



Cite this: *Phys. Chem. Chem. Phys.*,
2024, 26, 14960

Mechanistic understanding of the thermal-assisted photocatalytic oxidation of methanol-to-formaldehyde with water vapor over Pt/SrTiO₃†

Michel Deitermann,^{ab} Takuma Sato,^{ib} Yannik Haver,^{ib} Alexander Schnegg,^{ib} Martin Muhler^{ib} and Bastian Timo Mei^{ib}*^{ac}

Anaerobic thermal-assisted photocatalytic methanol conversion in the gas phase in the presence of water vapor has been suggested as an interesting way to generate formaldehyde as a valuable coupled product in addition to H₂ production. Here, the reaction mechanism and photocatalyst deactivation are investigated in detail using *in situ* diffuse reflectance infrared fourier transform (DRIFTS) and electron paramagnetic resonance (EPR) spectroscopy. EPR shows that paramagnetic oxygen vacancies are not involved in the reaction mechanism over undoped SrTiO₃. Instead, on an optimized 0.1 wt% Pt/SrTiO₃ photocatalyst, methoxy species are formed by dissociative adsorption of methanol leading to formaldehyde formation while the formation of CO, CO₂ (via a formate intermediate) and methyl formate occurs through three concurrent reactions from formyl species. Our findings suggest that CO adsorbed on Pt is a spectator species not perturbing the reaction kinetics, and deactivation is shown to be strongly correlated with the accumulation of formate groups on SrTiO₃, which is more pronounced at high reaction temperatures. The mechanistic understanding provided here forms the basis for the further optimization of photocatalysts to increase methanol conversion and improve formaldehyde selectivity.

Received 14th March 2024,
Accepted 3rd May 2024

DOI: 10.1039/d4cp01106f

rsc.li/pccp

Introduction

Photocatalytic conversion of small alcohols is conceived as a means to release chemically stored H₂ in a sustainable manner. Thus, photocatalytic methanol oxidation over metal oxide-based photocatalysts in both liquid^{1,2} and gas phase^{3–6} has been extensively studied with a strong emphasis on material design to improve H₂ evolution rates.^{1–6} Despite the possibility of co-creation of value-added alcohol oxidation products, there has been little emphasis on the precise distribution and selective formation of valuable oxidation products.

Recently, it was shown that performing methanol oxidation in the gas phase allows for modulation of the product distribution. For example, in the presence of O₂, oxidation of methanol

to formic acid and CO₂ is favored,^{7,8} but the reduction of O₂ is competing with the reduction of protons (H⁺) to H₂.^{7,8} Using anaerobic reaction conditions with water as oxidant, photocatalytic conversion of methanol in the gas phase over Pt/SrTiO₃ (Pt/STO) leads to the formation of valuable chemicals like formaldehyde or methyl formate in a distinctive temperature window <120 °C.⁹ Operation at temperatures above 120 °C resulted in a higher selectivity to CO at the expense of formaldehyde and methyl formate selectivity, and the photocatalyst was shown to suffer from severe deactivation with time on stream.⁹ Moreover, methanol conversion was found to be dependent on the Pt loading and 0.1 wt% of Pt deposited by reductive photodeposition resulted in the highest conversion of methanol.

Photocatalyst deactivation in photocatalytic methanol oxidation has been less of a concern in relevant mechanistic studies,^{10–16} and the long-term stability (hours–days) of photocatalysts at relevant reaction conditions like elevated temperatures has been barely addressed. If at all, catalyst deactivation is primarily explained by the accumulation of reaction intermediates.^{17,18} For example, formation of CO in photocatalytic methanol oxidation is frequently reported as side product, but the effects on the catalytic performance remain under debate.^{16,19–21} Based on temperature-programmed desorption

^a Laboratory of Industrial Chemistry, Ruhr University Bochum, 44780 Bochum, Germany. E-mail: Bastian.Mei@ruhr-uni-bochum.de

^b Max Planck Institute for Chemical Energy Conversion, 45470 Mülheim an der Ruhr, Germany

^c Photocatalytic Synthesis Group, Faculty of Science & Technology of the University of Twente, PO Box 217, Enschede, The Netherlands

† Electronic supplementary information (ESI) available: Experimental set-up and procedure of DRIFTS measurements, XP spectrum, DRIFTS spectra, EPR spectra with detailed analysis and simulations, additional pulse experiments, water gas shift experiments. See DOI: 10.1039/d4cp01106f



(TPD) experiments, a high coverage of CO on Pt was reported by Walenta *et al.*¹⁹ using Pt/TiO₂(110) catalysts. CO was suggested to be a spectator, thus not inhibiting the H₂ formation rate. In contrast, Chiarello *et al.*¹⁶ reported the irreversible adsorption of CO on Pt nanoparticles (NPs) (> 2 nm) during methanol reforming with water vapor causing catalyst poisoning.

Others^{20,21} argued that CO poisoning of the Pt sites is only problematic under dark conditions, as CO is constantly removed during illumination. Additionally, the water gas shift reaction (WGSR) was proposed to efficiently remove CO from the catalyst and lower the CO selectivity, which has been observed particularly for Cu-containing photocatalysts used at higher reaction temperatures.²²

Similarly, evidence for the presence of formate on the photocatalyst surface has already been provided with the extent of formate group formation and consumption being dependent on the external reaction conditions like temperature or methanol partial pressure in the reactive gas mixture.^{4,11,13} Yet formate groups are mainly considered to be the intermediate species in CO₂ and methyl formate formation, which are slow reaction steps compared with the formation of formate groups.^{11,13,14} Consequently, the accumulation of this species was observed but, so far, treated as reaction intermediate only.¹⁶ The influence of formate on catalyst deactivation by blocking of active sites was not further investigated. Matsubu *et al.*²³ did even report a positive effect of strongly bound HCO_x species as they limited the degree of the encapsulation of Rh particles by reducible supports and enabled the dynamic tuning of the selectivity during the thermally catalyzed CO₂ reduction.

Besides poisoning by surface species, deactivation in heterogeneous catalysis is often related to changes in the surface and bulk structure of the catalyst.²⁴ In photocatalysis, however, formation of defects, like oxygen vacancies (OVs), in the crystal structure is usually considered to enhance visible light absorption or enhance reactant adsorption, in turn enhancing their activity.^{25,26} For reducible oxides like TiO₂ and SrTiO₃ (STO), defects have been engineered using reducing agents like H₂,²⁵ solid carbon,²⁶ NaBH₄²⁷ or plasma treatments.²⁸ Haselmann *et al.*,¹⁷ however, also indicated that early-stage deactivation of Pt/TiO₂ photocatalysts in aqueous methanol reforming strongly correlates with the presence of OVs induced by either ultrasonic pretreatment or *in situ* UV illumination. Moreover, encapsulation of the Pt NPs was observed that might also result in photocatalyst deactivation.^{29,30} Thus, the role of defects in the deactivation of photocatalyst is yet to be explored in detail.

Considering the lack of a consistent description of the reaction mechanism of thermal-assisted photocatalytic (selective) methanol oxidation and photocatalyst deactivation, this work explores the reaction and deactivation mechanism of gas-phase methanol oxidation in the presence of water vapor using diffuse reflectance infrared Fourier transform spectroscopy (DRIFTS) over Pt/STO photocatalysts with a loading of 0.1 wt% Pt. *In situ* DRIFTS performed under illumination and at elevated reaction temperatures (100–200 °C) suggest that formate formation is strongly correlated with catalyst deactivation. Thus, it should be considered as inhibitor rather than spectator which is also supported by

transient pulse measurements. Additionally, detailed electron paramagnetic resonance (EPR) measurements were performed and critically discussed with literature evidencing the minor role of defects, like OVs, in the reaction mechanism and particular catalyst deactivation over undoped STO.

