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Mechanistic insights into boron-catalysed direct amidation reactions†

Sergey Arkhipenko,^a Marco T. Sabatini,^b Andrei S. Batsanov,^c Valerija Karaluka,^b Tom D. Sheppard,^b Henry S. Rzepa^d and Andrew Whiting^{id}*^a

The generally accepted monoacyloxyboron mechanism of boron-catalysed direct amidation is brought into question in this study, and new alternatives are proposed. We have carried out a detailed investigation of boron-catalysed amidation reactions, through study of the interaction between amines/carboxylic acids and borinic acids, boronic acids and boric acid, and have isolated and characterised by NMR/X-ray crystallography many of the likely intermediates present in catalytic amidation reactions. Rapid reaction between amines and boron compounds was observed in all cases, and it is proposed that such boron–nitrogen interactions are highly likely to take place in catalytic amidation reactions. These studies also clearly show that borinic acids are not competent catalysts for amidation, as they either form unreactive amino-carboxylate complexes, or undergo protodeboronation to give boronic acids. It therefore seems that at least three free coordination sites on the boron atom are necessary for amidation catalysis to occur. However, these observations are not consistent with the currently accepted ‘mechanism’ for boron-mediated amidation reactions involving nucleophilic attack of an amine onto a monomeric acyloxyboron intermediate, and as a result of our observations and theoretical modelling, alternative proposed mechanisms are presented for boron-mediated amidation reactions. These are likely to proceed *via* the formation of a dimeric B–X–B motif (X = O, NR), which is uniquely able to provide activation of the carboxylic acid, whilst orchestrating the delivery of the amine nucleophile to the carbonyl group. Quantum mechanical calculations of catalytic cycles at the B3LYP+D3/Def2-TZVPP level (solvent = CH₂Cl₂) support the proposal of several closely related potential pathways for amidation, all of which are likely to be lower in energy than the currently accepted mechanism.

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Introduction

Boron compounds have a rich, fascinating and complex structural chemistry, especially related to clusters, cages, complexes and anions associated with B–O¹ and B–N² containing systems. Indeed, such systems are well known to “have a distinct tendency to disproportionate, apparently more so than those of any other nonmetallic elements”.^{1a} This observation means that identifying and controlling the role played by boron in reactions

and catalytic processes is particularly challenging and has proved so on many occasions.

Elegant single and cooperative solutions, exemplified in Scheme 1, have enabled the precise reactivity control of boron in different reactions. For example, in the catalytic asymmetric CBS reduction system³ both nitrogen and oxygen ligands

^aCentre for Sustainable Chemical Processes, Department of Chemistry, Durham University, Science Site, South Road, Durham, DH1 3LE, UK. E-mail: andy.whiting@durham.ac.uk

^bDepartment of Chemistry, University College London, 20 Gordon Street, London, WC1H 0AJ, UK

^cDepartment of Chemistry, Durham University, Science Site, South Road, Durham, DH1 3LE, UK

^dDepartment of Chemistry, Imperial College London, South Kensington Campus, London, SW7 2AZ, UK

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Scheme 1 General structural role of boron in different catalytic processes.



control boron coordination to tune an external boron to behave cooperatively to control ketone Lewis-acid binding, activation and hence, reduction. Similarly, precise boron centre control can be employed to achieve highly efficient hydroxynaphthaquinone chiral Lewis-acid binding, activation and subsequent asymmetric Diels–Alder cycloaddition;⁴ the boron control and prevention of diborate-disproportionation deriving from use of hindered BINOL ligands, assembled by careful manipulation of unstable acyloxyboron species. Indeed, the model for these boron complexes was derived from ‘triaceoxyborate’ complexes of chiral juglone derivatives. However, triaceoxyborate is not stable at ambient temperatures and it rearranges to give tetra-acetyl diborate, as confirmed by X-ray crystallography.⁵ These seminal works led us to conclude that reaction mechanisms involving boron are non-trivial, because the structural chemistry of boron is itself non-trivial.

It is generally the case that in the absence of a catalyst, amines and carboxylic acids do not usually react to form amides at ambient temperatures, and forcing conditions are often required to promote this condensation reaction to access amides.⁶ Although a direct thermally accessible reaction is possible, direct amidation is better facilitated at lower temperatures by using catalysts,⁷ with boron-based systems being particularly effective.⁸ Useful catalysts for direct amidation include boric acid,⁹ borate esters¹⁰ and boronic acids.^{11–13} Boronic acids have also been claimed to promote direct amidation reactions under relatively mild conditions.¹⁴ Clearly, an understanding of the mechanism of these boron-mediated amidations is crucial for designing catalysts with improved activity, but more importantly, a fundamental understanding of the role of the boron in such systems is required. To this end, we have been interested in the structural and mechanistic chemistry associated with boron compounds employed in direct amidation for several years.^{12a} The diversity of the catalytic systems reported to date, from simple boric acid⁹ or boronic acids^{11–13} to carefully designed cooperative catalysts,¹⁵ including poly-boron-based systems,¹⁶ poses a challenge in terms of identifying key structural motifs which are necessary for effective catalysis. Even considering boronic acid catalysts alone, bifunctional systems containing nitrogen and boron,¹² or iodine/bromine and boron¹³ and electron deficient aryl boronic acids¹⁴ have all been demonstrated to be highly effective. Currently, a comprehensive understanding of how the catalyst structure affects the activity is clearly lacking.

Previously, an acyloxyboron-based intermediate was proposed as the reactive catalytic species,^{11a} with computational studies suggesting that it was energetically plausible (Scheme 2A).¹⁷ As shown in Scheme 2, this mechanism involved the formation of the acyloxyboron derivative which then had to undergo nucleophilic attack by an amine to give the amide, after loss of water. However, despite the fact that DFT calculations¹⁷ suggest that acyloxyboron species are feasible intermediates in direct amidation catalysis, actual definitive physical evidence of the mechanistic role of boron in such reactions is sparse and could be misleading due to the potential for numerous possible boron species to be present in such reactions.^{12a} Indeed, it should be noted that acyloxyboron species are particularly prone to dimerization/oligomerization, as mentioned above with regard to B(OAc)₃ (*vide supra* and Scheme 1).



Scheme 2 (A) Currently proposed mechanism *via* a monomeric acyloxyboron intermediate. (B) Summary of mechanisms proposed in this paper involving dimeric B–X–B motifs.

