

Cite this: *RSC Adv.*, 2019, 9, 18547

Kinetically-controlled laser-synthesis of colloidal high-entropy alloy nanoparticles†

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The single-step incorporation of multiple immiscible elements into colloidal high-entropy alloy (HEA) nanoparticles has manifold technological potential, but it continues to be a challenge for state-of-the-art synthesis methods. Hence, the development of a synthesis approach by which the chemical composition and phase of colloidal HEA nanoparticles can be controlled could lead to a new pool of nanoalloys with unparalleled functionalities. Herein, this study reports the single-step synthesis of colloidal CoCrFeMnNi HEA nanoparticles with targeted equimolar stoichiometry and diameters less than 5 nm by liquid-phase, ultrashort-pulsed laser ablation of the consolidated and heat-treated micropowders of the five constituent metals. Further, the scalability of the process with an unprecedented productivity of 3 grams of colloidal HEA nanoparticles per hour is demonstrated. Electrochemical analysis reveals a unique redox behavior of the particles' surfaces in an alkaline environment and a potential for future application as a heterogeneous catalyst for the oxygen evolution reaction.

Received 1st May 2019
Accepted 29th May 2019

DOI: 10.1039/c9ra03254a

rsc.li/rsc-advances

Introduction

High-entropy alloys (HEAs) have evolved to one of the most popular material classes in the last decades in both fundamental and applied material science.^{1–3} Unique mechanical,^{4–6} electromagnetic,^{7–9} and electrochemical^{10–12} properties originate from their multi-elemental, solid-solution character. However, research has been focused primarily on bulk HEA, ignoring HEA nanoparticles (NPs), due to the absence of a reliable, scalable, and straightforward synthesis method for HEA NPs. Very recently, HEA NPs with up to eight elements, *i.e.*, Pt, Pd, Co, Ni, Fe, Cu, Au, and Sn, were produced by the carbothermal shock synthesis.¹³ The remarkable method allows uniform elemental distributions in single NPs that thereafter are kinetically controllable in size by the shock duration with a minimum diameter of 3 nm. However, carbothermal shock synthesis only produces NPs immobilized on conductive, surface-oxidized carbon support materials, and that limits the possibilities for

industrial application. In another recent approach, CoCrCuFeNi-HEA NPs with 9 nm diameter immobilized on graphene were synthesized by mechanical ball milling of single metal powders mixed with graphene.¹⁴ This solid-state synthesis is more convenient than the carbothermal shock synthesis but cannot achieve the targeted chemical distribution of the NPs. A production route for CoCrCuNiAl-HEA NPs with an average diameter of 14 nm is based on the popular sol-gel autocombustion process.¹⁵ However, no colloidal NPs have been produced by the method so far, and the chemical compositions of the NPs that were synthesized were not studied in detail. By plasma arc discharge synthesis, CoCrCuFeNi-HEA NPs of average diameters larger than 80 nm were produced in a reactive gas atmosphere.¹⁶ This method is advantageous in terms of simplicity and scalability but is bound to the gas phase and yields only large HEA NPs.

In brief, none of the above-mentioned studies describes the synthesis of colloidal HEA NPs. There are only two known synthesis approaches for isolated, colloidal HEA NPs. The first approach was published in 2015.¹⁷ A common wet chemical approach was used and the obtained NPs had large diameters (26.7 nm) with strong deviations in the chemical composition from the targeted equimolar CoCrCuFeNi. The second approach yielded in ultrasmall CoCrFeMnNi NPs with diameters ranging from 2 to 4 nm with either amorphous or crystalline structure depending on the experimental parameters.¹⁸ By sputtering material from a nearly equimolar HEA target into an ionic liquid, NPs can be synthesized instead of films usually

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/c9ra03254a

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obtained by the method. However, the required ionic liquid and the complexity of the method strongly limit its applicability. Accordingly, since the first observation of spherical nanoprecipitations in bulk HEA,¹⁹ no successful, scalable synthesis of colloidal HEA NPs with targeted elemental composition has been reported.

In this work, we demonstrate the simple, reproducible, and scalable synthesis of isolated, colloidal NPs of the equimolar CoCrFeNiMn HEA, first investigated by Cantor *et al.*,²⁰ by picosecond-pulsed laser ablation of a solid CoCrFeNiMn HEA target immersed in a flow cell. Since the first report on laser synthesis of colloids was published,²¹ the method ultimately became a competitive method for the synthesis of NP colloids.^{22–25} With this synthesis approach, we were able to obtain highly stable colloids of isolated HEA NPs without the use of any ligands for stabilization. To demonstrate the industrial applicability, we also investigated the reproducibility and scalability of the laser-based method. The investigation of the reproducibility of the approach included the evaluation of the performance of the as-synthesized NPs of the CoCrFeNiMn HEA loaded on carbon black (CB) for the oxygen evolution reaction (OER) in alkaline environment.

Results and discussion

Colloidal state of the laser ablation product

Picosecond-pulsed laser ablation of CoCrFeMnNi HEA targets in ethanol yielded brownish-colored, transparent liquids, indicating the formation of a NP colloid. Fig. 1 shows a qualitative representation of the general formation mechanism of NPs during picosecond-pulsed laser ablation of

bulk targets immersed in liquids.²⁶ In the zoom into the atomic scale, we considered the likely ablation process and NP formation for the laser ablation of a HEA ablation target. In brief, a laser beam of sufficient energy fluence irradiates the surface of a bulk HEA (Fig. 1a), leading to the formation of a plume. Later a cavitation bubble is formed (Fig. 1b) by the ablation of hot atoms, ions and clusters from the bulk HEA and their mixing with the liquid's vapor (Fig. 1b, left zoom). Then, the plume cools and the ablated species condense (Fig. 1b, right zoom). Ultimately, the growth of the condensates is quenched and they are released as NPs to the liquid during the collapse of the cavitation bubble (Fig. 1c). The HEA ablation targets were produced by mixing, pressing, and heat-treating of micropowders of the single metals. Fig. S1 (ESI†) shows a simplified sketch of the different steps in the entire procedure, starting at single metal micropowders and resulting in NPs.

Analytic disc centrifugation (ADC) analysis of the liquids synthesized by laser-ablation of HEA targets revealed the presence of NPs with a bimodal volume-weighted size distribution with peak maxima at hydrodynamic diameters of 7.7 and 87.7 nm (Fig. 2a). It remained unclear from the ADC measurement, if particles larger than 300 nm existed in the colloid. By transmission electron microscopy, no particles with Feret diameters larger than 100 nm were found (Fig. S2 and S13, ESI†). Furthermore, a size fraction smaller than the limit for reliable detection of NPs with moderate gravimetric density by ADC showed up during TEM analysis. Analytical ultracentrifugation (AUC) was applied at different rotation speeds for a statistically reliable determination of the size distribution at small diameters.

