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Reductive dehydrocoupling of diphenyltin dihydride with LiAlH₄: selective synthesis and structures of the first bicyclo[2.2.1]heptastannane-1,4-diide and bicyclo[2.2.2]octastannane-1,4-diide†

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The reaction of diphenyltin dihydride with LiAlH₄ gives access to a set of charged tin cages as their lithium salts. Variation in the ratio of reactants provides a perstannabicyclooctane dianion and a perstananorbornane as the di- and monoanions. These compounds can be synthesised selectively by careful stoichiometric control and have been characterised by single crystal X-ray diffractometry, NMR and UV-vis spectroscopy. Computational exploration of the electronic structures of these compounds was undertaken and, in agreement with structural and spectroscopic features, indicated significant σ -delocalisation in the tin skeletons.

The chemistry of anionic oligo-tin cage and cluster compounds is dominated by a plethora of Zintl phases, discrete Zintl ions,¹ $[\text{Sn}_n]^{X-}$, or their derivatisation products, $[\text{R}_m\text{Sn}_n]^{X-}$.² Derivatisation reactions of the Zintl ion Sn_9^{4-} have led to the trianionic compounds RSn_9^{3-} .³ Reaction of thermolabile tin(i) halides with $[\text{Si}(\text{SiMe}_3)_3]^-$ resulted in the isolation of neutral $\text{Sn}_{10}[\text{Si}(\text{SiMe}_3)_3]_6$,⁴ the monoanion $\text{Sn}_{10}[\text{Si}(\text{SiMe}_3)_3]_5^-$,⁵ and the dianions $\text{Sn}_9[\text{Si}(\text{SiMe}_3)_3]_4^{2-}$ and $\text{Sn}_{10}[\text{Si}(\text{SiMe}_3)_3]_4^{2-}$.^{6,7} Moreover, the anionic cluster $\text{Sn}_9\text{R}_7(\text{NHC})^-$ ($\text{R} = \text{CH}(\text{SiMe}_3)_2$, $\text{NHC} = 1,2,3,4$ -tetramethylimidazol-2-ylidene) was isolated from the reaction of the trihydride RSnH_3 with the corresponding NHC.⁸ Alongside these, several neutral metalloid clusters of the form Sn_nR_m ($n > m$) have been obtained from reductive or dehydrogenative coupling methods.⁹ Closely related to metalloid compounds,^{10a} albeit with a superstoichiometric ratio of substituents to tin, are the elementoid^{10b} [1.1.1]pentastannapropellanes, Sn_5R_6 ,¹¹ tetra-cyclic Sn_7R_8 ,¹² and the hexastannabenzene isomer, Sn_6R_6 .¹³

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The aforementioned compounds share a non-classical, *i.e.* 3D-delocalised, bonding situation. This confers a narrow energy gap between frontier orbitals which facilitates interesting addition chemistry and redox behaviour.¹⁴ In contrast to the well represented class of simple monocyclic rings, $(\text{R}_2\text{Sn})_n$,¹⁵ purely σ -bonded tin cages, $(\text{RSn})_n$, are sparse. Examples include tricyclo- and pentacycloprians of the stoichiometry, R_6Sn_6 and $\text{R}_{10}\text{Sn}_{10}$,^{8,16} cubanes, R_8Sn_8 ,¹⁷ and a tetrahedral cage molecule with edge-bridging methanidyl substituents.¹⁸ Anionic, covalent, oligotin cages include $\text{Sn}_8\text{R}_6^{2-}$ ($\text{R} = \text{Si}-t\text{-Bu}_3$),^{16c} and $\text{Sn}_5(\text{CH}_3)\text{R}_6^-$ and the radical anion Sn_5R_6^- ($\text{R} = 2,6\text{-Et}_2\text{-C}_6\text{H}_3$). The latter species originate from the addition of methyl lithium to, and the electrochemical one-electron reduction of, a [1.1.1]pentastannapropellane.^{14b} Organotindihydrides are common starting materials in the synthesis of oligotin compounds. Their dehydrogenative coupling has also been widely applied in the synthesis of linear tin polymers and cyclic oligomers $(\text{R}_2\text{Sn})_n$,¹⁹ whilst the reaction of a tin dihydride with sodium in liquid ammonia led to the formation of Ph_2SnH_2 and $\text{NaPh}_2\text{SnSnPh}_2\text{Na}$.²⁰

