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## A highly active Pt/In<sub>2</sub>O<sub>3</sub> catalyst for CO<sub>2</sub> hydrogenation to methanol with enhanced stability†

Kaihang Sun,<sup>a,b</sup> Ning Rui,<sup>a,b</sup> Zhitao Zhang,<sup>a,b</sup> Zeyu Sun,<sup>a,b</sup> Qingfeng Ge<sup>c</sup> and Chang-Jun Liu  <sup>a,b</sup>

Supported Pt catalysts have been extensively investigated for CO<sub>2</sub> hydrogenation with methane and CO as the principal products. In this work, a Pt/In<sub>2</sub>O<sub>3</sub> catalyst prepared with decomposition precipitation was tested for CO<sub>2</sub> hydrogenation to methanol. The Pt/In<sub>2</sub>O<sub>3</sub> catalyst exhibited a highly improved activity towards CO<sub>2</sub> hydrogenation, with methanol selectivity of ca. 100% at temperatures below 225 °C, 74% at 275 °C and 54% at 300 °C, respectively, compared to the pure In<sub>2</sub>O<sub>3</sub> catalyst. This represents the highest methanol selectivity reported on Pt catalysts for CO<sub>2</sub> hydrogenation. The stability of the Pt/In<sub>2</sub>O<sub>3</sub> catalyst at elevated temperatures has also been higher than that of the pure In<sub>2</sub>O<sub>3</sub> catalyst, indicated by the methanol formation rate decreasing only to 95% of the initial rate after 5 h in the reaction stream and remaining largely constant thereafter. In contrast, the pure In<sub>2</sub>O<sub>3</sub> catalyst loses 20% of the initial methanol formation rate after 9 h in the reaction stream. The characterization of the catalysts confirms that the Pt nanoparticles are well dispersed on In<sub>2</sub>O<sub>3</sub> with a particle size below 3 nm. The strong metal–support interaction (SMSI) between Pt and In<sub>2</sub>O<sub>3</sub> improves the stability of the catalyst and prevents the over-reduction of In<sub>2</sub>O<sub>3</sub>. The synergy between the supported Pt nanoparticles and In<sub>2</sub>O<sub>3</sub> balances the hydrogen activation and the density of the surface oxygen vacancies in In<sub>2</sub>O<sub>3</sub>, resulting in the high activity for CO<sub>2</sub> hydrogenation and enhanced stability of the Pt/In<sub>2</sub>O<sub>3</sub> catalyst.

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## Introduction

With the development of renewable energy, CO<sub>2</sub> capture and utilization<sup>1–4</sup> have the potential to alleviate the elevated CO<sub>2</sub> level. Hydrogenation of CO<sub>2</sub> to methanol<sup>5</sup> has been considered as key to the so-called “methanol economy”.<sup>6</sup> Meanwhile, this process can also be considered as a means of hydrogen storage and/or energy storage.<sup>7</sup> Extensive and continual efforts have been made over the past decades to develop a highly active catalyst with high methanol selectivity. The investigated catalysts include those based on Cu/ZnO,<sup>8–10</sup> ZnO–ZrO<sub>2</sub>,<sup>11</sup> GaNi alloy,<sup>12</sup> bimetallic PdZn<sup>13</sup> and others. Our group predicted *via* density functional theoretical (DFT) studies, and then confirmed experimentally, that In<sub>2</sub>O<sub>3</sub> with oxygen vacancies exhibited a high activity for CO<sub>2</sub> activation and further high selectivity for hydrogenation to methanol.<sup>14–18</sup> Since then,

In<sub>2</sub>O<sub>3</sub>-based catalysts have attracted increasing attention as novel catalysts for CO<sub>2</sub> hydrogenation to methanol. Further studies include novel preparation and characterization of In<sub>2</sub>O<sub>3</sub>,<sup>19–21</sup> combination of In<sub>2</sub>O<sub>3</sub> with other oxides, such as ZrO<sub>2</sub>,<sup>22–25</sup> and the addition of metals, including copper,<sup>26,27</sup> cobalt,<sup>28</sup> palladium,<sup>18,29</sup> rhodium,<sup>30</sup> and nickel,<sup>31</sup> to further improve the catalytic activity for CO<sub>2</sub> activation and hydrogenation. It has also been demonstrated that the strong hydrogenation ability of the metal catalysts enhances the hydrogenation of the In<sub>2</sub>O<sub>3</sub>-supported metal catalysts.<sup>18,26–33</sup> Among the metals considered, platinum is an important one for its extensive applications in hydrogenation and photo-catalytic conversion.<sup>33</sup> Although a Pt/In<sub>2</sub>O<sub>3</sub> film catalyst has been shown to exhibit an enhanced activity for CO<sub>2</sub> hydrogenation to methanol, the reaction was performed under atmospheric pressure plasma conditions.<sup>32</sup> The In<sub>2</sub>O<sub>3</sub>-supported Pt catalyst has not been examined for thermal catalytic methanol synthesis from CO<sub>2</sub> and hydrogen.

The supported Pt catalysts normally exhibit a high activity for CO<sub>2</sub> hydrogenation to methane and CO. The metal dispersion,<sup>34–36</sup> the acidity of the support,<sup>35</sup> the effect of the promoter<sup>37</sup> and oxygen vacancies,<sup>36</sup> the metal–support interaction,<sup>34,35,38,39</sup> and the degree of Pt reduction<sup>40,41</sup> have been shown to affect the activity and selectivity. CO<sub>2</sub> hydrogenation to methanol has been investigated as well.<sup>42,43</sup> But the

<sup>a</sup>Tianjin Co-Innovation Center of Chemical Science & Engineering, Tianjin University, Tianjin 300072, China. E-mail: ughg\_cjl@yahoo.com

<sup>b</sup>School of Chemical Engineering and Technology, Tianjin University, Tianjin 300072, China

<sup>c</sup>Department of Chemistry and Biochemistry, Southern Illinois University, Carbondale, Illinois 62901, USA

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most selective products remain to be methane and CO over the Pt catalysts.<sup>34–36</sup> A Pt catalyst with high methanol selectivity from CO<sub>2</sub> hydrogenation is desired.