Experimental

Photocatalyst synthesis

The photocatalyst was prepared by photodeposition (PD) of Pt using commercial STO (Sigma Aldrich, 99%). PD was carried out in a continuously stirred double-walled liquid-phase reactor equipped with a Hg lamp (500 W, 59 mW cm⁻², UV Consulting Peschl) described in detail elsewhere.³¹ The PD of 0.1 wt% Pt was performed using a suspension of 2 g STO in 500 mL water (HPLC grade) and 50 mL methanol (VWR Chemicals, 100%) containing H₂PtCl₆ (Sigma Aldrich 99.995%) in the respective amount.⁹ During illumination (1 h), the temperature of the reactor walls was maintained at 30 °C while the immersion Hg lamp was cooled to 10 °C. The 0.1Pt/STO photocatalyst was obtained after filtration, washing (HPLC grade water, > 500 mL) and freeze-drying (−50 °C for 24 h).

Gas-phase photocatalytic measurements and pulse experiments

All gas-phase photocatalytic tests were performed with the set-up described elsewhere.⁹ The measurement procedure is described in detail in the ESI.† Long-term experiments under quasi steady-state conditions were performed by continuous dosing of methanol and/or water. Furthermore, the set-up allows for N₂ (99.999%, air liquide), CO₂ (99.995%, air liquide), CO (4% in N₂, 99.999%, air liquide), and O₂ (99.999%, air liquide) pulsing during activity testing using an integrated pulse loop (~7 mL pulse volume). The standard reaction conditions used for pulse experiments, mechanistic studies and general understanding of photocatalyst deactivation are summarized in Table 1.

EPR measurements

X-band EPR spectra were collected using an MS-5000 (Magnettech GmbH, Freiberg Instruments) operated at either 100 K or RT. The STO (Sigma Aldrich) or 0.1Pt/STO sample (~10 mg each) was placed in a 3.8 mm EPR tube. In selected measurements 50 μL of dry methanol (VWR Chemicals, 100%) was added. Illumination of the samples was carried out with a Hg UV lamp (Oriel instruments, 200 W, 280–400 nm) at RT and subsequent fast cooling in

Table 1 Standard reaction conditions used for catalytic tests and pulse experiments in the gas phase

Reaction condition	Value
Reactor temperature $T / ^\circ\text{C}$	100
UV light intensity $I / \text{mW cm}^{-2}$	52.9
Total volumetric gas flow $\dot{V}_{\text{total}} / \text{mL min}^{-1}$	50
Methanol mole fraction $y_{\text{MeOH}} / \%$	1.5
Water mole fraction $y_{\text{H}_2\text{O}} / \%$	4.5



LN₂ to avoid the possibility of defect passivation until the measurement was started. If not stated otherwise, the spectra were recorded at a microwave frequency of ~ 9.64 GHz, in a magnetic field (B_0) of 337 mT with a sweep range of ± 30 mT, field modulation of 0.55 mT, microwave power of 0.2 mW (27 dB attenuation), and a field modulation frequency of 100 kHz averaging 9 scans to increase the signal-to-noise ratio. The EPR spectra were simulated using the MATLAB (2021) package *EasySpin*.³² Reduction of 0.1Pt/STO was performed in an EPR tube connected to a Schlenk line for 2.5 h at 150 °C in flowing H₂ (100%). Additionally, the reduction of STO (250 mg) was carried out at 750 °C in flowing 2% H₂/N₂ (20 mL min⁻¹) for 12 h in a tubular furnace to achieve the formation of OV's similar to a procedure described in literature.²⁵ The sample was obtained as fine powder with grey color which was maintained after storage at RT for >2 days.

DRIFTS measurements

DRIFTS measurements were performed using a high temperature reaction chamber (HTRC). The spectrometer and the HTRC were purged with He (10 mL min⁻¹) for at least 30 min before the measurements. The actual measurement sequence included heating of the samples to 200 °C to efficiently remove surface adsorbed water and subsequent cooling of the photocatalyst in flowing He to the desired adsorption or reaction temperature. Afterwards, adsorption experiments of water and/or methanol on the 0.1Pt/STO surface in the dark were performed by allowing He to pass through a saturator containing water, methanol, or a water/methanol mixture. The photocatalyst was allowed to equilibrate with the respective gas phase for at least 30 min. Subsequently, pure He was used for purging the cell to efficiently remove weakly bound species. The difference spectra were obtained using background spectra obtained at the respective temperature in He only. Constant flow experiments using methanol/water mixtures were performed in the dark and under illumination (UV LEDs). Here, background subtraction was performed using a spectrum obtained after 1.5 h of constant purging in the dark with a saturated methanol/water in He mixture. Independent of the measurement procedure, a good signal-to-noise ratio was obtained by averaging 64 scans (3 min per spectrum). A more detailed description of the set-up and settings of the FTIR spectrometer used for DRIFTS experiments can be found in the ESI†

Results and discussion

Photocatalytic methanol oxidation in the gas phase

STO-based photocatalysts are widely employed in various (photocatalytic) reactions.^{33–36} Here, the thermal-assisted methanol oxidation in the gas phase has been performed using 0.1Pt/STO that was previously shown to be most active in methanol conversion and highly selective for formaldehyde formation.⁹ Extensive material characterization using N₂ physisorption measurements (BET method), X-ray diffraction (XRD), UV-Vis diffuse reflectance spectroscopy (UV-Vis DRS), inductively

coupled plasma optical emission spectroscopy (ICP-OES), high resolution transmission electron microscopy (HR-TEM), and energy-dispersive X-ray spectroscopy (EDX) confirmed the phase purity of STO and successful deposition of Pt NPs. X-ray photoelectron spectroscopy (XPS) measurements revealed that the Pt NPs are mainly in the Pt⁰ state (see ESI†).

The time-resolved evolution rates of relevant methanol oxidation products obtained with 0.1Pt/STO in the thermal-assisted photocatalytic methanol oxidation (100–175 °C, UV LED illumination) are summarized in Fig. 1 using freshly coated catalyst plates for each reaction temperature. Each measurement was performed after pretreatment in humid N₂ to remove organic residues that might originate from previous preparation steps (Fig. S8, ESI†). Clearly, with time-on-stream (TOS) deactivation is observed, which appears to be most pronounced for temperatures >100 °C. With increasing temperature, formaldehyde selectivity is compromised at the expense of CO/CO₂ formation and above a reaction temperature of 150 °C the evolution rate of formaldehyde after 6 h on stream is clearly inferior to the corresponding stable rate obtained at 100 °C. A similar trend, though with a more significant decrease in initial evolution rate, is also observed for methyl formate, while overall for CO and CO₂ a consistent increase in their respective evolution rates is revealed. Independent of the reaction temperatures deactivation of the photocatalyst resulted in a decrease in methanol conversion leading to a similar conversion after 6 h on stream (see Fig. S9, ESI†).

