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ARTICLE

Computational approaches to understanding reaction outcomes of organic processes in ionic liquids

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Sinead T. Keaveney,^a Jason B. Harper^{a,*} and Anna K. Croft^{b,*}Received 00th January 2012,
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This review considers how various computational methods have been applied to explain the changes in reaction outcome on moving from a molecular to an ionic liquid solvent. Initially, different conceptual approaches to modelling ionic liquids are discussed, followed by a consideration of the limitations and constraints of these approaches. A series of case studies demonstrating the utility of computational approaches to explain processes in ionic liquids are considered; some of these address the solubility of species in ionic liquids while others examine classes of reaction where the outcome in ionic liquids can be explained through the application of computational approaches. Overall, the utility of computational methods to explain, and potentially predict, the effect of ionic liquids on reaction outcome is demonstrated.

Introduction

Ionic liquids are defined as salts that are molten below 100°C;¹ the lower melting points of these liquids arise from the bulky, charge diffuse nature of their ions, most commonly the cation.² Ionic liquids can be composed of many different cations and anions; the ions that make up the ionic liquids discussed in this review are shown in Figure 1 (including the abbreviations that will be used throughout). These are representative of the types of ions commonly found in these solvents; pyridinium, tetraalkylammonium and tetraalkylphosphonium ions are also common.

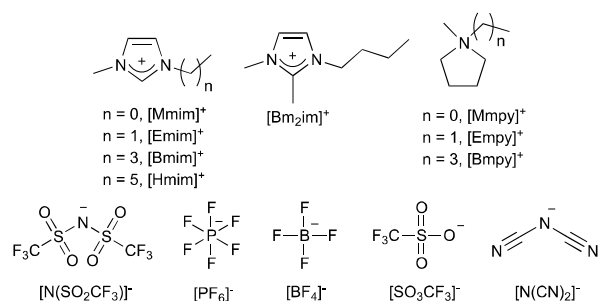


Figure 1: Cations and anions referred to in this review, and their abbreviations.

Ionic liquids have many properties that make them attractive alternatives to traditional solvents; negligible vapour pressure³⁻⁵ and low flammability,^{5, 6} as well as the prospect of recycling,⁷ can make them a safer and more environmentally-friendly option than volatile organic compounds. Their tuneable nature and ability to generate reaction outcomes different from those seen in traditional solvents is another factor driving

interest in the use of ionic liquids as solvents for organic reactions.^{2, 8-29} For such a replacement to be viable, their effect on reaction outcome needs to be understood and readily predicted, as the limited understanding of ionic liquid solvent effects is hindering further applications.^{30, 31}

The focus of this review is to introduce computational simulation techniques, and their underpinning methodologies, that have been able to provide valuable information in understanding the effects of ionic liquids on organic reaction outcomes, particularly when utilised in the context of experimental data. Computational methodologies have been used extensively in recent times to complement experimental studies in the ionic liquid area with a great deal of success. Although there have been a relatively limited number of studies that combine both experimental and computational techniques to understand the effect of ionic liquids on reaction outcome, there are a wealth of studies that have considered basic properties of the ionic liquids themselves, providing important physicochemical data. Due to the size of this emergent field, these underpinning studies will not be covered in this focussed review, except where they are directly related to organic reactions. Such studies include: investigating the existence of 'nano-domains' where the polar and non-polar moieties in the ionic liquid group together;³²⁻³⁶ the use of mixtures of different ionic liquids;³⁷⁻⁴⁰ the effect of using different chain lengths on the (imidazolium) cation of an ionic liquid;⁴¹⁻⁴⁵ as well as the use of di-cations.⁴⁶⁻⁴⁸

A number of reviews have been published in recent years addressing both the technical aspects of computational approaches used in ionic liquid modelling,⁴⁹⁻⁵² and the experimental work relating to chemical reaction mechanisms.⁵³

A related review, focussing on the wider issue of chemical reactivity in, and of, ionic liquids, assessed by computational methods, has recently appeared.⁵⁴ This review covers applications where the ionic liquid takes part directly in the reaction, often due to the acidic or basic nature of the solvent components. A representative example is the involvement of chloroaluminate ionic liquids, where the anion acts in a dual role as both part of the solvent and as a reagent.^{2, 55, 56} The applications discussed in the current review will consider reactions where the ionic liquid is acting as a solvent, rather than where the ionic liquid takes part in the reaction. There will be a particular focus on the utility of using a combined experimental and computational approach, when trying to delineate what effect the various interactions within ionic liquid mixtures have on reaction outcomes.

The major computational approaches and associated challenges

In addition to the standard computational considerations for molecular solvents, such as dispersion, π -stacking interactions, accurate reproduction of specific hydrogen-bonding interactions, and other polar interactions, when modelling ionic liquids, strong Coulombic forces also play a dominant role.⁵⁷⁻⁵⁹ Another challenge is that individual components will often have very different electronic distributions, which further adds to the complexity of the interactions present, as well as introducing the possibility of selective aggregation of like components.^{33, 60-64} Overall, it is very difficult to accurately describe the subtle balance of these interactions. This is further complicated by the nature of computational chemistry, for which assumptions need to be made in any molecular description. The extent of the assumptions made is dependent on the theoretical approach taken, as is the 'computational cost' of the method; that is, the amount of time and the computing resources required. As such, when modelling an ionic liquid the limitations of the available approaches must be matched with the information required from the simulation to arrive at the optimal technique.

When using computational techniques, there are three major conceptual approaches to treating a molecular system; these are depicted in Figure 2. The first is the use of electronic structure methods, which try to capture the most detailed view of a molecular system, short of solving the Schrödinger equation. *Ab initio* and semi-empirical calculations are highly accurate and reliable, as they are derived directly from quantum-mechanical principles; the energy is calculated from a wavefunction obtained using approximations of varying quality. As a result of this, there is generally a high computational cost associated with these methods and the simulations are restricted to smaller systems. Density functional theory (DFT), for which the underlying functionals can be either 'pure' or rely on parameterisation (*i.e.* are semi-empirical), is a good compromise between efficiency and accuracy for modelling ionic liquid systems. The energy is now expressed in terms of total electron density and this approach is less computationally demanding than traditional, good-quality *ab initio* methods of similar accuracy. This allows much larger systems to be handled with reasonable computation times. This

is particularly relevant when modelling ionic liquids as it allows multiple ion pairs to be modelled; ionic liquids have a complex mixture of short and long range interactions that can only be accurately captured through models incorporating multiple ion pairs.^{40, 65, 66} One of the main limitations of standard DFT methods is the lack of accuracy in describing dispersive interactions, the inclusion of which has been shown to be important for ionic liquid systems,^{49, 67-69} and techniques that circumvent this problem include dispersion-corrected DFT (DFT-D) functionals.

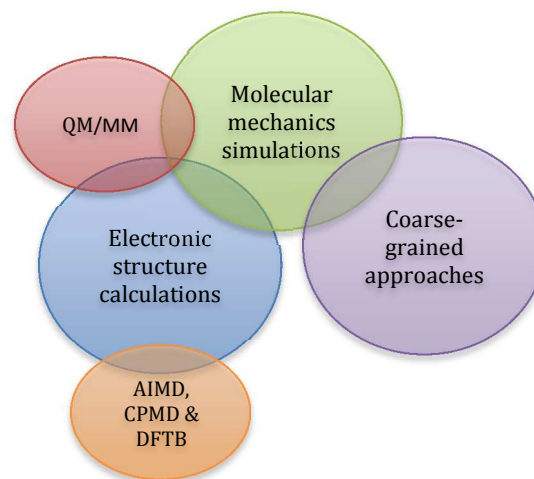


Figure 2: A representation of the three major conceptual approaches to computational chemistry used with ionic liquids, along with subsets of these methods.

In general, *ab initio* and DFT methods are used on static systems, but with recent advances in computing power, dynamic approaches have been emerging. These *ab initio* molecular dynamics (AIMD) approaches include the popular *Car Parinello* Molecular Dynamics (CPMD) methods. Currently such studies are limited by the size of the system that can be handled due to the quantity of computational power required, but provide a wealth of information not captured by static systems.⁷⁰⁻⁷⁶

Significant reductions in computational time relative to AIMD, whilst retaining DFT-level accuracy, can be achieved with Density Functional Tight Binding (DFTB) calculations,⁷⁷⁻⁷⁹ with this approach being recently pioneered for ionic liquid systems.⁸⁰⁻⁸³ Here, some parameterisation of the integrals is used to speed up the calculations, and dispersion corrections in the DFTB3 method help to describe long-range interactions with some accuracy. More importantly, this approach scales well for larger systems, so is able to handle a larger number of ion pairs for accurate calculations (see below). This is an emerging methodology that will no doubt be extremely useful when applied to reactive systems that include ionic liquid solvents.

The second conceptual approach is to use classical, atomistic molecular dynamics (MD) simulations, that rely on molecular mechanics (MM) treatments; this approach is valuable in visualising and interpreting the interactions that

exist between both the ions of an ionic liquid and any dissolved solutes, and between the ions themselves.^{50, 84} The MM treatment considers whole atoms, rather than the electrons, as discrete units, so is able to reduce computational time significantly at the cost of accuracy for systems where electronic interactions are less important.⁸⁴ The electronic properties of each atom are represented through a force field that characterises charges, interactions, bonds, angle bending and dihedral rotation using a series of parameters. The main factor that determines the accuracy of a molecular dynamics simulation is the quality of the force field that is used. There are currently a relatively limited number of force fields available that adequately describe all the physical and chemical properties of a wide range of ionic liquids,⁸⁵⁻⁸⁹ with many being specific for the ionic liquid (or class of ionic liquids) they are parametrised for (see Table 8 below).

To gain the best of both worlds and utilise both of the above conceptual frameworks, the hybrid approach of Quantum Mechanics/Molecular Mechanics (QM/MM) can be used to describe the core reactive centres using electronic structure (*ab initio*, DFT) approaches, with the impact of the ionic liquid environment described using molecular-mechanics based force fields.^{90, 91} An alternative approach is to use ReaxFF (reactive force field), which has an accuracy similar to semi-empirical calculations. Here, a specifically parameterised force field is used with a bond order approach, allowing bond-breaking and bond-making to be simulated.⁹² This methodology has been utilised in the ionic liquid field for examining carbon dioxide absorption,⁹³ and may have future applicability for reactions in ionic liquids.

Coarse-grained approaches form the third conceptual model. These are variations of classical molecular dynamics simulations, where atoms are grouped together into 'beads' or 'grains'.⁹⁴ These beads are typically parameterised from atomistic MD simulations, with the concomitant loss of chemical resolution, but greatly improved calculation times. This methodology has not yet been utilised for organic reactions in ionic liquids (and thus will not be described further in this review), but has been successfully applied to ionic liquids themselves^{95, 96} and in areas such as biomass dissolution, where very large molecular systems need to be considered.^{97, 98} Based on work in the biomolecular simulation field,⁹⁹ coarse-grained methods show significant promise for future applications in multiscale systems, where they can be combined with atomistic models to look at reactions on large systems and for the long time-scales that are required for proper treatment of slowly-diffusing ionic liquids.

Methodological limitations and considerations

Currently there is a strong focus on developing accurate and reliable computational techniques that can precisely model the static and transport properties of ionic liquids. Work in this area often involves complex parameterisation for each ion combination used, and is very computationally demanding. Other areas are focused on the development of force fields for

use in classical (non-polarisable) molecular dynamics, with an emphasis on transferability and efficiency, at the expense of accuracy.⁸⁵ Whilst these methods may not be able to reproduce the exact transport properties of specific ionic liquids, they can still be very useful for gaining a basic understanding of what is happening at the molecular level, and can provide the rationalisation needed to complement experimental work. It is arguable whether computationally intensive and exact simulations are always necessary, and the different computational approaches discussed can be useful for certain applications, as long as the limitations of the technique are taken into account.

Ab initio and DFT studies

The primary limitation of *ab initio* and DFT studies is that they are computationally 'expensive' for systems of more than a few atoms; that is, they require substantial amounts of computer time because they generate a detailed electronic description of the system under study. As such, the primary motivation when using these methods is on delineating which specific interactions play a governing role in determining ionic liquid properties and in influencing reactivity. This information can also be fed forward into improving force field parameters. An example interaction that requires electronic detail is hydrogen bonding, which is known to be heavily involved in directing a number of reactions and solvent properties.^{23, 100-103} Such interactions require a full quantum mechanical treatment, as the nuances behind them are not always fully captured in molecular dynamics force fields.^{68, 102, 104}

NUMBER OF ION PAIRS

Many *ab initio* and DFT calculations are performed on single ions or an ion pair due to the computational expense of these calculations, despite suggestions that multiple ion pairs are required to generate accurate electronic descriptions.^{65, 105} The importance of system size has been highlighted for quantum computational methods, for example modelling of single and multiple ion pairs of the ionic liquid [Mmim][Cl] indicated that the calculated dipole moment, binding energy and partial charges were substantially dependent on the number of ion pairs used in the calculation.¹⁰⁶ Considering the variety of short range interactions and long range Coulombic forces that exist in ionic liquids, it is unsurprising that in order to capture the true physical and chemical properties of the bulk ionic liquid phase, multiple ion pairs need to be used.

An interesting method that has emerged in recent years is DFTB.^{78, 79} While there will be a decrease in the accuracy of the calculations relative to *ab initio* methods, due to the greatly improved calculation times it is possible to perform calculations on multiple ion pairs. In a recent study investigating protic ionic liquids using DFTB systems containing up to 20 ion pairs were able to be modelled.⁸⁰ Another study was able to model clusters of imidazolium nitrate ionic liquids containing up to 15 ion pairs using DFTB, and found that the structure and stability of the ionic liquids could be accurately modelled using this technique.⁸³

One further approach to reducing the complexity of including a number of ion pairs is to utilise a continuum method (sometimes referred to as implicit solvent method) to mimic the bulk phase. This is driven by a static dielectric parameter, and leaves a minimum number of molecules to be fully calculated. Such continuum models provide a computationally cheap way to rapidly screen systems by providing the bulk of the forces to be experienced by the molecules of interest, but do not allow for specific interactions to be modelled beyond those explicitly described. For some systems, however, this approach will be adequate to qualitatively assess the impact of the ionic liquid on reaction barriers and energies.^{52, 54} Particularly notable is the proprietary COSMO-RS model, which has been simplified and parameterised to be numerically more efficient than standard polarisable continuum models, has proven popular in the ionic liquid field. However, it has not been extensively used for reaction mechanism studies, with only a few examples highlighted in the literature.^{107, 108} Use of this model has, however, highlighted the difference in charge transfer experienced by ions, relative to non-application of a continuum,¹⁰⁹ and this will have an impact on the properties and solvation of substrates, products and transition states.