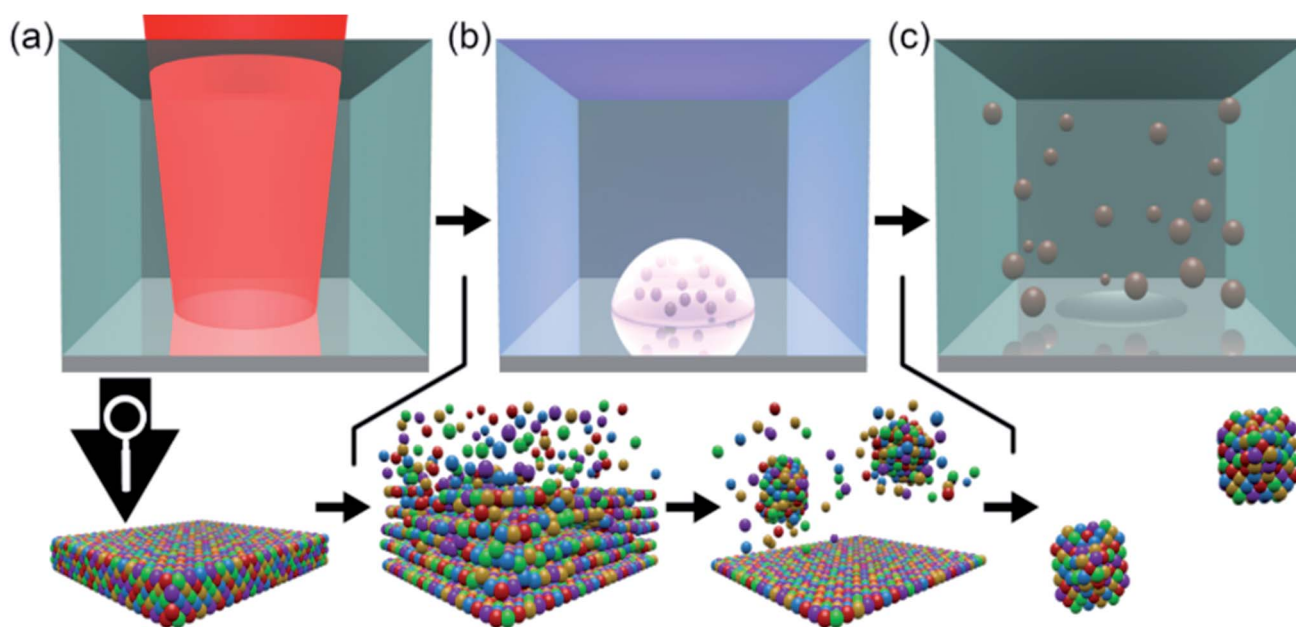


Fig. 1 Qualitative representation of the laser-based synthesis of high-entropy alloy nanoparticles. The synthesis method consisted of the following stages: ultrashort-pulsed laser irradiation of the bulk high-entropy alloy (HEA) (a), the atomization/ionization of the bulk causing the formation of a plume, and subsequent nucleation and condensation of the ablated matter in the vapor phase of the liquid (b) and the colloidal high-entropy alloy nanoparticles electrostatically stabilized in ethanol (c).



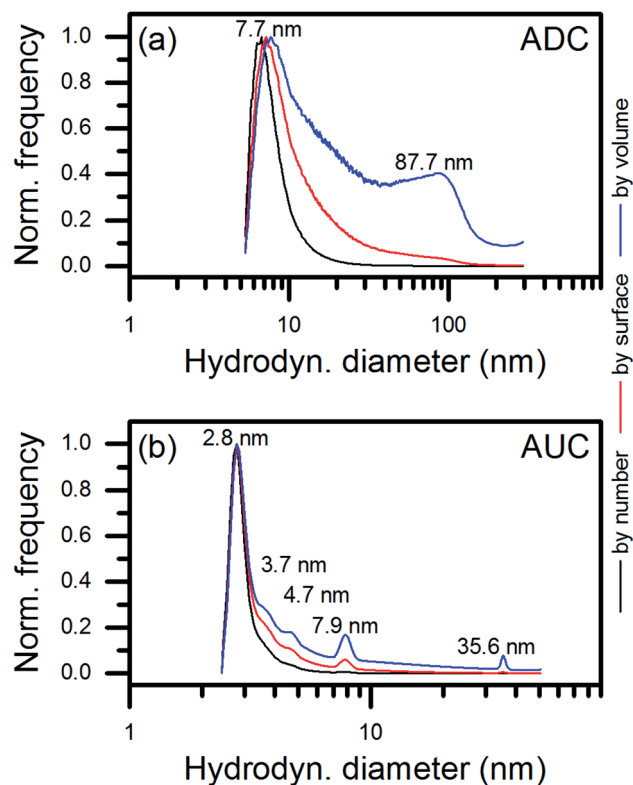


Fig. 2 Relative number, surface and volume frequencies vs. hydrodynamic diameter measured by analytical disc centrifugation (a) and analytical ultracentrifugation (b). The positions of the maxima of the different fractions found in the volume-related size distributions were marked in the diagrams.

In contrast to ADC, AUC enables the analysis of colloidal particles down to the sub-nanometer scale.²⁷ Therefore, we used this third technique of size measurement to determine whether the ultrasmall NPs observed in the TEM micrographs were present as non-agglomerated NPs in the colloid. Fig. 2b shows that most of the synthesized NPs had hydrodynamic diameters of 2.8 nm. Therefore, AUC confirmed the existence of isolated, ultrasmall NPs in the colloid. Interestingly, a second dominant size mode with an average hydrodynamic diameter of 7.9 nm was observed in the AUC measurement data. This mode is in good agreement with the main size mode detected in the ADC measurement at 7.7 nm (Fig. 2a) and the TEM analysis. An indication of a third dominant size mode at 35.6 nm appeared in the volume-related size distribution measured by AUC. This value was in good agreement with the onset of the second size mode in the ADC data. However, the size mode contributed differently to the results of both methods for unclear reasons. Conclusively, NPs of diameters smaller than 5 nm made the main product of the laser ablation of HEA targets in ethanol by number, surface, and volume. Minor amounts of particles of 5 to 10 nm and 30 to 100 nm contributed only to the surface- and the surface- and volume-related size distributions measured by AUC.

