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Metal-nanocluster science and technology: my personal history and outlook

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Metal nanoclusters (NCs) are one of the leading targets in research of nanoscale materials, and elucidation of their properties (science) and development of control techniques (technology) have been continuously studied for the past 60 years or so. I have been continuously studying metal NCs since 1995 while changing my own interests, and thereby research topics, according to development of the field and the global situation. Specifically, I started my research on gas-phase metal NCs, and then investigated systematic isolation of glutathionate-protected gold NCs. Thereafter, I worked on development of flexible and precise techniques for controlling ligand-protected metal NCs and creation of highly functional energy and environmental catalysts using ligand-protected metal NCs. This personal account summarizes my previous studies on metal NCs. Through this personal history, I would like to describe my motivation for each previous subject, what type of research I want to perform for each subject in the future, and how I consider the outlook of the field.

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1. Introduction

Research on nanoscale materials is generally believed to have begun with a talk by Feynman in 1959 entitled “There’s Plenty of Room at the Bottom”.¹ In this talk, he stated that in the nanoscale region of a few atoms, properties that are completely

different from the macroscopic world appear. He also claimed that it is very important for humanity to conduct research in this area, and he predicted that in the future we would be able to arrange atoms one by one as we wish to create desired materials. Later, high-resolution analysis devices, such as scanning tunneling microscopes² and atomic force microscopes,³ were developed, which made it possible to directly observe atoms on a surface. In addition, instruments such as lasers and molecular beams^{4,5} have been successively developed, and much knowledge has been gained about the stability and electronic/geometric structure of aggregates consisting of metal atoms in a vacuum, namely, gas-phase metal nanoclusters (NCs). Since 2000, many countries have adopted nanotechnology as a national policy,⁶ which has led to dramatic advancement in the technology and knowledge related to these materials, as well as a dramatic increase in the number of researchers involved in research of nanoscale materials. Through this knowledge, it is now possible to synthesize metal NCs consisting of certain metal elements with atomic precision.^{7–27} In addition, through these studies, it has been experimentally demonstrated that in the nanoscale region, novel electronic/geometric structures and physicochemical properties different from those of ordinary metals actually emerge. In recent years, a number of studies on creation of new materials that take advantage of the unique properties of such nanoscale materials have been reported. Thus, with regard to nanoscale materials, both elucidation of their properties (science) and development of control techniques (technology) have made great progress in the past 60 years or so.^{28–38}

I have been continuously studying metal NCs since 1995 (Fig. 1). In this time, I have witnessed dramatic progress of the

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Tsukuda). His current research interests include the precise synthesis of stable and functionalized metal nanoclusters and their applications in energy and environmental materials.



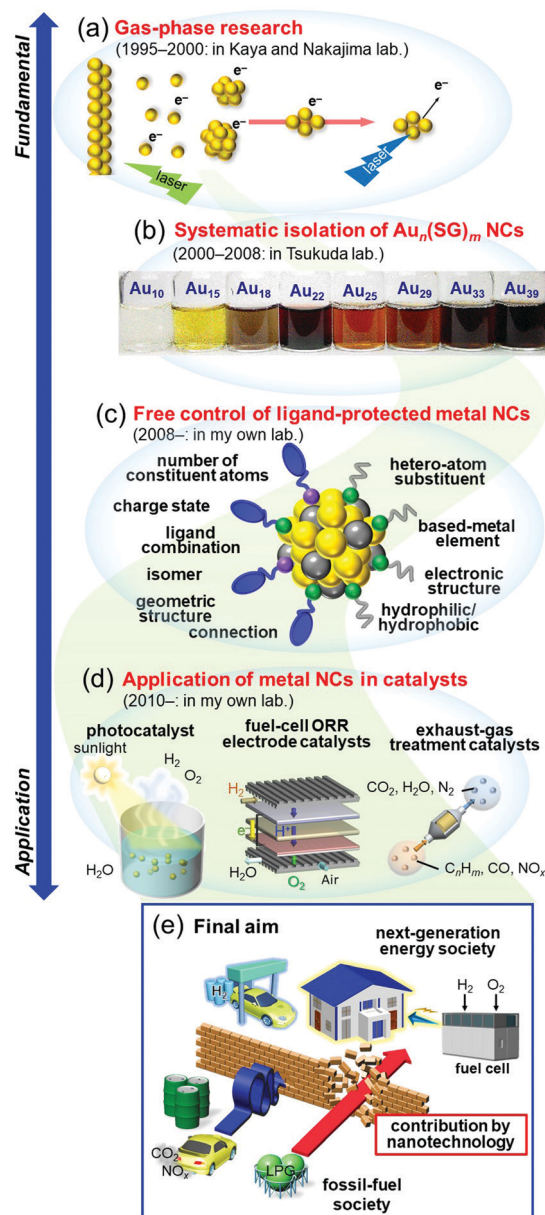


Fig. 1 (a–d) History of my research on NCs: (a) gas-phase research, (b) systematic isolation of glutathionate-protected Au NCs, (c) development of flexible and precise control techniques for ligand-protected metal NCs, and (d) application of metal NCs to energy and environmental catalysts. (e) Final aim of my research and what I would like to contribute through my studies. In this figure, LPG indicates the liquefied petroleum gas.

“science” and “technology” in related fields. In response to development of these fields and the state of the world, I have changed my own interests and research themes many times. Specifically, I started my research on gas-phase metal NCs (Fig. 1(a)),^{39–41} and then investigated systematic isolation of thiolate (SR)-protected gold (Au) NCs (Au_n(SR)_m; Fig. 1(b)).^{42–47} Since 2008, I have been working on development of flexible and precise control techniques for ligand-protected metal NCs (Fig. 1(c)) and their use in creation of highly functional energy and environment catalysts (Fig. 1(d)).⁴⁸ I have already

written several accounts^{49–57} and reviews^{58–66} on each research topic. Therefore, in this personal account, I will describe my personal history in metal NCs. Through this personal history, I will explain what has led me to perform each type of research, what type of research I would like to perform in the future, and what I consider to be the future prospects in the field.

2. Gas-phase research

Isolation of metal NCs in a vacuum where there is no interaction with the outside world and investigation of their formation distribution, structure, and physical properties can provide an essential and fundamental understanding of the origin of the stability and mechanism of the size-specific structures and properties of metal NCs. Investigation of isolated metal NCs and supported metal NCs in the gas phase has a great advantage in this respect, and it has therefore been carried out for many years, with various improvements in the equipment.^{67–80}

I conducted research on gas-phase NCs from 1995 to 2000 (Fig. 1(a)) under the guidance of Professor Kaya and Associate Professor Nakajima. In this research, NCs were produced by laser vaporization^{4,5} and their electronic structures were investigated by negative-ion photoelectron spectroscopy (Fig. 2(a)).⁸¹ This research revealed that in the bulk, group 14 elements are classified as semiconductors for silicon (Si) and germanium (Ge), semi-metals for tin (Sn), and metals for lead (Pb), whereas in the fine size range all NCs have highest occupied molecular orbital (HOMO)–lowest unoccupied molecular orbital (LUMO) gaps and overall similar electronic structures (Fig. 2(b)).^{40,82,83} In a study of Au–silver (Ag) alloy NCs, we experimentally demonstrated that not only controlling the number of constituent atoms, but also mixing the different elements is very effective for creating a new electronic structure (Fig. 2(c)).⁸⁴ Through this research, I confirmed that new materials can be created by controlling the number of constituent atoms and chemical composition of NCs with atomic precision.

However, through this research, I also strongly felt that it is essential to develop a large-scale synthesis method for NCs to create new materials that take advantage of their size-specific properties while understanding their fundamental properties. In addition, at this time, I got the chance to deposit germanium oxide (Ge_nO_m) NCs generated in a vacuum on a support and investigate their luminescence properties in air.⁸⁵ Because the Ge_nO_m NCs deposited at that time were originally oxides, their electronic structure did not significantly change even when exposed to air, and they exhibited blue luminescence in air. Through this experiment, I was impressed that it was essential to develop NCs that are stable in air to create NCs that can be used as materials.

3. Isolation of glutathionate-protected Au NCs

While my interest was shifting from gas-phase experiments to materials chemistry research, in 2000, I got the chance to



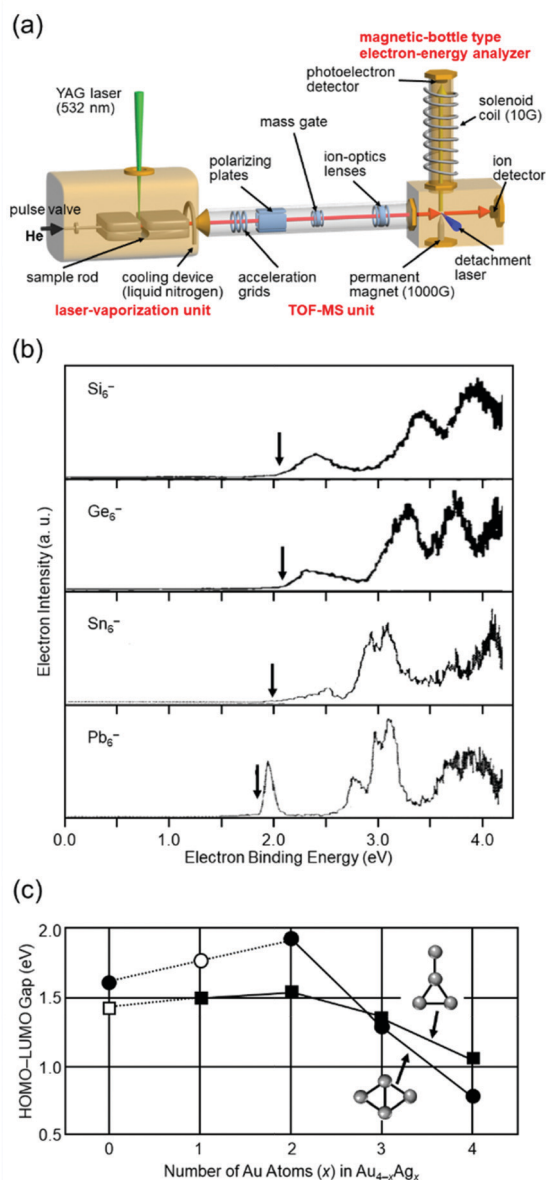


Fig. 2 (a) Schematic illustration of the apparatus used in my research on gas-phase NCs. (b) Photoelectron spectra of Si₆⁻, Ge₆⁻, Sn₆⁻, and Pb₆⁻ obtained by irradiating a 266 nm laser as a detachment laser. The vertical arrows indicate the thresholds energies.^{82,83} (c) Dependence of the HOMO–LUMO gaps of the rhombic and “T-form” isomers on the number of the Ag atoms (*x*) for Au_{4-*x*}Ag_{*x*} (*x* = 0–4). The solid circles are the rhombic isomers and the solid squares are the T-form isomers.⁸⁴ Reproduced with permission from ref. 82–84. Copyright 1997 Elsevier Science B.V., 2000 Elsevier Science B. V., and 2001 American Institute of Physics.

work as an Assistant Professor in the laboratory of Associate Professor Tsukuda, who had just established his own laboratory at the Institute for Molecular Science. Tsukuda also originally studied gas-phase NCs, but he started to work on liquid-phase synthesis of metal NCs at the end of his previous position.⁸⁶ After joining his laboratory, we worked together to develop a precise synthesis method for metal NC in solution (Fig. 1(b)). These efforts were not immediately successful, and it took

several years to achieve the first isolation. After a series of selections⁸⁷ and abandonments of various material systems, we finally arrived at Au_{*n*}(SG)_{*m*} NCs in 2003.^{42,88}

Au and SR form strong bonds, and thereby SR forms self-assembled monolayers on Au surfaces.^{89–91} For these reasons, many technologies that take advantage of the combination of Au and SR have been developed in nanotechnology, such as dip-pen nanolithography^{92,93} and microcontact printing.⁹⁴ In 1994, Brust *et al.*⁹⁵ reported that combination of Au and SR is also effective for formation of metal NCs. They found that Au_{*n*}(SC12)_{*m*} NCs (SC12 = dodecanethiolate) can be formed by reduction of Au salts with sodium borohydride in the presence of dodecanethiol in solution. The Au_{*n*}(SC12)_{*m*} NCs reported at this time had a particle-size distribution of 1–3 nm. Regarding metal NCs, synthesis of metal NCs using phosphine (PR₃), carbon monoxide (CO), and halogens as ligands, such as Au_{*n*}(PR₃)_{*m*}(Cl)_{*l*}, Pt_{*n*}(PR₃)_{*m*}(CO)_{*l*} (Pt = platinum) and Pd_{*n*}(PR₃)_{*m*}(CO)_{*l*} (Pd = palladium), had been reported before the report of Au_{*n*}(SC12)_{*m*} NCs. Those had been isolated with atomic precision, and their geometric structures had been determined by single-crystal X-ray diffraction (SC-XRD).^{23,96} However, the number of constituent atoms of Au_{*n*}(SC12)_{*m*} NCs could not be controlled at atomic precision at this stage. Because Au_{*n*}(SC12)_{*m*} NCs are easier to synthesize and more stable in air than Au_{*n*}(PR₃)_{*m*}(Cl)_{*l*}, Pt_{*n*}(PR₃)_{*m*}(CO)_{*l*}, and Pd_{*n*}(PR₃)_{*m*}(CO)_{*l*} NCs, Au_{*n*}(SR)_{*m*} NCs were easy for researchers who had not previously worked with NCs to investigate. Although the number of researchers working on metal NCs dramatically increased after 2000 when many countries had adopted nanotechnology as a national policy, this was probably the reason why the number of researchers working on Au_{*n*}(SR)_{*m*} NCs particularly increased.

