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## Ultrathin palladium nanosheets with selectively controlled surface facets†

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We report a facile bottom-up synthetic approach to preparing ultrathin two-dimensional (2D) palladium nanosheets (PdNSs) with selectively exposed surface facets. Our synthetic strategy is based on the utilization of the nanoconfined lamellar mesophases of amphiphilic functional surfactants to template the growth of PdNSs in aqueous solution. Preferential adsorption of functional groups (e.g., COOH, pyridyl and quaternary ammonium) and halide counter ions (e.g., Br<sup>-</sup> and Cl<sup>-</sup>) in the long-chain surfactants onto different Pd planes results in the epitaxial growth of {100}, {110} and {111}-exposed surface facets of ultrathin PdNSs. Our synthetic approach is a general, powerful and scalable method to precisely control the surface facets of ultrathin 2D PdNSs, thus providing an opportunity to evaluate facet-dependent catalytic performance of Pd nanocrystals. Ultrathin PdNSs have been examined as the electrocatalysts for hydrogen evolution reactions (HERs). We show that {100}-exposed PdNSs display superior catalytic activity and stability for HERs, compared to that of {110} and {111}-exposed ones as well as their bulk counterparts. Conceivably, our findings will offer a general guideline in rational design of surfactant templates for other 2D metal nanosheets.

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Ultrathin two-dimensional (2D) noble metal nanomaterials with atomic thickness have received considerable interest due to their unique electronic, catalytic, electrochemical, and optical properties, compared to their bulk nanoscale counterparts.<sup>1–5</sup> Those 2D nanosheets with a large portion of exposed surface atoms show superior catalytic activity for a number of chemical transformations.<sup>6,7</sup> The surfactant-directed synthetic approach has been utilized to engineer the nanostructures of ultrathin 2D noble metal nanosheets in the presence of specific adsorbates.<sup>8–11</sup> One example is the synthesis of Pd nanosheets with {111}-exposed crystal facet (denoted as PdNSs{111} hereafter) using cetyltrimethylammonium bromide and poly(vinylpyrrolidone) as templates and capping agents, and CO as reductant and facet-selective adsorbate.<sup>12–15</sup> Preferentially strong adsorption of CO onto Pd{111} plane inhibited the growth of Pd along {111} plane, and facilitated the epitaxial growth on uncovered {110} and {100} planes. This resulted in the formation of ultrathin PdNSs{111} with the hexagonal shape.<sup>16,17</sup> Other adsorbates, e.g., halide ions and amines, have

also been utilized to assist the control of exposed facets of Pd nanoparticles.<sup>18–20</sup> However, it is largely unsuccessful when extending the methodology to expose other low-index surface facets of ultrathin 2D PdNSs (e.g., {110} or {100}), although there are some examples on controlling the exposed facets in Pd nanocubes, polyhedrons and nanorods.<sup>21–24</sup> This may be originated from the complexity in tuning the interactions between the surfactants/adsorbates and specific Pd planes while simultaneously maintaining confined 2D nanostructures during the synthesis.<sup>19,25</sup> Therefore, a simple and general synthetic methodology to precisely control the specific exposed surface facets of ultrathin 2D PdNSs is exceedingly desired.<sup>18,26</sup>

We herein demonstrate a facile yet powerful bottom-up synthesis of ultrathin PdNSs with efficiently controlled surface facets through a template-assisted solution-phase growth. In comparison to the conventional surfactant templates, our design features the precious control of binding between metal precursors and surfactants by tuning the functionality on the head groups of the long-chain amphiphilic surfactants (see Fig. S1 and Table S1† for chemical structures of the surfactants). When varying the hydrophilic head groups, preferential adsorption on different crystal planes can direct the epitaxial growth along plane directions to control the exposed surface facets of PdNSs. Three functional groups, including carboxyl (COOH), pyridyl (Py) and quaternary ammonium (QA), are designed on surfactants having different lengths of alkyl tails to tune the surface crystal facets of 2D PdNSs. The functional surfactants used as the templates include three structural