The presence of different size fractions in the colloids that were produced can be assigned to distinct particle formation

mechanisms in the picosecond-pulsed laser synthesis of colloids.²⁸ Whereas ultrasmall NPs originate from nucleation and condensation in a low-density mixing region of the ablated metal species and the liquid environment in a supercritical state, larger NPs emerge from a highly dynamic molten layer on the surface of the ablation target due to Rayleigh–Taylor instability. However, the main product of the picosecond-pulsed laser ablation consisted of ultrasmall NPs, predicted to be formed in the low-density region.²⁶ The small hydrodynamic diameters of laser-generated NPs result from fast kinetics. Cooling rates of the condensates in the order of 10^{12} K s^{-1} can be expected,²⁶ and passivation of the surface of the NPs that occurs by the interaction with solvent species suppresses further growth.^{29,79}

Crystal structure of the laser-generated nanoparticles

Next, we investigated the crystal structure of NPs by X-ray diffraction (XRD), selected area electron diffraction (SAED), and high-resolution transmission electron microscopy (HRTEM). Generally, the single metals of CoCrFeMnNi crystallize in different structures, which was demonstrated by the diffractogram of the micropowder mixture (Fig. S3a and d, ESI†). The diffractograms of the ablation target and a dried colloid shown in Fig. 3 consisted mainly of three reflections, which could be assigned to the (111), (200), and (220) planes of a single fcc lattice. The *d*-values are presented in Table 1. They slightly diverged by about +0.01 Å from the plane distances for bulk equimolar CoCrFeMnNi reported in the literature,^{30–33} which was probably in the range of the uncertainty of the method. The results confirmed the HEA nature of the ablation target and the preservation or reformation of the HEA structure in the total volume of NPs during the laser-based synthesis.

Additional peaks showed up in the diffractogram of the dried colloid, and this was in good agreement with a spinel-type structure. The extracted lattice parameter of 8.44 Å was comparable to that of magnetite, which has been reported as

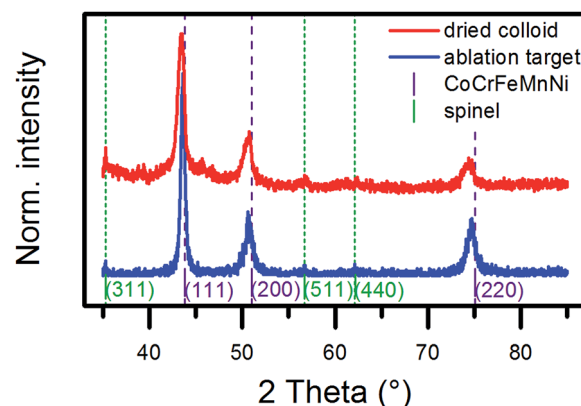


Fig. 3 X-ray diffractograms of the ablation target (blue line) and a dried colloid produced in ethanol (red line). The violet lines mark the reflections expected for a CoCrFeMnNi HEA with a lattice parameter of 3.58 Å.³⁰ The green lines mark the expected reflections for a spinel-type oxide with a lattice parameter of 8.44 Å, which is close to that of manganese chromite.³⁵



Table 1 Extracted *d*-spacing values of lattice analysis. *d*-spacing values extracted from X-ray diffractograms of the ablation target and a dried colloid produced in ethanol (Fig. 3), extracted from electron diffraction data of a nanoparticle agglomerate (Fig. 4a) and calculated values based on a lattice parameter of 3.58 Å³⁰

Plane	Calculation	Target, XRD	Particles, XRD	Particles, SAED
(111)	2.08 Å	2.08 Å	2.08 Å	2.08 Å
(200)	1.79 Å	1.80 Å	1.81 Å	1.81 Å
(220)	1.27 Å	1.27 Å	1.28 Å	1.28 Å

8.39 Å,³⁴ but it fitted best with manganese chromite.³⁵ The reflection of the (311) plane also appeared in the diffractogram of the ablation target. A reliable determination of the phase composition by a Rietveld analysis of the diffraction data of the dried colloid proved difficult because of the presence of single-crystalline NPs of different sizes in the ablation product. However, the low peak intensities of the oxide phase indicated only a minor share. The broadening in the base of the peak of the (111) reflection of the fcc lattice of the dried colloid most likely originated from a superposition of the diffractions of crystallites of different sizes. In conclusion, XRD analysis proved the HEA nature of the laser-generated NPs and strongly indicated the HEA nature of the ultrasmall NPs due to their dominant volume contribution (Fig. 2).

To verify the XRD results at the nanoscale, we applied selected area electron diffraction (SAED) at hundreds of NPs deposited on a Cu grid (Fig. 4a and b). Obviously, the NPs were crystalline, and, again, the diffraction pattern fitted the fcc lattice. Comparable lattice parameters to those determined by XRD resulted from SAED (Table 1). This finding was validated by investigations of single NPs from different size fractions. The single and larger crystalline NPs (*cf.* bright-field image in Fig. 4c) can be assigned to the fcc crystal structure, as indicated by the respective electron diffraction pattern (Fig. 4d, zone axis [101]). In addition, the fast Fourier transform (Fig. 4f) calculated from the HRTEM micrograph of a single and smaller NP (Fig. 4e) shows a corresponding [101] zone axis pattern. The periodic nature of the average structure also is supported by the HRTEM micrograph of Fig. 4g. These findings indicate the same crystal structure for NPs from the small size fraction and give clear evidence of the HEA nature of the NPs produced by pulsed laser ablation of the CoCrFeMnNi HEA target in ethanol. However, there is a fraction of large (>50 nm) NPs in the sample showing an incident of a defect structure in accordance with the XRD results (Fig. 3). The lamellar defects (Fig. 4h) of some larger NPs represented by diffuse [111]* streaks and reflection splitting in SAED (Fig. 4i) could be assigned to the formation of Σ3-twin due to relaxation similar to bulk observations.³⁶ SAED revealed a reduction of spacing of the (111) lattice, which coincided with the asymmetric shape of the (200) reflection in the XRD diffractogram towards smaller 2 theta angles (Fig. 3). Further characterization of the defect structure by HRTEM is provided in the ESI (Fig. S4).[†] Interestingly, the larger NPs, which showed a Σ3-twin formation, also exhibited surface