Diphenyltin dihydride can be synthesised by the reaction of LiAlH₄ with diphenyltin dichloride.²¹ During our repetition of this literature method, we noted the reaction mixture would often acquire a deep red or yellow colour upon the use of large excesses of LiAlH₄. This resulted in a significant decrease in yield of the desired diphenyltin dihydride. Thus, in order to elucidate the nature of possible over-reaction products, analytically pure Ph_2SnH_2 was reacted with LiAlH₄.²²

In an initial reaction, an equimolar mixture of the two reagents in THF was observed to provide a persistent bubbling and the formation of a dark red solution, as well as a flocculent white precipitate, identified as AlH_3 by IR spectroscopy. Addition of 12-crown-4 (hereafter, 12-Cr-4) and storage at $-30\text{ }^\circ\text{C}$ yielded a small amount of an intensely yellow material that was thought to be a single, crystalline product. X-ray diffraction analysis, however, indicated this material to be a disordered mixture of two oligostannane dianions (**1** and **2**). In order to access analytically



pure samples of **1** and **2**, the ratio of LiAlH_4 to Ph_2SnH_2 was optimised.

The reaction of 7 equivalents of diphenyltin dihydride with two equivalents of LiAlH_4 in THF or DME gave access to crystalline, analytically pure $[\text{Ph}_{10}\text{Sn}_7]^{2-}[\text{Li}(\text{12-Cr-4})_2]^{+}$, **1**, in moderate yields after the addition of 12-Cr-4 and storage at -30°C . Similarly, $[\text{Ph}_{12}\text{Sn}_8]^{2-}[\text{Li}(\text{12-Cr-4})_2]^{+}$, **2**, was obtained by applying analogous conditions albeit with a 8:2 ratio of $\text{Ph}_2\text{SnH}_2:\text{LiAlH}_4$. These optimised conditions yielded material suitable for X-ray crystallography as orange and red crystals, respectively. Higher tin to alumane ratios of 8:1 and above yielded the bicyclo[2.2.2]heptastanna-1-ide $[\text{Ph}_{13}\text{Sn}_8]^{2-}[\text{Li}(\text{12-Cr-4})_2]^{+}$, **3**. **3** was characterised by NMR and UV-vis spectroscopy as well as X-ray crystallography.²³ A $\text{Ph}_2\text{SnH}_2:\text{LiAlH}_4$ ratio of 1:0.75 and 1:2 led to the formation of elemental tin and compounds **4** and **5** which contain discrete $[\text{PhAlH}_3]^{2-}$ and $[\text{Ph}_3\text{AlH}]^{2-}$ anions and charge separated $[\text{Li}(\text{12-Cr-4})_2]^{+}$ or $[\text{Li}(\text{DME})(\text{12-Cr-4})]^{+}$ counterions. The formation of **4** and **5** demonstrates net phenyl group transfer from tin to aluminium which provides the tertiary tin bridgeheads in **1–3** with less than the original two phenyl groups per tin. NMR spectroscopic interrogation of the reaction mixtures indicated the initial formation of benzene and of the anion $[\text{HPh}_2\text{Sn}]^{2-}$ (^{119}Sn : -180.5 ppm, $^1J_{\text{H}-^{119}\text{Sn}} = 152$ Hz),²⁰ which undergoes subsequent dehydrogenative coupling with excess Ph_2SnH_2 to cause the intermediate formation of various unidentified oligostannyl anions. Isotopic labelling experiments showed the predominant formation of H_2 in the reaction of Ph_2SnH_2 with LiAlD_4 , while treatment of Ph_2SnD_2 with LiAlH_4 gave D_2 (see ESI† page S31).

