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Electroactive phase dependent triboelectric nanogenerator performance of PVDF–TiO₂ composites†

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This investigation explores the impact of the electroactive phase of a well-known tribonegative polymer, polyvinylidene fluoride (PVDF), on its triboelectric behaviour by compositing it with anatase, rutile, and mixed-phase TiO₂ nanoparticles. PVDF–TiO₂ polymer composite films with TiO₂ having different crystalline phases were prepared by spin coating. TENG specimens were fabricated using the prepared films and tested for their TENG properties in contact separation mode by pairing them with ITO-coated PET substrates. The XRD and FT-IR results show that the TiO₂ nanoparticles with rutile phase imparted the highest percentage of β crystalline phase in PVDF compared to that of the anatase and mixed phase. The difference in surface roughness of PVDF–TiO₂ composites was also observed with the change in the crystalline phase of the incorporated TiO₂ nanoparticles in the polymer matrix. The TENG studies suggest that the PVDF incorporated with rutile TiO₂ shows the highest output voltage (peak–peak ~ 105 V at 60 N force) compared to all the other PVDF–TiO₂ composites at specified contact forces and frequencies, whereas PVDF incorporated with anatase TiO₂ and a mix of anatase and rutile TiO₂ showed peak–peak voltages of ~ 82 V and ~ 33 V respectively. These results offer insights into the crystalline phase-dependent triboelectric behaviour of polymers and the enhancement of their TENG performance through the tuning of polymer crystalline phases using fillers.

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1. Introduction

The advancements in renewable energy technologies and the proliferation of sensor networks during recent years have laid the foundation for the steady progress of the internet of things (IoT) and for achieving the sustainable development of mankind. Consequently, the number of so-called small and smart electronic devices such as sensors, wireless transmitters, actuators *etc.* installed across the globe and even wearable smart

electronics has increased tremendously in the order of billions and trillions.^{1–4} However, the limited life spans and environmental concerns associated with the energy-storing units for powering these electronics are still far from being addressed and indicate the need for self-powered electronics. Since their discovery in 2012⁵ triboelectric nanogenerators (TENGs) have been an emerging area of research focused on self-powered electronics for diverse applications. TENGs can harvest the otherwise wasted energy from various forms of mechanical motion, such as human body movements and large-scale wind and ocean waves.³ The mechanism of working of TENGs is based on a combination of contact electrification and electromagnetic induction. TENGs have shown distinctiveness compared to other energy harvesting systems owing to their straightforward fabrication methods,^{6–8} high power density^{9,10} and feasibility of hybridization with other energy technologies^{11–13} *etc.* Recent research has validated TENG's capabilities as a reliable technology for next-generation electronics by demonstrating TENG-based electronic skin, flexible and touch-screen displays, electronic watches, biomechanical monitoring sensors, *etc.*^{4,14,15}

Current research is more focused on new designs and strategies for generating higher power output from TENGs

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composites. Furthermore, the composite films showed the difference in percentage crystallinity in PVDF by incorporating TiO_2 with different crystalline phases. The SEM images and 3D profilometry data of the composite films showed the difference in surface topography of the films mainly focusing on the surface roughness. TENG specimens were fabricated from these films by attaching Cu tape as back electrodes. The TENG studies of the fabricated specimens were carried out by contacting them with a commercially purchased ITO-coated PET substrate at different forces and frequencies. The results indicated that the crystalline phase of TiO_2 has a significant influence on modulating the phase characteristics of PVDF. This has a profound role in enhancing the triboelectric properties of the TENG. This work is expected to deliver direction to tune up the TENG performance of polymers based on the crystalline phase of the fillers incorporated in them.

2.1. Chemicals and reagents

2.2. Synthesis of TiO₂ nanoparticles

2.3. Preparation of PVDF-TiO₂ polymer composite films by spin coating

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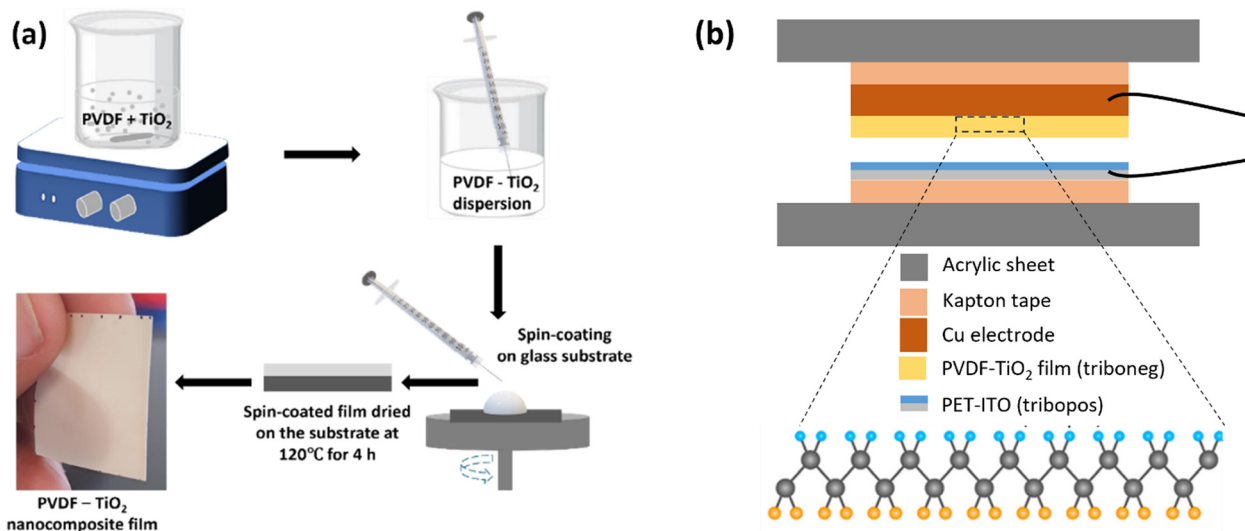


Fig. 1 Schematic representation of (a) the preparation of PVDF-TiO₂ nanocomposite films by spin coating and (b) the design of the fabricated TENG device.

TiO₂-PVDF films with TiO₂ nanoparticles calcined at different temperatures were prepared by first dispersing a required concentration of TiO₂ in 1:1 dimethyl formamide (DMF) and acetone solution. The particles were allowed to disperse uniformly for 24 h. Then PVDF (16.5 wt%) was added to this solution slowly over 30 minutes under constant stirring at 60 °C to avoid viscosity buildup and ensure homogeneity during the polymer dissolution in the solvent. The solution was stirred for 5 h to obtain a homogeneous polymer filler solution. After that, the solution was coated on a glass substrate by spin coating. For this ~1 mL of the homogeneous solution was added onto a 5 cm² clean and flat glass plate placed on the spinning chuck of the spin coater. All the film samples were prepared by coating at a spin speed of 1000 rpm for 35 s at room temperature (RT). The solution was uniformly coated over the glass plate after spinning. The films were then placed in a hot air oven at 120 °C for 4 h for complete drying. Freestanding films with approximately uniform thickness and smooth surfaces were obtained and stored for other characterizations and testing. PVDF film with no TiO₂ was also prepared using the same method. The PVDF film with no TiO₂ was designated as sPVDF and films containing TiO₂ were designated as PVT500, PVT600, PVT700, PVT800 and PVT900 based on the type of TiO₂ added. The solution preparation and spin coating parameters of the films were optimized through several trial-and-error experiments keeping the content of PVDF constant at 16.5 wt%. However, the viscosity build-up observed when adding 16.5 wt% of PVDF to the TiO₂ dispersed solution with TiO₂ content higher than 3 wt% indicates an important practical limitation in selecting the optimum TiO₂ percentage to get the best TENG output.