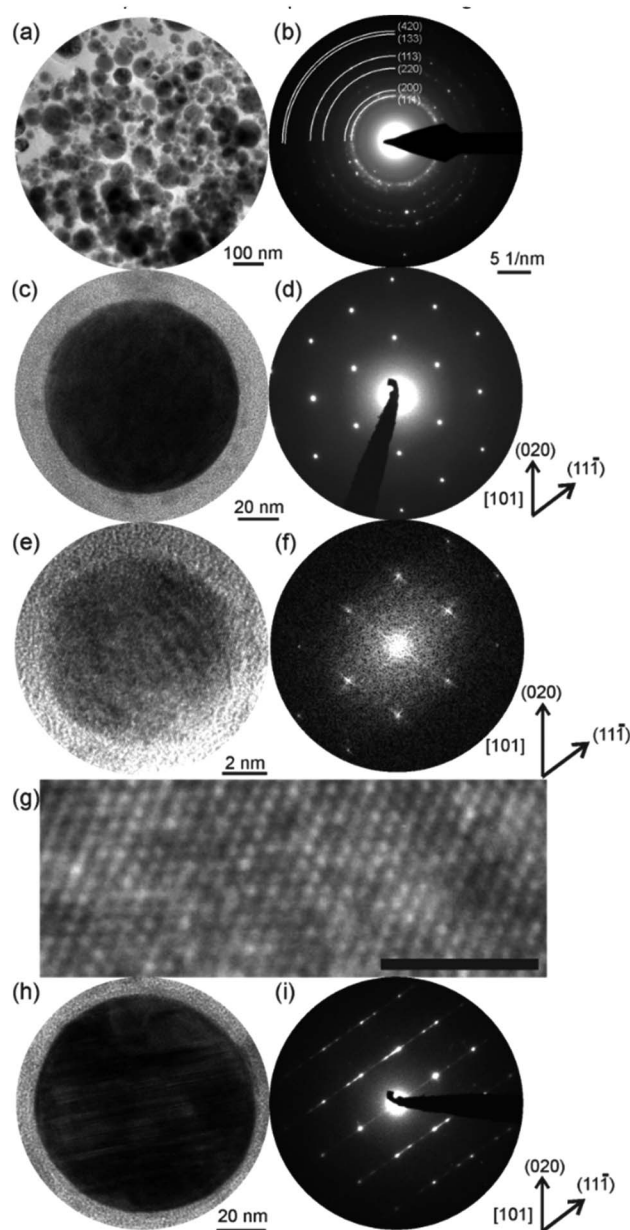


Fig. 4 Bright-field image of multiple CoCrFeMnNi HEA NPs (a) and the respective diffraction pattern (b) proving the fcc-type structure of the NPs. In (c) a single and larger crystalline nanoparticle is shown with the respective diffraction pattern in (d). The same crystal structure is determined for small nanoparticles by the high-resolution TEM micrograph (e) and the respective fast Fourier transform (f). An inset of (e) is magnified in (g) (scale bar is 1 nm). High-resolution TEM micrograph of a large disordered NP (>100 nm) (h) and the respective diffraction pattern with diffuse streaks and reflection splitting (i).

segregation of Cr and Mn (Fig. S5, ESI[†]). The surface segregation also corresponded to our observation by XRD of the Cr-, Mn-based spinel oxide phase (Fig. 3).

The size selectivity of the surface segregation of Cr and Mn, which probably occurred along with the oxidation in the spinel structure, indicated an origin in the formation of the NPs. This finding agreed with the results of the atomistic model on the



formation of NPs in the ultrashort-pulsed laser ablation of metals in liquids.^{26,28} The results of the model predict different residence times of larger NPs released from the molten surface layer into the hot, low-density mixing region of the ablated metal species and the supercritical solvent according to their kinetic energies. The movements of the laser-ablated species away from the ablation target at different velocities was also shown experimentally for the femtosecond-pulsed laser ablation of aluminum in vacuum.³⁷ A longer residence time in the low-density mixing region probably caused an intensified interaction with oxygen, which formed during the splitting of the solvent.³⁸ Subsequently, the surface segregation of Cr and Mn took place by selective oxidation in our experiments. Both metals naturally exhibit a higher oxidation potential than Co, Fe and Ni. Other larger NPs with higher kinetic energies left the low-density mixing region during their formation and cooled down faster without undergoing surface segregation. Consequently, we can conclude that surface segregation and oxidation were part of a late stage of the particle formation and only affected a small amount of larger NPs. Due to the lower heat capacity of smaller NPs their cooling may have taken place faster without surface segregation.

Chemical composition of laser-generated HEA nanoparticles

Initially, we investigated the chemical composition of the ablation target before and after the synthesis of HEA NPs by scanning electron microscopy (SEM) and energy dispersive X-ray (EDX) analysis. Fig. 5a shows a SEM image of the ablation target, depicting the border of the laser ablation area on the surface of the target. The rough surface of the target in the ablation zone on the left side of the image is clearly visible. Fig. 5b shows the exemplary EDX mapping of a $120 \times 120 \mu\text{m}$ -sized area of the ablation target for all five elements. We found a homogeneous distribution of each element. The ESI† also includes an EDX mapping of an ablation area (Fig. S6, ESI†). Table 2 provides the atomic fractions of the five single elements in the ablation area and on the non-ablated surface of the target. Each value represents the average of a mapping area of the size shown in Fig. 5b. The deviation of the elemental composition of the target, evaluated from three measurements at different positions on the surface of the target as

Table 2 Molar fractions of Co, Cr, Fe, Mn, and Ni measured in the ablated and non-ablated area of the surface of the ablation target by energy dispersive X-ray spectroscopy and the maximum deviation evaluated from three measurements at different positions of the surface of the target. Each measurement was performed as a mapping of a $120 \times 120 \mu\text{m}$ area and averaged afterwards. The average molar fractions of the single metals determined for HEA NPs are shown in comparison

	Co	Cr	Fe	Mn	Ni
Ablation	0.198	0.219	0.197	0.192	0.194
Non-ablation	0.209	0.196	0.197	0.201	0.197
Max. deviation	0.004	0.009	0.008	0.011	0.018
NPs	0.206	0.178	0.217	0.182	0.218
Std. deviation	0.018	0.009	0.005	0.001	0.006

a maximal deviation from the median, also is shown in Table 2. More SEM images and EDX results on the micropowders and ablation targets are provided in the ESI (Fig. S7–S12).† The variations in the composition of the ablated and non-ablated surface area fitted the maximum deviation between several measurements. We conclude that no selective ablation of single elements from the target took place, which ensured that equal molar amounts of each of the elements became part of the colloid.

Investigation of the elemental distribution within a single NP took place by EDX mapping. Fig. 5c demonstrates a homogeneous distribution of the five elements within a single NP. Despite some technical challenges and low counts on single NPs, the quantitative composition in Table 2 depicts the average of multiple NPs. The agglomerate of HEA NPs had a nearly equimolar composition.

A homogeneous chemical composition of single ultrasmall NPs only can be assumed based on the analysis of the crystal structure that was presented, because the EDX scanning area cannot be restricted to such small areas. In case of strong deviations from the chemical composition of the ablation target, a significant change in the phase composition, or at least the lattice parameter, should have occurred in the NPs. This was observed neither by SAED (Fig. 4) nor by XRD (Fig. 3) as summarized in Table 1, even though the ultrasmall NPs dominated the product by volume (Fig. 2). In addition, significant deviations from the equimolar composition only were observed within single larger NPs.