1 and **2** constitute the first reports of a bicyclo[2.2.1]-heptastannane-1,4-diide and bicyclo[2.2.2]octastannane-1,4-diide, respectively, and both crystallise in the presence of 2 $[\text{Li}(\text{12-Cr-4})_2]^{+}$ counterions as charge separated structures. In the solid state structure of the bicyclo[2.2.1]heptastannane-1,4-diide **1** (Fig. 1) the anionic bridgeheads $\text{Sn}1$ and $\text{Sn}7$ are separated by an intramolecular distance of 5.04 Å. The angles

around $\text{Sn}1$ and $\text{Sn}7$ sum to 265.8° ($\text{Sn}1$) and 264.6° ($\text{Sn}7$), suggesting high p-orbital contribution in bonding orbitals at $\text{Sn}1$ and $\text{Sn}7$ and implying significant s-orbital character for the lone pair at these anionic tin atoms.

The bicyclo[2.2.2]octastanna-1,4-diide **2** (Fig. 1), features a less constrained geometry in the solid state with angles around the bridgehead atoms $\text{Sn}1$ and $\text{Sn}8$ just above 90° (sum of angles around $\text{Sn}1 = 276.05^\circ$ and $\text{Sn}8 = 274.08^\circ$). The mono-anion **3** displays two distinctly different environments for the bridgehead tin atoms with a highly pyramidalised anionic tin atom $\text{Sn}1$ (sum of angles 287.69°) and a more tetrahedral geometry at the phenyl substituted, formally neutral bridgehead $\text{Sn}4$ (sum of Sn-Sn-Sn angles 318.62°). The bridgehead atoms, $\text{Sn}1$ and $\text{Sn}4$, are separated by 5.27 Å in **3**, which is *ca.* 0.67 Å less than in **2**. In contrast to **1**, compounds **2** and **3** are both twisted in the solid state (see ESI† Fig. S2.2, page S-7 and S2.3, page S-8). Compounds **4** and **5** (Fig. 1) containing discrete $[\text{PhAlH}_3]^{2-}$ and $[\text{Ph}_3\text{AlH}]^{2-}$ anions and charge separated $[\text{Li}(\text{12-Cr-4})_2]^{+}$ or $[\text{Li}(\text{DME})(\text{12-Cr-4})]^{+}$ counterions, respectively, display tetrahedral aluminium atoms with Al-C distances slightly greater than 2.0 Å.

The $^{119}\text{Sn}\{^1\text{H}\}$ NMR spectra of **1**, **2** and **3** indicate that these cages persist in solution, as evidenced by chemical shifts and the observed $^{119}\text{Sn}/^{117}\text{Sn}$ couplings (Fig. 2, ESI† Section 4 pages S-14–S-26). **1** displays three resonances at 200.3 (Ph_2Sn), 35.7 ($\text{Ph}_2\text{SnSnPh}_2$) and -857.3 (Sn^-) ppm for which the coupling pattern is in full agreement with the bicyclo[2.2.1]heptastannane-1,3-diide framework in solution. $^1J_{\text{Sn}-^{119/117}\text{Sn}}$ coupling constants originating from the anionic bridgehead atoms are large with values of 4640/4430 Hz for the interaction with the monotin-bridge. In the case of coupling between the bridgehead tin atoms and the ditin bridges, values of 5990/5710 Hz for the $^1J_{\text{Sn}-^{119/117}\text{Sn}}$ coupling and 1130/1080 Hz for the $^2J_{\text{Sn}-^{119/117}\text{Sn}}$ coupling are observed. The relatively large 1J values compare to a smaller $^1J_{\text{Sn}-^{119}\text{Sn}}$ coupling constant of 3940 Hz for chemically equivalent tin atoms within the same ditin bridge.