In terms of polarisable continuum models, the General Ionic Liquid solvation model (SMD-GIL) has recently demonstrated that small mean unsigned errors can be achieved for solvation free energies in a range of ionic liquids.¹¹⁰ The utility of this approach for exploring reaction pathways in more detail has been shown for the DABCO-enhanced reaction of the [Bmim]⁺ cation with benzaldehyde,¹¹¹ and the acid-catalysed hydrolysis of lignin using standard DFT approaches.¹¹²

METHOD CHOICE

The choice of which computational method to use has been discussed in a recent review,⁴⁹ where an assessment of the advantages and limitations of a range of different DFT methods, and the correlated second-order Moller-Plesset (MP2) methodology, concluded that the use of *ab initio* methods, specifically, MP2, rather than DFT methods, is desirable when performing energy calculations, whilst less computationally demanding theoretical approaches are most practical and cost-effective for geometry calculations.^{68, 106, 113-120} This is supported by a number of studies that have assessed the effectiveness of a variety of different DFT and *ab initio* methods; one such study calculated the ion-pair binding energies (IPBEs) of a range of pyrrolidinium based ionic liquids.¹²¹ As an example, Table 1 contains the mean average deviation of the calculated IPBEs from the benchmark method (MP2/6-311+G(3df,2p)) for the pyrrolidinium chloride and pyrrolidinium bistriflimide ionic liquids. As can be seen, there was a large variation in the calculated IPBEs when using different methods, with the KMLYP and M05-2X methods performing best. The authors generally concluded that the use of less costly methods, including the commonly used B3LYP method, was not sufficient to predict accurate ion pair binding energies for even single ion pairs.¹²¹

Table 1 The mean average deviation (in kJ mol⁻¹) of the calculated ion pair binding energies for pyrrolidinium chloride and bistriflimide ionic liquids ([Mmpy]⁺ and [Empy]⁺) for a select number of different levels of theory, compared with the benchmark MP2/6-311+G(3df,2p) method.¹²¹

Anion	Method						
	HF	PBE	B3LYP	KMLYP	M05-2X	SCS-MP2	SOS-MP2
[Cl] ^a	35.6	1.0	13.6	7.7	0.6	9.1	13.6
[N(SO ₂ CF ₃) ₂] ^b	43.2	21.5	27.2	3.0	5.1	10.7	16.0

^a Values for configuration 1, as reported in the paper.¹²¹

^b Values for configuration 1 – N, as reporting in the paper.¹²¹

Calculations of the standard molar enthalpy of formation for [Bmim][N(CN)₂] also highlighted the limitations of DFT methods; the G3MP2 calculated value of 359.63 kJ mol⁻¹ was much closer to the experimental value of (363.4 ± 2.7) kJ mol⁻¹, when compared to the more poorly performing B3LYP/6-31+G(d,p) method which gave a calculated value of 345.5 kJ mol⁻¹.¹¹³ Further, another investigation found that B3LYP calculations were sufficient to calculate bond length and angles for [Bmim][Cl] that were in good agreement with X-ray data, yet calculations of the interaction energies were very dependent on the method and basis set of choice (Table 2).¹¹⁷

Table 2 The interaction energies (in kJ mol⁻¹) for [Bmim][Cl], calculated using a number of different methods and basis sets.¹¹⁷

Interaction Energy ^a	Method / Basis Set				
	B3LYP / 6-31G*	B3LYP / 6-31++G**	B3LYP / 6-311++G**	HF / 6-31+G*	MP2 / 6-31+G*
	-405.86	-452.44	-457.08	-376.96	-371.21

^a Values for the monoclinic S1 calculations, as reported in the paper.¹¹⁷

While it can be generally concluded that the use of *ab initio* and composite approaches is desirable, as mentioned earlier, the use of multiple ion pairs is also necessary to model the nano-scale structure of the bulk phase of ionic liquids. This leads to a trade off between accuracy and system size, as calculations of multiple ions pairs using *ab initio* or DFT methods are often too computationally demanding to be practical. As discussed above, DFTB is a promising alternative that allows multiple ion pairs to be modelled relatively accurately with a lowered computational cost relative to traditional *ab initio* and DFT methods.

Another interesting approach is the use of DFT-D, which works by including an additional dispersive term in a standard DFT functional, so can better describe the weak interactions that are core to hydrogen bonding and π -interactions. Although this adds an overhead to calculation time, it has been shown to produce much improved results,^{122, 123} and such functionals have been applied with some success, particularly in the ionic liquid field.¹²²⁻¹²⁵

Molecular dynamics simulations

MD simulations provide a balance of accuracy and system size, where interactions that do not involve electron transfer from one component to another are to be considered. Underpinning the accuracy of such simulations are (i) the parameters used to

generate the system, which themselves are often now based upon *ab initio* and DFT calculations, and (ii) the simulation conditions used.

GENERATING THE STRUCTURES FOR THE SIMULATIONS

Before any dynamic simulations can be completed, it is necessary to obtain the relevant structure(s) either from X-ray crystal structures or, more commonly now, using either *ab initio* or DFT methods. Generally DFT methods are sufficient for the geometry optimisation of ionic liquids,¹²¹ but for calculations of atomic point charges (discussed below) a higher accuracy method is required,^{68, 85, 114} this is an important consideration with respect to the accurate re-parameterisation of force fields. A commonly used combination is to perform geometry calculations at the B3LYP level, followed by Restrained Electrostatic Potential (RESP) single point charge calculations, using the MP2 level of theory, to determine the atomic partial charges.^{42, 85, 126} Obviously, there are a number of other methods available, and the type of theory and size of the basis set used is dependent on the required accuracy of the calculation. As an example, less accurate methods may be sufficient to generate a structure to be used within a defined force field, whilst higher levels of theory would be desirable when the force field is to be parameterised with the calculated values.

ATOMIC PARTIAL CHARGE CALCULATIONS

Given the importance of electrostatic interactions in ionic liquids, it is essential that the atomic partial charges on the ions are accurately represented. The main factors that need to be considered are: (i) the method used to derive the partial charges, (ii) the type and accuracy of the theoretical approach, and (iii) whether calculations of a single atom in the gas phase truly represent the partial charge of the bulk liquid phase.¹²⁷ For the calculation of partial charges high accuracy *ab initio* methods are desirable, yet moderately sized basis sets can also perform adequately and are often compared with larger basis set calculations to validate the method.^{116, 128}

There are two major established methodologies to predict partial atomic charges; the one-particle density matrix method and the one-particle electron density method. The one-particle density matrix approach calculates the atomic partial charge by comparing the number of electrons residing in the ground state of a neutral, free atom and the atomic population on the atom within the molecule. The most commonly used methods that are based on this scheme are the Mulliken Population Analysis (MPA),¹²⁹ Löwdin Population Analysis (LPA)¹³⁰ and Natural Population Analysis (NPA)¹³¹ methods. While initially these methods were very popular, they are relatively inaccurate and considered somewhat dated; most studies now use the electron density approach.¹³² In these schemes the atomic partial charge is calculated by deriving the electrostatic potential of the system using *ab initio* methods, followed by fitting of the atomic charges to reproduce this electrostatic potential. These RESP techniques are now the method of choice in the calculation of ionic liquid partial charges.¹²⁷

A comprehensive study compared a number of commonly used atomic partial charge schemes and assessed their effectiveness.¹²⁷ The general conclusion was that all the methods had drawbacks, but the RESP methods performed considerably better than all of the population analysis schemes. There was also a surprisingly large variation in the calculated partial charges between the different methods; for example, the partial charge on the chlorine atom in [Mmim][Cl] varied from the commonly used value of -1, to as low as -0.63¹⁰⁵ when calculated by CPMD using the Blöchl method; chloride partial charges calculated using a variety of different methods are shown in Table 3.^{66, 105, 133-136}

Table 3 The partial charges on the chloride ion of [Mmim][Cl] calculated using different methods.

Charge ^a	Method				
	MP2, ChelpG	MP2, RESP	MP2, BA ^b	BP86, MPA	BP86, NPA
	-0.77 ⁶⁶	-0.78 ⁶⁶	-0.86 ⁶⁶	-0.67 ¹³³	-0.78 ¹³³

^a Average of the different conformations
^b Where BA = Bader Analysis

Another study calculated the partial charges on [Bmim][Br] using RESP, Natural Bonding Orbitals (NBO) and Shared Electron Number (SEN) methods, and found that the calculated charges within the imidazolium ring were very dependent on the method used.¹³⁷ For example, the charge on the nitrogen atom adjacent to the butyl chain was calculated to be three markedly different values; +0.07, -0.35 and +0.21; using three different methods. It is clear that when wanting to compare a number of ionic liquids, using the same method for the partial charge calculations will at least allow for a degree of internal consistency within the set.

Most partial charge calculations are currently done using a single ion in the gas phase, mainly due to the computational expense of modelling bulk systems; with arguments for and against this approach having been published.^{76, 119, 136} Recently, the use of ionic liquid-based continuum methods has indicated that significant improvements can be obtained in the calculated charges over single ion pairs in the gas phase.¹⁰⁹ Many studies suggest that, due to the presence of short and long range interactions, multiple ion pairs are required to accurately model both the electronic environment and the distribution of ions in ionic liquids.^{40, 65, 66, 138} This necessity to use multiple ion pairs is further reinforced by calculations which have demonstrated that different conformers of an ion pair can give markedly different partial charges (Table 4); in other words, the position of the ions relative to each other affects the resultant partial charge.^{135, 136} Clearly as the position of the ion affects the partial charge, having a number of ion pairs is best as it will allow for a more accurate average partial charge to be calculated, and the most accurate representation of the bulk liquid environment.

Another advantage of using multiple ion pairs is the charge transfer information that can be obtained from these calculations; the importance of considering extents of charge transfer between ions will be discussed further below. The

detailed study mentioned above highlighted the need to use clusters of ions to represent charge transfer within the ionic liquid, but also identified issues with this approach due to insufficient charge fitting as some atoms were being ‘buried’ within the cluster of ions.¹²⁷ Despite this, the advantages that ion pair(s) confer in partial charge calculations, with respect to capture of charge transfer information, has seen calculations utilising ion pair(s) becoming more prevalent than those of a single ion.^{66, 76, 105, 139-143}

Table 4 The partial charges on the chloride ion of [Mmim][Cl] calculated using different conformers of a single ion pair.

Method	Chloride ion charge, in different conformers			
CCSD/RESP ^{136 a}	-0.75	-0.73	-0.86	-0.80
PBE/Blöchl ^{136 a}	-0.68	-0.63	-0.80	-0.73
B3LYP/RESP ^{135 b}	-0.723	-0.751	-0.824	-0.662

^a Conformers 12, 13, 14 and 15 from the IC1A cluster from the paper¹³⁶
^b Conformers 5m, 4m, 44 and top from the paper.¹³⁵

CHARGE SCALING

An alternative proposed in order to avoid the use of cluster calculations is to scale the total ionic charge to below unity. Scaled charges allow for an extent of charge transfer between the ions to be taken into account, as well as mimicking the charge screening that results from the polarisability of the electron clouds in the system. While this method is not ideal, it has been shown to improve the calculated dynamic properties substantially without any additional computational cost, and is therefore becoming a very popular approach.^{66, 76, 105, 139-151} For example, ionic liquid densities can be calculated quite accurately using either unscaled (*i.e.* net ion charge = 1) or scaled charges (Table 5), yet generally other dynamic properties such as the self-diffusion coefficient (Table 6) and heat of vaporisation (Table 7) are better predicted when scaled charges are used.

Table 5 The calculated and experimentally determined densities (g cm⁻³) of a number of ionic liquids.

Ionic liquid	Scaled	Un-scaled	Experimental
[Emim][BF ₄] ^a	1.162 ¹⁴²	1.183 ¹⁵² 1.218 ¹⁵²	1.193 ¹⁵²
[Bmim][BF ₄] ^b	1.202 ¹⁴⁰ 1.190 ¹⁴²	1.178 ¹⁵³ 1.208 ¹⁵⁴	1.190 ¹⁵⁵ 1.204 ¹⁵⁶
[Bmim][N(SO ₂ CF ₃) ₂] ^b	1.446 ¹⁴⁰	1.48 ⁸⁷	1.437 ¹⁵⁷

^a Determined at 400 K. ^b Determined at 300 K.

Table 6 The calculated and experimentally determined self-diffusion coefficients (10⁻⁷ cm² s⁻¹) of the anion, in a number of ionic liquids.

Ionic liquid	Scaled	Un-scaled	Experimental
[Bmim][BF ₄] ^a	1.83 ¹⁴⁰	1.0 ¹⁵⁸ , 0.3, ¹⁵³ 0.08 ¹⁴²	1.70 ¹⁵⁷
[Emim][BF ₄]	3.7 ¹⁴² , 5.5 ¹⁴⁹	0.9, ¹⁵⁹ 2.0, ¹⁵⁴ 0.11 ¹⁴²	4.2 ¹⁶⁰
[Bmim][PF ₆] ^a	0.74 ¹⁴⁰	1.0 ¹⁵⁹	0.71 ¹⁵⁷
[Bmim][N(SO ₂ CF ₃) ₂] ^a	2.10 ¹⁴⁰		2.60 ¹⁵⁷

^a Determined at 300 K. ^b Determined at 298 K.

Table 7 The calculated and experimentally determined heat of vaporisation (kJ mol⁻¹) for a number of different ionic liquids.

Ionic liquid	Scaled	Un-scaled	Experimental
[Hmim][N(SO ₂ CF ₃) ₂] ^a	158 ¹⁴⁷	186 ¹⁴⁷	149 ¹⁶¹
[Emim][BF ₄] ^a	124 ¹⁴²	159, ¹⁴² 75, ¹⁶² 100 ¹⁶²	149 ¹⁶¹
[Bmim][BF ₄] ^a	135, ¹⁴² 133 ¹⁴⁰	161, ¹⁴² 116, ¹⁶² 126 ¹⁶²	141 ¹⁶³

^a Determined at 298 K.

There are two ways in which this method has been used: (i) by applying a relatively arbitrary scaling factor to the partial charges (either calculated or directly from the force field being used) to match a set of target experimental data,¹⁴⁴⁻¹⁵¹ and (ii) by determining the reduced charges for a particular anion-cation combination from quantum calculations of the ion pair(s).^{66, 76, 105, 139-143} In the majority of these studies, the transport properties of the ionic liquid are significantly closer to experimental values when reduced charges are used. The validity of this method has been supported by a number of experimental studies, where charge transfer between ions was reported in studies using X-ray Photoelectron Spectroscopy (XPS)^{40, 164-166} and Nuclear Magnetic Resonance (NMR) spectroscopy.¹⁶⁶ It has generally been found that the difference in the extent of charge transfer is more marked when changing the anion, than when changing the cation.^{76, 140} This observation highlights that when modelling a range of different ionic liquids it might be more important to consider the different charge transfer effects when varying the anion, than when the cation is varied. Overall, charge scaling is an effective way to introduce a degree of ‘polarisability’, although it is not as robust as the use of true polarisable models.¹⁶⁷

NUMBER OF SIMULATION ION PAIRS

As mentioned earlier, in order to accurately capture both short and long-range interactions that exist in the liquid phase a number of ion pairs would need to be used to simulate the true nature of the bulk ionic liquid. It is quite widely accepted that the use of multiple ion pairs is necessary when performing an MD simulation, as effects such as solvent ordering, the formation of nano-domains and long-range Coulombic interactions can only accurately modelled when a number of ion pairs are used. One study suggests that the use of 50 ion pairs is adequate for static properties, whilst 500 are required to accurately model dynamic properties.¹⁶⁸ A separate study on a phosphonium based ionic liquid found that the results from using either 64 or 512 ion pairs were quite different.¹⁴⁴ However, there have also been studies suggesting that a lower number of ion pairs is sufficient in some cases: a CPMD study on the ionic liquid [Mmim][Cl] indicated that near-order effects could be captured by using a smaller number of ion pairs (25, as compared to either 41 or 300), by comparison of the radial distribution functions;¹³³ note that for this case it was the near-order effects that were well represented. Overall, considering the complex number of interactions present in ionic liquids and the computational power now available, it is recommended that MD simulations consist of at least 200 ion pairs so that the long-range interactions and any ‘bulk liquid’ effects can be

captured, with most simulations now using an excess of 200 ion pairs.^{35, 42, 46, 143, 144}

FORCE FIELDS

The most commonly used force field frameworks are CHARMM,¹⁶⁹ AMBER^{170, 171} and OPLS,^{171, 172} with the major difference between these force field frameworks being way that the force fields are parameterised. For this reason, different force fields should not be mixed within a simulation. Force fields can either treat each atom explicitly, known as all atom (AA) force fields, or can group multiple atoms together (*e.g.* carbon and the attached hydrogen atoms in a methyl group) and treat them implicitly; these are known as united atom (UA) force fields and are often used to reduce computation time.

The OPLS-AA based CL&P^{85, 87-89, 173} force field is one of the most widely used and transferable force fields for ionic liquids currently available. As many ionic liquids are based on general organic moieties, many of the repulsive, angle and bond parameters could be transferred directly from the OPLS-AA to the CL&P force field. However, given the charged nature of ionic liquids, the focus was on re-parametrisation of the partial charges and dihedral angles so that ionic liquids were more accurately described.⁸⁵ This approach has been very successful, providing an adequate force field to be used for modelling a range of ionic liquids.^{85, 87-89, 173} The advantages and limitations of this framework lie in its transferability: while it can be used for multiple systems, the 'generic' nature of its parameters meant that a number of approximations were made, reducing its accuracy in describing specific ionic liquids.

