from A_{1g} ($\gg 690$ cm^{-1}), T_{2g} ($\gg 470$ cm^{-1}), and E_g ($\gg 320$ cm^{-1}) modes.^{36–39} As expected, the different elemental compositions of the prepared ferrites were mostly reflected in the small shifts and splitting of the observed Raman bands.⁴⁰ Based on the XRD and the Raman spectroscopy results, the synthesized catalysts were obtained as phase pure ferrites.

After confirming the phase composition of the catalysts, the cation distribution within the ferrite catalysts was studied. Notably, the spinel structure $AB_2\text{O}_4$ has tetrahedral (A) and octahedral (B) sites, and for ferrites with normal and inverse spinel crystal structures, the compositions $[M_1]^A\{Fe_2\}^B\text{O}_4$ and $[Fe_1]^A\{M_1Fe_1\}^B\text{O}_4$ are expected, respectively. To estimate the cation distribution in the synthesized catalysts, the products were analyzed by Mössbauer spectroscopy at 6 K (Fig. 1). The collected Mössbauer spectra for all four catalysts can be fitted with three magnetically split sextets, Q_1 , Q_2 , and Q_3 , and the extracted Mössbauer parameters, namely, the centroid shift, δ , quadrupole shift, ϵ , magnetic hyperfine field, B_{hf} , magnetic hyperfine field distribution, σ , and intensity, I , for all measurements are summarized in Table S2 (ESI[†]). Analysis of the Mössbauer data revealed that the cation distribution for CoFe_2O_4 , NiFe_2O_4 , ZnFe_2O_4 , and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$, within the experimental errors, are $[\text{Co}_{0.36}\text{Fe}_{0.64}]^A\{\text{Co}_{0.64}\text{Fe}_{1.36}\}^B\text{O}_4$, $[\text{Fe}]^A\{\text{NiFe}\}^B\text{O}_4$, $[\text{Zn}_{0.36}\text{Fe}_{0.64}]^A\{\text{Zn}_{0.64}\text{Fe}_{1.36}\}^B\text{O}_4$, and $[\text{M}_{0.01}\text{Fe}_{0.99}]^A\{\text{M}_{0.79}\text{Fe}_{1.21}\}^B\text{O}_4$, respectively (Table S3, ESI[†]). These distributions confirm that while CoFe_2O_4 , ZnFe_2O_4 , and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ catalysts are partially inverse ferrites, the NiFe_2O_4 catalyst is an ideal inverse ferrite.

TEM was next employed to gather information about the size of the obtained catalysts. Fig. 2 shows the representative low-magnification TEM images of the ferrites. The NPs were found

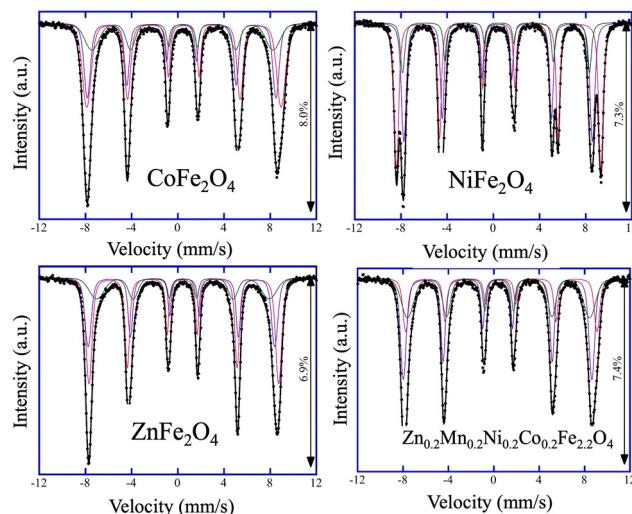


Fig. 1 Mössbauer spectra of CoFe_2O_4 , NiFe_2O_4 , ZnFe_2O_4 , and high-entropy $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ catalysts collected at 6 K. Experimental data: black circles. Calculated spectrum: black line. Q_1 , Q_2 , and Q_3 components: red, violet, and green, respectively.

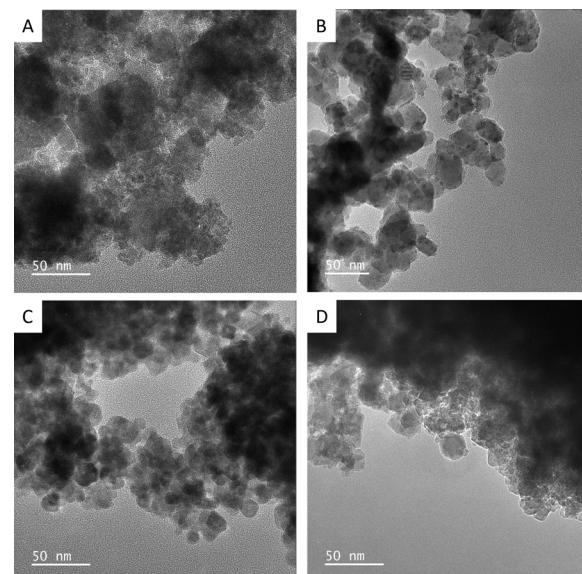


Fig. 2 Representative low-magnification TEM images of CoFe_2O_4 (A), NiFe_2O_4 (B), ZnFe_2O_4 (C), and high-entropy $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ (D) catalysts.