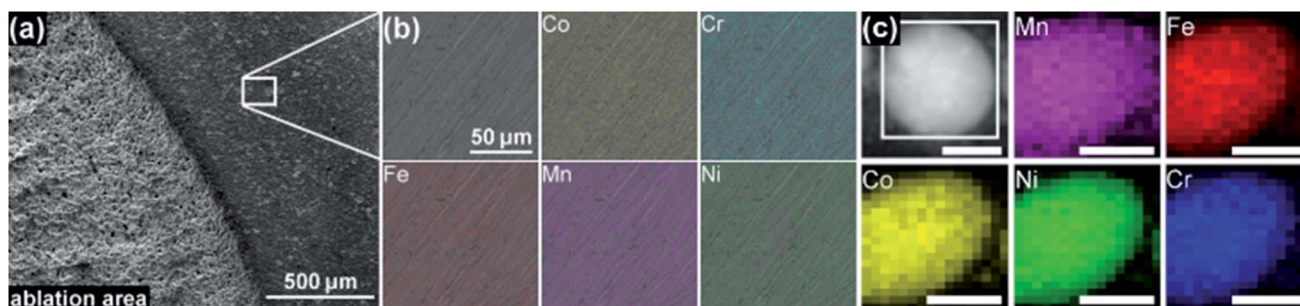


Fig. 5 Scanning electron microscope image (secondary electrons) of an ablation target near the border of an ablation area (a), elemental maps of Co, Cr, Fe, Mn, and Ni obtained by energy dispersive X-ray spectroscopy of a non-ablated target area (b), and scanning transmission electron microscopy image and elemental maps of Co, Cr, Fe, Mn, and Ni for a single nanoparticle (c). In (a), the laser ablation area is located on the left side of the image. In (c), all scale bars represent 25 nm.



Investigation of the relevance of the synthesis approach for industrial application in terms of scalability, reproducibility, and demand

In order to emphasize the relevance of the synthesis approach to the application of the laser-based synthesis of HEA NP colloids, we modified the synthesis process for use with a high-power laser system. (See also the description in the Experimental section.) As an initial step in modification, we replaced ethanol as the ablation liquid with water. This mainly was due to the higher rates of ablation using water instead of ethanol.^{38,39} In addition, water is both economically and ecologically advantageous compared to ethanol.

The XRD data of the NPs synthesized in different solvents indicated slight differences in the crystal structure (Fig. 6a). The HEA-related diffraction peaks of the NPs produced in water showed a shift in the reflection angle. In addition, relatively higher intensities of diffractions related to the before mentioned spinel oxide phase, especially for the (400) plane, occurred. However, it remained unclear from the results whether the amount of phase-segregated, larger NPs increased or smaller NPs underwent phase segregation during the laser-synthesis in water. The further adaption of the water-based synthesis to the high-power laser system, adapted from Streubel *et al.*,⁴⁰ yielded a mass ablation rate of 3 g h^{-1} , which was 50 times higher than the mass ablation rate using ethanol in the setup at the low-power laser system. As demonstrated by the XRD data, similar amounts of oxide showed up for NPs synthesized in water at both low and high laser power. A comparison of the size distribution of the NPs produced in ethanol at the low-power laser and those produced in the scale-up by TEM (Fig. S2, ESI†) exhibited a slight shift of the ultra-small size fraction of the NPs to larger diameters for the scaled-up NP colloids. This change probably originated from the larger extent of the oxidation of the surface of the NPs synthesized in water, which was also indicated by low-contrast shells around the particles in the TEM micrographs (Fig. S13, ESI†).

We investigated further the effect of the upscaling process on the performance of the NPs in a chosen application. Laser-generated NP colloids exhibit a high potential for various applications.²⁴ The NPs can be immobilized easily on conductive^{41–48} and non-conductive^{49–54} substrates and be used, for instance, in additive manufacturing,^{48,51} as sensors,^{41,54} or in energy conversion and storage.^{43,45} For colloids based on polar solvents, an optimization of the deposition process can be achieved by adjusting the static charge on the surfaces of the NPs and the substrate,⁵² or by applying an electric field between the substrate and a counter electrode.^{44,49} The ligand-free surfaces of the laser-generated NPs enable their effective adsorption on substrates.^{50,55} In addition, the surface charges of the laser-generated NPs cause their unique electrostatic stabilization properties.⁵⁶ The absence of stabilizing ligands promotes, in particular, the application of laser-generated NPs in heterogeneous catalysis.⁵⁷ Consequently, we prepared heterogeneous catalysts by immobilization of the laser-generated HEA NPs on CB (Fig. S14, ESI†). Due to the relevance of alloys and mixed oxides of the transition metals of Co,