Whetten and co-workers⁹⁷ and Murray and co-workers⁹⁸ had worked on high-resolution separation of Au_{*n*}(SR)_{*m*} NCs since the 1990s. Among their studies, the work on Au_{*n*}(SG)_{*m*} NCs (SG = glutathionate) reported by Whetten and co-workers in 1998⁹⁹ and 2000¹⁰⁰ deserves special mention. They successfully separated a series of Au_{*n*}(SG)_{*m*} NCs at high resolution by polyacrylamide gel electrophoresis (PAGE, Fig. 3(a)) and reported that the chemical composition of the main product was Au₂₈(SG)₁₆.⁹⁹

After these reports, Tsukuda and I started to work on Au_{*n*}(SG)_{*m*} NCs, where our first aim was to determine the chemical compositions of Au_{*n*}(SG)_{*m*} NCs other than the main product by mass spectrometry (MS). We had experienced development of the vacuum apparatus at our former positions.¹⁰¹ Therefore, we developed an electrospray ionization (ESI) unit that could deliver Au_{*n*}(SG)_{*m*} NCs synthesized in solution to a mass spectrometer without dissociation (Fig. 3(b)).¹⁰² Using a mass spectrometer equipped with the original ESI unit, we succeeded in obtaining the ESI-MS spectra of all of the Au_{*n*}(SG)_{*m*} NCs separated by PAGE. In a paper published in 2004, we assigned the Au_{*n*}(SG)_{*m*} NCs to Au₁₈(SG)₁₁, Au₂₁(SG)₁₂, Au_{25±1}(SG)_{14±1}, Au₂₈(SG)₁₆, Au₃₂(SG)₁₈, and Au₃₉(SG)₂₃.¹⁰³ However, further studies with modification of sample purification and the MS sections revealed that the chemical compositions



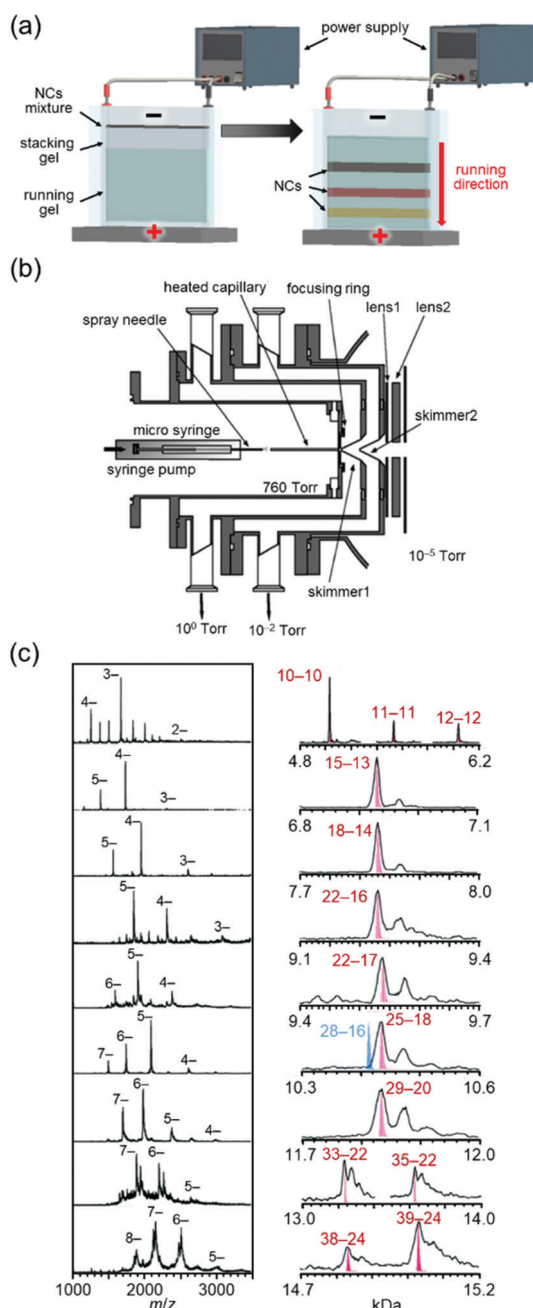


Fig. 3 (a) Schematic illustration of PAGE separation techniques.⁶¹ (b) Details of the ESI source. (c) Negative-ion ESI mass spectrum of each fraction separated by PAGE.¹⁰² In right figure, $Au_n(SG)_m$ is described as $n-m$. Reproduced with permission from ref. 61 and 102. Copyright 2020 The Royal Society of Chemistry and 2005 American Chemical Society.

reported in this paper were incorrect. Therefore, the chemical compositions of the series of $Au_n(SG)_m$ NCs were corrected to $Au_{10-12}(SG)_{10-12}$, $Au_{15}(SG)_{13}$, $Au_{18}(SG)_{14}$, $Au_{22}(SG)_{16}$, $Au_{22}(SG)_{17}$, $Au_{25}(SG)_{18}$, $Au_{29}(SG)_{20}$, $Au_{33}(SG)_{22}$, and $Au_{39}(SG)_{24}$ in a paper published in 2005 (Fig. 3(c)).¹⁰² Note that the chemical composition of the stable $Au_{25}(SR)_{18}$ NC^{46,105} was determined for the first time in this work. Determination of the chemical compositions of a series of Au NCs was also achieved for the Au NCs protected by mercaptosulfuric acid (S(SA)) or *N*-(2-mercaptopropionyl)glycine

(S(SG)).¹⁰⁴ Through these studies, we succeeded in demonstrating that $Au_n(SR)_m$ NCs can be treated as compounds with defined chemical compositions, similar to organic and complex compounds. Research of $Au_n(SR)_m$ NCs has provided a deep understanding of the correlation between the chemical composition, stability, electronic structure, and luminescence properties of $Au_n(SR)_m$ NCs.^{43,45,103-106}

4. Toward development of the ultimate nanotechnology and a deep understanding of ligand-protected metal NCs

After reporting $Au_n(SG)_m$ NCs, I got the opportunity to run my own laboratory at Tokyo University of Science in 2008. Therefore, after 2008, I began new projects. I had the feeling that atomic control of ligand-protected metal NCs would be the ultimate nanotechnology for metal-molecule complex systems. However, the NCs that we could control at my previous institute were limited to small $Au_n(SR)_m$ NCs ($n = 10-39$) protected by certain SRs (SG, S(SA), and S(SG)). To develop the ultimate nanotechnology for metal-molecule complex systems and thereby gain a better understanding of the correlation between the chemical composition, structure, and physical properties for ligand-protected metal NCs, it was necessary to develop new precise synthesis methods for (1) Au NCs with a large number of constituent atoms, (2) alloy NCs composed of several metal elements, (3) metal NCs composed of other elements, and (4) metal NCs protected by other ligands. According to the report by Teo *et al.*,¹⁰⁷ certain metal NCs could become the constituent units of “clusters of clusters”. To create new materials using such connections in the future, it was necessary to also make further progress in (5) the connection techniques of ligand-protected metal NCs.

4.1. Our strategy for flexible and precise control of ligand-protected metal NCs

Since 2008, we have been working in parallel to develop the following six techniques: (1) isolation of Au NCs with a large number of constituent atoms^{108,109} (Section 4.2 and Fig. 4(a)), (2) isolation of alloy NCs¹¹⁰⁻¹¹⁷ (Section 4.3 and Fig. 4(b)), (3) isolation of metal NCs consisting of Ag or Pt¹¹⁸⁻¹²¹ (Section 4.4 and Fig. 4(c)), (4) isolation of metal NCs protected by other ligands¹²²⁻¹²⁸ (Section 4.5 and Fig. 4(d)), (5) connection of metal NCs¹²⁹⁻¹³¹ (Section 4.6 and Fig. 4(e)), and (6) high-resolution separation of ligand-protected metal NCs by high-performance liquid chromatography (HPLC)^{53,55,61,108-110,132-138} (Section 4.6 and Fig. 4(f)). Through these six techniques, we have been attempting to develop flexible and precise control techniques for ligand-protected metal NCs, that is, the ultimate nanotechnology for metal-molecule complex systems (Fig. 4(g)), and thereby gaining a deeper understanding of the correlation between the chemical composition, structure, and physicochemical properties of ligand-protected metal NCs.



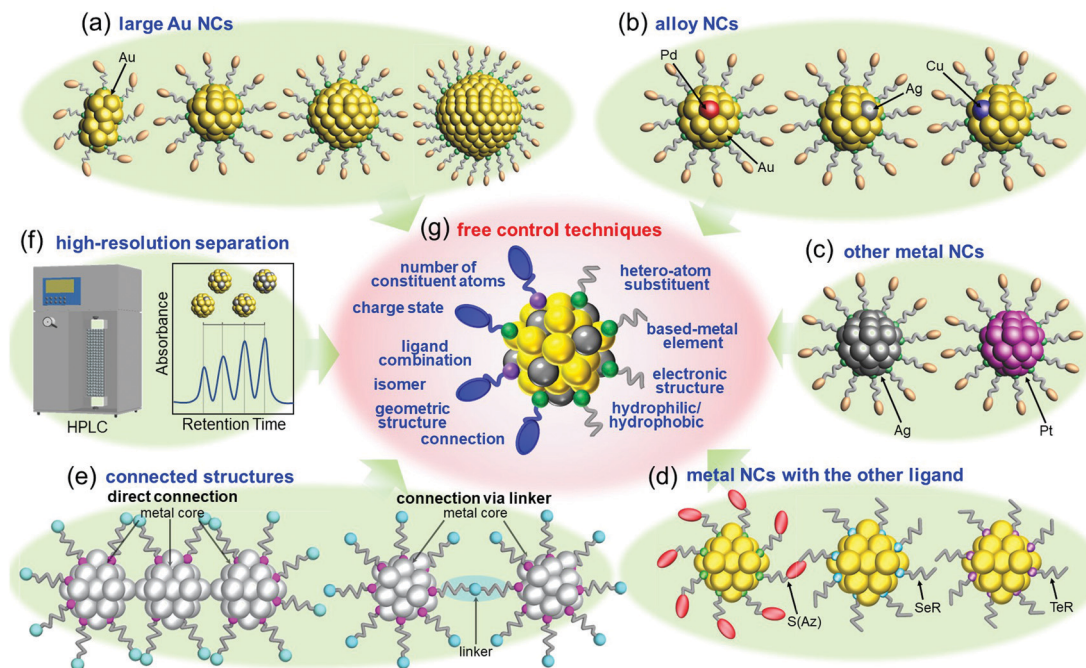


Fig. 4 Our strategy for developing ultimate control methods for ligand-protected metal NCs. Development of isolation methods for (a) large Au NCs, (b) Au-based alloy NCs, (c) Ag and Pt NCs, (d) metal NCs protected by S(Az), SeR, and TeR, (e) connected structures. (f) Development of high-resolution separation methods using HPLC. (g) Controllable parameters achieved by (a–f).

4.2. Isolation of relatively large Au NCs

In recent years, it has become possible to size-selectively synthesize relatively large $Au_n(SR)_m$ NCs.^{139,140} However, as of 2008, the largest $Au_n(SR)_m$ NCs that could be isolated with atomic precision were Au_{144} NCs,^{47,141,142} for which I was also involved in determining the chemical composition. Therefore, since 2008, we have been working on isolating $Au_n(SR)_m$ NCs larger than Au_{144} .

When the hydrophilic SRs (SG or S(SA)) described in Section 3 are used as ligands, relatively large $Au_n(SR)_m$ NCs are difficult to synthesize.¹⁴³ However, Whetten and co-workers⁹⁷ and Wolfe and Murray¹⁴⁴ reported that $Au_n(SR)_m$ NCs larger than Au_{144} NCs can be synthesized when hydrophobic SRs, such as alkanethiolate and 2-phenylethanethiolate (PET), are used as ligands. Therefore, we attempted to separate each $Au_n(SC12)_m$ NC from a mixture at high resolution by reversed-phase (RP) HPLC and then determine their chemical compositions by ESI-MS (Fig. 5(a)).^{108,109} Using these methods, we succeeded in systematically isolating $Au_n(SC12)_m$ NCs up to $Au_{\sim 520}(SC12)_{\sim 130}$ in 2014 (the paper was published in 2015).¹⁰⁹

In parallel with us, Jin and co-workers and Dass and co-workers were also working on isolation of relatively large $Au_n(SR)_m$ NCs, and as of 2014, they had reported isolation of $Au_{329}(PET)_{84}$ (or $Au_{333}(PET)_{79}$), $Au_{\sim 500}(PET)_{\sim 120}$, and $Au_{\sim 940}(PET)_{\sim 160}$.^{145–148} Through optical absorption spectroscopy of those NCs^{145–148} and density functional theory (DFT) calculations by Häkkinen and co-workers,¹⁴⁹ it had been predicted that non-bulk to bulk transition occurs at around $Au_{144}(SR)_{60}$ in the $Au_n(SR)_m$ -NC system. We worked with Tsukuda and Häkkinen to clarify the electronic/geometric structures of a series of isolated

$Au_n(SC12)_m$ NCs ($n = 38 \sim 520$) to gain a better understanding of the details of the size at which non-bulk to bulk transition occurs. The results showed that non-bulk to bulk transition occurs between $Au_{144}(SC12)_{60}$ and $Au_{187}(SC12)_{68}$ in the $Au_n(SC12)_m$ -NC system (Fig. 5(b)).¹⁰⁹ In subsequent papers by Jin and co-workers and Dass and co-workers,^{139,150–152} these electronic/geometric transitions were reported to occur in different size regions when the bulkier SR is used as the ligand.

