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# Mechanism of coupled phase/morphology transformation of 2D manganese oxides through Fe galvanic exchange reaction†

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Nanostructured manganese oxides have a rich variety of morphologies and crystal phases which can undergo transformations during synthesis and application. Although these structural features are crucial for their performance, the mechanisms behind such transitions are not well understood. Herein, we describe the mechanism of transformation from layered 2D  $\delta$ - $\text{MnO}_2$  nanosheets to the scarcely reported  $\gamma$ - $\text{MnO}_2$  nanocone morphology. Despite the common purpose of introducing Fe dopants to enhance the conductivity of layered manganese oxides, the Fe galvanic exchange reaction was found responsible for such coupled phase/morphology transition. Electrochemical characterization confirmed a distinct electrochemical behaviour of the nanocones, emphasizing the need to unravel the mechanism of 2D  $\text{MnO}_2$  transformation. Such mechanistic insights were gained by systematic and rigorous electron microscopy studies. The effect of the local chemical composition was determined by energy dispersive X-ray spectroscopy while electron energy loss spectroscopy unravelled the key influence of the oxidation state of Mn ions within nanosheets and nanocones. We propose and demonstrate a  $\text{Mn}^{2+}$ -mediated oxidative mechanism of coupled morphology/phase transformation subjected to the equilibrium of Fe and Mn ions during galvanic exchange reaction. These findings contribute to the understanding of the growth and morphology/phase transformations of manganese oxide nanostructures, providing insights for the rational design of nanomaterials.

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## Introduction

Nanostructured manganese oxides ( $\text{MnO}_2$ ) are promising candidates for energy storage (e.g. batteries and supercapacitors),<sup>1</sup> environmental (e.g. pollutant removal),<sup>2</sup> and catalysis applications (e.g. oxygen evolution reaction).<sup>3</sup> Their properties are strongly dependent on the structure, which shows high complexity due to the variety of nanomorphologies, crystal structures (distinguishing between tunnel- and layer-like structure types), chemical compositions and stoichiometries.<sup>3–5</sup>

The introduction of dopants is conducted to further improve the properties of  $\text{MnO}_2$ .<sup>5,6</sup> First described by Oh *et al.*,<sup>7</sup> the galvanic exchange reaction can be performed to introduce Fe dopant in nanostructured oxides, including 2D  $\text{MnO}_2$

nanosheets.<sup>8</sup> The presence of Fe enhances the electric conductivity of the  $\text{MnO}_2$  material, which is a requirement for electrochemical applications. According to previous reports, dopant percentages of Fe up to 5 at% optimized the specific capacitance of  $\text{MnO}_2$ .<sup>9–11</sup>

Morphology and/or phase transitions can occur not only during synthesis, but also during post-synthesis procedures such as dopant modification and application of the nanostructures.<sup>12,13</sup> It has been reported that the presence of certain cations can stabilize specific crystal structures, thus causing phase transformations that are often coupled with structural changes. This is the case for the layer-to-tunnel polymorphic transitions of  $\text{MnO}_2$ , which benefit from the presence of cations of similar size occupying the tunnel void.<sup>14–16</sup> Specific examples of layer-to-tunnel transformations are those triggered by the sorption of  $\text{Mn}^{2+}$ . Two mechanisms were hypothesized to explain them: on the one hand, auto-catalytic oxidation of the adsorbed  $\text{Mn}^{2+}$  on the  $\text{MnO}_2$  surface would account for transition to  $\text{Mn}^{3+}$  enriched oxides.<sup>17</sup> On the other hand, comproportionation between  $\text{Mn}^{2+}$  ions and  $\text{Mn}^{4+}$  of  $\text{MnO}_2$  would also explain  $\text{Mn}^{3+}$  presence.<sup>18,19</sup>

Although understanding these transformations is of utmost importance for the design of high-performing materials, the