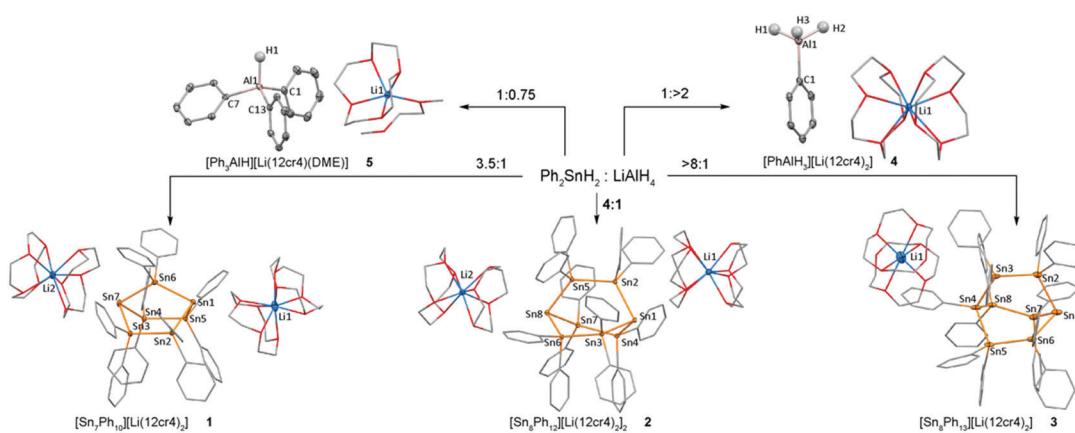


Fig. 1 Formation of anionic oligotin cages **1–3** and phenyl aluminium hydrides **4** and **5**. Metal atoms are drawn at 30% probability level, only hydrogen atoms attached to Al are shown, all other H atoms are omitted for clarity. For full geometric parameters, see ESI† pages S-7–S-9. Selected bond angles (°). **1**: $\text{Sn}_2-\text{Sn}1-\text{Sn}5$ 94.37(2); $\text{Sn}_2-\text{Sn}1-\text{Sn}6$ 86.75(1); $\text{Sn}_5-\text{Sn}1-\text{Sn}6$ 84.67(1); $\text{Sn}_3-\text{Sn}7-\text{Sn}4$ 93.52(2); $\text{Sn}_3-\text{Sn}7-\text{Sn}6$ 85.41(1); $\text{Sn}_4-\text{Sn}7-\text{Sn}6$ 85.69(1). **2**: $\text{Sn}_2-\text{Sn}1-\text{Sn}3$ 91.29(2); $\text{Sn}_2-\text{Sn}1-\text{Sn}4$ 91.41(2); $\text{Sn}_3-\text{Sn}1-\text{Sn}4$ 93.35(2); $\text{Sn}_5-\text{Sn}8-\text{Sn}6$ 90.95(2); $\text{Sn}_5-\text{Sn}8-\text{Sn}7$ 89.00(2); $\text{Sn}_6-\text{Sn}8-\text{Sn}7$ 94.13(2). **3**: $\text{Sn}_2-\text{Sn}1-\text{Sn}6$ 96.92(8); $\text{Sn}_2-\text{Sn}1-\text{Sn}7$ 95.47(8); $\text{Sn}_6-\text{Sn}1-\text{Sn}7$ 95.30(8); $\text{Sn}_3-\text{Sn}4-\text{Sn}5$ 106.93(8); $\text{Sn}_3-\text{Sn}4-\text{Sn}8$ 105.22(8); $\text{Sn}_5-\text{Sn}4-\text{Sn}8$ 106.47(8).