4.3. Isolation of Au-based alloy NCs

Even if only one metal element is used as a constituent element, various electronic structures and physicochemical properties can be obtained by controlling the number of constituent atoms in the metal NCs. However, as mentioned in Section 2, if the size effect is combined with the synergy effect caused by mixing different elements, it is expected that a variety of different electronic structures and physicochemical properties can be increased. Therefore, since 2008, we have been working on isolation of alloy NCs and elucidation of their geometric/electronic structures and physicochemical properties.^{110–117}

When I started this research in 2008, it was not easy to control the number of constituent atoms, even for metal NCs consisting of a single metal element. Furthermore, when I started my laboratory at Tokyo University of Science, I only had undergraduate students in my laboratory. Therefore, I thought that it was necessary to investigate metal-NC systems in which I had some experience to create new alloy NCs. Accordingly, $Au_n(SR)_m$ NCs, which I had experience in handling, were selected as the base metal NCs for alloying.^{46,47}



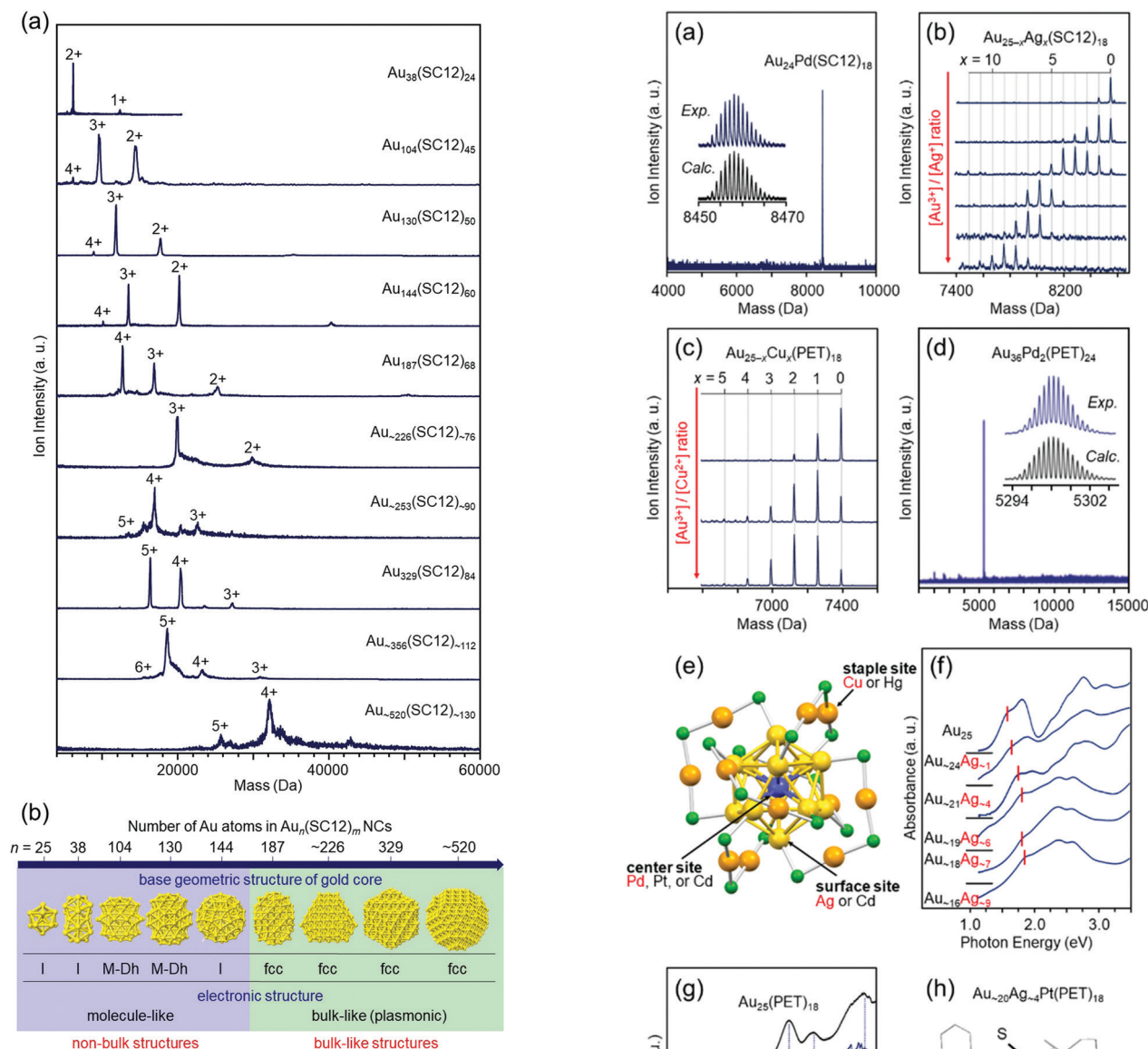


Fig. 5 (a) Positive-ion ESI mass spectra of $Au_n(SC12)_m$ NCs. (b) Non-bulk to bulk transition between $Au_{144}(SC12)_{60}$ and $Au_{187}(SC12)_{68}$ in the $Au_n(SC12)_m$ -NC system.¹⁰⁹ Reproduced with permission from ref. 109. Copyright 2015 American Chemical Society.

Specifically, we attempted to isolate alloy NCs in which some of the Au atoms in $[Au_{25}(SR)_{18}]^-$ and $[Au_{38}(SR)_{24}]^0$ ($SR = SC12$ or PET)^{153–157} were replaced by Pd, Ag, or Cu atoms.

Fortunately, we isolated $[Au_{24}Pd(SC12)_{18}]^0$ (Fig. 6(a)),¹¹⁰ $[Au_{25-x}Ag_x(SC12)_{18}]^-$ ($SR = SC12$ or PET , $x = 1–11$, Fig. 6(b)),¹¹¹ $[Au_{25-x}Cu_x(PET)_{18}]^-$ ($x = 1–5$, Fig. 6(c)),¹¹² and $[Au_{36}Pd_2(PET)_{24}]^0$ (Fig. 6(d))¹¹⁴ at an early stage. Studies on the obtained alloy NCs showed that (1) these three different elements have different preferential substitution sites (Fig. 6(e)),^{110–112,114,158} (2) Pd substitution is effective for creating more stable metal NCs than $Au_n(SR)_m$ NCs,^{110,114} (3) Ag substitution is effective for creating metal NCs with larger HOMO–LUMO gaps than $Au_n(SR)_m$ NCs (Fig. 6(f)),¹¹¹ (4) Cu substitution is effective for creating metal NCs with smaller HOMO–LUMO gaps than $Au_n(SR)_m$ NCs

Fig. 6 Negative-ion MALDI mass spectra of (a) $Au_{24}Pd(SC12)_{18}$,¹¹⁰ (b) $Au_{25-x}Ag_x(SC12)_{18}$,¹¹¹ and (c) $Au_{25-x}Cu_x(PET)_{18}$,¹¹² (d) Negative-ion ESI mass spectrum of $Au_{36}Pd_2(PET)_{24}$.¹¹⁴ (e) Preferential substitution site in $Au_{25}(SR)_{18}$ for each element.¹¹⁶ Blue, yellow, and orange spheres represent Au atoms, and green spheres represent S atoms. The R groups are omitted for clarity. Optical absorption spectra of (f) $[Au_{25-x}Ag_x(SC12)_{18}]^-$ and (g) $[Au_{24}Cu_x(PET)_{18}]^-$.^{112,158} (h) Geometric structure of $[Au_{20}Ag_4Pt(PET)_{18}]^0$.¹¹⁷ Reproduced with permission from ref. 110–112, 114, 116 and 117. Copyright 2010 The Royal Society of Chemistry, 2012 American Chemical Society, 2012 The Royal Society of Chemistry, 2016 The Royal Society of Chemistry, and 2018 American Chemical Society.



(Fig. 6(g)),^{112,158} (5) heteroatom substitution with Pd, Ag, or Cu enhances the speed of ligand-exchange reaction of NCS,^{52,159,160} and (6) Pd substitution improves the hydrogen evolution reaction (HER), oxygen evolution reaction, and oxygen reduction reaction (ORR) activities of the NCS.⁴⁸ In addition, studies on tri-metal¹¹⁵ and tetra-metal¹¹⁶ alloy NCS have revealed that simultaneous substitution with different elements at different substitution sites causes superimposed addition of each substitution effect on the base metal NCS.¹¹⁶ In these early studies, the substitution sites were predicted based on DFT calculations by Nobusada and/or X-ray absorption fine structure (XAFS) analysis by Tsukuda and Yamazoe. However, since 2016,^{128,161} we have also been able to determine the geometric structures of metal NCS by SC-XRD, and such studies have enabled us to (1) create $[\text{Au}_{\sim 20}\text{MAG}_{\sim 4}(\text{PET})_{18}]^{-}$ ($\text{M} = \text{Pd}$ or Pt), where both the number and position of Ag substitution are controlled (Fig. 6(h)),¹¹⁷ and (2) change the Cu substitution site by combining Cu and Pt substitution.¹⁶²

4.4. Isolation of Ag and Pt NCS

When I started my own laboratory, it was not easy to control the number of constituent atoms, even for $\text{Au}_n(\text{SR})_m$ NCS. Therefore, until 2010, the examples of isolated $\text{Ag}_n(\text{SR})_m$ NCS were limited to $\text{Ag}_7(\text{DMSA})_4$ (DMSA = meso-2,3-dimercaptosuccinic) reported by Jin and co-workers¹⁶³ and $\text{Ag}_9(\text{S}(\text{SA}))_8$ reported by Pradeep and co-workers.¹⁶⁴ Thus, we attempted to isolate larger $\text{Ag}_n(\text{SR})_m$ NCS.

Based on the report of Murray and co-workers,¹⁶⁵ it was expected that using 4-(*tert*-butyl)benzylmercaptan (SBB) as a ligand would result in synthesis of stable $\text{Ag}_n(\text{SR})_m$ NCS. Therefore, we first synthesized a mixture of $\text{Ag}_n(\text{SR})_m$ NCS and then applied the reaction with thiols, which had been frequently used in isolation of $\text{Au}_n(\text{SR})_m$ NCS,^{105,166} to the mixture to convert the unstable species to stable species. As a result, $\text{Ag}_n(\text{SBB})_m$ NCS with an average particle size of 2.1 nm were successfully obtained in a size-selective manner.¹¹⁸ The obtained stable $\text{Ag}_n(\text{SBB})_m$ NCS were assigned to $\text{Ag}_{\sim 280}(\text{SBB})_{\sim 120}$ based on the results of ESI-MS. However, in this experiment, which was conducted in 2010 (the paper was published in 2011), the chemical composition was determined by a home-made MS apparatus without high resolution. In recent years, many ESI-MS apparatuses with high sensitivity and resolution have been developed. It is expected that the chemical composition of the stable $\text{Ag}_n(\text{SBB})_m$ NCS obtained in that study will be determined more accurately using such ESI mass spectrometers in the future.

Ag_n NCS are attractive because of their high quantum yield of luminescence¹⁶⁷ and high selectivity for carbon dioxide reduction.¹⁶⁸ Since we have also been working on catalytic application of metal NCS, we have continued to work on development of isolation techniques for Ag_n NCS also after the report of $\text{Ag}_n(\text{SBB})_m$ NCS. As a result, for example, in 2013, we successfully isolated $\text{Ag}_{44}(\text{SePh})_{30}$ (SePh = phenylselenolate) in collaboration with Pradeep and co-workers (Fig. 7(a)).¹¹⁹ DFT calculations by Häkkinen and co-workers showed that $\text{Ag}_{44}(\text{SePh})_{30}$ has a geometric structure similar to that of $[\text{Ag}_{44}(\text{SR})_{30}]^{4-}$ ($\text{SR} = \text{SPhF}_2$ or 4-mercaptobenzoic acid (*p*-MBA)), whose geometric structure was determined just before our

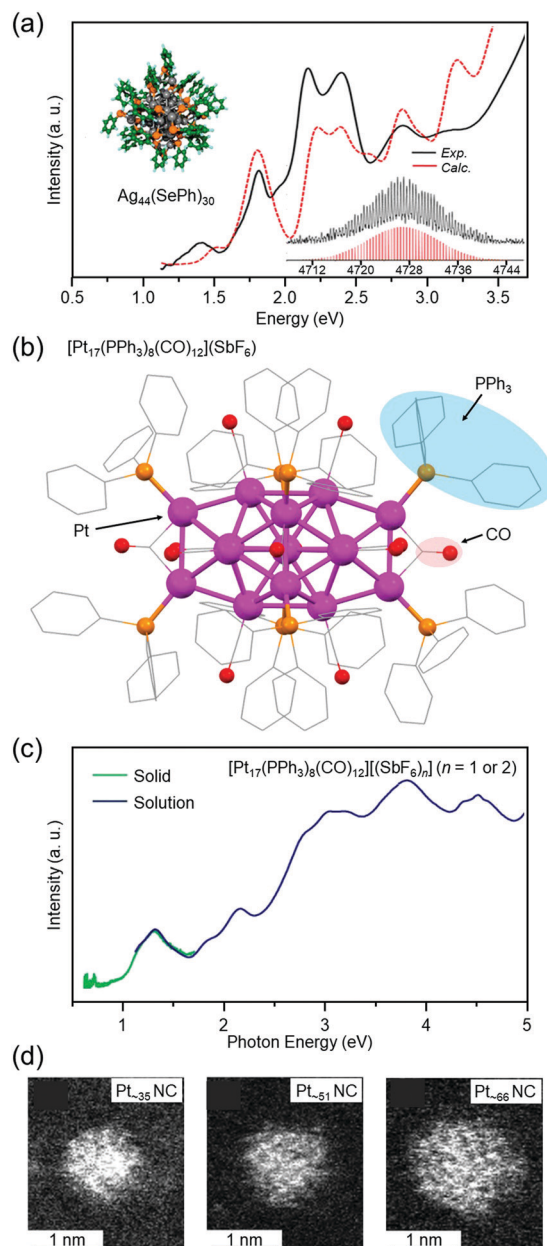


Fig. 7 (a) Optical absorption spectra, computationally relaxed structure, and expanded ESI mass spectrum of $\text{Ag}_{44}(\text{SePh})_{30}$.¹¹⁹ (b) Geometric structure of $[\text{Pt}_{17}(\text{PPh}_3)_8(\text{CO})_{12}](\text{SbF}_6)$ (SbF_6^- = hexafluoroantimonate) determined by SC-XRD.¹²⁰ (c) Optical absorption spectra of $[\text{Pt}_{17}(\text{PPh}_3)_8(\text{CO})_{12}][(\text{SbF}_6)_n]$ ($n = 1$ or 2) in dichloromethane solution (blue) and the solid state $[\text{Pt}_{17}(\text{PPh}_3)_8(\text{CO})_{12}][(\text{SbF}_6)_n]$ ($n = 1$ or 2) (green).¹²⁰ (d) High-angle scattering dark-field scanning transmission electron microscopy (HAADF-STEM) images of $\text{Pt}_n(\text{PET})_m(\text{CO})_l$ NCS with $n = \sim 35$ (left), ~ 51 (center), and ~ 66 (right).¹²¹ Reproduced with permission from ref. 119–121. Copyright 2013 American Chemical Society, 2017 American Chemical Society, and 2021 The Royal Society of Chemistry.

reports by Zheng and co-workers¹⁶⁹ and Bigioni and co-workers.¹⁷⁰ In addition to this NC, we recently succeeded in isolating several new Ag NCS¹⁷¹ using anions as templates (the paper is being prepared for submission).