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mechanisms behind them are not yet completely unravelled.<sup>20</sup> For instance, Tu *et al.* observed a coupled phase/morphology transition upon  $\text{Mn}^{2+}$  addition from layered birnessite to several different  $\text{MnO}_2$  structures.<sup>17</sup> However, their synthesis methodology required several days and the exact mechanism behind the layer-to-tunnel transformation could not be proven due to lack of local structural analyses. Among the nanostructures yielding from the layer-to-tunnel transformation,  $\text{MnO}_2$  nanocones were obtained. The nanocones consist of a hexagonal pyramid morphology with faceted faces and a ramsdellite  $\gamma\text{-MnO}_2$ <sup>21</sup> or the closely related  $\epsilon\text{-MnO}_2$ <sup>22</sup> crystal structure. This nanostructure has since rarely been reported either arising from self-assembly of  $\alpha\text{-MnOOH}$  nanorods<sup>21,23</sup> or other direct wet chemical synthesis.<sup>22</sup> Furthermore, no electrochemical structure–property relationship study was pursued for the nanocone structure.

In this work, we achieve rapid phase transformation of layered  $\text{MnO}_2$  to  $\gamma\text{-MnO}_2$  nanocones by *in situ* generation of Mn defects and  $\text{Mn}^{2+}$  ions within 2D manganese oxide nanosheets using Fe galvanic exchange reaction. To investigate the phase transformation mechanism, extensive local structural and chemical characterization of the nanostructures was conducted by electron microscopies (scanning electron microscopy (SEM), (scanning) transmission electron microscopy ((S)TEM) and spectroscopic techniques (energy dispersive X-ray (EDS), electron energy loss spectroscopy (EELS) and Raman spectroscopy). The results allowed us to describe a  $\text{Mn}^{2+}$ -mediated oxidative mechanism for the nanosheet to nanocone morphology *via* galvanic exchange reaction. Our findings shed light into the mechanism behind coupled phase/morphology transitions of  $\text{MnO}_2$ , which is crucial for building structure–property relationships of functional nanomaterials.

## Experimental section

### Materials

The materials used during synthesis were tetramethylammonium hydroxide (TMAOH, 25 wt% in water, Sigma-Aldrich), hydrogen peroxide ( $\text{H}_2\text{O}_2$ , 30 wt%, Merck), manganese chloride ( $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ , >98%, Sigma-Aldrich), hydrochloric acid (HCl, 37 wt%, VWR), iron sulphate ( $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ , >98%, Sigma-Aldrich), tetramethylammonium chloride (TMACl, >98%, Sigma-Aldrich) and ethanol (EtOH, >99.8%, Carl Roth).

To prepare the working electrodes 1-methyl-2-pyrrolidone (NMP, >99%, Alfa Aesar) and Nafion® 117 solution ( $\approx 5\%$  in mixture of lower aliphatic alcohols, Sigma-Aldrich) were used. The electrolyte for the electrochemical measurements was prepared from potassium sulphate ( $\text{K}_2\text{SO}_4$ , AnalaR NORMA-PUR, VWR).

### Synthesis

**$\text{MnO}_2$  nanosheets synthesis.**<sup>8,24</sup> To prepare  $\sim 60$  ml of product, a mixture of 43 ml of a solution of 0.6 M TMAOH and 3 wt%  $\text{H}_2\text{O}_2$  was added slowly to 16 ml of a 0.3 M  $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$  solution. The dark brown mixture was let to stir from 15 h (HMO, 1M-HMO) to 24 h (2M-HMO, 3M-HMO) and was

centrifuged for 8 min at 3800 rpm either immediately (1M-HMO, 2M-HMO) or after 24 h (FMO, 3M-HMO). The supernatant suspension was aggregated with HCl 36 wt% (*ca.* 0.8 ml). This solution was centrifugated for 4 min at 5000 rpm and the aggregates were redispersed in deionized water and in EtOH. Finally, the aggregates in the solution were dried for 6.5 h in an oven preheated at 110 °C and the powder was grinded in an agate mortar.