2.4. Fabrication of specimens for TENG testing

Spin-coated PVDF and PVDF-TiO₂ films (tribonegative layers) were cut into pieces with dimensions of 2.5 cm × 2.5 cm. Care was taken to eliminate any raised edges in the spin-coated films

to avoid the possibility of restricting the contact between the surface of materials during TENG testing. Conductive copper tape was attached as the back electrode to one side of all the films. The electrodes were then insulated using Kapton tape. Thin conductive copper lead wires were drawn out from the copper electrode to capture the output signals from the TENG during testing. ITO-coated polyethylene terephthalate (PET) (tribopositive layer) was used as the second contact layer. A schematic representation of the TENG is shown in Fig. 1(b).

2.5. Material characterization

An X-ray diffractometer (Siemens D500) was used to analyse the crystalline structure and phase of TiO₂ nanoparticles and PVDF-TiO₂ nanocomposite films. The diffractometer utilized a Cu Kα radiation ($\lambda = 0.15418$ nm) at 40 kV and 30 mA for recording the XRD patterns at the 2θ range 10°–80° with a scan rate of 0.2° s⁻¹. Fourier transform infrared spectroscopy (FT-IR) analysis of the polymer composite samples was performed to evaluate the electroactive phase content of the polymer films. A PerkinElmer spectrometer (Spectrum 100) was utilized for this, and the spectra were obtained in the region 1500–600 cm⁻¹. X-Ray photoelectron spectroscopy (XPS) was employed to analyse the elemental composition and bonding characteristics of TiO₂ nanoparticles. XPS, Thermo Fisher Scientific, (East Grinstead, UK) K-Alpha+ spectrometer was used for obtaining the XPS data. The surface topography of the films was analysed from SEM scans obtained using the instrument FEI Nova NanoSem 630 and a 3D optical profilometer (Alicona InfiniteFocus).

2.6. TENG testing rig and electrical measurements

The TENG studies of the samples were carried out in vertical contact-separation mode. A high-precision linear electrodynamic testing machine (Electropuls E3000, Instron UK) was adapted as the test force impactor. At the actuation drive, a load cell having a capacity of 5 kN (tension and compression) was



equipped. A dedicated holder with a levelled base was utilised to achieve a parallel contact between the tribolayers. The bottom friction layer was fitted onto a 360° movable platform to ensure proper alignment for maximum contact between the test surfaces. The bottom tribo-layer position was locked to avoid any further movement while testing. Normal force of 20 N, 40 N and 60 N was applied at frequencies varying from 4 Hz to 8 Hz to obtain the electrical output from the different TENG devices. A separation distance of ~2 mm between the active layers was used during the measurements. The voltage signals from the TENG were recorded using a digital oscilloscope (RS PRO, Bench-top Mixed Signal Oscilloscope, UK). Current output was retrieved using a low-noise current amplifier (SR570, Stanford Research Systems, USA) and the signals were recorded on a digital oscilloscope. The output power from the TENG was measured under external load resistances ranging from 1 KΩ to 5 GΩ. All TENG testing was performed at ambient environment conditions.

3. Results and discussion

Synthesized TiO₂ nanoparticles were calcined separately at a range of temperatures from 400 °C to 1000 °C to obtain a gradual transition from anatase to rutile phase in TiO₂ with the change in calcination temperature. Utmost care was taken to avoid any chances of contamination in the TiO₂ synthesis by carrying out the reaction in a laminar flow hood. Moreover, the calcination of the TiO₂ powders carried out at higher temperatures (400 °C to 1000 °C) has also removed any impurity that would have been present in the sample. This was further ensured by analysing the XRD patterns of TiO₂ nanoparticles and confirming that no peaks other than TiO₂ peaks were present in the sample. Consequently, the impact of TiO₂ nanoparticles containing a single crystalline phase (anatase or rutile) and mixed-phase on altering the electroactive (EA) phase of the PVDF matrix in PVDF-TiO₂ composites can be evaluated and correlated with the TENG output of these samples. The preparation of PVDF-TiO₂ polymer solution was found to be smooth and convenient when the TiO₂ nanoparticles were dispersed fully in the solvent first and then added the polymer for dissolution. This is because, owing to the viscosity build-up in the polymer solution, a uniform dispersion of TiO₂ was difficult to achieve leading to particle agglomeration. PVDF was dissolved in the solvent before the addition of TiO₂ in the initial trials of sample preparation. This will, in turn, hinder the effective interaction between the surface of TiO₂ particles and the single PVDF chains. The percentage loading of TiO₂ chosen for this study is not the optimized maximum loading for achieving the maximum TENG output.²⁰ There were practical limitations to increasing the loading of TiO₂ and PVDF above the amount used in this study as it led to viscosity increases in the PVDF-TiO₂ solution where the solution was unable to spin coat. By maintaining a constant speed and time of spin coating for all the prepared PVDF-TiO₂ samples, a consistent thickness of approximately

48 μm (Fig. S1 (ESI[†])) was obtained for all the samples. The surface of the spin-coated film which was in contact with the glass base (coating substrate) appeared glossy, and the opposite exposed surface had a comparatively rough finish as expected. For fabricating the specimens for TENG testing, the surface with a rough finish (exposed surface) was considered the surface of interest for all the films. Therefore, the glossy surface was attached to the back electrodes. Considering the reliability of comparison between samples, it is important to use the same surface as the test surface for all the samples. This is crucial because TENG output is susceptible to the amount of real contact area developed between the materials and it varies considerably based on the nature of the materials' surface.^{1,27} It is also worth mentioning that, as the comparison is between different TiO₂ crystalline phases, the percentage of TiO₂ in PVDF was kept constant for all the samples as the TiO₂ particles calcined at a range of temperatures were used in this study.