to be moderately agglomerated and reasonably polydisperse with average sizes of 20 ± 7 nm, 26 ± 6 nm, 18 ± 4 nm, and 21 ± 5 nm for CoFe_2O_4 , NiFe_2O_4 , ZnFe_2O_4 , and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$, respectively (Table S1 and Fig. S3, ESI[†]). Notably, the observed nanocrystalline nature of the catalysts gave rise to relatively high specific surface areas, S_{BET} , which were estimated to be 65, 42, 54, and 84 $\text{m}^2 \text{ g}^{-1}$ for CoFe_2O_4 , NiFe_2O_4 , ZnFe_2O_4 , and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$, respectively (Table S1, ESI[†]). Furthermore, SEM-EDX/STEM-EDX was used for the elemental analysis of the catalysts. The obtained data confirmed the chemical composition of the samples (Fig. S4 to S6, ESI[†]),



and indicated that the elements were uniformly distributed, even within the high-entropy $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ ferrite (Fig. S6, ESI†).

Since the catalysis takes place at the surface of the materials, X-ray photoelectron spectroscopy (XPS) was used to probe the surface of the catalysts. Fig. 3 and Fig. S7 (ESI†) show the survey and the high-resolution XPS data for the synthesized CoFe_2O_4 , NiFe_2O_4 , ZnFe_2O_4 , and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ materials. The deconvolution of the XPS data revealed that the as-synthesized ferrites display the expected elemental composition at the surface. In the ferrites containing Co, the detailed XPS peaks ($\text{Co } 2p_{3/2}$) located at 780.5 eV, 782.1 eV, and 787.5 eV can be attributed to Co^{2+} and Co^{3+} , together with the corresponding shakeup satellites,^{41,42} which is consistent with the Mössbauer data above. Here the peak located at around 782 eV is attributed to the Augur line of Fe LMM (its contribution is much more pronounced in the case of $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ due to the higher Fe:Co ratio). The same applies to the other synthesized ferrite catalysts (Fig. 3 and Fig. S7, ESI†), where the corresponding ionic species were observed, namely, for $\text{Ni}^{2+}/\text{Ni}^{3+}$,^{43,44} Zn^{2+} ,⁴⁵ and Mn^{2+} .⁴⁶ In the case of Fe 2p XPS spectra, binding energies (BEs) at 710.5 eV, 712.0 eV, and 716.0 eV can be attributed to Fe^{2+} and Fe^{3+} , together with the corresponding shakeup satellites.^{47,48} Moreover, the asymmetric nature of the Fe $2p_{3/2}$ bands for all the ferrites further suggests the existence of Fe^{3+} ions in the octahedral and tetrahedral sites of the structures.⁴⁹ In all the catalysts, in the O 1s XPS data, the bands located at around 529.2 eV, 530.4 eV,

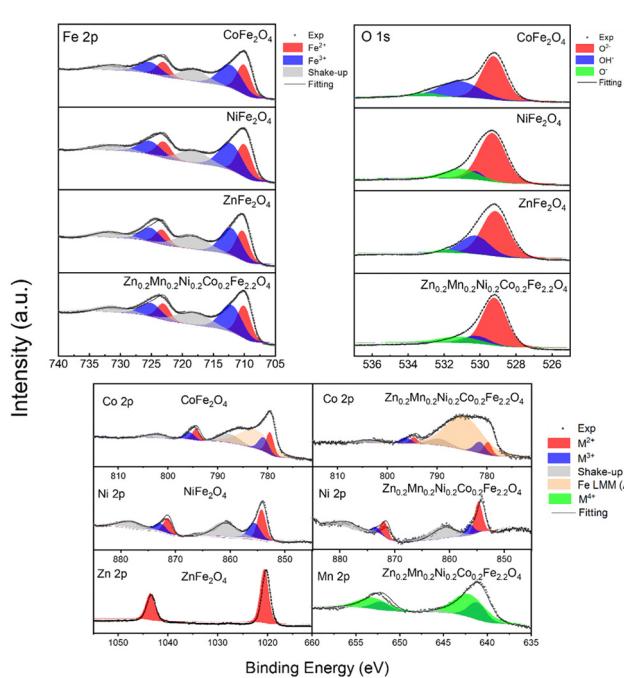


Fig. 3 XPS spectra collected from the CoFe_2O_4 , NiFe_2O_4 , ZnFe_2O_4 , and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ catalysts synthesized by the coprecipitation method followed by calcination and ball milling. High-resolution XPS data for the Fe 2p and O 1s region collected from all studied catalysts (top panel). High-resolution XPS data for the remaining metallic elements Co 2p, Ni 2p, Zn 2p, and Mn 2p for the catalysts (bottom panel).

and 531.5 eV are attributed to O^{2-} (metal oxide), OH^- (metal hydroxide) and O^- (deficiencies, or adsorbed oxygen), confirming the formation of M–O bonds.