There have been other ionic liquid force fields developed, *e.g.* an all atom force field based on the AMBER force field for use with imidazolium based ionic liquids,¹⁵⁹ as well as many other force fields for use with specific ionic liquids. Table 8 lists a number of the available ionic liquid force fields (although it is not comprehensive), and includes the force field framework on which the developed force field was based, the cations and anions contained in the force field, whether it is an all atom or united atom model, whether it is polarisable or non-polarisable, and for the non-polarisable cases whether the sum of the partial charges have been scaled below unity or not.

NON-POLARISABLE VS POLARISABLE FORCE FIELDS

Standard MD force fields are non-polarisable, meaning that they do not take into account the fluctuations in charge distribution (polarisation) experienced by one molecule in response to the approach of another molecule. When the approaching species is an ion, this response would be expected to be large, and have a strong influence on the dynamic interactions within an ionic liquid mixture. Despite the lack of inclusion of polarisability, the standard force fields have been successful at predicting a number of static quantities, including melting points, and radial and spatial distribution functions. Dynamic properties of ionic liquids are not always accurately predicted when using non-polarisable force fields,^{152, 174-176} and the fine details of the interactions are likewise not always accurately captured,¹⁷⁷ highlighting some of the limitations of

classical MD. As discussed above, substantial improvement in the calculated properties has been achieved when the force field is refined (Tables 5-7),^{140, 150, 178-180} yet many studies conclude that the use of a polarisable model is required for the accurate representation of ionic liquids,^{175, 176, 181, 182} despite the perceived high computational cost associated with these models.

There are two main types of polarisable models: point induced dipoles, many of which are based on the Thole model,¹⁸³ where a smeared charge distribution is used both on the atom and on 'off atom' sites; and Drude oscillators, which mimic dipoles by attaching a fixed point charge on an harmonic spring to each atom location.¹⁸⁴ Both methods have had some success. Studies using point induced dipoles have found that the calculated transport properties were much closer to experimental values than those calculated using classical MD,^{175, 181, 182} and found that long range ordering¹⁸⁵ and the formation of ion cages¹⁸⁶ in ionic liquids is important. A recent study calculated a number of anisotropic ionic polarisabilities, based on the Thole model, for a wide range of anions and cations, which can be utilised in MD simulations.¹⁸⁷ Recent efforts have also been made to develop a polarisable force field, based on the Thole model, that can be used for a wide range of ionic liquids, known as the Atomistic Polarizable Potential for Liquids, Electrolytes and Polymers (APPLE&P) force field.¹⁸⁸⁻¹⁹¹

There have also been a number of studies that have utilised the Drude oscillator model;^{167, 192-194} one demonstrated that the dynamics, dipole moments and dielectric constants are dependent on the extent of polarisation used when modelling the ionic liquid [Emim][SO₃CF₃].¹⁹² Another study found that relatively long lived, heterogeneous ion cages formed within the ionic liquid.¹⁹³

Another method, which is becoming more popular, is the Electronegativity Equalisation Method (EEM) (which can then be combined with molecular mechanics), where the polarisability of a system is mimicked by allowing the atomic partial charges to fluctuate.^{195, 196} Such an approach accounts for changes in the electrostatic potential of atoms in response to different environments, and can also represent hydrogen bonding interactions.¹⁹⁵

The results obtained from polarisable models describe properties that are important for explaining reactivity, and in the coming years it is likely that the use of polarisable force fields will become more widespread in investigating the impact of ionic liquids on reaction outcome.

Case studies: the utility of combining computational and experimental approaches

General solubility studies

Solubility studies are important as they highlight some of the underpinning concepts regarding the use of ionic liquids as solvents, along with helping to shape our understanding of ionic liquid mixtures. Having a sound understanding of the range of

Table 8 Summary of a number of the ionic liquid force fields that are currently available.

Force Field	Basis	Cations included	Anions Included	Type ^a	Charges ^b
CL&P ^{85, 87-89, 173, 197}	OPLS-AA	Many imidazolium, pyridinium, ammonium and phosphonium based cations	Halides, [NO ₃] ⁻ , [N(CN) ₂] ⁻ , [SO ₃ CF ₃] ⁻ , phosphate, and alkylsulfate, alkylsulfonate, fluoroalkylfluorophosphate and bis(sulfonyl)imide based anions	AA, Non-P	Unity
Acevedo <i>et al.</i> ¹⁶²	OPLS-AA	1-alkyl-3-methylimidazolium and <i>N</i> -alkylpyridinium cations, choline	[Cl] ⁻ , [NO ₃] ⁻ , [PF ₆] ⁻ , [BF ₄] ⁻ , [AlCl ₄] ⁻ , [Al ₂ Cl ₇] ⁻ , [SO ₃ CF ₃] ⁻ , saccarinate and acesulfamate	AA, Non-P	Unity
LHW AA ^{159, 198}	AMBER	1-alkyl-3-methylimidazolium cations	[Cl] ⁻ , [BF ₄] ⁻ and [PF ₆] ⁻	AA, Non-P	Unity
LHW UA ¹⁵⁴	Liu <i>et al.</i> AA ¹⁵⁹	1-alkyl-3-methylimidazolium cations	[BF ₄] ⁻ and [PF ₆] ⁻	UA, Non-P	Unity
Maginn <i>et al.</i> ¹⁹⁹	CHARMM	Triazolium cations	[NO ₃] ⁻ and [ClO ₄] ⁻	AA, Non-P	Unity
Maginn <i>et al.</i> ²⁰⁰	CHARMM	Pyridinium cations	[N(SO ₂ CF ₃) ₂] ⁻	AA, Non-P	Unity
Maginn <i>et al.</i> ²⁰¹	CHARMM	[Bmim] ⁺	[PF ₆] ⁻	AA, Non-P	Unity
Maginn <i>et al.</i> ²⁰²	CHARMM	[Emim] ⁺	[CH ₃ CH ₂ SO ₄] ⁻	AA, Non-P	Unity
Maginn <i>et al.</i> ²⁰³	CHARMM	Hydrazinium based cations	[NO ₃] ⁻	AA, Non-P	Unity
Ludwig <i>et al.</i> ¹⁸⁰	CL&P	1-alkyl-3-methylimidazolium cations	[N(SO ₂ CF ₃) ₂] ⁻	AA, Non-P	Unity
Ludwig <i>et al.</i> ²⁰⁴	Ludwig <i>et al.</i> ¹⁸⁰	1-alkyl-3-methylimidazolium cations	[N(SO ₂ CF ₃) ₂] ⁻	UA, Non-P	Unity
Soares <i>et al.</i> ²⁰⁵	GROMOS	[Bmim] ⁺	[NO ₃] ⁻ and [PF ₆] ⁻	UA, Non-P	Unity
Stassen <i>et al.</i> ²⁰⁶	AMBER	[Emim] ⁺	[AlCl ₄] ⁻	AA, Non-P	Unity
Stassen <i>et al.</i> ¹⁵⁸	AMBER	[Bmim] ⁺	[AlCl ₄] ⁻ and [BF ₄] ⁻	AA, Non-P	Unity
Zhang <i>et al.</i> ²⁰⁷	AMBER	1,1,3,3-tetramethylguanidium	Formate, lactate, [ClO ₄] ⁻ , [SO ₃ CF ₃] ⁻ , [CF ₃ COO] ⁻	AA, Non-P	Unity
Zhang <i>et al.</i> ¹⁷⁹	AMBER	Cyclic guanidium based cations	[NO ₃] ⁻	AA, Non-P	Unity
Zhang <i>et al.</i> ²⁰⁸	AMBER	Guanidium cations	[Cl] ⁻ , [BF ₄] ⁻ and [PF ₆] ⁻	AA, Non-P	Unity
Zhang <i>et al.</i> ²⁰⁹	AMBER	Guanidium cations	[NO ₃] ⁻ and [ClO ₄] ⁻	UA, Non-P	Unity
Zhang <i>et al.</i> ²¹⁰	AMBER	Tetrabutylphosphonium	Amino acids	AA, Non-P	Unity
Zhang <i>et al.</i> ²¹¹	AMBER	Phosphonium cations	[N(SO ₂ CF ₃) ₂] ⁻	AA, Non-P	Unity
Zhang <i>et al.</i> ²¹²	AMBER	1-allyl-3-methylimidazolium	[Cl] ⁻	AA, Non-P	Unity
Youngs <i>et al.</i> ²¹³	OPLS-AA	[Mmim] ⁺	[Cl] ⁻	AA, Non-P	Unity
Li <i>et al.</i> ²¹⁴	OPLS-AA	1,1,3,3-tetramethylguanidium	Lactate	AA, Non-P	Unity
Voth <i>et al.</i> ²¹⁵	AMBER	[Emim] ⁺	[NO ₃] ⁻	AA, Non-P	Unity
Voth <i>et al.</i> ⁹⁴	Yan <i>et al.</i> ¹⁷⁵	1-alkyl-3-methylimidazolium cations	[NO ₃] ⁻	UA, Non-P	Unity
Hu <i>et al.</i> ²¹⁶	AMBER	[Emim] ⁺	Lactate	UA, Non-P	Unity
Singh <i>et al.</i> ²¹⁷	AMBER	[Emim] ⁺	[CH ₃ COO] ⁻	AA, Non-P	Unity
Smit <i>et al.</i> ²¹⁸	AMBER and OPLS	Triazolium cations	[N(SO ₂ CF ₃) ₂] ⁻	AA, Non-P	Unity
Laaksonen <i>et al.</i> ²¹⁹	AMBER and DREIDING	Tetraalkylphosphonium cations	Chelated orthoborate anions	AA, Non-P	Unity
Chaban <i>et al.</i> ¹⁴²	LHW AA ¹⁵⁹	[Emim] ⁺ and [Bmim] ⁺	[BF ₄] ⁻	AA, Non-P	Scaled
Chaban <i>et al.</i> ²²⁰	CL&P	<i>N</i> -butylpyridinium	[BF ₄] ⁻ , [N(SO ₂ CF ₃) ₂] ⁻ , [PF ₆] ⁻ , [N(CN) ₂] ⁻ , [SO ₃ CF ₃] ⁻ and [Cl] ⁻	AA, Non-P	Scaled
Balabramanian <i>et al.</i> ¹⁵⁰	CL&P	[Bmim] ⁺	[PF ₆] ⁻	AA, Non-P	Scaled
Balabramanian <i>et al.</i> ¹⁴⁰	CL&P	1-alkyl-3-methylimidazolium cations	[Cl] ⁻ , [BF ₄] ⁻ , [NO ₃] ⁻ , [PF ₆] ⁻ , [SO ₃ CF ₃] ⁻ and [N(SO ₂ CF ₃) ₂] ⁻	AA, Non-P	Scaled
Han <i>et al.</i> ²²¹	CHARMM	[Bmim] ⁺	[PF ₆] ⁻ , [SO ₃ CF ₃] ⁻ , [CF ₃ COO] ⁻ , [C ₃ F ₇ COO] ⁻ and [C ₄ F ₉ SO ₃] ⁻	AA, Non-P	Scaled
Lisal <i>et al.</i> ^{143, 222}	CHARMM	Chiral imidazolium based cations	[Br] ⁻	AA, Non-P	Scaled
Youngs <i>et al.</i> ¹⁵¹	Youngs <i>et al.</i> ²¹³	[Mmim] ⁺	[Cl] ⁻	AA, Non-P	Scaled
Economou <i>et al.</i> ⁴¹	CHARMM	1-alkyl-3-methylimidazolium cations	[N(SO ₂ CF ₃) ₂] ⁻	AA, Non-P	Scaled
Economou <i>et al.</i> ²²³	AMBER and DREIDING	[Emim] ⁺	[B(CN) ₄] ⁻	UA, Non-P	Scaled
Maginn <i>et al.</i> ²⁰¹	CHARMM	[Bmim] ⁺	[PF ₆] ⁻	AA, Non-P	Scaled
Smit <i>et al.</i> ²²⁴	AMBER	1-alkyl-3-methylimidazolium cations	[Cl] ⁻	UA, Non-P	Scaled
Liu <i>et al.</i> ²²⁵	OPLS and AMBER	1-alkyl-3-methylimidazolium cations	[BF ₄] ⁻ , [PF ₆] ⁻ , [SO ₃ CF ₃] ⁻ , [SO ₃ CH ₃] ⁻ , [CF ₃ COO] ⁻ , [CH ₃ COO] ⁻ and [N(SO ₂ CF ₃) ₂] ⁻	UA, Non-P	Scaled
Holm <i>et al.</i> ¹³⁹	CL&P	1-alkyl-3-methylimidazolium cations	[Cl] ⁻ , [N(CN) ₂] ⁻ and [SCN] ⁻	AA, Non-P	Scaled
Chaban <i>et al.</i> ²²⁶	CHARMM	[Emim] ⁺	Amino acid based anions	AA, Non-P	Scaled
APPLE&P ¹⁸⁹		Many imidazolium, pyridinium, ammonium, morpholinium and phosphonium cations	[PF ₆] ⁻ , [NO ₃] ⁻ , and many fluoroborate, cyano and bis(sulfonyl)imide based anions	AA, P	-
Wu <i>et al.</i> ²²⁷	Wu <i>et al.</i> ¹⁹⁵	1-alkyl-3-methylimidazolium cations	Glycine	AA, P	-
Wu <i>et al.</i> ²²⁸	OPLS-AA	Guanidium cations	[NO ₃] ⁻ and [ClO ₄] ⁻	AA, P	-
Wu <i>et al.</i> ¹⁹⁵	AMBER	[Emim] ⁺	Glycine	AA, P	-
AMOEBAA ²²⁹		[Mmim] ⁺	[F] ⁻ , [Cl] ⁻ , [BF ₄] ⁻ , [NO ₃] ⁻	AA, P	-
Margulis <i>et al.</i> ²³⁰	OPLS-AA	1-methoxyethylpyridinium	[N(CN) ₂] ⁻	AA, P	-
Yan <i>et al.</i> ¹⁷⁵	Voth <i>et al.</i> ²¹⁵	[Emim] ⁺	[NO ₃] ⁻	AA, P	-
Yan <i>et al.</i> ^{181, 182}	AMBER	[Emim] ⁺	[NO ₃] ⁻	AA, P	-
Steinhauser <i>et al.</i> ¹⁹²	AMBER and CL&P	[Emim] ⁺	[SO ₃ CF ₃] ⁻	AA, P	-
Yethiraj <i>et al.</i> ²³¹	Schmidt <i>et al.</i> ²³²	[Bmim] ⁺	[BF ₄] ⁻	AA, P	-
Johnson <i>et al.</i> ⁹³	ReaxFF ⁹²	Tetrabutylphosphonium	Glycinate	ReaxFF	-

^a Where AA = all atom, UA = united atom, Non-P = non-polarisable and P = polarisable^b Indicates whether the sum of the partial charge on each ion is equal to 1 (unity), or is below 1 (scaled), either through calculation or 'arbitrary' scaling

interactions and microscopic phenomena within ionic liquids, and their mixtures with solutes, is necessary before embarking on any analysis of their effects on reaction outcome. As will be discussed later, interactions between the ionic liquid and species along the reaction coordinate (that is, the starting materials and transition state) are important in directing the outcome of many reactions.