In this work, we report recent studies in which we have probed the types of species potentially involved in boron-catalysed amidation reactions using a combination of experimental, structural and theoretical approaches. As a result, we postulate that acyloxyboron species like **I** (Scheme 2A) are unlikely to be sufficiently reactive as acylating agents, and we propose new mechanistic pathways for boron-mediated amidation which proceed through B–X–B systems, as shown in Scheme 2B. We also demonstrate their feasibility experimentally, supported by computational methods. Indeed, several routes for catalytic amidation are suggested, all of which were calculated to have lower energy barriers than the previously postulated single boron atom mechanisms.¹⁷ Importantly, they involve intermediates with two boron centres, and take account of the possible role of both boron–nitrogen and boron–oxygen interactions, which can be readily observed between amines and boronic or borinic acids, and between these boron compounds and carboxylic acids. These observations point to a generic structural motif which can potentially explain the success of cooperative,¹⁵ bifunctional^{12a} and polyboron amidation catalysts,¹⁶ and they may even explain catalysis in other, related reactions.¹⁸

Results and discussion

Observations on the acyloxy mechanism

In recent computational studies,¹⁷ both tetrahedral^{17a} and trigonal^{17b} acyloxyboron species were considered as possible active acylating agents, with the latter leading to a lower energy barrier





Fig. 1 X-ray structures of Lewis adducts **4a** (H-bonded dimer) and **4b**.



Fig. 2 X-ray structures of amino-carboxylate borinic complexes **5a** (showing a hydrogen-bonded dimer of two independent molecules) and **5b** (disorder of phenyl ring is omitted for clarity). For **5c** and **5d** see ESI.†

In contrast, when bis-3,4,5-trifluorophenylborinic acid **3a** was reacted with carboxylic acids alone, immediate formation of a boron “ate”-complex was observed according to ^{11}B NMR (Fig. 3), as shown in Scheme 5, and a subsequent slow protodeboronation reaction occurred which generated a boronic species (see Fig. 3). This was followed, in this case, by slow crystallisation of the corresponding boroxine **7a** (Fig. 4). Hence, we propose that the “ate”-complex observed by ^{11}B NMR was **6** (Scheme 5 and Fig. 3) based on systems of this type which have been isolated.²¹ In addition, the mass spectrum of the mixture showed an m/z 725.1340 (calcd m/z for **6a** $[\text{M} - \text{H}]$ 725.1343). Complex **6a** underwent protodeboronation–dehydration to give the notably uncomplexed boroxine **7a**, and its corresponding boronic acid **8a**. The fact that such reactions occur readily at room temperature shows that formation of B–O–B species is extremely facile. Both borinic and boronic acids can thus readily form bridged B–O–B systems such as **6** and **7** (and **11**, *vide infra*) which demonstrates that such motifs play an important role, both in boron chemistry in general, and likely also in direct amidation catalysis. It also exemplifies the ease with which a borinic acid undergoes protodeboronation in the presence of a carboxylic acid.



Fig. 3 ^{11}B NMRs of a mixture of bis-3,4,5-trifluorophenylborinic acid **3a** with benzoic acid after 5 min and after 6 days upon mixing (Ar = 3,4,5-trifluorophenyl).



Scheme 5 Reaction of 3,4,5-trifluorophenylborinic acid **3a** with 4-phenylbutyric or benzoic acid, resulting in boroxine formation.



Fig. 4 X-ray structure of 3,4,5-trifluorophenylboroxine **7a** in orthorhombic β -form at 120 K.

Having observed that borinic acids do not mediate amide bond formation, we were surprised by a recent report describing borinic acid-catalysed amide formation.¹⁴ However, when we attempted amidations under exactly the same conditions to those recently reported,¹⁴ amides could indeed be obtained in good yield; an explanation for which was required.

After further investigations, we discovered that the key to catalytic reactivity was the 15 minute “pre-stir” period, as



explained in Scheme 6. When the claimed catalyst, bis-2-chlorophenylboronic acid **3c**, was exposed to a carboxylic acid in the presence of activated 5 Å molecular sieves (MS), catalytic activity was observed. Without this “prestir” period, no amide formation occurred. The significant finding was that the ^{11}B and ^1H NMR spectra of the crude reaction mixtures were different depending on whether the “prestir” period was applied or not (Scheme 6A and B). This data indicated that if boronic acid **3c**, phenylacetic acid and benzylamine were mixed simultaneously without the “prestir” period, aminocarboxylate complex **5c** was formed and remained stable over at least one week in solution (*vide supra*, Scheme 4 and 6A). When analyzing the product mixtures obtained after the “prestir” period, significant amounts of boronic species were observed (Scheme 6B). We propose, therefore, that boronic acid **3c** acts as a pre-catalyst, and its protodeboronation occurs rapidly during the “prestir” period to give 2-chlorophenylboronic acid **8c**, which is well known to be an effective catalyst for direct amidation, as reported by Hall *et al.*^{13a} This finding was further verified by a separate experiment in which bis-2-chlorophenylboronic acid **3c** and phenylacetic acid were stirred for 15 minutes with 5 Å MS, leading to a 1 : 6 ratio of borinic to boronic species (Scheme 6C). Interestingly, the process of protodeboronation occurred much more slowly when activated 4 Å MS were used compared to 5 Å MS, so the particular sieves used play an important role in assisting the deboronation reaction (it is worth noting that generally either 3 or 4 Å MS are used for achieving dehydration in most amidation reactions as an alternative to azeotropic water removal. See ref. 13b for a relevant discussion).

To further support the fact that boronic acid **8c** is actually formed from boronic acid **3c** *in situ* under the “prestir” conditions and that **8c** is the active catalyst, we compared **3c** and commercial 2-chlorophenylboronic acid **8c** side-by-side under the same conditions (Scheme 6B). Similar amide conversions were observed in the two reactions with **3c** and **8c** (91% and 90% respectively), which supports our proposal that the boronic acid is the actual catalyst. It can, therefore, be concluded that boronic acids do not catalyse amidations directly, and can only do so through protodeboronation to the boronic acid.

The fact that aminocarboxylate complexes **5** could not be converted into amides seems to be rather inconsistent with the commonly accepted mechanism for boron-mediated amidation (Scheme 2A). Since boronic acids are stronger Lewis acids than boronic acids, it might be expected that acyloxyboronic species **5** should actually be stronger acylating agents than the acyloxyboronic analogues, and these complexes also have the ability to form internal hydrogen bonds between the protonated amine and the carbonyl. Since both boronic acids and boric acid catalyse direct amidation,⁸ but boronic acids do not, it implies that at least three free coordination sites are needed on the boron atom in order for amidation catalysis to be possible. This conclusion led us to consider, therefore, whether B–X–B dimer formation might be essential for the observed catalytic activity, and that a bridging X-group (*i.e.* O or N) could be essential to orchestrate the catalytic process. We therefore looked more closely at the speciation that might be occurring in boronic acid-mediated direct amidation catalysis, using both experimental and theoretical methods.



