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A novel Ta/Os heterobimetallic complex, [Ta(CH₂tBu)₃(μ-H)₃OsCp*], **2**, is prepared by protonolysis of Ta(CHtBu)(CH₂tBu)₃ with Cp*OsH₅. Treatment of **2** and its iridium analogue [Ta(CH₂tBu)₃(μ-H)₂IrCp*], **1**, with CO₂ under mild conditions reveal the efficient cleavage of CO₂, driven by the formation of a tantalum oxo species in conjunction with CO transfer to the osmium or iridium fragments, to form Cp*Ir(CO)H₂ and Cp*Os(CO)H₃, respectively. This bimetallic reactivity diverges from more classical CO₂ insertion into metal–X (X = metal, hydride, alkyl) bonds.

The design of synthetic bimetallic complexes associating different metals with complementary Lewis acidic/Lewis basic behaviour has raised interest for cooperative reactivity,^{1–10} including CO₂ activation.^{11–17} In many instances, these bifunctional complexes lead to CO₂ adducts or insertion products, in which a bent CO₂ fragment binds across the two metals.^{18–27} In contrast, only a few heterobimetallic complexes have clearly exhibited the capability to cleave the C–O bond within CO₂. Thomas and colleagues reported oxidative CO₂ cleavage across the early/late heterobimetallic complex Co(iPr₂PNMes)₃Zr(THF), yielding (OC)Co(iPr₂PNMes)₂(μ-O)Zr(iPr₂PNMes) at ambient temperature (Scheme 1a).²⁸ The Mazzanti group reported the potassium-assisted reductive cleavage of CO₂ by a U(III) siloxide complex, resulting in CO evolution and the formation of a pentavalent

CO₂ cleavage by tantalum/M (M = iridium, osmium) heterobimetallic complexes†

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uranium oxo complex (Scheme 1b). When the potassium cation is encapsulated in 18-crown-6, bimetallic cooperativity no longer occurs, and a carbonate complex is formed instead. Our group has developed an Ir/Al-based heterobimetallic complex proficient in CO₂ deoxygenation, yielding Cp*Ir(CO)H₂, Cp*IrH₄, and [Al(Py)(OAr)(iBu)₂(μ-O)] at room temperature (Scheme 1c).²⁹ Recently, Campos and coworkers reported the use of Al(C₆F₅)₃ for triggering the bimetallic cleavage of Fe-bound CO₂ moiety, to form an oxo carbonyl complex (Scheme 1d). The choice of the Al-based Lewis acid partner plays a pivotal role in initiating this reaction, as boron, zinc, or gold Lewis acids did not exhibit activity in this transformation.²² Finding right bimetallic combinations therefore remains a major challenge for controlling reactivity.



Scheme 1 Reports of CO₂ cleavage by heterobimetallic complexes relevant to the present study.^{22,28–30}

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The alkane elimination reaction between metal (poly)alkyls and (poly)hydride species has proven efficient for accessing heterobimetallic complexes.^{31–36} We used this strategy to synthesize compound $[\text{Ta}(\text{CH}_2\text{tBu})_3\text{IrH}_2\text{Cp}^*]$ **1** from the tantalum tris-neopentyl neopentylidene complex $\text{Ta}(\text{CHtBu})(\text{CH}_2\text{tBu})_3$ and the iridium tetrahydride complex Cp^*IrH_4 (Scheme 2-top).³⁷ This prompted us to extend this chemistry by investigating the reactivity of $\text{Ta}(\text{CHtBu})(\text{CH}_2\text{tBu})_3$ towards related 6d metal polyhydrides. Treating $\text{Ta}(\text{CHtBu})(\text{CH}_2\text{tBu})_3$ with Cp^*OsH_5 ^{38,39} in a 1:1 stoichiometric ratio in pentane at room temperature yields the heterobimetallic complex $[\text{Ta}(\text{CH}_2\text{tBu})_3(\mu\text{-H})_3\text{OsCp}^*]$ **2** in 98% isolated yield, accompanied by the elimination of one equivalent of neopentane (Scheme 2-middle). ¹H NMR monitoring of the reaction of **2** with Cp^*OsH_5 (1 equiv.) suggests the slow formation of a trinuclear TaOs₂ species (see Fig. S7, ESI†), analogous to the TaIr_2 species previously reported.⁴⁰ Surprisingly, Cp^*ReH_6 ⁴¹ shows no reactivity towards $\text{Ta}(\text{CHtBu})(\text{CH}_2\text{tBu})_3$ either in pentane at room temperature or in C_6D_6 at 80 °C. DFT calculations indicate that the Ta/Re analogue should be thermodynamically stable: this observed lack of reactivity is thus surprising, and might be due to the lack of available coordination site at Re (see ESI† for discussion).