3.1. Material properties and crystalline phase analysis of TiO₂ nanoparticles and PVDF-TiO₂ films

Material characterization of the synthesized TiO₂ nanoparticles was carried out using X-ray photoelectron spectroscopic (XPS) and X-ray diffraction (XRD) analyses (Fig. 2). The XPS spectra of the TiO₂ nanoparticles were analysed to understand their elemental composition. Fig. 2(a) shows the survey spectra of the TiO₂ 500, TiO₂ 600, and TiO₂ 900 samples. Peaks corresponding to titanium (Ti), oxygen (O) and carbon (C) were observed in the samples' survey spectrum. C 1s peaks were observed at binding energies of 285 eV in all the samples, which arise from contamination during synthesis and calcination of the nanoparticles. The two strong peaks from TiO₂ nanoparticles at around 465.2 eV and 459.5 eV (Fig. 2(b)) can be attributed to Ti 2p_{1/2} and Ti 2p_{3/2}, respectively. These binding energies indicate that Ti is in a +4 oxidation state. Considering Ti's elemental composition and +4 oxidation state, it is reasonable to attribute the Ti-O bonding to TiO₂. Fig. 2(c) presents the XPS spectra of the O 1s of the TiO₂ samples. The peak positions of O 1s are located at 530.2 eV, which is in agreement with the reported O 1s electron binding energy for TiO₂.²⁶ The peak positions of Ti 2p which corresponds to the binding energy of Ti metal in TiO₂, O 1s which corresponds to the binding energy of oxygen in TiO₂ and a peak separation of 5.8 eV of the Ti 2p doublet agree well with the binding energy reported for TiO₂ nanoparticles.^{26,27} Fig. 2(d) shows the percentage of elements present in each TiO₂ sample. XPS data identifies the presence of Ti-O bonding in the sample which confirms the formation of TiO₂ in all the particles and matches with the reported peak positions of Ti and O of TiO₂.²⁸ It is also worth mentioning that C, F and N are also observed (in Fig. 2(d)) in the samples due to contamination during the sample synthesis.

The identification of the crystalline phase of TiO₂ nanoparticles is crucial in the context of this study to assess the TiO₂ induced phase transformation in PVDF matrix and their effect on TENG output. XRD patterns of TiO₂ nanoparticles were analysed to investigate the crystallinity, crystal phases present and the change in crystallite size of the nanoparticles with an



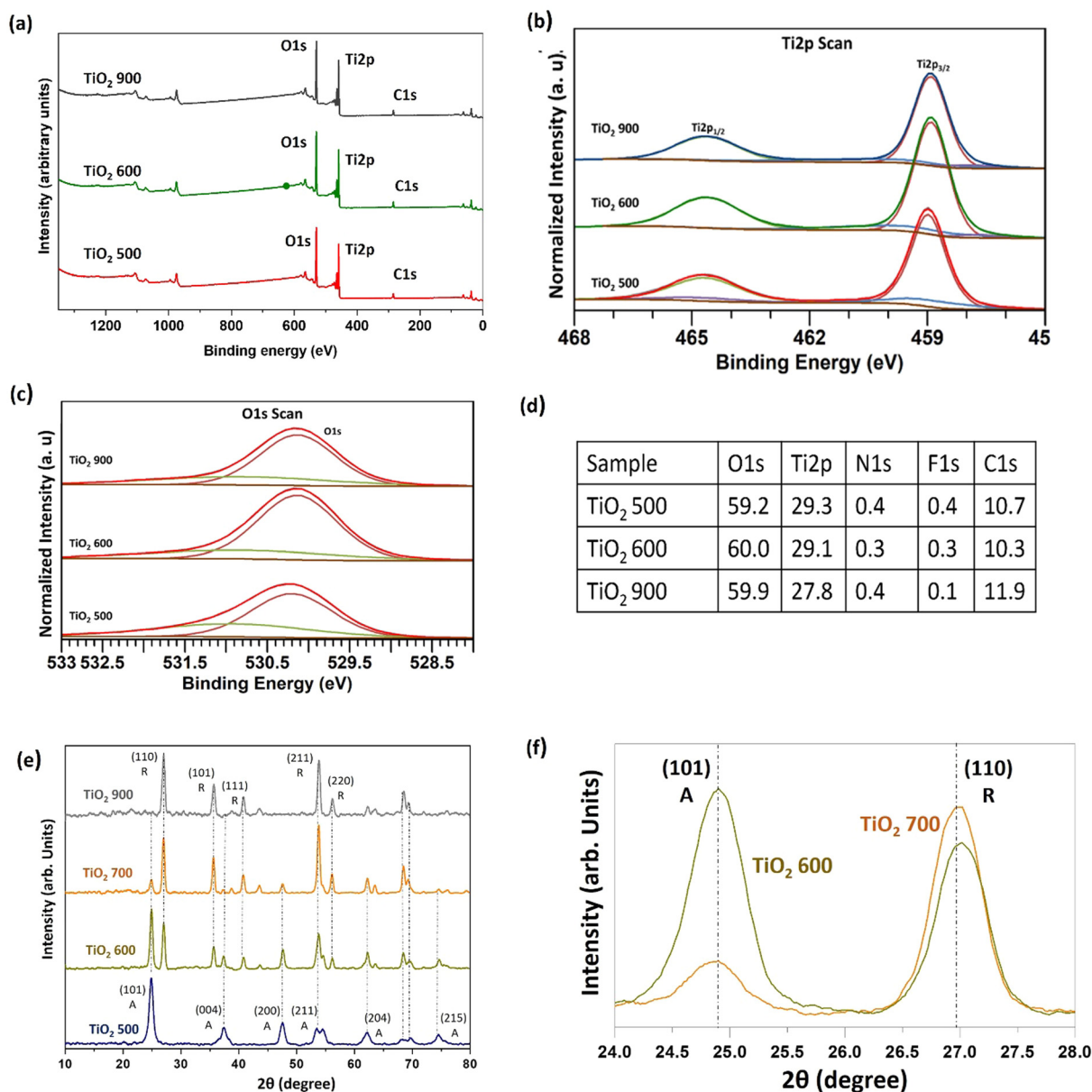


Fig. 2 XPS and XRD analysis of TiO₂ nanoparticles. (a) XPS survey spectra, (b) and (c) spectra of Ti 2p and O 1s BE region respectively and (d) tabulation of the composition of elements for TiO₂ 500, TiO₂ 600 and TiO₂ 900. XRD patterns of TiO₂ showing (e) comparison of the patterns of TiO₂ 500, TiO₂ 600, TiO₂ 700 and TiO₂ 900 and (f) changes in the anatase (101) and rutile (110) peak intensities with change in calcination temperature from 600 °C to 700 °C.

increase in calcination temperature. The XRD patterns of synthesized TiO₂ nanoparticles calcined at different temperatures varying from 500–900 °C are shown in Fig. 2(e) and (f). Characteristic reflection peaks of the anatase phase of TiO₂ (JCPDS card no. 21-1272) were observed in TiO₂ 500, TiO₂ 600 and TiO₂ 700 samples at $2\theta \sim 24.8^\circ$, 37.4° , 47.5° , 53.5° , 62.1° , 68.4° , 69.6° and 74.6° corresponding to the Bragg's reflections from (101), (004), (200), (211), (204), (116), (220) and (215) crystal planes respectively.^{27,29} Characteristic reflection peaks of the rutile phase of TiO₂ (JCPDS card no. 21-1276) were observed in TiO₂ 700 and TiO₂ 900 samples at $2\theta \sim 27.1^\circ$, 35.6° , 40.8° , 53.9° and 56.2° corresponding to the hkl planes (110), (101), (111), (211) and (220) respectively.^{27,29,30} The

crystallinity of the materials is observed to increase with the increase in calcination temperature. This can be attributed to the enhancement in the definition of the crystalline phase by the reduction of defects, edges, strain, irregularities, *etc.* present in the particles due to heat treatment.^{30,31} This can also be due to the transformation of the elongated side-to-side packed anatase phase to the closed packed octahedrally oriented rutile phase with increased calcination temperature.³¹ The XRD pattern of TiO₂ 500 shows that it is constituted of only the anatase phase of TiO₂. Whereas the XRD patterns of TiO₂ 600 and TiO₂ 700 samples show the presence of both anatase and rutile phases in these samples, but in different amounts, the XRD pattern of TiO₂ 900 shows it is fully rutile. The crystallite