Although In<sub>2</sub>O<sub>3</sub>-based catalysts are active and selective for methanol synthesis from CO<sub>2</sub> and H<sub>2</sub>, the over-reduction of In<sub>2</sub>O<sub>3</sub> under the reaction conditions causes deactivation and a decrease in selectivity because In<sub>2</sub>O<sub>3</sub> can be reduced readily in the presence of hydrogen and CO.<sup>19,44,45</sup> Over-reduced In<sub>2</sub>O<sub>3</sub> catalyses preferably the competing reverse water gas shift (RWGS) reaction, resulting in CO instead of methanol, especially at elevated reaction temperatures. CO is more potent at reducing In<sub>2</sub>O<sub>3</sub> than H<sub>2</sub>. Its production during the reaction exacerbates the problem, leading to further reduction of the activity and selectivity.<sup>19</sup>

In this work, we prepared the Pt/In<sub>2</sub>O<sub>3</sub> catalysts using the decomposition precipitation method. The resulting catalyst shows an improved activity over In<sub>2</sub>O<sub>3</sub> alone. High methanol selectivity was achieved on the Pt/In<sub>2</sub>O<sub>3</sub> catalyst with a methanol selectivity of ca. 100% at temperatures below 225 °C. The methanol selectivity was maintained at 74% at 275 °C and 54% at 300 °C. Inhibition of the over-reduction of In<sub>2</sub>O<sub>3</sub> was achieved due to the strong metal–support interaction (SMSI) between Pt and In<sub>2</sub>O<sub>3</sub>, resulting in the enhanced stability.

## Experimental

### Catalyst preparation

In<sub>2</sub>O<sub>3</sub> was prepared using the precipitation method. In (NO<sub>3</sub>)<sub>3</sub>·4H<sub>2</sub>O (HWRK Chem, 99.99%) and Na<sub>2</sub>CO<sub>3</sub>·10H<sub>2</sub>O (Tianjin Kermel Chemical Reagent, 99%) are the precursor and precipitant, respectively. Firstly, a desired amount of the precursor was dissolved in 100 mL of deionized water (0.15 mol L<sup>-1</sup>), followed by vigorous stirring at 70 °C until fully dissolved. Then 0.2 mol L<sup>-1</sup> of Na<sub>2</sub>CO<sub>3</sub> was added to the solution dropwise at the same temperature until the pH value reached 7. The mixture was aged at 80 °C for 2 h after being stirred for an additional 0.5 h. The precipitate was then washed with deionized water several times. The resulting solid was dried at 80 °C for 12 h prior to calcination in static air at 450 °C for 3 h.

The Pt/In<sub>2</sub>O<sub>3</sub> catalyst was prepared using the deposition–precipitation method. Platinum nitrate hydrate (Aladdin Chemicals, Shanghai, 18.02 wt% of metal platinum basis) was first dissolved in 50 mL of deionized water. The as-prepared In<sub>2</sub>O<sub>3</sub> was then added into the solution, and the resulting mixture was stirred vigorously for 1 h at room temperature. This was followed by adding an excess amount of urea (0.2 g) into the mixture under continuous stirring at 80 °C for 3 h while the precipitate formed. The precipitate was washed and filtered with 500 mL of deionized water. Prior to calcination in static air at 450 °C for 3 h, the precipitate was freeze-dried overnight.

### Characterization

The catalyst samples were characterized using X-Ray fluorescence (XRF) spectroscopy on a JEOL Supermini200 system,

equipped with a palladium X-ray source (Pd anode, 50 kV and 200 W). 77 K N<sub>2</sub> adsorption/desorption isotherms of the samples were measured on an AUTOSORB-1-C instrument (Quantachrome). The specific surface area (*S*<sub>BET</sub>) was calculated using the Brunauer–Emmett–Teller (BET) model.

The crystal structure of the catalyst was determined by powder X-ray diffraction (PXRD), using a Rigaku D/max 2500v/pc diffractometer with Cu Kα radiation (40 kV, 200 mA) at a scanning rate of 1° min<sup>-1</sup> within the 2θ range of 10°–90°. The phase identification was performed by comparing with the Joint Committee on Powder Diffraction Standards (JCPDS).

Transmission electron microscopy (TEM) images were collected on a JEOL JEM-2100F system equipped with an energy-dispersive X-ray spectrometer (EDX) operated at 200 kV. The sample powder was suspended in ethanol and dispersed ultrasonically for 10 min. A drop of the suspension was deposited on a copper grid coated with carbon.

Temperature-programmed reduction of hydrogen (H<sub>2</sub>-TPR) was conducted on a Micromeritics Autochem II 2920 chemisorption analyzer. About 200 mg of the sample was placed into a U-shaped quartz tube. The tube with the sample was purged by flowing helium for 1 h at 200 °C to remove the absorbed impurities. After cooling to –50 °C with liquid nitrogen under flowing helium, the sample was heated to 350 °C at a rate of 10 °C min<sup>-1</sup> under a gaseous mixture of 10% H<sub>2</sub> in N<sub>2</sub> as the reductant. The effluent gas was analysed with a TCD. The TCD signal was normalized to the weight of the sample for subsequent analysis.