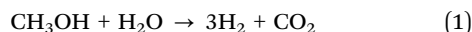
Considering the negligible deactivation observed at 100 °C, a long-term experiment was performed (49 h of continuous illumination, see Fig. S10, ESI†). A low activity due to thermal methanol reforming (MR) (eqn (1)) is observed prior to



Fig. 1 Temperature variation from 100 °C to 175 °C (single catalyst plates); temporal course of evolution rates of H₂, formaldehyde, CO and methyl formate during UV illumination using a 0.1Pt/STO coated glass substrate in a flat-plate reactor. Smoothing (moving average over 10 data points, 8.5 min) was applied to the formaldehyde and methyl formate evolution rates at 125 °C and 175 °C. The methanol conversion X_{MeOH} and CO₂ evolution rates can be found in Fig. S9 (ESI†).



illumination, while an immediate increase in formaldehyde and methyl formate is detected once the photocatalyst is illuminated.^{4,9}



Formation of carbon-based products is accompanied by an initial excessive increase in H₂ evolution rate which was likely caused by a high surface coverage of reaction intermediates established during the dark period prior to illumination. Although the peak in the H₂ evolution rate was not accompanied by an increase in the evolution rate of carbon-containing products, no O₂ was detected excluding the possibility of overall water splitting as source for H₂. Moreover, the extent of H₂ evolution appeared to be pre-treatment dependent (Fig. S11, ESI[†]). Terminating UV illumination after 49 h TOS, the product evolution rates dropped below the detection limit, or their respective thermal activity observed before illumination. Interestingly, the CO₂ evolution rate briefly increased (see Fig. S10, ESI[†]), and for both methyl formate and CO₂, a slower decrease (tailing) was observed. Methyl formate formation was not observed in the dark period before illumination, suggesting that intermediates are formed exclusively during illumination that further convert thermally. Considering that a decrease in activity has been noticed during extended TOS even at a reaction temperature of 100 °C, it is reasonable to assume that in the dark period abundant surface species are involved in methyl formate formation and CO₂ evolution. Those surface species are presumably exclusively formed during illumination and enrich on the photocatalyst surface leading to deactivation during illumination.

Two temperature regimes are defined allowing to study the reaction mechanism and photocatalyst deactivation: at ≤100 °C a quasi-stable methanol conversion (negligible deactivation) and selective formaldehyde formation allows for evaluation of the reaction mechanism; ≥100 °C total oxidation of methanol is preferred and photocatalyst deactivation can be accessed.

EPR spectroscopy

To monitor the role of paramagnetic defects and trap states in the photoinduced charge transfer processes, EPR spectra were recorded on pristine STO, after reduction at 750 °C and on illuminated STO and Pt/STO in the presence and absence of MeOH (see Fig. S2 and the ESI[†] for further details and simulations carried out to deconvolute the spectra). The 100 K EPR spectrum of STO (Fig. 2a) is dominated by an EPR line with nearly axial *g*-tensor (*g*₁ = 2.0021, *g*₂ = 2.0085 and *g*₃ = 2.0910), which is assigned to O₂[−] adsorbed on the STO surface.^{37–39} In addition, an isotropic signal from Ti³⁺ (*g*_{iso} = 1.978) in the STO lattice is observed.⁴⁰ STO and Pt/STO samples exhibit the same EPR spectra (see Fig. S12, ESI[†]). Reduction of the sample at 750 °C (Fig. 2b), drastically alters the EPR spectrum. The O₂[−] signal strongly reduces in intensity, while an intense isotropic EPR line centered at *g*_{iso} = 2.0039, appears. This signal has recently been ascribed to Ti³⁺/OV defects in STO annealed in H₂ at 500 °C.⁴¹ In addition, the isotropic Ti³⁺ signal

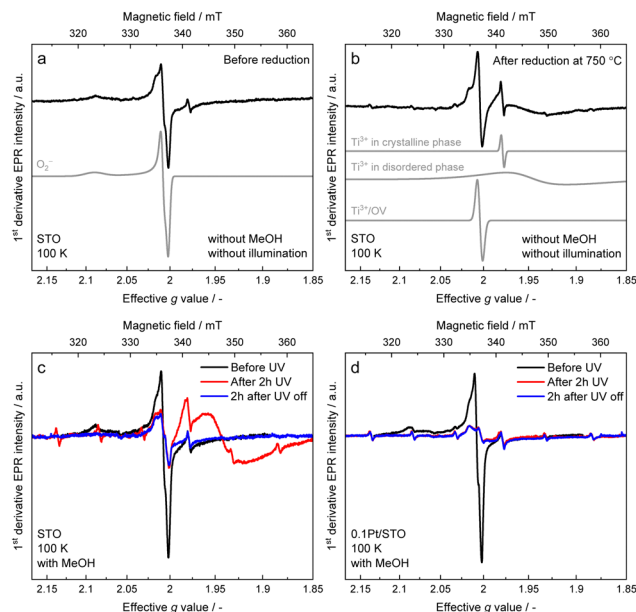


Fig. 2 X-band EPR spectra on pristine STO (a), after reduction at 750 °C (b), on illuminated STO and Pt/STO in the presence of MeOH (c) and (d). Simulations of O₂[−] (a), Ti³⁺ in crystalline phase, Ti³⁺ in disordered phases, and Ti³⁺/OV (b) are plotted offset as grey lines.

at *g*_{iso} = 1.9784 gains in intensity and an additional broad signal (*g*_⊥ = 1.95 and *g*_∥ = 1.93, line width of ~50–95 mT) contributes to the EPR spectrum. The latter originates from Ti³⁺-sites in disordered STO phases, where pronounced site-to-site disorder leads to a broad distribution of *g*-values (*g*-strain).^{40,42,43} Furthermore a six line spectrum with a hyperfine splitting of 237 MHz is observed, which matches very well with reported values for Mn²⁺ (*S* = 5/2, *I* = 5/2) substituting Ti⁴⁺ sites.⁴⁴ UV light exposure in the absence of MeOH only led to minimal changes of the EPR spectrum (see Fig. S12, ESI[†]). However, UV irradiation in the presence of MeOH significantly altered the EPR spectra of both STO (Fig. 2c) and Pt/STO (Fig. 2d). In STO in presence of methanol, UV irradiation led to the disappearance of the O₂[−] signal, which does not reappear after switching off the light and exposure of the sample to air. The most significant change is a reversible increase of the Ti³⁺ defects in disordered STO phases (broad EPR signal *g*_⊥ = 1.95 and *g*_∥ = 1.93). After switching off the light source, the broad Ti³⁺ signal disappears. The metastable Ti³⁺ is assigned to the reduction of Ti⁴⁺ sites in the disordered surface areas of the STO by light generated electrons.⁴⁰ Reduction by photoelectrons requires the presence of MeOH as a hole scavenger to avoid direct recombination of electrons and holes and is thus directly related to photocatalysis. This assignment is further supported by the spectral changes observed for the Pt/STO samples. While the Pt/STO spectrum after illumination and in the presence of MeOH also exhibits a strong decrease of the O₂[−] signal, no surface Ti³⁺ states are formed. The absence of the metastable Ti³⁺ may be rationalized by the fact that in Pt/STO photo-generated electrons are trapped at the Pt–STO interface. Thus, EPR confirms the role of methanol as hole scavenger and Pt as



electron trap. However, no evidence for localized paramagnetic OV in the reaction or photocatalyst deactivation is obtained. OVs in the crystal lattice (bulk or surface) of STO were supposed to strongly influence photocatalytic activity in various reactions.^{25–28,34,35} A high number of surface OVs is considered beneficial by introducing specific adsorption sites for reactants or by changing the light absorption properties. Typically, these crystal defects cannot be induced under mild conditions as applied in this work requiring high temperature treatments (>400 °C) and strong reducing agents. Additionally, surface OVs are expected to readily re-oxidize in the presence of water.²³ Nevertheless, it was reported that Pt NPs can activate hydrogen which can induce OV formation in STO even at much lower temperatures (~150 °C).²⁵ Depending on the Fermi energy, defects can take different charged or neutral states. In STO, paramagnetic OV (V_{O}^{\cdot}) are only predominant in a relatively narrow window of the Fermi energy around 2.5 eV, while diamagnetic states (V_{O}^0 and V_{O}^{2+}) are predominant over most of the Fermi energy region.⁹ The absence of a clearly observable OV-EPR signal in STO and Pt/STO before and after MeOH oxidation studied herein, could therefore either be due to a very low OV abundance or due to OVs in diamagnetic states.