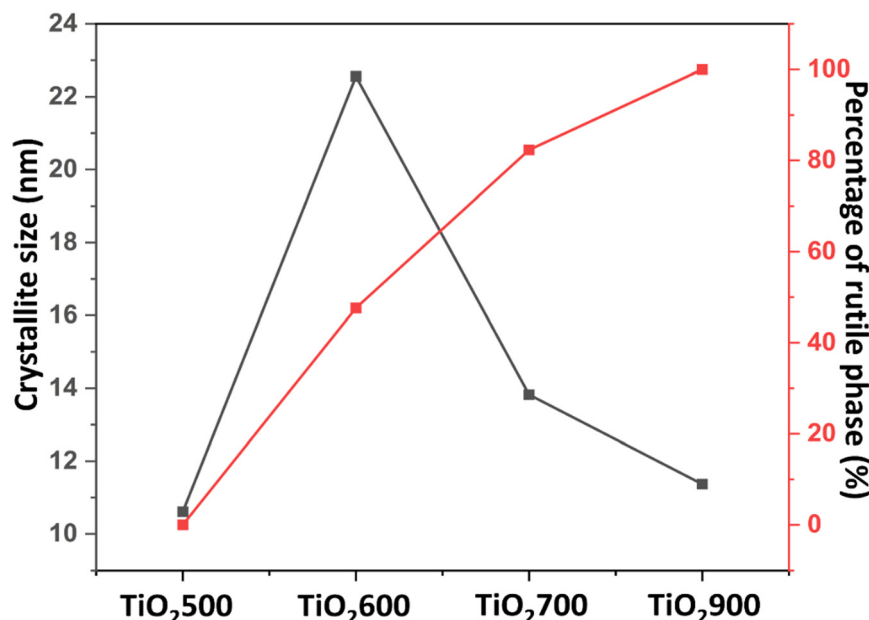


Fig. 3 Crystallite size and percentage of rutile phase of TiO_2 nanoparticles calcined at different temperatures.

size (φ) in nm, of TiO_2 nanoparticles was estimated by the Debye-Scherrer's equation (eqn (1))^{26,32} as

$$\varphi = \frac{K\lambda}{\beta \cos \theta} \quad (1)$$

where φ is the crystallite size, K (0.9) is the shape factor, λ is the wavelength (1.5406 Å) of the X-ray used, β is the full width at half maxima (FWHM) of the peaks and θ is the Bragg's angle. From Fig. 3, which shows the estimated values of the crystallite size of the particles, it can be noted that the crystallite size of the particles increases with calcination temperature from 500 °C to 600 °C and then it reduces as the calcination temperature increases further up to 900 °C. This variation in crystallite size is due to the change in the percentage of anatase and rutile phases in the TiO_2 nanoparticles. TiO_2 600 has both anatase and rutile phases present in them which is the reason why TiO_2 600 has an increased crystallite size compared to that of TiO_2 500. TiO_2 700 also shows the presence of both anatase and rutile phases. However, its crystallite size is lower than that of TiO_2 600 as the percentage of anatase phase is considerably decreased in TiO_2 700 and it is in the process of getting converted to a single crystalline phase TiO_2 . From this observation, it is clear that the presence of two different crystalline phases in TiO_2 results in an increase in their crystalline size compared to that of TiO_2 with only one crystalline phase. The percentage of anatase and rutile phases present in TiO_2 600 and TiO_2 700 (Fig. 3) were calculated by considering the intensities of anatase (101) and rutile (110) using the Spurr equation (eqn (2))³³ as

$$F_R = \frac{1}{1 + 0.8[I_A(101)/I_R(110)]} \quad (2)$$

where F_R is the fraction of the rutile phase in the TiO_2 sample,

I_A (101) is the intensity of (101) peak of the anatase phase and I_R (110) is the intensity of (110) of the rutile phase.

According to the Spurr equation, it was calculated that the percentage of the anatase phase is 52.4% and that of the rutile phase is 47.6% in the TiO_2 600 sample, and the percentage of the anatase phase is 17.7%, and that of the rutile phase is 82.3% in TiO_2 700 sample. The percentage of the anatase phase is reduced compared to that of the rutile phase as the calcination temperature is increased. On increasing the calcination temperature further from 700 °C to 900 °C it was found that the anatase phase has completely disappeared and TiO_2 900 contains only the rutile crystalline phase. From this, it can be inferred that the increase in calcination temperature favours the formation of a thermodynamically stable rutile phase. Also, the presence of both anatase and rutile phases in TiO_2 600 and TiO_2 700 shows that the phase transformation has taken place gradually.²⁹ XRD patterns of TiO_2 nanoparticles calcined at temperatures ranging from 400 °C to 1000 °C are shown in Fig. S2 (ESI†) to confirm that TiO_2 formation is not complete below 500 °C, there is no mixed phase present above 700 °C and both TiO_2 800 and TiO_2 1000 has only rutile phase.

Increasing the content of the electroactive phase (EA) in the PVDF matrix is crucial for many applications.^{19,34} Similarly, the TENG output of PVDF could also be impacted by the percentage of different electroactive phases present in PVDF. FT-IR analysis is widely used to identify the percentage of crystalline phases present in PVDF.^{35,36} FT-IR analysis of PVDF and PVDF- TiO_2 films (designated as PVT) with TiO_2 having different crystalline phases was carried out to evaluate the changes in the crystalline phase of PVDF and their percentage with the incorporation of TiO_2 . Since the spin coating parameters for all the prepared films were the same (Section 2.3), the comparative changes observed in the electroactive phases in PVDF between the films