Identification of **2** is confirmed through a range of analytical methods including infrared (IR) and multinuclei (¹H, ¹³C, ¹H–¹H COSY, ¹H–¹³C HSQC and HMBC) solution NMR spectroscopy, elemental analysis, and X-ray diffraction studies. In the ¹H-NMR spectrum of **2** obtained in a toluene-*d*₈ solution, the hydride resonance appears as a high-field singlet at $\delta = -6.90$ ppm, integrating for 3H. The hydride resonance in **2** exhibits a shift of $\Delta\delta = +4.3$ ppm from Cp^*OsH_5 which is reminiscent of the observed shift from Cp^*IrCH_4 to **1**, of approximately $\Delta\delta = +3.5$ ppm.³⁷ The IR spectrum of compound **2** displays a characteristic metal-hydride stretching vibration signal at 1961 cm^{-1} , consistent with bridging hydrides. This value deviates significantly from that of complex $[\text{Ta}(\text{CH}_2\text{tBu})_3\text{IrCp}^*(\text{H})_2]$, **1**, featuring two terminal hydrides ($\nu_{\text{Ir-H}} = 2061\text{ cm}^{-1}$, see Fig. S6, ESI†) and that of the Cp^*OsH_5 precursor, which exhibits a strong absorption at $2083\text{ (s)}\text{ cm}^{-1}$ with a minor



Scheme 2 Reactivity of $\text{Ta}(\text{CHtBu})(\text{CH}_2\text{tBu})_3$ towards Cp^*IrH_4 , Cp^*OsH_5 and Cp^*ReH_6 .



Fig. 1 Solid-state molecular structure of **2** (30% probability ellipsoids). Hydrogen atoms from the hydrocarbon ligands are omitted for clarity. Selected bond distances (Å) and angles (°): Ta1–Os1 2.4817(2), Ta1–C1 2.132(3), Ta1–C2 2.136(3), Ta1–C3 2.143(3), Os1–H1 1.45(5), Os1–H2 1.56(5), Os1–H3 1.48(8), Ta1–H1 2.07(5), Ta1–H2 2.03(5), Ta1–H3 2.07(8), Ta1–Os1– $\text{Cp}^*\text{centroid}$ 178.1(1).

one at $2214\text{ (w)}\text{ cm}^{-1}$.⁴² For comparison, the metal-hydride stretch is observed at 1982 cm^{-1} in $[\text{Hf}(\text{CH}_2\text{tBu})_3(\mu\text{-H})_3\text{IrCp}^*]$ ³⁴ and at 1952 and 1970 cm^{-1} for $[\text{Cp}_2\text{Zr}(\text{X})(\mu\text{-H})_3\text{Os}(\text{PMe}_2\text{Ph})_3]$ ($\text{X} = \text{Cl}$ or H respectively),⁴³ where the two metal centres are bridged by three hydride ligands.

Single crystals of **2** suitable for X-ray diffraction were grown from a saturated pentane solution at -40 °C . The solid-state structure is depicted in Fig. 1. The nearly linear Ta–Os– $\text{Cp}^*\text{centroid}$ angle ($178.1(1)^\circ$), indicates the presence of three bridging hydrides between the two metals, arranged in a tripod geometry around the $\{\text{Cp}^*\text{Os}\}$ core. This angle aligns well with values reported for systems featuring similar bridging hydride motifs, such as $[\text{CpRu}(\mu\text{-H})_4\text{OsCp}^*]$ ($179.2(9)^\circ$)⁴⁴ and $[\text{Hf}(\text{CH}_2\text{tBu})_3(\mu\text{-H})_3\text{IrCp}^*]$ ($179.2(3)^\circ$),³⁴ but starkly contrasts with that found in complex **1**, featuring two terminal Ir–H moieties ($151.3(1)^\circ$). The Ta– C_{Np} bond lengths (with an average value of $2.137(5)\text{ Å}$) are consistent with neopentyl groups.^{45–47} The Ta–Os distance in compound **2** is $2.4817(2)\text{ Å}$, which is 0.115 Å shorter than the sum of the metallic radii of tantalum (1.343 Å) and osmium (1.255 Å).⁴⁸ This difference results in a formal shortness ratio (FSR) slightly below unity (FSR = 0.95),⁴⁹ suggestive of some degree of metal–metal interaction, although the presence of bridging hydrides could also explain the proximity. This FSR value lies between those of complex $[\text{Hf}(\text{CH}_2\text{tBu})_3(\mu\text{-H})_3\text{IrCp}^*]$ (FSR = 0.99),³⁴ where the close proximity between the Hf and Ir centres likely results from bridging hydrides, and the Ta/Ir complex **1** (FSR = 0.90),³⁷ which exhibits clear double metal–metal bonding.