Scheme 6 Examining bis-2-chlorophenylboronic acid **3c** as a catalyst for direct amide formation and ^{11}B NMR spectra of product mixtures obtained in 3 examined cases.



Boronic acids in direct amide formation

Due to the complexity of carboxylic acid/amine/boronic acid systems, the reactivity of boronic acids with amines and carboxylic acids was initially explored separately. When treated with an amine, boronic acids appear to form complexes of general structures **9** and **10**.²⁰ 3,4,5-Trifluorophenyl boroxine-benzylamine complex **10a** (Fig. 5) and the adduct with ethylenediamine **10b** (see ESI† for X-ray) were formed at room temperature and crystallised (Scheme 7), while phenylboroxine complexes **10c** and **10d** with two or four benzylamine molecules (Scheme 8 and Fig. 5) formed upon reaction of **8a** with two equivalents of benzylamine at reflux in toluene (either with reflux condenser or with Dean–Stark water removal apparatus). An important observation was that the amine promotes dissolution of the boronic acid in organic solvent by formation of these boroxine complexes (see ESI†).

These examples further underlined (also observed for boronic acids, *vide supra*) the importance of B–N interactions, which should readily occur between the amine and any trigonal boron species present in catalytic amidation systems. This is also consistent with our hypothesis that any trigonal acyloxyboron species is likely to react at the boron atom with an amine and indeed, such interactions may facilitate interconversion to and from boroxine species.

Having investigated the various interactions between amines and boronic acids, we next examined reactions between carboxylic acids and boronic acids. It was observed that boronic acid **8c** did not react with phenylacetic acid unless 5 Å MS were introduced into the reaction mixture. In the presence of 5 Å MS alone, boronic acid **8c** was converted into the corresponding boroxine **7c** (Fig. 7A and B). With carboxylic acid present, initial



Scheme 7 Reactions between boronic acids **8a–d** and amines at r.t.



Scheme 8 Reactions between phenylboronic acid **8a** and benzylamine at reflux.

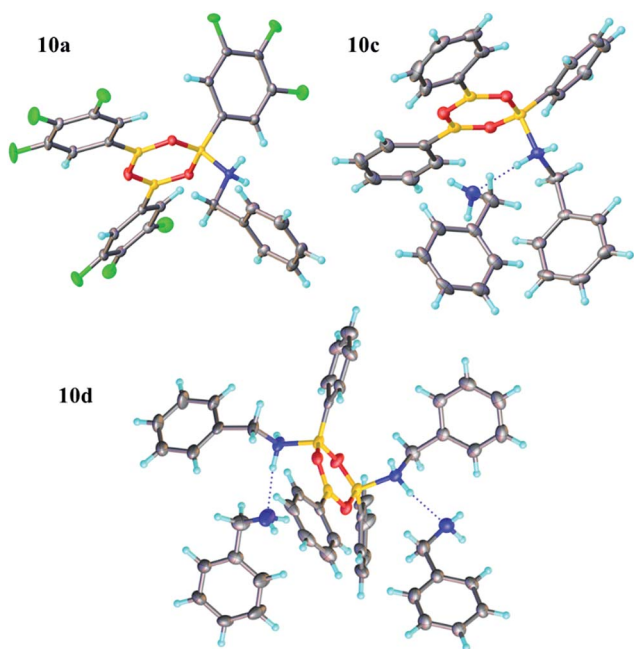


Fig. 5 X-ray structures of 3,4,5-trifluorophenylboroxine-benzylamine adduct **10a** and phenylboronic acid derived boroxine-amine adducts **10c** and **10d**.

formation of boroxine **7c** occurred (Fig. 7C), and as more MS were added as well as phenylacetic acid, a new “ate”-complex (δ 5 ppm, Fig. 7D and E) was formed, and the equilibrium between it and boroxine **7c** was pushed further towards the new species. Careful crystallisation from this heterogeneous system (containing solid 5 Å MS), allowed crystals to be separated which were suitable for X-ray characterisation. The product was unambiguously identified as the “ate”-complex **11c** (see Fig. 6) which is structurally related to analogous complexes reported elsewhere.²² Interestingly, in a separate NMR reaction, addition of an excess of phenylacetic acid to boronic acid **8c**, in the presence of 5 Å MS shifted the equilibrium even further towards the “ate”-complex as seen by ¹¹B NMR (see further examples in the ESI†).

The amount of MS used also had a profound effect, and with a large excess of MS it was possible to generate >85% of the “ate”-complex in the mixture (calculated by ¹¹B NMR, see Fig. 7E and ESI† for further examples). The resulting species **11c** is analogous to “ate”-complexes **6** in the boronic acid series, which undergo subsequent protodeboronation to give boronic acid derivatives rather than act as amidation catalysts. This again highlights the need for the correct boron oxidation state, so that there are sufficient free coordination sites on the boron centre: boronic acids being unreactive but boronic acids/boronic acid being catalytically active.

It should be noted that without 5 Å MS the formation of the diacyloxy B–O–B bridged complexes **11** did not occur to a significant extent; even when 4 Å MS were used, less than 5% of the diacyloxy B–O–B complex **11c** was observed. Interestingly though, a similar diacyloxy B–O–B bridged complex **11d** could



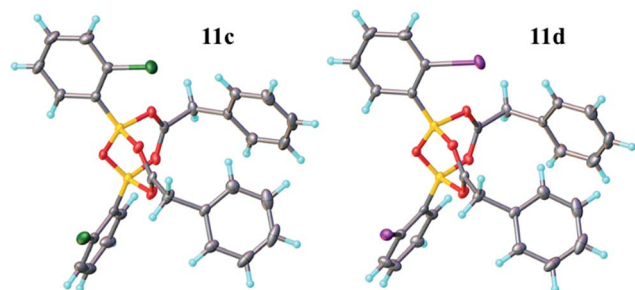


Fig. 6 X-ray structures of B–O–B “ate”-complexes **11c** and **d**.

be isolated by reaction of 2-iodophenylboronic acid **8d** and phenylacetic acid. In contrast, when boronic acids without an *ortho*-substituent, such as 3,4,5-trifluorophenylboronic acid **8a**, were reacted with phenylacetic acid, the “ate”-complex was formed, as shown by ^{11}B NMR, but did not crystallise and only boroxine **7a** crystallised from the reaction mixture. In the case of *ortho*-substituted boronic acids, the boroxines are likely to be destabilized due to steric effects, and thus, the formation of dimeric “ate”-complexes **11** may be more favourable, enabling the formation of catalytically competent species.