Temperature-programmed desorption of carbon dioxide (CO<sub>2</sub>-TPD) was conducted on the Micromeritics Autochem II 2920 chemisorption analyzer equipped with a Hiden HPR-20 mass spectrometer (MS). About 100 mg of the sample was placed in a U-shaped quartz tube and reduced with 10% H<sub>2</sub>/Ar for 1 h at 200 °C and then cooled to 50 °C under flowing helium, followed by CO<sub>2</sub> adsorption at the same temperature for 1 h. The tube with the sample was purged with flowing helium for 1 h to remove physically adsorbed CO<sub>2</sub>. It was then heated to 500 °C at a rate of 10 °C min<sup>-1</sup> under flowing helium. The effluent gas was analysed using the TCD and mass spectrometer (the signal of *m/z* = 44 was collected from the TCD outlet). The signal was also normalized to the weight of the sample for subsequent analysis.

Raman spectra were acquired using an inVia Reflex Renishaw Raman Spectroscopy System in the range of 200–800 nm, with a 532 nm laser as the excitation source. The laser power was set at 7 mW and the integration time was 5 s. The Raman signals were normalized to the weight of the sample for the subsequent analysis.

The UV–vis absorption spectra of the samples were recorded with a UV-2600 UV–vis spectrophotometer (Shimadzu Corporation) to investigate the formation of surface oxygen vacancies.

### Catalytic activity test

The catalytic activity for CO<sub>2</sub> hydrogenation to methanol over the Pt/In<sub>2</sub>O<sub>3</sub> catalyst and the pure In<sub>2</sub>O<sub>3</sub> catalyst was tested in a

vertical fixed bed reactor. 0.2 g of catalyst was diluted with 1.0 g of SiC before being loaded into the reactor. Prior to the test, the reactor with the catalyst was purged with N<sub>2</sub> (30 ml min<sup>-1</sup>) for 0.5 h at room temperature. The Pt/In<sub>2</sub>O<sub>3</sub> catalyst was pre-reduced under flowing hydrogen (H<sub>2</sub>/N<sub>2</sub> = 1/9, molar ratio, 30 ml min<sup>-1</sup>) at 200 °C for 1 h. The reactant mixture (H<sub>2</sub>/CO<sub>2</sub>/N<sub>2</sub> = 76/19/5, molar ratio) was then introduced into the reactor until the pressure reached 5 MPa. The catalytic activity was measured from 200 °C to 300 °C at a constant flow rate of 70 ml min<sup>-1</sup>. The gaseous hourly space velocity (GHSV) was adjusted to 21 000 cm<sup>3</sup> h<sup>-1</sup> g<sub>cat</sub><sup>-1</sup>. The effluent was analysed with an online gas chromatograph (Agilent 7890A) equipped with two columns, one connected to a flame ionized detector (FID) and the other to the thermal conductivity detector (TCD). The valves and lines between the reactor outlet and the GC inlet were maintained at 110 °C to prevent the condensation of methanol.

The CO<sub>2</sub> conversion ( $X_{\text{CO}_2}$ ), methanol selectivity ( $S_{\text{methanol}}$ ), and methanol formation rate were calculated according to the following equations:

$$X_{\text{CO}_2} = \frac{F_{\text{CO}_2,\text{in}} - F_{\text{CO}_2,\text{out}}}{F_{\text{CO}_2,\text{in}}} \times 100\%,$$

$$S_{\text{methanol}} = \frac{F_{\text{methanol,out}}}{F_{\text{CO}_2,\text{in}} - F_{\text{CO}_2,\text{out}}} \times 100\%,$$

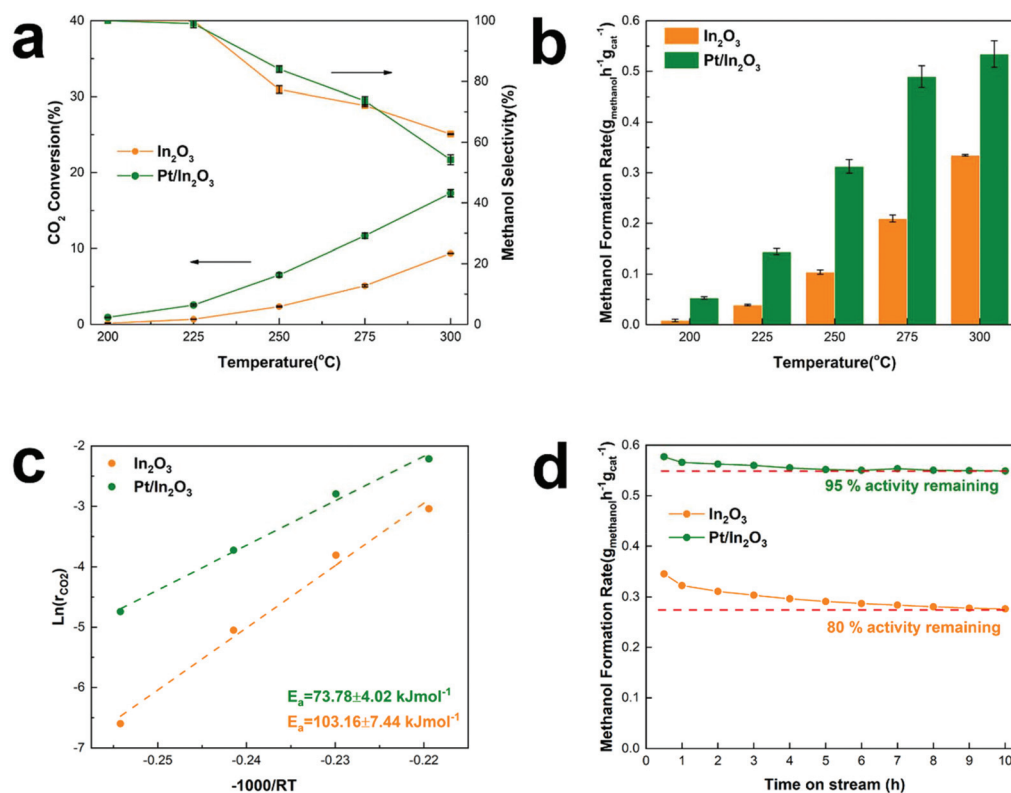
$$\text{Methanol formation rate} = \frac{F_{\text{CO}_2,\text{in}} \times X_{\text{CO}_2} \times S_{\text{methanol}}}{W} \times M,$$

where  $F$  is the molar flow rate,  $M$  the molar mass of methanol and  $W$  the weight of the catalyst.

