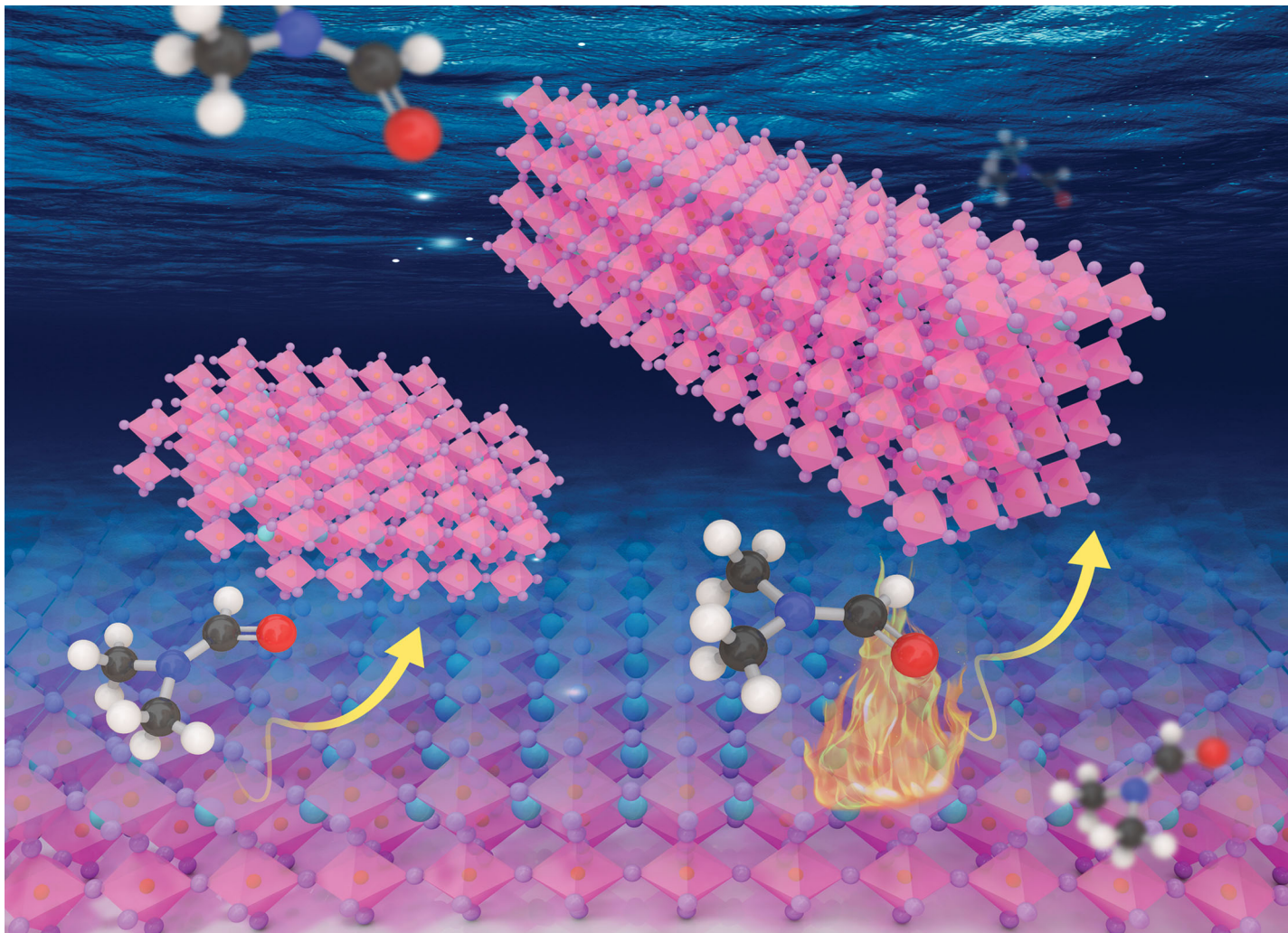




Showcasing research from Associate Professor Qiao Wen and Professor Junle Qu's laboratory, College of Physics and Optoelectronic Engineering, Shenzhen University, Shenzhen, China.

Photoelectronic mechanism investigation of the structural transformation of $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites from a subnanosheet to a microwire

A subnanosheet $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite was successfully structurally transformed into a microwire shape *via* temperature tuning and solvent effects. The photoelectronic characteristics of both types of $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite were investigated, and the structural transformation mechanism was demonstrated in experimental and theoretical. This study provides insight into the characteristic differences of $\text{CH}_3\text{NH}_3\text{PbI}_3$ upon structural transformation and offers suggestions for choosing a suitable perovskite structure for an application.

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Photoelectronic mechanism investigation of the structural transformation of $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites from a subnanosheet to a microwire

Yunjia Wang, Junle Qu * and Qiao Wen *

In this work, a subnanosheet $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite with a side length of 260 nm was successfully structurally transformed into a microwire shape with a radius of 1.8 μm and a length of 4.2 μm via temperature tuning and solvent effects. The photoelectronic characteristics of both types of $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite were investigated, and the structural transformation mechanism was demonstrated. TEM results showed that the crystal distance increased from 0.254 nm to 0.325 nm when the subnanosheet structure $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite transformed into the microwire shape. The rotation angle of the covalent bond decreased by 3.92° , based on the TEM results, which was attributed to the subnanosheet having a 0.03 eV lower energy than that of the microwire shape, as characterized by the fluorescence spectrum. Further analysis of the structural transformation was performed by employing a DC constant power source and a laser source at a wavelength of 446 nm. Experimental results showed that the resistance of the subnanosheet decreased from 14 Ω to 13.5 Ω under an applied voltage of 1 V and from 18 Ω to 15.5 Ω under an applied voltage of 5 V, which has the largest reduction ratio of resistance reaching 12.5% when it transformed into a microwire. Compared to the microwire-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite, the power absorption of the subnanosheet was 25% to 50% higher when the incident power ranged from 1.15 W to 0.25 W. This study provides insight into the characteristic differences of $\text{CH}_3\text{NH}_3\text{PbI}_3$ upon structural transformation and offers suggestions for choosing a suitable perovskite structure for an application.