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spacing of 1.96 Å (Fig. 2d), which can be assigned to the {200} plane of face-centered cubic (fcc) Pd (see wide-angle XRD in Fig. 2h). Fourier transform (FT) pattern further confirms the growth in [100] plane direction. We examined the lattice fringes of >20 different nanosheets (Fig. S9†) and all nanosheets show identical crystal structures, suggesting the same exposed surface facet of PdNSs. HRTEM images of different domains in an individual PdNS exhibits uniform spacings and orientations of the lattice fringes (Fig. S10†), indicating the single-crystalline feature of PdNSs. Selected-area electron diffraction (SAED) pattern in Fig. 2e displays a series of spots of single-crystalline structure, corresponding to the [100] zone diffraction of the fcc Pd. The thickness of PdNSs is measured to be *ca.* 2.5 nm with ~12 atomic layers, when taking TEM images perpendicular to the nanosheet (Fig. 2f). The fringes with a lattice spacing of 1.99 Å can be indexed to the Pd {002}, further indicating that the flat nanosheet planes have exposed {100} facet. The high-resolution X-ray photoelectron spectroscopy (XPS) of Pd 3d shows two asymmetric peaks at 335.4 and 340.6 eV (Fig. 2h), slightly higher than the values of PdNSs capped by quaternary ammonium groups or pyridyl groups (Fig. S11†) and the reported value of commercial Pd nanoparticles,<sup>34</sup> possibly because of strong surface dipole induced by carboxyl groups. Those findings suggest the first successful preparation of ultrathin PdNSs with {100}-exposed surface facet using the functional surfactant of C<sub>22</sub>N-COOH (Br<sup>-</sup>) simultaneously as the templates and capping agents.

To synthesize PdNSs with {110} and {111}-exposed surface facets, a set of functional surfactants were further examined by varying binding interactions of hydrophilic head groups with Pd crystal planes. Surfactants having functional Py groups that favor to interact the Pd{110} crystal planes,<sup>35–37</sup> were designed (C<sub>22</sub>-Py (Br<sup>-</sup>)). The as-designed functional surfactant of C<sub>22</sub>-Py (Br<sup>-</sup>) thus directed the growth of ultrathin PdNSs{110}. On the other hand, the surfactants with only QA groups and Cl<sup>-</sup> as counter ion (C<sub>22</sub>-QA (Cl<sup>-</sup>)) were designed to decrease the chemisorption interactions with the Pd{100} or {110} planes, and to expose the Pd{111} planes with low surface energy.<sup>38</sup> In addition, CO with strong preferential chemisorption onto the Pd{111} plane was also investigated as a control. The formation of ultrathin PdNSs with {111}-exposed facet regardless of the functional groups and halide ions in the surfactants (see Fig. S12 and S13†).

Low-magnification TEM images show that PdNSs{110} and PdNSs{111} are ultrathin with similar sheet nanostructures as PdNSs{100} (Fig. 3a and d). PdNSs{110} are irregular polygons with an average size of ~150 nm, while PdNSs{111} are nearly hexagonal of ~100 nm. HRTEM was used to characterize the crystallinity and exposed surface facets of PdNSs. As shown in Fig. 3b, for PdNSs{110}, there are two lattice fringes with the inter-planar spacings of 2.28 and 1.94 Å, ascribed to (111) and (200) planes of fcc Pd. This indicates the [110] growth direction of PdNSs. Meanwhile, the FT pattern (Fig. 3c) also confirms the [110] axes perpendicular to the nanosheet planes. More HRTEM observations of the PdNSs further demonstrated that all nanosheets were enclosed by the {110} plane (see Fig. S14†). For PdNSs{111}, the HRTEM image in Fig. 3e shows an interplanar



Fig. 3 Structural characterizations of (a–c) ultrathin PdNSs{110} and (d–f) PdNSs{111}. (a) TEM, (b) HRTEM images, and (c) corresponded FT pattern of PdNSs{110} synthesized using C<sub>22</sub>-Py (Br<sup>-</sup>). (d) TEM, (e) HRTEM images, and (f) corresponded FT pattern of PdNSs{111} synthesized using C<sub>22</sub>-QA (Cl<sup>-</sup>). The inset in (c) and (f) are schemes indicating the exposed surface facets.

lattice fringe of 2.42 Å, well-matched to the 1/3 (422) reflection of fcc Pd (Fig. 3e). The result indicates that the PdNSs have exposed {111} facets, similarly in the FT pattern (Fig. 3f).