**Fe galvanic exchange.**<sup>8</sup> Typically, 0.5 ml of a  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$  precursor solution of concentration in the range of 2–25  $\text{mg ml}^{-1}$  were injected to 5 ml of  $\text{MnO}_2$  nanosheets suspension (1  $\text{mg ml}^{-1}$ ). To test the effect of pH and cations,  $1.7 \times 10^{-5}$  mol of either TMAOH (6  $\mu\text{L}$ ) or TMACl (1.8 mg) were added to the  $\text{MnO}_2$  solution prior to Fe addition. In all cases, immediately after Fe addition the solution was heated to 90 °C (heating rate 3 °C  $\text{min}^{-1}$ ) while stirring for 2 h.

**Aging of the materials.** After Fe galvanic exchange reaction, the suspension was let to age for 24 h. The materials were subsequently purified with centrifugation and redispersion cycles in deionized water and EtOH.

### Electrochemical characterization

Cyclic voltammograms (CVs) were recorded with a three electrodes setup connected to a BioLogic VSP-300 potentiostat. A homemade Ag/AgCl, 3 M KCl electrode ( $E = 0.210$  V *vs.* SHE) equipped with 0.5 M  $\text{K}_2\text{SO}_4$  double junction was used as a reference electrode and a graphite rod ( $d = 8$  mm) was used as a counter electrode. The CVs were acquired using an analog scan rate of 25  $\text{mV s}^{-1}$  between 0.0 – 0.8 V *vs.* Ag/AgCl, 3 M KCl with a 0.5 M  $\text{K}_2\text{SO}_4$  solution as an electrolyte. The working electrodes were prepared on a glass slide substrate (0.2 mm thickness, 18 mm diameter) coated with a 200 nm layer of gold and with an 8 nm Ti adhesion layer. 10  $\mu\text{L}$  of a dispersion of 10  $\text{mg ml}^{-1}$  of the active material and a 1  $\text{mg ml}^{-1}$  of Nafion® in NMP were drop cast in the middle of the electrode slide in four 1.5  $\mu\text{L}$  and two 2.0  $\mu\text{L}$  steps. The functionalized electrodes were loaded in a custom-made teflon cell leaving exposed to the electrolyte solution the central area of the working electrode ( $d = 10$  mm).<sup>25</sup> A schematic of the cell is shown in Fig. S1.† The potentials were converted to the SHE scale for graph preparation.

### Structural characterization

**Electron microscopy.** SEM micrographs were acquired in a ZEISS Gemini with an acceleration voltage of 1.5 kV using an in-lens secondary electron detector. High resolution TEM (HRTEM) and scanning transmission electron microscopy high angle annular dark field (STEM-HAADF) were conducted in two FEI Titan Themis 60–300 with a  $C_s$  corrector for the objective and condenser lenses, respectively. EDS was performed with the Bruker Super X-EDX detector of the microscopes while EELS was acquired at the probe corrected STEM with a Gatan Quantum ERS energy filter. To check any possible reduction by the electron beam on the thin nanostructures, experiments were performed on nanosheets and nanocones of various thickness. The results can be found on Fig. S2.† Since no severe reduction was



As demonstrated by Lim *et al.*,<sup>8</sup> after Fe introduction and reaction at 90 °C the Fe cations are not only present in the interlayer space replacing the hydrated protons but are also partially incorporated to the MnO<sub>2</sub> lattice in Mn substitutional sites by galvanic exchange. For the MnO<sub>2</sub>/Mn<sup>2+</sup> and Fe<sup>3+</sup>/Fe<sup>2+</sup> couples, the galvanic exchange reaction is described *via*:

$$\text{MnO}_2 + 2\text{Fe}^{2+} + 4\text{H}^+ \rightarrow \text{Mn}^{2+} + 2\text{Fe}^{3+} + 2\text{H}_2\text{O}$$

Besides, our results show that the Fe galvanic exchange reaction also resulted in a morphology transformation of the initial HMO nanosheet assemblies to nanocones (Fig. 1c and d) in FMO1 and FMO2 as indicated with arrows in Fig. S3a-d.† Interestingly, this morphology was only observed in the materials with relatively lower amount of Fe<sup>2+</sup> ions used, thus indicating that the transformation is sensitive to the ratio of Fe<sup>2+</sup> ions used during galvanic exchange reaction.