Isolation of Pt_n NCS has been reported by Chini and co-workers and Longoni and co-workers.^{22,27} In their studies,



CO was used as the main ligand, but in later studies on Pt_n and its alloy NCs by Dahl, Slovokhotov, Zacchini, *etc.*, PR_3 was also used as a ligand.^{23,172,173} In these studies, the geometric structures of the metal NCs were determined by SC-XRD. For $Au_n(SR)_m$ NCs, in 2007, Kornberg, Ackerson and co-workers reported the geometric structure of $[Au_{102}(p\text{-MBA})_{44}]^0$, which was the first report of the geometric structure of $Au_n(SR)_m$ NCs determined by SC-XRD.¹⁷⁴ The advancement of research of Pt_n and its alloy NCs at that time was astonishing. However, synthesis and the subsequent handling of the reported Pt_n NCs had to be carried out under CO or an inert-gas atmosphere,^{23,172,173} which was not easy for those who were not used to using CO gas or Schlenk techniques. Therefore, we searched for Pt_n NCs that could be synthesized by mixing reagents in air, similar to $Au_n(SR)_m$ NCs.

When a Pt salt is dissolved in polyol and the resulting solution is heated, the Pt salt is reduced to generate Pt atoms (polyol reduction).^{175–179} The polyol generates CO, and the generated CO coordinates to the Pt surface. Addition of PR_3 to the solution of the resulting product is expected to give $Pt_n(PR_3)_m(CO)_l$ NCs without the need for a CO atmosphere. Indeed, we successfully isolated $[Pt_{17}(PPh_3)_8(CO)_{12}]^{2+}$ (PPh_3 = triphenylphosphine, $z = 1$ or 2) with atomic precision by this method (Fig. 7(b)).¹²⁰ The obtained NC had a geometric structure similar to previously reported $[Pt_{17}(PET_3)_8(CO)_{12}]^0$ (PET_3 = triethylphosphine)²³ and $[Pt_{13}Au_4(PPh_3)_8(CO)_{12}]^0$.¹⁷² Optical absorption spectroscopy revealed that $[Pt_{17}(PPh_3)_8(CO)_{12}]^{2+}$ ($z = 1$ or 2) has a discrete electronic structure similar to a small $Au_n(SR)_m$ NC (Fig. 7(c)).¹⁰²

For Pt_n NCs, we have also succeeded in obtaining a series of $Pt_n(PET)_m(CO)_l$ NCs ($n = \sim 35$, ~ 51 , and ~ 66) in high yields ($\sim 80\%$ yield for $Pt_{\sim 66}$) by changing the ligand from PPh_3 to PET (Fig. 7(d)).¹²¹ Unfortunately, the chemical compositions of these $Pt_n(PET)_m(CO)_l$ NCs cannot be controlled with atomic precision at present. Therefore, the chemical compositions and geometric structures of these NCs have not been accurately determined. Thus, we are now attempting to apply various size-focusing methods^{166,180} to $Pt_n(PET)_m(CO)_l$ NCs to isolate the $Pt_n(PET)_m(CO)_l$ NCs with a single chemical composition. If these efforts are successful and the chemical compositions and geometric structures of the isolated $Pt_n(PET)_m(CO)_l$ NCs are accurately determined, it is expected that a deeper understanding of the chemical compositions and geometric structures of the $Pt_n(PET)_m(CO)_l$ NCs that can be isolated will be obtained.

4.5. Isolation of metal NCs protected by functional ligands

For ligand-protected metal NCs, it is possible to impart new functions by controlling not only the metal core, but also the functional groups of the ligands. The solvent solubility, redox potential,^{181–184} luminescence properties,¹⁸⁵ optical activity,¹⁸⁶ and catalytic activity of $Au_n(SR)_m$ NCs are remarkably dependent on the functional group structure of SR. If $Au_n(SR)_m$ NCs with these characteristics are protected by SR, which changes its molecular structure in response to light irradiation, it is expected that the physicochemical properties of $Au_n(SR)_m$ NCs can be reversibly changed by light irradiation.

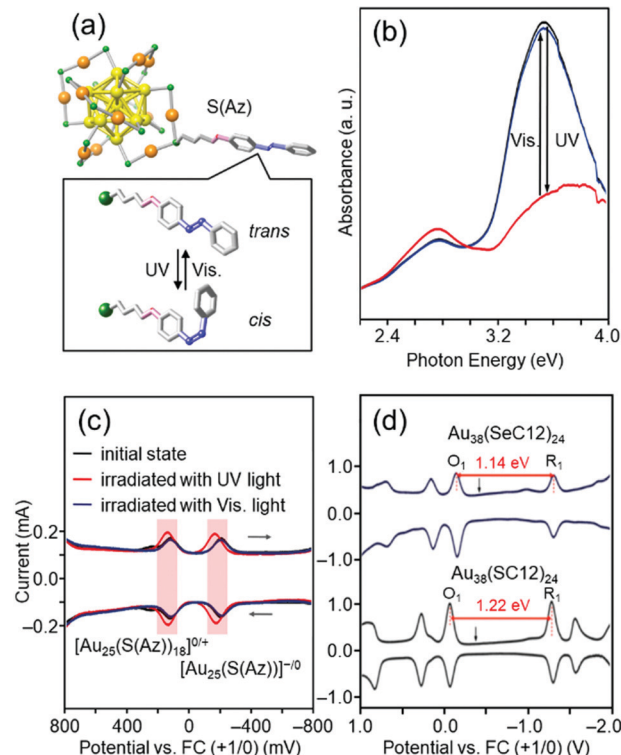


Fig. 8 (a) Proposed geometric structure of $[Au_{25}(S(Az))_{18}]^{-}$ and (b) the photoresponsive behavior in the optical absorption spectra of $[Au_{25}(S(Az))_{18}]^{-}$ in toluene solution in the energy range 2.2–4.0 eV.¹²² The visible photostationary state (blue line) overlaps with the initial state (black line). (c) Photoresponse of the differential pulse voltammetry (DPV) curve of $[Au_{25}(S(Az))_{18}]^{-}$.¹²² (d) DPV curves of $[Au_{38}(SeC12)_{24}]^0$ and $[Au_{38}(SC12)_{24}]^0$ in CH_2Cl_2 containing 0.1 M $(C_4H_9)_4NClO_4$.¹²⁶ O₁ and R₁ indicate the first oxidation and reduction peaks of the NCs, respectively. The black arrows indicate the solution open-circuit potentials. The red arrows indicate the electrochemical HOMO–LUMO gaps of $[Au_{38}(SeC12)_{24}]^0$ and $[Au_{38}(SC12)_{24}]^0$. Reproduced with permission from ref. 122 and 126. Copyright 2012 The Royal Society of Chemistry and 2013 American Chemical Society.

Azobenzene is a typical photoresponsive molecule.^{187,188} It changes from *trans* to *cis* by ultraviolet (UV)-light irradiation and from *cis* to *trans* by visible-light irradiation. When SR containing azobenzene is used as a protective ligand, $Au_n(SR)_m$ NCs, whose physicochemical properties repeatedly change by light irradiation, could be created. Accordingly, we have synthesized $Au_{25}(S(Az))_{18}$ protected by an azobenzene derivative thiolate (S(Az)) (Fig. 8(a)) to create such photoresponsive $Au_n(SR)_m$ NCs.¹²² When the obtained $Au_{25}(S(Az))_{18}$ was irradiated with light, S(Az) repeatedly underwent photoisomerization (Fig. 8(b)). In addition, the redox potential of $Au_{25}(S(Az))_{18}$ reversibly changed with such structural changes (Fig. 8(c)). These results demonstrated that the combination of azobenzene and Au_{25} NCs can be used to create $Au_n(SR)_m$ NCs with photoresponsive redox potentials.

$Au_{25}(SR)_{18}$ has a compact and symmetric metal core (Fig. 8(a)).^{153,154} Therefore, the metal-core structure of $Au_{25}(S(Az))_{18}$ remains almost unchanged even after isomerization of the ligand by photoirradiation. The change in the redox potential is considered to be because of the change in the dipole moment of the S(Az)



molecule caused by isomerization of azobenzene, which in turn changes the orbital energy of $\text{Au}_{25}(\text{S}(\text{Az}))_{18}$.^{189,190} In 2016, Jin and co-workers reported that $[\text{Au}_{28}(\text{SR})_{20}]^0$ can change the metal-core structure depending on the functional group structure of the ligand.¹⁹¹ Therefore, in the future, we will combine S(Az) with a $\text{Au}_n(\text{SR})_m$ NC whose metal core was easily deformable,¹⁹² and thereby create “intelligent” $\text{Au}_n(\text{SR})_m$ NCs in which the metal-core structure changes by light irradiation, leading to significant changes in the redox potential, luminescence properties, optical activity, and catalytic activity.

In the above study of $\text{Au}_{25}(\text{S}(\text{Az}))_{18}$, a new function was provided to the Au_n NCs by including a function in a functional group that was not directly bound to Au. However, changing the functional group directly bound to Au (*i.e.*, S) to Se is expected to change the binding mode between Au and the ligand,¹⁹³ thereby providing Au_n NCs with new physical and chemical properties. With this expectation, we precisely synthesized $[\text{Au}_{25}(\text{SeR})_{18}]^-$ (SeR = octaneselenolate (SeC8) or dodecaneselenolate (SeC12)),^{123,124} $[\text{Au}_{25}(\text{PET})_{18-y}(\text{SePh})_y]^-$ ($y = 1-18$),¹²⁸ $[\text{Au}_{25-x}\text{Cu}_x(\text{SeC8})_{18}]^-$ ($x = 1-9$),¹²⁵ and $[\text{Au}_{38}(\text{SeC12})_{24}]^-$ ¹²⁶ with selenolate (SeR) ligands. Studies on the obtained Au NCs revealed that changing the ligand from SR to SeR (1) changes the binding mode between Au and the ligand, thereby making the Au NCs more stable,¹²⁴ and (2) has the effect of narrowing the HOMO-LUMO gap of the Au NCs (Fig. 8(d)).^{124,126,128} We have also systematically synthesized $[\text{Au}_{25}(\text{SC8})_{18-y}(\text{TePh})_y]^-$ (TePh = phenyltelluroate, $y = 1-18$) containing a telluroate (TeR) in the ligand layer.^{127,128} We found that changing the ligand from SR to TeR further narrows the HOMO-LUMO gap of the Au NCs.^{127,128}

Based on a study of two-dimensional self-assembled monolayers,¹⁹⁴ changing the ligand from SR to SeR or TeR is expected to induce an increase in the conductivity between the Au core and the ligand. High conductivity is a desired feature in creation of single-molecule devices based on Au NCs,¹⁹⁵ and we hope to clarify this point in future studies by comparing the conductivities of $\text{Au}_n(\text{SR})_m$, $\text{Au}_n(\text{SeR})_m$, and $\text{Au}_n(\text{TeR})_m$ NCs.

4.6. Connection of ligand-protected metal NCs

Among ligand-protected metal NCs, icosahedral NCs have been reported to become constituent units of “clusters of clusters”.¹⁹⁶ For example, in the 1990s, Teo and co-workers reported synthesis of $[(p\text{-tol}_3\text{P})_{10}\text{Au}_{13}\text{Ag}_{12}\text{Cl}_7]^{2+}$ ($p\text{-tol}_3\text{P}$ = tris(4-methylphenyl)phosphine) consisting of two $\text{Au}_{13-x}\text{Ag}_x$ ¹⁹⁷ and $[(p\text{-tol}_3\text{P})_{12}\text{Au}_{18}\text{Ag}_{20}\text{Cl}_{14}]^0$ consisting of three $\text{Au}_{13-x}\text{Ag}_x$.¹⁹⁸ In 2007, I was also involved in isolation of $[\text{Au}_{25}(\text{PPh}_3)_{10}(\text{SCn})_5\text{Cl}_2]^{2+}$ (SCn = alkanethiolate, $n = 2-18$), which consists of two Au_{13} NCs.¹⁹⁹ Further development of techniques for connecting metal NCs could lead to fabrication of ultra-thin metal nanowires. Therefore, we have also been working on development of techniques to connect metal NCs.