To explore the potential of these heterobimetallic complexes in promoting cooperative reactivity, we investigated the reaction of **1** and **2** with CO_2 (1 atm, *ca.* 6 equiv.). The reactions were carried out in tetrahydrofuran (THF) at ambient temperature, resulting in rapid discoloration of the reaction mixture within *ca.* 10 minutes in both cases. Analysis of the crude reaction mixtures by ¹H NMR reveals the complete consumption of complexes **1** and **2**, with clean and quantitative formation of compounds $\text{Cp}^*\text{Os}(\text{CO})\text{H}_3$, **3** and $\text{Cp}^*\text{Ir}(\text{CO})\text{H}_2$, **4**, respectively derived from complexes **1** and **2**, alongside the generation of a tantalum oxo complex, $[\text{Ta}(\text{O})(\text{CH}_2\text{tBu})_3]_x$ **5** (refer to Fig. S9 and S10 in the ESI†).





Scheme 3 Reaction of compounds **1** and **2** with CO_2 , yielding $\text{Cp}^*\text{Ir}(\text{CO})\text{H}_2$ and $\text{Cp}^*\text{Os}(\text{CO})\text{H}_3$, respectively, together with the formation of $\text{Ta}(\text{O})(\text{CH}_2\text{tBu})_3$, **5**.

The insolubility of compound **5** in pentane facilitated its separation from the reaction mixtures by simple evaporation of THF followed by pentane extraction of **3** or **4**. Compound **3** was isolated in 96% yield; ^1H and ^{13}C NMR data are in agreement with the literature (Scheme 3).^{29,50}

The ^1H -NMR spectrum of **4**, recorded in THF-d_8 , indicates that the three hydrides are not equivalent in solution, resulting in two signals at -10.48 ppm and -12.50 ppm integrating for 1H and 2H, respectively and coupling in the ^1H - ^1H COSY NMR spectrum (Fig. S15, ESI[†]). These signals are assigned to hydrides in *-trans* and *-cis* positions relative to the CO group, respectively, which is consistent with literature data.⁵¹ The IR spectrum for **4** displays a broad terminal hydride stretching signal at 2075 cm^{-1} , and ν_{CO} bands at 1932 – 1898 cm^{-1} , as expected.⁵¹ Diluted THF solutions of complex **4** are stable at room temperature in the dark. Yet compound **4** is reported to be unstable in the solid-state,⁵¹ spontaneously eliminating H_2 upon drying, which could explain the moderate 45% isolated yield. Regardless, single crystals suitable for X-ray diffraction were obtained by avoiding visible light and crystallisation from pentane at $-40\text{ }^\circ\text{C}$. The solid-state structure of **4**, determined for the first time in this study, is shown in Fig. 2. The Os1–C1 (1.851(3) Å) and O1–C1 (1.162(4) Å) distances are consistent with those observed in compound $[\text{Cp}^*\text{Os}(\text{CO})(\mu\text{-H})]_2$, featuring Os–C bond lengths of 1.833(9) Å and C–O bond lengths of 1.18(1) Å.²⁹

The ^1H -NMR spectrum of **5** indicates that the three CH_2tBu groups are equivalent in solution, resulting in two signals at 0.55 ppm and 1.12 ppm for the CH_2 and *t*Bu moieties, respectively. Analysis of the $^{13}\text{C}\{^1\text{H}\}$ -NMR spectrum of **5** reveals three distinct characteristic resonances at 104.3, 35.1 and 34.4 ppm



Fig. 2 Solid-state molecular structure of **4** (30% probability ellipsoids). Hydrogen atoms from the Cp^* ligand are omitted for clarity. Selected bond distances (Å) and angles ($^\circ$): Os1–H1 1.53(4), Os1–H2 1.58(4), Os1–H3 1.51(4), Os1–C1 1.851(3), O1–C1 1.162(4), Os1– $\text{Cp}^*_{\text{centroid}}$ 1.916(1), C1–Os1– $\text{Cp}^*_{\text{centroid}}$ 132.5(1).