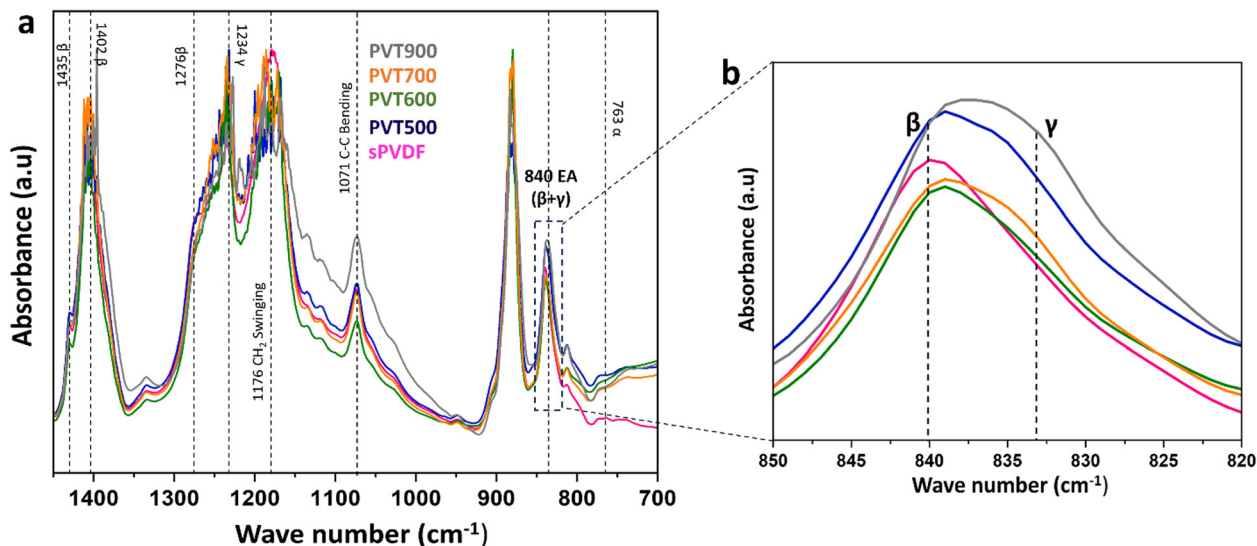


Fig. 4 FT-IR analysis of PVDF-TiO₂ composites. Comparison of FT-IR spectra (absorbance vs. wavenumber) tested in the range of (a) wavenumber 1450–700 cm⁻¹ and (b) wavenumber 850–820 cm⁻¹ of sPVDF, PVT500, PVT600, PVT700 and PVT900.

can be attributed to the influence of TiO₂ particles. Fig. 4(a) and (b) compare the FT-IR spectra of PVDF and PVDF-TiO₂ composites. The peak observed at 1071 cm⁻¹ corresponds to the bending vibration of C–C, the peak at 1176 cm⁻¹ corresponds to the swinging vibration of CH₂ and the peak at 1397 cm⁻¹ is due to CF₂ in PVDF.³⁷ The intensification of the vibration band at 1232 cm⁻¹ for PVDF-TiO₂ films shows the nucleation of β crystalline phase in PVDF with the incorporation of TiO₂. Also, the absence of identity peaks of the α phase of PVDF at 763 cm⁻¹ and 976 cm⁻¹ in the spectra is an indication of the conversion of the α phase to electroactive phases (β and γ) of PVDF. Also, the peaks at 1276 cm⁻¹ and 1402 cm⁻¹ are assigned to the β phase and the one at 1234 cm⁻¹ is attributed to the γ phase. Another major peak observed at 840 cm⁻¹ is considered the electroactive phase of PVDF since it is a combination of both β and γ phases and correlated with the intensities of the peaks at 1276 cm⁻¹ and 1234 cm⁻¹.^{21,35,38} For this, the percentage of electroactive phase present in each sample was calculated from the FT-IR spectra using eqn (3)³⁵ as

$$F_{EA} = \frac{A_{EA}}{\left(\frac{K_{840}}{K_{763}}\right)A_{763} + A_{EA}} \times 100\% \quad (3)$$

where A_{EA} and A_{763} are absorbances at 840 and 763 cm⁻¹ respectively and the absorption coefficients at these wave numbers (K_{840} and K_{763}) are 7.7×10^4 and 6.1×10^4 cm² mol⁻¹ respectively. The fraction of EA phases of PVT500, PVT600, PVT700 and PVT900 were calculated as 71.61%, 68.28%, 72.55% and 75.74% respectively. On comparing the spectra of PVDF-TiO₂ samples with that of PVDF, it can be observed that the intensity of electroactive phases in PVDF is increased. This shows that incorporating TiO₂ into the PVDF matrix favours the formation of more electroactive phases in PVDF. Furthermore, since the TiO₂ nanoparticles incorporated into the PVDF matrix occur in different crystalline phases (as explained in the XRD

analysis), a more detailed analysis of FT-IR is required to understand how the crystalline phase of TiO₂ impacts the phase transformation in PVDF. Fig. 4(b) shows that the peak at 840 cm⁻¹ which corresponds to the EA phase of PVDF is widening with the incorporation of TiO₂ nanoparticles with different crystalline phases. This is a clear indication of the change in the EA phase of PVDF with the addition of TiO₂ and also that the crystalline phase of TiO₂ affects how the EA phase of PVDF-TiO₂ is being manipulated. A shift in the peak towards the lower wave number is observed in Fig. 4(b) and this shows the increase in the contribution of the γ phase in the EA phase which is a mixture of β and γ phase.^{35,39} The intensification of the peak at 840 cm⁻¹ can also be attributed to the enhancement of γ phase.³⁹ Fig. 4(b) shows that the γ content is higher for films containing full rutile (PVT900) and fully anatase (PVT500) and comparatively lower for PVT films containing mixed-phase TiO₂ (PVT600 and PVT700). However, the percentage of γ phase is the highest for PVT900.

XRD patterns of spin-coated PVDF film and PVDF-TiO₂ nanocomposite films (Fig. 5(a)–(c)) were analysed to understand the crystalline phase of the PVDF in them. Fig. 5(a) and (b) compares the XRD patterns of PVDF in powder form and spin-coated PVDF film. The two intense peaks observed in PVDF powder at 18.70° and at 20.20° and the medium peak at 26.5°, which corresponds to (020), (110) and (021) planes, respectively, are the major peaks of the monoclinic α-crystalline phase of PVDF which is considered as the kinetically stable phase.^{16,23,40} In the case of spin-coated PVDF, the intensity of XRD peaks reduced and the peaks got broadened and shifted towards the right to 20.50°. This shows the enhancement of the β-crystalline phase in spin-coated PVDF. This is an indication of the transformation of the α-crystalline phase to the β-crystalline phase in PVDF by spin coating and the drying of samples at 120 °C. XRD peaks of both PVDF and TiO₂ are observed in Fig. 5(c) which shows the XRD patterns of all the