Next, the physical properties of the synthesized materials were studied. The magnetization *versus* magnetic field $M(H)$ dependence at 300 K is shown in Fig. 4. While NiFe_2O_4 , ZnFe_2O_4 , and high-entropy $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ samples demonstrate superparamagnetic-like behavior, CoFe_2O_4 features ferromagnetic-like behavior but with rather small coercivity and remanence in the $M(H)$ curve (Table S4, ESI†). The lowest saturation magnetization $M_s = 15 \text{ emu g}^{-1}$ was expectedly observed for ZnFe_2O_4 . At the same time, CoFe_2O_4 , NiFe_2O_4 , and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ exhibited reasonably high M_s values of 39, 44, and 50 emu g^{-1} . Furthermore, the optical bandgap E_g of the synthesized catalysts was experimentally estimated by means of UV/vis diffuse reflectance spectroscopy. From the Tauc plots, optical bandgap values of 1.48 eV, 1.58 eV, 1.70 eV, and 1.51 eV were obtained for CoFe_2O_4 , NiFe_2O_4 , ZnFe_2O_4 , and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$, respectively (Table S1 and Fig. S8, ESI†). Despite the limitation of this method for band gap determination,⁵⁰ the observed results are in line with reported data, wherein the optical band gap of the ferrites is strongly influenced by the preparation method, heat treatment procedures, mean particle size, and the degree of structural disorder in the lattice.^{51–55}

The magnetic and electronic properties of CoFe_2O_4 , NiFe_2O_4 , ZnFe_2O_4 , and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ were also investigated by the DFT+U+J method (see the Computational study section on spinel ferrites in the ESI† for further information). The cation distributions provided by Mössbauer spectroscopy were used as plausible input to build the computational bulk models for all the ferrites under study. Two possible models were constructed in the case of $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ (Fig. 5a) and NiFe_2O_4 (Fig. 5b). The $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ bulk models were built based on both the cation distribution provided by Mössbauer

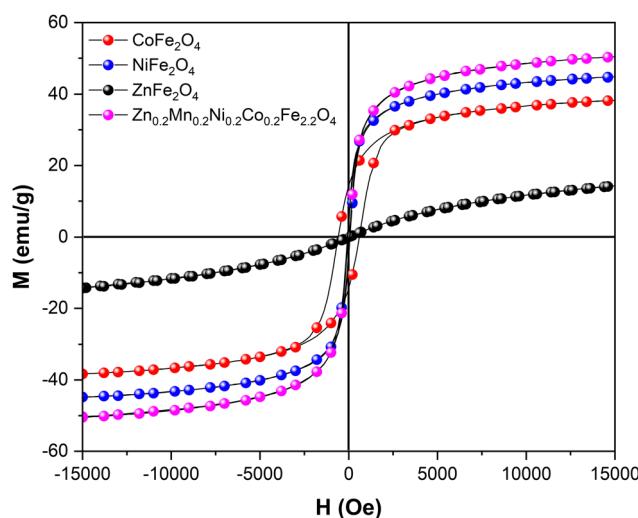


Fig. 4 Room temperature $M(H)$ dependence data for the synthesized catalysts.