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1 Introduction

$\text{CH}_3\text{NH}_3\text{PbI}_3$, as an organic–inorganic metal halide (OIMH) perovskite absorber, has drawn much attention due to its power conversion efficiency (PCE) and played a promising role in the development of next-generation two-dimensional materials.^{1–6} Most application research on the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite is correlated with its morphology or structure. For example, strip-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites have been mainly reported for solar cell applications, while subnanosheet $\text{CH}_3\text{NH}_3\text{PbI}_3$ have mostly been applied to mode locking of lasers and other optical applications.^{7–10} However, different constructions of $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites are basically dependent on the preparation methods.¹¹ To date, two common methods have been used for preparing $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites: one-step and two-step methods.^{12–14} Compared to the two-step method, the one-step method mainly involves mixing PbI_2 and $\text{CH}_3\text{NH}_3\text{I}$ in a solvent to form a $\text{CH}_3\text{NH}_3\text{PbI}_3$ film. Although the one-step

method is rather convenient for obtaining perovskites,^{15–17} changing the morphology or the structure by only dissolving organic and inorganic precursors appears to be difficult. In contrast, the two-step method involves first preparing a PbI_2 film by a spin coating method and subsequently reacting it with a $\text{CH}_3\text{NH}_3\text{I}$ film to obtain $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites.^{18–21}

Compared to the one-step method, the two-step method can be used to slightly tune the morphology or the structure by changing the composition and concentrations of the PbI_2 and $\text{CH}_3\text{NH}_3\text{I}$ films or by changing the spin coating method, but achieving structural transformation under complex conditions is still difficult.^{22–25} Therefore, more improved methods should be developed to obtain ideal $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites and apply them to optical and electrical applications. In 2013, a strip-shaped perovskite as a solar cell absorber with a PCE of 9.5% was prepared by an improved two-step method, and in the same year, a higher PCE of a solar cell with a strip-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite obtained by improved methods was also reported. However, the mechanism of this higher PCE solar cell with a strip-shaped absorber structure was seldom demonstrated in the research.^{26–28} In 2017, a $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite with a nanosheet structure was prepared by the

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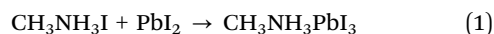


chemical vapor deposition (CVD) method and used for ultrafast pulsed fiber lasers, achieving stable soliton state mode locking.²⁹ In the next year, Dong *et al.* proposed a solvent-assisted thermal-pressure strategy to promote the secondary growth of $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite grains in the film, and as a result, photodetectors based on the high-quality perovskite film exhibited enhanced optoelectrical performance.³⁰ In the same year, a higher performance photodetector was also reported by applying a subnanosheet $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite, but the reason why the hexagonal shape was employed was hardly mentioned.³¹ Furthermore, the optoelectrical characteristics of $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites with different morphologies or structures for application were seldom considered or demonstrated when the structural transformation occurred.^{32–35} In addition, the growth mechanism of perovskite crystals has rarely been discussed, and the solvent effect was seldom noted in applications, nor were the optoelectrical differences of $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites with different morphologies or structures.^{36–38}

In our study, subnanosheet and microwire-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites were synthesized by an improved two-step method, and the morphology or the structure was tuned by the solvent effect and the reaction temperature of PbI_2 and $\text{CH}_3\text{NH}_3\text{I}$. The photoelectric properties of the different structure $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites were analyzed by UV-vis absorption spectroscopy, fluorescence spectroscopy, and Raman spectroscopy and using a DC constant power source. The subnanosheet perovskite showed better optical characteristics and weaker conductivity than the microwire-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite, and the results demonstrated the mechanism of optoelectronic change due to the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite structural transformation, which may be helpful for optimizing the optoelectronic character of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite and widening $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite applications.

2 Results and discussion

The flowchart in Fig. 1 shows the experimental processes involved in the preparation of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite. As depicted, certain amounts of $\text{CH}_3\text{NH}_3\text{I}$ crystals and PbI_2 crystals were dissolved in DMF solvent and reacted by the two-step method with stoichiometric coefficients. The reaction process was as follows:



The reaction was conducted in a nitrogen air atmosphere, not only to prevent the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite from reacting with O_2 , but also to avoid the DMF solvent from evaporating from the solution with the increasing temperature. The PbI_2 solution was added dropwise to ensure a smooth and steady structural transformation. After the reaction process, sample one was transferred to silica by a spin coating method for morphology characterization, as was sample two.

Fig. 2(a) shows a SEM image of sample one. As depicted, the prepared perovskite particle exhibited a regular hexagonal shape,

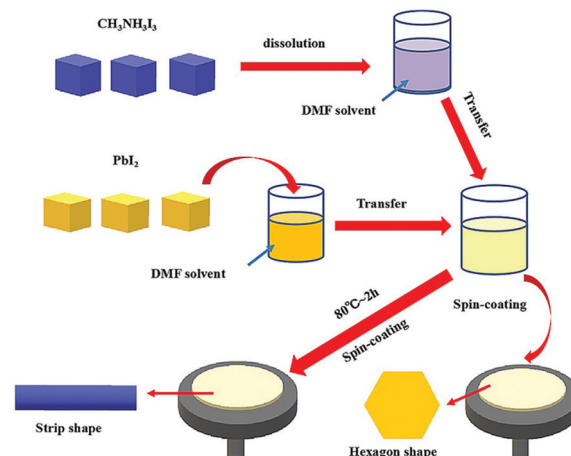


Fig. 1 Schematic illustration of the fabrication of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite.