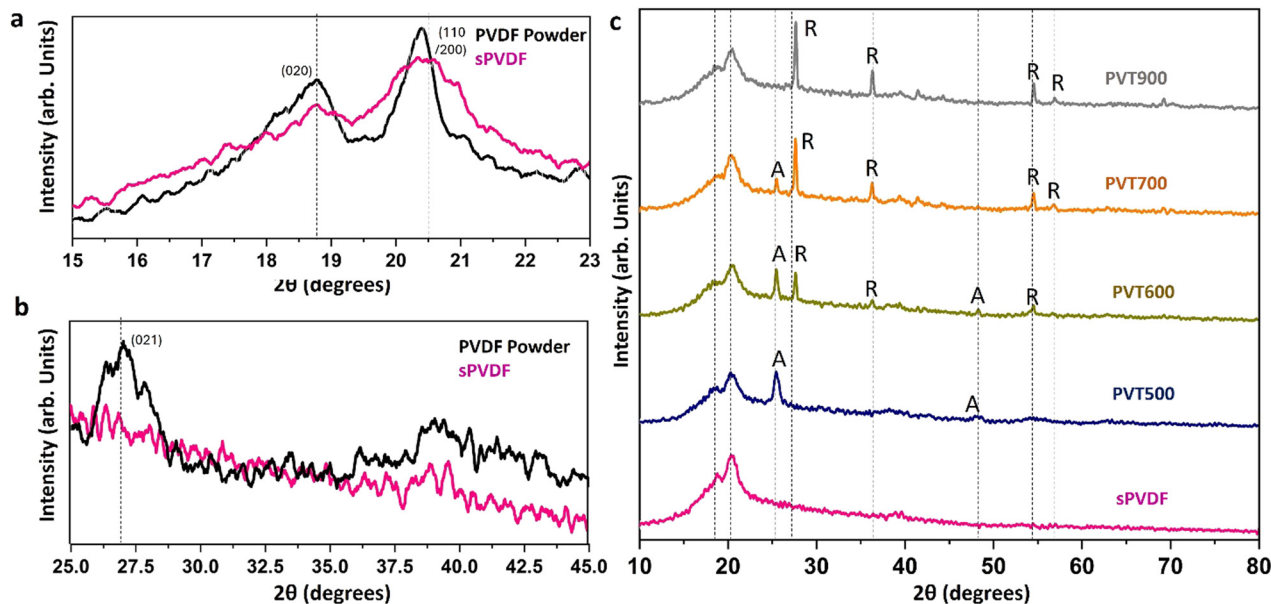


Fig. 5 XRD analysis of PVDF and PVDF–TiO₂ composite films. (a) and (b) Comparison of XRD patterns of PVDF powder and spin-coated PVDF (sPVDF) and (c) comparison of XRD patterns of sPVDF, PVT500, PVT600, PVT700 and PVT900.

PVDF–TiO₂ nanocomposite films prepared by spin coating. All the PVDF–TiO₂ composites have the same mass loading of TiO₂ (3 wt%). The presence of anatase and rutile phases in TiO₂ is also reflected in these XRD patterns of PVDF–TiO₂ composites. It can be observed that with the addition of TiO₂ in PVDF, the intensity of the (110/200) peak corresponding to the β -crystalline phase reduces showing further enhancement of the β -crystalline phase.⁴¹ Also, with the incorporation of TiO₂, a reduction in crystallinity and enhancement of amorphous regions in PVDF is observed owing to the interaction of polymer chains with TiO₂ nanoparticles.⁴⁰ Fig. 6 shows the comparison of the XRD pattern of the PVT films in the 2θ range 15°–30°. It can be observed that the intensity of the major peak of PVDF at 20.5° is similar for PVDF, PVT500 (anatase) and PVT600

(anatase and rutile). Whereas the intensity of this peak is considerably decreased in the case of PVT900 (fully rutile). This shows that the TiO₂ nanoparticles with rutile crystalline phase impart more enhancement of the PVDF β -crystalline phase than other TiO₂ nanoparticles.

3.2. Surface topography analysis of PVDF–TiO₂ nanocomposite films

SEM scans and 3D profilometry images of PVDF and PVDF–TiO₂ films were analysed to observe the surface morphology of the films and to understand the distribution of TiO₂ nanoparticles in the PVDF matrix in these films. The surface scans of the PVT500, PVT600 and PVT900 from a 3D optical profilometer and their SEM scans are shown in Fig. 7(a)–(f). It can be

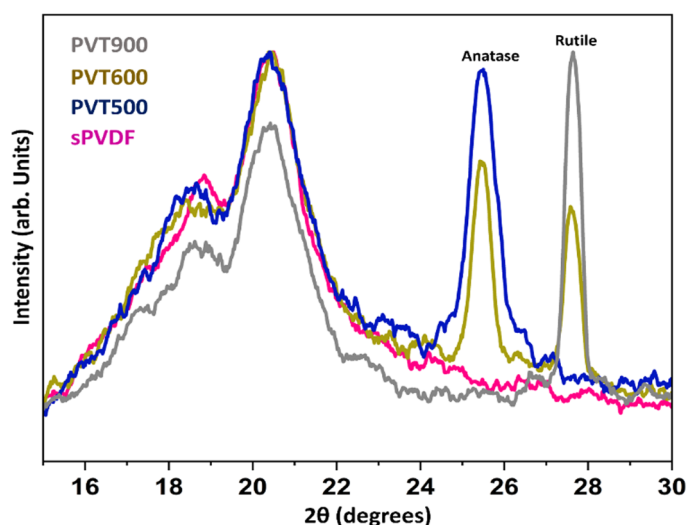


Fig. 6 Comparison of XRD patterns of sPVDF, PVT500, PVT600 and PVT900 in the 2θ range 15°–30°.