Such sterically-induced destabilization of boroxines was indeed confirmed by relative free energy calculations²⁹ (B3LYP+D3/Def2-TZVPP/solvent = dichloromethane) in which

cyclisation of *o*-chlorophenyl boronic acid **8c** to give **7c** + 3H₂O was found to be 2.55 kcal mol⁻¹ more endoenergetic than formation of the boroxine **7e** from *p*-chlorophenyl boronic acid **8e** (Scheme 9). For the former, this is due to reduced conjugation between the aryl rings and the boroxine ring as a result of twisting of $\sim 11^\circ$. At 298 K, this translates into a $\sim 10^5$ -fold decrease in the equilibrium concentration of the boroxine **7**. Indeed, the high catalytic activity of such *ortho*-substituted systems may be due to these key reactivity effects which can destabilise the boroxine, a species which is likely to be an off-cycle resting state of the catalyst.

Interestingly, when we attempted to reproduce the reported¹⁰ generation of an “acyloxy” boron derivative as described by Ishihara *et al.* using 3,4,5-trifluorophenylboronic acid **8a** heated at reflux with phenylacetic acid in toluene (Dean–Stark water



Scheme 9 Calculated ΔG_{298} values for the formation of boroxines from *o*-chlorophenylboronic **8c** acid and *p*-chlorophenylboronic acid **8e**.



Fig. 7 ^{11}B NMR spectra showing interactions between 2-chlorophenylboronic acid **8c**, phenylacetic acid and 5 Å MS. (A) 2-Chlorophenylboronic acid; (B) addition of 5 Å MS to (A), leading to boroxine formation; (C) addition of a tiny amount of 5 Å MS to a mixture of boronic and carboxylic acid; (D) addition of more 5 Å MS; (E) addition of even more 5 Å MS.



removal), there was no obvious reaction by either ^{11}B and ^1H NMR or ReactIR.

Having isolated the diacyloxy B–O–B complexes **11**, we attempted to react these systems with benzylamine (Fig. 8). Surprisingly, when 1 equivalent of amine per dicarboxylate complex **11c** was added, the ^{11}B NMR showed the appearance of a new species **12** with a signal at 3 ppm, but ^1H NMR showed no amide formation even after 2.5 h (see ESI†). However, when a second equivalent of amine was added, amide formation was seen within 20 minutes. This reaction of the diacyloxy B–O–B bridged complex **11c** with two equivalents of benzylamine did not reach completion, and the signals we assign to the new species **12** remained in the ^1H NMR even after 48 h. It is possible that this is due to the fact that there is an excess of boron compound (in total) with respect to amine in the reaction mixture, which indicates that an excess of amine and/or carboxylic acid per catalyst molecule is required for catalytic turnover and efficient reactivity.

We then went on to investigate the species present in catalytic reaction mixtures containing boronic acid, amine and carboxylic acid (Scheme 10). Interestingly, the behavior of these systems was dependent on the order of reagent addition. When the carboxylic acid was premixed with boronic acid **8c** in the presence of 5 Å molecular sieves, rapid formation of the “ate”-complex **11c** was observed (Scheme 10A); subsequent addition of

amine led to complete formation of amide within 2 h. During the amidation reaction, small quantities of boroxine–amine complex **10e** were also observed, along with a signal consistent with the unknown species **12** (δ_{B} 3 ppm). The amine complex **10e** became the dominant species after the amidation was complete. In contrast, when boronic acid **8c** was mixed with the amine in the presence of 5 Å molecular sieves (Scheme 10B), rapid formation of a boroxine amine complex (*e.g.* **10e**) was observed, though a lower ^{11}B NMR shift of 15 ppm suggests that **10e** was likely in equilibrium with a compound analogous to **10d**, in which two amine molecules are coordinated to the boroxine. Addition of carboxylic acid led to amide formation, but after 2 h the reaction had only reached 50% conversion. Again, during the amidation reaction both **10d** and **12** were observed, with the latter being the major species. Finally, a reaction was carried out in which amine and carboxylic acid were pre-mixed to give the ammonium carboxylate salt and boronic acid **8c** was then added (Fig. 10C). This exhibited similar behavior to the reaction shown in Fig. 10B, giving 53% conversion to amide after 2 h.

Finally, when either B_2O_3 or boric acid were mixed with benzylamine and phenylacetic acid (both insoluble in CDCl_3), some dissolution was observed and the resulting ^{11}B NMR showed the appearance of an “ate”-complex (see ESI†), suggesting that these boron compounds can be pulled into solution *via* complexation.