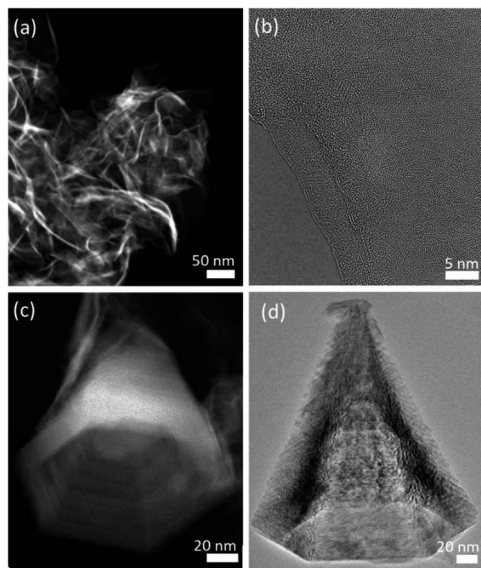
The nanocones clearly show the characteristic hexagonal pyramid shape with faceted faces. The average nanocone lengths of FMO1 and FMO2 are  $230 \pm 89$  nm and  $279 \pm 100$  nm with corresponding basal widths of  $87 \pm 32$  nm and  $184 \pm 73$  nm, respectively (Fig. S4<sup>†</sup>). Furthermore, unlike these hollow shaped nanocones, nanocones with a closed tip could be distinguished as minority (Fig. S5<sup>†</sup>). This can be a result of the pyramidal growth of nanocones one inside the other and subsequent break of the structure into individual cones.<sup>22</sup>

Since the pH is known to affect the morphology/phase transformation of manganese oxides<sup>17</sup> the pH of the MnO<sub>2</sub> and Fe solutions was measured prior to galvanic exchange reaction. In all cases, the pH was close to neutral (pH  $\approx$  7) (Fig. S6a<sup>†</sup>). Strong changes in pH could force different morphology/phase transitions as a result from the galvanic exchange reaction, as seen in Fig. S6b<sup>†</sup> for the case of conducting FMO3 in basic pH (pH  $\approx$  10.5, adjusted with TMAOH). Fig. S6c<sup>†</sup> demonstrates that the presence of TMA cations is not responsible for the different transition, as FMO3 experiments with equivalent presence of TMA<sup>+</sup> in neutral pH do not modify the nanosheet morphology.

### Transformation of MnO<sub>2</sub> through Fe galvanic exchange reaction

2D  $\text{MnO}_2$  nanosheets (HMO) were synthesized through a wet chemical reaction and subsequently assembled with protons by adding HCl, as explained in the Experimental section. The electrostatic interaction of the hydrated protons with the negatively charged nanosheets produces 3D assemblies (Fig. 1a) of individual nanosheets (Fig. 1b). Fe dopant was subsequently introduced to the assembled  $\text{MnO}_2$  with  $\text{Fe}^{2+}$  precursor to HMO ratios of 1 : 5, 3 : 5, 5 : 5 and 12.5 : 5 in weight and let 24 h to age. Electron micrographs of these materials, respectively named

CVs of 2D MnO<sub>2</sub> (HMO) and a Fe<sub>x</sub>Mn<sub>1-x</sub>O<sub>2</sub> nanocone-rich material (1M-FMO2, see Section 5) were measured in the potential range of 0.21–1.01 V vs. the standard hydrogen electrode (SHE). SEM micrographs of the two electrodes can be found in Fig. S7.† To investigate the effect of the morphology transformation on the capacitive properties, the specific differential capacitance was calculated from the CV data obtained in three distinct measurements, averaged and plotted against the potential (Fig. 2). Despite the presence of ohmic-drop-related distortions of the CVs (particularly for 1M-FMO2) obstructing a quantitative comparison between both data sets, higher specific capacitance of HMO is observed in most part of the analyzed potential range.