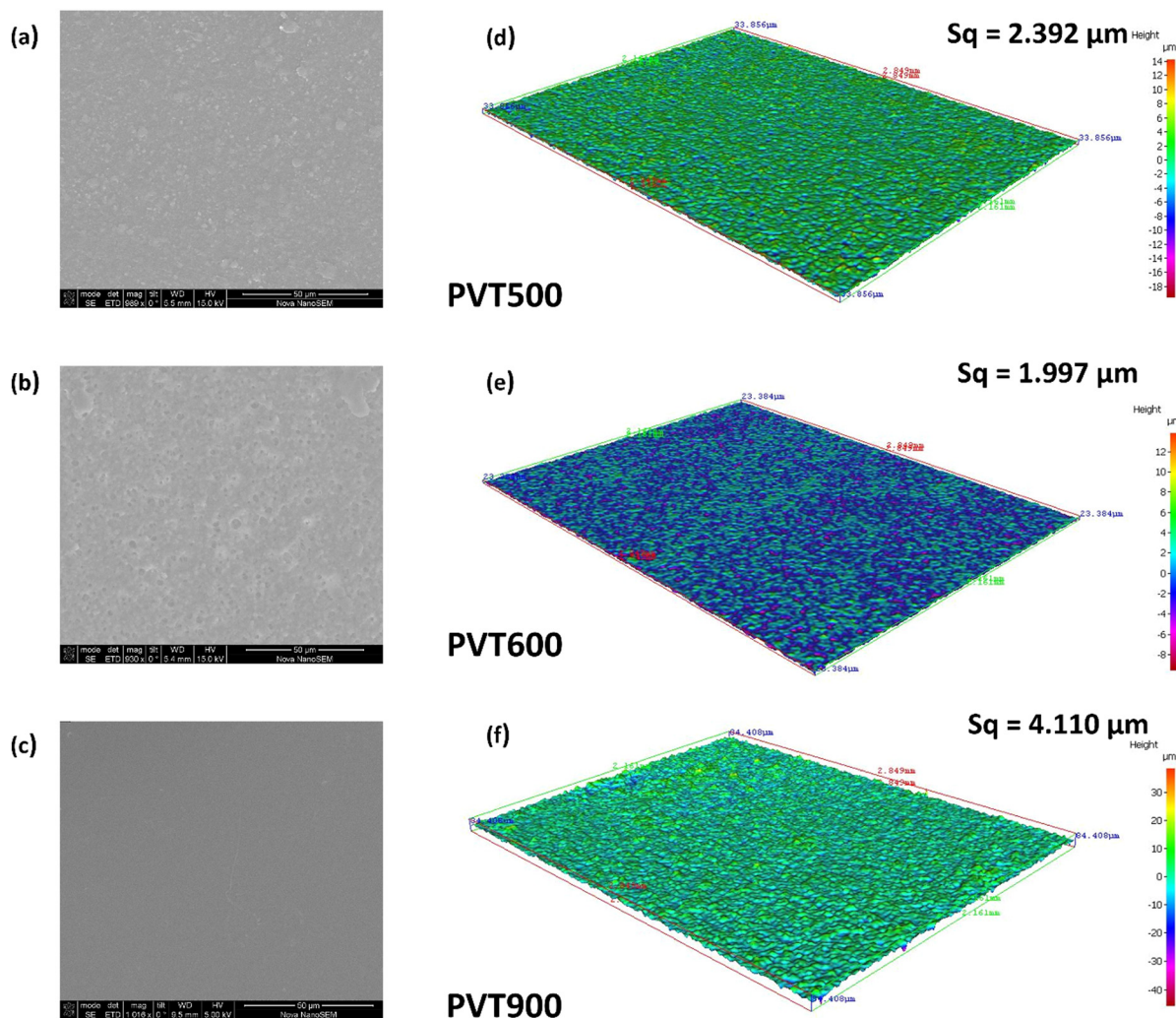


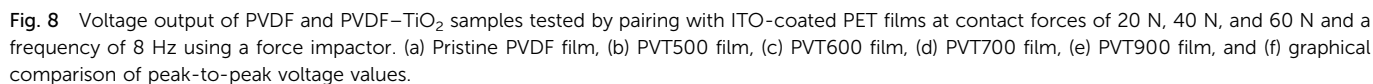
Fig. 7 SEM images (a–c) of the PVT500, PVT600 and PVT900 films and their surface scans (d–f) from a 3D optical profilometer.

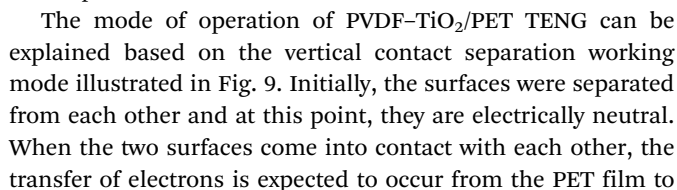
observed that the roughness of the surface of the PVT900 film is higher than that of the PVT500 and PVT600 film. PVT900 shows a root mean square (RMS) surface roughness of 4.110 μm , which is higher than the S_q 1.997 μm and 2.392 μm of PVT600 and PVT500 respectively. This can be again correlated to the influence of crystallinity and crystallite size of the TiO_2 nanoparticles incorporated in these films. Since the average crystallite size of rutile TiO_2 is higher than that of the anatase TiO_2 , the surface roughness of PVT900 is more than that of PVT600. Also, the RMS surface roughness of PVT500 falls in between that of PVT600 and PVT900. Fig. S3 (ESI[†]) shows the SEM scans of drop-casted and spin-coated PVDF films. On comparing the images of PVDF, spin-coated film and PVDF drop-casted films it can be observed that the spin-coated films show a more smooth and oriented film surface compared to that of drop-casted PVDF film. This can be an indication of the presence of a more crystalline phase than the amorphous phase in spin-coated PVDF compared to that of drop-casted PVDF film. Beta crystalline phase is reported to be favourable for triboelectric properties of PVDF compared to other crystalline

phases.^{20,42} From Fig. S4 (ESI[†]), TiO_2 nanoparticles embedded in the PVDF matrix are observed. It can be noted that the crystal size of TiO_2 embedded in the PVDF matrix is increasing with an increase in calcination temperature which agrees with the trend in crystallite size of TiO_2 nanoparticles estimated from their XRD patterns. Moreover, a homogeneous distribution of TiO_2 nanoparticles with a minimum agglomeration of reinforced particles is obtained in all the PVDF- TiO_2 films. This could be due to the overnight high-speed magnetic stirring employed to disperse the TiO_2 samples in PVDF solution. Moreover, the concentration of PVDF in the solution and the percentage loading of TiO_2 which was optimized for maintaining the uniformity of polymer nanocomposite was appropriate to achieve uniform distribution of the TiO_2 nanoparticles in the PVDF matrix. The SEM scans of PVDF and PVDF- TiO_2 samples clearly show the change in surface morphology of PVDF film with a change in crystallinity of the incorporated TiO_2 nanoparticles. It is visible that the surface roughness of the PVDF film is increasing compared to that of pristine PVDF film, with the incorporation of TiO_2 nanoparticles.



from Fig. 8(a)–(e). It is also expected that the electrical output signals (AC) may not be uniform and symmetric in the positive and negative cycle. This asymmetry is because of the nature of contact-separation between the materials which occurs from the operation of the testing device used. Therefore, there could be differences between the output signal generated during the contact cycles and during the separation cycles.^{17,21,43} For the PVDF/PET pairing, the output voltage obtained was ~ 25 V. It increased up to ~ 60 V (2.4 times) for the PVT900/PET sample which was the highest voltage value obtained among all the samples. The increase in voltage value under the tested conditions with the incorporation of TiO_2 in the PVDF matrix confirms that the triboelectric properties of PVDF film are increased with the loading of TiO_2 . However, the increase in voltage in PVDF- TiO_2 films was observed to be impacted by the crystalline phase of the TiO_2 in the PVDF matrix. Taking the example of the output values at 60 N and 8 Hz, the peak-peak voltage values obtained are ~ 35 V, 76 V, 33 V, 80 V and 105 V for PVDF, PVT500, PVT600 and PVT900 samples (Fig. 8(f)) respectively. The peak-peak voltage output of PVDF is doubled with the addition of anatase phase TiO_2 (PVT500) and it reduces considerably with mixed TiO_2 (PVT600) which has an almost equal percentage of anatase and rutile phase (Fig. 3). This can be attributed to the decrease in the beta crystalline phase observed in the PVT600 film compared to the other films. This can be also validated by the surface topography results obtained using the 3D profilometer. The PVT700 sample which also has a mix of anatase and rutile phases, but a considerably higher percentage of rutile than anatase, shows an increase in





power of the TENG device at different applied external loads was measured by the equation⁴⁵ as

$$P = \frac{I^2 R}{A} \quad (4)$$

The self-powering capability of the TENG was estimated by charging different commercial capacitors with capacitances of 1, 4.7, 10, 22 and 47 μF using the PVT900 sample. Fig. 11(b) shows that the 1 μF capacitor can be charged up to ~ 2.5 V,

Figure 1 consists of four panels. Panel (a) is a plot of Voltage (V) versus Time (s) from 0 to 2000 s. The voltage fluctuates between approximately 40 V and 80 V. An inset shows a zoomed-in view of the voltage spikes between 1000 and 1500 s. Panel (b) is a plot of Voltage (V) versus Time (s) from 0 to 100 s, showing the charging curves for different capacitors: 1 μF (black), 4.7 μF (red), 10 μF (blue), 22 μF (magenta), and 47 μF (green). Panel (c) is a photograph of the experimental setup, showing a dynamometer, a motor, and a sample. Panel (d) is a photograph of the sample, showing a grid of light spots.