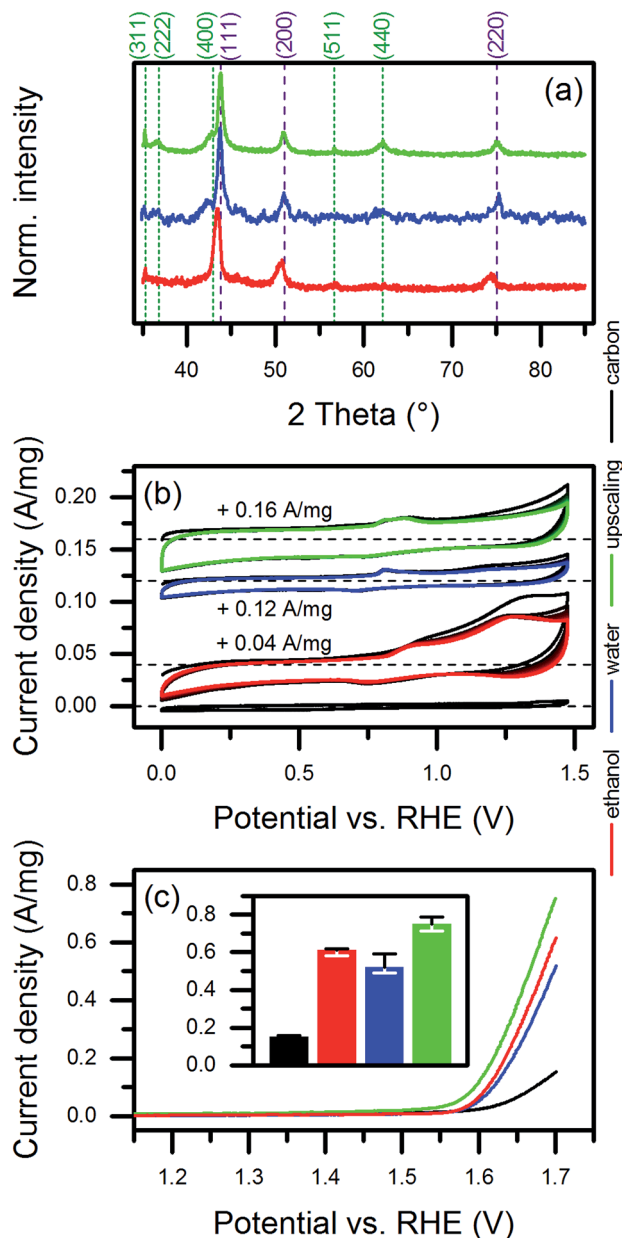


Fig. 6 X-ray diffraction analysis of dried colloids from the laser ablation in ethanol, in water, and from the upscaling approach (a), and electrochemical characterization of the nanoparticles immobilized on carbon black by cyclic (b) and linear sweep (c) voltammetry. The lattice planes marked in (a) correspond to diffractions expected for CoCr-FeMnNi with a lattice parameter of $3.58 \text{ \AA}$ (violet lines) and to diffractions expected for a spinel-type oxide with a lattice parameter of $8.44 \text{ \AA}$ (green lines). The inset in (c) shows the current densities extracted at 1.7 V vs. RHE. The error bars in (c) represent the maximum deviation evaluated from three measurements.

Fe, Mn and Ni for electrochemical water splitting,^{58–60} we tested the potential of the HEA for alkaline OER.

As shown in Fig. 6b, the catalysts produced from NPs synthesized in ethanol, water, and the upscaling approach responded differently in cyclic voltammetry, suggesting a change in the composition of the surface (discussed further below). The initial decrease in anodic currents observed for all



catalysts probably originated from the removal of impurities during cycling. A significantly higher decrease in the currents measured for the ethanol-based catalyst most likely was linked to organic leftovers from ethanol or laser-decomposition products of ethanol. Interestingly, the dominance and positions of the characteristic redox features varied slightly for the different catalysts and only two to three redox features appeared for each catalyst. The presence of five metals in the NPs complicated the assignment of contributions to the observed currents. For instance, a NiOOH/Ni(OH)₂ transition, which likely occurs at a voltage higher than 1.25 V vs. RHE at pH 13, can shift to an even higher voltage in case of alloying Ni with Fe.⁶¹ Alloying Ni with Co causes the opposite effect.⁶² Li *et al.*⁶³ described effects of exchanging the alloy partner in Co-based, binary NPs by Fe, Mn, and Ni on the redox behavior of the particles. Considering the observed surface segregation of Cr, the Cr(III)/Cr(VI) transition is another candidate that would fit at least the oxidation feature at higher voltages.⁶⁴ Due to the solubility of Cr(VI), probably no reduction peak would be assignable. However, in that case, the oxidation peak also should decrease during cycling, which we did not observe. Mn accompanied Cr in the surface segregation. Some redox transitions of Mn, *i.e.*, Mn₃O₄/MnOOH and MnOOH/MnO₂, could have lied in the cycled potential range.⁶⁵ Due to the close positions of the oxidation and reduction waves of the Mn reaction, the two expected peaks for the oxidation and the reduction reactions, respectively, may have overlapped in the voltammogram. Moreover, we were unable to detect any features close to the expected voltages for Fe-related reactions.^{64,66}

In conclusion, unique electrochemical properties of the surfaces of the catalysts appeared. Synergetic effects caused by the presence of different metals contributed to the redox features. In addition, the solvent that was used during the synthesis and the used laser influenced the potentials of the redox reactions. Further analysis of the particle surfaces, such as X-ray photoelectron spectroscopy, should shed light on the elemental compositions of the surfaces in future studies.

We extracted the current densities at 1.7 V vs. RHE from the linear sweep voltammograms to investigate the suitability for OER catalysis (Fig. 6c). The currents were related to the mass of NPs (1.96 µg) used to produce the amount of catalyst deposited on the glassy carbon electrode for each electrochemical measurement. In case of pure CB, the same mass was used. Interestingly, no significant difference showed up for NPs synthesized in ethanol or water. However, the NPs produced in the upscaling approach performed best, which was probably caused by the slightly different redox behavior of their surfaces. A comparison of the achieved mass activities to literature data (Fig. 7a) revealed a good performance of our catalysts. We considered only studies, which conducted the electrochemical characterization similar to ours. Basu⁶⁷ synthesized Co₃O₄ catalysts immobilized on differently prepared, carbon-based supports and achieved a maximum mass activity of 0.077 A mg⁻¹ at 1.7 V vs. RHE. The value was significantly lower than the 0.104 A mg⁻¹ reached in average for the water-based HEA NPs on CB. However, the author used a lower scan rate of 5 mV s⁻¹ and a stronger electrolyte (1 M NaOH) compared to our study

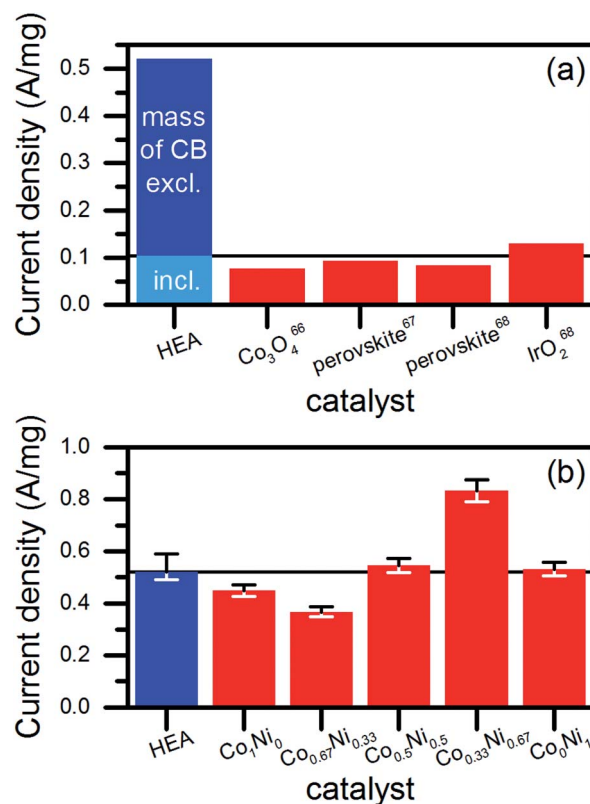


Fig. 7 Comparison of the mass-related current density or mass effectivity of HEA NPs (water-based synthesis at low-power laser) to literature data (a) and to catalysts based on NPs from the laser ablation of Co_xNi_{1-x} targets, which were produced the exact same way as the HEA NPs. The black line in both diagrams marks the mass activity of the HEA catalyst for alkaline OER. The error bars in (b) represent the maximum deviation evaluated from three measurements in case of the HEA sample and the systematic error of the electrochemical measurement of 5% in case of the other samples.