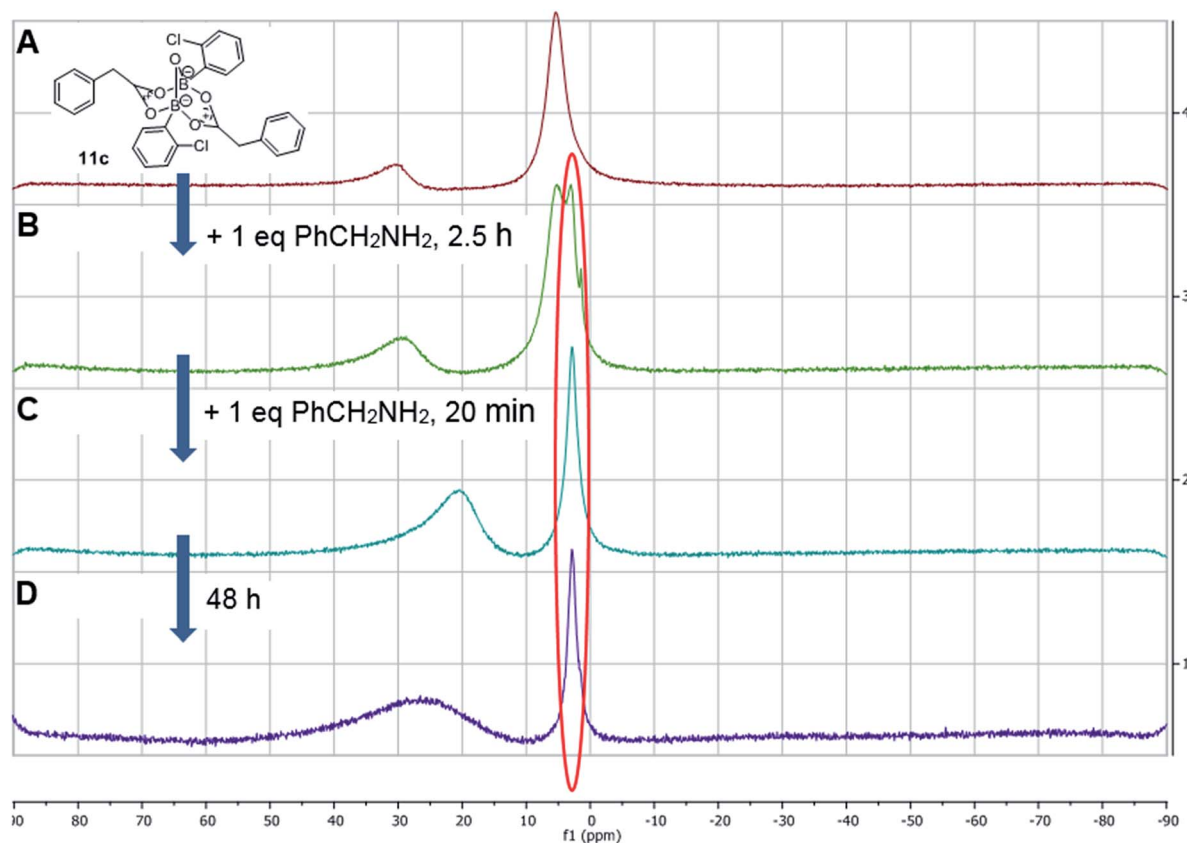


Fig. 8 ^{11}B NMRs of (A) mixture of 2-chlorophenylboronic acid **8c** and phenylacetic acid in the presence of 5 Å MS resulting in equilibrium between B–O–B dicarboxylate **11c** (δ 5 ppm) and boronic species **7c** (δ 30 ppm); (B) addition of 1 equivalent (per “ate”-complex **11c**) of benzylamine, 2.5 h, showing formation of a new complex **12**; (C) addition of 2nd equivalent of benzylamine, 20 min; (D) same mixture over 48 h.



exploration of transition states involving monomeric B–N species resulted in high overall energy barriers (>30.0 kcal mol $^{-1}$). Hence, pathfinder explorations of the potential energy surfaces were then carried out for the different putative B–X–B-mediated catalytic cycles, organised around five possible routes based on the number of amines involved (Fig. 9–13), in order to explore the mechanistic diversity of these systems. Key transition state structures in each cycle are shown together in Fig. 14. To enable a high quality basis set to be employed (Def2-TZVPP), substituents were all modelled at this stage using methyl groups. Each cycle included the reaction free energies of the pre-step involving species used to assemble the reactant used as the start point in the cycle, and the energies of all subsequent steps in the cycle, ending in the final products. A final reaction free energy is included to re-generate the starting reactant ready for a second catalytic cycle.

The cycle involving one amine is illustrated in Fig. 9. The pre-assembly condensation step (blue) indicates a modestly endoenergetic step ($\Delta\Delta G_{298} +9.4$ kcal mol $^{-1}$) producing water, which agrees with the need for water removal using molecular sieves for the reaction to proceed. With the reactant in the cycle set to a relative energy of 0.0, the high point is reached with C–O cleavage (TS3) corresponding to $\Delta G_{298} +23.4$ kcal mol $^{-1}$, corresponding to a relatively slow ambient temperature reaction. The product in this cycle is $+0.3$ kcal mol $^{-1}$ relative to the reactant, but -4.2 kcal mol $^{-1}$ is recovered in the next regeneration step (red). Such a computational exploration can only demonstrate the viability of a single pathway to product; it does not preclude, of course, that lower energy pathways might not be discovered upon a more complete exploration of the potential energy

surface. These variants might include lower energy conformations of the various species explored. Rather more certain is that the free energy barrier surmounted during the cycle is probably an upper calculated bound to any true value. It should also be noted that the model used for the substituents (methyl) may introduce further uncertainty, but since the purpose here was a broad exploration of various mechanistic alternatives, we did not undertake the much more computationally resource-intensive recalculation using larger substituents.

A cycle involving two amine species is illustrated in Fig. 10. The pre-assembly step (blue) is rather more endoenergetic ($\Delta G_{298} +25.4$ kcal mol $^{-1}$) than the previous cycle, but the energetic maximum in the cycle (TS3, $\Delta G_{298} +24.8$ kcal mol $^{-1}$) relative to the assembled reactants is very similar. The product of the first cycle is exoenergetic by $\Delta G_{298} 22.3$ kcal mol $^{-1}$, offset in part by the regeneration step for the second cycle being $\Delta G_{298} +16.5$ endoenergetic.

The three-amine cycle (Fig. 11) again shows similar overall behavior. The condensation reaction resulting in the initial cycle reactant is endoenergetic by $\Delta G_{298} +24.4$ kcal mol $^{-1}$, matched by the exoenergetic products ($\Delta G_{298} -24.8$ kcal mol $^{-1}$), and with a regenerative step for the start of the second cycle being endoenergetic by $\Delta G_{298} +14.1$ kcal mol $^{-1}$. The rate limiting TS3 has $\Delta G_{298} 21.1$ kcal mol $^{-1}$, slightly more favourable than the previous two cycles.

The four-amine cycle (Fig. 12) has a highly endoenergetic condensation step ($\Delta G_{298} +37.9$ kcal mol $^{-1}$), again offset by the products being exoenergetic by 32.4 kcal mol $^{-1}$. The regeneration step for the second cycle is endoenergetic by $+24.9$ kcal mol $^{-1}$ and the overall barrier originating from TS1 is 25.4 kcal mol $^{-1}$. This makes this model cycle less favourable.



Fig. 9 Catalytic cycle involving one amine (data²⁹ sub-collection DOI: 10.14469/hpc/1854). The total computed free energy for each stationary point is shown in Hartree, with the relative energies shown in parentheses in kcal mol $^{-1}$. The DOI for data repository entries for individual species are shown in the form e.g. 10.14469/hpc/1885. Energies in blue are for the pre-assembly stage, in black for the cycle proper and in red for the regeneration step to initiate a new cycle. All DOIs in this figure are available as clickable links (final paginated PDF only).





Fig. 10 Catalytic cycle involving two amines (data²⁹ sub-collection DOI: 10.14469/hpc/1622; the labels have the same meaning as in Fig. 9). All DOIs in this figure are available as clickable links (final paginated PDF only).