The first connected structure that was successfully isolated was $[\text{Au}_{24}\text{Pd}(\text{PPh}_3)_{10}(\text{PET})_5\text{Cl}_2]^+$ (Fig. 9(a)).¹²⁹ This connected structure was size-selectively synthesized under appropriate conditions. In this connected structure, only one central Au atom is substituted by a different atom, unlike most of the connected structures reported in previous papers.²⁰⁰⁻²⁰³ DFT calculations by Jiang and co-workers showed that $[\text{Au}_{24}\text{Pd}(\text{PPh}_3)_{10}(\text{PET})_5\text{Cl}_2]^+$ NCs

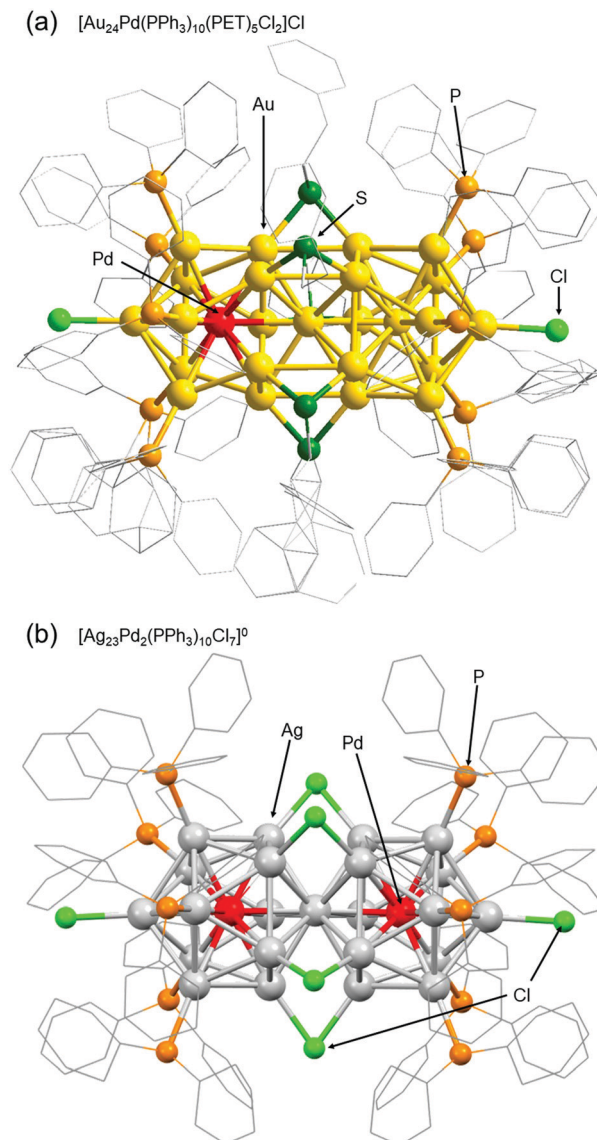


Fig. 9 Geometric structures of (a) $[\text{Au}_{24}\text{Pd}(\text{PPh}_3)_{10}(\text{PET})_5\text{Cl}_2]\text{Cl}$ and (b) $[\text{Ag}_{23}\text{Pd}_2(\text{PPh}_3)_{10}\text{Cl}_7]^0$.^{129,131} Reproduced with permission from ref. 129 and 131. Copyright 2018 The Royal Society of Chemistry and 2021 American Institute of Physics.

with these hetero bi-icosahedral cores have dipole moments in the NC.¹²⁹ Very recently, we also succeeded in synthesizing $[\text{Ag}_{23}\text{Pd}_2(\text{PPh}_3)_{10}\text{Cl}_7]^0$ consisting of two Ag_{12}Pd cores (Fig. 9(b)).¹³¹ Stability experiments and DFT calculations by Iwasa and co-workers indicated that it is very important to strengthen the framework of individual icosahedral cores and increase their stability to form the connected structures described by the chemical formula $[\text{Ag}_{23}\text{M}_2(\text{PPh}_3)_{10}\text{Cl}_7]^z$ ($\text{M} = \text{Ag}, \text{Pd}, \text{or Pt}, z = 2+ \text{ or } 0$).¹³¹

In the above study, two metal NCs were connected. On the other hand, in the study on $[\text{Au}_4\text{Pt}_2(\text{SR})_8]^0$ (SR = PET (Fig. 10(a)) or 2-chlorobenzylmercaptane (SCH_2PhCl)), we successfully connected them continuously in one dimension.¹³⁰ This was the second report of formation of a one-dimensional connected structure, following that reported by Maran and co-workers for $[\text{Au}_2\text{M}(\text{SR})_{18}]^0$ ($\text{M} = \text{Au}, \text{Hg}, \text{or Cd}, \text{SR} = \text{butanethiolate or}$



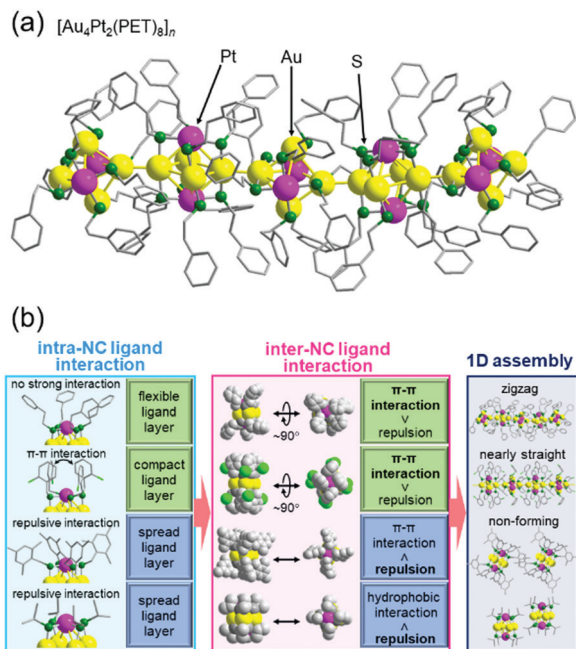


Fig. 10 (a) Crystal structure of $[\text{Au}_4\text{Pt}_2(\text{PET})_8]_n$. (b) Relationships between intra-NC ligand interactions, which are related to the distribution of the ligands within each NC, inter-NC ligand interactions, and one-dimensional assembly. Reproduced with permission from ref. 130. Copyright 2020 The Royal Society of Chemistry.

pentanethiolate)^{204–206} (note that although it was not mentioned in the paper,²⁰⁷ $[\text{Au}_4\text{Pd}_2(\text{PET})_8]^{10}$ reported by Wu and co-workers also seemed to form a similar connected structure). A series of studies with SR ligands revealed that it is important to design the ligand interaction within $[\text{Au}_4\text{Pt}_2(\text{SR})_8]^{10}$ to form one-dimensional connected structures and control their geometry (Fig. 10(b)).¹³⁰ The obtained one-dimensional connected structures have a semiconductor band gap, and the band gap size can be modulated by the ligand structure. These findings are expected to provide clear design guidelines for creation of metal-NC connected structures with desired geometric structures and physicochemical properties. We are now also working on connection of these metal NCs *via* linker molecules.^{62,208–213}

4.7. Development of high-resolution separation methods

Most of the ligand-protected metal NCs and their connected structures discussed in Sections 4.2–4.6 can be isolated by only mixing the reagents and washing the by-products in the ambient atmosphere. However, all NCs cannot be necessary isolated by such a simple process in the case of $\text{Au}_n(\text{SC12})_m$ NCs¹⁰⁹ described in Section 4.2. Therefore, it was necessary to develop new high-resolution separation techniques for ligand-protected metal NCs to isolate the metal NCs precisely and systematically, and thereby to understand the correlation between the chemical composition and the structure/properties.

PAGE (Fig. 3(a)), which is described in Section 3, is suitable for high-resolution separation of hydrophilic compounds, but it is not so suitable for high-resolution separation of hydrophobic compounds. On the other hand, HPLC (Fig. 11(a)) is suitable for

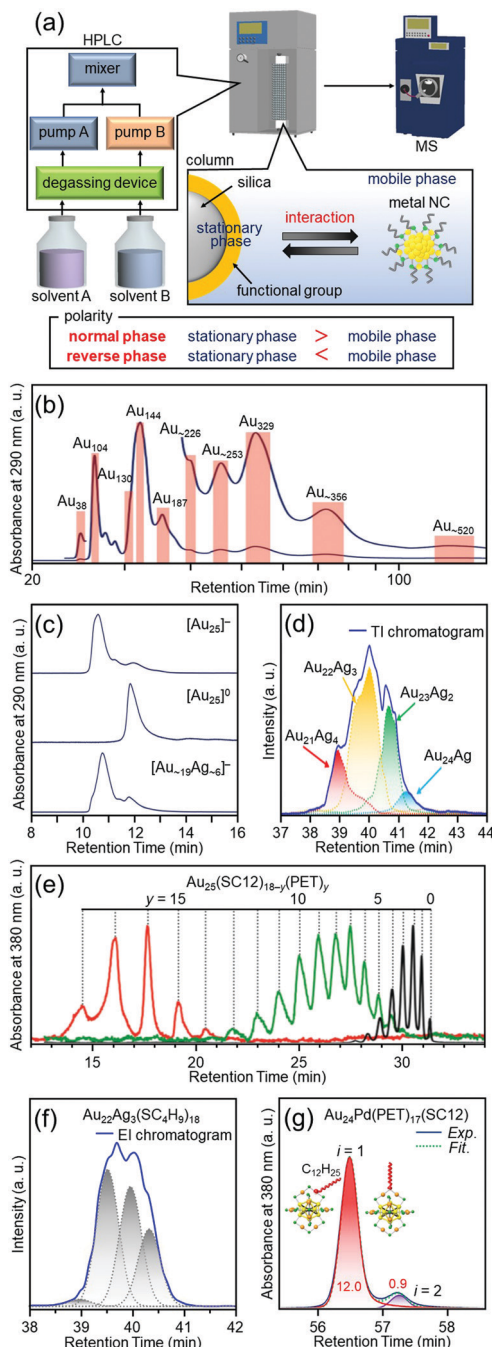


Fig. 11 (a) Schematic illustration of HPLC separation techniques.⁶¹ Chromatograms of (b) the crude sample of $\text{Au}_n(\text{SC12})_m$ NCs¹⁰⁹ and (c) $[\text{Au}_{25}(\text{SC12})_{18}]^-$, $[\text{Au}_{25}(\text{SC12})_{18}]^0$, and $[\text{Au}_{\sim 19}\text{Ag}_{\sim 6}(\text{SC12})_{18}]^-$.¹¹¹ (d) Extracted ion (EI) chromatogram of $[\text{Au}_{25-x}\text{Ag}_x(\text{SC}_4\text{H}_9)_{18}]^0$.¹³² Tl means total ion. (e) Chromatogram of $\text{Au}_{25}(\text{SC12})_{18-y}(\text{PET})_y$.¹³⁴ (f) EI chromatogram of $\text{Au}_{22}\text{Ag}_3$, which represents separation of several types of isomers.¹³² (g) Chromatogram of $\text{Au}_{24}\text{Pd}(\text{PET})_{17}(\text{SC12})$, which represents separation of coordination isomers.¹³⁷ Reproduced with permission from ref. 61, 109, 111, 132, 134 and 137. Copyright 2020 The Royal Society of Chemistry, 2015 American Chemical Society, 2010 The Royal Society of Chemistry, 2018 American Chemical Society, and 2013 American Chemical Society.

separation of both hydrophilic and hydrophobic compounds. Indeed, in 2003, Murray and co-workers succeeded in isolating



Au NCs, such as Au₃₈, Au₁₄₀, and Au₂₂₅ (the chemical composition is reported one in the papers), from a mixture of Au_n(SC6)_m NCs (SC6 = hexanethiolate) by RP-HPLC.^{144,214,215}

Inspired by the work of Murray and co-workers, we attempted to separate metal NCs, which were difficult to isolate by only washing the by-products, at the atomic and molecular levels by advanced use of RP-HPLC.⁵³ As a result, in addition to separation depending on the number of constituent atoms (Fig. 11(b)),^{108,109} ligand-protected metal NCs were also successfully separated depending on the charge state (Fig. 11(c)),¹¹⁰ alloy composition (Fig. 11(d)),^{132,138} ligand combination (Fig. 11(e)),^{133,134} and structural isomer (Fig. 11(f and g)).^{132,137} The separation techniques described in Fig. 11(f and g) led to not only control of the structural isomers, but also elucidation of the reaction mechanisms of the metal-exchange¹³⁸ and ligand-exchange¹³⁷ reactions. Although these high-resolution separations were achieved for hydrophobic metal NCs, in later studies, we also succeeded in separating hydrophilic metal NCs depending on their constituent atoms with high resolution by ion-pair chromatography¹³⁵ or hydrophilic-interaction chromatography.¹³⁶

In the future, we would like to use two-dimensional HPLC²¹⁶ for such high-resolution separations. It allows the use of two different separation modes. We expect that this will make it possible, for example, to first separate the products generated by ligand exchange depending on the number of exchanged ligands,¹³⁴ and then to separate the products of the same chemical composition depending on the optical isomers²¹⁷ generated by the position of the exchanged ligands.

4.8. Summary of control of ligand-protected metal NCs

We have succeeded in developing several new techniques for (1) isolation of large Au NCs, (2) isolation of alloy NCs, (3) isolation

of metal NCs consisting of Ag or Pt, (4) isolation of metal NCs protected by functional ligands, (5) connection of metal NCs, and (6) high-resolution separation of metal NCs. Other groups have also been simultaneously working in these areas, and they have developed several other isolation methods.^{7,218} Through these studies, the control techniques for ligand-protected metal NCs, that is, “bottom-up nanotechnology for complexes composed of metals and molecules”, has been dramatically enhanced in recent years. In the future, it is expected that the “science of ligand-protected metal NCs” will be further deepened by further development of these control techniques.

5. Toward creation of high-performance energy and environmental catalysts

When I started investigating gas-phase NCs in 1995, I was not concerned about the applications of metal NCs as materials. However, after I had deposited Ge_nO_m NCs on a support and observed their luminescence in air with my own eyes, I thought to contribute to society by applying metal NCs as materials in future studies. Because energy and environmental issues have become major global challenges, I decided to contribute to the field of energy and environmental materials.

