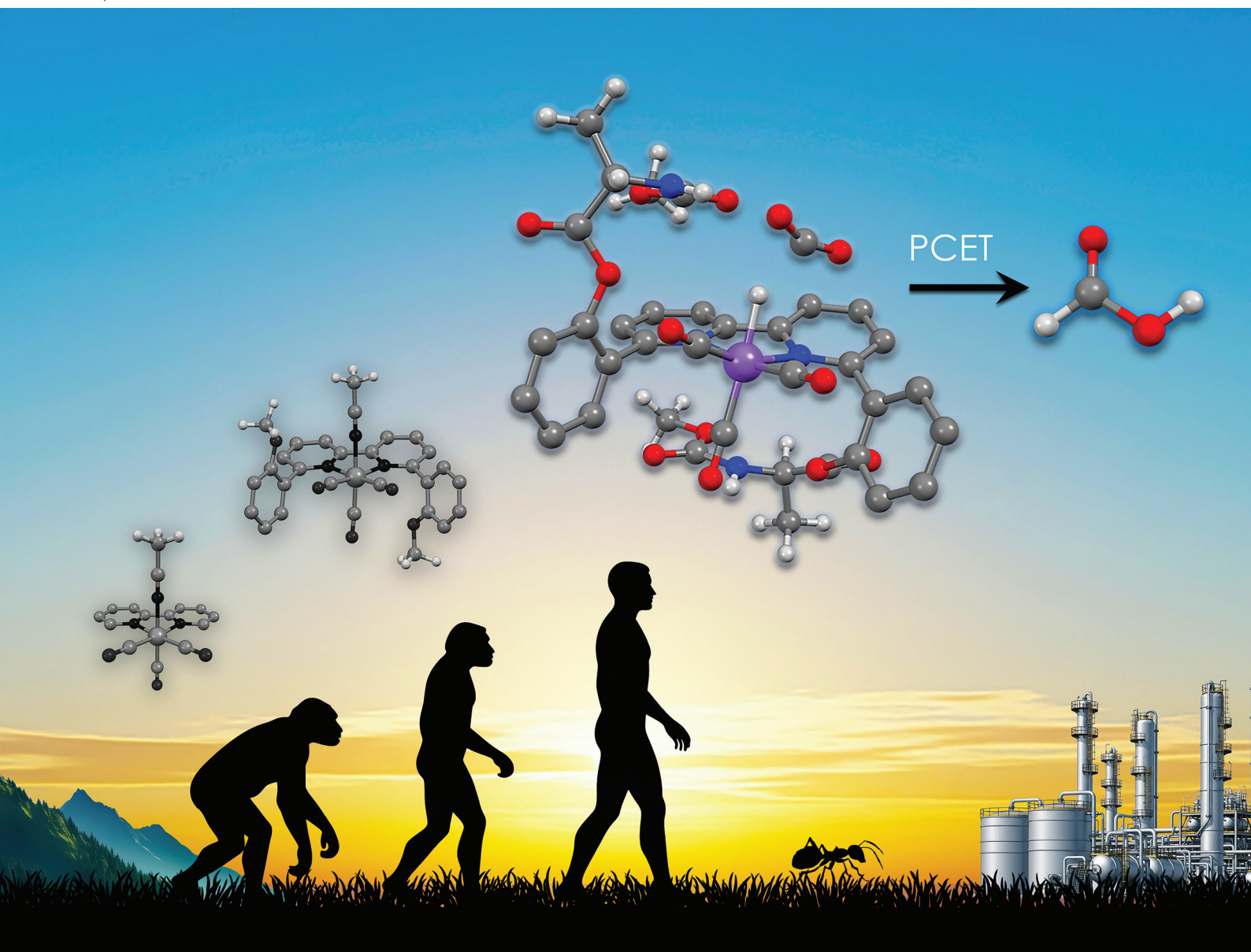


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The impact of second coordination sphere functional group extension on product selectivity for manganese bipyridyl CO₂ reduction electrocatalysts†

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Utilizing the well established manganese bipyridyl class of homogeneous electrocatalyst, four new ligands are studied to probe the influence of distal, outer coordination sphere, H-bonding and steric effects on product selectivity for proton-coupled electrocatalytic CO₂ reduction. The presence of a simple acetate functional group in the second coordination sphere provides a high selectivity for CO₂-to-CO conversion irrespective of proton source (H₂O vs. PhOH) or applied potential. The *o*-(methoxybenzoate)phenyl second/outer coordination sphere at the bipyridyl 6,6'-positions imparts poor product selectivity. In contrast, upon conjugation of the acetate functional group with the *N*-Boc-alanine moiety, a CO : HCO₂⁻ product selectivity of ~1 : 1 is observed at the high overpotential catalytic wave (for both H₂O and PhOH acids). Computed enthalpy and free energy of activation parameters suggest that selective CO₂ insertion at the manganese hydride transition state is favored, over protonation, consistent with negligible hydrogen production during controlled potential electrolysis studies.