Finally, a variation of the one amine cycle involving utilization of a second amine to form a B–N–B catalytic intermediate was examined (Fig. 13), for which the catalyst assembly step is similarly endoenergetic ($\Delta G_{298} +26.0$ kcal mol⁻¹) to the B–O–B bridged two amine route above. The highest free energy point in the catalytic cycle (TS2, $\Delta G_{298} +23.2$ kcal mol⁻¹) is now relative to the

initial tetrahedral intermediate (TI1), which forms exoenergetically from the bicyclic reactant related to the previous cycles, but with a transition state of presumed low barrier corresponding to reactant reorganization which we were unable to locate. This suggests that the barriers resulting from species with B–N–B bridges are similar to those with B–O–B bridges.



Fig. 11 Catalytic cycle involving three amines (data²⁹ sub-collection DOI: 10.14469/hpc/1621; the labels have the same meaning as in Fig. 9). All DOIs in this figure are available as clickable links (final paginated PDF only).





Fig. 12 Catalytic cycle involving four amines (data²⁹ sub-collection DOI: 10.14469/hpc/1728; the labels have the same meaning as in Fig. 9). All DOIs in this figure are available as clickable links (final paginated PDF only).

These five pathfinding explorations have thus identified a model corresponding to a thermal reaction at 298 K, with the caveat that further optimization of these very complex catalytic combinations may always reveal even lower energy routes. In the future, we plan to exploit these core models to study different substituent effects on the energies.

All five of the catalytic cycles investigated have reasonable thermal barriers for the catalytic mechanism (Fig. 14). It is difficult, however, at least purely on the basis of the computed energetics of such catalytic processes, to confidently discriminate between these. Importantly, all these systems have lower overall energetics than the monomeric acyloxyboron pathway

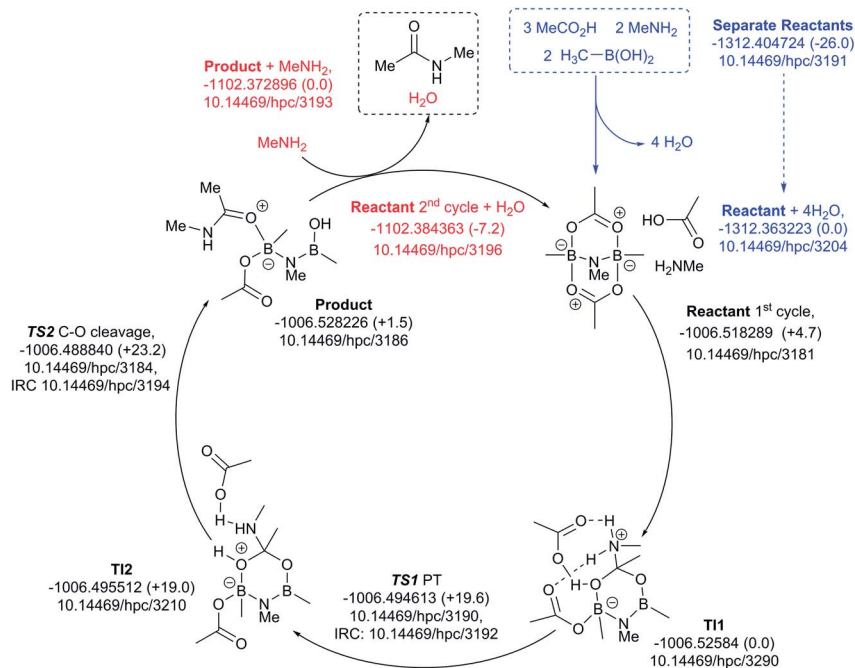


Fig. 13 Catalytic cycle involving B–N–N bridging group (data²⁹ sub-collection DOI: 10.14469/hpc/3188; the labels have the same meaning as in Fig. 9). All DOIs in this figure are available as clickable links (final paginated PDF only).



previously proposed^{14a} and computationally examined.¹⁷ Finally, on the basis of these calculations, the resting state of the catalyst would be expected to be a bridged (B–O–B or B–N–B) system containing tetrahedral boron atoms (e.g. TI2 in Fig. 13), consistent with the signals observed from NMR analysis of catalytic reaction mixtures (δ_B 3 ppm).

Kinetic analysis of a catalytic amidation reaction

Preliminary studies of the reaction kinetics for a boronic acid catalyzed amidation reaction revealed some interesting findings (Scheme 12 and ESI†). The selection of reaction conditions and amine/carboxylic acid reactants was guided by the need for a homogenous reaction mixture in order to obtain useful kinetic data. Reaction of 4-phenylbutylamine with benzoic acid^{12b} was carried out in *tert*-amyl methyl ether (TAME) under Dean–Stark conditions^{10f} using **8c** as the amidation catalyst. Analysis of reaction profiles was carried out using the Burés method (see ESI†).²³ The reaction was first order in catalyst up to a loading of 5 mol%, but a catalytic loading of 10 mol% did not provide a significant rate acceleration over 5 mol%. Interestingly, the first order dependence on catalyst was observed up to 10 mol% loading, when 1.4 equivalents of amine were present. This may suggest that the amine is important for generating the active catalytic species, perhaps related to dimer **12**. A positive order of reaction with respect to the amine concentration was also observed at 3.5 mol% catalyst loading. Unexpectedly, increasing the concentration of carboxylic acid had a negative effect on the reaction rate, and even a slight excess of 1.2 equivalents of carboxylic acid led to almost complete inhibition of the reaction at 3.5 mol% catalyst loading; the effect was less pronounced at higher catalyst loadings.

At present, this observation is difficult to explain, but it appears that excess carboxylic acid prevents formation of the active catalytic species, e.g. *via* formation of off-cycle species such as the ammonium carboxylate salt or dimer **11**. Similar observations regarding catalyst inhibition by carboxylic acid were noted in Zr-catalysed amidation reactions.²⁴

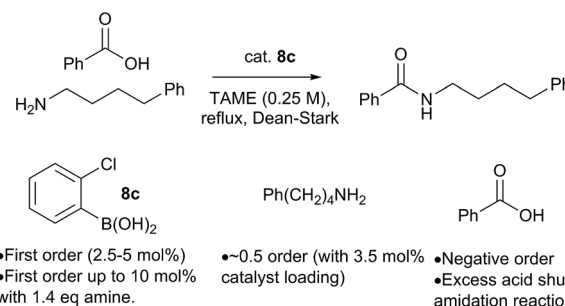
All the mechanisms outlined above, as well as those previously proposed in the literature, involve an essentially unimolecular turnover limiting step, so reaction kinetics do not provide a useful tool for distinguishing between them. The fact that the reaction kinetics show a first order dependence on the catalyst concentration is consistent with either a monomeric or dimeric active catalyst, provided it remains as a monomer/dimer throughout the catalytic cycle.^{23c,24} The positive interaction between the concentration of amine and boronic acid catalyst provides some support for mechanisms involving an active species containing both of these components, especially if its formation is dependent on B–N interactions. Crucially, there appears to be a restriction on the quantity of active catalyst that can be generated from free boronic acid (somewhere between 5–10 mol%), which can be lifted at higher amine concentrations.