assigned to the TaCH_2 , $\text{C}(\text{CH}_3)_3$ and $\text{C}(\text{CH}_3)_3$ moieties, respectively. These assignments are confirmed by the 2D ^1H - ^1H COSY and ^1H - ^{13}C HSQC and HMBC data (Fig. S19–S21, ESI[†]). Unfortunately, we were unable to determine the XRD structure of **5**, which probably adopts oligomeric structures, given that terminal Ta-oxo species are rare in the literature.^{52–54} To confirm the identity of **5**, we thus carried out a high resolution mass spectrometry analysis using an APCI source, which shows a clear signal for the ion $[\text{Ta}(\text{O})(\text{CH}_2\text{tBu})_3 + \text{H}]^+$ at $411.2086\text{ }m/z$ (see Fig. S22, ESI[†]).

The computed reaction mechanism (DFT, B3PW91) is similar for **1** and **2**. CO_2 undergoes first a kinetically accessible (13 kcal mol^{-1} for **1**, 11 kcal mol^{-1} for **2**) nucleophilic attack by the Ir (or Os) center, which is assisted by oxygen-coordination to Ta. This results in 4-member metallacyclic intermediates shown on Fig. 3. The next step is a C–O bond breaking TS (barrier of 9 kcal mol^{-1} for **1** and 14 kcal mol^{-1} for **2**) to yield to products **3** (or **4**) and **5**, which formation is strongly exothermic (see ESI[†] for reaction profiles).

In summary, the reaction between Cp^*OsH_3 and $\text{Ta}(\text{CHtBu})(\text{CH}_2\text{tBu})_3$ affords a heterobimetallic Ta–Os complex, **2**, in high yields *via* alkane elimination. Complex **2**, along with its Ta–Ir analogue, **1**, exhibit clean CO_2 cleavage reactivity, driven by the formation of a tantalum oxo species in conjunction with late metal carbonyls. Given the propensity of related transition metal alkyls and hydrides for CO_2 insertion,^{55–59} the selective, divergent bimetallic reactivity observed herein is notable. These results clearly further demonstrate how the synergistic action of early/late metal assemblies – particularly those based on tantalum – can facilitate the deoxygenation of CO_2 . This understanding contributes to advancing knowledge in CO_2 activation and could lead to future applications in deoxygenative chemistry.

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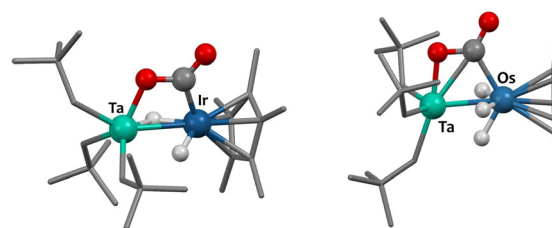


Fig. 3 Computed (DFT) structures of the metallacyclic reaction intermediates.



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Data availability

The data supporting this article have been included as part of the ESI.† CCDC 2351192 and 2351850 contain supplementary crystallographic data for this article; they can be obtained free of charge from The Cambridge Crystallographic Data Centre.

Conflicts of interest

There are no conflicts to declare.