## Results and discussion

### Catalytic activity tests

Fig. 1a shows CO<sub>2</sub> conversion and methanol selectivity on the pure In<sub>2</sub>O<sub>3</sub> and Pt/In<sub>2</sub>O<sub>3</sub> catalysts. Obviously, the conversion of CO<sub>2</sub> increases when the reaction temperature is increased. The methanol selectivity is close to 100% at reaction temperatures below 225 °C. It decreases when the temperature is increased further as the competing RWGS reaction is endothermic. On the other hand, the selectivity toward methanol is still higher than 70% at 275 °C and 50% at 300 °C. These results indicate that the Pt/In<sub>2</sub>O<sub>3</sub> catalyst remains highly selective for methanol synthesis from CO<sub>2</sub> hydrogenation. Furthermore, as shown in Fig. 1b, the Pt/In<sub>2</sub>O<sub>3</sub> catalyst shows a significantly higher methanol formation rate than the pure In<sub>2</sub>O<sub>3</sub> catalyst at the same temperature. The CO<sub>2</sub> conversion reaches 17.3% at 300 °C for the Pt/In<sub>2</sub>O<sub>3</sub> catalyst with a methanol formation rate of 0.542 g<sub>methanol</sub> h<sup>-1</sup> g<sub>cat</sub><sup>-1</sup>, significantly higher than 0.335 g<sub>methanol</sub> h<sup>-1</sup> g<sub>cat</sub><sup>-1</sup> on the pure In<sub>2</sub>O<sub>3</sub> catalyst with a CO<sub>2</sub> conversion of 9.4%. These results clearly demonstrate that the



**Fig. 1** Catalytic activity tests of the Pt/In<sub>2</sub>O<sub>3</sub> and pure In<sub>2</sub>O<sub>3</sub> catalysts. (a) CO<sub>2</sub> conversion and methanol selectivity, (b) methanol formation rate as a function of reaction temperature, (c) apparent activation energy of CO<sub>2</sub> conversion and (d) methanol formation rate versus time of the catalyst in the reaction stream (TOS).

Pt loading (1.07 wt% according to the results of XRF) on  $\text{In}_2\text{O}_3$  significantly increases the  $\text{CO}_2$  conversion and the methanol formation rate. The comparison of typically reported catalysts for  $\text{CO}_2$  hydrogenation to methanol is shown in Table S1.† The Pt/ $\text{In}_2\text{O}_3$  catalyst shows a higher activity than most of the reported catalysts. Compared to the supported Ni and Rh catalysts in Table S1,† the Pt/ $\text{In}_2\text{O}_3$  catalyst shows better  $\text{CO}_2$  conversion with higher selectivity towards methanol at lower temperatures of 250 °C and 275 °C.

The methanol selectivity of the Pt/ $\text{In}_2\text{O}_3$  catalyst at 300 °C is slightly lower than that of the pure  $\text{In}_2\text{O}_3$  catalyst at the same temperature, due largely to the agglomeration of the Pt nanoparticles at elevated temperatures. The average size of the Pt nanoparticles increases from 1.47 nm to 2.42 nm after reaction at 300 °C and 5 MPa. The large-sized Pt nanoparticles facilitate RWGS reaction according to the literature.<sup>29</sup> We speculate that hydrogenation of  $\text{CO}_2$  over Pt nanoparticles prefers to undergo the RWGS + CO-hydrogenation pathway. Weakening the binding of CO on the Pt nanoparticles leads to higher selectivity to CO.<sup>38</sup> Thus, the methanol selectivity of the Pt/ $\text{In}_2\text{O}_3$  catalyst is lower than that over the pure  $\text{In}_2\text{O}_3$  catalyst at 300 °C.

The apparent activation energy of the pure  $\text{In}_2\text{O}_3$  and the Pt/ $\text{In}_2\text{O}_3$  catalysts was calculated based on the Arrhenius equation at a  $\text{CO}_2$  conversion of less than 15% to minimize the effect of mass and heat transfer. As shown in Fig. 1c, the activation energy of  $\text{CO}_2$  conversion over the Pt/ $\text{In}_2\text{O}_3$  catalyst is 73.78  $\text{kJ mol}^{-1}$ , much lower than that over the pure  $\text{In}_2\text{O}_3$  catalyst (103.16  $\text{kJ mol}^{-1}$ ). This indicates that the Pt loading on  $\text{In}_2\text{O}_3$  significantly reduces the activation barrier for  $\text{CO}_2$  conversion. The apparent activation energy for methanol and CO production was calculated and is shown in Fig. S1.† For methanol production, the activation energies on the Pt/ $\text{In}_2\text{O}_3$  and pure  $\text{In}_2\text{O}_3$  catalysts are 64.08  $\text{kJ mol}^{-1}$  and 91.83  $\text{kJ mol}^{-1}$ , respectively. This indicates that the supported Pt nanoparticles reduce the activation energy of methanol production

significantly from that on the pure  $\text{In}_2\text{O}_3$  catalyst. In contrast, the activation energies for CO production are 102.18  $\text{kJ mol}^{-1}$  and 96.48  $\text{kJ mol}^{-1}$  on the Pt/ $\text{In}_2\text{O}_3$  and pure  $\text{In}_2\text{O}_3$  catalysts, respectively. These results demonstrate that supported Pt over  $\text{In}_2\text{O}_3$  promotes methanol production, while it suppresses the RWGS reaction.