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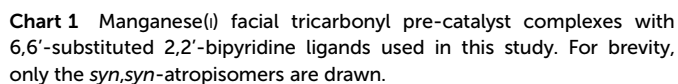
Introduction

The [Ni-Fe]CODH enzyme has inspired many studies of molecular transition metal catalyst designs for electrocatalytic CO₂ reduction,^{1–3} in particular toward the role of secondary and outer coordination sphere effects on enhancing catalytic performance.^{4–7} By extending the second coordination sphere functionality one can potentially mimic the qualities of enzymatic co-factors and allosteric effects found in nature by including extended hydrogen bonding (H-bonding) networks, and through-space electrostatic effects. This strategy has seen successful application in electrocatalytic proton,^{8–10} CO₂^{11–24} and CO₂²⁵ reduction as well as CO₂ hydrogenation^{26–32} catalysis. One of the first studies highlighting the role of the outer coordination sphere was for the purpose of electrocatalytic H₂ oxidation where [Ni^{II}(P^{Cy}₂N^{Gly}₂)₂]²⁺ catalysts (where P₂N₂ = 1,5-diaza-3,7-diphosphacyclooctane) were functionalized with glycine. This was found to facilitate proton transfer from the

outer carboxylate groups to the pendant amino groups in the second coordination sphere, enabling electrocatalytic H₂ oxidation over a wide range of pH values (0.1–9.0).³³ The outer coordination sphere of a homogenous transition metal complex is understandably less well defined than the inner and second coordination spheres. An outer coordination sphere may take the form of bulk solvent and electrolyte and any co-factors dissolved therein, each of which can potentially participate with the inner and second coordination sphere of a transition metal complex along its reaction coordinate.^{19,22,34,35} For example, the 1-ethyl-3-methylimidazolium tetracyanoborate, [EMIm][TCB], ionic liquid electrolyte was found to increase the rate of CO₂ to CO reduction for the [*fac*-ReCl(2,2'-bipyridine)(CO)₃] pre-catalyst at a reduced overpotential through an outer sphere [EMIm]⁺bpy⁻ π -stacking interaction.¹² Gotico *et al.* later demonstrated a similar rate enhancement upon covalently incorporating a urea second coordination sphere at an Fe(III)Cl *meso*-tetraphenylporphyrin pre-catalyst.¹⁷ This was later attributed to multi-point hydrogen-bond templating of the bicarbonate anion in the second coordination sphere at the rate-determining C–OH bond-cleavage transition state.³⁶ Also noteworthy is how Liu and McCrory encapsulated the cobalt phthalocyanine (CoPc) complex within a hydrophobic poly-4-vinylpyridine (P4VP) membrane thereby increasing CO product selectivity through optimized proton delivery in the outer coordination

^aDepartment of Chemistry, University of Massachusetts Boston, 100 Morrissey Boulevard, Boston, MA 02125, USA. E-mail: jonathan.rochford@umb.edu^bChemistry Division, Brookhaven National Laboratory, Upton, New York 11973-5000, USA†Electronic supplementary information (ESI) available: Synthesis, HRMS and NMR data, additional voltammetry and IRSEC data, computed geometries. See DOI: <https://doi.org/10.1039/d5dt00157a>

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Scheme 1 Summary of the reaction sequence employed for ligand synthesis.

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confirming an ECE activation mechanism (eqn (1)–(3)). Unfortunately, the quality of IRSEC spectra was poor due to reduced solubility of all reduction products for each pre-catalyst. For this reason, identification of $[3\text{-CH}_3\text{CN}]^0$ and $[4\text{-CH}_3\text{CN}]^0$ was not feasible. Exemplar IRSEC spectra are provided in Fig. 3 for the *N*-boc-ala functionalized pre-catalyst $[2\text{-CH}_3\text{CN}]^+$.