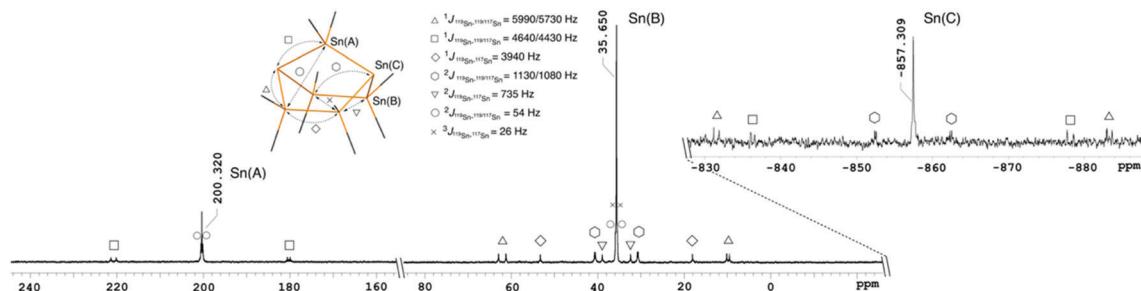


Fig. 2 $^{119}\text{Sn}\{^1\text{H}\}$ NMR spectrum of **1** in $\text{d}_3\text{-acetonitrile}$ showing Sn–Sn couplings indicative of the solution integrity of these compounds.

ng to its higher molecular symmetry, only two resonances are observed in the $^{119}\text{Sn}\{^1\text{H}\}$ NMR spectrum of **2**, with peaks at -316.8 ($\text{Ph}_2\text{SnSnPh}_2$) and -585.0 (Sn^-) ppm. The ^1J and the $^2\text{J}_{\text{Sn}-119/\text{Sn}-117}$ coupling constants between the bridgehead tin atoms and the $\text{Ph}_2\text{SnSnPh}_2$ bridges are $5020/4800$ Hz and $807/770$ Hz, respectively. Again, a smaller $^1\text{J}_{\text{Sn}-117}$ coupling constant of 3610 Hz is observed for neighbouring diphenyltin fragments. The monoanion **3** displays four distinct signals in the $^{119}\text{Sn}\{^1\text{H}\}$ NMR spectrum at -183.1 (Ph_2SnSn^-), -238.6 (Ph_2SnSnPh), -470.8 (SnPh) and -757.9 (Sn^-) ppm. Notably, the coupling constant of the tricoordinate, anionic tin atom in **3** is larger ($^1\text{J}_{\text{Sn}-119/\text{Sn}-117} = 6410/6090$ Hz) in comparison to the signal of the neutral tin bridgehead, which is only $1150/1090$ Hz.

Experimentally observed NMR chemical shifts were replicated by DFT calculations. To support the bonding model proposed based on the structural analysis of **1–3**, the anions in these compounds were interrogated by DFT calculations employing different methods and basis sets. (For computational details and references see ESI† page S-84). The highest occupied molecular orbitals of **1–3** are shown in Fig. 3. These orbitals are in each case associated with the lone pairs on the anionic bridgehead tin atom(s) with contribution of the tin skeleton and show significant s-character.

This supports the conclusion of significant p-orbital involvement in the Sn–Sn bonding of the bridgehead tin atoms inferred from their experimental solid state structures. The frontier orbitals of **1–3** (Fig. 3 and ESI† pages S-37–S-40) display distinct groupings into sets of orbitals with σ -, σ^* - and π^* -character interpreted as the onset of band-like behaviour. The HOMO to HOMO–8 ($\Delta E_{\text{HOMO-HOMO-8}} = 1.78$ eV) for **2** and HOMO to HOMO–7 orbitals for **1** ($\Delta E_{\text{HOMO-HOMO-7}} = 2.03$ eV) and **3** ($\Delta E_{\text{HOMO-HOMO-7}} = 1.66$ eV) essentially represent the

σ -bonded tin cores with only minor orbital contributions from the phenyl substituents. Within HOMO–LUMO energy gaps of only 2.33 eV (**1**), 2.15 eV (**2**) and 2.47 eV (**3**), the character of the frontier orbitals changes to phenyl based π^* character.