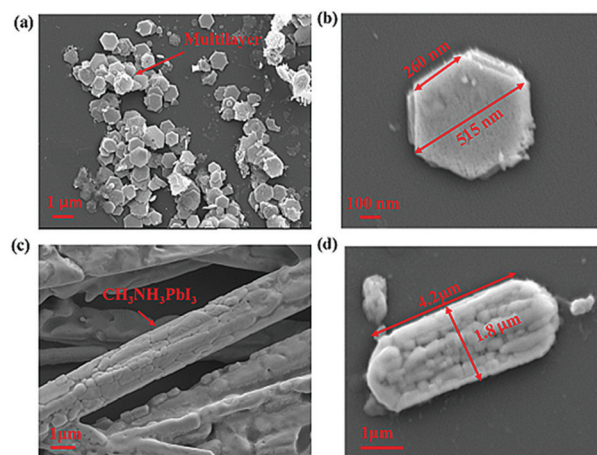


Fig. 2 (a and b) SEM image of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite prepared at room temperature; (c and d) SEM image of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite prepared at 80 °C.

and the side length of the subnanosheet perovskite was approximately 260 nm, as shown in Fig. 2(b). However, with the increase in the temperature to 80 °C, the structure was successfully transformed, as shown in the SEM image in Fig. 2(c). As depicted, the prepared perovskite particle exhibited a strip shape, with a length of approximately 4.2 μm and a width of approximately 1.8 μm, as shown in Fig. 2(d). We observed this phenomenon for two reasons: first, the DMF solvent applied during the preparation restricted the crystal growth, which ensured the overall shape of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite, but such restriction may differ in different directions; second, the thermodynamic factor played an important role in the crystal growth of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite, which accelerated the structural transformation due to the thermal energy. More details of the transformation mechanism will be discussed below.

Fig. 3 shows the lattice fringes of the two different shaped perovskites, characterized by TEM to further study the inner structure. As depicted, the distance between the lattice fringes



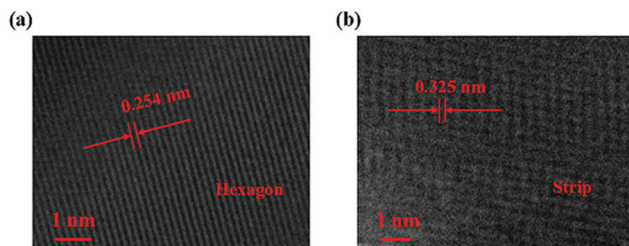


Fig. 3 (a and b) Lattice fringes of different $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite crystals characterized by TEM.

was 0.254 nm for the subnanosheet-shaped perovskite, while the distance was increased to 0.325 nm for the microwire shape. The results suggested that the inner structure of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite was modified due to the thermodynamic factor in the reaction.

Fig. 4 shows a schematic illustration of the structural transformation of the perovskites influenced by the thermodynamic and solvent effects. As depicted, the purple balls, blue balls, and yellow balls represent the iodine atoms, CH_3NH_3 groups, and lead atoms, respectively. The $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite forms a tetragonal crystal with such atoms in a temperature range from 0 °C to 120 °C.²³ In this structure, the iodine atoms located at the four sides of the tetragonal body are less stable than those on the upper and lower surfaces, resulting in easy rotation of the CH_3NH_3 group covalent bond with extra energy. In our work, the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite was dispersed in DMF solvent, and the hexagonal shape tended to more easily form at room temperature because the direction restriction for the solvent was the same as that for the perovskite. However, the structure of the perovskite was modified with the increasing temperature and with the help of the DMF solvent due to the growth of the crystal, resulting in an increase in the lattice distance from 0.254 nm to 0.325 nm. This result indicated that such a change in the structure may alter the photoelectronic properties and stability of the perovskite, and further investigation of these two $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites should be performed in a future study.

Fig. 5 shows the XRD analysis of these perovskites. The two samples were dried to a powder and subjected to incident X-rays at different diffraction angles ranging from 20° to 80°.

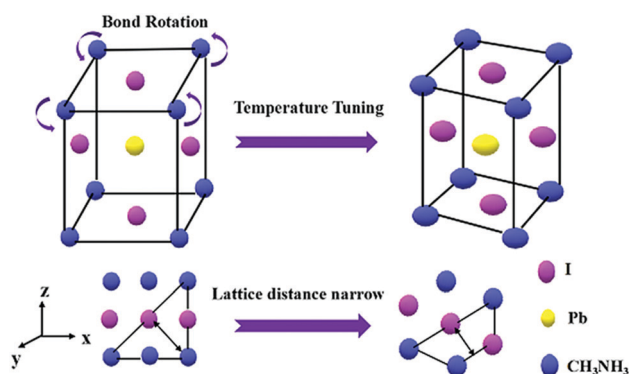


Fig. 4 Schematic illustration of the structural transformation of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite.