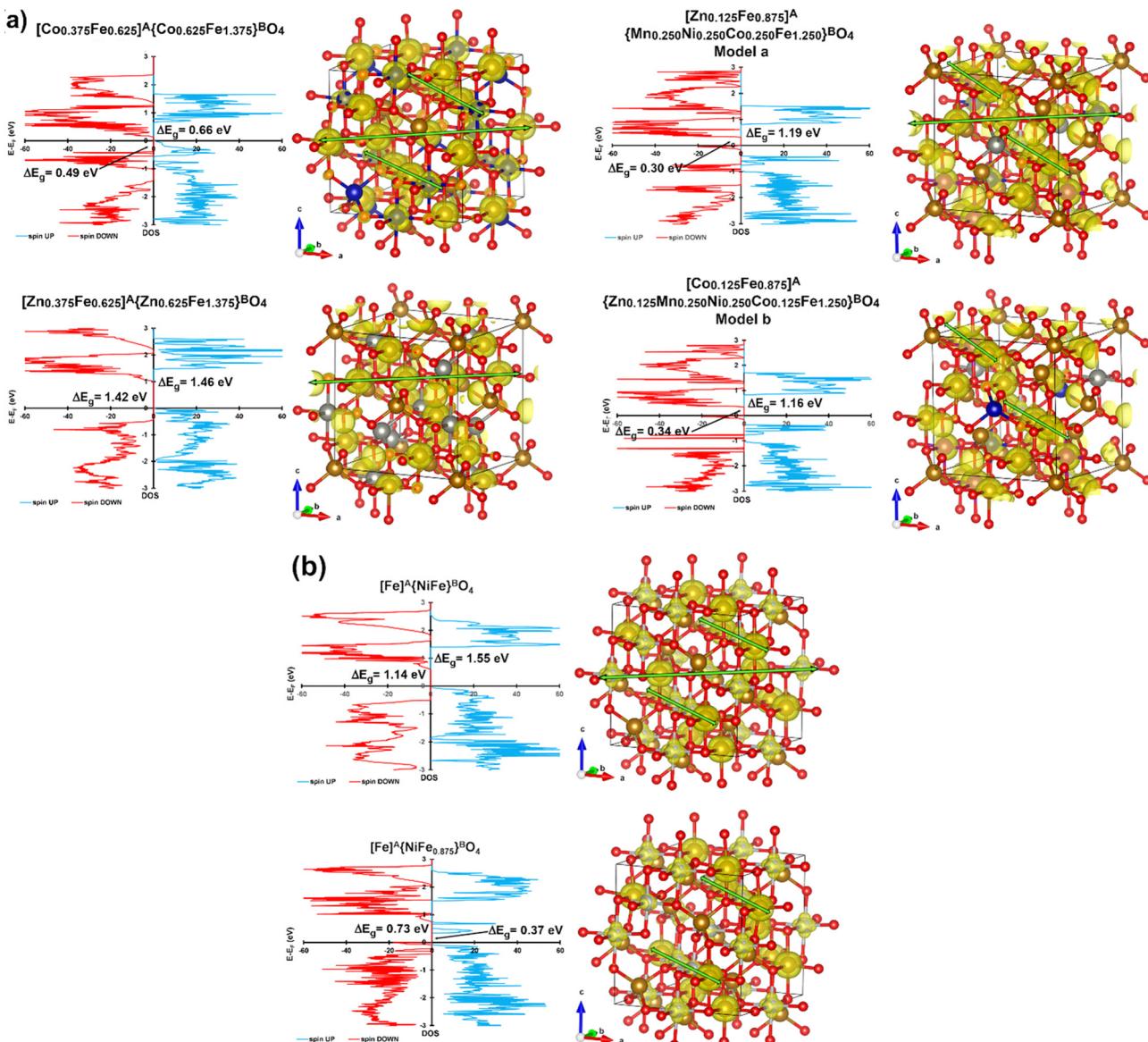


Fig. 5 (a) Density of states (DOS, left) and $\alpha - \beta > 0$ spin density (spin up, $0.04 a_0^{-3}$, right) for the most stable ground states of partially inverse spinel model a and model b of $Zn_{0.2}Mn_{0.2}Ni_{0.2}Co_{0.2}Fe_{2.2}O_4$, and (b) DOS (left) and $\alpha - \beta > 0$ spin density (spin up, $0.04 a_0^{-3}$, right) for the most stable ground states of perfect $NiFe_2O_4$ and defective $NiFe_{2-x}O_4$. Possible spin channels are marked with green arrows.

spectroscopy ($T = 6$ K) and the high-resolution XPS data shown in Fig. 3, in which both Zn^{2+} and Co^{2+} can occupy tetrahedral positions. Thus, $[Zn_{0.125}Fe_{0.875}]^A\{M_{0.750}Fe_{1.250}\}^B O_4$ ($M = Co, Ni$, and Mn) and $[Co_{0.125}Fe_{0.875}]^A\{M_{0.750}Fe_{1.250}\}^B O_4$ ($M = Co, Ni, Mn$, and Zn) are model a and model b, respectively, for $Zn_{0.2}Mn_{0.2}Ni_{0.2}Co_{0.2}Fe_{2.2}O_4$. These models align with the available experimental data on $Zn_{0.2}Mn_{0.2}Ni_{0.2}Co_{0.2}Fe_{2.2}O_4$, which report challenges in including Zn^{2+} and Ni^{2+} species into the spinel structure.⁵⁶ Regarding $NiFe_2O_4$, Mössbauer spectroscopy showed an intensity of 46% at the B-sites (see the first sub-spectrum Q_1 in Table S2, ESI[†]). This value is slightly smaller than the expected 50%. A possible explanation for the observed small reduction in Q_1 could be the presence of some Fe

vacancies ($\approx 12\%$) at the octahedral sites of $NiFe_2O_4$. This has also been suggested experimentally in the available literature data, since the cation distribution and the cation occupancy for nickel ferrite depends on several factors, such as the preparation method, calcination temperature, and particle size.⁵⁷ Thus, two computational bulk models for the inverse $NiFe_2O_4$ ferrite were built, one with no Fe vacancies at the octahedral sites and one with 15% of Fe vacancies at the B-sites (Fig. 5b). The calculated lattice parameters at 0 K for all the investigated ferrites are in good agreement with the experimental ones presented in Table S1 (ESI[†]). Regarding the magnetic properties, all possible electronic ground states (*i.e.*, ferromagnetic, antiferromagnetic, ferrimagnetic, and non-magnetic) were



computationally explored for each ferrite (Tables S8–S13 and Fig. S9–S20, ESI†). Calculation data show that all the bulk structures possess ferrimagnetic ordering as the most stable electronic ground state at 0 K. Indeed, all the samples have non-zero magnetization at 300 K (Fig. 3 and Table S4, ESI†), suggesting that the blocking temperature falls between 1–300 K in the case of superparamagnetic samples. Moreover, the calculated minimum band gaps (Fig. 5a and b) show that all the ferrites under study are semiconductors, in good agreement with the experimental optical band gaps estimated by UV/vis diffuse reflectance spectroscopy, which are known to depend on the particle size.^{58,59}

Regarding conductivity properties (Table S14, ESI†), defective $\text{NiFe}_{1.875}\text{O}_4$ is the best among the investigated bulk models, exhibiting the smallest minimum band gaps (0.37 and 0.73 eV as spin \uparrow and spin \downarrow band gaps, respectively). Partially inverse $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ (model a and model b) and CoFe_2O_4 are the two ferrites displaying the smallest minimum band gaps within the stoichiometric ferrites, demonstrating 1.19–1.16 eV (spin \uparrow)/0.30–0.34 eV (spin \downarrow) and 0.66 eV (spin \uparrow)/0.49 eV (spin \downarrow), respectively.