Tsoukalou *et al.*<sup>19</sup> reported that  $\text{In}_2\text{O}_3$  was not stable over the course of the  $\text{CO}_2$  hydrogenation reaction and could not maintain its initial activity with a 200 min test. They concluded that the formation of  $\text{In}^0$  as a result of partial reduction of  $\text{In}_2\text{O}_3$  caused deactivation,<sup>19</sup> which typically occurs at elevated temperatures. The present results show that the stability of the pure  $\text{In}_2\text{O}_3$  catalyst can be significantly improved in the presence of well-dispersed Pt nanoparticles. As shown in Fig. 1d, the methanol formation rate on the Pt/ $\text{In}_2\text{O}_3$  catalyst only decreased slightly from its initial value and was stabilized at 95% of the initial rate after a 5 h reaction at 300 °C and 5 MPa. In contrast, the pure  $\text{In}_2\text{O}_3$  catalyst needs 9 h to become stable under the same reaction conditions with a much more significant reduction (20%) in the methanol formation rate.

### X-ray diffraction and TEM analysis

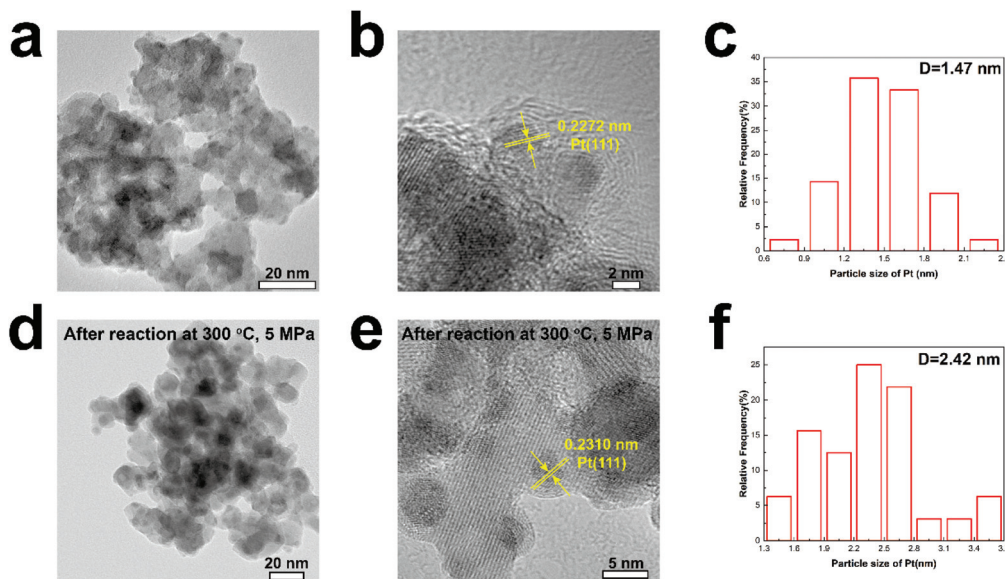
Based on the results of  $\text{N}_2$  adsorption, the specific surface areas of the Pt/ $\text{In}_2\text{O}_3$  and pure  $\text{In}_2\text{O}_3$  catalysts were determined to be 63.8  $\text{m}^2 \text{g}^{-1}$  and 65.8  $\text{m}^2 \text{g}^{-1}$ , respectively. The XRD patterns of the Pt/ $\text{In}_2\text{O}_3$  and pure  $\text{In}_2\text{O}_3$  catalysts are shown in Fig. 2. The Pt/ $\text{In}_2\text{O}_3$  catalyst before hydrogen reduction was named “Pt/ $\text{In}_2\text{O}_3$ -BT”. The samples after reaction at 300 °C and 5 MPa were labelled as “Pt/ $\text{In}_2\text{O}_3$ -AR” and “ $\text{In}_2\text{O}_3$ -AR”. The diffraction peaks at 21.5°, 30.7°, 35.5°, 45.7°, 51.0° and 60.7° were assigned to the diffraction from the (211), (222), (400), (431), (440) and (622) planes of the  $Ia\bar{3}$  (206) space group according to PDF#06-0416. No characteristic diffraction peaks of the Pt species could be observed in the XRD patterns, indicating that Pt is highly dispersed on  $\text{In}_2\text{O}_3$  with sizes below 3 nm. Fig. 2 also indicates that the crystal structure of the Pt/ $\text{In}_2\text{O}_3$  catalyst remains intact after hydrogen reduction and reaction.

Figs. 3a and b show the TEM images of the Pt/ $\text{In}_2\text{O}_3$  catalyst. As shown in Fig. 3c, the well-dispersed Pt nanoparticles, with an average diameter of 1.47 nm, on  $\text{In}_2\text{O}_3$  can be clearly observed. The lattice spacing of the Pt particle on  $\text{In}_2\text{O}_3$  is 0.2272 nm, which can be attributed to the (111) planes. As shown in Figs. 3d and e, the Pt nanoparticles exposing the (111) facet with a lattice spacing of 0.2310 nm can also be observed after reaction at 300 °C and 5 MPa. The increase can be attributed to the diffusion of Pt nanoparticles towards the bulk region of  $\text{In}_2\text{O}_3$ .<sup>29</sup> The average diameter increases to 2.42 nm but is still below 3 nm as shown in Fig. 3f. The samples annealed at high temperature (>700 K) appear to undergo a classical SMSI transition, as has been observed for Pt and  $\text{TiO}_2$ .<sup>46</sup> Therefore, a thin  $\text{In}_2\text{O}_3$  overlayer covering the Pt nanoparticles indicates that the SMSI between Pt and  $\text{In}_2\text{O}_3$  exists before and after the reaction.<sup>47–49</sup>

Figs. 4a and b show the TEM images of the pure  $\text{In}_2\text{O}_3$  catalyst before the reaction and Figs. 4c and d show those after the



Fig. 2 XRD patterns of the Pt/ $\text{In}_2\text{O}_3$  and pure  $\text{In}_2\text{O}_3$  catalysts.