References

- P. Buchwalter, J. Rosé and P. Braunstein, *Chem. Rev.*, 2015, **115**, 28–126.
- E. K. Van Den Beuken and B. L. Feringa, *Tetrahedron*, 1998, **54**, 12985–13011.
- J. A. Mata, F. E. Hahn and E. Peris, *Chem. Sci.*, 2014, **5**, 1723–1732.
- A. Lachguar, A. V. Pichugov, T. Neumann, Z. Dubrawski and C. Camp, *Dalt. Trans.*, 2023, **53**, 1393–1409.
- J. Campos, *Nat. Rev. Chem.*, 2020, **4**, 696–702.
- N. P. Mankad, *Chem. Commun.*, 2018, **54**, 1291–1302.
- B. G. Cooper, J. W. Napoline and C. M. Thomas, *Catal. Rev.*, 2012, **54**, 1–40.
- B. Chatterjee, W. C. Chang, S. Jena and C. Werlé, *ACS Catal.*, 2020, **10**, 14024–14055.
- T. S. Hollingsworth, R. L. Hollingsworth, R. L. Lord and S. Groyzman, *Dalt. Trans.*, 2018, **47**, 10017–10024.
- C. Z. Ye, I. Del Rosal, S. N. Kelly, I. J. Brackbill, L. Maron, C. Camp and J. Arnold, *Chem. Sci.*, 2024, **15**, 9784–9792.
- C. Zhang, P. Gotico, R. Guillot, D. Dragoe, W. Leibl, Z. Halime and A. Aukauloo, *Angew. Chem., Int. Ed.*, 2023, **62**, e202214665.
- D. Ghosh, S. Sinhababu, B. D. Santarsiero and N. P. Mankad, *J. Am. Chem. Soc.*, 2020, **142**, 12635–12642.
- M. Pérez-Jiménez, H. Corona, F. de la Cruz-Martínez and J. Campos, *Chem. – A Eur. J.*, 2023, **29**, e202301428.
- Z. B. G. Fickenscher, P. Lönnecke, A. K. Müller, O. Hollóczki, B. Kirchner and E. Hey-Hawkins, *Molecules*, 2023, **28**, 2574.
- J. Ye, R. C. Cammarota, J. Xie, M. V. Vollmer, D. G. Truhlar, C. J. Cramer, C. C. Lu and L. Gagliardi, *ACS Catal.*, 2018, **8**, 4955–4968.
- J. R. Prat, C. A. Gaggioli, R. C. Cammarota, E. Bill, L. Gagliardi and C. C. Lu, *Inorg. Chem.*, 2020, **59**, 14251–14262.
- A. Lachguar, I. Del Rosal, L. Maron, E. Jeanneau, L. Veyre, C. Thieuleux and C. Camp, *J. Am. Chem. Soc.*, 2024, DOI: [10.1021/jacs.4c02172](https://doi.org/10.1021/jacs.4c02172).
- S. Sinhababu, M. R. Radzhabov, J. Telser and N. P. Mankad, *J. Am. Chem. Soc.*, 2022, **144**, 3210–3221.
- J. Hicks, A. Mansikkamäki, P. Vasko, J. M. Goicoechea and S. Aldridge, *Nat. Chem.*, 2019, **11**, 237–241.
- M. Devillard, R. Declercq, E. Nicolas, A. W. Ehlers, J. Backs, N. Saffon-Merceron, G. Bouhadir, J. C. Slootweg, W. Uhl and D. Bourissou, *J. Am. Chem. Soc.*, 2016, **138**, 4917–4926.
- G. Fachinetti, C. Floriani and P. F. Zanazzi, *J. Am. Chem. Soc.*, 1978, **100**, 7405–7407.
- H. Corona, M. Pérez-Jiménez, F. de la Cruz-Martínez, I. Fernández and J. Campos, *Angew. Chem., Int. Ed.*, 2022, **61**, e202207581.
- C. Yoo and Y. Lee, *Chem. Sci.*, 2016, **8**, 600–605.
- E. G. Lundquist, J. C. Huffman and K. G. Caulton, *J. Am. Chem. Soc.*, 1986, **108**, 8309–8310.
- T. A. Hanna, A. M. Baranger and R. G. Bergman, *J. Am. Chem. Soc.*, 1995, **117**, 11363–11364.
- J. R. Pinkes, B. D. Steffey, J. C. Vites and A. R. Cutler, *Organometallics*, 1994, **13**, 21–23.
- N. J. Hartmann, G. Wu and T. W. Hayton, *Chem. Sci.*, 2018, **9**, 6580–6588.
- J. P. Krogman, B. M. Foxman and C. M. Thomas, *J. Am. Chem. Soc.*, 2011, **133**, 14582–14585.
- L. Escomel, I. Del Rosal, L. Maron, E. Jeanneau, L. Veyre, C. Thieuleux and C. Camp, *J. Am. Chem. Soc.*, 2021, **143**, 4844–4856.