We extended our investigations to understand whether the alkyl lengths and reaction temperatures affect the nanostructures of as-resultant PdNSs. As shown in Fig. 4a, the longer alkyl length favors to stabilize lamellar structures assembled from functional surfactants, promoting the formation of ultrathin 2D nanosheets (see Fig. S15–S17† for TEM images). Similarly, the more ordered nanosheet structures were obtained at the lower reaction temperature (Fig. S18†). This may be because the shorter alkyl chains and the higher reaction temperature will destabilize the initial lamellar mesophases thermodynamically, thus disrupting the epitaxial growth of ultrathin PdNSs along plane direction.<sup>39,40</sup>

In addition, the effect of functional head groups and halide counter ions of the surfactants on surface facets of the PdNSs were systematically investigated. To this end, surfactants with long alkyl chain of C<sub>22</sub> were used to form stable lamellar

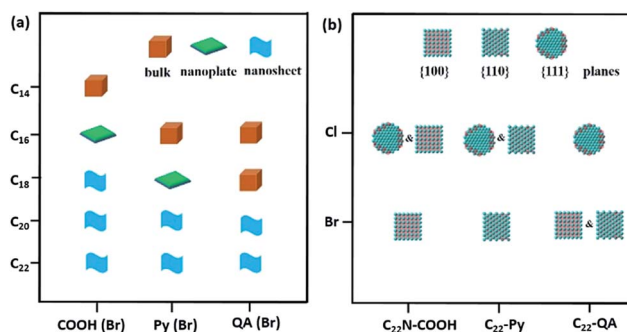


Fig. 4 Growth diagrams of Pd nanocrystals synthesized using different surfactants. (a) Structural "phase-like" diagram of Pd nanocrystals obtained from different alkyl lengths and functional groups. (b) Exposed surface facet "phase-like" diagram of PdNSs collected from different functional groups and halide counter ions.



mesophases; while the head groups and halide counter ions were varied (Fig. 4b). As reported previously, halide ion of  $\text{Br}^-$  preferentially chemisorbed onto  $\text{Pd}\{100\}$  and  $\text{Pd}\{110\}$  planes, while no favorable binding for  $\text{Cl}^-$ .<sup>16,32,41</sup> We found that  $\text{C}_{22}\text{-QA}$  ( $\text{Br}^-$ ) led to the formation of PdNSs with the mixed surface facets of  $\{100\}$  and  $\{110\}$  planes with a ratio of 39 : 61, confirming the preferential interaction of  $\text{Br}^-$  ions with  $\text{Pd}\{100\}$  or  $\{110\}$  planes (Fig. S19<sup>†</sup>). Those findings suggest that the weak adsorption of both quaternary ammonium groups and  $\text{Cl}^-$  ions on Pd crystal planes is the key to the formation of PdNSs with the lower crystal facets of  $\text{Pd}\{111\}$ . On the contrary, when  $\text{Cl}^-$  was used instead of  $\text{Br}^-$  in the surfactants of  $\text{C}_{22}\text{-N-COOH}$  ( $\text{Cl}^-$ ) and  $\text{C}_{22}\text{-Py}$  ( $\text{Cl}^-$ ), PdNSs $\{111\}$  mixed with PdNSs $\{100\}$  ( $\{111\} : \{100\} = 46 : 54$ ) and PdNSs $\{110\}$  ( $\{111\} : \{110\} = 32 : 68$ ) were simultaneously observed (Fig. S20<sup>†</sup>). Given slight difference of  $\Delta E_b$  in COOH and Py-containing surfactants (Fig. 1c), the results further suggested that the halide ion of  $\text{Br}^-$  has a synergistic effect on controlling exposed facets of the resultant PdNSs.