Initial electrolysis of $[2\text{-CH}_3\text{CN}]^+$ at -1.53 V exhibits depletion of the parent complex at 2043 and 1957 cm^{-1} with concurrent grow-in of the one-electron reduced six-coordinate species $[2\text{-CH}_3\text{CN}]^0$ at 2023 , 1931 and 1910 cm^{-1} (eqn (1)). Weak evidence of the two-electron reduced species $[2]^-$ appeared to grow-in over an extended electrolysis time at 1807 cm^{-1} . Quantitative formation of $[2]^-$ was observed at 1895 and 1807 cm^{-1} by applying a more negative potential of -2.00 V. A weak depletion of the parent *N*-boc $\nu(\text{C}=\text{O})$ stretches was also observed upon formation of $[2]^-$, however this region of the spectrum was mostly obscured by strong electrolyte absorption. Comparable IRSEC results were recorded for all four complexes (Table 2) and is provided in the ESI (Fig. 25–27†).

Voltammetry under 1 atm CO_2 with Brønsted acid

Each catalyst was studied under 1 atm CO_2 in the presence of optimum H_2O and PhOH. Brønsted acid titration experiments were performed to determine the concentration of each acid at which proton dependence would not be rate-limiting. Furthermore, once optimum Brønsted acid concentrations were reached, scan rate dependence experiments were performed to reach steady-state conditions at which CO_2 consumption is also not rate-limiting to allow determination of the maximum turnover frequencies (TOF_{max}). The turnover frequency was determined using the ratio of catalytic current (i_{cat}) and non-catalytic faradaic current (i_{p}) according to eqn (6)

$$\text{TOF} = \frac{F\nu n_p^3}{RT} \left(\frac{0.4463}{n_{\text{cat}}} \right)^2 \left(\frac{i_{\text{cat}}}{i_{\text{p}}} \right)^2 \quad (6)$$

where F is the Faraday constant (96485 s A mol^{-1}), ν is the scan rate (V s^{-1}), R is the universal gas constant (8.3145 V A s

Table 2 IRSEC-derived FTIR absorption data recorded in 0.1 M $[\text{Bu}_4\text{N}][\text{PF}_6]$ acetonitrile electrolyte for all complexes, summarizing $\nu(\text{CO})$ stretching frequencies

	$\nu(\text{CO})$ cm^{-1}
$[1\text{-CH}_3\text{CN}]^+$	2043, 1958, 1949(sh), 1768 ^a , 1752(sh) ^a
$[1\text{-CH}_3\text{CN}]^0$	2023, 1925
$[1]^-$	1911, 1812
$[2\text{-CH}_3\text{CN}]^+$	2043, 1957, 1772 ^a , 1713 ^a
$[2\text{-CH}_3\text{CN}]^0$	2023, 1931, 1910
$[2]^-$	1895, 1807
$[3\text{-CH}_3\text{CN}]^+$	2043, 1957, 1747 ^a , 1729(sh) ^a
$[3]^-$	1912, 1808
$[4\text{-CH}_3\text{CN}]^+$	2044, 1958, 1948(sh), 1769 ^a , 1755(sh) ^a
$[4]^-$	1913, 1813

^a Low energy ligand $\nu(\text{C}=\text{O})$ absorption bands are obscured by strong electrolyte absorption for one- and two-electron reduced species in IRSEC experiments.

$\text{K}^{-1} \text{mol}^{-1}$), T is the temperature (298 K).⁶³ The first reduction wave was used as a reference point for both i_{p} and n_{p} (*i.e.* $n_{\text{p}} = 2$ for a two electron reduction). The number of electrons required for a single catalytic cycle is equivalent for each of the CO , HCO_2^- and H_2 producing catalytic cycles (*i.e.* $n_{\text{cat}} = 2$). The maximum turnover frequency (TOF_{max}) was estimated by plotting TOF vs. scan rate.

As discussed above, all four catalysts are activated by a two-electron reduction to the five coordinate 18-electron $[\text{Mn}]^-$ anions, primed for CO_2 and/or proton binding to form the key $[\text{Mn}\text{-CO}_2\text{H}]$ and $[\text{Mn}\text{-H}]$ intermediates. Optimum voltammetry at $\nu = 0.1$ V s^{-1} for all four catalysts comparing both H_2O and PhOH proton sources is summarized in Table 4 below and illustrated in Fig. 4. Two catalytic waves grow-in in the presence of optimum PhOH (1.5 M) as the proton source. Consistent with our prior studies using a bulky proton donor, growth of the low-overpotential protonation-first pathway appears to be most prominent for $[1\text{-CH}_3\text{CN}]^+$ and $[2\text{-CH}_3\text{CN}]^+$, which arguably possess lesser steric bulk.