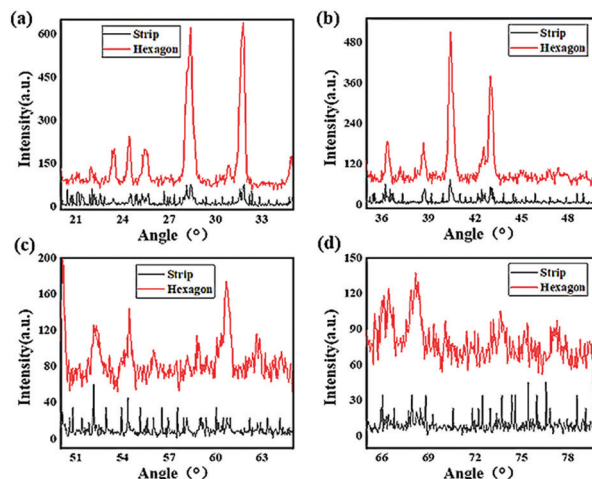


Fig. 5 XRD spectral analysis of different $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites. (a) Diffraction angle range from 20° to 35°; (b) diffraction angle range from 35° to 50°; (c) diffraction angle range from 50° to 65°; and (d) diffraction angle range from 65° to 80°.

All the characteristic diffraction peaks of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite are exhibited in the spectra, confirming the successful synthesis of $\text{CH}_3\text{NH}_3\text{PbI}_3$. The diffraction intensity of the hexagonal perovskite crystal was significantly higher than that of the strip shape crystal, which indicated that good crystallizability was achieved for the hexagonal perovskite, and the lattice plane reflection was more intense due to the occurrence of covalent bond rotation and crystal growth. In addition, rather sharp peaks could be found in the spectra of the hexagonal perovskite, which were at diffraction angles of approximately 23°, 24.5°, 26°, 28°, 31.5°, 36.5°, 39°, 41°, and 43.5°, indicating a higher regularity and symmetry of the hexagonal shape than that of the strip shape; however, evidently broader peaks were found at diffraction angles of 52.5°, 55°, 61°, 62.5°, 67°, 68.5°, 74° and 77° for the hexagonal structure, and certain diffraction peaks, for example, those at angles of approximately 52.5° and 55°, showed little change compared to those of the strip-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite. These results suggested that the increase in the lattice distance may be due to the covalent bond rotation and growth of the perovskite crystal, resulting in a variation in the diffraction of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite crystal lattice, but more details need to be researched.

A schematic illustration of XRD characterization is shown in Fig. 6. As depicted, X-ray with a certain wavelength impinges on the atoms of the perovskite crystal, and the propagation direction is changed due to the surface reflection of the atoms. Additionally, incident diffraction also occurs when the X-rays pass through the space between two atoms. However, the reflection angle is altered with the position variation of the atoms, which results in a change in the lattice distance that correlates with the diffraction angle according to Bragg's law, expressed as follows:²⁴

$$k\lambda = 2d \sin \theta, \quad (2)$$

where $k\lambda$ represents an integral multiple of the wavelength, d is the lattice distance and θ is half the diffraction angle.



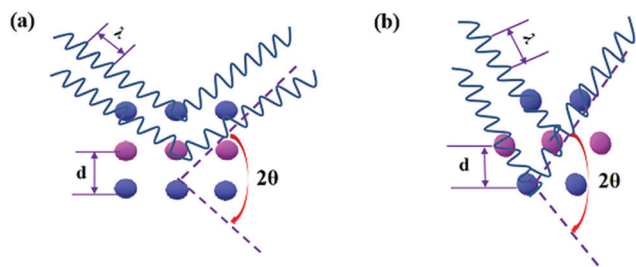


Fig. 6 Schematic illustration of XRD for the structural transformation of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite.

We assumed that the diffraction angles of sample one and sample two were θ_1 and θ_2 , respectively. The deviation of the diffraction angle was $\Delta\theta$, which was correlated with the covalent bond rotation angle of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite. Here, the value of k was regarded as 1, and the corresponding deviation value could be calculated as follows:

$$\Delta\theta = \theta_1 - \theta_2 = \arcsin \frac{0.154}{2 \times 0.254} - \arcsin \frac{0.154}{2 \times 0.325} = 3.93^\circ. \quad (3)$$

As shown, the diffraction angle decreased by 3.93° due to the increase in the lattice distance, as characterized by TEM.

To further investigate the photoelectronic difference of these two $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites, more experiments were performed using spectroscopy. Fig. 7(a) and (b) shows the emission spectra and excitation spectra of these two samples. As depicted, the two emission spectra exhibited similar shapes, and two emission peaks were evidently found in sample one and sample two when excited by an incident wavelength of 360 nm. Compared to those for the strip-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite, the two emission peaks for the hexagonal-shaped perovskite, which were at approximately 480 nm and 515 nm, were both enhanced; in addition, a redshift of 5 nm was also found for the peak located at 515 nm. However, a much

higher excitation intensity was achieved for the hexagonal $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite than for the stripe-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite, as shown in the excitation spectra, suggesting that a higher excitation efficiency was achieved with the hexagonal $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite. Fig. 7(c) and (d) shows a schematic illustration of the fluorescence emission mechanisms. The $\text{CH}_3\text{NH}_3\text{PbI}_3$ molecule jumps from the ground state to the excitation state *via* energy absorption when excited by the incident beam. Due to the rotation of the covalent bond and crystal growth in the y -axis direction for the strip-shaped perovskite, the excitation cross-section became narrowed, contributing to the lower excitation efficiency. After being in the excitation state, the $\text{CH}_3\text{NH}_3\text{PbI}_3$ molecule rapidly returns to the ground state with radiation and energy loss during emission, along with the occurrence of internal conversion or vibrational relaxation. Due to the higher symmetry and stability of the hexagonal structure, it possessed a 0.03 eV lower energy than the strip-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite, calculated according to the Planck equations as follows²⁵

$$\Delta E = E_1 - E_2 = \frac{hc}{\lambda_1} - \frac{hc}{\lambda_2} = 0.03 \text{ eV}. \quad (4)$$

The calculation result indicated that the redshift in the emission spectra around a wavelength of 515 nm further demonstrated the structural transformation of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite, resulting in a stability difference.