With phase-pure, nanocrystalline, and chemically uniform ferrite materials in hand, we investigated the anodic OER half-reaction of electrochemical water splitting in alkaline electrolyte at room temperature through half-cell laboratory measurements. First, we studied the OER properties of the synthesized materials in the kinetic regime with low catalyst mass loading (0.3 mg cm^{-2}) while using the RDE testing mode so that the effects of mass transport limitations are minimized.⁶⁰ As shown in the inset of Fig. 6a, the basic CV characteristics of all the ferrite catalysts were quite similar: (i) capacitor-like behavior over a wide potential range, which may be ascribed to the capacitance current of both the ferrite catalyst and the conductive carbon additive; (ii) a pair of redox peaks located at $\approx 1.3 \text{ V}_{\text{RHE}}$, which may be ascribed to the redox behavior of metal hydroxide/oxyhydroxide (A^{2+} -to- A^{3+} , *e.g.*, $\text{Ni}(\text{OH})_2$ to NiOOH in the case of NiFe_2O_4) at potentials prior to the OER onset potential.

In Fig. 6a, the OER current may overlap with the oxidation current of the ferrites, as shown by the redox peaks at $\approx 1.3 \text{ V}_{\text{RHE}}$ (inset of Fig. 6a). According to the literature,^{31,61,62} the onset potential for the OER over ferrite catalysts (*e.g.*, NiFe_2O_4 and CoFe_2O_4) is expected to be $\approx 1.5 \text{ V}_{\text{RHE}}$. This is also the case with most of our ferrites, with CoFe_2O_4 and ZnFe_2O_4 featuring a slightly earlier onset. Nevertheless, for the precise determination of the OER onset potential (*i.e.*, to track the onset potential for the production of O_2 rather than the anodic current), the application of *in situ* techniques, such as differential electrochemical mass spectrometry^{63,64} is necessary.

The electrochemical activity towards OER was evaluated by comparing the current densities (normalized by electrochemically active surface area) at $1.6 \text{ V}_{\text{RHE}}$ (Fig. 6b). Since the measured overall current is $\approx 0.1 \text{ mA}$, the Ohmic drop corrections (*viz.* we applied 80%) will not make a large difference.⁶⁵ All the synthesized ferrite catalysts have similar OER activities in alkaline media except for CoFe_2O_4 , the activity of which was

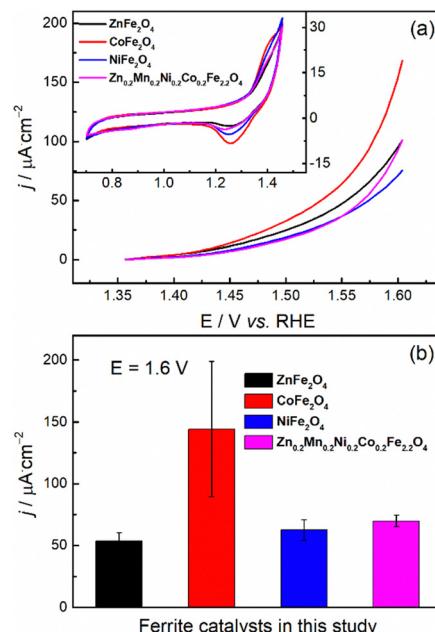


Fig. 6 (a) Representative positive going linear sweep voltammograms at OER potentials and (b) the current densities (normalized by ECSA) at $1.6 \text{ V}_{\text{RHE}}$ towards the OER for the synthesized catalysts. The corresponding basic cyclic voltammograms are given as the inset in (a).

measured to be about twice that of the others, albeit with significant uncertainty. The observed higher OER activity of CoFe_2O_4 in the kinetic regime could originate from the suitable binding strength of the reactive intermediates during the OER.⁶⁶

Second, to be closer to realistic AEMEL conditions, we examined the OER properties of the synthesized ferrites with a higher catalyst mass loading of 1 mg cm^{-2} on a glassy carbon plate to drive higher current densities.⁶⁷ For this purpose, we employed commercially available Ni foam. From the recorded LSV voltammograms (Fig. 7a), the OER activity decreases in the order of $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4 > \text{NiFe}_2\text{O}_4 > \text{CoFe}_2\text{O}_4 > \text{ZnFe}_2\text{O}_4$, with overpotential values of 432, 455, 484, and 500 mV, respectively, at a current density of 10 mA cm^{-2} . Similarly, the Tafel slopes increase in the same trend (Fig. 7b). Also, $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ shows a very low charge transfer resistance (0.66 Ohm) in the Nyquist plot (Fig. 7c). Fig. 7d shows the capacitive current at different scan rates recorded in the non-faradaic potential window, where $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ again shows the highest positive slope, implying the highest roughness factor. From the slope, the calculated roughness values are 61, 47, 37, and 21 for $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$, NiFe_2O_4 , CoFe_2O_4 , and ZnFe_2O_4 , respectively.