Electrocatalytic HERs were further used as a model reaction to evaluate the surface facet-dependent catalytic performances using PdNSs as the electrocatalysts. The commercial Pt (cPt) and Pd black (PdB) (see TEM images in Fig. S21<sup>†</sup>) were also tested as controls. Fig. 5a shows HER linear sweep voltammetry (LSV) curves under acidic condition (0.5 M  $\text{H}_2\text{SO}_4$ ) at a scan rate of  $5 \text{ mV s}^{-1}$ . PdNSs with different exposed surface facets exhibited completely different HER activities, following an order of PdNSs $\{100\} > \text{PdNSs}\{110\} > \text{PdNSs}\{111\}$ . The lowest overpotential of 67 mV was seen for PdNSs $\{100\}$  at a current density of  $10 \text{ mA cm}^{-2}$ , which is 91 and 160 mV lower than that of PdNSs $\{110\}$  and PdNSs $\{111\}$ , respectively (Fig. 5b). Besides,

PdNSs $\{100\}$  exhibited a comparable HER activity to the state-of-the-art catalyst of cPt (the overpotential of PdNSs $\{100\}$  is only 49 mV higher than that of cPt at  $10 \text{ mA cm}^{-2}$ ). The similar trends of HER activities were also observed at current densities at 50 and  $100 \text{ mA cm}^{-2}$  (Fig. 5b). The better electrocatalytic performance of PdNSs $\{100\}$  may be originated from the optimal balance between adsorption and desorption of H on the  $\text{Pd}\{100\}$  planes during HER.<sup>42,43</sup> By contrast, commercial PdB exhibited the lowest HER activity due to their large sizes.

The electrocatalytic stability of ultrathin 2D PdNSs $\{100\}$  was further examined. As shown in Fig. 5c, negligible shifts of LSV curves were observed for PdNSs $\{100\}$  after continuous potential sweeps for 15 000 cycles. The result indicated the superior stability of ultrathin PdNSs. No significant change on the nanosheets shape of the PdNSs $\{100\}$  was also observed after test for 15 000 cycles (see the TEM image as the inset in Fig. 5c), further confirming the superior stability. The current-time ( $i-t$ ) chronoamperometric response was also recorded by a continuous potential of 42 mV at the current density of  $50 \text{ mA cm}^{-2}$  (Fig. 5d). A respectable current retention of 96.7% was maintained for 21 h, further indicating the superior electrocatalytic durability of PdNSs $\{100\}$  in an acidic solution. In contrast, only 65.4% of current retention was maintained ( $\sim 20 \text{ h}$ ) for commercialized PdB nanoparticles. The improved electrocatalytic stability of ultrathin PdNSs can be ascribed to the anisotropic ultrathin 2D nanostructures which retarded the dissolution and ripening processes of PdNSs and strengthened the physical interaction between catalysts and the carbon support/electrode.<sup>1,44</sup>

In summary, we reported a novel but facile surfactant-assisted synthetic approach for epitaxial growth of ultrathin 2D nanosheets with well-controlled exposed surface facets, for the first time. The key to the synthesis of ultrathin PdNSs with specific exposed facets is the utilization of surfactants having functional hydrophilic head groups as the templates and capping agents. The change in hydrophobic carbon chain lengths, functional head groups and halide counter ions of the surfactants resulted in specifically exposed surface facets of ultrafine metal nanosheets. As a result, all three low-index exposed crystal facets of  $\{100\}$ ,  $\{110\}$  and  $\{111\}$  in ultrathin PdNSs were synthesized using the surfactant of  $\text{C}_{22}\text{-N-COOH}$  ( $\text{Br}^-$ ),  $\text{C}_{22}\text{-Py}$  ( $\text{Br}^-$ ) and  $\text{C}_{22}\text{-QA}$  ( $\text{Cl}^-$ ), respectively. Lastly, the surface facet-dependent electrocatalytic activity of ultrathin PdNSs was also examined by HERs. We found that the PdNS $\{100\}$  exhibited much better electrocatalytic activity and stability, compared to the PdNS $\{110\}$  and PdNS $\{111\}$ . This work represents the first successful example for facet-controlled synthesis of ultrathin PdNSs. We believe that the concept of 2D-templated strategy deriving from the self-assembly of novel designed functional surfactants could provide new insight in the prospective construction of other kinds of ultrathin noble metals or related inorganic nanomaterials.