The LUMO to LUMO $+n$ orbitals ($n = 20$ in **1**, $n = 23$ in **2**, $n = 25$ in **3**) are similar in orbital energy and all localised on the phenyl substituents. This manifold of π^* orbitals is then followed by a set of orbitals with predominant σ^* character of the tin framework, which span energy ranges of 0.78 eV (**1**), 0.81 eV (**2**) and 1.34 eV (**3**). These closely spaced MOs are reflected in the electronic absorptions displayed by **1–3**. The UV-visible spectra of these species do not display distinct absorption maxima and are instead broad and tailing, suggestive of weak but extensive absorptions. Nevertheless, the visibly observed colours for crystals of **1** (orange-yellow), **2** (orange-red) and **3** (bright yellow) are consistent with the expected effects of cage size and charge upon σ -delocalisation²⁴ inferred from the DFT calculations.

In summary, the reaction of Ph_2SnH_2 with LiAlH_4 provides facile access to the charge-separated species, **1**, **2** and **3**, comprising an unprecedented set of structural motifs in anionic oligostannane cages. **1** and **2** constitute new dianionic covalent tin cages and we propose that they will provide convenient synthons in further transformations because of their charged nature and solution integrity. The electronic structure of **1–3** was interrogated computationally and support the presence of σ -delocalisation.

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Conflicts of interest

The authors declare no conflict of interest.

Notes and references

- (a) S. C. Sevov and J. M. Goicoechea, *Organometallics*, 2006, **25**, 5678–5692; (b) T. F. Fässler, *Coord. Chem. Rev.*, 2001, **215**, 347–377.
- (a) A. Schnepf, *Chem. Soc. Rev.*, 2007, **36**, 745–758; (b) A. Schnepf, *Struct. Bonding*, 2016, 193–223.

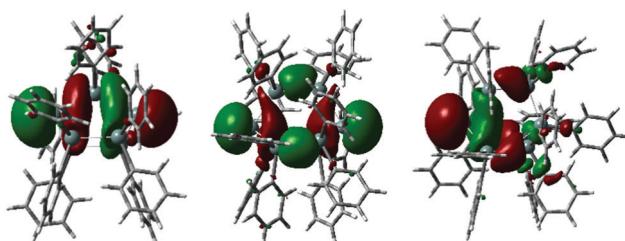


Fig. 3 Highest occupied molecular orbitals of **1–3**, for full computational details see ESI† Section 5, pages S-34–S-83.

3 (a) F. S. Kocak, D. O. Downing, P. Zavalij, Y. F. Lam, A. N. Vedernikov and B. Eichhorn, *J. Am. Chem. Soc.*, 2012, **134**, 9733–9740; (b) F. S. Kocak, P. Y. Zavalij, Y.-F. Lam and B. W. Eichhorn, *Chem. Commun.*, 2009, 4197–4199; (c) F. S. Kocak, D. O. Downing, P. Zavalij, Y. F. Lam, A. N. Vedernikov and B. Eichhorn, *J. Am. Chem. Soc.*, 2012, **134**, 9733–9740; (d) D. J. Chapman and S. C. Sevov, *Inorg. Chem.*, 2008, **47**, 6009–6013; (e) C. B. Benda, M. Waibel and T. F. Fässler, *Angew. Chem., Int. Ed.*, 2015, **54**, 522–526.

4 (a) C. Schrenk, A. Kubas, K. Fink and A. Schnepf, *Angew. Chem., Int. Ed.*, 2011, **50**, 7273–7277; (b) C. Schrenk, I. Schellenberg, R. Pöttgen and A. Schnepf, *Dalton Trans.*, 2010, **39**, 1872–1876; (c) C. Schrenk and A. Schnepf, *Chem. Commun.*, 2010, **46**, 6756–6758.