**Fig. 3** (a) and (b) TEM images of the Pt/In<sub>2</sub>O<sub>3</sub> catalyst, (c) particle size distribution for the Pt/In<sub>2</sub>O<sub>3</sub> catalyst, (d) and (e) TEM images of Pt/In<sub>2</sub>O<sub>3</sub> after reaction, and (f) particle size distribution for Pt/In<sub>2</sub>O<sub>3</sub> after reaction.

reaction. The In<sub>2</sub>O<sub>3</sub> particles can be clearly identified with lattice fringes attributable to the (222) and (211) planes, suggesting that the catalyst exists in the form of In<sub>2</sub>O<sub>3</sub> and retains good crystallinity, consistent with the XRD results.

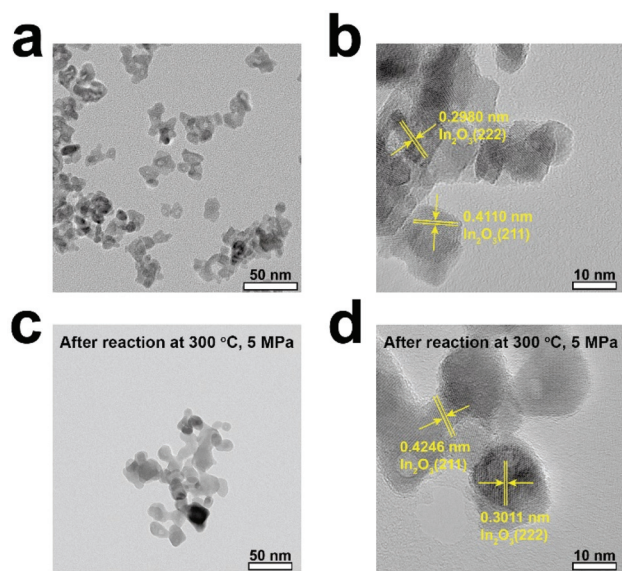
#### Raman spectroscopy

Raman spectroscopy has been believed to provide direct spectroscopic evidence of the oxygen vacancies in In<sub>2</sub>O<sub>3</sub>.<sup>50,51</sup> In the present study, the catalysts were analysed using room temperature-visible Raman scattering to probe the oxygen vacancies in

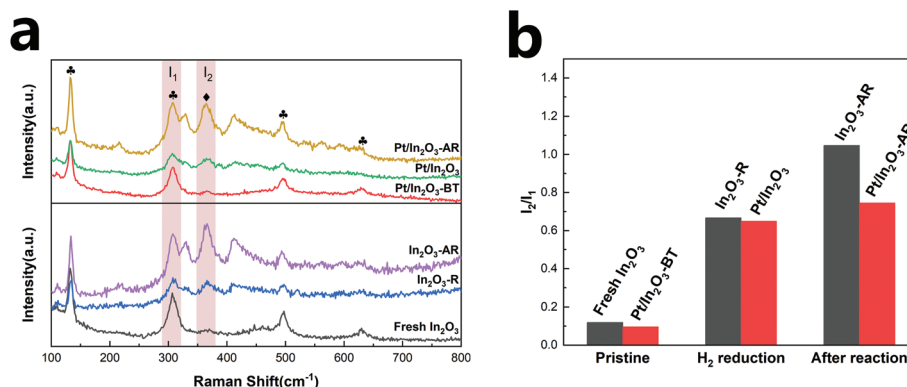
In<sub>2</sub>O<sub>3</sub>. The Raman spectra are shown in Fig. 5a. The peak at 131 cm<sup>-1</sup> is assigned to the In-O vibration of the InO<sub>6</sub> structural units.<sup>52</sup> The peak at 307 cm<sup>-1</sup> is usually interpreted as the bending vibration  $\delta(\text{InO}_6)$  of octahedra.<sup>53</sup> The two peaks at 497 cm<sup>-1</sup> and 629 cm<sup>-1</sup> are attributed to the stretching vibration of the same  $\nu(\text{InO}_6)$  octahedra.<sup>50</sup> The scattering feature at 367 cm<sup>-1</sup> can be attributed to the stretching vibration of In-O-In ( $\nu(\text{In-O-In})$ ) and it reflects the presence of oxygen vacancies in the In<sub>2</sub>O<sub>3</sub>.<sup>15,50,51,54,55</sup> The intensities of the 367 cm<sup>-1</sup> peak of Pt/In<sub>2</sub>O<sub>3</sub> and In<sub>2</sub>O<sub>3</sub>-R (In<sub>2</sub>O<sub>3</sub> after hydrogen reduction) increase significantly compared to those of the fresh samples, consistent with the expectation that hydrogen reduction creates more oxygen vacancies.

In order to quantify the oxygen vacancies, we integrated the peaks at 307 cm<sup>-1</sup> (*I*<sub>1</sub>) and 367 cm<sup>-1</sup> (*I*<sub>2</sub>) and used the *I*<sub>2</sub>/*I*<sub>1</sub> ratio to characterize the amount of the oxygen vacancies on In<sub>2</sub>O<sub>3</sub>.<sup>56</sup> As shown in Fig. 5b, the oxygen vacancies in the Pt/In<sub>2</sub>O<sub>3</sub>-BT sample are 81% of those of the pure In<sub>2</sub>O<sub>3</sub> catalyst, likely due to the presence of the PtO<sub>x</sub> species covering some oxygen vacancies. However, the amount of the oxygen vacancies on the Pt/In<sub>2</sub>O<sub>3</sub> catalyst increases to 92% of the In<sub>2</sub>O<sub>3</sub>-R sample after hydrogen reduction. This suggests that the well-dispersed Pt nanoparticles have a favourable effect on the formation of the oxygen vacancies over In<sub>2</sub>O<sub>3</sub>.