Controlled potential electrolysis

To probe the product selectivity of each catalyst, controlled potential electrolysis (CPE) experiments were conducted at $E_{\text{cat}/2}$ of each catalytic wave under 1 atm CO_2 in the presence of an optimum concentration of H_2O or PhOH Brønsted acid proton sources. A summary of CPE parameters and faradaic yields is presented below in Table 3.

Of the four catalysts investigated only the simple methyl-ester second coordination sphere functional group in $[1\text{-CH}_3\text{CN}]^+$ maintains a high selectivity for CO production across all CPE experimental conditions. Specifically, $[1\text{-CH}_3\text{CN}]^+$ exhibits $\text{FE}_{\text{CO}} = 99\%$ at high-overpotential ($E_{\text{cat}/2} = -1.96$ V) in the presence of 6.0 M H_2O . In the presence of just 1.5 M PhOH, $[1\text{-CH}_3\text{CN}]^+$ exhibits $\text{FE}_{\text{CO}} = 97\%$ and 100% at its low overpotential ($E_{\text{cat}/2} = -1.55$ V) and high overpotential ($E_{\text{cat}/2} = -1.93$ V) catalytic waves, respectively. With such high selectivity for CO production the low and high overpotential catalytic waves of $[1\text{-CH}_3\text{CN}]^+$ can be definitively attributed to the PT-ET (aka pro-

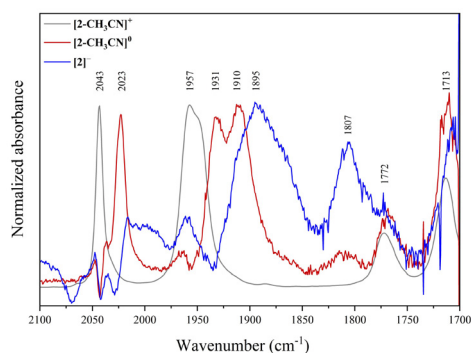
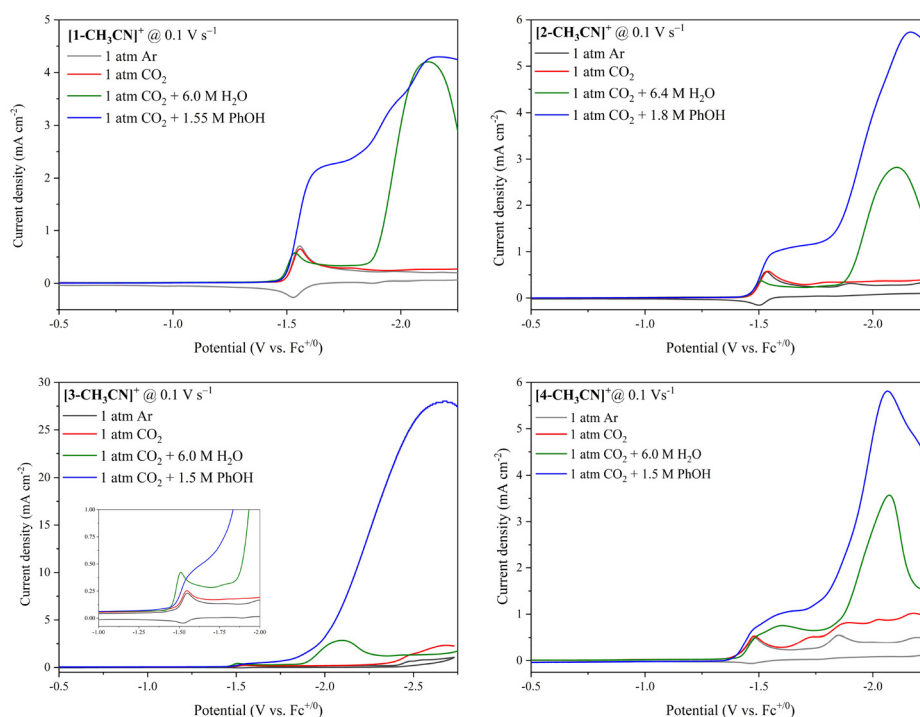


Fig. 3 IRSEC results upon controlled potential electrolysis of $[2\text{-CH}_3\text{CN}]^+$ recorded in argon-saturated 0.1 M $[\text{Bu}_4\text{N}][\text{PF}_6]$ acetonitrile supporting electrolyte.