5 (a) C. Schrenk, J. Helmlinger and A. Schnepf, *Z. Anorg. Allg. Chem.*, 2012, **638**, 589–593; (b) C. Schrenk, B. Gerke, R. Pöttgen, A. Clayborne and A. Schnepf, *Chem. – Eur. J.*, 2015, **21**, 8222–8228.

6 C. Schrenk, F. Winter, R. Pöttgen and A. Schnepf, *Inorg. Chem.*, 2012, **51**, 8583–8588.

7 C. Schrenk, F. Winter, R. Pöttgen and A. Schnepf, *Chem. – Eur. J.*, 2015, **21**, 2992–2997.

8 J.-J. Maudrich, C. P. Sindlinger, F. S. W. Aicher, K. Eichele, H. Schubert and L. Wesemann, *Chem. – Eur. J.*, 2017, **23**, 2192–2200.

9 (a) E. Rivard, J. Steiner, J. C. Fettinger, J. R. Giuliani, M. P. Augustine and P. P. Power, *Chem. Commun.*, 2007, 4919–4921; (b) G. Prabusankar, A. Kempter, C. Gemel, M. K. Schröter and R. A. Fischer, *Angew. Chem., Int. Ed.*, 2008, **47**, 7234–7237; (c) B. E. Eichler and P. P. Power, *Angew. Chem., Int. Ed.*, 2001, **40**, 796–797; (d) A. F. Richards, B. E. Eichler, M. Brynda, M. M. Olmstead and P. P. Power, *Angew. Chem., Int. Ed.*, 2005, **44**, 2546–2549; (e) J. Wiederkehr, C. Wölper and S. Schulz, *Chem. Commun.*, 2016, **52**, 12282–12285; (f) M. Brynda, R. Herber, P. B. Hitchcock, M. F. Lappert, I. Nowik, P. P. Power, A. V. Protchenko, A. Růžička and J. Steiner, *Angew. Chem., Int. Ed.*, 2006, **45**, 4333–4337; (g) C. P. Sindlinger, A. Stasch, H. F. Bettinger and L. Wesemann, *Chem. Sci.*, 2015, **6**, 4737–4751; (h) C. P. Sindlinger, W. Grahneis, F. S. W. Aicher and L. Wesemann, *Chem. – Eur. J.*, 2016, **22**, 7554–7566; (i) C. P. Sindlinger, F. S. W. Aicher, H. Schubert and L. Wesemann, *Angew. Chem., Int. Ed.*, 2017, **56**, 2198–2202; (j) L. R. Sita, *Acc. Chem. Res.*, 1994, **27**, 191–197.

10 (a) A. Schnepf and H. Schnöckel, *Angew. Chem., Int. Ed.*, 2002, **41**, 3532–3554; (b) A. Schnepf and H. Schnöckel, Nanostructural Element Modifications: Synthesis and Structure of Elementoid Gallium Clusters, in *ACS Symposium Series 822, Group 13 Chemistry*, ed. P. J. Shapiro and D. A. Atwood, American Chemical Society, Washington DC, 2002.

11 (a) L. R. Sita and R. D. Bickerstaff, *J. Am. Chem. Soc.*, 1989, **111**, 6454–6456; (b) C. Drost, M. Hildebrand and P. Lönnecke, *Main Group Met. Chem.*, 2002, **25**, 93–98.

12 L. R. Sita and I. Kinoshita, *J. Am. Chem. Soc.*, 1992, **114**, 7024–7029.

13 C. P. Sindlinger and L. Wesemann, *Chem. Sci.*, 2014, **5**, 2739–2746.

14 (a) L. R. Sita and I. Kinoshita, *J. Am. Chem. Soc.*, 1991, **113**, 5070–5072; (b) L. R. Sita and I. Kinoshita, *J. Am. Chem. Soc.*, 1990, **112**, 8839–8843; (c) D. Nied, E. Matern, H. Berberich, M. Neumaier and F. Breher, *Organometallics*, 2010, **19**, 6018–6037; (d) P. Vasko, S. Wang, H. M. Tuononen and P. P. Power, *Angew. Chem., Int. Ed.*, 2015, **54**, 3802–3805; (e) D. Nied and F. Breher, *Chem. Soc. Rev.*, 2011, **40**, 3455–3466.