IR spectroscopic investigations

To study the type, relative amount and stability of carbon-based adsorbates, DRIFTS measurements were performed under thermal and thermal-assisted photocatalytic conditions. The band assignment summarized in Table 2 is based on adsorption experiments with the single components methanol and water (Fig. S16 and S17, ESI[†]), ATR-IR measurements with pure components (Fig. S18, ESI[†]) and supported by literature reports.

The spectra recorded during adsorption of methanol and water at 100 °C on 0.1Pt/STO and subsequent reactor purging are shown in Fig. 3. Upon contact of 0.1Pt/STO with the

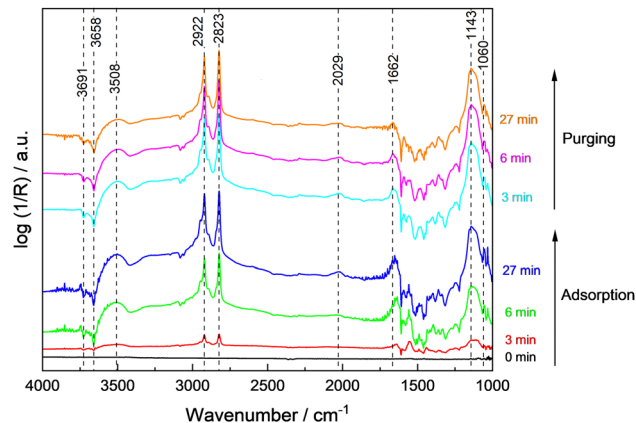


Fig. 3 DRIFTS spectra obtained during adsorption of methanol and water on 0.1Pt/STO at 100 °C and subsequent purging with He. As background spectra, the clean surface before the contact to methanol and water at 100 °C was selected.

methanol–water mixture, bands related to physisorbed methanol, water and methoxy groups were observed suggesting that dissociative adsorption of methanol occurs. All bands were readily formed, and their maximum intensity was obtained already after ~6 min pointing to a fast dissociative adsorption of methanol and water on STO.

During adsorption, the bands assigned to hydroxyl groups on the surface rapidly decreased indicating their consumption or replacement. Experiments performed with bare STO (Fig. S19, ESI[†]) indicate that adsorption and further reaction of adsorbates mainly occurs on the bare STO surface rather than at Pt sites in agreement with the general perception that thermal dissociation of methanol to methoxy groups is feasible on all STO facets in the presence of surface hydroxyl groups.^{52–54} It is also frequently suggested that the formation of methoxy

Table 2 IR bands of the identified species adsorbed on 0.1Pt/STO and STO with the respective assignments based on literature. The exact band positions may differ slightly depending on the sample and measurement. (ν_s and ν_{as} : symmetric and antisymmetric stretching vibrations, δ : in-plane deformation modes)

Species	Absorption mode	Wavenumber / cm^{-1}	Ref.
Hydroxyl groups	$\nu(\text{O-H})$	3659	16 and 45
		3691	
Physisorbed water Combination band	$\nu(\text{H-O-H})$	3800–3200	16 and 45–47
	$\delta(\text{H-O-H})$	~2040	
		1635	
Gas-phase water	$\nu(\text{H-O-H})$	3950–3500	48
	$\delta(\text{H-O-H})$	1900–1300	
Physisorbed methanol	$\nu_{as}(\text{C-H})$	2950	11
	$\nu_s(\text{C-H})$	2842	
	$\nu(\text{C-O})$	1033	
	$\nu_{as}(\text{C-H})$	2921	
Methoxy groups	$\nu_s(\text{C-H})$	2814	11, 16, 49 and 50
	$\nu(\text{C-O})$	1060	
	$\nu(\text{C-O})$	1157/1131	
	$\nu(\text{C-H})$	2864	
Formate groups	$\nu_{as}(\text{C-O})$	1610 + 1569	16 and 22
	$\nu_s(\text{C-O})$	1368, 1310	
	$\nu(\text{M-C-O})$	2060	
Pt-CO (linear)		2060	46 and 47
Gas-phase CO_2	$\nu_{as}(\text{O=C=O})$	2347 (double band)	48
SrCO_3	$\nu_{as}(\text{C-O})$	1451 (see Fig. S16, ESI)	51
Not assigned	—	1493	—



groups occurs primarily on highly defective/reduced surfaces which seems unlikely based on the presented EPR measurements (Fig. 2). The band at $\sim 1033\text{ cm}^{-1}$ disappeared indicating the (partial) desorption of weakly bound methanol. Also, the bands related to physisorbed water gradually decreased. These results point to a competitive adsorption of methanol and water as methoxy groups and hydroxyl groups on the same STO surface sites. In addition to methoxy groups, a weak and broad band at $\sim 2030\text{ cm}^{-1}$ was observed. As this band was observed for 0.1Pt/STO (Fig. 3) and STO (Fig. S19, ESI[†]) during methanol/water adsorption, it is likely that the band is not related to CO adsorbed on Pt. Moreover, CO adsorbed on cationic sites like Ti^{4+} or Sr^{2+} may cause a more significant shift to even higher wavenumbers.⁴⁶ Thus, we assigned the band to a combination band of the O–H–O scissoring and a broad liberation band.⁴⁷ Overall, methoxy appears to be the predominant adsorbate present in the dark at $100\text{ }^\circ\text{C}$.

The thermal-assisted photocatalytic reaction at elevated temperatures ($100\text{ }^\circ\text{C}$) using UV light irradiation was studied with a constant flow of methanol and water in He (Fig. 4). No significant changes of the methanol- and water-related bands were observed due to the fast adsorption and regeneration of the consumed surface species from the gas phase. The band related to CO adsorbed on Pt at 2060 cm^{-1} slowly increased, and formate bands at 1610 cm^{-1} , 1569 cm^{-1} and 1368 cm^{-1} were detected at an increasing intensity with illumination time. Thus, methoxy species are considered important intermediates while both formate and Pt-adsorbed CO seem to be involved in deactivation. Additionally, the gradually increasing negative band at 1450 cm^{-1} points to the consumption of carbonate as residue in commercial STO on the surface during the photocatalytic reaction. The missing reflexes of SrCO_3 in the XRD pattern of 0.1Pt/STO after PD (see Fig. S18, ESI[†]) indicate that only a minor amount of SrCO_3 is present not significantly influencing the photocatalytic experiments.⁵⁵

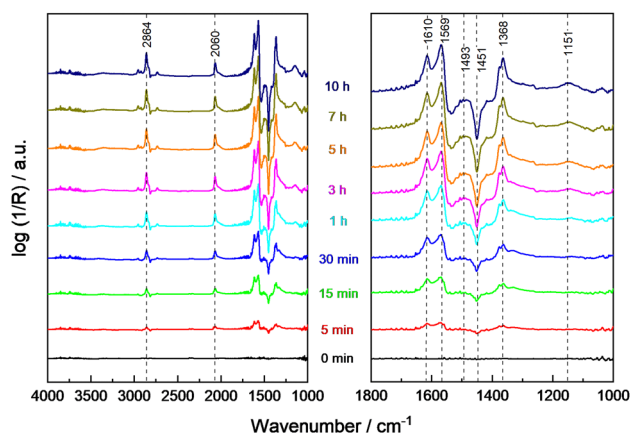


Fig. 4 DRIFT spectra during illumination of 0.1Pt/STO exposed to a constant flow of methanol (0.5%) and water (1.5%) in He (left). The spectral region $1800\text{--}1000\text{ cm}^{-1}$ is enlarged (right). As background spectra, the spectra before illumination at $100\text{ }^\circ\text{C}$ were chosen revealing the differences to the experiments before illumination.