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Table 1 Comparison of the electrical output of the TENG device with the TENGs with similar research methods reported in the literature

TENG system	Material(s) used	Output performance	Application(s)	Ref.
Electrospun flexible membrane based TENG	PVDF–ZnO NWs/nylon–ZnO NWs	Power density = 3.0 W m^{-2}	Mechanical energy harvesting	17
Phase inversion enabled film based TENG	PVDF–TiO ₂ /cellulose acetate	Output voltage $\sim 60 \text{ V}$ Output current $\sim 450 \text{ nA}$	Benzene monitoring system	21
Droplet based TENG	PVDF – lead free potassium sodium niobate (KNN–PVDF)	Output voltage = 1.56 V Output current = $9.91 \text{ }\mu\text{A}$	Energy harvesting using sea waves	47
Electrospun TENG	Copper oxide (CuO) nanoparticles doped PVDF	Output voltage = 7.5 V	Wearable and biomedical applications	48
Thermally drawn polymer nanocomposite fibre based TENG	Graphene integrated PVDF	Output voltage $\sim 35 \text{ V}$ Output current $\sim 20 \text{ }\mu\text{A}$	Application in extreme environmental conditions	43
Stretchable, transparent TENG	MXene nanosheets, silver nanowires, PU, PDMS	$V_{oc} = 38 \text{ V}$ Current density = 1.67 mA m^{-2}	Motion sensor	49
Fully stretchable and durable TENG	Au nanosheets embedded PDMS matrix	$V_{oc} = 98.9 \text{ V}$ $I_{sc} = 2.8 \text{ }\mu\text{A}$	Medical diagnostics and wireless sensors	50
Highly flexible TENGs	TiO ₂ nanorod arrays, titanium substrate	$V_{oc} = 40 \text{ V}$ Current density = 1 mA cm^{-2}	Biosensors, photo detectors	51
Electroactive phase dependant PVDF–TiO ₂ nanocomposite based TENG	PVDF–TiO ₂ /ITO coated PET	Output voltage $\sim 60 \text{ V}$ Output current $\sim 3 \text{ }\mu\text{A}$	Mechanical energy harvesting	This work

4.7 μF capacitor up to $\sim 2.5 \text{ V}$, 10 μF capacitor up to $\sim 1.5 \text{ V}$, 22 μF capacitor up to $\sim 1.0 \text{ V}$ and 47 μF capacitor up to $\sim 0.45 \text{ V}$ within 100 seconds of charging. The rectifier circuit used to charge the capacitors using the TENG is shown in Fig. 10(d). Thus, the fabricated TENG can be used to power electronic devices by storing the generated output in a capacitor. Additionally, as a demonstration of the TENG ability of the fabricated device, 46 LEDs connected in series were illuminated as shown in Fig. 11(d) and Fig. S5 (ESI[†]). One of the main limitations of TENGs is their low current output (microampere range) owing to the high internal impedance in them (mega Ohm range). Those limitations are also observed in the TENG shown in the current study. The maximum current output obtained was approximately $3 \text{ }\mu\text{A}$ (Fig. 10(a)) and the internal impedance of TENG was observed to be in the mega Ohm range from the load-matching studies (Fig. 11(b)).

Considerable change in the TENG output performance was observed based on the crystalline phase of TiO₂ nanoparticles incorporated in the PVDF matrix. A gradual change in the crystalline phase of TiO₂ was observed by calcining the samples over a range of temperatures from $500 \text{ }^\circ\text{C}$ to $1000 \text{ }^\circ\text{C}$ and the percentage of the rutile phase increased with an increase in the calcination temperature of the TiO₂ nanoparticles.^{29,46} The difference in crystalline phase modulation of PVDF based on the crystalline phase of the incorporated TiO₂ was observed from the XRD and FT-IR analysis of the films.⁴³ It was shown that the electroactive phase (EA) percentage of PVT500 and PVT900 films with pure anatase and pure rutile TiO₂ respectively was higher than that of the mixed phase TiO₂ incorporated PVDF films. This change in the EA phase content of PVDF–TiO₂ films influenced the electrical output of the PVDF–TiO₂ films when paired with PET. The pristine PVDF film showed a peak-to-peak voltage of $\sim 35 \text{ V}$ and this increased up to $\sim 77 \text{ V}$ for PVT500 films which have anatase phase TiO₂ nanoparticles. The peak-to-peak voltage of PVT600 which has a mix of anatase (52.4%) and rutile (47.6%) phase was found to be considerably decreased to $\sim 33 \text{ V}$. This shows that apart

from the crystallinity and percentage loading of TiO₂, the crystalline phase of TiO₂ has an impact on the TENG output of PVDF–TiO₂ composites.^{21,22} This was further confirmed by evaluating the voltage output of PVT700 which also had a mixed phase of TiO₂ but with a higher percentage of rutile phase. The peak-to-peak voltage value of PVT700 increased up to $\sim 82 \text{ V}$. The maximum peak-to-peak voltage was attained for PVT900 ($\sim 105 \text{ V}$) with a pure rutile phase TiO₂. From these results, it can be inferred that the rutile phase of TiO₂ imparts a higher content of EA phase in PVDF–TiO₂ nanocomposites compared to that of other phases. Owing to this, the TENG output of rutile phase TiO₂ containing PVDF is higher than that of anatase and mixed phase containing TiO₂ nanoparticles. This study showed comparable voltage output to the PVDF nanocomposite-based and other flexible and stretchable TENGs reported in the literature (Table 1). Moreover, the fabrication of the proposed TENG involves low-cost techniques and materials, making the TENG worth approximately 50 EUR. A simple spin coating method was used to prepare the tribonegative film (Section 2.3), and low-cost materials such as Kapton tape and acrylic sheets were used to assemble this 2.5 cm^2 TENG (Section 2.4). However, in TENGs the comparison of electrical output with similar reported TENGs is not a reliable practice for some obvious reasons. Firstly, the output will vary depending on the properties of the opposite contact layer used as well. Also, the conditions of contact separation, force and frequency at which the TENGs were analysed, humidity conditions *etc.* will affect the output. Furthermore, concerning the force impactor instrument/rig used for testing, the contact mechanics of the TENG will vary from rig to rig as well as due to the two different materials involved in the operation of TENG.

4. Conclusions

The impact of crystallinity and crystalline phase of TiO₂ nanoparticles on the crystalline phase modification of the PVDF



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The data supporting this article have been included as part of the ESI.†

There are no conflicts to declare.

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