Table 4 Summary of electrocatalysis data derived from linear sweep voltammogram experiments for both low and high overpotential pathways

		$[1\text{-CH}_3\text{CN}]^+$		$[2\text{-CH}_3\text{CN}]^+$		$[3\text{-CH}_3\text{CN}]^+$		$[4\text{-CH}_3\text{CN}]^+$	
		H ₂ O	PhOH	H ₂ O	PhOH	H ₂ O	PhOH	H ₂ O	PhOH
Low η	$E_{\text{cat}/2}$ (V)	—	−1.55	—	−1.51	−1.47	−1.50	—	−1.45
	[HA] (M)	—	1.5	—	1.8	6.0	1.5	—	1.5
	TOF _{max} (s ^{−1})	—	58	—	35	6	5	—	82
High η	$E_{\text{cat}/2}$ (V)	−1.96	−1.93	−1.96	−1.94	−1.96	−2.25	−1.92	−1.93
	[HA] (M)	6.0	1.5	6.4	1.8	6.0	1.5	6.0	1.5
	TOF _{max} (s ^{−1})	167	235	127	598	78	10 451	586	1846

**Fig. 4** Linear sweep voltammetry of all four complexes recorded at 1 mM concentration in 0.1 M [Bu₄N][PF₆] acetonitrile electrolyte. Each plot includes cyclic voltammetry under 1 atm Ar, overlaid with linear sweep voltammetry in CO₂ saturated electrolyte with the addition of optimum concentrations of H₂O and PhOH.

tonation-first) and ET-PT (aka reduction-first) pathways of the CO₂-to-CO conversion catalytic cycle (re. Scheme 2 below). This observation further emphasizes the unique capacity of an *o*-aryl aprotic O-atom Brønsted base at promoting the highly desired low-overpotential PT-ET pathway for CO₂-to-CO conversion by Mn polypyridyl electrocatalysts.⁵⁸

Discussion

To correctly discern the data here presented the relevant competing catalytic pathways exhibited by the Mn polypyridyl class of CO₂ reduction electrocatalysts must first be fully appreciated. Presented below are the established low- and high-overpotential catalytic pathways for each of the CO, HCO₂[−] and H₂ products generated from the two-electron activated catalyst

[Mn][−]; for brevity the core five-coordinate [*fac*-Mn(R₂bpy)(CO)₃] structure is abbreviated as [Mn] with [Mn-X] indicating the variable sixth coordination site throughout the catalytic cycle (Scheme 2). Formal Mn and R₂bpy oxidation states throughout the catalytic cycle are also illustrated in Scheme 2. The CO producing pathways originate from CO₂ binding to the active catalyst [Mn][−], in the presence of a proton source, to generate the metallocarboxylic acid intermediate [Mn-CO₂H]⁰. This intermediate may directly undergo rate-determining proton induced C-OH bond cleavage to evolve H₂O alongside the [Mn^I-CO]⁺ cation. Reduction of this cation to [Mn^I-CO]⁰ requires a potential less negative than initial catalyst activation, hence the onset of catalytic current for this low-overpotential PT-ET (aka protonation-first) pathway for CO evolution being coincident with the concerted two-electron catalyst activation reduction wave. The corresponding high-overpotential



Table 3 Summary of controlled potential electrolysis data. Experimental conditions: 5 mL of 1 mM catalyst in 0.1 M [Bu₄N][PF₆] acetonitrile supporting electrolyte with optimum H₂O and PhOH concentrations under 1 atm CO₂

	HA ^a (M)	E _{applied} (V vs. Fc ^{+/0})	Faradaic yield CO : HCO ₂ ⁻ : H ₂ (%)
[1-CH ₃ CN] ⁺	H ₂ O (6.0 M)	-1.96	99 : 0 : 1
	PhOH (1.5 M)	-1.55	97 : 0 : 1
	PhOH (1.5 M)	-1.93	100 : 0 : 0
[2-CH ₃ CN] ⁺	H ₂ O (6.4 M)	-1.96	42 : 48 : 1
	PhOH (1.8 M)	-1.51	100 : 0 : 0
	PhOH (1.8 M)	-1.94	54 : 43 : 1
[3-CH ₃ CN] ⁺	H ₂ O (6.0 M)	-1.60 ^b	4 : 12 : 20
	H ₂ O (6.0 M)	-1.96	16 : 25 : 58
	PhOH (1.5 M)	-1.50	32 : 19 : 7
	PhOH (1.5 M)	-2.25	14 : 17 : 67
[4-CH ₃ CN] ⁺	H ₂ O (6.0 M)	-1.92	63 : 0 : 36
	PhOH (1.5 M)	-1.60 ^b	66 : 0 : 12
	PhOH (1.5 M)	-1.93	59 : 0 : 34