5.1. Social problems to be solved by chemistry

In the 21st century, there are serious issues regarding energy and the environment, such as depletion of fossil fuels and destruction of the environment on a global scale. Therefore, chemistry in the 21st century is expected to address these issues, and thereby build a clean and sustainable society. The energy conversion system shown in Fig. 12 is one of the systems

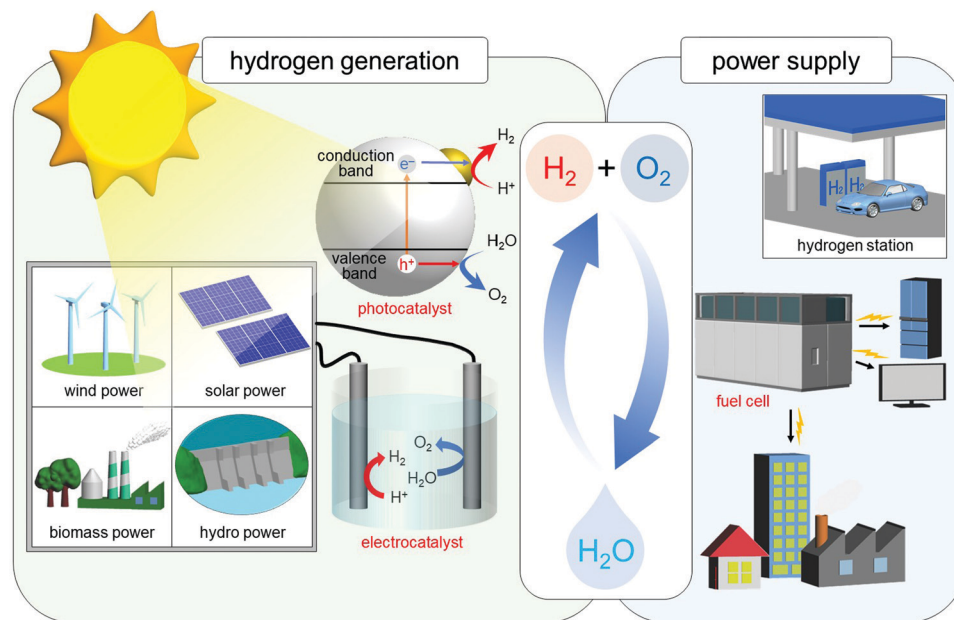


Fig. 12 Schematic of the energy-conversion system for constructing a sustainable society. Note that in addition to solar power, sunlight also produces wind, biomass, and hydro power. Reproduced with permission from ref. 56. Copyright 2021 The Royal Society of Chemistry.



necessary to build such a society. In this system, hydrogen (H_2) is produced by photocatalysts or electrocatalysts, and the obtained H_2 is converted to electricity by fuel cells. Using such a system, it is possible to obtain energy from only water and sunlight, which are almost unlimited on earth. In addition, this system allows the energy medium (H_2) to be circulated, thus preventing the issues of energy depletion. Furthermore, this system does not generate carbon dioxide, which leads to global warming. In this system, the properties of the photocatalyst,²¹⁹ electrocatalyst,^{63,220–232} and fuel cell^{60,224,233–238} are highly dependent on the chemistry. Therefore, modern chemistry is expected to improve the functions of these components. Considering the fact that the shift to H_2 society cannot be realized immediately, it is necessary to improve the current system at the same time. In other words, automotive exhaust-gas purification catalysts,²³⁹ which inhibit emission of harmful gases, need to be further improved. Therefore, we have been working on development of highly functionalized water-splitting photocatalysts, fuel-cell electrocatalysts, and automotive exhaust-gas purification catalysts (Fig. 1(d and e)) using precise metal NCs, in addition to developing flexible and precise control techniques for metal NCs.

5.2. Our strategy for activation of energy and environmental catalysts

For preparation of water-splitting photocatalysts, fuel-cell electrocatalysts, and automotive exhaust-gas purification catalysts, metal nanoparticles (NPs) are generally loaded on the support by impregnation (Fig. 13(a))²⁴⁰ or photodeposition method (Fig. 13(b)).²⁴¹ In the impregnation method, metal NPs are immobilized on the support by stirring the support and metal salt in solution followed by calcination (Fig. 13(a)). In the photodeposition method, the support is dispersed in a solution containing the precursor metal salts of the metal NPs, which is then irradiated with light. The electrons and holes generated by photoexcitation are used to reduce or oxidize the precursor metal salts of the metal NPs, thereby immobilizing the metal NPs on

the surface of the support (Fig. 13(b)). These loading methods are widely used because of their simplicity. However, the size distributions of the metal NPs obtained by these methods are wide because the metal atoms are aggregated on the support. Therefore, it is extremely difficult to load precisely controlled metal NPs on the support by these methods.

On the other hand it is possible to load metal NCs with a controlled particle size on the support by adsorbing ligand-protected metal NCs synthesized in the liquid phase onto the support and then removing the ligand by calcination (Fig. 13(c)).^{242–244} If the chemical composition of the loaded metal NCs can be controlled with atomic precision, it could be possible to create high-performance heterogeneous catalysts based on selective loading of highly active metal NCs.⁶⁶ For atomically controlled loaded metal NCs, a deep understanding of the electronic/geometric structure of the loaded metal NCs and the electronic/geometric interaction between the loaded-metal NCs and the catalyst surface can be obtained by various high-resolution analyses and theoretical calculations (Fig. 14(a)).^{11,245–255} It is expected that this will provide a deep understanding of the important factors for high activation and clear design guidelines for further activation (Fig. 14(b)). Design and creation of appropriate metal NCs based on the obtained information are expected to lead to further functionalization of water-splitting photocatalysts, fuel-cell electrocatalysts, and automotive exhaust-gas purifying catalysts (Fig. 14). However, when we started this work, there were no examples of atomically precise control of loaded metal NCs on these catalysts using ligand-protected metal NCs as precursors.

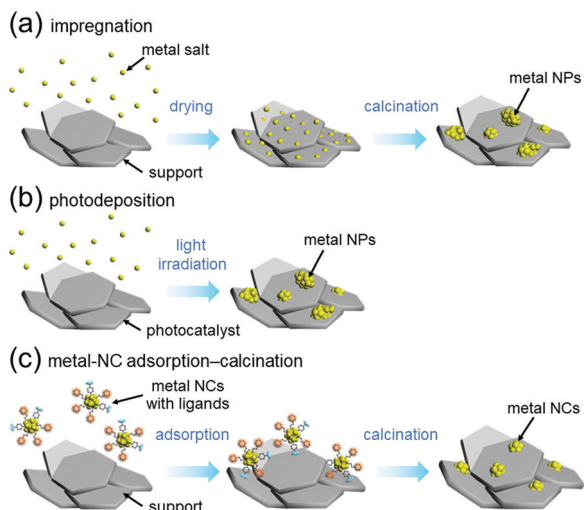


Fig. 13 Comparison of (a) impregnation, (b) photodeposition, and (c) metal-NC adsorption–calcination methods.

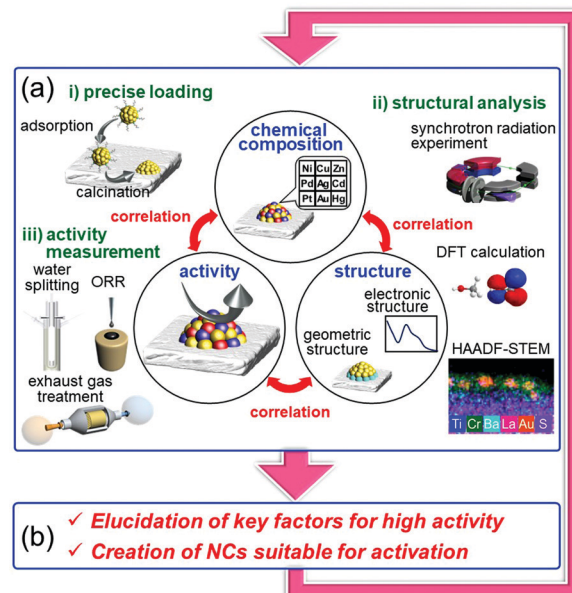


Fig. 14 Our strategy for activation of energy and environmental catalysts: elucidation of (a) the correlation between the chemical composition, structure, and activity and (b) the key factors for the activity and therefore for creation of NCs suitable for activation. This cycle is repeated to create high-performance energy and environmental catalysts.



5.3. Activation of water-splitting photocatalysts

Using photocatalytic reactions (Fig. 12),^{256–267} it is possible to produce H₂ directly from water and sunlight. Therefore, water-splitting photocatalysis is considered to be one of the cleanest energy production reactions. However, at present, the energy conversion efficiency from sunlight to H₂ (STH) is only about 1.1%.²⁶⁸ To realize practical use of water-splitting photocatalysts, it is essential to improve the STH to about 5–10% in the future.²⁶⁴

In many cases, metal/metal-oxide NPs/NCs (cocatalysts) should be loaded on water-splitting photocatalysts as the active sites (Fig. 12). Controlling their particle sizes is very effective for improving the water-splitting activity.^{64,258,269} Therefore, we have focused on (1) precisely controlling the loaded metal NCs on the photocatalyst using precise ligand-protected metal NCs as precursors, (2) elucidating the details of how control of the cocatalyst affects the water-splitting activity, and (3) actually improving the functionality of the water-splitting photocatalyst.

The first step was to precisely load Au_n NCs onto BaLa₄Ti₄O₁₅ (an UV-light-driven water-splitting photocatalyst),²⁷⁰ which is one of the most advanced photocatalysts developed by our collaborator, Kudo and co-workers.²⁵⁶ We successfully determined the appropriate precursor NCs for precise loading (Fig. 15(a)), the appropriate adsorption conditions, and the appropriate calcination conditions.^{271,272} Activity measurement (Fig. 15(b)) of the obtained Au₂₅ NC-loaded BaLa₄Ti₄O₁₅ (Au₂₅/BaLa₄Ti₄O₁₅) revealed that the small Au cocatalyst accelerates not only the HER, but also the oxygen (O₂)-photoreduction reaction, which is one of the reverse reactions.²⁷³ Therefore, we next attempted to form a layer to prevent the reverse reaction on the Au₂₅ cocatalyst. Maeda, Domen and co-workers reported that formation of a chromium oxide (Cr₂O₃) layer is extremely effective for preventing the reverse reaction.^{244,274–283} However, using their formation method, it was difficult to form the Cr₂O₃ layer on the Au₂₅ cocatalyst while maintaining the number of constituent atoms of the cocatalyst. Therefore, we developed a new method for formation of the Cr₂O₃ layer.²⁷³ Specifically, the Cr₂O₃ layer was first loaded on BaLa₄Ti₄O₁₅ (Cr₂O₃/BaLa₄Ti₄O₁₅), and then Au₂₅(SG)₁₈ was adsorbed on Cr₂O₃/BaLa₄Ti₄O₁₅. The obtained catalyst (Au₂₅(SG)₁₈/Cr₂O₃/BaLa₄Ti₄O₁₅) was calcined in an electric furnace under low vacuum to remove the ligands of Au₂₅(SG)₁₈ and embed Au₂₅ into the Cr₂O₃ layer (Fig. 15(c)). The obtained Cr₂O₃/Au₂₅/BaLa₄Ti₄O₁₅ photocatalyst showed high water-splitting activity. Thus, we succeeded in creating a highly active water-splitting photocatalyst with the advantage of high HER activity of fine Au NCs by combining miniaturization of the Au cocatalyst with formation of the reverse-reaction prevention layer (Fig. 15(c)). In a later study using Au₂₅ as the cocatalyst, it was found that even higher activity can be obtained by substituting one Pt atom for one Au atom in the Au₂₅ cocatalyst.²⁸⁴ On the basis of the various experimental results, we concluded that Au₂₄Pt/BaLa₄Ti₄O₁₅ showed higher water-splitting activity than Au₂₅/BaLa₄Ti₄O₁₅, mainly because (1) photoexcited electrons can more easily transfer from BaLa₄Ti₄O₁₅ to the metal NCs in Au₂₄Pt/BaLa₄Ti₄O₁₅ than in Au₂₅/BaLa₄Ti₄O₁₅ and (2) Au in Au₂₄Pt is more negatively charged than Au in Au₂₅ (Fig. 15(d)).

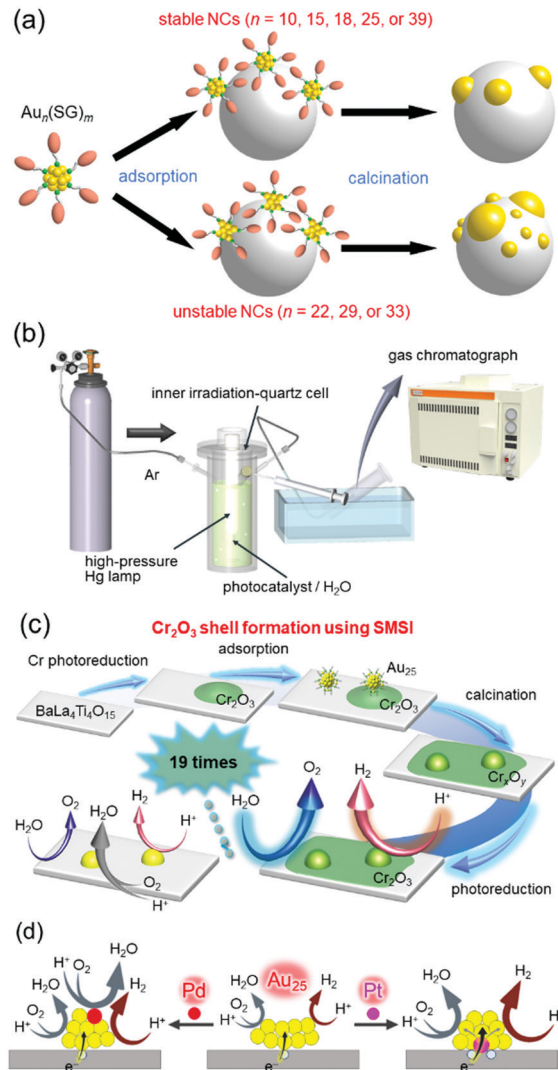


Fig. 15 (a) Au_n(SG)_m suitable for loading Au_n NCs on the photocatalyst while retaining the number of constituent atoms.²⁷² (b) Schematic of the system used to estimate the photocatalytic activity.²⁷³ (c) Schematic of preparation of Cr₂O₃/Au₂₅/BaLa₄Ti₄O₁₅ and its water-splitting activity.²⁷³ (d) Proposed structures of Au₂₄M/BaLa₄Ti₄O₁₅ for M = Au, Pd, and Pt and the effect of the geometric/electronic structure on each reaction over metal-NC cocatalysts.²⁸⁴ Reproduced with permission from ref. 272, 273 and 284. Copyright 2015 American Chemical Society, 2018 American Chemical Society, and 2019 American Chemical Society.