Ionic liquids have been found to be surprisingly good at dissolving a range of compounds. Early work examined the solubility of aromatic compounds, particularly benzene, when compared to their aliphatic counterparts (for benzene, this is cyclohexane).^{177, 233-238} A classical molecular dynamics study concluded that the high solubility of aromatics is due to the local electrostatic interactions between the ionic liquid (specifically [Mmim][Cl] and [Mmim][PF₆] in this case) and the quadrupole moment of the aromatic species,²³⁵ analogous to cation- π interactions.²³⁹ A subsequent study looking at the solubility of benzene and hexafluorobenzene in the ionic liquid [Mmim][PF₆], found that the specific grouping of the ions is dependent on the orientation of the quadrupole moment; in benzene, which has a region of electron density above and below the molecule and positive charge about the equator, the cations grouped above and below the plane of the solute whilst the anions were found about the equator. Hexafluorobenzene has an inverted quadrupole moment relative to benzene due to the electron withdrawing substituents; that is, positive regions above and below the ring, and electron density about the equator; resulting in opposite grouping of the ions.²³³ A similar molecular dynamics study, looking at a range of fluorinated benzenes, found that when both quadrupole and dipole moments exist there is an antagonistic effect; the solubility will be decreased when both exist, with higher solubilities found when there are solely either quadrupole or dipole moments within the compound.²³⁶

The importance of considering the dipole and quadrupole moments of solutes was also highlighted when looking at the solubility of gases in ionic liquids; this area has gained much attention in recent years, mainly focused around capture and storage of greenhouse gases.²⁴⁰⁻²⁴³ A classical molecular dynamics study investigated the solubility of a number of different gases in the ionic liquid [Emim][N(SO₂CF₃)₂]; they found that both hydrogen sulphide (which has a significant dipole moment) and carbon dioxide (no dipole moment but with a large quadrupole moment), were highly soluble. In contrast, molecular hydrogen and nitrogen, with no dipole or quadrupole moments, were much less soluble. This trend was supported by experimental work, and reinforces the importance of considering the overall electronic distribution of species dissolved in ionic liquids.²⁴⁴

It has also been demonstrated that the nature of the cation and anion is important; a number of experimental and computational studies on the solubility of carbon dioxide and sulfur dioxide have generally found that the solubility of the gas is more dependent on the anion than the cation, and that the anion-carbon dioxide (and anion-sulfur dioxide²⁴⁵) interactions

reduce the extent of anion-cation interactions within the mixture.^{70, 116, 120, 246-249} The anion-solute interactions thus disrupt the ordering of the ionic liquid components, which results in a decreased viscosity of the solution and an increase in the diffusion coefficients of both the ionic liquid components and the dissolved gas.^{241, 242, 247, 249} This brings about another important point of reflection; considering the balance of ion-solute and ion-ion interactions within the mixture. In the cases described here, the balance of these interactions resulted in a disruption in ordering of the ionic liquid upon dissolution of the gas.

Another area which has received considerable attention is the dissolution of lignocellulosic biomass in ionic liquids; lignocellulosic biomass is defined as any material obtained from currently or previously living flora, and its dissolution is an area of increasing interest due to the possibility of utilising waste materials for energy production and as a source of aromatic starting materials.²⁵⁰ A number of studies have demonstrated that lignin, cellulose and other biomass components are much more soluble in ionic liquids than traditional solvents.²⁵¹⁻²⁵⁴ A range of computational studies in this area have sought to understand the role of the anion and cation in the dissolution process.^{255, 256} A combination of experimental and DFT studies found that the greater the hydrogen bond accepting ability of the anion, the greater the solubility of lignin (a highly complex, aromatic species). Interestingly it was found that there was a subtle balance between the ionic liquid ions interacting with themselves, and with the biomass components; increasing the coordinating ability of the anion would result in an increase in both these interactions.²⁵⁶ An experimental study found that the nature of the cation also had an effect on lignin solubility,²⁵⁷ with the importance of $\pi-\pi$ interactions between an aromatic cation and lignin also demonstrated through DFT calculations.¹²⁵

For cellulose, a series of classical molecular dynamics simulations found that the ionic liquid cations interact more with the non-polar regions of cellulose, whilst the anions interact with the polar regions.²⁵⁵ The general conclusion of this study is that the solvation of cellulose is driven by the strong interaction between the anion and the polar regions of cellulose; this idea is supported by other computational^{258, 259} and experimental²⁶⁰ studies where it has been concluded that the anion disrupts the intermolecular forces between the cellulose polymers, resulting in increased solubility.

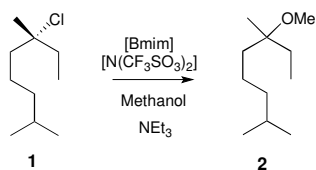
Highlighted by the examples above, there are some important considerations that become apparent when beginning to analyse the effects of ionic liquids on reaction outcome. These include consideration of: interactions between the ionic liquid and specific sites on the species along the reaction coordinate based on electronic distribution (such as charged species, and dipole and quadrupole moments); the competing interaction between the ions themselves and between ions and any dissolved species; the existence of hydrogen bonding interactions, if any; and disruption in the ordering of the system, whether it be disruption of the ionic liquid ordering or

of any strong interactions that exist with the reagents. The following discussion will be with reference to, and in the context of, these points.

Studies on reaction outcome

UNIMOLECULAR SUBSTITUTION REACTIONS

Ionic liquids have been investigated with respect to their impact on the outcome of substitution reactions proceeding through a unimolecular rate determining step involving a neutral nucleophile; generally, high mole fractions of ionic liquid slow the reaction down when compared to polar protic solvents.^{8, 9, 16, 261, 262} An important example is the reaction between the linalool derivative **1** and methanol (Scheme 1), the rate constant of which was determined in mixtures of the ionic liquid [Bmim][N(SO₂CF₃)₂] and methanol.⁸



Scheme 1: The methanolysis reaction of the chloride **1** to give the ether **2**. The rate constant was found to vary when the mole fraction of [Bmim][N(SO₂CF₃)₂] in methanol was changed.^{8, 9}

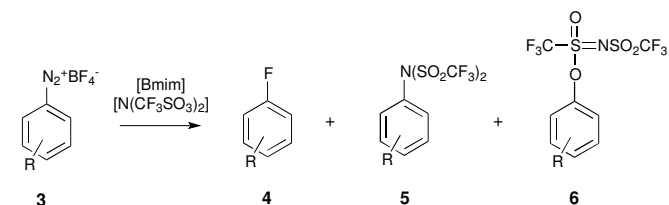
It was found that on addition of a small amount of [Bmim][N(SO₂CF₃)₂] to methanol (χ_{IL} ca. 0.02) there was a two-fold increase in the rate constant. Further increases in the ionic liquid concentration resulted in a decreased rate constant, and at χ_{IL} 0.72 the rate constant was half that found in methanol.⁸ These observations were rationalised using activation parameter data determined from temperature-dependent kinetic experiments, molecular dynamics simulations, and by considering the electrostatic interactions that act to stabilise the incipient charges in the transition state.⁹ At low concentrations, this charge stabilisation results in a faster rate through a lowering of the enthalpy of activation - an enthalpic benefit. With increasing amounts of ionic liquid there is a degree of solvent organisation when moving to the transition state, resulting in a decrease in the entropy of activation - an entropic cost. At higher concentrations this substantial entropic cost outweighs the enthalpic benefit, causing an overall decrease in reaction rate. It is interesting to note that in both kinetic¹⁶ and experimental²⁶² studies on unimolecular substitution reactions that involve less charge localisation in the transition state, the use of ionic liquids increased the reaction rate relative to molecular solvents.

Classical molecular dynamics simulations were used to further investigate the microscopic origin of these effects found for the reaction shown in Scheme 1. Geometry optimisations were performed at the B3LYP/6-31G(d) level to obtain the substrate and intermediate structures, and the simulations were done using the CL&P force field. [Mmim][PF₆] (an ionic liquid comparable to [Bmim][N(SO₂CF₃)₂], used because it was small and tractable) was modelled about both the chloride **1** and the carbocation intermediate of the process. The simulations

showed a significant increase in ordering about the intermediate, specifically organisation of the ionic liquid cation about the chloride leaving group, and the anion of the ionic liquid about the carbocation intermediate, when compared to the non-dissociated species **1**.⁹ This is consistent with the experimental observations.

It should be noted that the above case involves consideration of the intermediate in the process, not the transition state, due to computational limitations in modelling a transition state. This is not an unreasonable approximation (consider Hammond's postulate), given that how the transition state would vary between solvents would be hard to model.

A few reports have also examined the effect of ionic liquids on reactions involving a charged nucleophile,²⁶³⁻²⁶⁶ these studies all focused on a dediazonation reaction, where it was found that the non-nucleophilic anion [N(SO₂CF₃)₂]⁻ was reacting in preference to the more nucleophilic halide reagents (Scheme 2), with the product **6** forming predominantly.



Scheme 2: The fluorodediazotiation of the diazonium salt **3** to give either product **4** due to reaction with the fluoride nucleophile, or products **5** and **6** due to reaction with the anion of the ionic liquid either at the nitrogen or at the oxygen, respectively.²⁶³

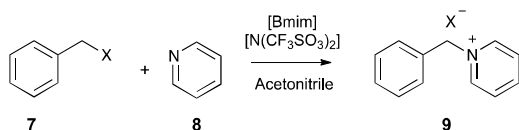
DFT calculations showed that the product **6** was also less thermodynamically stable, indicating that the selectivity was kinetically driven.²⁶⁴ Further work^{265, 266} was able to demonstrate that the apparent 'increased nucleophilicity' of [N(SO₂CF₃)₂]⁻ is actually a result of a decrease in the relative nucleophilicity of the halide in the ionic liquid. The charge dense and coordinating halides interact more strongly with the [Bmim]⁺ cation, relative to [N(SO₂CF₃)₂]⁻; as a result there is much more 'free' [N(SO₂CF₃)₂]⁻ available to react with the aromatic diazonium salt **3**, hence the apparent increase in the nucleophilicity of [N(SO₂CF₃)₂]⁻. This concept is supported by a number of computational^{105, 139, 140} and experimental^{40, 166} studies that show that more coordinating anions will interact more strongly with, and will have a greater degree of charge transfer to, the cation, when compared to less coordinating anions. This reinforces the importance of considering the competing interactions between the components of the ionic liquid themselves, and the ions and reagent.

Overall, studies so far on unimolecular substitution processes have revealed two key points: firstly, the importance of considering the electrostatic interactions between the ionic liquid and the charge-separated transition state, leading to the ionic intermediate. The magnitude of such an interaction has a marked effect on the reaction rate when compared to molecular solvents; with studies suggesting that strong intermediate-ionic liquid interaction can decrease the rate constant,^{8, 9} while

lessened extents of intermediate-ionic liquid interaction can have a positive influence on the rate constant.^{16, 262} Secondly, that having an understanding of the different coordination strength of the ions in the reaction mixture is essential, as when using a mixture of ionic species there can be preferential interaction of ionic components.²⁶³⁻²⁶⁶

BIMOLECULAR SUBSTITUTION REACTIONS

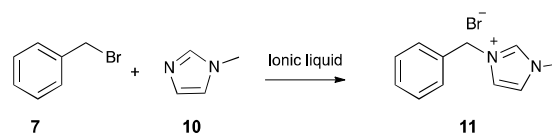
There have also been a number of investigations into the effect of ionic liquids on the outcome of substitution reactions that proceed through a single step. This has included the effect on reactions containing both charged and neutral electrophiles and/or nucleophiles, although more attention has been given to those involving neutral reagents. Investigations into the bimolecular substitution reaction between benzyl halides **7** (Br and Cl) and pyridine **8** (Scheme 3) found that the rate constant of the reaction gradually increased as the mole fraction of the ionic liquid [Bmim][N(SO₂CF₃)₂] in the solvent mixture (with acetonitrile) was increased,^{16, 19} with both the enthalpy and entropy of activation increasing when the ionic liquid was added.^{12, 15, 16}



Scheme 3: The bimolecular substitution reaction between benzyl halide **7** (where X = Cl or Br) and pyridine **8**, performed in different mole fractions of [Bmim][N(SO₂CF₃)₂] in acetonitrile.^{12, 15, 16, 19}

This trend suggests that there is an increase in disorder on moving to the transition state, with the increased enthalpy arising from either stabilisation of the starting materials or destabilisation of the transition state. Given the high solubility of aromatic compounds in ionic liquids (as discussed previously), it was initially suspected that the increased enthalpy was due to ordering about the delocalised π systems on reagents **7** and **8**.¹² A series of deconvolution and classical molecular dynamics studies using the CL&P force field showed that, whilst there is a significant cation-quadrupole interaction, the extent of organisation about the starting material **8** does not change significantly with the removal of the delocalised π system.¹⁵ Hence, the organisation about the nucleophilic nitrogen centre in the amine **8** was key as, on moving to the transition state, interaction with this centre is no longer possible.

In a study on a similar Menshutkin reaction (Scheme 4),²⁶⁷ the effect of a selection of ionic liquids on the rate constant was found to be comparable to that discussed above. DFT calculations of a series of single ion pairs, using the B3LYP/CEP-121G(d,p) level of theory, were used to investigate interactions with the transition state. This study found that the primary interaction in the transition state was that between the ionic liquid cation and the halide leaving group.



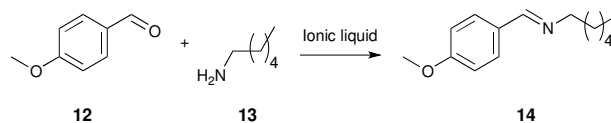
Scheme 4: The bimolecular substitution reaction between benzyl bromide **7** and N-methylimidazole **10**, performed in a range of ionic liquids.²⁶⁷

By considering both of these studies^{12, 15, 267} it is clear that whilst the cation interacts with both the nucleophile starting material and the anionic leaving group in the transition state, the trend in the activation parameters indicate that organisation about, and stabilisation of, the nucleophile is more substantial, resulting in the significant entropic benefit responsible for the increased rate constant.^{14, 184}

Another aspect that is interesting to consider is the effect that changing the ionic liquid itself has on reaction outcome. Whilst many experimental studies have focussed on these, there is less related computational data.

The effect of changing both the cation and anion on the substitution reaction between benzyl bromide **7** and pyridine **8** has been investigated in a range of ionic liquids.^{18, 20} An important outcome when altering the cation was that, when systemically replacing the imidazolium ring hydrogens with methyl groups, there was no change in the enthalpy of activation, indicating that there is no directional interaction between the ring hydrogens and the nucleophile **8**.¹⁸ This is supported by calculations on imidazolium and methylated imidazolium ionic liquids,^{100, 103, 268} along with other computational studies, suggesting that all the ring hydrogens, as well as the methylene and methyl protons in the alkyl substituents, are capable of hydrogen bonding, and therefore methyl substitution has little effect.^{102, 269}

The condensation reaction between 4-methoxybenzaldehyde **12** and hexylamine **13** (Scheme 5) is not a substitution process, but the first step involves nucleophilic attack on a carbonyl centre and, hence, it is mechanistically related to the previous example. Further, it is affected by ionic liquids in the same manner as the substitution case shown in Scheme 3, with a comparable interaction between the cation and the nitrogen lone pair on the amine **13** inferred by the enthalpic benefit and entropic cost on moving to the ionic solvent.²¹ This case was more affected by modification of the ionic liquid cation; this was rationalised by the fact that, due to the less bulky nature of the nucleophile **13** when compared to the more rigid pyridine **8**, the nucleophile can better access the charged cationic centre.



Scheme 5: The condensation reaction between 4-methoxybenzaldehyde **12** and hexylamine **13**, performed in a range of different ionic liquids where the charge accessibility and density of the cation, and the coordinating ability of the anion, was varied.^{21, 270}

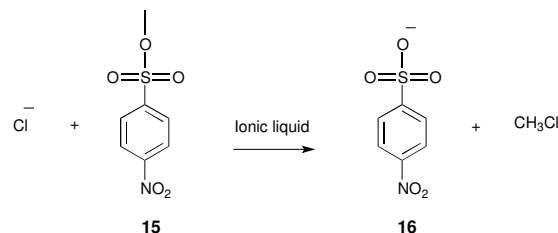
An interesting result from this work was that the extent of cation–nucleophile interaction could be controlled in a systematic fashion by varying the nature of both the cation²¹ and anion.²⁷⁰ Generally the more accessible and charge dense the cation, and the lower the coordinating ability of the anion, the larger the associated enthalpy and entropy of activation.^{21, 270} A more subtle observation was that the enthalpy of activation was slightly less for [Bm₂im][N(SO₂CF₃)₂] than [Bmim][N(SO₂CF₃)₂], due to the steric bulk of the methyl group inhibiting access to the charged centre. Yet the entropy of activation was the same for both ionic liquids, rather than [Bmim][N(SO₂CF₃)₂] being higher as would be expected.