Fig. 8(a) shows the results of further investigation of these two samples using Raman spectroscopy excited by an incident laser with a wavelength of 632 nm. As depicted, the overall Raman scattering signal was weaker for the hexagonal-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ than for the strip-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$. In addition, redshifts of the C–C and C–N vibrations at approximately 1500 cm^{-1} and 1750 cm^{-1} , respectively, were observed in the hexagonal-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ compared to the strip-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$. The electrons around the perovskite molecules interact with photons when excited, leading to 3D atomic vibration, and polarization occurs, as shown in Fig. 8(b). In addition to the 3D vibration, the electrons also travel around the hexagonal plane that consists of atoms but not around the microwires, which could stabilize the $\text{CH}_3\text{NH}_3\text{PbI}_3$ molecule and weaken the polarization. The results suggested that lower scattering was achieved for the hexagonal-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$.

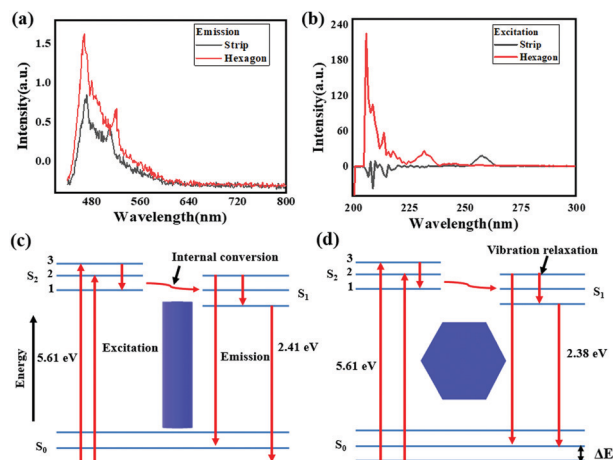


Fig. 7 (a) Emission spectra of different shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites; (b) excitation spectra of different shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites; and (c and d) schematic illustration of the fluorescence mechanisms for different shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites.

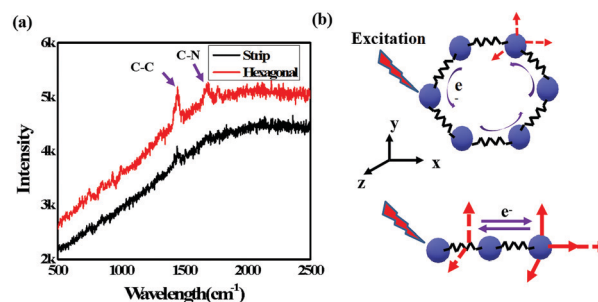


Fig. 8 (a) Raman spectra of different shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites and (b) schematic illustration of the Raman scattering for different shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites.



perovskite and further confirmed the higher stability of the hexagonal-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite due to its symmetry.

Photon capture is a critical factor for a $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite when applied to optoelectronic research, such as solar cell applications or photodetector applications.^{26–31} Fig. 9(a) shows the UV-vis spectra of the two samples. As depicted, broad absorption across the UV-vis region was found for both samples; however, the hexagonal-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ exhibited higher absorption than the strip-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$. In addition, the absorption peak originated from the transition of unbound electrons, which was located at approximately 380 nm for the strip-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$, while a redshift of ~ 20 nm was observed for the hexagonal-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$. The energy deviation after structural transformation can be calculated according to the following equation:

$$\Delta E = E_2 - E_1 = \frac{hc}{\lambda_2} - \frac{hc}{\lambda_1}$$

$$= \frac{6.63 \times 10^{-34} \times 3.0 \times 10^8}{3.8 \times 10^{-7}} - \frac{6.63 \times 10^{-34} \times 3.0 \times 10^8}{4.0 \times 10^{-7}}$$

$$= 0.16 \text{ eV.} \quad (5)$$

The calculation results suggested that the redshift of the absorption peak for the hexagonal $\text{CH}_3\text{NH}_3\text{PbI}_3$ was attributed to the electron conjugation effect and the higher symmetry of the structure, which may explain the higher fluorescence excitation efficiency. Fig. 9(b) shows the absorption measurement results of these two samples when subjected to an incident laser with a wavelength of 468 nm for further demonstration of the absorption mechanism. Here, we defined the absorption ratio α_R as follows:

$$\alpha_R = \frac{P_1}{P_2}, \quad (6)$$

where P_1 is the absorbed power and P_2 is the initial power of the incident laser. As depicted, the absorption ratio increased for both perovskites with the increasing initial power from 0.25 W to 1.15 W. However, the absorption ratios of the hexagonal-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ were all higher than those of the strip-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$, and the highest absorption increase of 50% was achieved for the hexagonal-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ for an initial incident power of 0.25 W. These results well coincided

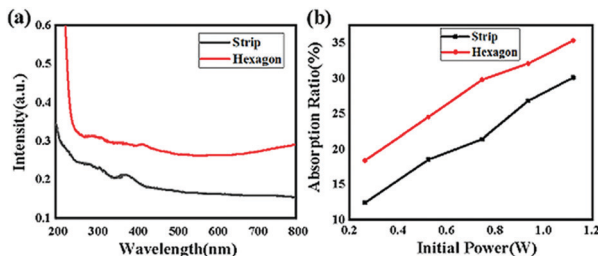


Fig. 9 (a) UV-vis spectra of two different shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites and (b) absorption measurement of two different shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites, subjected to an incident laser with a wavelength of 468 nm.