(20 mV s⁻¹, 0.1 M NaOH). In two studies on multi-element perovskite catalysts,^{68,69} the authors determined maximal mass activities of 0.093 (ref. 68) and 0.085 A mg⁻¹ (ref. 69) at a potential close to 1.7 V vs. RHE. Interestingly, Xu *et al.*⁶⁹ measured also the activity of a commercial IrO₂ catalyst as 0.13 A mg⁻¹, which was 25% higher than the HEA NP synthesized in water at the low-power laser but 13% lower compared to those synthesized at the high-power system.

However, such comparisons always lack comparability due to differences in the preparation of the electrodes and the measurement procedure. Therefore, we synthesized NPs by the laser ablation of binary alloy targets of Co and Ni in water (Fig. S15, ESI†). Mixed materials of Co and Ni represent efficient OER catalysts.^{70,71} The production of the ablation targets, the NP synthesis and the electrochemical characterization took place at identical parameters compared to the HEA. As shown in Fig. 7b, the HEA NPs achieved similar mass activities than the Co_xNi_{1-x} NPs. Exceptions appeared for the Co-rich and Ni-rich compositions, which showed significantly lower (30%) and higher (60%) mass activities related to the activity of the HEA NPs,



respectively. This dependency of the catalytic performance on the ratio of Co to Ni agreed with the literature.^{72,73}

For the investigation of the reproducibility of the NP synthesis, we varied different process parameters within the ablation of CoCrFeMnNi HEA targets in ethanol using the low-power laser system and within the target and catalyst preparation. We monitored the effect on the light extinction properties of the colloids (Fig. S16–S20, ESI†) and their catalytic performance represented by the current density at 1.7 V vs. RHE (Fig. S21 and S22, ESI†). From the wavelength-dependent extinction of light, we extracted information on the productivity⁷⁴ and the size distribution⁷⁵ of the NP colloid. The parameters that were varied included the duration of the heat-treatment of the ablation targets (Fig. S18, ESI†), the distance between the focal plane and the target surface during the synthesis (Fig. S19, ESI†), and the repeated use of a single ablation target for the synthesis of NP colloids (Fig. S20, ESI†). We observed that the laser synthesis of NP colloids from ablation targets of the CoCrFeMnNi HEA was highly reproducible (with variations of less than 10%) in terms of the productivity and size distribution.

In using the catalytic performance of the NPs as the figure of merit for the synthesis process, the highest deviation of 12% was observed for catalysts produced from three different targets heat-treated for 20 h (Fig. S21b, ESI†) and from three targets heat-treated for different durations (Fig. S21c, ESI†). We conclude that the preparation of the ablation targets, and most likely the mixing step of the metal micropowders, introduced the highest deviation in the electrochemical performance of the catalysts. However, the deviation was still in the range of acceptance for industrial application.

The demonstrated procedure for the laser-based synthesis of HEA NPs proved scalability and reproducibility. Moreover, a brief outlook on an application of the NPs as heterogeneous catalysts for OER indicated an existing potential. However, NPs produced by the laser ablation of binary alloys of Co and Ni revealed comparable and in one case even higher mass activities. Consequently, we will focus on variations of the composition of the CoCrFeMnNi alloy to identify the most active one in future studies. In addition, the quantitative and structural analysis with relation to the NP surface needs to be extended to investigate the origin of the catalytic performance.

Experimental

Pulsed laser ablation in liquid

Laser-based synthesis of HEA NPs was performed by using a 10 ps pulsed Nd:YAG laser (Atlantic, Ekspla, Vilnius, Lithuania) at a wavelength of 1064 nm, a repetition rate of 100 kHz, and a power of 8.8 W behind all optics. We used a thermopile sensor (PowerMax PM30, Coherent, Santa Clara, USA) to measure the power of the laser. After leaving the laser source, the laser beam with a raw beam diameter of about 2 mm (manufacturer's specification) enters a galvanometer scanner (SCANcube 10, SCANLAB, Puchheim, Germany) and afterwards passes through an f-theta lens with a focal length of 100 mm that focuses the beam on an HEA sheet mounted in a flow chamber. A constant

volumetric flow of 20 mL min^{−1} was applied to the ablation fluid, *i.e.*, ethanol (99.8%, Thermo Fisher Scientific, Waltham, USA) streaming through the ablation chamber.

A demonstration of the scalability of the synthesis method was performed using a 3 ps pulsed high-power laser system (500 flex, AMPHOS, Herzogenrath, Germany) in steady-state conditions of a liquid flow of 500 mL min^{−1}.⁴⁰ The wavelength of the laser was 1030 nm and the pulse repetition rate was 10 MHz. Spatial separation of single laser pulses on the ablation target was achieved by applying a scanning speed of 484 m s^{−1}. The optimal working distance for the ablation was adjusted by determining the maximized ablation rate *via* UV-vis extinction measurements (Evolution 201, Thermo Fisher Scientific, Waltham, USA) of colloids produced at different working distances. The ablation at the high-power laser system was performed in water (Milli-Q, Merck, Darmstadt, Germany).

Preparation of the ablation targets

Ablation targets of the CoCrFeNiMn HEA were produced by weighting, homogenizing, pressing, and heat-treating (in an argon atmosphere) of micropowders of the single metals. All powders were purchased from Alfa Aesar (Haverhill, USA). Purities and grain sizes are as follows (supplier specifications): Co (99.8%, 1.6 μm), Cr (99.2%, <10 μm), Fe (99.5%, <10 μm), Mn (99.6%, <10 μm), Ni (99.9%, 3–7 μm). To achieve an equimolar composition, the powders were weighed in the mass ratio 1 : 0.88 : 0.95 : 0.93 : 1, following the preceding order. The powders were homogenized by a pestle in a mortar for 20 min before pressing the powder mixture at 200 MPa to sheets of 20 × 20 × 2 mm. The sheets were heat-treated in an argon atmosphere for 20 h at 1000 °C to trigger the formation of the fcc HEA atomic lattice by diffusion. A thin oxide layer on the heat-treated sheets, which likely was caused by oxygen from the oxidized surfaces of metal micrograins, was removed by sanding. The ablation targets were mechanically processed to fit in the ablation chamber.