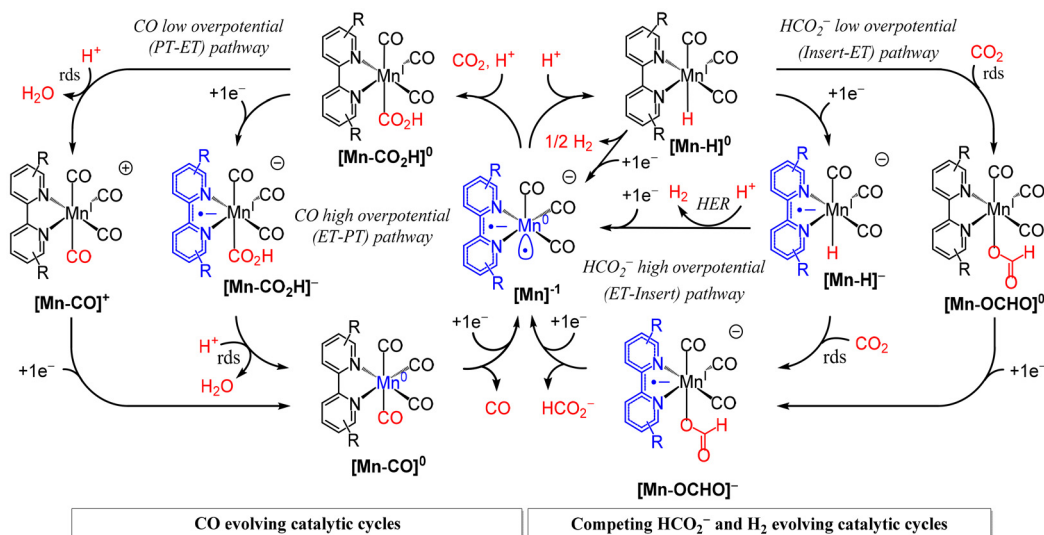
^a Refers to the concentration of Brønsted acid [HA] not be confused with [H⁺]. ^b E_{applied} was taken from the peak maximum due to the weak catalytic current.

ET-PT (aka reduction-first) CO producing pathway exhibits an onset of catalytic current at higher overpotential concomitant with one-electron reduction of the metallocarboxylic acid intermediate to generate the [Mn-CO₂H]⁻ anion. Rate-determining proton induced C-OH bond cleavage at the [Mn-CO₂H]⁻ anion is more facile hence the consistently higher catalytic rates for this ET-PT pathway relative to is low-overpotential PT-ET pathway. Should protonation of the [Mn]⁻ active catalyst be favoured over CO₂ binding, [Mn-H]⁰ is formed initiating entry to the competing HCO₂⁻ and H₂ evolving pathways. CO₂ insertion can occur, prior to reduction, along the low-overpotential (insert-ET) pathway, with an overpotential dictated by the

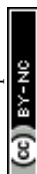
[Mn-CO₂H]^{0/-} reduction potential. Alternatively, the high-overpotential HCO₂⁻ producing pathway has an overpotential dictated by the [Mn-H]^{0/-} reduction, followed by CO₂ insertion (ET-insert). The low-overpotential insert-ET formate pathway is kinetically inferior but has the advantage of eliminating competitive H₂ production relative to protonation of the anionic [Mn-H]⁻ intermediate along the ET-insert pathway. It should also be noted that bimolecular H₂ production is known to occur from two equivalents of the neutral [Mn-H]⁰ intermediate, at least with the non-sterically hindered bpy ligand.¹²

The first take away from this study is how the simplest *o*-arylester second coordination sphere functionality in pre-catalyst [1-CH₃CN]⁺ exhibits a rare example of the highly sought after low overpotential PT-ET catalytic pathway for selective CO production. However, despite the presence of an optimized PhOH concentration (1.5 M), the TOF_{max} values for both its low overpotential PT-ET (TOF_{max} = 58 s⁻¹, E_{cat/2} = -1.55 V) and high overpotential ET-PT (TOF_{max} = 235 s⁻¹, E_{cat/2} = -1.93 V) CO producing pathways are relatively low in comparison to our previously reported *o*-arylether analogues. That being said, the goal of this study was primarily to probe the impact of extending this second coordination sphere functionality on catalyst performance from the perspective of thermodynamics, kinetics and selectivity.