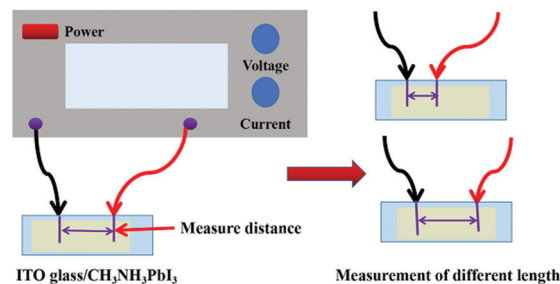


Fig. 10 Experimental scheme of resistance measurement for different $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites.

with the UV-vis spectrum characterization and further supported the above studies.

In addition to photon capture, the conductivity is another important factor for solar cell application.^{33–36} Fig. 10 shows the experimental scheme of the conductivity investigation, which was performed on the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite combined with ITO glass under different voltage excitation. As shown in Fig. 11(a), the resistance of the hexagonal-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ decreased from 14 Ω to 13.5 Ω under an applied voltage of 1 V when it transformed into a microwire shape, and the largest decrease of 2.5 Ω , from 18 Ω to 15.5 Ω , was achieved when the applied voltage was 5 V. To the best of our knowledge, the resistance of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite would increase with the increasing voltage excitation, and we attributed the reason to the higher heat energy accumulated due to the electric field, resulting in more collision of excitons and an increase in resistance. This result suggested that the structural transformation optically changed the resistivity of the $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite, resulting in better conductivity of the strip-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ compared to the hexagonal-shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$, which may be due to the strip-shaped structure being more suitable for electron migration.

Based on the above study, the resistivity, which was deeply impacted by the structure of the material, was a critical reflection of the conductivity. Regardless of this, different resistivity perovskite materials may be suitable for different applications, for example, lower resistivity and higher photon absorption perovskite materials are more suitable for solar cell applications, but higher resistivity materials would be more suitable

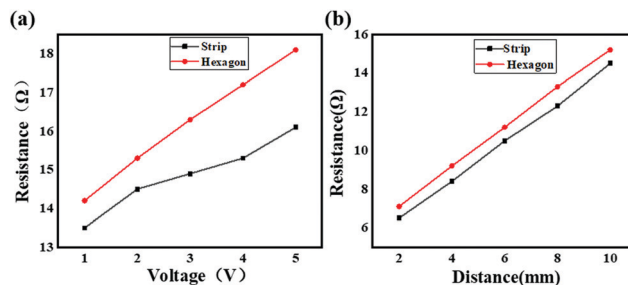


Fig. 11 (a) Resistance–voltage curves of the two different shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites on ITO glass substrates and (b) resistances of the two different shaped $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskites on ITO glass, with the length of the ITO glass ranging from 2 mm to 10 mm.



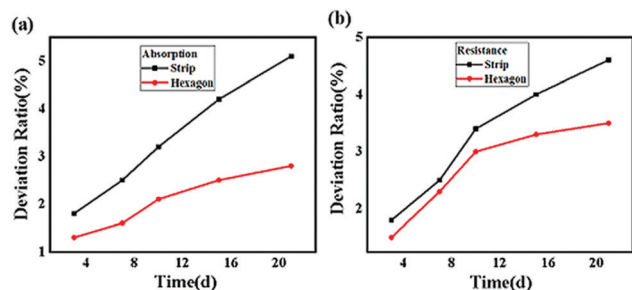


Fig. 12 (a) Absorption deviation ratio of different shaped CH₃NH₃PbI₃ perovskites on different days and (b) resistance deviation ratio of different shaped CH₃NH₃PbI₃ perovskites on different days.

for optical applications.^{37–39} To thoroughly analyze the conductivity of these two different shaped CH₃NH₃PbI₃ perovskites coated on ITO glass, the resistance of the ITO glass coated with different perovskites was tested for different lengths. The experimental scheme is shown in Fig. 10, and the results are shown in Fig. 11(b). As depicted, the resistances of both the hexagonal-shaped and strip-shaped CH₃NH₃PbI₃ crystals linearly increased with the increasing length of the ITO glass, which indicated that a well-distributed CH₃NH₃PbI₃ perovskite was coated on the ITO glass. In addition, the resistances of the hexagonal-shaped CH₃NH₃PbI₃ perovskite were all higher than those of the strip-shaped structure for lengths ranging from 2 mm to 10 mm, as shown in Fig. 11(a). This further confirmed that the resistivity of the CH₃NH₃PbI₃ perovskite can be effectively altered through structural transformation.

The stability of a perovskite material directly affects the operation life of a device when applied as an absorber in solar cells.⁴⁰ In general, the stability of a perovskite, which includes the conductivity and photon capture, decreases with the increasing storage time due to the redox reaction.⁴¹ To analyze the stability of these two samples, absorption experiments and resistance examinations were both performed. In the absorption experiment, the CH₃NH₃PbI₃ perovskite samples exposed to air for several days were subjected to an incident laser with a wavelength of 468 nm and compared with the samples not exposed to air, and the result is shown in Fig. 12(a). Here, we defined the absorption deviation σ_P expressed as follows:

$$\sigma_P = \frac{P_1 - P_2}{P_1}, \quad (7)$$

where P_1 is the absorbed power of the CH₃NH₃PbI₃ perovskite not exposed to air, and P_2 is the absorbed power of that exposed to air for several days. As depicted, the absorbed power decreased as the exposure time increased from 3 days to 21 days. However, the absorbed power deviation of the hexagonal-shaped CH₃NH₃PbI₃ could be controlled at 2.5%, while that of the strip-shaped CH₃NH₃PbI₃ was approximately 5%. Fig. 12(b) shows the resistance deviation of these two types of CH₃NH₃PbI₃ perovskite, which was defined as follows:

$$\sigma_R = \frac{R_2 - R_1}{R_1}, \quad (8)$$

where R_1 is the resistance of the CH₃NH₃PbI₃ perovskite not exposed to air and R_2 is the resistance of that exposed to air for several days. As depicted, the resistance increased as the exposure time increased from 3 days to 21 days. However, the resistance deviation of the hexagonal-shaped CH₃NH₃PbI₃ could be controlled at 3.3%, which was lower than that of the strip-shaped CH₃NH₃PbI₃ with a resistance deviation of approximately 4.8%. These results suggested that the stability of the hexagonal-shaped CH₃NH₃PbI₃ was better than that of the strip-shaped CH₃NH₃PbI₃.