Summary and conclusions

We have investigated carboxylic acid/amine/boron catalyst systems experimentally and computationally and have shown



Fig. 14 The computed structure for: (a) TS3 shown in Fig. 9; (b) TS3 shown in Fig. 10; (c) TS3 shown in Fig. 11; (d) TS1 shown in Fig. 12 and (e) TS2 shown in Fig. 13. Indicated bond length are in Å. The activation free energies ΔG_{298} for the five transition states are respectively 23.4, 25.2, 21.1, 25.4 and 23.2 kcal mol⁻¹.



Scheme 12 Preliminary kinetics studies on a boronic acid catalyzed amidation reaction.

the potential importance of both the formation of B–X–B species and of B–N interactions in these systems, leading to the proposal of new catalytic cycles for boron-catalysed amidation reactions involving boronic acids, though by extrapolation, boric acid is highly likely to behave similarly. We have demonstrated that whereas boronic acids can catalyse amidation reactions, boronic acids are ineffective as catalysts



themselves, and in fact, undergo protodeboronation instead to derive boronic acids *in situ*. In turn, the boronic acid analogue is catalytically competent and causes downstream amidation catalysis. This work clearly indicates that at least three free coordination sites are needed on the boron atom to enable effective catalysis to take place, and hence, boronic acids are precluded from amidation reactivity. Indeed, our proposed mechanisms can account for this and other observations regarding amidation catalysis. Firstly, the formation of an active dimeric boron catalyst appears to be impossible at the boronic acid oxidation level, as rapid complexation of an acid and amine molecule leads to the formation of a catalytically inactive complex **5**. The success of *ortho*-substituted boronic acids,¹³ including bifunctional systems,¹² as catalysts can also potentially be rationalized by their preference for the formation of dimeric species over trimeric boroxines. Furthermore, this supports our previous observation that cooperative catalysis using two different boronic acids can lead to enhanced yields of amides in challenging cases.¹⁵ Our proposed mechanisms are also consistent with a recent report that a complex B/N/O heterocyclic system displays high activity in catalytic amidation reactions.¹⁶

The findings in this study indicate the clear importance of bringing together multiple techniques to understand the various entities involved in amidation catalysis, *i.e.* the combined use of NMR (¹H, ¹³C and especially ¹¹B), IR, X-ray crystallography, reactivity studies and the importance of examining multiple potential reaction pathways by theoretical methods, as exemplified by the five related key transition states summarised in Fig. 14. In this case, this comprehensive approach has allowed us to move away from a seemingly 'accepted' mechanism¹⁷ which arguably has little supporting experimental evidence,^{11a} and to propose instead closely related low energy alternatives. We consider that the mechanism shown in Fig. 13 may be the most likely to be effective under catalytic conditions using boronic acids, due to our experimental observations of the reactivity of dimer **11** with amine, and the direct ¹¹B NMR analysis of catalytic reaction mixtures. The key point though is that this is a guide, and individual systems may change between different modes of action depending upon both catalyst and substrate properties (electronics, substituents, *etc.*). However, we anticipate that these mechanistic insights will provide valuable assistance for the design of more active boron amidation catalysts in the future.

Experimental

Computational methods

Pathfinder explorations of the potential energy surfaces for putative catalytic cycles all utilised the Def2-TZVPP basis set²⁵ and the B3LYP density functional augmented with the Grimme D3 dispersion energy correction,²⁶ a combination selected after earlier evaluation for catalytic cycles²⁷ and as implemented in the Gaussian 09, versions D01 and E.01 and Gaussian 16 version A.03 programs. A solvent correction using the CPCM method and dichloromethane parameters was employed throughout and all energies are corrected for thermal contributions and

entropy by calculation of normal mode frequencies, with transition states exhibiting the required single negative force constant. All free energies correspond to a standard state at 298 K of 0.0445 M (1 atm). Intrinsic reaction coordinates (IRCs) were used to define the start and end species of all transition states. Research data management (RDM) is *via* deposition into a collection-based data repository²⁸ where access is *via* the appropriate FAIR data DOI.²⁹

General experimental methods

General experimental details and data files are available *via* a data repository.²⁹ Selected experimental procedures and data are included in the ESI.†

Synthetic methods

All synthetic methods and compound characterization data is available *via* a data repository or in the ESI.†²⁶

X-ray crystallography and NMR spectroscopy

Crystallographic and NMR full data files (in Mpublish format) are available *via* a data repository.²⁹

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- (a) M. F. Lappert, *Chem. Rev.*, 1956, **56**, 959–1064; (b) G. Heller, *Top. Curr. Chem.*, 1986, **131**, 42–78.
- M. J. D. Bosdet and W. E. Piers, *Can. J. Chem.*, 2009, **87**, 8–29.
- E. J. Corey, R. K. Bakshi and S. Shibata, *J. Am. Chem. Soc.*, 1987, **109**, 5551–5553.
- T. R. Kelly, A. Whiting and N. S. Chandrakumar, *J. Am. Chem. Soc.*, 1986, **108**, 3510–3512.
- (a) B. M. Trost, D. O'Krongly and J. L. Belletire, *J. Am. Chem. Soc.*, 1980, **102**, 7595–7596; (b) A. Dal Negro, L. Ungaretti and A. Perotti, *J. Chem. Soc., Dalton Trans.*, 1972, 1639–1643.
- (a) H. Charville, D. Jackson, G. Hodges and A. Whiting, *Chem. Commun.*, 2010, **46**, 1813–1823; (b) H. Charville, D. A. Jackson, G. Hodges, A. Whiting and M. R. Wilson, *Eur. J. Org. Chem.*, 2011, 5981–5990.
- (a) H. Lundberg, F. Tinnis, N. Selander and H. Adolfsen, *Chem. Soc. Rev.*, 2014, **43**, 2714–2742; (b) R. M. de Figueiredo, J.-S. Suppo and J.-M. Campagne, *Chem. Rev.*, 2016, **116**, 12029–12122.