- O. Cooper, C. Camp, J. Pécaut, C. E. Kefalidis, L. Maron, S. Gambarelli and M. Mazzanti, *J. Am. Chem. Soc.*, 2014, **136**, 6716–6723.
- M. V. Butovskii, C. Döring, V. Bezugly, F. R. Wagner, Y. Grin and R. Kempe, *Nat. Chem.*, 2010, **2**, 741–744.
- C. J. Isaac, F. M. Miloserdov, A. F. Pécharman, J. P. Lowe, C. L. McMullin and M. K. Whittlesey, *Organometallics*, 2022, **41**, 2716–2730.
- L. Escomel, E. Jeanneau, C. Thieuleux and C. Camp, *Inorganics*, 2024, **12**, 72.
- S. Lassalle, J. Petit, R. L. Falconer, V. Hérault, E. Jeanneau, C. Thieuleux and C. Camp, *Organometallics*, 2022, **41**, 1675–1687.
- C. Z. Ye, I. Del Rosal, M. A. Boreen, E. T. Ouellette, D. R. Russo, L. Maron, J. Arnold and C. Camp, *Chem. Sci.*, 2022, **14**, 861–868.
- L. Escomel, N. Soulé, E. Robin, I. Del Rosal, L. Maron, E. Jeanneau, C. Thieuleux and C. Camp, *Inorg. Chem.*, 2022, **61**, 5715–5730.
- S. Lassalle, R. Jabbour, P. Schiltz, P. Berruyer, T. K. Todorova, L. Veyre, D. Gajan, A. Lesage, C. Thieuleux and C. Camp, *J. Am. Chem. Soc.*, 2019, **141**, 19321–19335.
- C. L. Gross and G. S. Girolami, *Organometallics*, 2007, **26**, 160–166.
- C. L. Gross, S. R. Wilson and G. S. Girolami, *J. Am. Chem. Soc.*, 1994, **116**, 10294–10295.
- S. Lassalle, R. Jabbour, I. Del Rosal, L. Maron, E. Fonda, L. Veyre, D. Gajan, A. Lesage, C. Thieuleux and C. Camp, *J. Catal.*, 2020, **392**, 287–301.
- W. A. Herrmann, H. G. Theiler, P. Kiprof, J. Tremmel and R. Blom, *J. Organomet. Chem.*, 1990, **395**, 69–84.
- C. L. Gross and G. S. Girolami, *Organometallics*, 2007, **26**, 160–166.
- J. W. Bruno, J. C. Huffman, M. A. Green and K. G. Caulton, *J. Am. Chem. Soc.*, 1984, **106**, 8310–8312.
- T. Shima and H. Suzuki, *Organometallics*, 2005, **24**, 3939–3945.
- L. J. Guggenberger and R. R. Schrock, *J. Am. Chem. Soc.*, 1975, **97**, 2935.
- R. Srivastava, E. A. Quadrelli and C. Camp, *Dalt. Trans.*, 2020, **49**, 3120–3128.
- R. Srivastava, R. Moneuse, J. Petit, P.-A. A. Pavard, V. Dardun, M. Rivat, P. Schiltz, M. Solari, E. Jeanneau, L. Veyre, C. Thieuleux, E. A. Quadrelli and C. Camp, *Chem. – Eur. J.*, 2018, **24**, 4361–4370.
- L. Pauling, *J. Am. Chem. Soc.*, 1947, **69**, 542–553.
- F. A. Cotton, *Acc. Chem. Res.*, 1978, **11**, 225–232.
- D. M. Heinekey, D. A. Fine, T. G. P. Harper and S. T. Michel, *Can. J. Chem.*, 1995, **73**, 1116–1125.
- J. K. Hoyano and W. A. G. Graham, *J. Am. Chem. Soc.*, 1982, **104**, 3722–3723.
- J. I. Fostvedt, M. A. Boreen, R. G. Bergman and J. Arnold, *Inorg. Chem.*, 2021, **60**, 9912–9931.
- S. M. Mullins, R. G. Bergman and J. Arnold, *Organometallics*, 1999, **18**, 4465–4467.
- P. Horrillo-Martinez, B. O. Patrick, L. L. Schafer and M. D. Fryzuk, *Dalt. Trans.*, 2012, **41**, 1609–1616.
- M. K. Whittlesey, R. N. Perutz and M. H. Moore, *Organometallics*, 1996, **15**, 5166–5169.
- O. R. Allen, S. J. Dalgarno, L. D. Field, P. Jensen, A. J. Turnbull and A. C. Willis, *Organometallics*, 2008, **27**, 2092–2098.
- J. Sánchez-Nieves and P. Royo, *J. Organomet. Chem.*, 2001, **621**, 299–303.
- M. A. Rankin and C. C. Cummins, *J. Am. Chem. Soc.*, 2010, **132**, 10021–10023.
- J. M. Mörsdorf and J. Ballmann, *Inorg. Chem.*, 2021, **60**, 18291–18295.