With the promising OER data in hand and to gain insight into structure–property relationships of the synthesized ferrite catalysts, we next implemented the materials in the AEMEL cell to study their system-level performance. Fig. 8 shows the polarization curves of AEMEL cells with the four different ferrite anode catalysts, operated at 60°C (Fig. 8a) and 80°C (Fig. 8b).

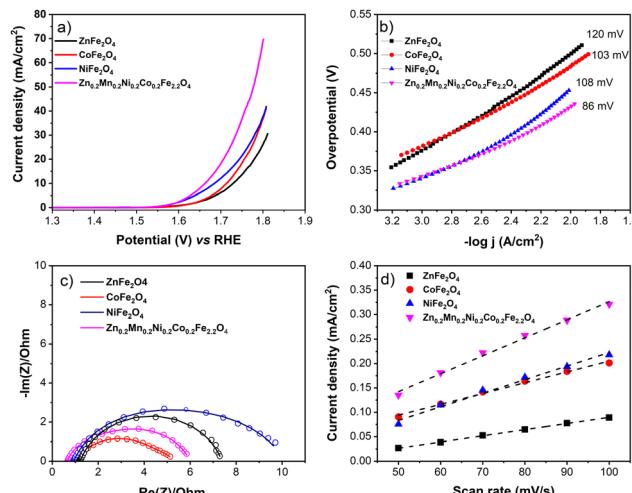


Fig. 7 (a) Representative LSV voltammograms recorded at the 5 mV s^{-1} scan rate and (b) Tafel plots extracted from the respective LSV curves; (c) Nyquist plot; (d) capacitive current density as a function of scan rate for the synthesized ZnFe₂O₄, CoFe₂O₄, NiFe₂O₄, and Zn_{0.2}Mn_{0.2}Ni_{0.2}Co_{0.2}Fe_{2.2}O₄ catalysts.

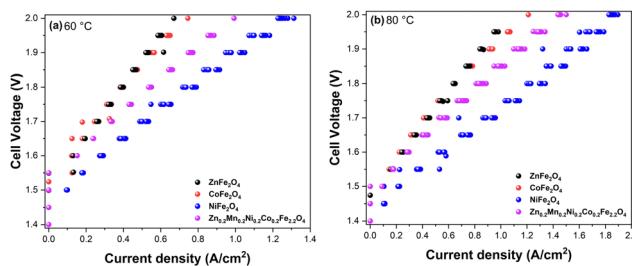


Fig. 8 Polarization curves of AEMEL cells with the synthesized ZnFe₂O₄, CoFe₂O₄, NiFe₂O₄, and Zn_{0.2}Mn_{0.2}Ni_{0.2}Co_{0.2}Fe_{2.2}O₄ anode catalysts recorded at 60 °C (a) and 80 °C (b) under atmospheric pressure in 1 M KOH electrolyte.

At a cell potential of 2.0 V, the AEMEL cells with ZnFe₂O₄ and CoFe₂O₄ anode catalysts were found to deliver the lowest current densities of ≈ 0.7 and $\approx 1.2 \text{ A cm}^{-2}$ at 60 and 80 °C, respectively. The AEMEL cells with high-entropy Zn_{0.2}Mn_{0.2}Ni_{0.2}Co_{0.2}Fe_{2.2}O₄ as the anode showed a significantly higher performance, delivering current densities of $\approx 1.0 \text{ A cm}^{-2}$ (60 °C) and $\approx 1.5 \text{ A cm}^{-2}$ (80 °C) at a cell potential of 2.0 V. Finally, at a cell potential of 2.0 V, AEMEL cells with the NiFe₂O₄ anode catalyst were found to deliver ≈ 1.3 and $\approx 1.9 \text{ A cm}^{-2}$ at 60 and 80 °C, respectively.

After determining the viability of the synthesized ferrites as anode catalysts on a single-cell AEMEL level, we then investigated the short-term durability of the AEMEL cells with different ferrite anodes. Fig. 9 shows the cell-voltage changes under a constant current density as a function of time. Notably, the recorded cell voltages demonstrate merely small fluctuations over 20 h of chronoamperometric measurements, thus evidencing good short-term durability of the AEMEL cells with ferrite anode catalysts. At a constant current density of 0.5 A cm^{-2} , the AEMEL cell performance decreased in the order of NiFe₂O₄ > Zn_{0.2}Mn_{0.2}Ni_{0.2}Co_{0.2}Fe_{2.2}O₄ > CoFe₂O₄ > ZnFe₂O₄, reinforcing