**Fig. 1** (a) Scanning transmission electron microscopy high angle annular dark field (STEM-HAADF) micrograph of HMO 3D nanosheets assembly composed by individual nanosheets shown by HRTEM in (b). (c) STEM-HAADF and (d) TEM images of FMO2 hollow nanocones.





Fig. 2 Averaged specific differential capacitance of 2D  $\text{MnO}_x$  nano-sheets (HMO) and nanocone-rich  $\text{Fe}_x\text{Mn}_{1-x}\text{O}_2$  (1M-FMO2) sample. The standard deviation was used to obtain the error bar reported in the plot.

We hypothesize that this difference may arise from a different accessibility of the material for charge compensating ions. The hexagonal pyramid morphology of the nanocones present in 1M-FMO2 likely results in a decreased accessible surface area when compared to the 2D nanosheet morphology of HMO. The Fe content could also play a role in the differences and mechanism of the capacitive (dis-)charging (Fig. S8†). Furthermore, the lower capacitances could also indicate a transition to a different crystal structure with less favorable capacitance properties than  $\delta\text{-MnO}_2$ .<sup>27,28</sup> The distinctive capacitance behavior of the nanocones compared to that of nanosheets highlights the importance of understanding the paths of morphology transitions for the rational design of functional materials. To gain mechanistic insights, the transformation was further studied by thorough characterization of the nanostructures.

### Composition and oxidation state analyses

EDS was conducted in STEM mode on all the FMO materials to determine the Fe contents of the different nanostructures, *i.e.* assembled nanosheets or nanocones. The results are summarized on Fig. 3a. The average at% Fe of the assembled nanosheets were 20, 20, 56 and 75% for FMO1, FMO2, FMO3 and FMO4, respectively. Although the Fe content on FMO3 and FMO4 was relatively homogeneous, FMO1 and FMO2 showed higher scattering of Fe concentration from various regions of analysis. Smaller 3D nanosheet assemblies (<500 nm) tend to possess higher Fe content than larger assemblies. Additionally, in case of larger assemblies, nanosheets in outer area (or surface) showed higher Fe content than those in inner part (or core) (Fig. S9†). From this we believe that upon low amount of  $\text{Fe}^{2+}$ , these ions could not diffuse into the inner part of the thicker nanosheet assemblies.

Analyses of the nanocone structure in both FMO1 and FMO2 revealed that those barely contain an average of 5 at% Fe, a much lower value compared to the 20 at% Fe of the assembled nanosheets of the same samples. Local differences of up to



Fig. 3 (a) Normalized at% Fe with respect to at% Mn for each material. (b–d) HAADF micrograph and corresponding EDS mapping after PCA analysis showing the distribution of Fe and Mn on nanosheets and nanocones in FMO2.

a  $\sim 50$  at% Fe between nanocones and nanosheets are shown quantitatively with a principle component analysis (PCA) of the EDS map data<sup>29</sup> on Fig. 3b–d. This suggests that nanocones may consist of a different crystal structure than the assembled nanosheets by terms of chemical component as well as morphology.

EELS measurements were conducted in STEM mode on all the  $\text{Fe}_x\text{Mn}_{1-x}\text{O}_2$  (FMO) materials to determine the oxidation state of the metal ions in the structure. The intensity ratios between  $L_2$  and  $L_3$  white lines ( $I(L_3)/I(L_2)$ ) as well as the Fe  $L_{2,3}$  and Mn  $L_{2,3}$  peak positions are indicators of the average oxidation state of the element, with shifts towards larger energy losses indicating higher oxidation states.<sup>30,31</sup> However, the Fe  $L_3$  peak position is maintained stable at  $\sim 710.2$  eV for all materials while the Fe  $I(L_3)/I(L_2)$  ratio only suffers from a slight change (Fig. S10,† Table 1). All the intensity ratio values are close to the reported  $I(L_3)/I(L_2) \approx 5.1$  for  $\alpha\text{-Fe}_2\text{O}_3$ ,<sup>32</sup> thus suggesting a dominance of  $\text{Fe}^{3+}$  oxidation state in all nanosheet specimens. This result is consistent with the galvanic exchange reaction transforming all  $\text{Fe}^{2+}$  precursor to  $\text{Fe}^{3+}$ .