To rationalize the strong dependence of the product selectivity and deactivation (compare Fig. 1 and Fig. S9, ESI[†]) on the reaction temperature, *in situ* methanol oxidation has been performed at temperatures up to $200\text{ }^\circ\text{C}$ (Fig. 5).⁹ In the absence of illumination, the amount of physisorbed water decreased as expected indicated by the lower intensity of the characteristic broad band at 3506 cm^{-1} , but otherwise no significant changes were observed with increasing temperature. Thus, even at $200\text{ }^\circ\text{C}$, no change of the bands from hydroxyl groups (3659 cm^{-1} and 3691 cm^{-1}) or of methoxy groups (2928 cm^{-1} , 2825 cm^{-1} , 1152 cm^{-1}) were observed indicating a high stability or their fast regeneration from gas-phase water and methanol. The band of CO adsorbed on Pt at $\sim 2056\text{ cm}^{-1}$ slightly increased with temperature which points to an enhanced thermal decomposition of methanol forming CO. However, due to the low Pt content (0.1 wt%), this band remains small.

During sample illumination, the methoxy bands (2928 cm^{-1} , 2825 cm^{-1} , 1152 cm^{-1}) were generally weaker pointing to a lower surface coverage in agreement with a fast thermal-assisted photocatalytic reaction. A fast desorption of methoxy groups due to the elevated temperatures as origin of the less intense bands is considered unlikely because of the negligible effect of temperature changes in the absence of illumination. Importantly, in all spectra rather intense formate bands were observed. Considering that formate bands at 2864 cm^{-1} , 1615 cm^{-1} and 1565 cm^{-1} were only obtained during illumination, formate formation is considered as a kinetically limited photocatalytic step. In the temperature range $<140\text{ }^\circ\text{C}$, the amount of methoxy groups was increasing with temperature, whereas the surface coverage of formate was constantly decreasing. This trend reversed at temperatures $\geq 140\text{ }^\circ\text{C}$ and formate groups appear to be of higher surface abundance. Importantly, no bands related to formaldehyde or methyl

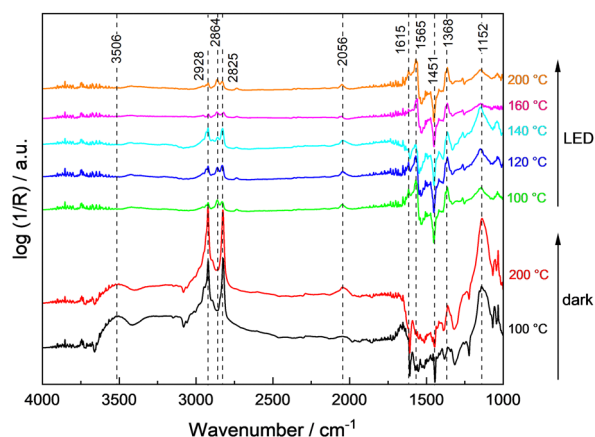


Fig. 5 DRIFT spectra of 0.1Pt/STO obtained in the temperature range from $100\text{ }^\circ\text{C}$ to $200\text{ }^\circ\text{C}$ with flowing methanol (0.5%) and water (1.5%) in the dark and under UV illumination. The reaction temperature was increased in 10 K steps (not all shown here) and kept constant for 90 min for each step; spectra were recorded at the end of each step. The temperature ramp was carried out three times (background recording, dark reaction, and photocatalytic reaction) with the same sample.



formate were detected during the *in situ* DRIFTS analysis suggesting their fast desorption.

In comparison to the data obtained in the continuously-operated flat-plate reactor (see Fig. 1) which revealed severe changes in catalyst activity, stability and selectivity at temperatures ≥ 120 °C, a mismatch in temperature of about 20 °C was obtained, likely caused by the indirect temperature measurements in DRIFTS studies.⁹ Still in general, the low methoxy surface coverage at high temperatures agrees well with the high methanol conversion. The change in selectivity to CO at the expense of formaldehyde and methyl formate formation is in line with the more intense CO band observed in the DRIFTS measurements, and the intensity increase of bands associated to formate suggest its contribution to deactivation of 0.1Pt/STO at higher reaction temperatures. Nevertheless, CO adsorbed on Pt sites cannot be fully excluded as deactivating species based on the DRIFTS measurements.

Reaction mechanism, deactivation and regeneration

Based on the DRIFTS results discussed above and the evolved gas-phase products, a mechanism for the thermal-assisted photocatalytic methanol oxidation in the gas phase is proposed, which is also supported by recent literature.^{4,9,11,13} Frequently, thermal and photocatalytic reaction steps were considered to be concurrent, with UV illumination accelerating specific reaction steps. However, the experimental evidence presented in this work suggests that certain reaction steps proceed exclusively under illumination. Due to the observed displacement of hydroxyl groups by methoxy groups, it is assumed that methanol and water adsorb competitively on the same sites,¹⁶ and that all oxidation reactions occur on STO as illustrated in Fig. 6.

The initial step of the proposed methanol oxidation mechanism (Fig. 6) is the thermal dissociative adsorption of methanol on the STO surface which is supported by the DRIFTS analysis discussed above (see Fig. 3 and Fig. S19, ESI†) (A). This step requires the presence of hydroxyl groups which are continuously regenerated by gas-phase water (B). Our DRIFTS results suggest that both molecular methanol and methoxy groups are present during the reaction (Fig. 3). However, Shen *et al.*¹⁵ provided evidence based on TPD experiments under UHV conditions that methoxy groups are oxidized much faster by photogenerated holes with formaldehyde being the primary product.^{13,15} Formaldehyde either desorbs or is further oxidized to a formyl intermediate (A). Although no spectroscopic evidence for the presence of formyl species was found here, its presence has been proposed before^{11,14} and, as highlighted below, agrees well with the observation of the evolution of formaldehyde and consecutive products (Fig. 1). Formyl species are converted in three concurrent consecutive reactions on the STO surface: oxidation to CO₂ *via* formate species as a relatively stable intermediate which is supported by the increasing formate bands during the photocatalytic reaction and requires the presence of adjacent hydroxyl groups (Fig. 4) (C); further dehydrogenation of formyl groups to CO (D); or cross coupling between formyl and methoxy species yielding methyl formate (E). Previously, it has been shown that with decreasing methanol and constant water partial pressure an increase in the selectivity to CO/CO₂ and a simultaneous decrease in methyl formate selectivity is obtained.⁹ A similar change of the product distribution was obtained when the reaction temperature in the flat-plate reactor was increased to > 125 °C.⁹ In agreement with the decreasing methoxy bands and stable hydroxy bands observed here at temperatures ≥ 140 °C (Fig. 5), a lower coverage of methoxy groups can be correlated with a lower selectivity