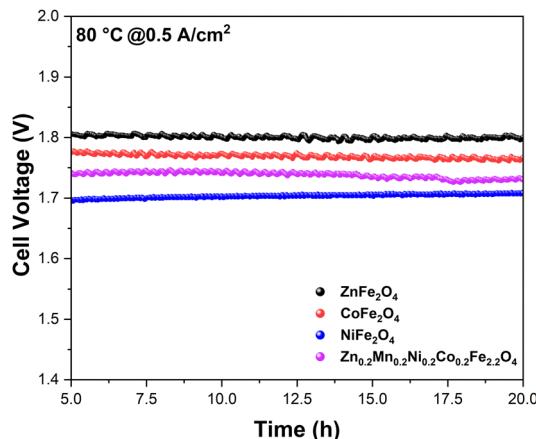


Fig. 9 Cell–voltage changes as a function of testing time for the AEMEL cells with the synthesized ZnFe₂O₄, CoFe₂O₄, NiFe₂O₄, and Zn_{0.2}Mn_{0.2}Ni_{0.2}Co_{0.2}Fe_{2.2}O₄ anode catalysts, recorded at 80 °C under atmospheric pressure in 1 M KOH electrolyte.

the I – V measurement results presented in Fig. 8. Overall, from all the prepared ferrites with different chemical compositions and material properties, NiFe₂O₄ showed the best performance at the single-cell AEMEL level, affording the highest current density of $\approx 1.9 \text{ A cm}^{-2}$ with a cell voltage of 2.0 V at 80 °C while demonstrating high stability.

Discussion

The spinel ferrites, such as NiFe₂O₄ or CoFe₂O₄, are among the most explored earth-abundant OER catalysts because of their high OER activity.^{31,43,68–71} For example, Ni_xCu_{1-x}Fe₂O₄ ferrites anchored onto S-doped graphitic carbon nitride have been reported to exhibit low overpotential $\eta_{10} = 250$ mV at current density $j = 10 \text{ mA cm}^{-2}$ and good stability.⁷² In another study, doped CoFe₂O₄-based hollow nanospheres displayed high OER activity due to heteroatom substitution and vacancy engineering through Cr³⁺ doping and S²⁻ exchange.⁷³ Doping with Cr³⁺ at the octahedral Fe sites induces a Co vacancy that activates the adjacent Fe³⁺ present at the tetrahedral sites. The S²⁻ exchange resulted in a structural distortion of tetrahedral Fe, leading to enhanced adsorption of intermediate *OOH species in the OER. Another work described the preparation of mesoporous first-row transition metal ferrites based on Mn²⁺, Fe²⁺, Co²⁺, Ni²⁺, Cu²⁺, and Zn²⁺,⁷⁴ and found NiFe₂O₄ to show the highest OER activity with low $\eta_{10} = 278$ mV. The best OER activity was also found for NiFe₂O₄ in a study of tubular ferrite microstructures (AFe₂O₄, A = Fe²⁺, Co²⁺, and Ni²⁺),⁷⁵ where the high activity was attributed to the large specific surface area, high active surface area, and low charge-transfer resistance. In this context, MnFe₂O₄ has also been reported as an alternative OER catalyst by fine-tuning the Mn occupancy at the octahedral sites along with the conversion of Mn²⁺–Mn³⁺ through heat treatment.⁷⁶ The mixed valence of Mn resulted in improved alkaline OER activity, wherein Mn³⁺ was suggested to be more catalytically active than Mn²⁺.



In this work, we show that high-quality ferrite nanoparticle catalysts can be prepared at large-scale (5 g) using an automated synthesis system. The magnetic property measurements confirmed that the synthesized materials largely exhibit a superparamagnetic state at room temperature as a consequence of the small size of the particles, allowing the spontaneous flipping of their magnetization at room temperature. We further found that the OER properties of the synthesized ferrites are quite similar in the kinetic regime at low current densities of $\gg 0.1$ mA (Fig. 6). At the same time, the alkaline OER measurements with higher catalyst mass loading over glassy carbon plate, *i.e.*, at higher current densities $\gg 0.1$ mA, clearly show that the catalysts demonstrate different performance compared to that observed in the kinetic regime. Specifically, NiFe_2O_4 and CoFe_2O_4 ferrites feature OER activities that align with the literature data (Table S5, ESI[†]), with overpotentials of 455 and 484 mV at 10 mA cm⁻², respectively. In addition, high-entropy $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ ferrite outperforms the two-metal ferrites showing an overpotential of $\eta_{10} = 432$ mV at $j = 10$ mA cm⁻². Interestingly, the number of reports regarding the OER activity of high-entropy spinels is still limited as compared to one- or two-metal spinel oxides (Table S5, ESI[†]), and the results in Fig. 7 highlight the potential of such materials as highly active OER catalysts.^{77,78}

Previously, many descriptors have been identified for the OER,⁷⁹ such as e_g electrons, M–O–M bond angle, and M–O bond length. The magnetic moment is also considered as one of the descriptors.⁸⁰ In the case of spinel ferrites, therefore, the saturation magnetization can be identified as an important macroscopic experimental descriptor, as anticipated from theory,^{23,81,82} as the increase in the saturation magnetization results in an increase in the OER activity in the case of ferrites AFe_2O_4 (A = Mg, Ni, Mn, and Co).⁸³ This trend is clearly observed in the current study (Fig. 10), with $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ exhibiting the highest saturation magnetization and showing the highest OER activity at moderate current densities (Fig. 6). On the other hand, the ZnFe_2O_4