More importantly, the amount of the oxygen vacancies on the pure In<sub>2</sub>O<sub>3</sub> and the Pt/In<sub>2</sub>O<sub>3</sub> catalysts varies quite differently after the reaction at 300 °C and 5 MPa. In<sub>2</sub>O<sub>3</sub>-AR has many more oxygen vacancies than Pt/In<sub>2</sub>O<sub>3</sub>-AR. However, the catalytic activity and stability of the pure In<sub>2</sub>O<sub>3</sub> catalyst are worse than those of the Pt/In<sub>2</sub>O<sub>3</sub> catalyst. These results demonstrate that the additional oxygen vacancies created at high temperature in In<sub>2</sub>O<sub>3</sub> had a negative effect on the reaction, causing In<sub>2</sub>O<sub>3</sub> over-reduction and destabilization of the cata-



**Fig. 4** TEM images of (a and b) the fresh In<sub>2</sub>O<sub>3</sub> catalyst and (c and d) the pure In<sub>2</sub>O<sub>3</sub> catalyst after reaction.



**Fig. 5** (a) Visible Raman spectra of the Pt/In<sub>2</sub>O<sub>3</sub> and pure In<sub>2</sub>O<sub>3</sub> catalysts, (b) comparison of the relative amount of the oxygen vacancies according to the I<sub>2</sub>/I<sub>1</sub> ratio. ♠: InO<sub>6</sub> octahedron, ♦: the stretching vibrations of In-O-In. In<sub>2</sub>O<sub>3</sub>-R represents the In<sub>2</sub>O<sub>3</sub> catalyst after hydrogen reduction (H<sub>2</sub>/N<sub>2</sub> = 1/9, molar ratio, at 200 °C for 1 h).

lyst. This is, in general, consistent with Tsoukalou *et al.*,<sup>19</sup> who showed that the over-reduction of In<sub>2</sub>O<sub>3</sub> results in the formation of metallic indium, which is inactive for hydrogen activation.

#### Hydrogen temperature programmed reduction

H<sub>2</sub>-TPR was used to characterize the reduction behaviour of the catalyst and the formation of oxygen vacancies. As shown in Fig. 6a, the reduction peak at 133 °C of pure In<sub>2</sub>O<sub>3</sub> can be assigned to the reduction of the surface, corresponding to the formation of the surface oxygen vacancies.<sup>15,18,54</sup> The reduction peaks appear at much lower temperatures of 11 °C and 41 °C for the Pt/In<sub>2</sub>O<sub>3</sub>-BT sample. These peaks can be assigned to the reduction of the PtO<sub>x</sub> species with the assistance of the Pt nanoparticles.<sup>29</sup> The formation of PtO<sub>x</sub> also indicates the existence of the SMSI between Pt and In<sub>2</sub>O<sub>3</sub>, which in turn promotes dispersion of the Pt nanoparticles.<sup>47,57,58</sup> Furthermore, the active H adatoms spilled over from the Pt

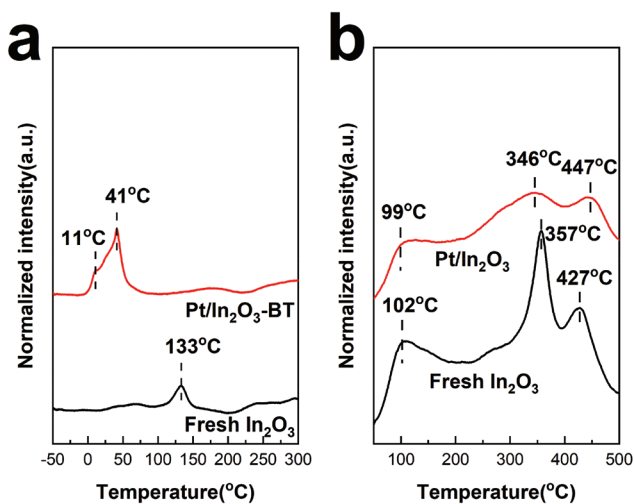
nanoparticles to In<sub>2</sub>O<sub>3</sub> can also cause the reduction of the In<sub>2</sub>O<sub>3</sub> surface,<sup>18,29,59</sup> making the In<sub>2</sub>O<sub>3</sub> surface reduction peak disappear. The normalized intensity of the reduction peak of the Pt/In<sub>2</sub>O<sub>3</sub>-BT sample is much higher than that of In<sub>2</sub>O<sub>3</sub>. This also confirms that the well-dispersed Pt nanoparticles promote the formation of the surface oxygen vacancies of In<sub>2</sub>O<sub>3</sub> during hydrogen reduction. This is in line with the results of the Raman analyses discussed above.

#### CO<sub>2</sub> temperature programmed desorption

Generally, the surface oxygen vacancies on In<sub>2</sub>O<sub>3</sub> promote CO<sub>2</sub> adsorption and activation.<sup>15</sup> Therefore, CO<sub>2</sub> can be used as a probe molecule to characterize the surface oxygen vacancies on In<sub>2</sub>O<sub>3</sub> through CO<sub>2</sub>-TPD.<sup>60</sup> The profiles of CO<sub>2</sub>-TPD in Fig. 6b confirm that less CO<sub>2</sub> is adsorbed on the Pt/In<sub>2</sub>O<sub>3</sub> catalyst than on the pure In<sub>2</sub>O<sub>3</sub> catalyst. The peak located at *ca.* 100 °C for these two samples is assigned to the physically adsorbed CO<sub>2</sub>. It is noteworthy that the normalized intensity of the desorption peak at 357 °C for fresh In<sub>2</sub>O<sub>3</sub> is much stronger than that for the Pt/In<sub>2</sub>O<sub>3</sub> catalyst. This peak can be attributed to CO<sub>2</sub> adsorption in the hydrogen-induced surface oxygen vacancies.<sup>18</sup> The peaks at 447 °C for the Pt/In<sub>2</sub>O<sub>3</sub> catalyst and 427 °C for the pure In<sub>2</sub>O<sub>3</sub> catalyst belong to CO<sub>2</sub> adsorption in the thermally induced oxygen vacancies.<sup>14,15,18</sup> These results confirm that the total amount of the surface oxygen vacancies on the Pt/In<sub>2</sub>O<sub>3</sub> catalyst is smaller than that of the pure In<sub>2</sub>O<sub>3</sub> catalyst after hydrogen reduction, which is consistent with the results of the Raman spectra above.