These results can be explained by the numerous computational studies on [Bmim]⁺ and [Bm₂im]⁺ based systems, aimed at explaining the anomalous melting points and viscosities found for [Bm₂im]⁺ ionic liquids.^{100, 101, 103, 268, 271, 272} Methyl substitution at the C2 position was anticipated to decrease the extent of hydrogen bonding in the ionic liquid, causing a decrease in the melting point and viscosity. Yet [Bm₂im]⁺ ionic liquids have higher melting points and viscosities than their [Bmim]⁺ counterparts.^{101, 271, 272} A number of quantum chemical calculations (B3LYP and MP2) of clusters of ionic liquids found that the C2-methylated systems have fewer stable configurations than their protonated versions, with larger energy differences between the stable conformers when compared to the protonated cases. This restricts movement of the anion and results in the observed increased viscosities and melting points, as there is an increase in ordering of the system.^{100, 103, 268} The main conclusion from these studies is that entropy is more significant for the C2 methylated ionic liquids than for the corresponding unsubstituted cases.

Linking this back to the condensation reaction discussed above, the similar entropies can now be explained by the increased solvent ordering in [Bm₂im][N(SO₂CF₃)₂], compared to [Bmim][N(SO₂CF₃)₂]. Ordering within the [Bm₂im][N(SO₂CF₃)₂] solvent shell about the nucleophile **13** will be more substantial than expected if *only* considering the ordering caused by cation–nucleophile interaction. When moving to the transition state, where the cation will move away from the nucleophile, the ordering within this solvent shell will be disrupted significantly. This change in ordering is greater than what would be expected if only considering the extent of cation–nucleophile interaction, resulting in the entropies of activation for [Bm₂im][N(SO₂CF₃)₂] and [Bmim][N(SO₂CF₃)₂] being comparable.

There have also been studies on bimolecular substitution processes that involve charged reagents; of particular interest here is the reaction of methyl *p*-nitrobenzenesulfonate **15** with a number of charged nucleophiles, which has been investigated both experimentally^{25, 273–275} and computationally.²⁷⁶ When performed in molecular solvents this process can proceed through either the discrete anion, where the nucleophile is not coordinated to a counterion, or as an ion pair when a higher reagent concentration is used.²⁷⁷ When using an ionic liquid solvent it was found that the reaction proceeded faster than

when in the highly polar, protic hexafluoropropan-2-ol but slower than the relatively non-polar solvent dichloromethane. The activation parameters for the reaction with a chloride anion (Scheme 6) were determined in a number of ionic liquids, and compared with both the ion pairing and discrete anion mechanism in dichloromethane to gain an understanding of the microscopic origin of this effect.²⁵



Scheme 6: The reaction of *p*-nitrobenzenesulfonate **15** with chloride to give the product **16**, performed in a number of molecular solvents and ionic liquids.^{25, 273, 276}

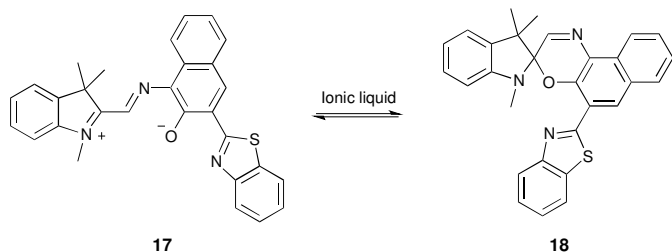
It was found that the entropy of activation for the ionic liquids cases was close to that of the free ion in dichloromethane, while the enthalpy of activation more closely resembled the ion pairing mechanism.²⁵ A QM/MM study²⁷⁶ investigated this reaction further in [Bmim][PF₆], and found that the [Bmim]⁺ cations are strongly associated with the chloride nucleophile; such an interaction would stabilise the nucleophile and introduce an enthalpic cost, resulting in an enthalpy of activation comparable to the ion pairing mechanism, as an ion will need to be ‘removed’ from the nucleophile for the reaction to proceed. It was also suggested that there is a hydrogen-bonding interaction between the ionic liquid cation and [Cl]⁻;²⁷⁶ this idea is supported by the experimental work that found that the rate constant increased as the hydrogen bond donor ability of the cation decreased.²⁵ Further, the computational study suggested that this interaction was entropically favourable as the [Bmim]⁺ - [Cl]⁻ distance increased on moving to the transition state, yet this was offset by interaction between [Bmim]⁺ and the leaving group in the transition state, resulting in the entropy of activation more closely resembling that of the free ion in dichloromethane. In other words, the entropic benefit will be minimal as the ionic liquid cation can interact with both the charged nucleophile and the charge separated transition state.

For these bimolecular processes the importance of the extent of electrostatic interactions between the ionic liquid components and the species along the reaction coordinate, as seen for the unimolecular case, is once again demonstrated. Interestingly, the combination of experimental and computational work has provided significant insight into the subtle interactions between the ionic liquid and the reagents’ lone pair, and the effect that ordering within the ionic liquid itself can have on reaction outcome. Importantly, when using a neutral nucleophile use of an ionic liquid resulted in an entropically driven rate increase, yet when using a charged nucleophile the enthalpic barrier for the process was higher due to the strong interactions with the charged nucleophile.

Further, for the charged nucleophile cases there was significant interaction between the ionic liquid and the more charge-separated transition state, resulting in any entropic effects to be minimal.

CYCLISATION

The cyclisation of the merocyanine **17** to the spiroxazine **18** (Scheme 7) proceeds through an intramolecular nucleophilic attack of a phenoxide onto an iminium moiety. This can be compared to the condensation case discussed above and is a slightly unusual example of a nucleophilic 1,2-addition. This reaction was found to proceed slower in a number of ionic liquids, compared to molecular solvents.^{278, 279}



Scheme 7: Cyclisation of the merocyanine **17** to give the spiroxazine **18**.²⁷⁸⁻²⁸¹

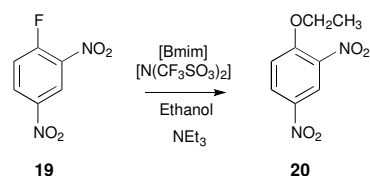
Whilst it generally appears that the decreased rate is due to an increased enthalpy of activation, partially offset by an increased entropy of activation, when compared to polar aprotic molecular solvents, no uncertainties in the data were reported and the activation parameters covered a large range of values for different ionic liquids.²⁷⁸ They also noted that the observed effects could not be correlated with any measured solvent parameters. Through a combined UV-Vis and computational (*ab initio* and DFT) approach on a related system, it was suggested that the nature of both the cation²⁸⁰ and the anion²⁸¹ played a role. It was found that the ionic liquid cation could interact with the phenolate oxygen, stabilising the derivative of **17** and shifting the equilibrium position to the left.²⁸⁰ Considering this cation-phenolate interaction, and the increased enthalpy and entropy of activation for the reaction in Scheme 7 when using an ionic liquid, it is likely that there is a similar effect to that discussed above for the bimolecular substitution cases. Interaction between the cation and the phenolate on **17** would stabilise this reagent, decreasing its nucleophilicity and hence increase the enthalpic barrier for the reaction. Whilst there is an entropic benefit associated with such an interaction, due to the increased disorder on moving to the transition state and considering that the rate constant is decreased relative to molecular solvents, the entropic advantage for this reaction is not sufficient to offset the enthalpic cost (in contrast to what was seen for the reactions in Schemes 3-5, and again emphasises the importance of the balance of these interactions).

NUCLEOPHILIC AROMATIC SUBSTITUTION REACTIONS

Nucleophilic aromatic substitution reactions are of particular interest given the high solubility of aromatic compounds in

ionic liquids; as strong interactions between the quadrupole moment of the electrophile and the ionic liquid might be expected in these cases.

The ethanolysis of 1-fluoro-2,4-dinitrobenzene **19** (Scheme 8) was conducted in ethanol and the ionic liquid [Bmim][N(SO₂CF₃)₂] (χ_{IL} ca. 0.5) across a range of temperatures to allow the activation parameters to be determined. When using the ionic liquid, there was a small enthalpic cost that was offset by a more substantial entropic benefit when compared to ethanol, with an overall increase in the rate constant.¹⁴

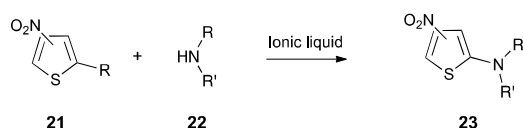


Scheme 8: The nucleophilic aromatic substitution reaction between the substituted benzene **19** and ethanol, to give the product **20**.¹⁴

Molecular dynamics simulations, using the CL&P force field for the ionic liquid components, were performed on the starting material **19** and the corresponding Meisenheimer intermediate with either [Bmim][N(SO₂CF₃)₂] or ethanol. The radial distribution functions showed that [Bmim][N(SO₂CF₃)₂] was more ordered about the starting material than the intermediate, and that ethanol showed little organisation about either species under the same simulation conditions. Particularly, it was noted that the anion – π interaction was most significant, with a decreased extent of ordering about the intermediate, whilst the extent of ordering of the cation changed little. The importance of the anion – π interaction can be rationalised by recognising that the inclusion of electron-withdrawing substituents will reorient the quadrupole moment of the aromatic system relative to benzene, resulting in grouping of the anion above and below the plane of the reagent **19**, and the cations about the equator.²³³ This phenomena is reflected in the molecular dynamics simulations of this fluoro- and nitro-substituted benzene, and confirms that the trend in the activation parameters is due to increased ordering of the anion about the delocalised π -system of the reagent, relative to the disrupted π -system of the transition state.¹⁴

This work was extended to investigate the effect of changing the cation and anion of the ionic liquid.¹⁷ Generally, there were minimal changes in the activation parameters when varying the cation; a logical result given the localisation of cations about the aromatic-equator, an interaction which is unlikely to be largely affected by a disruption in aromaticity.^{14, 233} It was anticipated that changing the coordinating ability of the anion would have a more significant effect on the activation parameters, due to the likely ordering above and below the plane of the aromatic species **19**. This expected trend did hold for single ionic liquid systems, but not for binary mixtures of ionic liquids.¹⁴

Another interesting example is the amination of activated thiophenes **21** using a series of different secondary cyclic amines **22** (for example, piperidine, pyrrolidine and morpholine), shown in Scheme 9.²⁷ It was found that the reaction proceeded faster in the ionic liquids [Bmim][BF₄], [Bmim][PF₆] and [Bm₂m][BF₄] than it did in molecular solvents. The experimentally-determined activation parameters indicated that this rate enhancement was enthalpically driven, and it was suggested that this was due to greater stabilisation of the transition state by the ionic liquids, when compared to molecular solvents. Interestingly, it was also suggested that the situation was likely more complex, and factors such as differing extents of nucleophile solvation needed to be taken into account.²⁷



Scheme 9: The nucleophilic aromatic substitution reaction between the substituted thiophenes **21** and a number of cyclic, secondary amines **22** (for example, R/R' = -(CH₂)₄-, -(CH₂)₅-, -(CH₂)₂O(CH₂)₂-).^{27, 282}

A recent QM/MM study has appeared, which aimed to provide further insight into the microscopic origin of the rate changes described above for the reaction in Scheme 9. The calculated activation parameters closely matched the experimentally derived parameters, once again concluding that the rate increase was due to an enthalpic benefit.²⁸² The computational work also suggested an increased amine nucleophilicity in the ionic liquids, compared to methanol, as well as significant π - π interactions between the [Bmim]⁺ cation and nitrothiophene **21** substrate. Interestingly, they also found that the transition state in the ionic liquid more closely resembled the Meisenheimer complex than it did in methanol, and that there was a significant increase in the electrostatic interactions between the ionic liquid and the transition state, compared to the reagents. This increased interaction with the transition state, when using an ionic liquid, resulted in the enthalpic benefit and entropic cost seen.²⁸²

The main points to take away from studies on nucleophilic aromatic substitution processes are that: firstly, (again) the importance of considering electrostatic interactions between the ionic liquid and charged species along the reaction coordinate, similar to that seen for unimolecular substitution processes, is identified. And secondly, for processes that involve a change in aromaticity in the rate-determining step, differing extents of solvation of these aromatic species along the reaction pathway can affect reaction outcome when using ionic liquids. This concept links to the solubility studies discussed earlier, where the importance of ion-quadrupole and ion-dipole interactions have been realised.^{233, 235, 236, 244}

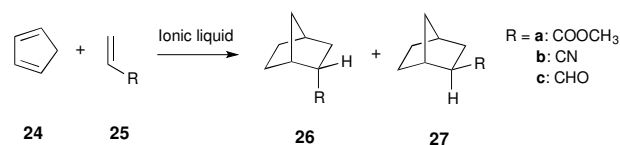
DIELS-ALDER CYCLOADDITIONS

There is much literature on the Diels-Alder [4+2] cycloaddition reaction, with many publications focusing on the effect that

ionic liquids have on the rate and the selectivity of the process. The general conclusion from the many experimental studies is that the rate constant and *endo* selectivity is increased relative to non-polar solvents, but decreased when compared to water.^{24, 283, 284}

A detailed study on the reaction between cyclopentadiene **24** and methyl acrylate **25a** (Scheme 10) in a number of ionic liquids concluded that the increased rate and selectivity for isomer **26** over isomer **27**, when compared to a non-polar solvent, is due to hydrogen bonding between the ionic liquid and the dienophile **25**. The greater the extent of hydrogen bonding with the dienophile **25**, the more selective the reaction was and the more the reaction was accelerated.²³

The complex balance between the ions interacting with themselves, and with the dienophile, was also highlighted and has been raised by many authors.^{247, 249, 265, 285} The highest selectivity and rate were achieved when using a cation with a strong hydrogen bond donating ability, and an anion that is a poor hydrogen bond acceptor.²³ Attempts to develop Linear Solvation Energy Relationships (LSERs) when using dienophiles **25a**, **25b** and **25c** (Scheme 10) found that whilst the selectivity fit was quite good, correlation with the rate constant was not as successful.²⁸⁶ The limitations of this approach were also highlighted in a study on an intramolecular case, where simple use of LSERs was not sufficient to describe the observed rate constants and selectivity, and much better fits were seen when viscosity was taken into account.²⁸⁷



Scheme 10: The Diels Alder reaction between cyclopentadiene **24** and a number of dienophiles **25**.^{23, 286}

A number of computational studies have been used to further investigate the observed selectivity and reactivity; DFT calculations (B3LYP/cep-121g(d,p)) of the cation with the reagents, transition state and product were performed on the reaction between **24** and dienophiles **25** (Scheme 10).²⁸⁸ There were two key conclusions from this work; first, that the cation coordinates to the dienophile, which could activate it towards cycloaddition as has previously been suggested,²⁸⁹ and second, that the dienophile is held rigid due to this interaction. This essentially 'clamps' the dienophile in place, increasing the probability of dienophile-diene stacking that is necessary for the reaction to proceed.^{285, 288} The interaction of the dienophile and the cation supports the hydrogen bonding argument discussed above. Further investigations (using KS-DFT/3D-RISM-SCF^{290, 291}) of the reaction of cyclopentadiene **24** and acrolein **25c** in the ionic liquid [Mmim][PF₆] showed that coordination of the cation was not essential, and in some cases the preferential interaction of the ionic liquid components with each other were responsible for the changes in reaction outcome.²⁹² Interestingly, it was noted

that interactions between the carbonyl oxygen of acrolein **25c** and the C4, C5 and methyl hydrogens of the cation [Mmim]⁺ were more significant than interaction with the C2 hydrogen, reinforcing the idea that interaction with the C2 proton is not fundamental.^{18, 269} Overall, the solvent effects of the Diels Alder reaction are complex, and a complete understanding of the role that an ionic liquid plays in dictating reaction outcome is still developing.