catalyst with the lowest saturation magnetization showed the lowest OER activity among the spinel ferrites (Fig. 10). Also, the surface state of the high-entropy ferrite is unique compared to the ferrites with fewer elements. Namely, the presence of a higher number of elements leads to highly dispersed active sites, and may provide various active sites for the OER.⁸⁴ In the present case, $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ has very low charge transfer resistance, a low overpotential, a low Tafel slope, a low bandgap, high saturation magnetization, and a high roughness factor, leading to the observed highest OER activity in the moderate current density range (Fig. 7). Interestingly, in full-cell AEMEL testing at high current densities, NiFe_2O_4 outperforms $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ (Fig. 8 and 9), and the observed AEMEL performance for these two anodes compares favorably with the literature (Table S6, ESI[†]). Two explanations can be proposed for the superior activity of NiFe_2O_4 and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ at high current densities. First, the computational calculations revealed that nickel and high-entropy ferrites show the lowest band gaps, *i.e.*, exhibit high conductivity (Fig. 5a and b and Table S13, ESI[†]), due to the presence of Fe vacancies that help to reduce the band gap and enhance the electron transfer. Second, after careful examination of the spin polarization of the oxygen atoms along the spin channels (Fig. 5a and b), we can suggest that nickel ferrite should advantageously possess more catalytically active centers along the spin channels (*i.e.*, more metal sites with high spin-polarized oxygen atoms) than $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ ferrite, which showed spin polarization on the oxygen atoms only at very few metal sites.

Conclusions

The spinel ferrites, CoFe_2O_4 , NiFe_2O_4 , and ZnFe_2O_4 , and high entropy $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ were successfully synthesized on a large scale by the coprecipitation method. XRD, Mössbauer spectroscopy, and XPS confirmed the mixed spinel compositions of the samples. The magnetic measurements evidenced that the synthesized nanocrystalline materials predominantly exhibit a superparamagnetic ground state at room temperature. In the kinetic regime at very low current densities, the ferrites demonstrate similar alkaline OER properties, while high entropy $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ ferrite was found to be the best-performing OER catalyst in the moderate current density range. An increase in the OER activity was observed with increasing saturation magnetization, indicating that saturation magnetization can be considered a macroscopic experimental descriptor for magnetic spinels. Finally, the AEMEL measurements at high current densities revealed NiFe_2O_4 and $\text{Zn}_{0.2}\text{Mn}_{0.2}\text{Ni}_{0.2}\text{Co}_{0.2}\text{Fe}_{2.2}\text{O}_4$ as the most promising anode catalysts among the studied spinel ferrites, most likely due to their favorable electronic and magnetic structures, as corroborated theoretically. Future efforts should focus on (i) the electrode level optimization of AEMEL, since this hydrogen generation technology is in the developing stage at present; and (ii) investigation of the OER properties of ferrimagnetic/ferromagnetic counterparts of the ferrites presented in the current study.

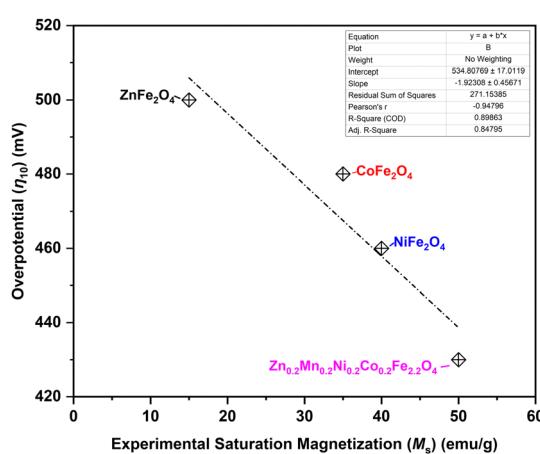


Fig. 10 Alkaline OER overpotential η_{10} needed to drive current density $j = 10$ mA cm⁻² as a function of saturation magnetization of the synthesized ferrite catalysts.



Author contributions

Tiago Fernandes: conceptualization, data curation, formal analysis, investigation, methodology, validation, visualization, and writing – original draft. Ramsundar Mohan: data curation, investigation, methodology, and writing – original draft. Laura Donk: investigation and methodology. Wei Chen: investigation, methodology, and writing – original draft. Chiara Biz: investigation, methodology, and writing – original draft. Mauro Fianchini: investigation, methodology, and writing – original draft. Anna Kitayev: investigation and writing – review & editing. Saeed Kamali: investigation, data curation, and writing – review & editing. Siavash Mohammad Alizadeh: data curation. Aviv Ashdot: investigation and writing – review & editing. Miles Page: conceptualization, funding acquisition, methodology, and writing – review & editing. Laura M. Salonen: writing – review & editing. Sebastian Kopp: writing – review & editing. Ervin Tal Gutelmacher: funding acquisition, resources, supervision, validation, and writing – review & editing. José Gracia: funding acquisition, resources, supervision, validation, and writing – review & editing. Marta Costa Figueiredo: funding acquisition, resources, supervision, validation, and writing – review & editing. Yury V. Kolen'ko: funding acquisition, resources, supervision, validation, and writing – review & editing.

Data availability

The data supporting this article have been included as part of the ESI.† Data related to materials synthesis and characterization are openly available in a public repository that issues datasets with DOIs.

Conflicts of interest

There are no conflicts to declare.

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