#### UV-vis absorption spectra

Fig. 7a shows the UV-vis absorption spectra of the Pt/In<sub>2</sub>O<sub>3</sub> catalyst. Fresh In<sub>2</sub>O<sub>3</sub> shows a strong UV absorption because of its semiconducting nature. The band gap can be determined from absorption spectra using the Tauc relationship. Based on the plot of  $(ah\nu)^2$  versus the photoenergy,  $h\nu$ , in Fig. 7b, the band gap of the Pt/In<sub>2</sub>O<sub>3</sub>-BT sample is determined to be 2.8 eV. The band gap is reduced to 2.57 eV for the Pt/In<sub>2</sub>O<sub>3</sub> catalyst. This result indicates that the creation of the oxygen



**Fig. 6** (a) H<sub>2</sub>-TPR profiles and (b) CO<sub>2</sub>-TPD profiles of the Pt/In<sub>2</sub>O<sub>3</sub> and pure In<sub>2</sub>O<sub>3</sub> catalysts.

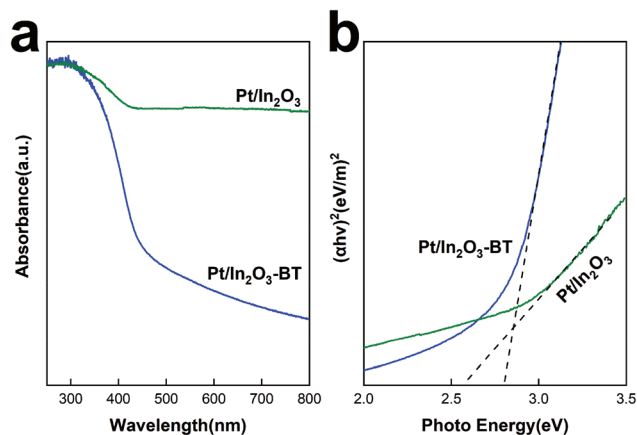


Fig. 7 (a) UV-vis absorption spectra and (b) Tauc plot of the Pt/In<sub>2</sub>O<sub>3</sub> catalyst.

vacancies shifts the onset of light absorption from the UV to the visible region. It is also consistent with the Raman analyses discussed above: oxygen vacancies can be created over the Pt/In<sub>2</sub>O<sub>3</sub> catalyst after hydrogen reduction. According to the literature,<sup>50,51</sup> the increase in visible light absorption can be attributed to the electronic transition from the In<sub>2</sub>O<sub>3</sub> valence band to the newly created oxygen vacancy level. The sensitivity of visible light absorption to the oxygen vacancies in In<sub>2</sub>O<sub>3</sub> also makes Pt/In<sub>2</sub>O<sub>3</sub> potentially useful in exploiting visible light for photocatalysts.

## Conclusions

In this work, we prepared a novel Pt/In<sub>2</sub>O<sub>3</sub> catalyst using the deposition-precipitation method. The catalyst was tested for methanol synthesis from CO<sub>2</sub> and hydrogen and the methanol selectivities of ca. 100% at temperatures below 225 °C, 74% at 275 °C and 54% at 300 °C on the Pt/In<sub>2</sub>O<sub>3</sub> catalyst have been achieved. These results represent a significant improvement from the pure In<sub>2</sub>O<sub>3</sub> catalyst and are the highest methanol selectivity reported on Pt catalysts. For example, the Pt/In<sub>2</sub>O<sub>3</sub> catalyst shows a CO<sub>2</sub> conversion above 17% with a methanol formation rate of 0.542 g<sub>methanol</sub> h<sup>-1</sup> g<sub>cat</sub><sup>-1</sup> at 300 °C and 5 MPa, while the CO<sub>2</sub> conversion is only 9.4% and the methanol formation rate is 0.335 g<sub>methanol</sub> h<sup>-1</sup> g<sub>cat</sub><sup>-1</sup> for In<sub>2</sub>O<sub>3</sub>. In addition to the improved activity, the stability of the Pt/In<sub>2</sub>O<sub>3</sub> catalyst is also outstanding. The methanol formation rate on the Pt/In<sub>2</sub>O<sub>3</sub> catalyst after 5 h in the reaction stream only decreases to 95% of the initial rate and remains constant thereafter. This is significantly higher than 80% of the initial rate on the pure In<sub>2</sub>O<sub>3</sub> catalyst reached after 9 h in the reaction stream. The textural characterization (by BET, XRD and TEM) of the catalysts indicates that the Pt nanoparticles are well-dispersed with an average particle size of 1.47 nm. More importantly, TEM and H<sub>2</sub>-TPR showed that there were SMSIs between Pt and In<sub>2</sub>O<sub>3</sub>, resulting in the improved surface stability. In addition, the results of Raman spectra confirm that the

SMSI between Pt and In<sub>2</sub>O<sub>3</sub> inhibits the over-reduction of In<sub>2</sub>O<sub>3</sub>. Consequently, the Pt/In<sub>2</sub>O<sub>3</sub> catalyst showed a much higher catalytic activity and better stability than the pure In<sub>2</sub>O<sub>3</sub> catalyst.

## Conflicts of interest

There are no conflicts to declare.

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