Second coordination sphere extension with the *N*-*boc* arylester system in pre-catalyst [2-CH₃CN]⁺ also exhibits highly selective CO production at low overpotential consistent with the PT-ET pathway in the presence of 1.8 M PhOH (TOF_{max} = 35 s⁻¹, E_{cat/2} = -1.51 V). The slight reduction in catalytic rate relative to [1-CH₃CN]⁺ is likely due to steric bulk and crowding of the catalyst active site. Intriguingly, however, at higher overpotential (E_{cat/2} = -1.94 V) under otherwise identical conditions this same catalyst exhibits an almost equal faradaic



Scheme 2 A summary of competing catalytic pathways exhibited by the Mn polypyridyl class of CO₂ reduction electrocatalysts for CO, HCO₂⁻ and H₂ production. Left side: Low-overpotential (PT-ET) and high-overpotential (ET-PT) CO evolving pathways. Right side: Low-overpotential (insert-ET) and high-overpotential (ET-insert) pathways for HCO₂⁻ production, and competing hydrogen evolution reaction (HER) pathways.



yield for HCO_2^- production with negligible evidence for H_2 evolution. This suggests an almost 1 : 1 split with respect to the selectivity of the activated catalyst $[2]^-$ at binding CO_2 or H^+ to form the critical metallocarboxylic acid $[2\text{-CO}_2\text{H}]$ or metal-hydride $[2\text{-H}]$ intermediates along the CO or HCO_2^- vs. H_2 catalytic pathways (Scheme 2). $[2\text{-CH}_3\text{CN}]^+$ exhibits a very similar product distribution of $\text{FE}_{\text{CO:HCO}_2^-:\text{H}_2} = 42 : 1 : 48\%$ in the presence of optimum H_2O (6.4 M) at high overpotential ($E_{\text{cat}/2} = -1.96$ V). Of great significant for each of these high overpotential observations of equitable CO vs. HCO_2^- formation is the negligible production of H_2 . As the high-overpotentials required are consistent with HCO_2^- production along the ET-insert pathway, *via* the one-electron reduced manganese hydride intermediate $[2\text{-H}]^-$, this suggests that the *N*-*boc* terminated *o*-arylester second coordination sphere of $[2\text{-CH}_3\text{CN}]^+$ favors CO_2 insertion over protonation of $[2\text{-H}]^-$ to maintain a high selectivity for HCO_2^- production. This observation is indeed consistent with related complexes which also contain pendant amino functionalities in their SCS.^{19,22,48} Estimated hydricities of the neutral $[\text{Mn-H}]$ catalyst, using Kubiak's linear correlation method with two-electron reduction potentials, of all four complexes are in the range of $\Delta G_{\text{H}}^\circ \sim 46 - 47 \text{ kcal mol}^{-1}$.⁶⁴ This is slightly endergonic of the 44 kcal mol^{-1} assumed necessary for HCO_2^- formation and may also explain the low yields of H_2 at low overpotential. Hydricity of the reduced $[\text{Mn-H}]^-$ intermediate has been estimated at $\sim 14 \text{ kcal mol}^{-1}$ for a catalyst that exhibited 71% HCO_2^- selectivity and which bears pendant amine second coordination sphere, comparable in many respects to $[2\text{-CH}_3\text{CN}]^+$.⁴⁸

In contrast, the decreased selectivity imparted by the *o*-methoxybenzoate second coordination sphere in $[3\text{-CH}_3\text{CN}]^+$ relative to the *N*-*boc* arylester in $[2\text{-CH}_3\text{CN}]^+$, in the presence of either PhOH or H_2O , is striking. It is also worth noting the more negative $E_{\text{cat}/2}$ of the large high-overpotential catalytic wave of $[3\text{-CH}_3\text{CN}]^+$ at -2.25 V in the presence of PhOH. Considering how the three-electron reduced derivative may be accessed at $E_{\text{pc}} = -2.48$ V (Table 1) this contrasting reactivity could possibly be associated with a three-electron activated catalyst thereby mitigating any desired control of product selectivity. However, an alternative explanation may be that the product selectivity of $[3\text{-CH}_3\text{CN}]^+$ is simply dictated by the steric bulk of the extended second coordination sphere structure with limited influence of the second coordination sphere.⁵⁷ Controlled potential electrolysis studies of the bulky $[4\text{-CH}_3\text{CN}]^+$ catalyst do indicate that CO_2 binding is favoured but significant H_2 evolution is still observed in this case. Although computed transition state geometries suggest the Mn active site is accessible in the *syn,anti* atropisomers of $[3\text{-CH}_3\text{CN}]^+$ and $[4\text{-CH}_3\text{CN}]^+$ (ESI Fig. 42–47†), not knowing the ratio of atropisomers in solution prevents any real conclusions to be made at this time with respect to the balance of steric vs. second coordination sphere effects on their product distribution.