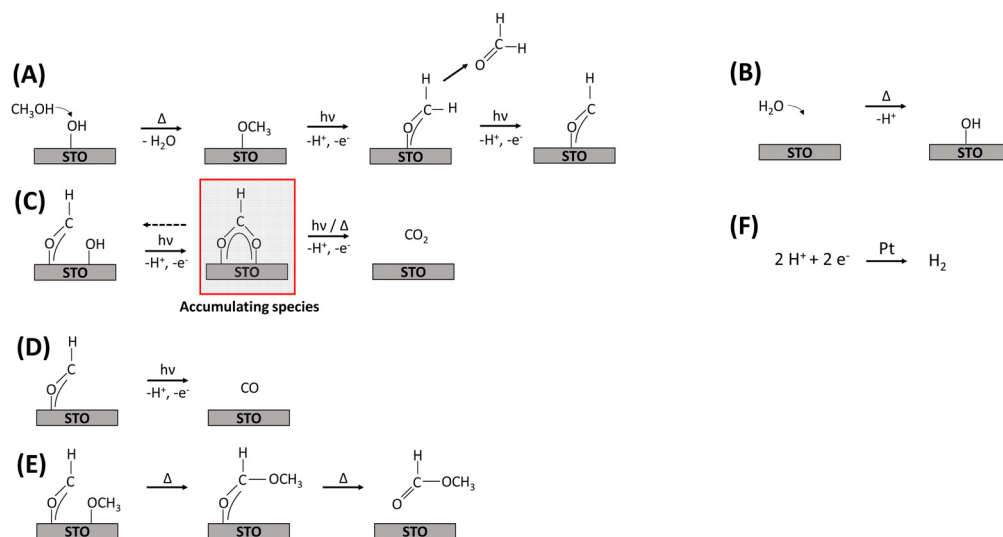
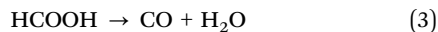
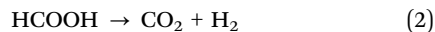


Fig. 6 Proposed reaction mechanism for the thermal-assisted photocatalytic methanol oxidation in the gas phase (steps A)–(F) including the main carbon-based reactants, intermediates, and products as well as the reduction of H⁺ at Pt NPs. The accumulated formate species was highlighted (see (C)) as it is proposed to be the main deactivating species. Note that, the regeneration of hydroxyl and methoxy groups must be considered to fulfil the hydrogen balance. Step F does not take into account the total amount of H⁺ produced by other steps. The trace compounds CH₄, dimethoxy methane and dimethyl ether were omitted.



to methyl formate supporting the mutual influence of reaction C, D and E depending on the surface coverage of methoxy and hydroxy groups. Additionally, the evolution of formic acid (*via* the formate intermediate) may occur even if it is not detected due to the easy decomposition of formic acid on noble metal-containing catalysts at elevated temperatures either to CO₂ and H₂ or to CO and water according to eqn (2) and (3).^{4,56,57} These steps were omitted in the mechanism (Fig. 6), as no experimental evidence for the presence of formic acid was found here.



Considering that the CO₂ evolution peaked after illumination was discontinued may stem from the thermal oxidation of accumulated formate species. Consequently, the formate oxidation to CO₂ during illumination may include both a thermal and a photocatalytic pathway. An equilibrium between formyl and formate species has been reported before,¹¹ which is important to explain the tailing of the methyl formate evolution rate after the photocatalytic reaction. In that case, a part of the formate groups may have been converted back to formyl species which reacted to methyl formate through a thermal mechanism. The thermal reaction of two methoxy groups to methyl formate is excluded here as methyl formate was not observed in the absence of illumination.

Considering the presented reaction mechanism and the DRIFTS data discussed above, changes in selectivity towards CO formation with increasing temperature are rationalized by a decrease in neighboring formyl and methoxy groups caused by the lower surface coverage of methoxy groups (Fig. 5). Moreover, dehydrogenation of the formyl groups to CO was favored over oxidation to CO₂ again due to the generally slow oxidation of formate to CO₂ being in agreement with formate accumulation at the surface.

H₂ evolution (step F) proceeds alongside methanol oxidation. For H₂ evolution, two fundamentally different approaches regarding the type of hydrogen surface species are currently discussed in literature. In analogy to electrochemical H₂ evolution, H⁺ released at the oxidation sites are subsequently reduced to H₂ by trapped photogenerated electrons at the metal cocatalyst. In contrast, the Heiz group suggested a fundamentally different hole-mediated disproportionation mechanism where the metal cocatalyst acts as recombination center for atomic hydrogen when deposited on a reducible semiconductor which enables diffusion of atomic hydrogen rather than acting as reduction site. Furthermore, the Heiz group suggested that photoexcited electrons are not trapped by cocatalyst particles but on the photocatalyst leading to negatively charged states which act as starting point of the catalytic cycle. The strict separation of half reactions, often made in photocatalysis, is thus not applicable in that case.^{19,58,59} However, we observed a strong dependence of the reaction kinetics on the presence of water being the H⁺ transport medium, and for the Pt loading an optimum Pt amount for every photocatalytic system was derived.⁹ Additionally, EPR measurements

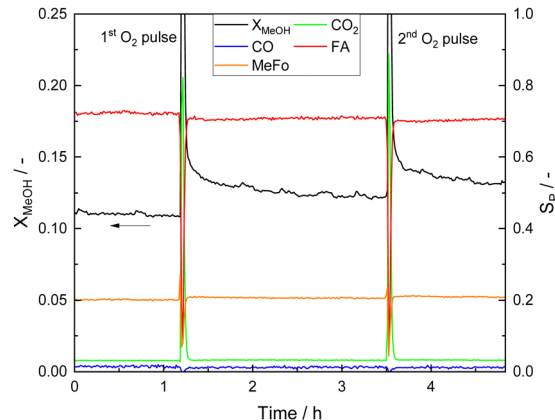


Fig. 7 X_{MeOH} profiles and product distribution before, after and during the first and second O₂ pulse. The O₂ pulses were performed after the quasi-steady state was reached at standard reaction conditions (Table 1) which corresponds to $t = 0$ h.

indicated that Pt NPs capture photogenerated electrons supporting the commonly accepted mechanism of H⁺ reduction at the surface of Pt NPs after surface migration *via* hydroxyl groups being favored in the presence of water.⁵⁷ Thus, H₂ formation is depicted here as reduction of H⁺ by trapped photogenerated electrons on the Pt NPs but overall the mechanistic sequence also allows to consider both oxidation and reduction half-reactions as interconnected and rationalizes the formation of the observed products while closing the carbon and hydrogen balances.

Photocatalyst deactivation appears to be strongly related to formate surface coverage on STO and CO on Pt sites and is shown to be reversible which was revealed using O₂ pulses (Fig. 7). Methanol conversion increased by pulsing O₂, and, as a consequence, all product evolution rates increased while the product distribution remained unchanged.