Preparation and electrochemical characterization of the catalysts

To prepare heterogeneous catalysts from the laser-generated NPs, colloids were mixed with a dispersion of carbon black (VULCAN XC72R, Cabot Corporation, Boston, USA) in equal volumes. CB was dispersed in the same liquid that was used for the laser synthesis of the NP colloids. To achieve a specific mass loading of NPs on carbon black, the mass loss of the ablation target during the synthesis of the NPs was determined, and CB was weighed accordingly. Mixing of the dispersions occurred 1 h after the synthesis of the colloids of the NPs. The mixture of the two dispersions was stored for 20 min in an ultrasonic bath. Afterward, the liquid was evaporated to obtain a dry catalyst powder.

The electrochemical investigation of the catalysts was performed by utilizing a three-electrode setup at a potentiostat (VersaSTAT F3, Princeton Applied Research, Ametek, Berwyn, USA). All measurements were performed in 0.1 M NaOH (pH = 13) by using a Pt counter electrode and an Ag/AgCl reference



electrode. All potentials were corrected with respect to reversible hydrogen electrode (RHE). Inks of the HEA/CB catalyst were prepared by dispersing HEA/CB (4.9 mg) in a 1 : 1 (by volume) mixture (5 mL) of Milli-Q water and isopropyl alcohol (99.6%, Thermo Fisher Scientific, Waltham, USA) with 5 wt%-Nafion ionomer solution (40 μL , 15–20% water, Sigma Aldrich, St. Louis, USA) with subsequent sonication. Then, a volume (10 μL) of a catalyst ink was drop casted on a glassy carbon electrode and dried for 15 min by rotating at 700 rpm. Before each measurement, the electrolyte was degassed for 30 min with argon. Blank cyclic voltammetry was performed at a sweep of 100 mV s^{-1} (without any rotation of the electrode). All linear sweep measurements were conducted at a sweep of 20 mV s^{-1} , and current densities were extracted at 1.7 V (electrode rotation of 1600 rpm) for sample comparison. All currents were related to the mass of NPs (1.96 μg) deposited on the glassy carbon electrode (0.196 cm^2).

Analytical ultracentrifugation (AUC)

A precise, high-resolution analysis of the size distribution of the NPs was conducted by AUC of the colloids based on ethanol. The ultracentrifuge (ProteomeLab XL-I, Beckmann Coulter, Brea, USA) was used at speeds of 5000; 15 000 and 20 000 rpm for the colloid analysis. The different speeds allowed a detailed investigation of the hydrodynamic diameters of the NPs in the range of 1 to 50 nm. The used rotor type was an AN-60 Ti with 12 mm path length and a radius of 65 mm. Sedimentation velocities were detected in 200 scans at 20 °C. The absorbance was detected at 400 nm.

Analytical disc centrifugation (ADC)

ADC measurement was used for a statistically reliable determination of the sizes of the HEA NPs produced in ethanol and water using the low-power laser system. Measurements were performed using a disc centrifuge (DC 24000, CPS instruments, Prairieville, USA) using the maximum speed of 24 000 rpm and a water-based sucrose gradient. The ethanol-based samples were mixed with water in the ratio of 1 : 5 to ensure a proper measurement.

Scanning electron microscopy (SEM) and energy dispersive X-ray spectroscopy (EDX)

A scanning electron microscope (XL 30, Philips, Amsterdam, Netherlands) equipped with a Si(Li)-EDX-detector was used for the investigation of the uniformity of the chemical composition of the ablation targets.

X-ray diffraction analysis (XRD)

XRD analysis was performed on the rasped ablation targets and the dried colloids using a diffractometer (Empyrean, PANalytical, Almelo, Netherlands) with Bragg–Brentano geometry and Cu K α radiation.

Transmission electron microscopy (TEM)

Analyses were performed of the morphologies and sizes of the HEA NPs produced by the low- and high-power laser system and the catalysts using a transmission electron microscope (CM12 Philips, Amsterdam, Netherlands). All samples were drop casted on a carbon-coated Cu mesh. High-resolution transmission electron microscopy (HRTEM) and EDX investigations were performed with a Tecnai F30 STwin G2 300 kV acceleration voltage and a Si(Li) detector (EDAX system for elemental maps). Scanning TEM Z-contrast images were recorded using a high-angle, annular, dark-field detector, and a Philips CM30 ST microscope (LaB6 cathode, 300 kV) was used for the HRTEM and selected area electron diffraction (SAED) studies.

UV-vis extinction spectroscopy

The spectroscopic analysis was performed on colloids filled in a quartz glass cuvette (10 mm beam way). Each extinction spectrum was recorded at an extinction-calibrated spectrometer (Evolution 201, Thermo Fisher Scientific, Waltham, USA) in the range of 190 to 1100 nm *versus* a background spectrum of pure ethanol.

Conclusions

In summary, the ultrafast laser synthesis of colloids proved feasible to synthesize colloidal NPs of CoCrFeMnNi HEA with a targeted chemical composition. An fcc lattice fitting the equimolar CoCrFeMnNi HEA was determined by XRD, SAED, and in HRTEM imaging for NPs of different diameters. In addition, on average, a nearly equimolar elemental composition was determined for the HEA NPs. Furthermore, we verified the colloidal state of ultrasmall (<5 nm) HEA NPs by using centrifugal size measurement techniques. The isolated, colloidal nature of our HEA NPs in the absence of stabilizers enabled the immobilization of laser-synthesized NPs on different substrates. An exemplary application in heterogeneous catalysis of the alkaline OER by immobilization of the HEA NPs on CB was performed successfully in this study. The remarkable stability of the laser-synthesis approach was verified based on the low deviations in the catalytic performance of the HEA NPs that were produced. Further, we provided evidence of the industrial applicability by scaling up of the synthesis process and producing a yield of 3 g h $^{-1}$. The catalytic performance of the synthesized HEA NPs is promising, but a comparison to laser-generated CoNi NPs reveals the potential for further optimization by variation of the alloy composition.

Prospectively, the capability for rapid production of different functional materials by the variation of the alloy composition of the ablation target,⁷⁶ by the combination of the target material and the liquid environment,⁷⁷ and by exchanging the immobilization substrate, as discussed earlier, make the laser-based synthesis superior for screening nanomaterials. In particular, varying the amount of single elements in multinary, non-noble metal alloy NPs holds the chance of finding a specific composition, which may, for instance, in future challenge platinum as



fuel cell catalyst.⁷⁸ The laser-based synthesis may cover both the material screening and technical scale production.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

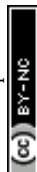
The authors gratefully acknowledge the German Federal Ministry of Education and Research (03SF0497B) and the German Research Foundation (GO 2566/2-1, BA 3580/18-1 and KI 1263/15-1) for funding. E. B. gratefully acknowledges the support of the Alexander von Humboldt Foundation for a post-doctoral fellowship.

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