Unlike Fe, the oxidation state of Mn varies. Fig. 4a contains the Mn  $L_3$  EELS spectra for assembled nanosheets of HMO,

Table 1 EELS values for Fe and Mn  $L_3$  white lines maxima and Fe and Mn  $I(L_3)/I(L_2)$  intensity ratio for nanosheets and nanocones of HMO, FMO1, FMO2, FMO3 and FMO4

Material	Fe $L_3$ (eV)	Fe $I(L_3)/I(L_2)$	Mn $L_3$ (eV)	Mn $I(L_3)/I(L_2)$
HMO nanosheet	—	—	641.5	2.9
FMO1 Nanosheet	710.2	5.4	641.1	3.2
FMO1 Nanocone	—	—	642.2	2.3
FMO2 Nanosheet	710.2	5.4	641.1	3.1
FMO2 Nanocone	—	—	642.4	2.4
FMO3 nanosheet	710.2	5.2	640.8	3.4
FMO4 nanosheet	710.2	5.5	640.7	3.5





Fig. 4 (a) Left-Y axis: EELS Mn  $L_3$  white line of the nanosheets of each material. Right-Y axis: intensity ratio of Mn  $L_3$  and  $L_2$  peaks (b) Left-Y axis: EELS Mn  $L_{2,3}$  white lines of nanosheets and nanocones of FMO1 and FMO2. Right-Y axis: intensity ratio of Mn  $L_3$  and  $L_2$  peaks for FMO1 and FMO2 nanocones and nanosheets. (c and d) HAADF and corresponding EELS mapping after PCA analysis showing the local energy loss corresponding to Mn  $L_3$  of the nanosheets and nanocones.

FMO1, FMO2, FMO3 and FMO4. The spectra show a shift of Mn  $L_3$  between the materials with Fe (Table 1) and HMO, indicating that upon galvanic exchange there are enhanced contributions of lower Mn valences ( $\text{Mn}^{3+}/\text{Mn}^{2+}$ ) due to the partial substitution of  $\text{Mn}^{4+}$  by  $\text{Fe}^{3+}$ . These qualitative results were also confirmed by  $I(L_3)/I(L_2)$  ratios, the reference values of which decrease from  $I(L_3)/I(L_2)_{\text{MnO}} \approx 4.0$ ,  $I(L_3)/I(L_2)_{\text{Mn}_2\text{O}_3} \approx 2.5$  to  $I(L_3)/I(L_2)_{\text{MnO}_2} \approx 1.9$ .<sup>30</sup> The calculated ratios shown in Fig. 4a and Table 1 increase from 2.9 to 3.5 with Fe, thus showing more reduction following the higher Fe content.

The value of  $I(L_3)/I(L_2)_{\text{HMO}} \approx 2.9$  indicates that the Mn ions in HMO before addition of Fe are already in a very low valence, with significant  $\text{Mn}^{2+}/\text{Mn}^{3+}$  contributions. This is compatible with a birnessite structure containing hydrated cations in the interlayer to compensate  $\text{Mn}^{3+}/\text{Mn}^{4+}$  or  $\text{Mn}^{2+}/\text{Mn}^{4+}$  of the structure, characteristic of the synthesis methodology.<sup>33,34</sup>