The coverage with formate groups increased with reaction temperature (see Fig. 5) and clear evidence for CO adsorbed on Pt was observed by DRIFTS. After a certain illumination time, the intensity of the corresponding CO bands was not further increasing suggesting a constant CO coverage under the applied reaction conditions. Estimation of the released amount of additional CO₂ produced during O₂ pulsing reveals that the amount of CO₂ (1.92×10^{-5} mol) exceeds the number of Pt atoms by one order of magnitude (1.54×10^{-6} mol, even without considering the surface-to-volume ratio of the Pt atoms in the NPs). Therefore, oxidation of adsorbed CO on Pt is not sufficiently explaining deactivation, being also in agreement with CO and CO₂ pulses that are not affecting photocatalytic performance (Fig. S20, ESI[†]). Furthermore, the low CO conversion during the WGS over 0.1Pt/STO under dark and illuminated conditions (Fig. S21, ESI[†]) indicates that a fast removal of CO from Pt as it was proposed before^{20,21} is unlikely under the conditions applied here. Other organic adsorbates on the STO surface must be the main source of CO₂ during the O₂ pulse and be responsible for catalyst deactivation. Thus, deactivation is primarily caused by adsorbed formate on STO with CO_{ads} on Pt being a spectator species present at high amounts on the



surface of the photocatalyst. The additional formation of polymeric moieties with high molecular weight from formaldehyde which might further deactivate the catalyst is not likely due to the elevated reaction temperatures,⁶⁰ and, correspondingly, no related bands were observed. Though, not evidenced by the spectroscopic data provided here, formate formation might also occur as consequence of the WGS on Pt nanoparticles and may, in addition to CO_{ads}, block Pt sites.^{61,62}

Obviously, the type and binding strength of the STO surface sites is a crucial factor. Therefore, controlling formaldehyde desorption, formate formation and its further oxidation to CO₂ through, among other factors, surface site engineering will likely contribute to lower catalyst deactivation and high formaldehyde selectivity even at elevated reaction temperatures.

Conclusions

This work presents a detailed *in situ* DRIFTS study to deduce the reaction mechanism of the thermal-assisted anaerobic gas-phase methanol photooxidation over Pt-modified SrTiO₃ providing insights into the deactivation mechanism. The primary oxidation product formaldehyde produced in conjunction with H₂ as coupled product agrees with the formation of formyl species. Formyl species are thus identified as key intermediate for three concurrent reactions resulting in CO₂, CO and methyl formate formation. Moreover, formation of methoxy groups from methanol and surface hydroxyl groups are found to occur under thermal conditions, while oxidation of methoxy groups to formaldehyde is exclusively driven by photocatalysis. Deactivation of the photocatalytic process occurs particularly at higher temperatures and is caused by the sluggish kinetics of the oxidation of stable formate groups to CO₂ and its corresponding accumulation at the surface of 0.1Pt/STO which is shown to be reversible by O₂ pulsing. Using EPR spectroscopy, participation of paramagnetic OV_s in the thermal-assisted photocatalytic reaction mechanism is excluded over undoped SrTiO₃. This study contributes to the development of an efficient procedure for catalyst regeneration and to enhancing the long-term stability of the photocatalytic reaction.

Author contributions

M. Deitermann: conceptualization, data curation, formal analysis, investigation, methodology, visualization, writing – original draft. T. Sato: data curation, formal analysis, investigation, methodology. Y. Haver: data curation, formal analysis, investigation. A. Schnegg: investigation, methodology, supervision. M. Muhler: conceptualization, funding acquisition, resources, supervision, writing – review & editing. B. Mei: conceptualization, supervision, writing – review & editing.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the International Max Planck Research School on Reactive Structure Analysis for Chemical Reactions (IMPRS-RECHARGE) and the “Center for Solvation Science ZEMOS” funded by the German Federal Ministry of Education and Research BMBF, German Research Foundation (DFG) under CRC/TRR 247 and by the Ministry of Culture and Research of North Rhine-Westphalia.

Notes and references

- 1 L. S. Al-Mazroai, M. Bowker, P. Davies, A. Dickinson, J. Greaves, D. James and L. Millard, *Catal. Today*, 2007, **122**, 46–50.
- 2 G. L. Chiarello, E. Selli and L. Forni, *Appl. Catal., B*, 2008, **84**, 332–339.
- 3 M. Bingham and A. Mills, *J. Photochem. Photobiol., A*, 2020, **389**, 112257.
- 4 A. Caravaca, H. Daly, M. Smith, A. Mills, S. Chansai and C. Hardacre, *React. Chem. Eng.*, 2016, **1**, 649–657.
- 5 G. L. Chiarello, L. Forni and E. Selli, *Catal. Today*, 2009, **144**, 69–74.
- 6 G. L. Chiarello, M. H. Aguirre and E. Selli, *J. Catal.*, 2010, **273**, 182–190.
- 7 L. M. Ahmed, I. Ivanova, F. H. Hussein and D. W. Bahnemann, *Int. J. Photoenergy*, 2014, **2014**, 1–9.
- 8 D. A. Armstrong, R. E. Huie, W. H. Koppenol, S. V. Lymar, G. Merényi, P. Neta, B. Ruscic, D. M. Stanbury, S. Steenken and P. Wardman, *Pure Appl. Chem.*, 2015, **87**, 1139–1150.
- 9 M. Deitermann, Y. Haver, B. T. Mei and M. Muhler, *Adv. Sustainable Syst.*, 2023, 2300329.
- 10 T. H. Chiang, H. Lyu, T. Hisatomi, Y. Goto, T. Takata, M. Katayama, T. Minegishi and K. Domen, *ACS Catal.*, 2018, **8**, 2782–2788.
- 11 M. El-Roz, P. Bazin, M. Daturi and F. Thibault-Starzyk, *Phys. Chem. Chem. Phys.*, 2015, **17**, 11277.
- 12 M. El-Roz, M. Kus, P. Cool and F. Thibault-Starzyk, *J. Phys. Chem. C*, 2012, **116**, 13252.
- 13 J. Schnee, M. Daturi and M. El-Roz, *Catal. Sci. Technol.*, 2020, **10**, 5618–5627.
- 14 K. R. Phillips, S. C. Jensen, M. Baron, S.-C. Li and C. M. Friend, *J. Am. Chem. Soc.*, 2013, **135**, 574–577.
- 15 M. Shen and M. A. Henderson, *J. Phys. Chem. Lett.*, 2011, **2**, 2707–2710.
- 16 G. L. Chiarello, D. Ferri and E. Selli, *Appl. Surf. Sci.*, 2018, **450**, 146–154.
- 17 G. M. Haselmann and D. Eder, *ACS Catal.*, 2017, **7**, 4668–4675.
- 18 G. M. Haselmann, B. Baumgartner, J. Wang, K. Wieland, T. Gupta, C. Herzig, A. Limbeck, B. Lendl and D. Eder, *ACS Catal.*, 2020, **10**, 2964–2977.
- 19 C. A. Walenta, C. Courtois, S. L. Kollmannsberger, M. Eder, M. Tschurl and U. Heiz, *ACS Catal.*, 2020, **10**, 4080–4091.
- 20 M. Bowker, D. James, P. Stone, R. Bennett, N. Perkins, L. Millard, J. Greaves and A. Dickinson, *J. Catal.*, 2003, **217**, 427–433.