To further investigate the influence of extended functionality at the second-coordination sphere, we performed density

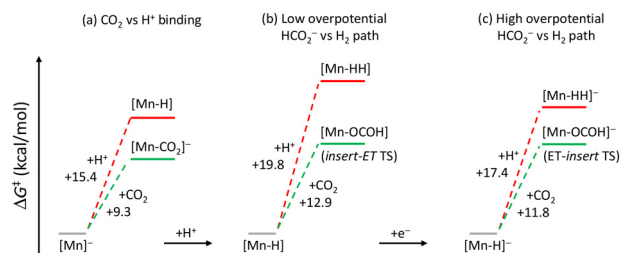


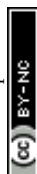
Fig. 5 A comparison of relative free energies of activation ($\Delta G^\ddagger$) values in kcal mol^{-1} for selected kinetic steps in Scheme 2 for $[2\text{-CH}_3\text{CN}]^+$ with PhOH as the Brønsted acid source.

functional theory (DFT) calculations at the MN15 level of theory in conjunction with the SMD continuum solvation model for acetonitrile. For computational efficiency only the *syn,anti* atropisomers were calculated, and the terminal *tert*-butyl substituents of the *N*-*boc* protecting groups were modelled as methyl groups (see Computational methods in the ESI† for details). Using theory to predict product selectivities at high overpotential is extremely challenging due to the competition between several pathways leading to CO, HCO_2^- and H_2 production. This is further compounded by the unknown ratio of atropisomers in each CPE experiment which likely exhibit varying product distributions. Thus, our primary aim in utilizing theoretical calculations in this work focussed on a qualitative assessment for the disparate product selectivities of just $[2\text{-CH}_3\text{CN}]^+$ and $[3\text{-CH}_3\text{CN}]^+$. The energetics of their transition state structures for $[\text{Mn-CO}_2]^-$ and $[\text{Mn-H}]$ formation, and Insert-ET *versus* ET-insert pathways for HCO_2^- and H_2 production, were analyzed. The free energy of activation ($\Delta G^\ddagger$) values, assuming fully separated reactants in solution, are illustrated in Fig. 5 for $[2\text{-CH}_3\text{CN}]^+$, (for tabulated data vs. $[3\text{-CH}_3\text{CN}]^+$ see ESI Table 1†).

The computed $\Delta G^\ddagger$ s indicate that CO_2 binding to doubly reduced $[2]^-$ is significantly favored over protonation by PhOH to generate $[2\text{-H}]$, however, the activation enthalpies (ESI Table 1,† $\Delta H^\ddagger$) are nearly isothermic which is more consistent with the almost equal product distribution of CO : HCO_2^- in controlled potential electrolysis. Once the $[2\text{-H}]^-$ intermediate is formed, HCO_2^- formation by electrophilic attack of CO_2 is favored over H_2 formation in line with our experimental observations (Fig. 5c). Theoretical data for $[3\text{-CH}_3\text{CN}]^+$ was inconsistent with our CPE experiments, however, as pointed out earlier, this may likely be due to three-electron activation of pre-catalyst $[3\text{-CH}_3\text{CN}]^+$ leading to excessive H_2 production, or alternatively steric crowding of the active site hindering CO_2 binding at $[3]^-$ or insertion at $[3\text{-H}]^-$ thereby enhancing the H_2 yield.

Conclusions

This study demonstrates that the introduction of distal, outer coordination sphere, H-bonding functionality at the periphery of a homogeneous transition metal complex active site, can have a strong influence on product selectivity for proton-



coupled electrocatalytic CO₂ reduction. Although the bulky *o*-methoxybenzoate group diminishes the product selectivity for CO₂ reduction, extension of the CO selective acetate SCS with the tertiary amine *N*-Boc-ala group shifts the reaction pathway toward a 1:1 ratio of the [Mn-CO₂H] and [Mn-H] intermediates with a surprising selectivity for CO₂ insertion at [Mn-H] along the competing HCO₂⁻ vs. H₂ pathways as corroborated by computational studies.

Author contributions

The manuscript was drafted through the contributions of all authors: L. S., V. B. and R. S. experimental studies and data analysis; M. Z. E. computational analysis; J. R. data analysis, L. S., V. B. and R. S. supervision and project management.

Data availability

The data supporting this article have been included as part of the ESI.†

Conflicts of interest

There are no conflicts to declare.

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