3 Conclusions

In conclusion, we successfully synthesized microwire- and subnanosheet-shaped CH₃NH₃PbI₃ perovskites *via* temperature tuning and solvent effects, which were characterized by SEM and TEM, suggesting that the lattice distance increased from 0.254 nm to 0.354 nm when the hexagonal structure was transformed into a microwire structure. Further analysis showed that the structural transformation originated from the covalent bond rotation and crystal growth in the y-axis direction, and the calculation results showed that the rotation angle decreased by 3.93° after the structural transformation, which was attributed to the subnanosheet having a 0.03 eV lower energy than that of the microwire shape. Compared to the microwire-shaped CH₃NH₃PbI₃, the power absorption of the subnanosheet increased by 25% to 50% when the incident power ranged from 1.15 W to 0.25 W. This result suggested that the optical characteristics of the hexagonal-structured CH₃NH₃PbI₃ perovskite were better than those of the microwire-shaped CH₃NH₃PbI₃ perovskite. The electrical measurement results further indicated that better conductivity was achieved for the microwire-shaped CH₃NH₃PbI₃ perovskite than for the subnanosheet-shaped CH₃NH₃PbI₃ perovskite. However, the stability test showed that the absorbed power deviation of the subnanosheet CH₃NH₃PbI₃ could be controlled at 2.5% when exposed to air for 21 days, while that of the strip-shaped CH₃NH₃PbI₃ was approximately 5%. In addition, the resistance deviation of the hexagonal-shaped CH₃NH₃PbI₃ could be controlled at 3.3%, which was lower than that of the strip-shaped CH₃NH₃PbI₃ with a resistance deviation of approximately 4.8%. These results have experimentally and theoretically demonstrated the structural transformation mechanism of a CH₃NH₃PbI₃ perovskite and suggest that the hexagonal structure CH₃NH₃PbI₃ perovskite is more suitable for optical applications, whereas the microwire structure may be more suitable for electrical applications.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- J. Y. Jeng, Y. F. Chiang, M. H. Lee, S. R. Peng, T. F. Guo, P. Chen and T. C. Wen, *Adv. Mater.*, 2013, **25**, 3727.
- H. S. Kim, J. W. Lee, N. Yantara, P. P. Boix, S. A. Kulkarni, S. Mhaisalkar, M. Gratzel and N. G. Park, *Nano Lett.*, 2013, **13**, 2412.
- E. Strelcov, Q. F. Dong, T. Li, J. S. Chae, Y. C. Shao, Y. H. Deng, A. Gruverman, J. S. Huang and A. Centrone, *Sci. Adv.*, 2017, **3**, 1602165.
- T. Oku, Y. Y. Ohishi and N. Ueoka, *RSC Adv.*, 2018, **19**, 12389.
- D. Webber, C. Clegg, A. W. Mason, S. A. March, I. G. Hill and K. C. Hall, *Appl. Phys. Lett.*, 2017, **111**, 121905.
- M. Becker and M. Wark, *J. Phys. Chem. C*, 2018, **122**, 3548.
- L. H. Zhang and P. H. L. Sit, *J. Mater. Chem. A*, 2017, **5**, 9042.
- Y. Yu, Y. T. Zhang, Z. Zhang, H. T. Zhang, X. X. Song, M. X. Cao, Y. L. Che, H. T. Dai, J. B. Yang, J. L. Wang, H. Zhang and J. Q. Yao, *J. Phys. Chem. Lett.*, 2017, **8**, 445.
- J. Ma and L. W. Wang, *Nano Lett.*, 2017, **17**, 3646.
- W. D. Zhu, L. Kang, T. Yu, B. H. Lv, Y. R. Q. Wang, X. Y. Chen, X. Y. Wang, Y. Zhou and Z. G. Zou, *ACS Appl. Mater. Interfaces*, 2017, **9**, 6104.
- A. Kovalsky, L. L. Wang, G. T. Marek, C. Burda and J. S. Dyck, *J. Phys. Chem. C*, 2017, **121**, 3228.
- R. A. Kerner and B. P. Rand, *J. Phys. Chem. Lett.*, 2017, **8**, 2298.
- G. W. Ji, G. H. J. Zheng, B. Zhao, F. Song, X. N. Zhang, K. C. Shen, Y. G. Yang, Y. M. Xiong, X. Y. Gao, L. Cao and D. C. Ci, *Phys. Chem. Chem. Phys.*, 2017, **19**, 6546.
- Y. Liu, F. S. Li, C. P. Veeramalai, W. Chen, T. L. Guo, C. X. Wu and T. W. Kim, *ACS Appl. Mater. Interfaces*, 2018, **9**, 11662.
- W. G. Zhao, Z. Yao, F. Y. Yu, D. Yang and S. Z. Liu, *Adv. Sci.*, 2018, **5**, 1700131.
- L. C. Chen, K. L. Lee, W. T. Wu, C. F. Hsu, Z. L. Tseng, X. H. Sun and Y. T. Kao, *Nanoscale Res. Lett.*, 2018, **6**, 13140.
- Z. B. Yao, W. L. Wang, H. P. Shen, Y. Zhang, Q. Luo, X. W. Yin, X. Z. Dai, J. B. Li and H. Li, *Sci. Technol. Adv. Mater.*, 2017, **18**, 254.
- J. H. Heo, M. Jahandar, S. J. Moon, C. E. Song and S. H. Im, *J. Mater. Chem. C*, 2017, **5**, 2883.
- M. Afzaal, B. Salhi, A. A. Ahmed, H. M. Yates and A. S. Hakeem, *J. Mater. Chem. C*, 2017, **5**, 8366.
- L. C. Chen, Y. S. Lin, P. W. Tang, C. Y. Tai, Z. L. Tseng, J. H. Lin, S. H. Chen and H. C. Kuo, *Nanoscale*, 2017, **9**, 17802.
- W. C. Lai, K. W. Lin, T. F. Guo, P. Chen and Y. Y. Liao, *Appl. Phys. Lett.*, 2018, **112**, 071103.
- Z. Hong, J. Zhao, S. J. Li, B. C. Cheng, Y. H. Xiao and S. J. Lei, *Nanoscale*, 2019, **11**, 3360.
- J. Y. Jeng, Y. F. Chiang, M. H. Lee, S. R. Peng, T. F. Guo, P. Chen and T. C. Wen, *Adv. Mater.*, 2013, **35**, 3727.
- J. H. Im, I. H. Jang, N. Pellet, M. Gratzel and N. G. Park, *Nat. Nanotechnol.*, 2014, **9**, 927.
- Y. X. Zhao and K. Zhu, *J. Phys. Chem. C*, 2014, **118**, 9412.
- Y. Y. Dang, Y. Liu, Y. X. Sun, D. S. Yuan, X. K. Liu, W. Q. Lu, G. F. Liu, H. B. Xia and X. T. Tao, *CrystEngComm*, 2015, **17**, 665.
- W. Y. Tan, P. P. Cheng, Y. W. Zhang, J. M. Liang, X. D. Chen, Y. D. Liu and Y. Min, *J. Mater. Chem. C*, 2019, **7**, 6004.
- G. Kakavelakis, E. Gagaoudakis, K. Petridis, V. Petromichelaki, V. Binas, G. Kiriakidis and E. Kymakis, *ACS Sens.*, 2018, **3**, 135.
- P. F. Li, Y. Chen, T. S. Yang, Z. Y. Wang, H. Lin, Y. H. Xu, L. Li, H. R. Mu, B. N. Shivananju, Y. P. Zhang, Q. L. Zhang, A. L. Pan, S. J. Li, D. Y. Tang, B. H. Jia, H. Zhang and Q. L. Bao, *ACS Appl. Mater. Interfaces*, 2017, **9**, 12759.
- N. Dong, X. W. Fu, G. Lian, S. Lv, Q. L. Wang, D. L. Cui and C. P. Wong, *ACS Appl. Mater. Interfaces*, 2018, **10**, 8393.
- R. K. Singh, R. Kumar, N. H. Jain, M. T. Kuo, C. P. Upadhyaya and J. Singh, *Appl. Nanosci.*, 2019, **6**, 6078.
- M. U. Rothmann, W. Li, Y. Zhu, U. Bach, L. Spiccia, I. Etheridge and Y. B. Cheng, *Nat. Commun.*, 2017, DOI: 10.1038/ncomms14547.
- E. Horvath, M. Spina, Z. Szekrenyes, K. Kamaras, R. Gaal, D. Gachet and L. Forro, *Nano Lett.*, 2014, **14**, 6761.
- D. Y. Liu, S. S. Li, F. Bian and X. Y. Meng, *Materials*, 2018, **11**, 1141.
- W. R. Zhou, P. C. Zhou, X. Y. Lei, Z. M. Fang, Q. Liu, T. Chen, H. L. Zeng, L. M. Ding, J. Zhu, S. Y. Dai and S. F. Yang, *ACS Appl. Mater. Interfaces*, 2018, **10**, 1897.
- F. Qin, W. Meng, J. C. Fan, C. Ge, B. W. Luo, R. Ge, L. Hu, F. Y. Jiang, T. F. Liu, Y. Y. Jiang and Y. H. Zhou, *ACS Appl. Mater. Interfaces*, 2017, **9**, 26045.
- M. Becker and M. Wark, *J. Phys. Chem. C*, 2018, **122**, 3548.
- L. H. Li, J. L. Mi, Y. C. Yong, B. D. Mao and W. D. Shi, *J. Mater. Chem. C*, 2018, **6**, 8226.
- Z. Y. Wang, Q. D. Ou, Y. P. Zhang, Q. H. Zhang, H. Y. Hoh and Q. L. Bao, *ACS Appl. Mater. Interfaces*, 2018, **10**, 24258.
- M. Kumar, H. S. Kim, D. Y. Park, M. S. Jeong and J. D. Kim, *RSC Adv.*, 2018, **8**, 25206.
- L. Ciric, K. Ashby, M. Spina, M. Duchamp, B. Nafradi, M. Kollar, L. Forro and E. Horvath, *Appl. Phys. Lett.*, 2018, **112**, 111901.