EELS spectra show a Mn  $L_3$  peak shift as high as  $\sim 1.1$  eV for FMO1 and  $\sim 1.3$  eV for FMO2 between nanocones and nanosheets (Fig. 4b, Table 1). This difference is also seen by local PCA of an EELS mapping of a nanocone and nanosheets<sup>29</sup> (Fig. 4c and d) with data displaying the energy loss of the Mn  $L_3$  peak maximum as indicator of oxidation state. Such results clearly reveal that Mn ions in the nanocones are more oxidized than those in assembled nanosheets. This is confirmed by the distinct changes of intensity ratios from  $I(L_3)/I(L_2)_{\text{FMO1}} \approx 3.2$  and  $I(L_3)/I(L_2)_{\text{FMO2}} \approx 3.1$  for nanosheets to  $I(L_3)/I(L_2)_{\text{FMO1}} \approx 2.3$  and  $I(L_3)/I(L_2)_{\text{FMO2}} \approx 2.4$  for nanocones. The magnitude of the difference between Mn oxidation state in assembled nanosheets and nanocones can neither be explained solely by the different Fe content nor by a difference in their thicknesses (Fig. S2†). Therefore, these results strongly indicate the oxidation of the Mn ions during morphology transformation from 2D  $\text{MnO}_2$  nanosheet to nanocones.

## Crystal structure

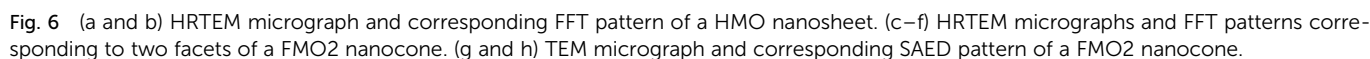
To elucidate the crystal structure of the materials, confocal Raman spectroscopy measurements were conducted (Fig. 5). All the spectra show four active modes:  $\nu_1$  ( $721.6\text{--}744.9\text{ cm}^{-1}$ ),  $\nu_2$  ( $636.2\text{--}643.0\text{ cm}^{-1}$ ),  $\nu_3$  ( $566.9\text{--}573.0\text{ cm}^{-1}$ ), and  $\nu_4$  ( $487.1\text{--}500.8\text{ cm}^{-1}$ ) (Table S1†). These modes can be assigned to the layered birnessite  $\delta\text{-MnO}_2$  structure for all FMO1, FMO2, FMO3 and FMO4 despite the different Fe contents.<sup>8,34,35</sup> X-ray diffraction (XRD) also confirmed such phase in HMO (Fig. S11†).

Moreover, Raman  $\nu_1$  and  $\nu_2$  modes appear broader in the materials with higher Fe contents (FMO3 and FMO4). This result is compatible with a higher extend of galvanic exchange reaction with higher Fe ratios, which induces the formation of more Mn defects and  $\text{Mn}^{2+}$  cations. Such defects reduce the order of the lattice, resulting in a polycrystalline structure with smaller grain sizes which broaden the Raman peaks.<sup>34</sup> The smaller grain size of the materials with higher Fe content was also confirmed by TEM imaging the materials (Fig. S12†).

Local crystal structure investigations of the nanocones and the nanosheets were conducted by HRTEM, Fast Fourier Transform (FFT) (Fig. 6a–f and S13†) and selected area aperture electron diffraction (SAED) (Fig. 6g and h). The data indicates that the assembled nanosheets of HMO as well as nanosheets containing Fe (FMO1, FMO2, FMO3 and FMO4) show a polycrystalline birnessite structure (Fig. S14†) while nanocones are highly crystalline *e.g.* much larger crystal size. The nanocone



Fig. 5 Raman spectra of HMO, FMO1, FMO2, FMO3 and FMO4.



Finally, since the nanocone transformation occurs through an oxidative step, the original oxidation state of the 2D  $\text{MnO}_2$  (HMO) can be of utmost importance. To assess its effect, three 2D  $\text{MnO}_2$  materials named 1M-HMO, 2M-HMO and 3M-HMO were synthesized by modifying the synthesis and aging time (see Experimental section) in order to change their Mn oxidation state. This was monitored by EELS (Fig. 8a), demonstrating





Fig. 7 Schematic mechanism governing the synthesis of  $\text{Fe}_x\text{Mn}_{1-x}\text{O}_2$  nanosheets and nanocones through galvanic exchange reaction.