## Conflicts of interest

There are no conflicts to declare.



Fig. 5 Electrocatalytic HER performances of different PdNSs in 0.5 M  $\text{H}_2\text{SO}_4$ . (a) LSV curves of PdNSs $\{100\}$ , PdNSs $\{110\}$ , PdNSs $\{111\}$ , PdB and cPt at a scan rate of  $5 \text{ mV s}^{-1}$ . (b) Electrocatalytic HER activities of PdNSs $\{100\}$ , PdNSs $\{110\}$ , PdNSs $\{111\}$ , PdB and cPt at current density of 10, 50 and  $100 \text{ mA cm}^{-2}$ , respectively. (c) LSV curves of PdNSs $\{100\}$  before and after repeating CV scans of 15 000 cycles at  $50 \text{ mV s}^{-1}$ . Inserted in (c) is the TEM image of PdNSs $\{100\}$  after test for 15 000 cycles. (d) Time-dependent  $i-t$  curves of PdNSs $\{100\}$  and commercial PdB under a constant overpotential of 42 mV at  $50 \text{ mA cm}^{-2}$ .



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## Notes and references

- Z. Fan, X. Huang, C. Tan and H. Zhang, *Chem. Sci.*, 2015, **6**, 95–111.
- A. Gupta, T. Sakthivel and S. Seal, *Prog. Mater. Sci.*, 2015, **73**, 44–126.
- S. Hu and X. Wang, *Chem. Soc. Rev.*, 2013, **42**, 5577–5594.
- P. K. Jain, X. Huang, I. H. El-Sayed and M. A. El-Sayed, *Acc. Chem. Res.*, 2008, **41**, 1578–1586.
- T. Ling, J. J. Wang, H. Zhang, S. T. Song, Y. Z. Zhou, J. Zhao and X. W. Du, *Adv. Mater.*, 2015, **27**, 5396–5402.
- Y. Sun, S. Gao, F. Lei, C. Xiao and Y. Xie, *Acc. Chem. Res.*, 2015, **48**, 3–12.
- Y. Jia, Y. Jiang, J. Zhang, L. Zhang, Q. Chen, Z. Xie and L. Zheng, *J. Am. Chem. Soc.*, 2014, **136**, 3748–3751.
- H. Duan, N. Yan, R. Yu, C. R. Chang, G. Zhou, H. S. Hu, H. Rong, Z. Niu, J. Mao, H. Asakura, T. Tanaka, P. J. Dyson, J. Li and Y. Li, *Nat. Commun.*, 2014, **5**, 3093.
- J. Niu, D. Wang, H. Qin, X. Xiong, P. Tan, Y. Li, R. Liu, X. Lu, J. Wu, T. Zhang, W. Ni and J. Jin, *Nat. Commun.*, 2014, **5**, 3313.
- A. X. Yin, W. C. Liu, J. Ke, W. Zhu, J. Gu, Y. W. Zhang and C. H. Yan, *J. Am. Chem. Soc.*, 2012, **134**, 20479–20489.
- M. S. Bakshi, *Cryst. Growth Des.*, 2016, **16**, 1104–1133.
- X. Huang, S. Tang, X. Mu, Y. Dai, G. Chen, Z. Zhou, F. Ruan, Z. Yang and N. Zheng, *Nat. Nanotechnol.*, 2011, **6**, 28–32.
- Y. Li, W. Wang, K. Xia, W. Zhang, Y. Jiang, Y. Zeng, H. Zhang, C. Jin, Z. Zhang and D. Yang, *Small*, 2015, **11**, 4745–4752.
- Y. T. Pan, X. Yin, K. S. Kwok and H. Yang, *Nano Lett.*, 2014, **14**, 5953–5959.
- P. F. Siril, L. Ramos, P. Beaunier, P. Archirel, A. Etcheberry and H. Remita, *Chem. Mater.*, 2009, **21**, 5170–5175.
- M. Chen, B. Wu, J. Yang and N. Zheng, *Adv. Mater.*, 2012, **24**, 862–879.