- 21 A. Dickinson, D. James, N. Perkins, T. Cassidy and M. Bowker, *J. Mol. Catal. A: Chem.*, 1999, **146**, 211–221.
- 22 F. Platero, A. López-Martín, A. Caballero and G. Colón, *ChemCatChem*, 2021, **13**, 3878–3888.
- 23 J. C. Matsubu, S. Zhang, L. DeRita, N. S. Marinkovic, J. G. Chen, G. W. Graham, X. Pan and P. Christopher, *Nat. Chem.*, 2017, **9**, 120.
- 24 A. J. Martín, S. Mitchell, C. Mondelli, S. Jaydev and J. Pérez-Ramírez, *Nat. Catal.*, 2022, **5**, 854.
- 25 S. Peng, C. Gan, Y. Yang, S. Ji and Y. Li, *Energy Technol.*, 2018, **6**, 2166–2171.
- 26 Y. Fan, Y. Liu, H. Cui, W. Wang, Q. Shang, X. Shi, G. Cui and B. Tang, *Nanomaterials*, 2020, **10**, 2572.
- 27 H. Tan, Z. Zhao, W.-B. Zhu, E. N. Coker, B. Li, M. Zheng, W. Yu, H. Fan and Z. Sun, *ACS Appl. Mater. Interfaces*, 2014, **6**, 19184–19190.
- 28 Y. Kim, M. Watanabe, A. Takagaki, J. Matsuda and T. Ishihara, *ChemCatChem*, 2019, **11**, 6270–6274.
- 29 B. Mei, K. Han and G. Mul, *ACS Catal.*, 2018, **8**, 9154.
- 30 B. Mei, C. Wiktor, S. Turner, A. Pougin, G. van Tendeloo, R. A. Fischer, M. Muhler and J. Strunk, *ACS Catal.*, 2013, **3**, 3041–3049.
- 31 G. W. Busser, B. Mei and M. Muhler, *ChemSusChem*, 2012, **5**, 2200–2206.
- 32 S. Stoll and A. Schweiger, *J. Magn. Reson.*, 2006, **178**, 42–55.
- 33 K. Han, D. M. Haiber, J. Knöppel, C. Lievens, S. Cherevko, P. Crozier, G. Mul and B. Mei, *ACS Catal.*, 2021, **11**, 11049–11058.
- 34 G. Zhao, G. W. Busser, C. Froese, B. Hu, S. A. Bonke, A. Schnegg, Y. Ai, D. Wei, X. Wang, B. Peng and M. Muhler, *J. Phys. Chem. Lett.*, 2019, **10**, 2075–2080.
- 35 Y. Hu, G. Zhao, Q. Pan, H. Wang, Z. Shen, B. Peng, G. W. Busser, X. Wang and M. Muhler, *ChemCatChem*, 2019, **11**, 5139–5144.
- 36 T. Rath, M. Deitermann, G. Zhao, G. Wilma Busser, H. Jansen, P. Schwiderowski, W. Xia and M. Muhler, *ChemPhotoChem*, 2021, **5**, 521–525.
- 37 V. N. Belevskii, I. Volnov and S. A. Tokareva, *Consultants Bureau*, 1971, 1366–1367.
- 38 S. M. Prokes, J. L. Gole, X. Chen, C. Burda and W. E. Carlos, *Adv. Funct. Mater.*, 2005, **15**, 161–167.
- 39 M. Anpo, M. Che, B. Fubini, E. Garrone, E. Giamello and M. C. Paganini, *Top. Catal.*, 1999, 189–198.
- 40 R. F. Howe and M. Gratzel, *J. Phys. Chem.*, 1985, **89**, 4495–4499.
- 41 X. Zhou, N. Liu, T. Yokosawa, A. Osvet, M. E. Miehllich, K. Meyer, E. Spiecker and P. Schmuki, *ACS Appl. Mater. Interfaces*, 2018, **10**, 29532.
- 42 S. Livraghi, M. Chiesa, M. C. Paganini and E. Giamello, *J. Phys. Chem. C*, 2011, **115**, 25413.
- 43 K. Lau, F. Niemann, K. Abdiaziz, M. Heidelmann, Y. Yang, Y. Tong, M. Fechtelkord, T. C. Schmidt, A. Schnegg, R. K. Campen, B. Peng, M. Muhler, S. Reichenberger and S. Barcikowski, *Angew. Chem., Int. Ed.*, 2023, **62**, e202213968.
- 44 R. A. Maier, A. C. Johnston-Peck and M. P. Donohue, *Adv. Funct. Mater.*, 2016, **26**, 8325.
- 45 K. I. Hadjiivanov, *Adv. Catal.*, 2014, **57**, 99–318.
- 46 K. I. Hadjiivanov, *J. Chem. Soc., Faraday Trans.*, 1998, **94**, 1901–1904.
- 47 B. L. Mojet, S. D. Ebbesen and L. Lefferts, *Chem. Soc. Rev.*, 2010, **39**, 4643–4655.
- 48 U.S. Secretary of Commerce, “NIST Chemistry Webbook”, 2023, <https://webbook.nist.gov/chemistry/>. last accessed 22.12.23.
- 49 F. Boccuzzi, A. Chiorino and M. Manzoli, *J. Power Sources*, 2003, **118**, 304–310.
- 50 A. Nuhu, J. Soares, M. Gonzalez-Herrera, A. Watts, G. Hussein and M. Bowker, *Top. Catal.*, 2007, **44**, 293–297.
- 51 N. Tipcompor, T. Thongtem, A. Phuruangrat and S. Thongtem, *Mater. Lett.*, 2012, **87**, 153–156.
- 52 N. Aas, T. J. Pringle and M. Bowker, *J. Chem. Soc., Faraday Trans.*, 1994, **90**, 1015–1022.
- 53 L.-Q. Wang, K. F. Ferris, S. Azad and M. H. Engelhard, *J. Phys. Chem. B*, 2005, **109**, 4507–4513.
- 54 R. Huang, V. Fung, Y. Zhang, D. R. Mullins, Z. Wu and D.-E. Jiang, *J. Phys. Chem. C*, 2018, **122**, 7210–7216.
- 55 B. Boga, N. Steinfeldt, N. G. Moustakas, T. Peppel, H. Lund, J. Rabeah, Z. Pap, V.-M. Cristea and J. Strunk, *Catal*, 2022, **12**, 978.
- 56 H. Bahruji, M. Bowker, P. R. Davies and F. Pedrono, *Appl. Catal., B*, 2011, **107**, 205.
- 57 G. L. Chiarello, D. Ferri and E. Selli, *J. Catal.*, 2011, **280**, 168–177.
- 58 C. A. Walenta, S. L. Kollmannsberger, C. Courtois, R. N. Pereira, M. Stutzmann, M. Tschurl and U. Heiz, *Phys. Chem. Chem. Phys.*, 2019, **21**, 1491–1496.
- 59 M. Eder, M. Tschurl and U. Heiz, *J. Phys. Chem. Lett.*, 2023, **14**, 6193–6201.
- 60 W. M. D. Bryant and J. B. Thompson, *J. Polym. Sci., Part A-1: Polym. Chem.*, 1971, **9**, 2523.
- 61 J. Zhao, Y. Bai, Z. Li, J. Liu, W. Wang, P. Wang, B. Yang, R. Shi, G. I. N. Waterhouse, X.-D. Wen, Q. Dai and T. Zhang, *Angew. Chem., Int. Ed.*, 2023, **62**, e202219299.
- 62 A. A. Gokhale, J. A. Dumesic and M. Mavrikakis, *J. Am. Chem. Soc.*, 2008, **130**, 1402–1414.