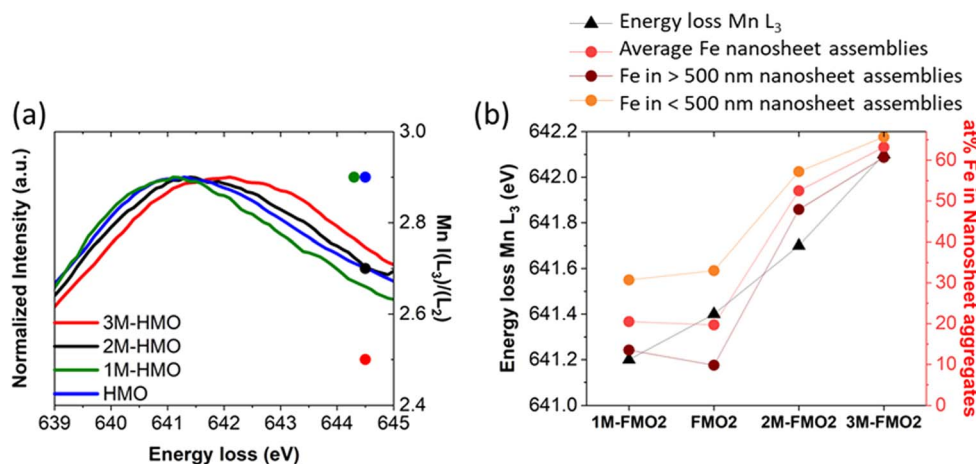


Fig. 8 (a) Left-Y Axis: EELS Mn  $L_3$  peaks of HMO, 1M-HMO, 2M-HMO and 3M-HMO. Right-Y axis: corresponding Mn  $L_3/L_2$  values. (b) Trends of oxidation state of xM-HMO (Left-Y axis) and Fe introduced by galvanic exchange reaction (right-Y axis) for xM-FMO2.

that the longer the synthesis and aging times, the more oxidized the Mn species of the 2D  $\text{MnO}_2$  nanosheets.

Galvanic exchange reaction was subsequently conducted with these xM-HMO materials emulating FMO2 conditions (xM-FMO2). The results showed that the materials produced with more oxidized  $\text{MnO}_2$  nanosheets (2M-FMO2 and 3M-FMO2) contained  $\text{Fe}_x\text{Mn}_{1-x}\text{O}_2$  nanosheets with larger contents of Fe than FMO2 but containing similar ratio of nanocones (Fig. 8b and S17†). On the contrary, reducing the oxidation state of Mn species (1M-HMO) by shorter synthesis and aging time resulted on materials with the same 20 at% Fe than HMO after galvanic exchange reaction. However, the difference in oxidation state of 1M-HMO contributed to the major enhancement of the

nanocone morphology with respect to nanosheet after reaction (Fig. S17†). These results can be related to the galvanic exchange reaction potentials with different  $\text{Mn}^{n+}$  species. The reaction between  $\text{Fe}^{3+}/\text{Fe}^{2+}$  and  $\text{Mn}_2\text{O}_3/\text{Mn}^{2+}$  couples is thermodynamically more favourable than that of  $\text{Fe}^{3+}/\text{Fe}^{2+}$  and  $\text{MnO}_2/\text{Mn}^{2+}$  ( $\Delta E^0_{\text{cell}} = 0.714 \text{ V}$  vs.  $\Delta E^0_{\text{cell}} = 0.459 \text{ V}$ , respectively). This difference potentially indicates larger local concentration of produced  $\text{Mn}^{2+}$  during galvanic exchange reaction with the more reduced 1M-HMO due to the higher thermodynamic driving force, explaining the larger nanocone to nanosheet ratio. Moreover, the nanocone size is also larger, showing an average length  $391 \pm 197 \text{ nm}$  and average width  $228 \pm 91 \text{ nm}$  distributed in two different family sizes. (Fig. S4†).





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