- B. Hammer, Y. Morikawa and J. K. Nørskov, *Phys. Rev. Lett.*, 1996, **76**, 2141–2144.
- J. Lai, W. Niu, R. Luque and G. Xu, *Nano Today*, 2015, **10**, 240–267.
- A. R. Tao, S. Habas and P. Yang, *Small*, 2008, **4**, 310–325.
- X. Huang, Z. Zhao, J. Fan, Y. Tan and N. Zheng, *J. Am. Chem. Soc.*, 2011, **133**, 4718–4721.
- H. Zhang, M. Jin, Y. Xiong, B. Lim and Y. Xia, *Acc. Chem. Res.*, 2012, **46**, 1783–1794.
- M. Jin, H. Zhang, Z. Xie and Y. Xia, *Angew. Chem., Int. Ed.*, 2011, **50**, 7850–7854.
- Y.-H. Chen, H.-H. Hung and M. H. Huang, *J. Am. Chem. Soc.*, 2009, **131**, 9114–9121.
- C.-K. Tsung, J. N. Kuhn, W. Huang, C. Aliaga, L.-I. Hung, G. A. Somorjai and P. Yang, *J. Am. Chem. Soc.*, 2009, **131**, 5816–5822.
- A. Chen and C. Ostrom, *Chem. Rev.*, 2015, **115**, 11999–12044.
- W. W. Xiong and Q. Zhang, *Angew. Chem., Int. Ed.*, 2015, **54**, 11616–11623.
- H. J. Monkhorst and J. D. Pack, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1976, **13**, 5188–5192.
- D. Vanderbilt, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1990, **41**, 7892–7895.
- J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- Y. Jiang, Y. Yan, W. Chen, Y. Khan, J. Wu, H. Zhang and D. Yang, *Chem. Commun.*, 2016, **52**, 14204–14207.
- Y. Sun and Y. Xia, *Science*, 2002, **298**, 2176–2179.
- Y. J. Xiong, H. G. Cai, B. J. Wiley, J. G. Wang, M. J. Kim and Y. N. Xia, *J. Am. Chem. Soc.*, 2007, **129**, 3665–3675.
- B. Veisz and Z. Király, *Langmuir*, 2003, **19**, 4817–4824.
- K. Noack, H. Zbinden and R. Schlögl, *Catal. Lett.*, 1990, **4**, 145–155.
- F. P. Netzer, *Vacuum*, 1990, **41**, 49–53.
- D. Xu, Y. Ma, Z. Jing, L. Han, B. Singh, J. Feng, X. Shen, F. Cao, P. Oleynikov, H. Sun, O. Terasaki and S. Che, *Nat. Commun.*, 2014, **5**, 4262.
- D. Xu, Y. Liu, S. Zhao, Y. Lu, M. Han and J. Bao, *Chem. Commun.*, 2017, **53**, 1642–1645.
- L. Vitos, A. V. Ruban, H. L. Skriver and J. Kollár, *Surf. Sci.*, 1998, **411**, 186–202.
- K. Na, M. Choi, W. Park, Y. Sakamoto, O. Terasaki and R. Ryoo, *J. Am. Chem. Soc.*, 2010, **132**, 4169–4177.
- Q. Huo, D. I. Margolese, U. Ciesla, P. Feng, T. E. Gier, P. Sieger, R. Leon, P. M. Petroff, F. Schuth and G. D. Stucky, *Nature*, 1994, **368**, 317–321.
- X. Xia, S. I. Choi, J. A. Herron, N. Lu, J. Scaranto, H. C. Peng, J. Wang, M. Mavrikakis, M. J. Kim and Y. Xia, *J. Am. Chem. Soc.*, 2013, **135**, 15706–15709.
- L. L. Jewell and B. H. Davis, *Appl. Catal., A*, 2006, **310**, 1–15.
- W. Dong, V. Ledentu, P. Sautet, A. Eichler and J. Hafner, *Surf. Sci.*, 1998, **411**, 123–136.
- S. Hu and X. Wang, *Chem. Soc. Rev.*, 2013, **42**, 5577–5594.