15 C. Marschner and J. Hlina, Catenated Compounds – Group 14 (Ge, Sn, Pb), in *Comprehensive Inorganic Chemistry II*, ed. J. Reedijk and K. Poepelmeier, Elsevier, Oxford, 2013, vol. 1.

16 (a) S. Masamune and L. R. Sita, *J. Am. Chem. Soc.*, 1985, **107**, 6390–6391; (b) L. R. Sita and I. Kinoshita, *J. Am. Chem. Soc.*, 1991, **113**, 1856–1857; (c) N. Wiberg, H.-W. Lerner, H. Nöth and W. Ponikwar, *Angew. Chem., Int. Ed.*, 1999, **38**, 1103–1105.

17 (a) L. R. Sita and I. Kinoshita, *Organometallics*, 1990, **9**, 2865–2867; (b) N. Wiberg, H. W. Lerner, S. Wagner, H. Nöth and T. Seifert, *Z. Naturforsch. B*, 1999, **54**, 877–880.

18 M. Wagner, M. Lutter, B. Zobel, W. Hiller, M. H. Prosenc and K. Jurkschat, *Chem. Commun.*, 2015, **51**, 153–156.

19 (a) T. Imori, V. Lu, H. Cai and T. D. Tilley, *J. Am. Chem. Soc.*, 1995, **117**, 9931–9940; (b) P. Braunstein, J. Durand, X. Morise, A. Tiripicchio and F. Ugozzoli, *Organometallics*, 2000, **19**, 444–450; (c) P. Braunstein and X. Morise, *Chem. Rev.*, 2000, **100**, 3541–3552; (d) N. R. Neale and T. D. Tilley, *J. Am. Chem. Soc.*, 2002, **124**, 3802–3803; (e) F. Choffat, P. Smith and W. Caseri, *J. Mater. Chem.*, 2005, **15**, 1789–1792; (f) F. Choffat, S. Käser, P. Wolfer, D. Schmid, R. Mezzenga, P. Smith and W. Caseri, *Macromolecules*, 2007, **40**, 7878–7889.

20 M. Trummer and W. Caseri, *Organometallics*, 2010, **29**, 3862–3867.

21 (a) A. E. Finholt, A. C. Bond, Jr., K. E. Wilzbach and H. I. Schlesinger, *J. Am. Chem. Soc.*, 1947, **69**, 2692–2696; (b) N. Braia, Y. Saihi, A. E. Azzouzi and F. Ferkous, *Asian J. Chem.*, 2009, **21**, 4628–4634.

22 The use of analytically pure LiAlH₄ (recryst. from diethyl ether) is crucial for the successful control of the stoichiometry.

23 The anionic part in 3 was well resolved and is fully consistent with NMR data. The X-ray structure determination of the cationic part and residual solvent molecules, however, was hampered by twinning and disorder. See ESI† pages S-4 and S-5.

24 (a) M. S. Hill, Homocatenation of Metal and Metalloid Main Group Elements, in *Metal–Metal Bonding: Structure and Bonding*, ed. G. Parkin, Springer, Berlin, 2010, vol. 136; (b) M. Jovanovic, D. Antic, D. Rooklin, A. Bande and J. Michl, *Chem. – Asian J.*, 2017, **12**, 1250–1263; (c) M. Jovanovic and J. Michl, *J. Am. Chem. Soc.*, 2018, **140**, 11158–11160; (d) M. Jovanovic and J. Michl, *J. Am. Chem. Soc.*, 2019, **141**, 13101–13113.

