



A focus on coordination chemistry at chlorine

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The first crystallographic characterization of chloronium cations stabilized by pyridine ligands (P. Pröhm, W. Berg, S. M. Rupf, C. Müller and S. Riedel, *Chem. Sci.*, 2023, <https://doi.org/10.1039/D2SC06757A>) is discussed in the context of coordination chemistry at chlorine.

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Coordination chemistry has most frequently been associated with classic transition metal–ligand complexes, where ligands donate two electrons to metals to form sigma bonds, often designated as coordinate or dative bonds. There is also rich coordination chemistry in the main group, with a classic example being the many adducts to boron or the heavier, more metallic, p-block elements.

Generally, more electronegative atoms are more electron rich and less likely to act as Lewis acids. The halogens (F, Cl, Br, I) are an electronegative group of elements and their chemistry is dominated by gaining electrons and existing in the -1 oxidation state. As such there is relatively less coordination chemistry for the halogens as compared to the other groups, nonetheless they do act as Lewis acids in compounds where they are found in higher oxidation states. Charge transfer complexes of the type Nu–X–X are well-studied.¹ They are most stable and common for iodine and several compounds have been crystallographically characterized, including for

pyridine² and phosphine ligands,³ for example. They are less common for bromine but can be observed⁴ and in more rare cases isolated.^{5,6} For chlorine there are few examples where the charge transfer complex can be observed,⁷ although they have been well studied computationally.⁸

Unsurprisingly, iodine as the least electronegative member of the group (radioactive astatine aside) has the best known and most stable coordination compounds where the halogen is found in the higher $+1$ oxidation state. The easy synthesis of [bis(pyridine)iodine][nitrate] ([Pyr-I-Pyr][NO₃]), from I₂, pyridine and AgNO₃ has even been implemented as undergraduate experiments at some institutions. The [BF₄][−] salt, known as Barluenga's reagent, is commercially available and widely used.⁹ The bromine analogue has been known for a very long time and has been crystallographically characterized,^{10–12} but rarely used, with a SciFinder search returning only a few dozen papers over 60 years compared to a few hundred for iodine. The chlorine analogue was first detected in solution at -80 °C by Erdélyi and co-workers in 2014 and they found that, as predicted by theoretical studies, the analogue forms a symmetric [Pyr–Cl–Pyr]⁺ species.¹³ [Pyr–F]⁺ is a commercially available electrophilic fluorinating reagent and Erdélyi found that addition of a second pyridine to this compound resulted in an asymmetric environment at low temperature. Additionally our group found that

attempts to observe any bis-pyridine adduct at room temperature, or perform ligand exchange, resulted in decomposition to complex mixtures.¹⁴

Which brings us to the remarkable recent results from Riedel and co-workers who were able to crystallographically characterize both mono and bis-pyridine adducts of chloronium cations [Cl]⁺ as well as a pyridine–Cl–Cl charge transfer complex for the first time.¹⁵

For the syntheses, the problem of obtaining stoichiometric amounts of Cl₂ gas was solved by first condensing and weighing Cl₂ into pressure tubes. This was then condensed onto the appropriate amount of pyridine or lutidine in propionitrile at -196 °C. Solutions were warmed to -40 °C to allow for reactions and cooled back to -80 °C to obtain crystals for X-ray analysis. The care required to achieve this is remarkable!

For the reactions of pyridine and lutidine with Cl₂, differing results were observed. For pyridine a Pyr–Cl–Cl complex was crystallized, while for lutidine, a [Lut–Cl–Lut][Cl₃] salt was obtained. This is a good illustration of the divergent results that can be obtained with subtle changes in systems containing weak bonds. The reactions begin with pyridine interacting with the σ^* orbital of chlorine, forming the complex observed for pyridine. This also induces weakening and polarization of the Cl–Cl bond. For the slightly stronger Lewis base lutidine,

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