In the above study, Au was used as the metal element for the cocatalyst. However, based on the volcano plot²⁸⁵ for H₂ adsorption and desorption, rhodium (Rh) is predicted to have higher catalytic activity than Au for the HER. Therefore, it is expected that highly active water-splitting photocatalysts can be created by loading a fine cocatalyst consisting of rhodium and chromium oxides (Rh_{2-x}Cr_xO₃)^{286–289} on the photocatalysts. Therefore, we also attempted to load fine Rh_{2-x}Cr_xO₃ NCs on BaLa₄Ti₄O₁₅. Unfortunately, there have been no reports of precise synthesis of Rh_{2-x}Cr_xO₃ NCs. Therefore, in this experiment, we used Rh₂(SG)₂ complexes (Rh-SG) as precursors, which were aggregated on BaLa₄Ti₄O₁₅ and then calcined. As a result, we succeeded in loading monodispersed Rh_{2-x}Cr_xO₃ NCs



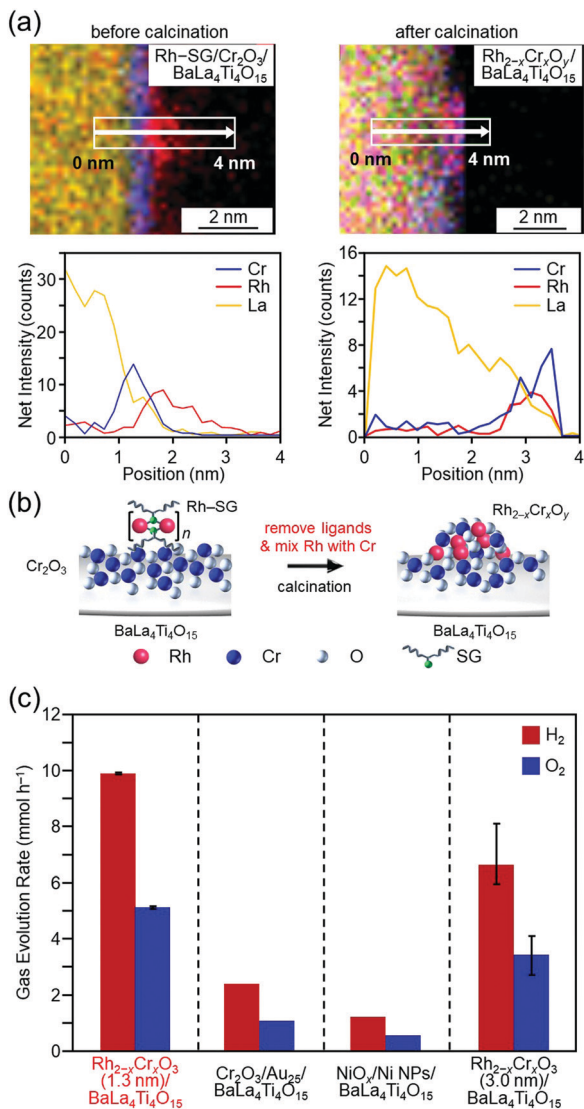


Fig. 16 HAADF-STEM images of Rh-SG complexes adsorbed on the Cr₂O₃/BaLa₄Ti₄O₁₅ photocatalyst before and after calcination (Rh-SG/Cr₂O₃/BaLa₄Ti₄O₁₅ and Rh_{2-x}Cr_xO_y/BaLa₄Ti₄O₁₅, respectively). (a) Line analysis of elemental mapping of Rh-SG/Cr₂O₃/BaLa₄Ti₄O₁₅ (left) and Rh_{2-x}Cr_xO_y/BaLa₄Ti₄O₁₅ (right). (b) Schematic of the phenomenon that occurred during the calcination process. (c) Comparison of the gas-evolution rates over different photocatalysts: Rh_{2-x}Cr_xO₃ (1.3 nm)/BaLa₄Ti₄O₁₅ (0.09 wt% Rh and 0.10 wt% Cr), Cr₂O₃/Au₂₅/BaLa₄Ti₄O₁₅ (0.10 wt% Au and 0.50 wt% Cr), Ni NiO_x/NPs/BaLa₄Ti₄O₁₅ (0.50 wt% Ni), and Rh_{2-x}Cr_xO₃ (3.0 nm)/BaLa₄Ti₄O₁₅ (0.10 wt% Rh and 0.15 wt% Cr). Reproduced with permission from ref. 290. Copyright 2020 Wiley-VCH.

with particle sizes of 1.3 ± 0.3 nm on BaLa₄Ti₄O₁₅ (Fig. 16(a and b)).²⁹⁰ The obtained photocatalyst showed the highest apparent quantum yield of 16% for BaLa₄Ti₄O₁₅ (excitation wavelength = 270 nm, Fig. 16(c)). The loading method of Rh_{2-x}Cr_xO₃ NCs developed in this study can in principle be applied to other photocatalysts. In addition, Rh_{2-x}Cr_xO₃ NCs are useful cocatalysts for many water-splitting photocatalysts.^{286–289,291,292} In the future, it is expected that high quantum yields will be obtained for many other water-splitting photocatalysts using this technique.

For practical use of a water-splitting photocatalyst, it is essential that the photocatalyst causes water splitting by visible light (visible-light-driven water-splitting photocatalyst), which accounts for about 40% of solar energy. Currently, there are only a few semiconductor photocatalysts that enable complete water splitting in one step using visible light.²⁶⁴ However, overall water splitting by visible light can also be achieved by photocatalytic systems using a two-step reaction called the Z-scheme, which mimics plant photosynthesis.^{257,260} We are currently working on improving the functionality of both of these visible-light-driven water-splitting photocatalysts based on the knowledge obtained in research of BaLa₄Ti₄O₁₅.

5.4. Activation of ORR electrode catalysts

The polymer electrolyte fuel cell (PEFC, Fig. 12) is already in practical use. However, a significant amount of Pt is used in current PEFCs, which makes the current PEFCs extremely expensive. This has impeded widespread use of PEFCs, and thereby transition to a society using H₂ as an energy source (Fig. 12) has stagnated. Even if the cost of PEFCs can be lowered by improving the parts other than the Pt catalyst, we will inevitably encounter a shortage of Pt if the current Pt catalyst is used. Therefore, it is essential to reduce the amount of Pt used in PEFC electrocatalysts to build a clean and sustainable society.

In PEFCs, the ORR at the cathode is the rate-limiting reaction. Pt NPs with particle sizes of 2–3 nm loaded on carbon black (CB) (Pt NPs/CB) are widely used ORR cathodes. Recently, the following facts were reported by Yamamoto and co-workers,^{293,294} Nakajima and co-workers,²⁹⁵ and Fischer and co-workers²⁹⁶ for smaller ~1 nm particle size Pt_n NCs: (1) the Pt_n NCs have higher mass activity than the currently used Pt NPs and (2) the mass activity of the Pt_n NCs dramatically varies depending on the number of constituent atoms. Therefore, it is expected that if highly active ~1 nm Pt_n NCs can be size-selectively synthesized and then loaded on CB, highly active ORR catalysts can be created, which will lead to cost reduction of PEFCs.

As described in Section 4.4, we have already developed a method to isolate [Pt₁₇(PPh₃)₈(CO)₁₂]²⁺ ($z = 1$ or 2) with atomic precision by simple manipulation in air.¹²⁰ In addition, by further modification of this method, we successfully synthesized a series of Pt_n(PET)_m(CO)_i NCs ($n = \sim 35, \sim 51, \text{ and } \sim 66$) in a size-selective manner with a narrow particle-size distribution.¹²¹ Pt_n/CB catalysts were then prepared by adsorbing these Pt_n NCs on CB and partially removing the ligands at appropriate temperatures (Fig. 17(a and b)). We found that the mass activities of Pt_n/CB ($n = \sim 35, \sim 51, \text{ and } \sim 66$) were higher than that of commercial Pt NPs/CB (Fig. 17(c)).¹²¹ Among the Pt_n/CB catalysts, Pt_{~51}/CB showed 2.1 times higher mass activity than the commercial Pt NPs/CB catalyst. On the basis of estimation of the electrochemical activity area and specific activity, we concluded that Pt_n/CB shows higher ORR mass activity than commercial Pt NPs/CB because both the percentage of surface Pt atoms and specific activity of each surface Pt atom are higher for Pt_n/CB than for commercial Pt NPs/CB.



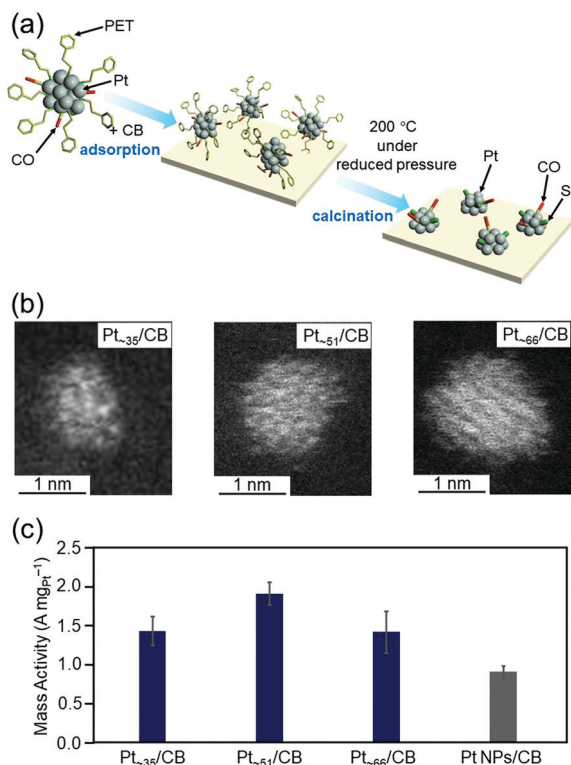


Fig. 17 (a) Schematic illustration of controlled loading of ~ 1 nm Pt NCs on CB. (b) HAADF-STEM images of Pt_{*n*}/CB with *n* = ~ 35 (left), ~ 51 (center), and ~ 66 (right). (c) ORR mass activity at 0.6 V vs. the reversible hydrogen electrode for Pt_{~35}/CB, Pt_{~51}/CB, Pt_{~66}/CB, and commercial Pt NPs/CB. In this figure, the average values obtained by performing the experiments three times are reported. Reproduced from ref. 121. Copyright 2021 The Royal Society of Chemistry.

Regarding the ORR activity of Pt NCs, it has been reported that the presence of Pt atoms in the generalized coordination number (GCN) region of 7.7–8.3 on the surface of the Pt NCs induces high activity.^{297,298} The GCN is a coordination number that takes into account the second nearest neighbor atoms in addition to the nearest neighbor atoms. When surface atoms are removed and concave cavities are created on the Pt-NC surface, Pt atoms with such GCNs are generated on the surface of the Pt NC. Unfortunately, in our previous studies, we did not obtain detailed information about the geometric structures of Pt_{*n*}/CB (*n* = ~ 35 , ~ 51 , and ~ 66). Therefore, we are currently working on elucidating the details of their geometric structures by obtaining high-resolution high-angle scattering dark-field scanning transmission electron microscopy (HAADF-STEM) images and analyzing the GCNs in the obtained geometric structures by DFT calculations. This is expected to provide a deep understanding of the factors that lead to high activity and clear design guidelines for further enhancement of the activity.

5.5. Activation of automobile exhaust-gas purification catalysts

It is expected that electric vehicles and fuel-cell vehicles will be widely used in the future. However, conventional vehicles (gasoline and diesel vehicles) are also estimated to account for 30–50% of the total market until 2030. Pt is widely used as

an exhaust-gas purification catalyst, along with Pd and Rh, because of its catalytic activity to purify the harmful substances contained in exhaust gases.²³⁹ However, reduction of Pt use is a critical issue at present because Pt is a rare and expensive precious metal. In addition, with the widespread use of automobiles, such as hybrid vehicles, where the engine is repeatedly stopped and restarted, high activation at low temperatures is also an issue that needs to be overcome.²³⁹

Studies on Pt-free catalysts are now being performed to overcome the above issues. However, the activity and durability of Pt are overwhelmingly superior to those of non-metallic materials. Reducing the particle size of the Pt catalyst is an extremely effective way to both reduce the amount of Pt used and improve the functionality of the Pt catalyst. Indeed, Anderson and co-workers demonstrated that Pt_{*n*} NCs (*n* = 14–18) loaded on alumina (Pt_{*n*}/Al₂O₃) have high catalytic activity for CO oxidation, one of the automotive exhaust-gas purification reactions, using a vacuum apparatus.⁷⁰

We have prepared Pt₁₇/γ-Al₂O₃ catalysts by adsorbing [Pt₁₇(PPh₃)₈(CO)₁₂]²⁺ (*z* = 1 or 2)¹²⁰ on γ-Al₂O₃ and then partially removing the ligands at appropriate temperatures.²⁹⁹ HAADF-STEM images and XAFS measurements showed that Pt does not become oxidized³⁰⁰ in Pt₁₇/γ-Al₂O₃, and it maintains the skeletal structure of Pt₁₇ NC (Fig. 18(a–c)). To evaluate the catalytic performance of the obtained catalysts under conditions as close as possible to those of actual vehicles, Pt₁₇/γ-Al₂O₃ was coated on a honeycomb substrate and the catalytic performance for CO or propylene (C₃H₆) oxidation was evaluated under these conditions (Fig. 18(d)). The measurements of the HAADF-STEM images and catalytic activity were performed in collaboration with Nagaoka of Johnson Matthey Japan.²⁹⁹ We found that Pt₁₇/γ-Al₂O₃ has higher CO oxidation activity at each temperature than Pt NPs/γ-Al₂O₃ loaded with Pt NPs with average particle size of 3.1 ± 3.1 nm by the impregnation method, and, therefore, it is capable of purifying CO at lower temperature than Pt NPs/γ-Al₂O₃ (Fig. 18(e)). Previous studies have suggested that in oxidation of CO, CO and O₂ are activated at Pt on the terrace and Pt in the step, respectively.^{72,301,302} In the structures shown in Fig. 18(b and c), most of the terrace Pt atoms are located near the Pt step. This seems to be one of the factors that leads to effective reaction of CO and O₂, that is, oxidation of CO, over Pt₁₇/γ-Al₂O₃. High oxidation capacity of Pt₁₇/γ-Al₂O₃ at low temperatures was also observed for oxidation of C₃H₆. These results indicate that Pt₁₇/γ-Al₂O₃ is a promising catalyst for reducing the amount of Pt used and improving the catalytic performance at low temperatures.