In summary, the extensive experimental and computational work over recent years has contributed significantly to the understanding of ionic liquids, and their effect on the outcome of organic processes. The major interactions identified so far as governing ionic liquid solvent effects are: 1) the electrostatic interactions between the ionic liquid and charged species along the reaction coordinate; 2) interactions between the ionic liquid, and dipole and quadrupole moments of the reagents; and 3) interactions between the ionic liquid and lone pairs, particularly in reagents. Once these features have been identified, careful consideration of the magnitude of these different interactions, as well as their position along the reaction coordinate, can allow for predictions to be made about how the ionic liquid will affect the activation parameters for the process, and hence the rate constant. While such a predictive framework is by no means comprehensive and is still developing, the growing interest in this area is bringing us closer to having a full understanding of ionic liquid solvent effects.

Conclusions

This review has summarised the major computational techniques that can be used to investigate ionic liquid solvent effects, and identifies both the utility and limitations of the computational approaches commonly used. There are many different forms of computational studies that can be used to compliment experimental work; whether it be studies focusing on the nature of ionic liquids themselves, such as the complex landscape of interactions that exist in ionic liquids; or on interactions between ionic liquids and dissolved species. The combination of experimental and computational work has greatly increased our fundamental understanding of ionic liquids as solvents, and has contributed substantially to the developing predictive framework for using ionic liquids as solvents for organic reactions. The increasing popularity of computational approaches, as well as advancing computing technologies, will hopefully see the emergence of more combined computational and experimental studies, as such an approach is crucial for gaining a thorough understanding of ionic liquid media and their effect on reaction outcome.

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Notes and references

^a School of Chemistry, University of New South Wales, Sydney, NSW 2052, Australia. E-mail: j.harper@unsw.edu.au

^b Department of Chemical and Environmental Engineering, University of Nottingham, University Park, Nottingham, NG7 2RD, United Kingdom. E-mail: anna.croft@nottingham.ac.uk

1. S. Z. El Abedin and F. Endres, *Acc. Chem. Res.*, 2007, **40**, 1106-1113.
2. J. P. Hallett and T. Welton, *Chem. Rev.*, 2011, **111**, 3508-3576.
3. K. R. Seddon, *Kinet. Catal. Engl. Transl.*, 1996, **37**, 693-697.
4. C. L. Hussey, *Pure Appl. Chem.*, 1988, **60**, 1763-1772.
5. M. J. Earle, J. M. S. S. Esperanca, M. A. Gilea, J. N. Canongia Lopes, L. P. N. Rebelo, J. W. Magee, K. R. Seddon and J. A. Widegren, *Nature*, 2006, **439**, 831-834.
6. D. M. Fox, J. W. Gilman, A. B. Morgan, J. R. Shields, P. H. Maupin, R. E. Lyon, H. C. De Long and P. C. Trulove, *Ind. Eng. Chem. Res.*, 2008, **47**, 6327-6332.
7. B. Wu, W. Liu, Y. Zhang and H. Wang, *Chem. Eur. J.*, 2009, **15**, 1804-1810.
8. B. Y. W. Man, J. M. Hook and J. B. Harper, *Tetrahedron Lett.*, 2005, **46**, 7641-7645.
9. H. M. Yau, S. A. Barnes, J. M. Hook, T. G. A. Youngs, A. K. Croft and J. B. Harper, *Chem. Commun.*, 2008, 3576-3578.
10. C. E. Rosella and J. B. Harper, *Tetrahedron Lett.*, 2009, **50**, 992-994.
11. H. M. Yau, S. J. Chan, S. George, J. Hook, A. Croft and J. Harper, *Molecules*, 2009, **14**, 2521-2534.
12. H. M. Yau, A. G. Howe, J. M. Hook, A. K. Croft and J. B. Harper, *Org. Biomol. Chem.*, 2009, **7**, 3572-3575.
13. S. R. D. George, G. L. Edwards and J. B. Harper, *Org. Biomol. Chem.*, 2010, **8**, 5354-5358.
14. S. G. Jones, H. M. Yau, E. Davies, J. M. Hook, T. G. A. Youngs, J. B. Harper and A. K. Croft, *Phys. Chem. Chem. Phys.*, 2010, **12**, 1873-1878.
15. H. M. Yau, A. K. Croft and J. B. Harper, *Faraday Discuss.*, 2012, **154**, 365-371.
16. S. T. Keaveney and J. B. Harper, *RSC Adv.*, 2013, **3**, 15698-15704.
17. E. E. L. Tanner, R. R. Hawker, H. M. Yau, A. K. Croft and J. B. Harper, *Org. Biomol. Chem.*, 2013, **11**, 7516-7521.
18. E. E. L. Tanner, H. M. Yau, R. R. Hawker, A. K. Croft and J. B. Harper, *Org. Biomol. Chem.*, 2013, **11**, 6170-6175.
19. H. M. Yau, S. T. Keaveney, B. J. Butler, E. E. L. Tanner, M. S. Guerry, S. R. D. George, M. H. Dunn, A. K. Croft and J. B. Harper, *Pure Appl. Chem.*, 2013, **85**, 1979.
20. S. T. Keaveney, Francis, D. V., Cao, W., Haines, R. S., and Harper, J. B., *Aust. J. Chem.*, 2015, **68**, 31-35.
21. S. T. Keaveney, K. S. Schaffarczyk McHale, R. S. Haines and J. B. Harper, *Org. Biomol. Chem.*, 2014, **12**, 7092.
22. B. J. Butler and J. B. Harper, *New J. Chem.*, 2015, **39**, 213-219.
23. A. Aggarwal, N. L. Lancaster, A. R. Sethi and T. Welton, *Green Chem.*, 2002, **4**, 517-520.

24. T. Fischer, A. Sethi, T. Welton and J. Woolf, *Tetrahedron Lett.*, 1999, **40**, 793-796.
25. N. L. Lancaster, P. A. Salter, T. Welton and G. B. Young, *J. Org. Chem.*, 2002, **67**, 8855-8861.
26. I. Newington, J. M. Perez-Arlandis and T. Welton, *Org. Lett.*, 2007, **9**, 5247-5250.
27. F. D'Anna, V. Frenna, R. Noto, V. Pace and D. Spinelli, *J. Org. Chem.*, 2006, **71**, 5144-5150.
28. F. D'Anna, V. Frenna, V. Pace and R. Noto, *Tetrahedron*, 2006, **62**, 1690-1698.
29. F. D'Anna, S. La Marca, P. Lo Meo and R. Noto, *Chem. Eur. J.*, 2009, **15**, 7896-7902.
30. S. T. Keaveney, Haines, R. S., Harper, J. B., in *Encyclopedia of Physical Organic Chemistry*, ed. U. Wille, Wiley, in press.
31. M. J. Earle, S. P. Katdare and K. R. Seddon, *Org. Lett.*, 2004, **6**, 707-710.
32. Y. Wang and G. A. Voth, *J. Am. Chem. Soc.*, 2005, **127**, 12192-12193.
33. K. Shimizu, C. E. S. Bernardes, A. Triolo and J. N. Canongia Lopes, *Phys. Chem. Chem. Phys.*, 2013, **15**, 16256-16262.
34. A. B. Pereira, M. J. Pastoriza-Gallego, K. Shimizu, I. M. Marrucho, J. N. C. Lopes, M. M. Piñeiro and L. P. N. Rebelo, *J. Phys. Chem. B*, 2013, **117**, 10826-10833.
35. K. Shimizu, C. E. S. Bernardes and J. N. Canongia Lopes, *J. Phys. Chem. B*, 2013, **118**, 567-576.
36. B. Schwenzler, S. N. Kerisit and M. Vijayakumar, *RSC Adv.*, 2014, **4**, 5457-5464.
37. G. Chatel, J. F. B. Pereira, V. Debbeti, H. Wang and R. D. Rogers, *Green Chem.*, 2014, **16**, 2051-2083.
38. H. Niedermeyer, J. P. Hallett, I. J. Villar-Garcia, P. A. Hunt and T. Welton, *Chem. Soc. Rev.*, 2012, **41**, 7780-7802.
39. R. S. Payal and S. Balasubramanian, *Phys. Chem. Chem. Phys.*, 2013, **15**, 21077-21083.
40. I. J. Villar-Garcia, K. R. J. Lovelock, S. Men and P. Licence, *Chem. Sci.*, 2014, **5**, 2573-2579.
41. G. E. Logotheti, J. Ramos and I. G. Economou, *J. Phys. Chem. B*, 2009, **113**, 7211-7224.
42. G. Adamova, J. N. Canongia Lopes, L. P. N. Rebelo, L. M. N. B. Santos, K. R. Seddon and K. Shimizu, *Phys. Chem. Chem. Phys.*, 2014, **16**, 4033-4038.
43. S. M. Urahata and M. C. C. Ribeiro, *J. Chem. Phys.*, 2004, **120**, 1855-1863.
44. E. Androulaki, N. Vergadou and I. G. Economou, *Mol. Phys.*, 2014, 1-13.
45. S. P. Verevkin, D. H. Zaitsau, V. N. Emel'yanenko, A. V. Yermalaye, C. Schick, H. Liu, E. J. Maginn, S. Bulut, I. Krossing and R. Kalb, *J. Phys. Chem. B*, 2013, **117**, 6473-6486.
46. S. Li, K. L. Van Aken, J. K. McDonough, G. Feng, Y. Gogotsi and P. T. Cummings, *J. Phys. Chem. C*, 2014, **118**, 3901-3909.
47. S. Yeganegi, A. Soltanabadi and D. Farmanzadeh, *J. Phys. Chem. B*, 2012, **116**, 11517-11526.
48. Z. Fei, D.-R. Zhu, N. Yan, R. Scopelliti, S. A. Katsuba, G. Laurenczy, D. M. Chisholm, J. S. McIndoe, K. R. Seddon and P. J. Dyson, *Chem. Eur. J.*, 2014, **20**, 4273-4283.
49. F. Dommert, K. Wendler, R. Berger, L. Delle Site and C. Holm, *ChemPhysChem*, 2012, **13**, 1625-1637.
50. E. J. Maginn, *J. Phys.: Condens. Matter*, 2009, **21**, 373101.
51. M. Salanne, *Phys. Chem. Chem. Phys.*, 2015, DOI: 10.1039/C1034CP05550K.
52. B. Kirchner, O. Hollóczki, J. N. Canongia Lopes and A. A. H. Pádua, *WIREs Comput. Mol. Sci.*, 2015, **5**, 202-214.
53. C. D. Hubbard, P. Illner and R. van Eldik, *Chem. Soc. Rev.*, 2011, **40**, 272-290.
54. C. Chiappe and C. S. Pomelli, *Phys. Chem. Chem. Phys.*, 2013, **15**, 412-423.
55. J. S. Wilkes, J. A. Levisky, R. A. Wilson and C. L. Hussey, *Inorg. Chem.*, 1982, **21**, 1263-1264.
56. K. R. Seddon, *J. Chem. Technol. Biotechnol.*, 1997, **68**, 351-356.
57. K. Fumino and R. Ludwig, *J. Mol. Liq.*, 2014, **192**, 94-102.
58. R. P. Matthews, T. Welton and P. A. Hunt, *Phys. Chem. Chem. Phys.*, 2014, **16**, 3238-3253.
59. S. Zahn, M. Brehm, M. Brüssel, O. Hollóczki, M. Kohagen, S. Lehmann, F. Malberg, A. S. Pensado, M. Schöppke, H. Weber and B. Kirchner, *J. Mol. Liq.*, 2014, **192**, 71-76.
60. S. Puttick, A. L. Davis, K. Butler, L. Lambert, J. El harfi, D. J. Irvine, A. K. Whittaker, K. J. Thurecht and P. Licence, *Chem. Sci.*, 2011, **2**, 1810-1816.
61. C. C. Weber, A. F. Masters and T. Maschmeyer, *Angew. Chem. Int. Ed.*, 2012, **51**, 11483-11486.
62. J. N. A. Canongia Lopes and A. A. H. Pádua, *J. Phys. Chem. B*, 2006, **110**, 3330-3335.
63. A. Triolo, O. Russina, B. Fazio, R. Triolo and E. Di Cola, *Chem. Phys. Lett.*, 2008, **457**, 362-365.
64. T. Pott and P. Meleard, *Phys. Chem. Chem. Phys.*, 2009, **11**, 5469-5475.
65. F. Dommert, J. Schmidt, C. Krekeler, Y. Y. Zhao, R. Berger, L. Delle Site and C. Holm, *J. Mol. Liq.*, 2010, **152**, 2-8.
66. J. Schmidt, C. Krekeler, F. Dommert, Y. Zhao, R. Berger, L. D. Site and C. Holm, *J. Phys. Chem. B*, 2010, **114**, 6150-6155.
67. W. T. Borden and E. R. Davidson, *Acc. Chem. Res.*, 1996, **29**, 67-75.
68. P. A. Hunt and I. R. Gould, *J. Phys. Chem. A*, 2006, **110**, 2269-2282.
69. E. Rezabal and T. Schäfer, *J. Phys. Chem. B*, 2012, **117**, 553-562.
70. B. L. Bhargava and S. Balasubramanian, *J. Phys. Chem. B*, 2007, **111**, 4477-4487.
71. M. Brussel, M. Brehm, T. Voigt and B. Kirchner, *Phys. Chem. Chem. Phys.*, 2011, **13**, 13617-13620.
72. K. Wendler, M. Brehm, F. Malberg, B. Kirchner and L. Delle Site, *J. Chem. Theory Comput.*, 2012, **8**, 1570-1579.
73. D. S. Firaha and B. Kirchner, *J. Chem. Eng. Data*, 2014, **59**, 3098-3104.
74. M. Brehm, H. Weber, A. S. Pensado, A. Stark and B. Kirchner, *Phys. Chem. Chem. Phys.*, 2012, **14**, 5030-5044.
75. M. G. Del Pópolo, R. M. Lynden-Bell and J. Kohanoff, *J. Phys. Chem. B*, 2005, **109**, 5895-5902.
76. Y. Zhang and E. J. Maginn, *J. Phys. Chem. B*, 2012, **116**, 10036-10048.
77. P. Koskinen and V. Mäkinen, *Comput. Mater. Sci.*, 2009, **47**, 237-253.
78. M. Elstner and G. Seifert, *Density functional tight binding*, 2014.