- 8 (a) R. M. Lanigan and T. D. Sheppard, *Eur. J. Org. Chem.*, 2013, 7453–7465; (b) K. Ishihara, *Tetrahedron*, 2009, **65**, 1085–1109; (c) B. M. Monks and A. Whiting, *Sustainable Catalysis: Challenges and Practices for the Pharmaceutical and Fine Chemical Industries*, ed. P. J. Dunn, K. K. Hii, M. J. Krische and M. T. Williams, Wiley, Hobokon, NJ, USA, 2013, ch. 5; (d) F. Ferdousi and A. Whiting, *Green Catalytic Direct Amide Bond Formation in Green and Sustainable Medicinal Chemistry: Methods, Tools and Strategies for the 21st Century Pharmaceutical Industry*, ed. L. Summerton, H. F. Sneddon, L. C. Jones and J. H. Clark, RSC Green Chemistry series, Cambridge, 2016, ch. 13.
- 9 P. Tang, *Org. Synth.*, 2005, **81**, 262.
- 10 (a) P. Starkov and T. D. Sheppard, *Org. Biomol. Chem.*, 2011, **9**, 1320–1323; (b) R. M. Lanigan, P. Starkov and T. D. Sheppard, *J. Org. Chem.*, 2013, **78**, 4512–4523; (c) V. Karaluka, R. M. Lanigan, P. M. Murray, M. Badland and T. D. Sheppard, *Org. Biomol. Chem.*, 2015, **13**, 10888–10894; (d) R. M. Lanigan, V. Karaluka, M. T. Sabatini, P. Starkov, M. Badland, L. Boulton and T. D. Sheppard, *Chem. Commun.*, 2016, **52**, 8846–8849; (e) T. Maki, K. Ishihara and H. Yamamoto, *Org. Lett.*, 2006, **8**, 1431–1434; (f) M. T. Sabatini, L. T. Boulton and T. D. Sheppard, *Sci. Adv.*, 2017, **3**, e1701028.
- 11 (a) K. Ishihara, S. Ohara and H. Yamamoto, *J. Org. Chem.*, 1996, **61**, 4196–4197; (b) K. Ishihara, S. Ohara and H. Yamamoto, *Macromolecules*, 2000, **33**, 3511–3513; (c) K. Ishihara, S. Ohara and H. Yamamoto, *Org. Synth.*, 2002, **79**, 176–185; (d) T. Maki, K. Ishihara and H. Yamamoto, *Synlett*, 2004, 1355–1358; (e) A. Sakakura, R. Yamashita, T. Ohkubo, M. Akakura and K. Ishihara, *Org. Lett.*, 2011, **13**, 892–895; (f) A. Sakakura, R. Yamashita, M. Akakura and K. Ishihara, *Aust. J. Chem.*, 2011, **64**, 1458–1465; (g) R. Yamashita, A. Sakakura and K. Ishihara, *Org. Lett.*, 2013, **15**, 3654–3657; (h) K. Ishihara and Y. Lu, *Chem. Sci.*, 2016, **7**, 1276–1280.
- 12 (a) K. Arnold, B. Davies, R. L. Giles, C. Grosjean, G. E. Smith and A. Whiting, *Adv. Synth. Catal.*, 2006, **348**, 813–820; (b) K. Arnold, A. S. Batsanov, B. Davies and A. Whiting, *Green Chem.*, 2008, **10**, 124–134; (c) K. Arnold, B. Davies, D. Héroult and A. Whiting, *Angew. Chem., Int. Ed.*, 2008, **47**, 2673–2676.
- 13 (a) R. M. Al-Zoubi, O. Marion and D. G. Hall, *Angew. Chem., Int. Ed.*, 2008, **47**, 2876–2879; (b) N. Gernigon, R. M. Al-Zoubi and D. G. Hall, *J. Org. Chem.*, 2012, **77**, 8386–8400.
- 14 T. M. El Dine, J. Rouden and J. Blanchet, *Chem. Commun.*, 2015, **51**, 16084–16087.
- 15 S. Liu, Y. Yang, X. Liu, F. K. Ferdousi, A. S. Batsanov and A. Whiting, *Eur. J. Org. Chem.*, 2013, 5692–5700.
- 16 H. Noda, M. Furutachi, Y. Asada, M. Shibasaki and N. Kumagai, *Nat. Chem.*, 2017, **9**, 571–577.
- 17 (a) T. Marcelli, *Angew. Chem., Int. Ed.*, 2010, **49**, 6840–6843; (b) C. Wang, H.-Z. Yu, Y. Fu and Q.-X. Guo, *Org. Biomol. Chem.*, 2013, **11**, 2140–2146.
- 18 P. Zhang, R. M. Krieger and J. W. Frost, *ACS Sustainable Chem. Eng.*, 2016, **4**, 6991–6995.
- 19 (a) P. Paetzold, P. Bohm, A. Richter and E. Scholl, *Z. Naturforsch.*, 1976, **31b**, 754–764; (b) G. H. L. Nefkens and B. Zwanenburg, *Tetrahedron*, 1983, **39**, 2995–2998.
- 20 B. Garrigues and M. Mulliez, *J. Organomet. Chem.*, 1986, **314**, 19–24.
- 21 A. Schäfer, W. Saak, D. Haase and T. Müller, *Angew. Chem., Int. Ed.*, 2012, **51**, 2981–2984.
- 22 (a) B. Wrackmeyer, E. Khan and R. Kempe, *Z. Naturforsch., B: J. Chem. Sci.*, 2008, **63**, 275–279; (b) R. Köster, A. Sporzyński, W. Schüßler, D. Bläser and R. Boese, *Chem. Ber.*, 1994, **127**, 1191–1199.
- 23 (a) J. Burés, *Angew. Chem., Int. Ed.*, 2016, **55**, 2028–2031; (b) J. Burés, *Angew. Chem., Int. Ed.*, 2016, **55**, 16084–16087; (c) J. Burés, *Top. Catal.*, 2017, **60**, 631–633.
- 24 H. Lundberg, F. Tinnis, J. Zhang, A. G. Algarra, F. Himo and H. Adolfsson, *J. Am. Chem. Soc.*, 2017, **139**, 2286–2295.
- 25 F. Weigend and R. Ahlrichs, *Phys. Chem. Chem. Phys.*, 2005, **7**, 3297–3305.
- 26 S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, **132**, 154104–154118.
- 27 A. Armstrong, R. A. Boto, P. Dingwall, J. Contreras-García, M. J. Harvey, N. J. Mason and H. S. Rzepa, *Chem. Sci.*, 2014, **5**, 2057–2071.
- 28 M. J. Harvey, A. McLean and H. S. Rzepa, *J. Cheminf.*, 2017, **9**, 4.
- 29 S. Arkhipenko, A. S. Batsanov, M. T. Sabatini, V. Karaluka, T. D. Sheppard, H. S. Rzepa and A. Whiting, *Imperial College Research Computing Services data repository*, 2016, DOI: 10.14469/hpc/1620 and sub-collections therein.