To use Pt₁₇/γ-Al₂O₃ as a practical catalyst, it is necessary to investigate the catalytic activity and durability of the catalyst with the loading amount in the actual vehicle and at the exhaust-gas mixture ratio^{302–305} during actual operation. We are currently performing measurements under such conditions in collaboration with industry.

5.6. Elucidation of the ligand-removal mechanism

In our catalyst preparation process (Fig. 13(c)), (1) metal NCs are precisely synthesized using ligands and then (2) the



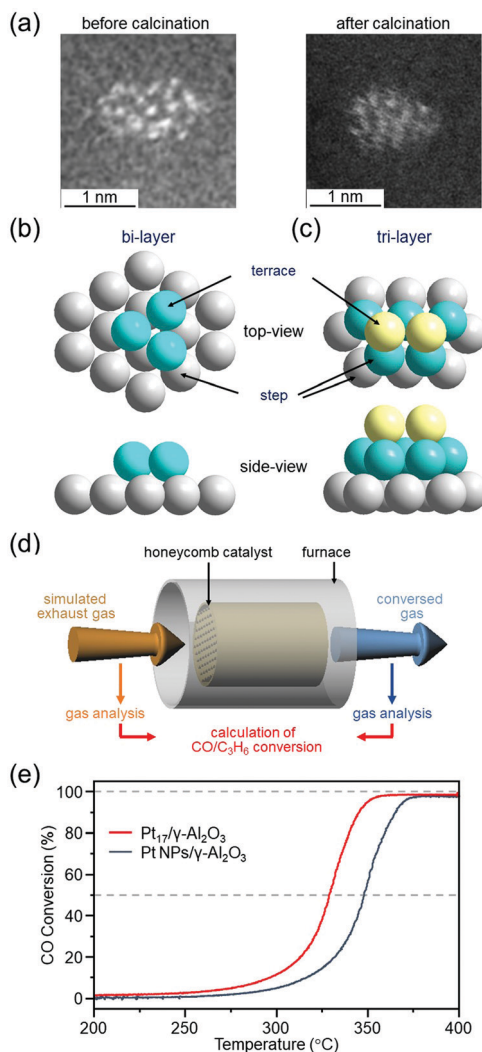


Fig. 18 (a) HAADF-STEM images of $\text{Pt}_{17}(\text{PPh}_3)_8(\text{CO})_{12}/\gamma\text{-Al}_2\text{O}_3$ (left) and $\text{Pt}_{17}/\gamma\text{-Al}_2\text{O}_3$ (right). Two proposed structures for Pt_{17} on $\gamma\text{-Al}_2\text{O}_3$ based on the HAADF-STEM images of $\text{Pt}_{17}/\gamma\text{-Al}_2\text{O}_3$: (b) bi-layer and (c) tri-layer. In (b) and (c), both top and side views are shown. (d) Schematic illustration of estimation of CO and C_3H_6 conversion over $\text{Pt}_{17}/\gamma\text{-Al}_2\text{O}_3$ or Pt NPs/ $\gamma\text{-Al}_2\text{O}_3$ coated on a cordierite honeycomb substrate. (e) CO conversion over $\text{Pt}_{17}/\gamma\text{-Al}_2\text{O}_3$ (red curve) and Pt NPs/ $\gamma\text{-Al}_2\text{O}_3$ (gray curve). Reproduced from ref. 299. Copyright 2020 The Royal Society of Chemistry.

resulting NCs are adsorbed on a support. The presence of the ligand on the surface of the metal NCs generally leads to a decrease in the activity^{306,307} because it prevents the approach of the reactants to the metal-NC surface and induces modulation of the electronic structure of the metal NCs. Therefore, (3) some or all of the ligands are removed by calcination to produce high-activity metal NCs. However, removal of the ligands simultaneously induces aggregation of the metal NCs and a resulting decrease in the catalytic activity. Thus, in calcination, it is extremely important to remove only the ligands while maintaining the number of constituent atoms of the metal NCs (Fig. 19(a)).

Regarding removal of the ligands, in the studies reported in Section 5.3–5.5, the calcination conditions were determined by confirming the presence or absence of ligands in the catalysts

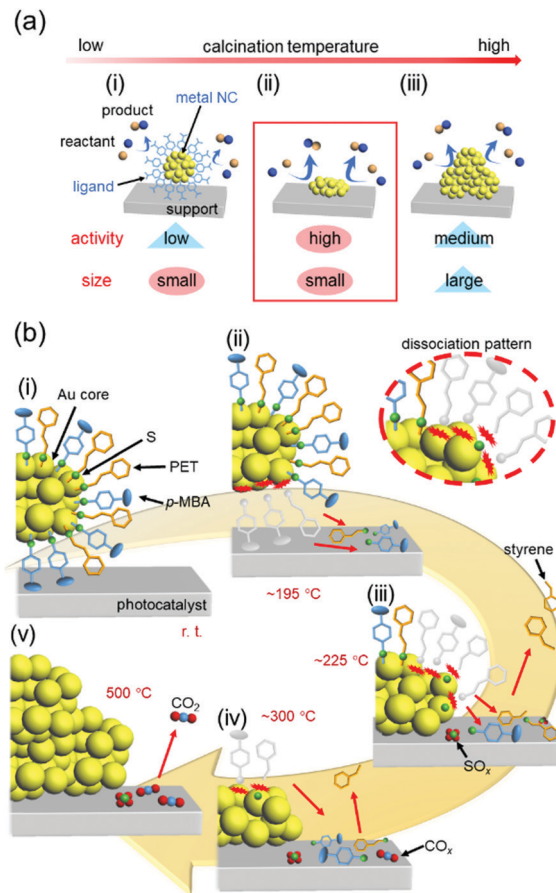


Fig. 19 (a) Schematic illustrations of the typical phenomena caused by increasing the calcination temperature in metal-oxide-supported ligand-protected metal NCs: (i) a small size is maintained but the activity is low, (ii) high activity is achieved while maintaining a small size, and (iii) the activity decreases owing to aggregation. (b) Proposed phenomenon occurring at each temperature during calcination of $\text{Au}_{25}(\text{PET})_9(\text{p-MBA})_9/\text{BaLa}_4\text{Ti}_4\text{O}_{15}$ at (i) room temperature, (ii) ca. 195 °C, (iii) ca. 225 °C, (iv) ca. 300 °C, and (v) 500 °C. Reproduced from ref. 308. Copyright 2021 Wiley-VCH.

after calcination by X-ray photoelectron spectroscopy (XPS) and XAFS. However, through these studies, we felt that it was essential to understand the reaction mechanisms occurring during calcination to perform calcination under appropriate conditions and thereby create high-performance heterogeneous catalysts. Therefore, we recently worked on elucidation of the reaction mechanism during ligand calcination.³⁰⁸ Specifically, the ligand desorption process in $\text{Au}_{25}(\text{PET})_9(\text{p-MBA})_9/\text{BaLa}_4\text{Ti}_4\text{O}_{15}$ during calcination was followed by five techniques: direct insertion probe-mass spectrometry, XAFS, Fourier-transform infrared spectroscopy, XPS, and TEM observation. As a result, we succeeded in elucidating that the ligand desorption process consists of dissociation of the ligand on the metal-NC surface, adsorption of the generated compound on the support, and desorption of the compound from the support, and at what temperatures these processes occur (Fig. 19(b)). Based on the obtained knowledge, we also succeeded in developing a method to form the Cr_2O_3 shell on the surface of Au NCs while preventing



their aggregation, and thereby obtained a highly active and stable water-splitting photocatalyst.³⁰⁸

The desorption temperature of SR from the Au-NC surface is strongly related to the interactions between the ligands on the Au-NC surface, as well as the strengths of the Au-S and S-C bonds.³⁰⁸ In addition, the temperature for desorption from the support is related to the strength of the compound-support interaction.³⁰⁹ Furthermore, the easiness of dissociation/desorption of the ligands and the resulting aggregation of the Au NCs slightly varies depending on the calcination atmosphere.^{309,310} However, previous calcinations with different SR functional groups, supports, and atmospheres have often suggested occurrence of the phenomena in Fig. 19(b).^{243,271,309,311–314} Therefore, it is assumed that a phenomenon similar to Fig. 19(b) also occurs in calcination of the other Au_n(SR)_m NCs/metal oxides, although the temperature at which it occurs is different. There had not been a unified view on the behavior in Au_n(SR)_m NCs/metal oxides during calcination before this study.^{312,315} We succeeded in elucidating the details of the phenomena occurring during calcination of Au₂₅(PET)₉(*p*-MBA)₉/metal-oxide catalysts by combining five measurements. The knowledge gained in this study is expected to provide clear design guidelines for creation of high-performance heterogeneous catalysts.

5.7. Summary of activation of energy and environmental catalysts using controlled metal NCs

We succeeded in precisely controlling the metal NCs on the support with atomic precision using precise ligand-protected metal NCs as precursors. Studies of the obtained energy and environmental catalysts revealed the correlation between the number of constituent atoms/chemical composition of the loaded metal NCs and the material function of each catalyst. Through these studies, each catalyst has been successfully highly functionalized. Specifically, we have achieved producing the highest water-splitting activity for BaLa₄Ti₄O₁₅,²⁹⁰ one of the most advanced water-splitting photocatalyst, creating electrode catalysts with higher ORR activity than the Pt catalysts currently used in fuel cells¹²¹ and improving the functionality of automotive exhaust-gas purification catalysts.²⁹⁹ These results strongly suggest that the bottom-up nanotechnology established in the metal-NC field can contribute to construction of a clean and sustainable society, that is, provide solutions for the issues faced by modern society.

For these catalytic applications, in the future, we will work on alloying³¹⁶ with various metal elements for the metal NC contained in each catalyst. In addition, the geometric and electronic structures of the loaded metal NCs during the reaction³¹⁷ will be revealed to gain a better understanding of the structure-property relationship of the loaded metal NCs and provide design guidelines for further enhancement of their activity. We believe that these catalysts could be further enhanced by repeating the cycle shown in Fig. 14 while incorporating the obtained knowledge. Regarding these applied studies, we would also like to continue our study while deepening both the science and technology of metal NCs.

6. Conclusions

Nanoscale materials have the following interesting and important features: (1) they possess great potential and value that cannot be found in bulk materials and (2) the nanoscale is the level of emergence of material functions, and thus if materials can be controlled at the nanoscale, we can control the properties of the larger materials. In this paper, I have described the contents of and motivation for my own studies in this field. The knowledge and skills required to study metal NCs are likely to change in the future, depending on the stage of the study and the situation in the world. However, the importance of understanding and controlling metal NCs will not change because they are deeply related to understanding and controlling various materials in the fields of energy, environment, and materials. I hope that the science and technology in the metal-NC field will be further advanced in the future through collaboration among researchers in many fields across physics, chemistry, and biology.

7. Outlook

The studies that we will perform in the future or are currently being performed have already been described at the end of each section. Therefore, in this section, I will describe the studies to be performed in the future for further development of the field of metal NCs (Fig. 20).

7.1. Elucidation of the synthesis mechanism

Recent work by Xie and co-workers has revealed the details of the reaction mechanism for synthesis of hydrophilic SR-protected metal NCs.^{318,319} In addition, recent studies by Pradeep and co-workers,^{320–322} Bürgi and co-workers,³²³ and Pei and co-workers³²⁴ have advanced our understanding of the reaction mechanism for synthesis of ligand-protected alloy NCs through metal exchange. However, the mechanism for synthesis of hydrophobic ligand-protected metal NCs in a two-phase system, which was first reported by Brust and co-workers,⁹⁵ has yet to be elucidated at the atomic and molecular levels. Similarly, ligand-exchange-induced structure transformation³²⁵ is not yet understood at the atomic or molecular level with respect to the chemical composition that induces the structural transformation. In the future, it is expected that high-resolution separation, mass spectrometry, ion-mobility experiments, and theoretical calculations will be used to clarify the details of reactions whose mechanisms are currently unknown (Fig. 20(a)), which will make it possible to further control synthesis of metal NCs.

7.2. Preparation of a textbook for students

For ligand-protected metal NCs, there was not a good textbook until 2014.¹⁰ However, taking into account the number of syntheses, reactions, and applications reported so far, it seems that metal-NC chemistry has reached a stage where many textbooks for undergraduate and graduate students should be published, like for organic chemistry³²⁶ and coordination chemistry.³²⁷ Publication of such a textbook would dramatically



developed in the future (Fig. 20(e)). I expect that such control can be realized if a size-convergence method can be found for loaded metal NCs, similar to the case of ligand-protected metal NCs dispersed in solution. In other words, it is expected that the number of constituent atoms and chemical composition of the loaded metal NCs can be controlled by combining the conventional loading method and size-convergence method. If such a method can be developed, it might also be possible to control the geometric structure of the loaded metal NCs. Although control of ligand-protected metal NCs dispersed in solution was once thought to be extremely difficult to achieve with atomic precision, many control methods have been developed.⁷ I hope that control of the number of constituent atoms and chemical composition of the loaded metal NCs on the support by an easier method will also be achieved through the enthusiasm of the readers of this paper.

7.6. Collaboration between academia and industry

The researchers in the field of metal NCs have discovered potential applications of metal NCs in various fields, as well as new facts hidden in nanoscale materials. For example, the applications described in Section 5 are such examples. Collaboration between industry and academia seems to be essential for successfully connecting the scientific discoveries found in these academic fields to development of industry. It is not necessary for all researchers in the academic field to collaborate with industry because the challenges that need to be overcome in the academic field and in industry are probably different. Thus, I would like to encourage young researchers to continue to make great effort to explore the new science hidden in the metal-NC field. However, the field of metal NCs is important in terms of both science and technology, and the techniques established in the field of metal NCs also have the potential to contribute to construction of the next-generation society (Section 5). For this reason, I urge researchers who already have the seeds of technology to also engage in industry–university collaborations so that the knowledge and techniques discovered in the field of metal NCs can be successfully linked to development of industry and society (Fig. 20(f)).

Conflicts of interest

There are no conflicts to declare.

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