79. A. F. Oliveira, G. Seifert, T. Heine and H. A. Duarte, *J. Braz. Chem. Soc.*, 2009, **20**, 1193-1205.
80. M. A. Addicoat, R. Stefanovic, G. B. Webber, R. Atkin and A. J. Page, *J. Chem. Theory Comput.*, 2014, **10**, 4633-4643.
81. T. Carstens, R. Gustus, O. Höfft, N. Borisenko, F. Endres, H. Li, R. J. Wood, A. J. Page and R. Atkin, *J. Phys. Chem. C*, 2014, **118**, 10833-10843.
82. A. J. Page, A. Elbourne, R. Stefanovic, M. A. Addicoat, G. G. Warr, K. Voitchovsky and R. Atkin, *Nanoscale*, 2014, **6**, 8100-8106.
83. M. A. Addicoat, S. Fukuoka, A. J. Page and S. Irle, *J. Comput. Chem.*, 2013, **34**, 2591-2600.
84. E. J. Maginn, in *Reviews in Computational Chemistry*, John Wiley & Sons, Inc., 2009, pp. 421-493.
85. J. Canongia Lopes and A. H. Pádua, *Theor. Chem. Acc.*, 2012, **131**, 1129.
86. J. N. Canongia Lopes, J. Deschamps and A. A. H. Pádua, *J. Phys. Chem. B*, 2004, **108**, 11250-11250.
87. J. N. Canongia Lopes and A. A. H. Pádua, *J. Phys. Chem. B*, 2004, **108**, 16893-16898.
88. J. N. Canongia Lopes and A. A. H. Pádua, *J. Phys. Chem. B*, 2006, **110**, 19586-19592.
89. J. N. Canongia Lopes, A. A. H. Pádua and K. Shimizu, *J. Phys. Chem. B*, 2008, **112**, 5039-5046.
90. A. Warshel and M. Levitt, *J. Mol. Biol.*, 1976, **103**, 227-249.
91. M. J. Field, P. A. Bash and M. Karplus, *J. Comput. Chem.*, 1990, **11**, 700-733.
92. A. C. T. van Duin, S. Dasgupta, F. Lorant and W. A. Goddard, *J. Phys. Chem. A*, 2001, **105**, 9396-9409.
93. B. Zhang, A. C. T. van Duin and J. K. Johnson, *J. Phys. Chem. B*, 2014, **118**, 12008-12016.
94. Y. Wang, S. Feng and G. A. Voth, *J. Chem. Theory Comput.*, 2009, **5**, 1091-1098.
95. Y. Wang, W. Jiang, T. Yan and G. A. Voth, *Acc. Chem. Res.*, 2007, **40**, 1193-1199.
96. B. L. Bhargava, R. Devane, M. L. Klein and S. Balasubramanian, *Soft Matter*, 2007, **3**, 1395-1400.
97. J. Wohlert and L. A. Berglund, *J. Chem. Theory Comput.*, 2011, **7**, 753-760.
98. C. A. López, A. J. Rzepiela, A. H. de Vries, L. Dijkhuizen, P. H. Hünenberger and S. J. Marrink, *J. Chem. Theory Comput.*, 2009, **5**, 3195-3210.
99. S. Riniker, J. R. Allison and W. F. van Gunsteren, *Phys. Chem. Chem. Phys.*, 2012, **14**, 12423-12430.
100. P. A. Hunt, *J. Phys. Chem. B*, 2007, **111**, 4844-4853.
101. K. Fumino, T. Peppel, M. Geppert-Rybczynska, D. H. Zaitsau, J. K. Lehmann, S. P. Verevkin, M. Kockerling and R. Ludwig, *Phys. Chem. Chem. Phys.*, 2011, **13**, 14064-14075.
102. I. Skarmoutsos, D. Dellis, R. P. Matthews, T. Welton and P. A. Hunt, *J. Phys. Chem. B*, 2012, **116**, 4921-4933.
103. Y. Zhang and E. J. Maginn, *Phys. Chem. Chem. Phys.*, 2012, **14**, 12157-12164.
104. I. Skarmoutsos, T. Welton and P. A. Hunt, *Phys. Chem. Chem. Phys.*, 2014, **16**, 3675-3685.
105. K. Wendler, S. Zahn, F. Dommert, R. Berger, C. Holm, B. Kirchner and L. Delle Site, *J. Chem. Theory Comput.*, 2011, **7**, 3040-3044.
106. S. Kossmann, J. Thar, B. Kirchner, P. A. Hunt and T. Welton, *J. Chem. Phys.*, 2006, **124**, 174506.
107. M. Gonzalez-Miquel, M. Talreja, A. L. Ethier, K. Flack, J. R. Switzer, E. J. Biddinger, P. Pollet, J. Palomar, F. Rodriguez, C. A. Eckert and C. L. Liotta, *Ind. Eng. Chem. Res.*, 2012, **51**, 16066-16073.
108. J. Vökl, K. Müller, L. Mokrushina and W. Arlt, *Chem. Eng. Technol.*, 2012, **35**, 579-583.
109. O. Holloczki, F. Malberg, T. Welton and B. Kirchner, *Phys. Chem. Chem. Phys.*, 2014, **16**, 16880-16890.
110. V. S. Bernales, A. V. Marenich, R. Contreras, C. J. Cramer and D. G. Truhlar, *J. Phys. Chem. B*, 2012, **116**, 9122-9129.
111. X. Wei and D. Zhang, *Comp. Theor. Chem.*, 2013, **1014**, 24-28.
112. B. G. Janesko, *Phys. Chem. Chem. Phys.*, 2014, **16**, 5423-5433.
113. V. N. Emel'yanenko, S. P. Verevkin and A. Heintz, *J. Am. Chem. Soc.*, 2007, **129**, 3930-3937.
114. P. A. Hunt, B. Kirchner and T. Welton, *Chem. Eur. J.*, 2006, **12**, 6762-6775.
115. P. Nockemann, B. Thijs, K. Driesen, C. R. Janssen, K. Van Hecke, L. Van Meervelt, S. Kossmann, B. Kirchner and K. Binnemans, *J. Phys. Chem. B*, 2007, **111**, 5254-5263.
116. C. S. Pomelli, C. Chiappe, A. Vidis, G. Laurenczy and P. J. Dyson, *J. Phys. Chem. B*, 2007, **111**, 13014-13019.
117. Y. Wang, H. Li and S. Han, *J. Chem. Phys.*, 2005, **123**, 174501.
118. Y. Wang, H. Li and S. Han, *J. Phys. Chem. B*, 2006, **110**, 24646-24651.
119. C. Krekeler, J. Schmidt, Y. Y. Zhao, B. Qiao, R. Berger, C. Holm and L. Delle Site, *J. Chem. Phys.*, 2008, **129**, 174503.
120. B. L. Bhargava and S. Balasubramanian, *Chem. Phys. Lett.*, 2007, **444**, 242-246.
121. E. I. Izgorodina, U. L. Bernard and D. R. MacFarlane, *J. Phys. Chem. A*, 2009, **113**, 7064-7072.
122. E. I. Izgorodina, D. Golze, R. Maganti, V. Armel, M. Taige, T. J. S. Schubert and D. R. MacFarlane, *Phys. Chem. Chem. Phys.*, 2014, **16**, 7209-7221.
123. S. Grimme, W. Hujo and B. Kirchner, *Phys. Chem. Chem. Phys.*, 2012, **14**, 4875-4883.
124. S. Zahn, D. R. MacFarlane and E. I. Izgorodina, *Phys. Chem. Chem. Phys.*, 2013, **15**, 13664-13675.
125. B. G. Janesko, *Phys. Chem. Chem. Phys.*, 2011, **13**, 11393-11401.
126. O. Borodin and G. D. Smith, *J. Phys. Chem. B*, 2006, **110**, 11481-11490.
127. J. Rigby and E. I. Izgorodina, *Phys. Chem. Chem. Phys.*, 2013, **15**, 1632-1646.
128. Y. Dai, Y. Qu, S. Wang and J. Wang, *Chem. Phys. Lett.*, 2014, **608**, 366.
129. R. S. Mulliken, *J. Chem. Phys.*, 1955, **23**, 1833-1840.
130. P. O. Löwdin, *J. Chem. Phys.*, 1950, **18**, 365-375.
131. A. E. Reed, R. B. Weinstock and F. Weinhold, *J. Chem. Phys.*, 1985, **83**, 735-746.
132. F. Martin and H. Zipse, *J. Comput. Chem.*, 2005, **26**, 97-105.
133. M. Bühl, A. Chaumont, R. Schurhammer and G. Wipff, *J. Phys. Chem. B*, 2005, **109**, 18591-18599.

134. R. M. Lynden-Bell, *Phys. Chem. Chem. Phys.*, 2010, **12**, 1733-1740.
135. Z. Liu, T. Chen, A. Bell and B. Smit, *J. Phys. Chem. B*, 2010, **114**, 4572-4582.
136. K. Wendler, F. Dommert, Y. Y. Zhao, R. Berger, C. Holm and L. Delle Site, *Faraday Discuss.*, 2012, **154**, 111-132.
137. M. Kohagen, M. Brehm, J. Thar, W. Zhao, F. Müller-Plathe and B. Kirchner, *J. Phys. Chem. B*, 2010, **115**, 693-702.
138. E. Bodo, S. Mangialardo, F. Ramondo, F. Ceccacci and P. Postorino, *J. Phys. Chem. B*, 2012, **116**, 13878-13888.
139. F. Dommert, K. Wendler, B. Qiao, L. Delle Site and C. Holm, *J. Mol. Liq.*, 2014, **192**, 32-37.
140. A. Mondal and S. Balasubramanian, *J. Phys. Chem. B*, 2014, **118**, 3409-3422.
141. W. Beichel, N. Trapp, C. Hauf, O. Kohler, G. Eickerling, W. Scherer and I. Krossing, *Angew. Chem. Int. Ed.*, 2014, **126**, 3207-3210.
142. V. V. Chaban, I. V. Voroshylova and O. N. Kalugin, *Phys. Chem. Chem. Phys.*, 2011, **13**, 7910-7920.
143. M. Lísal, Z. Chval, J. Storch and P. Izák, *J. Mol. Liq.*, 2014, **189**, 85-94.
144. M. A. Carignano, *J. Phys. Chem. B*, 2013, **117**, 15176-15183.
145. F. Chen, L. Jin, S. W. de Leeuw, J. M. Pringle and M. Forsyth, *J. Chem. Phys.*, 2013, **138**, -.
146. H. Liu, Z. Zhang, J. E. Bara and C. H. Turner, *J. Phys. Chem. B*, 2013, **118**, 255-264.
147. C. M. Tenney, M. Massel, J. M. Mayes, M. Sen, J. F. Brennecke and E. J. Maginn, *J. Chem. Eng. Data*, 2014, **59**, 391-399.
148. D. T. Bowron, C. D'Agostino, L. F. Gladden, C. Hardacre, J. D. Holbrey, M. C. Lagunas, J. McGregor, M. D. Mantle, C. L. Mullan and T. G. A. Youngs, *J. Phys. Chem. B*, 2010, **114**, 7760-7768.
149. Z. L. Terranova and S. A. Corcelli, *J. Phys. Chem. B*, 2013, **117**, 15659-15666.
150. B. L. Bhargava and S. Balasubramanian, *J. Chem. Phys.*, 2007, **127**, 114510.
151. T. G. A. Youngs and C. Hardacre, *ChemPhysChem*, 2008, **9**, 1548-1558.
152. F. Dommert, J. Schmidt, B. Qiao, Y. Zhao, C. Krekeler, L. Delle Site, R. Berger and C. Holm, *J. Chem. Phys.*, 2008, **129**, 224501.
153. C. E. Resende Prado and L. C. Gomide Freitas, *Journal of Molecular Structure: THEOCHEM*, 2007, **847**, 93-100.
154. Z. Liu, X. Wu and W. Wang, *Phys. Chem. Chem. Phys.*, 2006, **8**, 1096-1104.
155. S. Zhang, N. Sun, X. He, X. Lu and X. Zhang, *J. Phys. Chem. Ref. Data*, 2006, **35**, 1475-1517.
156. H. Tokuda, K. Hayamizu, K. Ishii, M. A. B. H. Susan and M. Watanabe, *J. Phys. Chem. B*, 2004, **108**, 16593-16600.
157. H. Tokuda, S. Tsuzuki, M. A. B. H. Susan, K. Hayamizu and M. Watanabe, *J. Phys. Chem. B*, 2006, **110**, 19593-19600.
158. J. de Andrade, E. S. Böes and H. Stassen, *J. Phys. Chem. B*, 2002, **106**, 13344-13351.
159. Z. Liu, S. Huang and W. Wang, *J. Phys. Chem. B*, 2004, **108**, 12978-12989.
160. A. Noda, K. Hayamizu and M. Watanabe, *J. Phys. Chem. B*, 2001, **105**, 4603-4610.
161. A. Deyko, K. R. J. Lovelock, J.-A. Corfield, A. W. Taylor, P. N. Gooden, I. J. Villar-Garcia, P. Licence, R. G. Jones, V. G. Krasovskiy, E. A. Chernikova and L. M. Kustov, *Phys. Chem. Chem. Phys.*, 2009, **11**, 8544-8555.
162. S. V. Sambasivarao and O. Acevedo, *J. Chem. Theory Comput.*, 2009, **5**, 1038-1050.
163. W.-G. Xu, L. Li, X.-X. Ma, J. Wei, W.-B. Duan, W. Guan and J.-Z. Yang, *J. Chem. Eng. Data*, 2012, **57**, 2177-2184.
164. B. B. Hurisso, K. R. J. Lovelock and P. Licence, *Phys. Chem. Chem. Phys.*, 2011, **13**, 17737-17748.
165. S. Men, K. R. J. Lovelock and P. Licence, *Phys. Chem. Chem. Phys.*, 2011, **13**, 15244-15255.
166. T. Cremer, C. Kolbeck, K. R. J. Lovelock, N. Paape, R. Wölfel, P. S. Schulz, P. Wasserscheid, H. Weber, J. Thar, B. Kirchner, F. Maier and H.-P. Steinrück, *Chem. Eur. J.*, 2010, **16**, 9018-9033.
167. C. Schroder, *Phys. Chem. Chem. Phys.*, 2012, **14**, 3089-3102.
168. S. Gabl, C. Schröder and O. Steinhauser, *J. Chem. Phys.*, 2012, **137**, 94501.
169. A. D. MacKerell, D. Bashford, Bellott, R. L. Dunbrack, J. D. Evanseck, M. J. Field, S. Fischer, J. Gao, H. Guo, S. Ha, D. Joseph-McCarthy, L. Kuchnir, K. Kuczera, F. T. K. Lau, C. Mattos, S. Michnick, T. Ngo, D. T. Nguyen, B. Prodhom, W. E. Reiher, B. Roux, M. Schlenkrich, J. C. Smith, R. Stote, J. Straub, M. Watanabe, J. Wiórkiewicz-Kuczera, D. Yin and M. Karplus, *J. Phys. Chem. B*, 1998, **102**, 3586-3616.
170. W. D. Cornell, P. Cieplak, C. I. Bayly, I. R. Gould, K. M. Merz, D. M. Ferguson, D. C. Spellmeyer, T. Fox, J. W. Caldwell and P. A. Kollman, *J. Am. Chem. Soc.*, 1995, **117**, 5179-5197.
171. G. Kaminski and W. L. Jorgensen, *J. Phys. Chem.*, 1996, **100**, 18010-18013.
172. W. L. Jorgensen, D. S. Maxwell and J. Tirado-Rives, *J. Am. Chem. Soc.*, 1996, **118**, 11225-11236.
173. J. N. Canongia Lopes, J. Deschamps and A. A. H. Pádua, *J. Phys. Chem. B*, 2004, **108**, 2038-2047.
174. A. Bagno, F. D'Amico and G. Saielli, *J. Mol. Liq.*, 2007, **131-132**, 17-23.
175. T. Yan, C. J. Burnham, M. G. Del Pópolo and G. A. Voth, *J. Phys. Chem. B*, 2004, **108**, 11877-11881.
176. C. E. R. Prado, M. G. D. Pópolo, T. G. A. Youngs, J. Kohanoff and R. M. Lynden-Bell, *Mol. Phys.*, 2006, **104**, 2477-2483.
177. R. M. Lynden-Bell, M. G. Del Pópolo, T. G. A. Youngs, J. Kohanoff, C. G. Hanke, J. B. Harper and C. C. Pinilla, *Acc. Chem. Res.*, 2007, **40**, 1138-1145.
178. F. Dommert and C. Holm, *Phys. Chem. Chem. Phys.*, 2013, **15**, 2037-2049.
179. X. Liu, G. Zhou, S. Zhang, G. Wu and G. Yu, *J. Phys. Chem. B*, 2007, **111**, 5658-5668.
180. T. Köddermann, D. Paschek and R. Ludwig, *ChemPhysChem*, 2007, **8**, 2464-2470.
181. T. Yan, Y. Wang and C. Knox, *J. Phys. Chem. B*, 2010, **114**, 6905-6921.
182. T. Yan, Y. Wang and C. Knox, *J. Phys. Chem. B*, 2010, **114**, 6886-6904.
183. B. T. Thole, *Chem. Phys.*, 1981, **59**, 341-350.
184. F. Wang and K. D. Jordan, *J. Chem. Phys.*, 2001, **114**, 10717-10724.

185. T.-M. Chang and L. X. Dang, *J. Phys. Chem. A*, 2009, **113**, 2127-2135.
186. T. Taylor, M. Schmollngruber, C. Schröder and O. Steinhauser, *J. Chem. Phys.*, 2013, **138**, -.
187. Y. Gu and T. Yan, *J. Phys. Chem. A*, 2012, **117**, 219-227.
188. O. Borodin and G. D. Smith, *J. Phys. Chem. B*, 2006, **110**, 6279-6292.
189. O. Borodin, *J. Phys. Chem. B*, 2009, **113**, 11463-11478.
190. D. Bedrov, O. Borodin, Z. Li and G. D. Smith, *J. Phys. Chem. B*, 2010, **114**, 4984-4997.
191. J. B. Hooper, O. N. Starovoytov, O. Borodin, D. Bedrov and G. D. Smith, *J. Chem. Phys.*, 2012, **136**, 194506/194501-194506/194507.
192. C. Schröder and O. Steinhauser, *J. Chem. Phys.*, 2010, **133**, 154511.
193. C. Schröder, *J. Chem. Phys.*, 2011, **135**, -.
194. C. Schroder, T. Sonnleitner, R. Buchner and O. Steinhauser, *Phys. Chem. Chem. Phys.*, 2011, **13**, 12240-12248.
195. Y. Wu, Y. Li, N. Hu and M. Hong, *Phys. Chem. Chem. Phys.*, 2014, **16**, 2674-2685.
196. A. d. O. Cavalcante, M. C. C. Ribeiro and M. S. Skaf, *J. Chem. Phys.*, 2014, **140**, -.
197. K. Shimizu, D. Almantariotis, M. F. C. Gomes, A. A. H. Padua and J. N. Canongia Lopes, *J. Phys. Chem. B*, 2010, **114**, 3592-3600.
198. X. Wu, Z. Liu, S. Huang and W. Wang, *Phys. Chem. Chem. Phys.*, 2005, **7**, 2771-2779.
199. C. Cadena and E. J. Maginn, *J. Phys. Chem. B*, 2006, **110**, 18026-18039.
200. C. Cadena, Q. Zhao, R. Q. Snurr and E. J. Maginn, *J. Phys. Chem. B*, 2006, **110**, 2821-2832.
201. T. I. Morrow and E. J. Maginn, *J. Phys. Chem. B*, 2002, **106**, 12807-12813.
202. M. S. Kelkar, W. Shi and E. J. Maginn, *Ind. Eng. Chem. Res.*, 2008, **47**, 9115-9126.
203. K. E. Gutowski, B. Gurkan and E. J. Maginn, *Pure Appl. Chem.*, 2009, **81**, 1799-1828.
204. T. Köddermann, D. Reith and R. Ludwig, *ChemPhysChem*, 2013, **14**, 3368-3374.
205. N. M. Micaelo, A. M. Baptista and C. M. Soares, *J. Phys. Chem. B*, 2006, **110**, 14444-14451.
206. J. de Andrade, E. S. Boees and H. Stassen, *J. Phys. Chem. B*, 2002, **106**, 3546-3548.
207. X. Liu, S. Zhang, G. Zhou, G. Wu, X. Yuan and X. Yao, *J. Phys. Chem. B*, 2006, **110**, 12062-12071.
208. X. Liu, G. Zhou and S. Zhang, *Fluid Phase Equilib.*, 2008, **272**, 1-7.
209. X. M. Liu, X. C. Zhang, G. H. Zhou, X. Q. Yao and S. J. Zhang, *Sci. China: Chem.*, 2012, **55**, 1573-1579.
210. G. Zhou, X. Liu, S. Zhang, G. Yu and H. He, *J. Phys. Chem. B*, 2007, **111**, 7078-7084.
211. X. Liu, Y. Zhao, X. Zhang, G. Zhou and S. Zhang, *J. Phys. Chem. B*, 2012, **116**, 4934-4942.
212. Y. Zhao, X. Liu, J. Wang and S. Zhang, *ChemPhysChem*, 2012, **13**, 3126-3133, S3126/3121-S3126/3125.
213. T. G. A. Youngs, M. G. Del Popolo and J. Kohanoff, *J. Phys. Chem. B*, 2006, **110**, 5697-5707.
214. Y. Wang, H. Pan, H. Li and C. Wang, *J. Phys. Chem. B*, 2007, **111**, 10461-10467.
215. M. G. Del Pópolo and G. A. Voth, *J. Phys. Chem. B*, 2004, **108**, 1744-1752.
216. Y. Hu, H. Jiang, X. Zhang, Z. Liu, J. Ren and R. Liu, *J. Mol. Struct.: THEOCHEM*, 2009, **915**, 132-140.
217. H. Liu, K. L. Sale, B. M. Holmes, B. A. Simmons and S. Singh, *J. Phys. Chem. B*, 2010, **114**, 4293-4301.
218. F. Yan, M. Lartey, K. Damodaran, E. Albenze, R. L. Thompson, J. Kim, M. Haranczyk, H. B. Nulwala, D. R. Luebke and B. Smit, *Phys. Chem. Chem. Phys.*, 2013, **15**, 3264-3272.
219. Y.-L. Wang, F. U. Shah, S. Glavatskih, O. N. Antzutkin and A. Laaksonen, *J. Phys. Chem. B*, 2014, **118**, 8711-8723.
220. I. V. Voroshylova and V. V. Chaban, *J Phys Chem B*, 2014, **118**, 10716-10724.
221. S. U. Lee, J. Jung and Y.-K. Han, *Chem. Phys. Lett.*, 2005, **406**, 332-340.
222. M. Lísal, *J. Chem. Phys.*, 2013, **139**, 214701.
223. T. Koller, J. Ramos, N. M. Garrido, A. P. Froba and I. G. Economou, *Mol. Phys.*, 2012, **110**, 1115-1126.
224. Z.-P. Liu, T. Chen, A. Bell and B. Smit, *J. Phys. Chem. B*, 2010, **114**, 4572-4582.
225. X. Zhong, Z. Liu and D. Cao, *J. Phys. Chem. B*, 2011, **115**, 10027-10040.
226. E. E. Fileti and V. V. Chaban, *Chem. Phys. Lett.*, 2014, **616-617**, 205-211.
227. Y. Wu, N. Hu, L. Yue, L. Wei and W. Guan, *J. Chem. Phys.*, 2015, **142**, 064503.
228. M. Klahn, A. Seduraman and P. Wu, *J. Phys. Chem. B*, 2008, **112**, 10989-11004.
229. O. N. Starovoytov, H. Torabifard and G. A. Cisneros, *J. Phys. Chem. B*, 2014, **118**, 7156-7166.
230. Z. Hu, X. Huang, H. V. R. Annapureddy and C. J. Margulis, *J. Phys. Chem. B*, 2008, **112**, 7837-7849.
231. E. Choi, J. G. McDaniel, J. R. Schmidt and A. Yethiraj, *J. Phys. Chem. Lett.*, 2014, **5**, 2670-2674.
232. J. G. McDaniel and J. R. Schmidt, *J. Phys. Chem. A*, 2013, **117**, 2053-2066.
233. J. B. Harper and R. M. Lynden-Bell, *Mol. Phys.*, 2004, **102**, 85-94.
234. L. A. Blanchard and J. F. Brennecke, *Ind. Eng. Chem. Res.*, 2000, **40**, 287-292.
235. C. G. Hanke, A. Johansson, J. B. Harper and R. M. Lynden-Bell, *Chem. Phys. Lett.*, 2003, **374**, 85-90.
236. K. Shimizu, M. F. Costa Gomes, A. A. H. Pádua, L. P. N. Rebelo and J. N. Canongia Lopes, *J. Phys. Chem. B*, 2009, **113**, 9894-9900.
237. M. B. Shiflett and A. Yokozeki, *J. Chem. Eng. Data*, 2008, **53**, 2683-2691.
238. J. Lachwa, J. Szydłowski, A. Makowska, K. R. Seddon, J. M. S. S. Esperanca, H. J. R. Guedes and L. P. N. Rebelo, *Green Chem.*, 2006, **8**, 262-267.
239. D. J. Cram and M. Goldstein, *J. Am. Chem. Soc.*, 1963, **85**, 1063-1074.
240. X. Zhang, X. Zhang, H. Dong, Z. Zhao, S. Zhang and Y. Huang, *Energy Environ. Sci.*, 2012, **5**, 6668-6681.
241. Z. Lei, C. Dai and B. Chen, *Chem. Rev.*, 2013, **114**, 1289-1326.

242. A. Maiti, *Int. J. Quantum Chem.*, 2014, **114**, 163-175.
243. E. Torralba-Calleja, J. Skinner and D. Gutierrez-Tauste, *Journal of Chemistry*, 2013, **2013**, 16.
244. H. Liu, S. Dai and D.-e. Jiang, *J. Phys. Chem. B*, 2014, **118**, 2719-2725.
245. M. Mohammadi and M. Foroutan, *J. Mol. Liq.*, 2014, **193**, 60-68.
246. S. N. V. K. Aki, B. R. Mellein, E. M. Saurer and J. F. Brennecke, *J. Phys. Chem. B*, 2004, **108**, 20355-20365.
247. J. Xu, S. Wang, W. Yu, Q. Xu, W. Wang and J. Yin, *Chin. J. Chem. Eng.*, 2014, **22**, 153-163.
248. C. Cadena, J. L. Anthony, J. K. Shah, T. I. Morrow, J. F. Brennecke and E. J. Maginn, *J. Am. Chem. Soc.*, 2004, **126**, 5300-5308.
249. T. C. Lourenço, M. F. C. Coelho, T. C. Ramalho, D. van der Spoel and L. T. Costa, *Environ. Sci. Technol.*, 2013, **47**, 7421-7429.
250. A. Brandt, J. Grasvik, J. P. Hallett and T. Welton, *Green Chem.*, 2013, **15**, 550-583.
251. D. A. Fort, R. C. Remsing, R. P. Swatloski, P. Moyna, G. Moyna and R. D. Rogers, *Green Chem.*, 2007, **9**, 63-69.
252. L. Kyllonen, A. Parviainen, S. Deb, M. Lawoko, M. Gorlov, I. Kilpelainen and A. W. T. King, *Green Chem.*, 2013, **15**, 2374-2378.
253. L. J. A. Conceicao, E. Bogel-Lukasik and R. Bogel-Lukasik, *RSC Adv.*, 2012, **2**, 1846-1855.
254. M. M. Hossain and L. Aldous, *Aust. J. Chem.*, 2012, **65**, 1465-1477.
255. B. D. Rabideau, A. Agarwal and A. E. Ismail, *J. Phys. Chem. B*, 2014, **118**, 1621-1629.
256. N. Sun, R. Parthasarathi, A. M. Socha, J. Shi, S. Zhang, V. Stavila, K. L. Sale, B. A. Simmons and S. Singh, *Green Chem.*, 2014, **16**, 2546-2557.
257. W. E. S. Hart, J. B. Harper and L. Aldous, *Green Chem.*, 2015, **17**, 214.
258. N. P. Novoselov, E. S. Sashina, V. E. Petrenko and M. Zaborsky, *Fibre Chem.*, 2007, **39**, 153-158.
259. Z. Jarin and J. Pfaendtner, *J. Chem. Theory Comput.*, 2014, **10**, 507-510.
260. J. Zhang, H. Zhang, J. Wu, J. Zhang, J. He and J. Xiang, *Phys. Chem. Chem. Phys.*, 2010, **12**, 1941-1947.
261. X. Creary, E. D. Willis and M. Gagnon, *J. Am. Chem. Soc.*, 2005, **127**, 18114-18120.
262. L.-y. Liu, B. Wang, H.-m. Yang, W.-x. Chang and J. Li, *Tetrahedron Lett.*, 2011, **52**, 5636-5639.
263. K. K. Laali and V. J. Gettwert, *J. Fluorine Chem.*, 2001, **107**, 31-34.
264. K. K. Laali, T. Okazaki and S. D. Bunge, *J. Org. Chem.*, 2007, **72**, 6758-6762.
265. R. Bini, C. Chiappe, E. Marmugi and D. Pieraccini, *Chem. Commun.*, 2006, 897-899.
266. R. Bini, O. Bortolini, C. Chiappe, D. Pieraccini and T. Siciliano, *J. Phys. Chem. B*, 2007, **111**, 598-604.
267. R. Bini, C. Chiappe, C. S. Pomelli and B. Parisi, *J. Org. Chem.*, 2009, **74**, 8522-8530.
268. E. I. Izgorodina, R. Maganti, V. Armel, P. M. Dean, J. M. Pringle, K. R. Seddon and D. R. MacFarlane, *J. Phys. Chem. B*, 2011, **115**, 14688-14697.
269. O. Acevedo, W. L. Jorgensen and J. D. Evanseck, *J. Chem. Theory Comput.*, 2007, **3**, 132-138.
270. S. T. Keaveney, R. S. Haines and J. B. Harper, *Org. Biomol. Chem.*, 2015, DOI: 10.1039/C1034OB02482F.
271. Y. Yoshida, O. Baba, C. Larriba and G. Saito, *J. Phys. Chem. B*, 2007, **111**, 12204-12210.
272. Z. Ren, A. S. Ivanova, D. Couchot-Vore and S. Garrett-Roe, *J. Phys. Chem. Lett.*, 2014, **5**, 1541-1546.
273. N. L. Lancaster, T. Welton and G. B. Young, *J. Chem. Soc., Perkin Trans. 2*, 2001, 2267-2270.
274. N. L. Lancaster and T. Welton, *J. Org. Chem.*, 2004, **69**, 5986-5992.
275. L. Crowhurst, R. Falcone, N. L. Lancaster, V. Llopis-Mestre and T. Welton, *J. Org. Chem.*, 2006, **71**, 8847-8853.
276. G. M. Arantes and M. C. C. Ribeiro, *J. Chem. Phys.*, 2008, **128**, 114503.
277. S. Alunni, A. Pero and G. Reichenbach, *J. Chem. Soc., Perkin Trans. 2*, 1998, 1747-1750.
278. S. Coleman, R. Byrne, S. Minkovska and D. Diamond, *Phys. Chem. Chem. Phys.*, 2009, **11**, 5608-5614.
279. S. Zhang, Q. Zhang, B. Ye, X. Li, X. Zhang and Y. Deng, *J. Phys. Chem. B*, 2009, **113**, 6012-6019.
280. R. Byrne, K. J. Fraser, E. Izgorodina, D. R. MacFarlane, M. Forsyth and D. Diamond, *Phys. Chem. Chem. Phys.*, 2008, **10**, 5919-5924.
281. R. Byrne, S. Coleman, K. J. Fraser, A. Raduta, D. R. MacFarlane and D. Diamond, *Phys. Chem. Chem. Phys.*, 2009, **11**, 7286-7291.
282. C. Allen, B. W. McCann and O. Acevedo, *J. Phys. Chem. B*, 2014.
283. D. A. Jaeger and C. E. Tucker, *Tetrahedron Lett.*, 1989, **30**, 1785-1788.
284. S. Tiwari and A. Kumar, *Angew. Chem. Int. Ed.*, 2006, **45**, 4824-4825.
285. C. Chiappe, M. Malvaldi and C. S. Pomelli, *Green Chem.*, 2010, **12**, 1330-1339.
286. R. Bini, C. Chiappe, V. L. Mestre, C. S. Pomelli and T. Welton, *Org. Biomol. Chem.*, 2008, **6**, 2522-2529.
287. S. Tiwari, N. Khupse and A. Kumar, *J. Org. Chem.*, 2008, **73**, 9075-9083.
288. R. Bini, C. Chiappe, V. Mestre, C. Pomelli and T. Welton, *Theor. Chem. Acc.*, 2009, **123**, 347-352.
289. H. Sun, D. Zhang, C. Ma and C. Liu, *Int. J. Quantum Chem.*, 2007, **107**, 1875-1885.
290. A. Kovalenko and F. Hirata, *Chem. Phys. Lett.*, 1998, **290**, 237-244.
291. A. Kovalenko and F. Hirata, *J. Chem. Phys.*, 1999, **110**, 10095-10112.
292. C. Chiappe, M. Malvaldi and C. S. Pomelli, *J. Chem. Theory Comput.*, 2010, **6**, 179-183.